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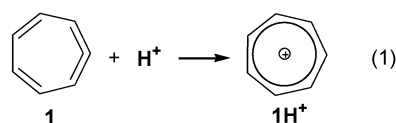
A DFT study of endocyclic allenes: unprecedentedly superbasic hydrocarbons†

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Density functional theory (DFT) calculations using the B3LYP, M05-2X and M06-2X functionals were employed in the study of basicity of allenes incorporated in seven-membered rings. These highly unstable species upon protonation give thermodynamically very stable tropylium cations, which are the origin of their extremely high basicity. An endocyclic allene → tropylium cation sequence was used for the design of novel hydrocarbon superbases, with calculated gas phase proton affinities exceeding well above 300 kcal mol^{−1}.

Neutral organic superbases are of theoretical and synthetic interest, in particular their importance as aprotic and metal-free catalysts in organic reactions increases rapidly.¹ Therefore, the design of novel superbasic molecules is an active research field.^{2,3} As a result, several functional groups with superbasic properties were identified: amidines, guanidines,⁴ biguanides,⁵ phosphazenes,⁶ and carbenes.^{7,8} Further increase in the basicity was achieved by imposing unfavourable interactions in the neutral form of the molecules, which are weakened upon protonation (proton sponges).⁹ Other effects that are employed in the design of superbases are the stabilization of protonated forms by resonance or intramolecular hydrogen bonding.¹⁰ For additional increase in basicity, various substituents were attached to delocalise the positive charge.¹¹

We report here an *in silico* design of novel extremely basic hydrocarbon molecular systems. It is based on the cyclic allene → tropylium cation concept, where the incorporation of an allene moiety in a cycloheptadiene ring distorts the allene moiety from the ideal linear geometry to a less stable species. Upon protonation at the central carbon atom of the allene moiety the effective positive charge stabilization by the resonance of the tropylium cation and its aromatic character giving a highly stable system favors this process (Scheme 1). Cyclohepta-1,2,4,6-tetraene **1** is a



Scheme 1 Formation of tropylium cation from allene **1**.

known compound¹² and it was shown that free **1** instantaneously yields 7-ethoxycyclohepta-1,3,5-triene if generated in ethanolic solution through the protonation of **1**.¹³ Apparently, the basicity of **1** is high enough to cause its smooth conversion by HOEt and DOEt into (deutero)tropylium ethoxide. In addition to the estimation of basicities of **1** and its derivatives, investigation of their electronic structure is of interest, since the nature of binding in small ring endocyclic allenes (allene *vs.* carbene) was the subject of intensive discussion.¹⁴ These cyclic bent allenes are of importance in various organometallic reactions.¹⁵

Structures of **1** and azuleno¹⁶ allenes **2–19** (Fig. 1) and their protonated forms were optimized by the B3LYP/6-31G(d), M052X/6-31G(d) and M062X/6-31G(d) methods and their gas phase basicities were calculated (for details see the ESI†). The results clearly favour cyclic allene-type structures over carbenes. These computationally studied exocyclic allenes possess extremely high basicity, as assessed by gas phase proton affinities (PAs) and basicities (GB) (Table 1).

The proton affinities of the bases defined as a negative value of the enthalpy change for the protonation reaction (Scheme 1) of studied molecules vary from 262.5 to 350.9 kcal mol^{−1} at the B3LYP/6-311+G(d,p)//B3LYP/6-31G(d) level (Table 1). These values are significantly larger than 245 kcal mol^{−1} (PA of dimethylaminonaphthalene) thus designating allenes **1–14** as organic superbases in the gas phase. The borderline of 300 kcal mol^{−1} signifies the threshold for hyperbasicity,³ indicating that allenes **15–19** belong to this category. For comparison, the calculated (B3LYP/6-311+G(d,p)) and experimental enthalpies of protonation for the parent allene are 185.9 and 185.3 kcal mol^{−1}, respectively,¹⁷ whereas the calculated PA of 1,4-dimethyl allene is 206.6 kcal mol^{−1}

