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New dinuclear hydrido-carbonyl rhenium complexes designed as photosensitizers in dye-sensitized solar cells

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The possible use of some dinuclear rhenium complexes as sensitizers for dye sensitized solar cells (DSSCs) has been investigated. They have general formula $[\text{Re}_2(\mu\text{-X})(\mu\text{-Y})(\text{CO})_6(\mu\text{-pyridazine-4-COOH})]$, with $\text{X} = \text{Y} = \text{Cl}$ (**1**), $\text{X} = \text{H}$, $\text{Y} = \text{benzoato}$ (**2**), and $\text{X} = \text{H}$, $\text{Y} = 4\text{-diphenylaminobenzoato}$ (**3**). An original synthetic strategy has been set for preparing the hydrido-carboxylato derivatives **2** and **3**. They have been indicated by DFT and TD-DFT computations as the most promising dyes, endowed with good light harvesting capability. The complexes have absorption maxima in the range of 405–443 nm, on TiO_2 films, arising from metal-to-ligand-charge transfer transitions. Cyclic voltammetry experiments have been performed on the derivatives containing the methyl ester of the pyridazine-4-COOH acid, showing electrochemical band gaps in the range of 2.25–1.63 eV. The best DSSC results have been obtained using complex **3**, with an overall solar-to-electric conversion efficiency of 1.0%. Noteworthy the presence of a hydrido ligand did not show any detrimental effect on the stability of the sensitizers under the operating conditions.

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Introduction

A new family of dinuclear rhenium complexes with general formula $[\text{Re}_2(\mu\text{-X})(\mu\text{-Y})(\text{CO})_6(\mu\text{-diazine})]$, containing two “ $\text{Re}(\text{CO})_3$ ” units connected by a bridging 1,2-diazine ligand and two anionic ancillary ligands, such as halides,¹ hydrides,² alcohoxides or sulphides,³ has been recently reported. Upon optical excitation, many of these complexes exhibit broad and featureless photoluminescence, centred between 520 and 660 nm, arising from the lowest-lying triplet-manifold excited state ($T_1 \rightarrow S_0$), with metal-to-ligand charge transfer (³MLCT) nature. Emission quantum yields up to 0.5, both in solution and in the solid state, have been measured for some members of this family,⁴ which have also found applications in OLED devices as a dopant in the emissive layer.⁵

Moreover very recently the first molecular dyad containing a dinuclear rhenium(i) moiety has been reported.⁶ In this species

a $[\text{Re}_2(\mu\text{-Cl})_2(\text{CO})_6(\mu\text{-pyridazine})]$ complex constitutes the chromophoric subunit, covalently linked to a fullerene unit by a carbocyclic molecular bridge. In this dyad the rhenium complex does not exhibit any emission, since its strongly emissive triplet charge-transfer state is efficiently quenched by the fullerene subunit, *via* oxidative electron transfer, with a time constant of about 110 ps.

These results prompted us to investigate the possible use of dinuclear rhenium complexes as dyes in dye sensitized solar cells (DSSCs).^{7–9} Actually sensitizers suitable for such applications must be able to transfer electrons from an excited state to the conductive band of an oxide (most often TiO_2) acting as the anode of the solar cell. The most commonly used dyes are ruthenium complexes containing poly-pyridine ligands functionalized with carboxylic substituents for anchoring to TiO_2 .¹⁰ These complexes exhibit broad absorptions, due to MLCT transitions, long excited state lifetime, and high chemical stability.¹¹ The highest efficiency reported for small area DSSCs based on the ruthenium dye and iodide/triiodide electrolyte is 12.1%,^{12,13} while an efficiency as high as 13% was achieved using a molecularly engineered porphyrin dye in combination with a cobalt-based electrolyte.¹⁴

Alternative dyes based on different metal complexes, such as iron,¹⁵ copper,¹⁶ and platinum,¹⁷ mainly containing bipyridine ligands, have been investigated so far.¹⁸ In this framework, the use of rhenium polypyridine complexes has not been largely investigated.

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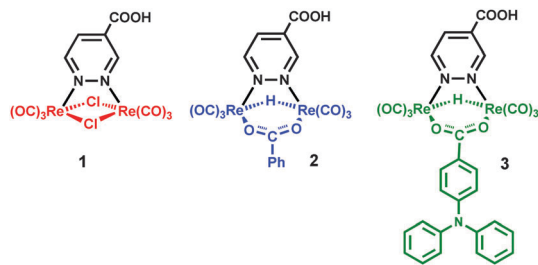


Chart 1 Molecular structure of the three dyes.

Indeed, beside some preliminary studies concerning the electron transfer from Re polypyridyl complexes to nanocrystalline TiO_2 ,^{19,20} only one study briefly exploring the photovoltaic performances of tricarbonyl rhenium complexes can be found.²¹

The dirhenium complexes here presented (Chart 1) have been designed by the use of combined density functional (DFT) and time-dependent density functional (TD-DFT) studies. All of them show ¹MLCT transitions, in which the excited electron is transferred to the diazine ligand, where the LUMO is mainly localized.^{1–4} The diazine has been functionalized by a carboxylic group, to allow the anchoring of the dye to the TiO_2 surface and the fast injection of the electron from the diazine to the TiO_2 conduction band. The ancillary ligands have been selected in order to maximize the light harvesting capability of the dye. The computational studies showed that complexes containing one hydrido and one benzoate bridging ligands were the most promising dyes. The synthesis of hydrido-carboxylato derivatives required the setting of a new general synthetic procedure, which has been exploited for the preparation of complexes 2 and 3 of Chart 1. Their electrochemical and spectroscopical properties have been measured and compared with those of the previously known complex 1, containing two bridging chloride anions. Finally, solar cells with each of the three complexes have been prepared and their performances are discussed herein, in relation to the molecular structure of the dyes and to the nature of the ancillary ligands.

Experimental section

Synthesis

All reactions were performed under a nitrogen atmosphere unless specified otherwise. All solvents were deoxygenated and dried by standard methods before use, while toluene has been distilled on Na(s)-benzophenone and dichloromethane on P_2O_5 , both under a N_2 atmosphere. The reagents were purchased from Aldrich, Fluka and Lancaster and used as received. Commercial deuterated solvents were used as received. Column chromatography was performed using Alfa Aesar silica gel 60 (0.032–0.063 mm). $[\text{Re}_4(\mu_3\text{-H})_4(\text{CO})_{12}](\text{C}_6\text{H}_6)_2$,²² $[\text{Re}_2(\mu\text{-H})_2(\text{CO})_6(\mu\text{-ppd})]$ (4, ppd = 2,5-diphenyl-1,3,4-oxadiazole)²³ and $[\text{Re}_2(\mu\text{-Cl})_2(\text{CO})_6(\mu\text{-pydz-4-COOH})]$ (1, pydz-4-COOH = pyridazine-4-carboxylic acid)²⁴ were synthesized according to the literature procedures. ¹H NMR spectra were recorded on Bruker DRX-300 and Bruker DRX-400 NMR spectrometers. Solution and thin-film UV/Vis absorption

spectra were obtained using an Agilent 8453 spectrophotometer, fused quartz cuvettes (10 mm optical path), and CHROMASOL V grade solvents. IR spectra were acquired on a Bruker Vector 22 FT spectrophotometer. MS spectrum of 1 was carried out on a FinniganTM LCQTM Advantage MAX ion trap instrument using methanol as the solvent (ESI-MS: m/z 736). FAB spectrum of 3 was carried out on a mass spectrometer VG Autospec M246 using nitrobenzyl alcohol as the matrix.

