## **Nanoscale**



## CORRECTION

View Article Online
View Journal | View Issue



Cite this: Nanoscale, 2016, 8, 19174

## Correction: Reduced graphene oxide nanosheets decorated with Au-Pd bimetallic alloy nanoparticles towards efficient photocatalytic degradation of phenolic compounds in water

Gitashree Darabdhara, a,b Purna K. Boruah, a,b Priyakshree Borthakur, a,b Najrul Hussain, a,b Manash R. Das, \*a,b Tansir Ahamad, Saad M. Alshehri, \*c Victor Malgras, Kevin C.-W. Wu\*e and Yusuke Yamauchid

DOI: 10.1039/c6nr90225a

www.rsc.org/nanoscale

Correction for 'Reduced graphene oxide nanosheets decorated with Au–Pd bimetallic alloy nanoparticles towards efficient photocatalytic degradation of phenolic compounds in water' by Gitashree Darabdhara, et al., Nanoscale, 2016, **8**, 8276–8287.

The authors would like to draw the attention of the readers to some corrected aspects of the published article:

On page 8281, section 3.2: the pH value in the sentence beginning "About 94.4% phenol degradation..." should be corrected to 7

On page 8282, section 3.2.1: the five different catalyst loadings should be corrected to 0.1, 0.3, 0.5, 0.8 and 1 g  $L^{-1}$ .

On page 8282, section 3.2.2: the four different concentrations of the phenolic components should be corrected to 0.2, 0.3, 0.5 and 0.8 mM.

On page 8282, section 3.2.2: the degradation efficiencies of the phenolic compounds at various initial concentrations for a fixed amount of catalyst and pH are actually displayed in Fig. S4, not Fig. S3.

On page 8283, section 3.2.4: the two sentences "The electron donating  $-NO_2$  group....reduces it" should be corrected to "The electron donating group on the aromatic ring possessing an (+I) inductive effect (*i.e.*, an electron donating effect) increases the negative charge on the aromatic ring of phenol. On the other hand, the presence of a -Cl and  $-NO_2$  group having an (-I) inductive effect (*i.e.*, electron accepting effect) reduces it".

On page 8284, section 3.2.6: all instances of  $H_2O^*$  should be corrected to  $HO_2^*$  in the discussion of the mechanism.

On page 8284, section 3.2.6: eqn (2) and (3) should be corrected to:

$$O_2^{\bullet -} + H_2O \rightarrow HO_2^{\bullet} + OH^- \tag{2}$$

$$e^- + HO_2^{\bullet} + H^+ \rightarrow H_2O_2 \tag{3}$$

<sup>&</sup>lt;sup>a</sup>Materials Science Division, CSIR-North East Institute of Science and Technology, Jorhat 785006, Assam, India. E-mail: mnshrdas@yahoo.com

<sup>&</sup>lt;sup>b</sup>Academy of Scientific and Innovative Research, India

<sup>&</sup>lt;sup>c</sup>Department of Chemistry, College of Science, King Saud University, Riyadh 11451, Saudi Arabia. E-mail: alshehri@ksu.edu.sa

<sup>&</sup>lt;sup>d</sup>National Institute for Materials Science (NIMS), 1-1 Namiki, Tsukuba, Ibaraki 305-0044, Japan

<sup>&</sup>lt;sup>e</sup>Department of Chemical Engineering, National Taiwan University, No. 1, Sec. 4, Roosevelt Road, Taipei 10617, Taiwan. E-mail: kevinwu@ntu.edu.tw

Nanoscale

On page 8284: Fig. 12(a), with corrected legend, should be:

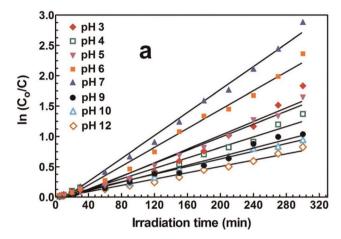


Fig. 12 Kinetic plot of  $ln(C_0/C)$  as a function of time for (a) phenol, (b) 2-CP and (c) 2-NP.

Lastly, the authors would like to correct the following references:

- 15. A. Fujishima and K. Honda, Nature, 1972, 238, 37.
- 37. P. Zhu, A. S. Nair, P. Shengjie, Y. Shengyuan and S. Ramakrishna, ACS Appl. Mater. Interfaces, 2012, 4, 581.
- 63. O. Akhavan, E. Ghaderi and A. Esfandiar, J. Phys. Chem. B, 2011, 115, 6279.
- 67. M. Aslam, I. M. I. Ismail, S. Chandrasekaran and A. Hameed, J. Hazard. Mater., 2014, 276, 120.

The Royal Society of Chemistry apologises for these errors and any consequent inconvenience to authors and readers.