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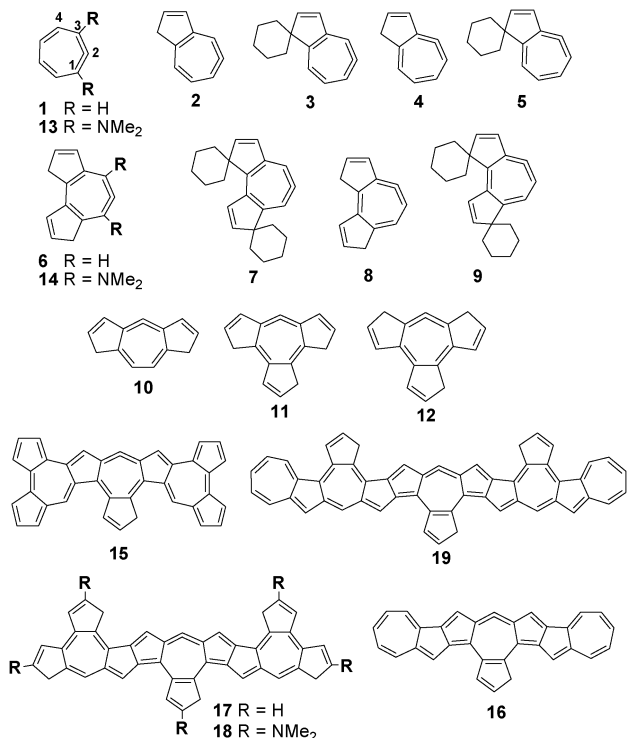


Fig. 1 Superbases 1–19 studied in this work.

Table 1 Gas phase proton affinities and basicities (kcal mol^{−1})

Base	PA _{B(2)} ^a	GB _B ^b	pK _a (MeCN) ^c	PA _{M05} ^c	GB _{M05}	PA _{M06} ^d	GB _{M06}
1	262.5	257.8	31.8	251.2	250.4	249.2	248.5
2	269.3	270.0	35.0	263.1	262.2	261.2	260.4
3	273.3	273.8	34.6	266.6	265.4	265.0	264.0
4	274.4	275.1	37.4	269.6	268.9	267.6	266.9
5	277.6	278.3	36.5	272.4	271.5	271.3	269.9
6	278.7	279.3	37.7	272.7	271.7	270.8	269.9
7	284.0	284.1	37.0	278.0	276.3	276.7	275.1
8	287.7	284.1	37.7	279.5	279.4	277.3	275.9
9	288.4	288.6	42.6	284.2	282.7	282.8	281.6
10	286.2	287.3	40.5	283.7	283.5	281.9	281.2
11	291.9	293.1	41.8	290.2	289.6	287.9	287.5
12	292.8	294.1	42.5	290.8	290.6	289.1	288.8
13	278.7	278.1	36.4	275.9	275.4	274.9	273.6
14	280.3	280.6	33.0	277.1	276.4	282.8	274.9
15	313.5	316.1	48.0	310.6	311.7	308.7	309.7
16	321.9	323.5	54.7	317.7	317.3	314.7	308.5
17	329.0	330.8	54.9	326.1	325.9	323.7	324.2
18	350.9	352.7	60.3	347.7	348.0	345.0	345.7
19	318.5	320.7	46.1	302.9	302.8	301.4	301.5

^a PA_{B3LYP(2)}: B3LYP/6-311+G(d,p)//B3LYP/6-31G(d) + *E*_{vib}(B3LYP/6-31G(d)).⁹^b B3LYP/6-31G(d). ^c PA_{M052X(2)}: M052X/6-311+G(d,p)//M052X/6-31G(d) + *E*_{vib}(M052X/6-31G(d)). ^d PA_{M062X(2)}: M062X/6-311+G(d,p)//M062X/6-31G(d) + *E*_{vib}(M062X/6-31G(d)).