$[\text{Re}_2(\mu\text{-H})(\mu\text{-OOCPh})(\text{CO})_6(\mu\text{-ppd})]$ 5. A sample of benzoic acid (6.4 mg, 0.052 mmol) was added to a solution of $[\text{Re}_2(\mu\text{-H})_2(\text{CO})_6(\mu\text{-ppd})]$ (4, 40 mg, 0.052 mmol) previously dissolved in 15 mL of anhydrous toluene. The reaction mixture was stirred overnight, and the completeness of the reaction was monitored by IR spectroscopy. Then the solution was evaporated to dryness under vacuum, and the solid was dissolved in CH_2Cl_2 . The addition of *n*-hexane caused the precipitation of the product, which was purified through column chromatography (eluent ethylacetate/hexane 9 : 1), then dried under vacuum, yielding 30 mg (0.034 mmol) of white powder (isolated yield 65%). IR (CH_2Cl_2) $\nu(\text{CO})$: 2041 (m), 2022 (vs), 1935 (vs), 1914 (s) cm^{-1} . ¹H NMR (CD_2Cl_2 , 300K, 400 MHz) δ_{H} (ppm) 8.15 (d, 4H, H_{ortho}), 7.87 (t, 2H, H_{para}), 7.77 (t, 4H, H_{meta}), 8.03 (d, 2H, H_{ortho}), 7.51 (t, 1H, H_{para}), 7.41 (dt, 2H, H_{meta}), –7.08 (s, 1H, hydride). Elemental anal. calcd for $\text{C}_{27}\text{H}_{16}\text{N}_2\text{O}_9\text{Re}_2$: C 36.65, H 1.82, N 3.17. Found: C 36.70, H 1.96, N 3.14.

$[\text{Re}_2(\mu\text{-H})(\mu\text{-OOCPh})(\text{CO})_6(\mu\text{-pydz-4-COOH})]$ 2. A sample of $[\text{Re}_2(\mu\text{-H})(\mu\text{-OOCPh})(\text{CO})_6(\mu\text{-ppd})]$ (5, 30 mg, 0.034 mmol) dissolved in freshly distilled THF solution (5 mL) was treated with 4-pyridazine-carboxylic acid (6.3 mg, 0.050 mmol). The temperature was set at 333 K for 2 h. The color of the solution progressively becomes deep red and the progress of the reaction was monitored by IR spectroscopy. Then the solution was evaporated to dryness under vacuum, the crude product dissolved in CH_2Cl_2 and addition of *n*-hexane affords the precipitation of red powder. The solution was removed and the remaining powder was washed with *n*-hexane (5 times with 3 mL), then dried under vacuum. The solid was dissolved again in CH_2Cl_2 and complex 2 was collected as red powder after precipitation with *n*-hexane (13 mg, 0.165 mmol, isolated yield 56%). IR (CH_2Cl_2) $\nu(\text{CO})$: 2039 (m), 2018 (s), 1935 (s), 1918 (s) cm^{-1} . ¹H NMR (THF, 300K, 400 MHz) δ_{H} (ppm) 9.75 (dd, 1H, H2 4-COOH-pydz), 9.51 (d, 1H, H3 4-COOH-pydz), 8.34 (dd, 1H, H5 4-COOH-pydz), 7.88 (dd, 2H, H_{ortho}), 7.50 (t, 1H, H_{para}), 7.38 (dt, 2H, H_{meta}), –6.71 (s, 1H, hydride). Elemental anal. calcd for $\text{C}_{18}\text{H}_{10}\text{N}_2\text{O}_{10}\text{Re}_2$: C 27.48, H 1.28, N 3.56. Found: C 27.42, H 1.30, N 3.55.

4-(Diphenylamino)benzoic acid (TPA-COOH). In a dried flask, a sample of 4-(diphenylamino)benzaldehyde (500 mg, 1.8 mmol) was dissolved in 60 mL of acetone. Separately, 40 mL of aqueous solution of KMnO_4 (0.157 M, 3.5 equiv.) was prepared, filtered, and then added dropwise to the solution containing the aldehyde. The reaction mixture was left at room temperature for 72 h, then filtered on celite to remove MnO_2 . The filtered solution was collected in an open vessel and left at room temperature until excess MnO_4^- was decomposed and the solution becomes yellow and cloudy. Then it was filtered again on celite. The yellow solution was treated with 2.5 mL of $\text{HCl}/\text{H}_2\text{O}$ 1 : 1, a pale yellow precipitate appears, then stirred for 30 minutes in an ice bath, and



filtered again. The resulting solution was dried under vacuum affording 270 mg (0.93 mmol) of pure microcrystalline product (isolated yield 50%). ^1H NMR (CD_2Cl_2 , 300 K, 400 MHz) δ_{H} (ppm) 7.91 (d, 2H), 7.35 (t, 4H), 7.17 (m, 6H), 7.01 (d, 2H).

[Re₂(μ -H)(μ -4-OOC-TPA)(CO)₆(μ -ppd)] 6. A sample of [Re₂(μ -H)₂(CO)₆(μ -ppd)] (4, 42 mg, 0.055 mmol) was dissolved in freshly distilled toluene (20 mL) and treated with a slight excess of 4-(diphenylamino)benzoic acid (19 mg, 0.066 mmol). The reaction mixture was refluxed overnight, then the solution was evaporated to dryness under vacuum. The white residue was dissolved in CH_2Cl_2 and precipitated with *n*-hexane, affording a white powder which was further purified through column chromatography (eluent toluene/ethyl acetate 9:1), affording the desired product as white powder (32 mg, 0.034 mmol, isolated yield 55%). IR (CH_2Cl_2) $\nu(\text{CO})$: 2041 (m), 2022 (vs), 1936 (vs), 1915 (s) cm^{-1} . ^1H NMR (CD_2Cl_2 , 300 K, 400 MHz) δ_{H} (ppm) 8.16 (d, 4H), 7.77 (m, 4H), 7.89 (t, 2H), 7.84 (dd, 2H), 7.32 (t, 4H), 7.14 (m, 6H), 6.95 (dd, 2H), -7.09 (s, 1H, hydride). Elemental anal. calcd for $\text{C}_{39}\text{H}_{25}\text{N}_3\text{O}_9\text{Re}_2$: C 44.52, H 2.40, N 3.99. Found: C 44.93, H 2.45, N 3.96.

[Re₂(μ -H)(μ -4-OOC-TPA)(CO)₆(μ -pydz-4-COOH)] 3. A sample of [Re₂(μ -H)(μ -4-OOC-TPA)(CO)₆(μ -ppd)] (6, 12 mg, 0.011 mmol) dissolved in freshly distilled THF (5 mL) was treated at room temperature with 4-pyridazine-carboxylic acid (1.7 mg, 0.014 mmol). The solution was refluxed for 3 h. The solution was evaporated to dryness under vacuum and the crude product was dissolved in CH_2Cl_2 and precipitated with *n*-hexane. The supernatant solution was removed and the remaining powder was washed with *n*-hexane (five times with 3 mL), then dried under vacuum, affording the desired product 3 (9 mg, 0.009 mmol, isolated yield 83%). IR (CH_2Cl_2) $\nu(\text{CO})$: 2038 (m), 2018 (s), 1936 (s), 1917 (s) cm^{-1} . ^1H NMR (THF, 300 K, 400 MHz) δ_{H} (ppm) 9.70 (s, 1H), 9.45 (d, 1H), 8.33 (d, 1H), 7.74 (d, 2H), 7.31 (m, 4H), 7.12 (m, 6H), 6.91 (d, 2H), -6.70 (s, 1H, hydride). FAB-MS: m/z 955, elemental anal. calcd for $\text{C}_{30}\text{H}_{19}\text{N}_3\text{O}_{10}\text{Re}_2$: C 37.77, H 2.01, N 4.41. Found: C 37.56, H 2.01, N 4.37.

Methyl ester of pydz-4-COOMe and of complexes 1-COOMe, 2-COOMe and 3-COOMe. A sample of pyridazine-4-carboxylic acid (pydz-4-COOH, 100 mg, 0.805 mmol) was dissolved in MeOH (2 mL) and was treated with H_2SO_4 (40 μL , 95% w/w) at room temperature. The reaction mixture was heated at reflux temperature and was stirred overnight. Then the solution was cooled at room temperature and the reaction was quenched by the addition of saturated solution of Na_2CO_3 until pH 8. The product was extracted with Et_2O and the organic fractions were collected, washed with brine, dried with Na_2SO_4 and evaporated to dryness to leave a pale yellow solid (isolated yield 30%). ^1H NMR (DMSO, 300 K, 400 MHz) δ_{H} (ppm) 9.58 (s, 1H, H_{ortho}), 9.51 (d, 1H, H_{ortho}), 8.09 (dd, 1H, H_{meta}), 3.94 (s, 3H, CH_3). This pyridazine was used to synthesize the corresponding derivatives 1-COOMe, 2-COOMe and 3-COOMe, according to the previous procedures.

Electrochemical measurements

The cyclovoltammetric study of the complexes has been performed at scan rates typically ranging from 0.02 to 10 V s^{-1} in HPLC-grade acetonitrile (MeCN) solutions at 2.5×10^{-4} M concentration in each substrate, deaerated by N_2 bubbling,

with 0.1 M TBAPF₆ (TBA = NBu_4 , Aldrich) as the supporting electrolyte, at 298 K. The ohmic drop has been compensated by the positive feedback technique. The experiments were carried out using an AUTOLAB PGSTAT potentiostat (EcoChemie, The Netherlands) run by a PC with GPES software. The working electrode was a glassy carbon one (AMEL, diameter 1.5 mm) cleaned by diamond powder (Aldrich, diameter 1 μm) on a wet cloth (STRUERS DP-NAP); the counter electrode was a platinum wire; the reference electrode was an aqueous saturated calomel electrode, having in our working medium a difference of -0.385 V vs. the Fc^+/Fc couple (the intersolvent redox potential reference currently recommended by IUPAC).

Computational details

Ground state geometries were optimized by means of density functional calculations. The parameter-free hybrid functional PBE0^{25,26} was employed along with the standard valence double- ζ polarized basis set 6-31G(d,p) for C, H, Cl, N, O and S. For Re the Stuttgart-Dresden effective core potentials were employed along with the corresponding valence triple- ζ basis set. Preliminary calculations were done without imposing any symmetry. However, almost all the optimized structures are symmetric (apart from 2, 3 and TPA-COOH), as reported in Table 1. The nature of all the stationary points was checked by computing vibrational frequencies and all the geometries were found to be true minima. In order to simulate the absorption electronic spectrum down to 300 nm the lowest 20 singlet excitation energies were computed by means of time-dependent density functional calculations. For complexes 1–3 calculations were done also in the presence of solvent (toluene, used in the photophysical characterizations) described by the conductor-like polarizable continuum model (CPCM). All the calculations were done with Gaussian 09.²⁷

Fabrication and characterization of DSSCs

FTO glass substrates (Pilkington, 8 U sq^{-1}) were ultrasonically cleaned sequentially in acetone and in isopropyl alcohol for 10 min and then immersed in a 40 mM aqueous solution of TiCl_4 at 70 $^\circ\text{C}$ for 30 min. The as treated substrates were rinsed with deionized water and ethanol and dried at room temperature. An opaque TiO_2 single-layer was screen-printed on clean FTO glass substrates using a commercial TiO_2 paste (18 NR-AO paste, Dyesol). Over each $4 \times 2 \text{ cm}^2$ glass sample, three 8 μm -thick TiO_2 active areas of $0.5 \times 0.5 \text{ cm}^2$ were deposited and electrically isolated one from another by laser scribing the FTO. The electrodes were then sintered in an oven with a final step at 525 $^\circ\text{C}$ for 30 min. The sintered photoanodes were left in a 40 mM aqueous solution of TiCl_4 at 70 $^\circ\text{C}$ for 30 min. After rinsing with deionized water and ethanol, the samples were treated at 500 $^\circ\text{C}$ for 30 min. For each dye, two 0.3 mM ethanol solutions were prepared, respectively, with and without adding 3 mM of chenodeoxycholic acid (CDCA from Sigma Aldrich). After 2 h of immersion in the dye solution, photoelectrodes were assembled with platinized counter electrodes, prepared by pyrogenic decomposition at 420 $^\circ\text{C}$ for 15 min of a screen-printed platinum-based paste (Pt10 004F-05, Chimet), using a thermoplastic gasket (Surlyn 25, Solaronix). For comparison, a reference sample was sensitized



by a commercial N719 dye (Dynamo AB). Finally, either a commercial electrolyte (Dyesol's EL-HSE High Stability Electrolyte) or the EL14 electrolyte, containing smaller amounts of the redox couple I^-/I_3^- , was injected by vacuum back-filling. All devices were assembled in air and no secondary encapsulants were utilized. Photovoltaic performance of the cells was measured using a computer-controlled Keithley 2420 source meter, under a solar simulator (SolarTest 1200 KHS Class B) at 1000 W m^{-2} , AM 1.5. The incident power was measured using a Skye SKS 1110 sensor; before measurement, a black $0.7 \times 0.7 \text{ cm}^2$ adhesive mask was applied over each cell to avoid overestimations by light collection from areas surrounding the active area of the cell. Incident photon-to-current efficiency (IPCE) measurements were carried out using a set-up consisting of a 150 W xenon lamp (Newport Model 70612) coupled with a monochromator (Cornerstone 130) and a Keithley 2420 source meter.

Results and discussion

Computational study

DFT and TD-DFT computations have been previously used to investigate in detail the frontier orbitals and the electronic transitions of many $[\text{Re}_2(\mu\text{-X})(\mu\text{-Y})(\text{CO})_6(\mu\text{-diazine})]$ complexes.^{1–4} None of them contained a carboxylic group on the bridging diazine ligand, therefore the first aim of the present study was to understand the effects caused by the presence of a COOH substituent on the diazine. The β position was chosen, since, as previously reported,¹ the presence of substituents in the α position has detrimental effects on the stability of the complex, due to the steric hindrance with the equatorial carbonyl ligands.

To enlighten the effects of the COOH substituent on the molecular energy levels, we have investigated the $[\text{Re}_2(\mu\text{-Cl})_2(\text{CO})_6(\mu\text{-pydz-4-COOH})]$ complex (**1**) and compared its data with those of the progenitor $[\text{Re}_2(\mu\text{-Cl})_2(\text{CO})_6(\mu\text{-pydz})]$ complex.¹

In agreement with the results previously reported for some other members of this family of complexes,^{1,4b} DFT calculations on **1** show that the six highest occupied molecular orbitals

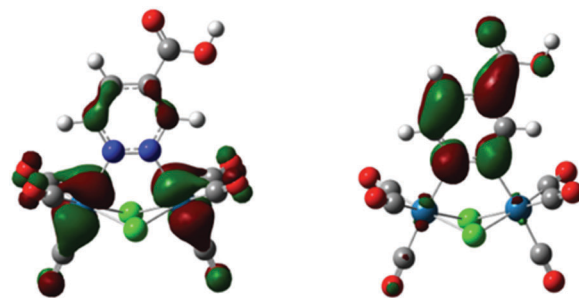


Fig. 1 Isodensity surface plots of some relevant molecular orbitals of $[\text{Re}_2(\mu\text{-Cl})_2(\text{CO})_6(\mu\text{-pydz-4-COOH})]$ (**1**): HOMO-3 (left) and LUMO (right).

correspond to the “ t_{2g} ” set of the two rhenium atoms in a pseudo-octahedral environment (three of them are combination of non-bonding d orbitals and the remaining three are $\text{Re}(\mu\text{-Cl}) \pi^*$ orbitals) while the two lowest unoccupied molecular orbitals are the lowest-lying π^* orbitals of the diazine. These orbitals lie at much lower energy than in the parent compound containing unsubstituted pyridazine (-0.47 eV , Table 1), in accordance with the electron-withdrawing nature of the carboxylic group. In agreement with the nature of the frontier orbitals, the HOMO level is poorly affected (-0.07 eV).