at the same computational level. PAs obtained in our study are well above previously calculated values for allenes.^{8,17,18} Analysis of results obtained by DFT calculations reveals that the fusion of five and seven-member rings leads to increased basicity by better delocalization of positive charge upon protonation. The enhanced basicity of 1,1-spiroalkylated azulenoallene 3 is attributed in the literature¹⁹ to the stabilization of both the inductive and σ-π

conjugation effects of the exocyclic ring at the 1-position and the π-π conjugation of the double bond at the 2,3-positions in the respective 1*H*-azulenium cation. Presumably, the same effects stabilize the cations of 5, 7 and 9. Further increase in the basicity of endocyclic allenes could be achieved by the strategic attachment of substituents, which efficiently stabilizes positive charges such as the dimethylamino group (1/13, 6/14, and 17/18). In addition, the positioning of allene terminal carbon atom at the ring junction of five and seven-membered rings was found to further destabilize the endocyclic allene system (by 5 kcal mol^{−1}, 2/4 and 6/8 pairs).

A similar relative basicity order was obtained for gas phase basicities (GB) reaching extremely high value (352.7 kcal mol^{−1}) for 18 at the B3LYP/6-31G(d) level, while the lack of experimental values precludes the establishment of p*K*_a correlation for allenes. However, an estimate of p*K*_a values in acetonitrile using the correlation devised for guanidines²⁰ still follows the trend of gas phase basicities, indicating that allenes listed in Table 1 have p*K*_a values in acetonitrile (B3LYP) within the range of 32–60. The most basic is allene 18 with the p*K*_a value certainly well above 50 p*K*_a units.

Geometrical distortions imposed by the cyclic framework on the allene moiety and structural changes caused by protonation at the central carbon atom of the allene moiety are illustrated in Fig. 2 for azulenoallene 4.

The C₁=C₂=C₃ angles in endocyclic allenes 1–19 vary from 131.2° to 146.4°, and the out-of-plane angles (C₁=C₂=C₃-C₄) are within a range of 17.8–33.9°, indicating a large deviation from the ideal allene geometry (180° and 45°, respectively). As a consequence of this distortion, the orbital overlap is not ideal and the double bonds are weakened; this is evidenced by the average allene C₁=C₂/C₂=C₃ distances of 1.330–1.375 Å, which are significantly elongated in comparison to the parent allene (1.309 Å). These geometrical features are reflected on the C=C bond orders and hybridisation of the central allene carbon atom, which changes, for instance, from sp¹ in parent 1,4-dimethylallene to sp^{1.29} in 15 (Table S1, ESI[†]). Upon protonation at the central allene carbon and its rehybridisation, effective cyclic conjugation is increased, which is demonstrated by the decreased bond alternation in the tropylium ring, such as for 4/4*H*⁺, as well as by the planarization of the cationic structure.

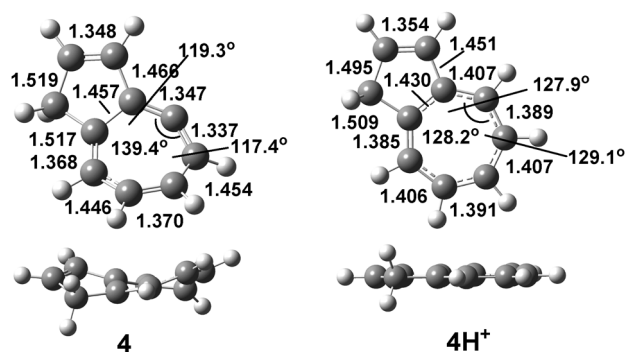


Fig. 2 Neutral and protonated structures of 4 (B3LYP/6-31G(d)).



In conclusion, endocyclic allenes are recognized as potent neutral organic superbases and their unprecedented basicity was estimated by DFT calculations. Results suggest that they could effect *in situ* generation of anions and deprotonation of common organic solvents such as acetonitrile or DMSO.