The orbitals involved in the strongest MLCT electronic transition computed for **1** (see Table 1) are represented in Fig. 1. This transition is red-shifted with respect to the corresponding transition computed for the pyridazine derivative, due to the stabilization of the LUMO. In addition, Fig. 1 clearly shows that the LUMO in **1** is delocalized over the carboxylic substituent, used to anchor the sensitizer to the nanocrystalline TiO_2 , possibly allowing a fast charge-injection process into the surface. We have then investigated the possibility to further shift to longer wavelengths the absorption band, in order to increase the light harvesting ability of the dye.

Taking into account the possibility to tune the energy-gap for these species not only modifying the substituents on the diazine (*i.e.* shifting the LUMOs) but also changing the nature of the ancillary ligands (*i.e.* shifting the HOMOs), we have

Table 1 Computed orbital energies (E_{HOMO} and E_{LUMO}), energy gap (ΔE), most intense MLCT excitation energies (λ) and oscillator strengths (f , in parentheses) for the compounds **1–3** here presented and for a series of related rhenium complexes (each molecule is labeled with its symmetry group)

Compound	E_{HOMO} [eV]	E_{LUMO} [eV]	ΔE [eV]	λ [nm, eV] (f)	Transition ^a
C_s - $[\text{Re}_2\text{Cl}_2(\text{CO})_6(\text{pydz-4-COOH})]$ (1)	−6.91	−3.99	2.92	428, 2.90 (0.178)	H−3 → L
C_{1v} - $[\text{Re}_2\text{H}(\text{OOCPh})(\text{CO})_6(\text{pydz-4-COOH})]$ (2)	−6.53	−3.78	2.75	526, 2.36 (0.118)	H−1 → L
C_{1v} - $[\text{Re}_2\text{H}(\text{OOC-TPA})(\text{CO})_6(\text{pydz-4-COOH})]$ (3)	−5.48	−3.71	1.77	530, 2.34 (0.079)	H−3 → L
				344, 3.60 (0.672) ^b	H → L+2
C_{2v} - $[\text{Re}_2\text{Cl}_2(\text{CO})_6(\text{pydz})]$	−6.84	−3.52	3.32	385, 3.22 (0.136)	H−3 → L
C_{2v} - $[\text{Re}_2(\text{OPh})_2(\text{CO})_6(\text{pydz})]$	−6.31	−3.36	2.95	387, 3.20 (0.130)	H−9 → L
C_2 - $[\text{Re}_2(\text{SPh})_2(\text{CO})_6(\text{pydz})]$	−6.27	−3.22	3.05	415, 2.98 (0.059)	H−5 → L
C_2 - $[\text{Re}_2(\text{OOCPh})_2(\text{CO})_6(\text{pydz})]$	−6.50	−3.19	3.32	446, 2.78 (0.091)	H−2 → L
				411, 3.01 (0.090)	H−4 → L
C_s - $[\text{Re}_2\text{HCl}(\text{CO})_6(\text{pydz})]$	−6.69	−3.45	3.24	406, 3.06 (0.113)	H−3 → L
C_s - $[\text{Re}_2\text{H}(\text{OPh})(\text{CO})_6(\text{pydz})]$	−6.39	−3.38	3.00	397, 3.12 (0.096)	H−6 → L
C_s - $[\text{Re}_2\text{H}(\text{SPh})(\text{CO})_6(\text{pydz})]$	−6.27	−3.29	2.97	386, 3.21 (0.128)	H−1 → L+1
C_s - $[\text{Re}_2\text{H}(\text{OOCPh})(\text{CO})_6(\text{pydz})]$	−6.44	−3.31	3.13	469, 2.64 (0.117)	H−1 → L
TPA-COOH	−5.48	−1.04	4.44	327, 3.79 (0.473) ^b	H → L

^a A description of the electronic transition in terms of one-electron excitation between the pair of ground-state orbitals mainly involved (H = HOMO, L = LUMO). ^b $\pi\text{-}\pi^*$ transition.



systematically computed the energies of the frontier molecular orbitals and the electronic transition energies for two series of complexes containing different ancillary ligands and an unsubstituted pyridazine: $[\text{Re}_2(\mu\text{-X})_2(\text{CO})_6(\mu\text{-pydz})]$ and $[\text{Re}_2(\mu\text{-H})(\mu\text{-X})(\text{CO})_6(\mu\text{-pydz})]$ ($\text{X} = \text{Cl}$, OPh , SPh , OOCPh). It is important to compute both the energy gap and the strongest MLCT electronic transition energy because in these species the HOMO \rightarrow LUMO transition is either symmetry forbidden or shows a very low oscillator strength. Therefore no clear correlation between the energy gap (as measured for instance from electrochemical data) and the maximum of the absorption spectra can be found. Moreover, the “ t_{2g} ” set of orbitals of the two rhenium atoms (involved in the MLCT transition) are not necessarily the six HOMOs for these species, since the presence of aromatic moieties in the ancillary ligands (OPh , SPh and OOCPh) determines the intercalation of some π orbitals between them. The data reported in Table 1 show that a more red-shifted absorption can be attained by using the hydrido derivative containing a μ -benzoato bridging ligand.

On this ground, we focused our attention on complex **2**, in which the pyridazine-4-carboxylic acid ligand is combined with one hydrido and one benzoato ligand, and on the analogous derivative containing the 4-diphenylamino-benzoate anion (compound **3**), due to the recognized ability²⁸ of the triarylamino moiety to promote charge separation after UV excitation. Fig. 2 shows the electronic density difference maps for the two most intense electronic transitions computed for **3**. In the MLCT transition it is clearly recognizable the electronic transfer from the $\text{Re}_2(\mu\text{-H})(\mu\text{-X})$ moiety to the pyridazine ligand (and its carboxylic substituent). Apparently the diphenylamino substituent is not directly involved in this transition, however it drastically changes the energy of the HOMO (not a rhenium centred orbital anymore). Its presence gives rise to an additional $\pi\text{-}\pi^*$ transition, that can be described as a charge transfer from the diphenylamino moiety to the benzoato bridging ligand, and is very similar to the corresponding transition computed

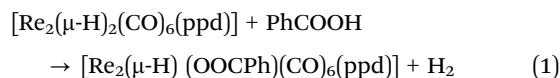
for the free ligand, even if their energies are quite different, as a consequence of their different charge distribution.

Synthesis of the designed molecules

A new synthetic strategy was designed for obtaining the hydrido-carboxylato derivatives **2** and **3** (whereas the dichloride complex **1** was prepared according to a literature procedure, involving the high temperature reaction of $\text{ReCl}(\text{CO})_5$ with 0.5 equivalents of pyridazine-4-carboxylic acid).^{1,24}

The three step procedure shown in Scheme 1 was followed, which is based on the peculiar reactivity of the tetrahedral hydrido-carbonyl cluster $[\text{Re}_4(\mu_3\text{-H})_4(\text{CO})_{12}]$ with bridging donor ligands.² Indeed this tetrahedral cluster, which possesses four electrons less than that required by the effective atomic number rule (56 instead of 60 valence electrons, v.e.s),²⁹ easily reacts with any (even weakly) donor species L .³⁰ By using pyridazine as a donor ligand, the main reaction products are the tetrametallic derivatives $[\text{Re}_4(\mu\text{-H})_4(\text{CO})_{12}(\mu\text{-pydz})_2]$, with a square geometry of the $\text{Re}(\mu\text{-H})\text{Re}$ skeleton and two pyridazine molecules bridging on opposite edges of the squares, in *trans* or *cis* positions with respect to the plane of the cluster.² The reaction produces also the dinuclear unsaturated complex $[\text{Re}_2(\mu\text{-H})_2(\text{CO})_6(\mu\text{-pydz})]$, which possesses 32 v.e.s, instead of the 34 v.e.s required by the effective atomic number rule.² Such dinuclear species is not formed in high yields in this reaction. In contrast dinuclear species of this kind are the only reaction products using 3,6-disubstituted pyridazine ligands, indicating the occurrence of a [2+2] fragmentation pathway of the parent tetranuclear cluster, without any spectroscopically recognizable intermediate. This is probably due to steric hindrance, which prevents the formation of the square planar species by destabilizing some key intermediate of the addition reaction. Analogously the selective formation of a dinuclear product is observed by reacting $[\text{Re}_4(\mu_3\text{-H})_4(\text{CO})_{12}]$ with 2,5-diphenyl-1,3,4-oxadiazole (ppd), which contains two aromatic substituents in the α position of the oxadiazole ring.²³ The labile ppd ligand can be easily replaced by a diazine molecule, and this provides a general route to $[\text{Re}_2(\mu\text{-H})_2(\text{CO})_6(\mu\text{-diazine})]$ complexes.

Therefore, at first the complex $[\text{Re}_2(\mu\text{-H})_2(\text{CO})_6(\mu\text{-ppd})]$ (**4**) was prepared in high yields by reacting $[\text{Re}_4(\mu_3\text{-H})_4(\text{CO})_{12}]$ with 2 equiv. of ppd, at room temperature in dichloromethane solution.²³ Then the mixed hydrido-carboxylato complex $[\text{Re}_2(\mu\text{-H})(\mu\text{-OOCPh})(\text{CO})_6(\mu\text{-ppd})]$ (**5**, Scheme 1) was obtained by reaction 1, which exploits the reactivity of the hydride ligands of **4** towards the acidic proton of the carboxylic acid.



Finally, substitution of the oxadiazole by the pyridazine-4-carboxylic acid afforded the desired product **2**. It is important to note that the introduction of the diazine ligand must be the last step of the synthesis, since both $[\text{Re}_4(\mu_3\text{-H})_4(\text{CO})_{12}]$ and $[\text{Re}_2(\mu\text{-H})_2(\text{CO})_6(\mu\text{-ppd})]$ (**4**) could react with the carboxylic substituent of the diazine ring, affording undesired byproducts.

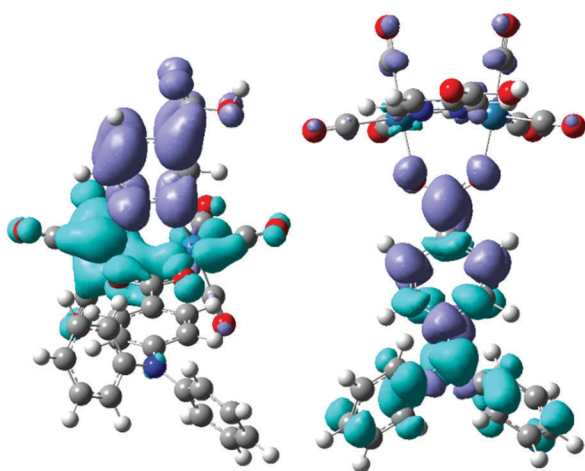
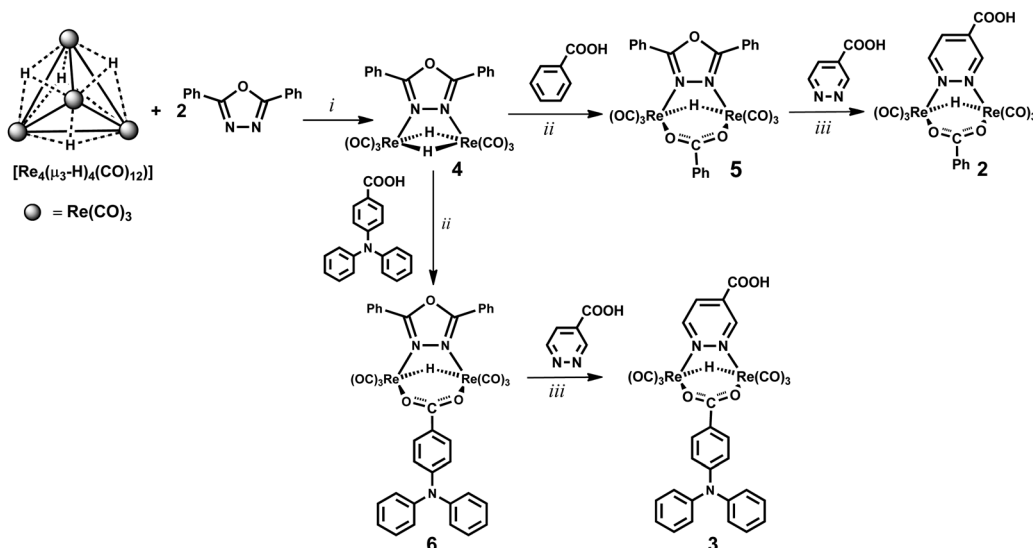


Fig. 2 Electronic density difference maps for the two most intense electronic transitions computed for the rhenium complex $[\text{Re}_2\text{H}(\text{OOC-TPA})(\text{CO})_6(\text{pydz-4-COOH})]$ (**3**): the MLCT at 530 nm (left) and the $\pi\text{-}\pi^*$ at 344 nm (right). Light blue indicates a decrease in electron density, blue indicates an increase.





Scheme 1 Synthetic pathways for the preparation of rhenium complexes **2** and **3**. Reaction conditions: (i) CH_2Cl_2 , room temperature, 24 h; (ii) toluene reflux, 2 h; (iii) THF reflux, 12 h.

The same synthetic procedure has been exploited to carry out the synthesis of complex **3**, containing the 4-diphenylaminobenzoate anion as an ancillary ligand.

Electrochemical characterization

The cyclic voltammetry analysis has been performed on the derivatives **1-COOMe**, **2-COOMe** and **3-COOMe**, containing the methyl ester of the pydz-4-COOH ligand, in order to avoid the reduction of the acidic proton, affording H_2 evolution. The cyclic voltammetry behaviour in acetonitrile, at 298 K, is reported in Fig. 3 and the peak potentials, together with the related HOMO and LUMO energy values, are collected in Table 2. In the cathodic region all the complexes show one monoelectronic reduction peak at about -0.9 V (vs. Fc^+/Fc , see Table 2), which is centred on the diazine ligand and whose chemical reversibility is confirmed by the presence of its anodic counterpart. Moreover, the position of this peak is unaffected by the scan rate, in agreement with the electrochemical reversibility of this

Table 2 First reduction and oxidation peak potentials ($E_{\text{p,c}}$ and $E_{\text{p,a}}$) and electrochemical (ΔE_{e}) and spectroscopic (ΔE_{s}) energy gaps of the complexes. Potentials are referred to the Fc^+/Fc couple^b in the operating medium (MeCN, 0.1 M [TBA][PF₆]). Scan rate 0.2 V s^{-1}

Compound	$E_{\text{p,c}}$ [V]	$E_{\text{p,a}}$ [V]	ΔE_{e} [eV]	ΔE_{s} [eV]	E_{LUMO} [eV]	E_{HOMO} [eV]
1-pydz	−1.34	1.31	2.66	3.31	−3.46	−6.11
1-COOMe	−0.91	1.34	2.25	2.98	−3.88	−6.14
2-COOMe	−0.89	1.11, 1.32	2.00	2.56	−3.90	−5.91
3-COOMe	−0.97	0.66, ^c 1.07, 1.32	1.63	2.56	−3.82	−5.46

^a The spectroscopic (ΔE_{s}) energy gaps is the energy associated with the electronic transition determined from the maximum of the $^1\text{MLCT}$ absorption band. ^b Fc^+/Fc potential is 0.385 V vs. SCE in acetonitrile solution. ^c First oxidation peak potential not located on the metal core.

reductive process. The peak potential is slightly modulated by the different nature of the ancillary ligand (see Table 2). Comparing the CV curve of the complexes **1–3** with those of analogous complexes of this family containing an unsubstituted pydz ligand, a shift in the positive direction of the reduction potential can be observed (see for instance **1-pydz** in Table 2),³¹ consistently with the presence of the electron-withdrawing substituent on the aromatic ring. These features are in agreement with the localization of the reduction process on the diazine ring, as indicated also by the description of the LUMO provided by the DFT computations.