Computational methodology

Calculations were performed using the Gaussian03 suite of programs. Geometry optimizations were carried out at the DFT level employing B3LYP, M052X and M062X hybrid functionals in conjunction with the 6-31G(d) basis set. Harmonic vibration frequencies were calculated to verify that the stationary structures are minima. Single point calculations were conducted, at the B3LYP/6-311+G(d,p) level, using B3LYP/6-31G(d) optimized structures. Gas phase basicity (GB) was calculated as the Gibbs free energy change of the protonation reaction. Proton affinities in the gas phase and pK_a values in acetonitrile were calculated using equations given in the ESI.†

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Notes and references

- 1 *Superbases for Organic Synthesis: Guanidines, Amidines, Phosphazenes and Related Organocatalysts*, ed. T. Ishikawa, Wiley, Chichester, 2009.
- 2 D. Margetić, *Physico-chemical properties of organosuperbases*, in Ref. 1, 2009, pp. 9–48.
- 3 Z. B. Maksić, B. Kovačević and R. Vianello, *Chem. Rev.*, 2012, **112**, 5240–5270.
- 4 Z. B. Maksić and B. Kovačević, *J. Org. Chem.*, 2000, **65**, 3303–3309.
- 5 Z. Glasovac, P. Trošelj, I. Jušinski, D. Margetić and M. Eckert-Maksić, *Synlett*, 2013, 2540–2544.
- 6 B. Kovačević, D. Barić and Z. B. Maksić, *New J. Chem.*, 2004, **28**, 284–288.
- 7 R. Lo and B. Ganguly, *Chem. Commun.*, 2011, **47**, 7395–7397; M. Liu, I. Yang, B. Buckley and J. K. Lee, *Org. Lett.*, 2010, **21**, 4764–4767.
- 8 R. Tonner and G. Frenking, *Angew. Chem., Int. Ed.*, 2007, **46**, 8695–8698.
- 9 D. Margetić, T. Ishikawa and T. Kumamoto, *Eur. J. Org. Chem.*, 2010, 6563–6572; D. Margetić, P. Trošelj, T. Ishikawa and T. Kumamoto, *Bull. Chem. Soc. Jpn.*, 2010, **83**, 1055–1057; A. Singh and B. Ganguly, *Eur. J. Org. Chem.*, 2007, 420–422; A. Singh and B. Ganguly, *New J. Chem.*, 2009, **33**, 583–587; V. Raab, K. Harms, J. Sundermeyer, B. Kovačević and Z. B. Maksić, *J. Org. Chem.*, 2003, **68**, 8790–8797.
- 10 B. Kovačević, Z. Glasovac and Z. B. Maksić, *J. Phys. Org. Chem.*, 2002, **15**, 765–774.
- 11 B. Kovačević and Z. B. Maksić, in *NIC Symposium Proceedings, John von Neumann Institute for Computing, Jülich, NIC series 20*, ed. D. Wolf, G. Münster and M. Kremer, 2003, pp. 71–80.
- 12 M. Christl, in *Modern Allene Chemistry*, ed. N. Krause and A. S. K. Hashmi, Wiley, Weinheim, 2004, ch. 6, vol. 1, pp. 243–357.
- 13 W. Kirmse, K. Loosen and H.-D. Sluma, *J. Am. Chem. Soc.*, 1981, **103**, 5935–5937.
- 14 I. Fernández, C. A. Dyker, A. DeHope, B. Donnadieu, G. Frenking and G. Bertrand, *J. Am. Chem. Soc.*, 2009, **131**, 11875–11881; H. V. Huynh and G. Frison, *J. Org. Chem.*, 2013, **78**, 328–338; E. Soriano and I. Fernández, *Chem. Soc. Rev.*, 2014, **43**, 3041–3105, and references cited therein.
- 15 M. Melaimi, M. Soleilhavoup and G. Bertrand, *Angew. Chem., Int. Ed.*, 2010, **49**, 8810–8849.
- 16 T. Okazaki and K. K. Laali, *Org. Biomol. Chem.*, 2003, **1**, 3078–3093.
- 17 S. Gronert and J. R. Keefe, *J. Org. Chem.*, 2007, **72**, 6343–6352.
- 18 D. S. Patel and P. V. Bharatam, *J. Org. Chem.*, 2011, **76**, 2558–2567.
- 19 M. Oda, N. Nakajima, N. C. Thanh, K. Kitahara, R. Miyatake and S. Kuroda, *Eur. J. Org. Chem.*, 2008, 5301–5307.
- 20 B. Kovačević and Z. B. Maksić, *Org. Lett.*, 2001, **3**, 1523–1526.