In all cases, the first reduction peak is followed by another reversible reduction peak, which is probably centred on the diazine ligand, as suggested by the invariance of the potential on varying the ancillary ligands. This agrees with DFT computations, that indicated that the two lowest LUMOs are both π^* orbitals of the diazine.

In the anodic region significant differences have been observed for the three complexes. **1-COOMe** exhibits an oxidation potential at 1.34 V (vs. Fc^+/Fc), which corresponds to a metal-centred bi-electronic oxidation and appears to be irreversible.¹

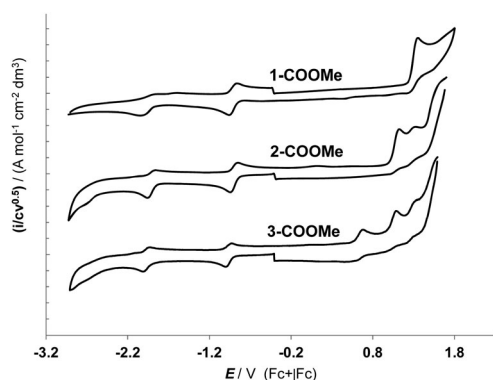


Fig. 3 Normalized CV features of the methyl-ester derivatives of the three dyes on GC electrodes, in MeCN +0.1 M $(\text{NBu}_4)\text{PF}_6$ solution, at 0.2 V s^{-1} with ohmic drop compensation.



However, it becomes partially reversible at a higher scan rate, as demonstrated by the presence of the return peak.¹ This clearly points to an electrochemically quasi-reversible electron transfer step and a subsequent chemical step. The resulting electrochemical HOMO–LUMO gap was 2.25 eV and the HOMO and the LUMO energy levels of **1** are calculated to be –6.14 eV and –3.88 eV respectively.

Interesting differences appear in the anodic behaviour of the complexes **2-COOMe** and **3-COOMe**. As already observed in the derivatives containing alcoholate or phenolate anions as ancillary ligands,³ also in these complexes a close sequence of two monoelectronic oxidation peaks is observed, instead of the bi-electronic one. These two oxidation processes are clearly localized on the metal core (as confirmed also by the computational analysis)³ and strongly indicate that the oxidation process is markedly different in the case of the oxygen-bridged derivatives with respect to the dichloride ones, also when only one oxygenated ancillary ligand is present. This might be related to the closer Re(μ -X)₂Re scaffold and to the harder nature of the bridging ligands in the case of the derivatives containing oxygen ligands. This feature could afford a less-efficient stabilization of the cationic products, with respect to the softer anions, such as halides.

The first oxidation peak for **2-COOMe** is observed at +1.11 V (vs. Fc^+/Fc), and it appears chemically irreversible. It is followed by a second peak, at 1.32 V, which is reversible in the whole scan rate range explored. This lower oxidation potential affords a higher HOMO level, in agreement with the DFT calculations, and the resulting electrochemical energy gap is reduced to 2.00 eV.

On the other hand, for complex **3-COOMe**, beside the two oxidation peaks centred on the metal core (at 1.07 V and at 1.32 V), another chemically reversible oxidation peak at +0.66 V is observed. This peak is clearly attributed to the formation of a radical cation on the triarylamine (TPA) moiety, as indicated by the comparison with the electrochemical behaviour of the free 4-diphenylamino-benzoic acid ($E_{\text{p,a}} = 0.70$ V). This confirms the lack of communication between the pyridazine ring and the triarylamine moiety, which is verified also in the UV-Vis absorption data (see the following section).³² In spite of the lack of electronic communication, the presence of the TPA moiety affects (even if slightly) the position of the reduction of the pyridazine ring, which results at more negative potential values than in **2**. The slight destabilization of the LUMO level is also confirmed by the computational data (see Table 1).

Even so, the introduction of the TPA moiety in **3** affords a strong reduction of the electrochemical energy gap, which does not parallel the spectroscopic (absorption) gap, as in the case of **1** and **2**. This is consistent with the fact that the electronic absorption arises from a ¹MLCT transition that in **3** does not involve the HOMO level, which is centred on the TPA moiety. In the case of dyes **1** and **2**, in contrast, the electrochemical and spectroscopical energy gaps does parallel each other, although the spectroscopical gaps are always higher than the electrochemical ones, since the electronic transition responsible for the ¹MLCT absorption band is not a pure HOMO–LUMO transition.

Photophysical characterization

The UV-Vis absorption spectra of the three dyes in diluted toluene and EtOH solution are reported in Table 3, while Fig. 4 shows the spectra in toluene solution. In toluene solution all the complexes exhibit a broad and featureless absorption band that covers a large part of the visible spectrum between 350 nm and 550–600 nm. The predicted red-shift in the absorption maximum (see the Computational study section) for **2** and **3**, with respect to complex **1**, is observed. This feature indicates that this band is ascribable to the metal-to-ligand $d\pi(\text{Re}) \rightarrow \pi^*(\text{diazine})$ charge transfer transition (¹MLCT). This attribution is also supported by the weak intensity of these bands and by the observed solvatochromic behaviour. Indeed a significant blue-shift of the absorption maximum has been observed in a more polar solvent as EtOH (see Table 3). This broad band arises from the convolution of multiple transitions, as testified by the more or less pronounced shoulders observed at longer wavelengths. As evidenced by TD-DFT computations, the electronic transitions responsible for the MLCT absorption band involve as starting orbitals the metal-centred HOMO–*n* orbitals and, as final orbitals, the two low lying π^* orbitals of the diazine, LUMO and LUMO+1.

Complex **3** displays another high energy absorption band, at 350 nm, whose position is independent of the polarity of the solvent. In agreement with the electrochemical and computational data, this band is attributed to the $\pi \rightarrow \pi^*$ transition involving the triarylamine moiety.

DSSC performances

Photovoltaic measurements have been carried out to evaluate the potential of the three new rhenium complexes as dyes in DSSC devices. The main photovoltaic parameter performances of the solar cells under AM 1.5 G at 1 Sun (1000 W m^{–2}) illumination are presented in Table 4 and the current–voltage (*J*–*V*) curves are reported in Fig. 5. Two different electrolytes were used, both based on the iodide/triiodide redox couple. The first one (HSE, Dyesol) is commercially available, while the latter one (labeled El14) was developed in-house, reducing the concentration of iodine (and thus of iodide), which strongly absorbs below 450 nm, competing with the rhenium complexes absorption. Actually, an increase in the photogenerated currents can be observed (see Table 4 and Fig. 5) for all the rhenium-based dyes when El14 substitutes HSE and it is confirmed by the incident photon-to-current conversion efficiency (IPCE) plots for the DSSCs sensitized by **1**, **2** and **3**, in the presence of the two different electrolytes shown in Fig. 6.

Table 3 UV-Vis MLCT absorption data of the dyes in solution and on TiO₂ film

Compound	$\lambda_{\text{max}}^a/\text{nm}$	$\epsilon^a (10^4 \text{ M}^{-1} \text{ cm}^{-1})$	$\lambda_{\text{max}}^b/\text{nm}$	$\lambda_{\text{max}}^c/\text{nm}$
1	416	0.65	375	405
2	487	0.67	430	440
3	487	0.43	425	443

^a Toluene solution (1×10^{-4} M). ^b EtOH solution (5×10^{-4} M). ^c TiO₂ film.



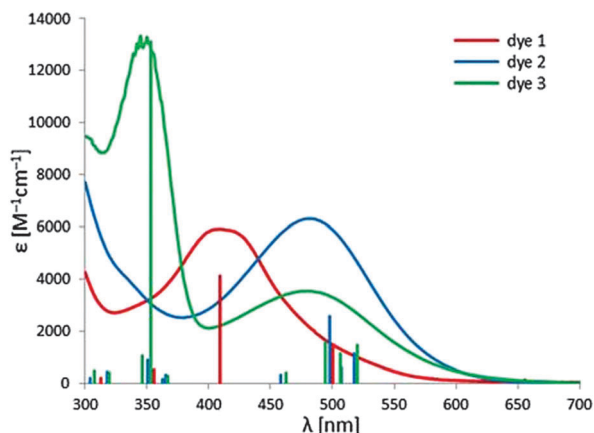


Fig. 4 UV-Vis absorption spectra of the three dyes in toluene solution. A comparison with excitation energies and oscillation strengths computed in toluene is reported (vertical lines).

Table 4 Photovoltaic parameters using homemade E114 electrolyte (first line for each entry) and commercial HSE electrolyte (Dyesol)

Compound	J_{SC} (mA cm ⁻²)	V_{OC} (mV)	Fill factor	η (%)
N719	12	0.78	0.32	3.0
	13.6 ^a	0.73 ^a	0.63 ^a	6.3 ^a
1	2.2	0.54	0.61	0.8
	1.8 ^a	0.53 ^a	0.70 ^a	0.6 ^a
2	1.7	0.48	0.6	0.45
	1.2 ^a	0.46 ^a	0.67 ^a	0.38 ^a
3	3.3	0.51	0.57	0.97
	2.9 ^a	0.51 ^a	0.70 ^a	1.0 ^a

^a Photovoltaic performances based on the commercial HSE electrolyte (Dyesol).

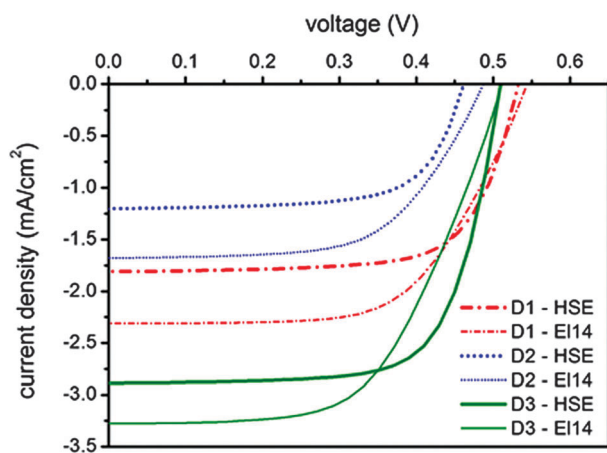


Fig. 5 Current–voltage curves for the DSSCs sensitized by the rhodium-based dyes under 1 Sun illumination, AM 1.5. Two different iodide/triiodide electrolytes were used, a commercially available one (HSE, Dyesol) and another developed in-house (labeled E114).

The best performing cell, sensitized by complex 3, gives an overall power conversion efficiency of 1% (with short-circuit photocurrent density $J_{SC} = 2.9$ mA cm⁻², open-circuit photovoltage $V_{OC} = 0.51$ mV, fill factor $ff = 0.70$). These values are far

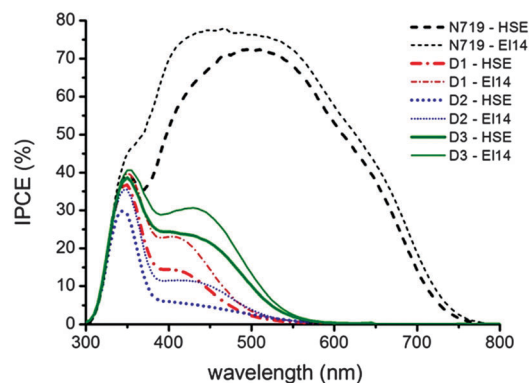


Fig. 6 IPCE spectra for DSSC devices sensitized by different dyes: **1** with a HSE electrolyte (thick red dash-dotted line) and with an E114 electrolyte (thin red dash-dotted line); **2** (dotted blue lines); **3** (solid green lines); **N719** as reference dye (dashed black lines). Two different iodide/triiodide electrolytes were used, a commercially available one (HSE, Dyesol) and another developed in-house (labeled E114).

from those obtained under similar conditions by cells assembled using the conventional ruthenium N719 dye. Indeed these dyes absorb in a narrow spectral range compared to N719 and in addition, below 450 nm, they compete with iodide/triiodide-based electrolytes in the absorption of light. However, the maximum of the absorption spectrum is not the only factor which determines the overall characteristics of the cell. Indeed, despite the red-shift observed in the absorption spectrum, mainly due to the raise of the HOMO level, complex 2 displays the worst performances. This result is probably due to the particularly low LUMO level (see Table 2) which, as in the case of complex 1, hampers the electron injection into the TiO₂ conduction band (−3.9 eV).

The influence of the ancillary ligand is confirmed by the results obtained with complex 3. The replacement of the simple benzoate anion with the carboxylate derivative of the triarylamine passing from 2 to 3 has a double role. The LUMO level of the complex was slightly raised, then allowing the injection on the TiO₂. Most importantly, the hole transport unit, now localized on the TPA moiety, is placed far away from TiO₂ and can better interact with the electrolyte. Moreover the higher steric hindrance of the TPA unit, with respect to the simple benzoate anion, reduces the recombination of the electrons, with the redox species, on the surface of TiO₂. This insulating effect is in agreement with the slightly higher V_{OC} value obtained for complex 3.

Conclusions

Dirhenium complexes designed as potential sensitizers in DSSCs have been prepared, characterized and tested in operating solar cells. Light-to-current conversion has been really observed, although with a low efficiency. The best results have been obtained for the hydrido-carboxylato complex 3, in which the introduction of the triarylamine moiety further suppresses the recombination of the injected electron with the oxidized state of the complex, thus improving the charge separation on



TiO₂. Moreover the bulky triarylamine moiety also reduces the back reaction of the injected electron with the electrolyte.

Compared to the standard ruthenium-based N719 dye, the maximum power conversion efficiency of 1% obtained for **3** is obviously too low. However, the data here collected have provided a sufficiently clear picture of the critical factors and of the lines to follow for improving the performances. The future molecular design will take advantage from the possibility to independently tune the ground-state energy levels of the dirhenium diazine complexes, since their frontier orbitals are clearly localized on different regions of the molecule.

So a different choice of the diazine ligand will permit to raise the LUMO level, to overcome the most serious problem concerning the electron injection into TiO₂, which is certainly represented by the excessive stabilization of the π^* orbitals of the diazine in the complexes here investigated.

On the other hand, wider light harvesting and efficient hole-transfer towards the periphery of the molecule could be obtained through the introduction of a more conjugated moiety on the triarylamine ancillary ligand. Moreover metal oxide photoelectrodes other than TiO₂, such as SnO₂, and different electrolytes, such as those based on cobalt complexes, could be used. Work towards these directions is in progress in our laboratory.

However, it must be pointed out that, due to the novelty of the employed dyes, the present work did not aim to optimize the cell performances, but rather to understand the relationships between the molecular structure of the complexes and their spectral and photovoltaic properties. This result has been attained.

Two final observations, from the point of view of coordination chemistry. The survival of hydrido complexes in the media and under the operating conditions of solar cells is noteworthy. Clearly the bridging coordination of the hydride contributes to this peculiar stability. Moreover, the three-step synthetic procedure here developed has a wider scope than the preparation of the compounds here described and opens the way to the obtainment of a large number of dinuclear complexes of this family, tailored for specific applications.

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Notes and references

- 1 D. Donghi, G. D'Alfonso, M. Mauro, M. Panigati, P. Mercandelli, A. Sironi, P. Mussini and L. D'Alfonso, *Inorg. Chem.*, 2008, **47**, 4243–4255.
- 2 M. Panigati, D. Donghi, G. D'Alfonso, P. Mercandelli, A. Sironi and L. D'Alfonso, *Inorg. Chem.*, 2006, **45**, 10909–10921.
- 3 A. Raimondi, M. Panigati, D. Maggioni, L. D'Alfonso, P. Mercandelli, P. Mussini and G. D'Alfonso, *Inorg. Chem.*, 2012, **51**, 2966–2975.
- 4 (a) E. Quartapelle Procopio, M. Mauro, M. Panigati, D. Donghi, P. Mercandelli, A. Sironi, G. D'Alfonso and L. De Cola, *J. Am. Chem. Soc.*, 2010, **132**, 14397–14399; (b) M. Panigati, M. Mauro, D. Donghi, P. Mercandelli, P. Mussini, L. De Cola and G. D'Alfonso, *Coord. Chem. Rev.*, 2012, **256**, 1621–1643.
- 5 (a) M. Mauro, E. Quartapelle Procopio, Y. Sun, C.-H. Chien, D. Donghi, M. Panigati, P. Mercandelli, P. Mussini, G. D'Alfonso and L. De Cola, *Adv. Funct. Mater.*, 2009, **19**, 2607–2614; (b) M. Mauro, C.-H. Yang, C.-Y. Shin, M. Panigati, C.-H. Chang, G. D'Alfonso and L. De Cola, *Adv. Mater.*, 2012, **24**, 2054–2058.
- 6 F. Nastasi, F. Puntoriero, M. Natali, M. Mba, M. Maggini, P. Mussini, M. Panigati and S. Campagna, *Photochem. Photobiol. Sci.*, 2015, **14**, 909–918.
- 7 B. O'Regan and M. Grätzel, *Nature*, 1991, **353**, 737–740.
- 8 T. M. Brown, F. De Rossi, F. Di Giacomo, G. Mincuzzi, V. Zardetto, A. Reale and A. Di Carlo, *J. Mater. Chem. A*, 2014, **2**, 10788–10817.
- 9 A. Fakhruddin, R. Jose, T. M. Brown, F. Fabregat-Santiago and J. Bisquert, *Energy Environ. Sci.*, 2014, **7**, 3952–3981.
- 10 (a) M. Nazeeruddin, A. Kay, I. Rodicio, R. Humphry-Baker, E. Müller, P. Liska, N. Vlachopoulos and M. Grätzel, *J. Am. Chem. Soc.*, 1993, **115**, 6382–6390; (b) K. Kalyanasundaram and M. Grätzel, *Coord. Chem. Rev.*, 1998, **177**, 347–414.
- 11 S. Campagna, F. Puntoriero, F. Nastasi, G. Bergamini and V. Balzani, *Top. Curr. Chem.*, 2007, **280**, 117–214.
- 12 M. Nazeeruddin, F. De Angelis, S. Fantacci, A. Selloni, G. Viscardi, P. Liska, S. Ito, B. Takeru and M. Grätzel, *J. Am. Chem. Soc.*, 2005, **127**, 16835–16847.
- 13 D. Kuang, S. Ito, B. Wenger, C. Klein, J. Moser, R. Humphry-Baker, S. Zakeeruddin and M. Grätzel, *J. Am. Chem. Soc.*, 2006, **128**, 4146–4154.
- 14 S. Mathew, A. Yella, P. Gao, R. Humphry-Baker, B. F. E. Curchod, N. Ashari-Astani, I. Tavernelli, U. Rothlisberger, M. K. Nazeeruddin and M. Grätzel, *Nat. Chem.*, 2014, **6**, 242–247.
- 15 (a) S. Ferrere and B. A. Gregg, *J. Am. Chem. Soc.*, 1998, **120**, 843–844; (b) P. Balraju, M. Kumar, M. S. Roy and G. D. Sharma, *Synth. Met.*, 2009, **159**, 1325–1331.
- 16 (a) T. Bessho, E. C. Constable, M. Grätzel, A. H. Redondo, C. E. Housecroft, W. Kylberg, M. K. Nazeeruddin, M. Neuburger and S. Schaffner, *Chem. Commun.*, 2008, 3717–3719; (b) C. L. Linfoot, P. Richardson, T. E. Hewat, O. Moudam, M. M. Forde, A. Collins, F. White and N. Robertson, *Dalton Trans.*, 2010, **39**, 8945–8956.
- 17 (a) W. Wu, X. Xu, H. Yang, J. Hua, X. Zhang, L. Zhang, Y. Longa and H. Tian, *J. Mater. Chem.*, 2011, **21**, 10666–10671; (b) E. A. M. Geary, L. J. Yellowlees, L. A. Jack, I. D. H. Oswald, S. Parsons, N. Hirata, J. R. Durrant and N. Robertson, *Inorg. Chem.*, 2005, **44**, 242–250.
- 18 B. Bozic-Weber, E. C. Constable and C. E. Housecroft, *Coord. Chem. Rev.*, 2013, **257**, 3089–3106.
- 19 (a) J. B. Asbury, E. Hao, Y. Wang and T. Lian, *J. Phys. Chem. B*, 2000, **104**, 11957–11964; (b) N. A. Anderson, X. Ai, D. Chen, D. L. Mohler and T. Lian, *J. Phys. Chem. B*, 2003, **107**, 14231–14239.



- 20 G. M. Hasselmann and G. J. Meyer, *J. Phys. Chem. B*, 1999, **103**, 7671–7675.
- 21 Y. Chen, W. Liu, J.-S. Jin, B. Liu, Z.-G. Zou, J.-L. Zuo and X.-Z. You, *J. Organomet. Chem.*, 2009, **694**, 763–770.
- 22 (a) M. A. Andrews, S. W. Kirtley and H. D. Kaesz, *Inorg. Chem.*, 1977, **16**, 1556–1561; (b) J. R. Johnson and H. D. Kaesz, *Inorg. Synth.*, 1978, **18**, 60–62.
- 23 M. Mauro, M. Panigati, D. Donghi, P. Mercandelli, P. Mussini, A. Sironi and G. D'Alfonso, *Inorg. Chem.*, 2008, **47**, 11154–11165.
- 24 E. Ferri, D. Donghi, M. Panigati, G. Prencipe, L. D'Alfonso, I. Zanoni, C. Baldoli, S. Maiorana, G. D'Alfonso and E. Licandro, *Chem. Commun.*, 2010, **46**, 6255.
- 25 C. Adamo and V. Barone, *J. Chem. Phys.*, 1999, **110**, 6158–6170.
- 26 J. P. Perdew, K. Burke and M. Ernzerhof, *Phys. Rev. Lett.*, 1997, **78**, 1396.
- 27 *Gaussian 09 (revision D.01)*, Gaussian Inc., Wallingford, CT, 2013.
- 28 B.-G. Kim, K. Chung and J. Kim, *Chem. – Eur. J.*, 2013, **19**, 5220–5230.
- 29 (a) R. Saillant, G. Barcelo and H. Kaesz, *J. Am. Chem. Soc.*, 1970, **92**, 5739–5741; (b) R. D. Wilson and R. Bau, *J. Am. Chem. Soc.*, 1976, **98**, 4687–4689; (c) N. Masciocchi, G. D'Alfonso, W. Kockelmann, W. Schäfer and A. Sironi, *Chem. Commun.*, 1997, 1903–1904.
- 30 (a) T. Beringhelli and G. D'Alfonso, *J. Chem. Soc., Chem. Commun.*, 1994, 2631–2632; (b) T. Beringhelli, G. D'Alfonso and M. Zarini, *J. Chem. Soc., Dalton Trans.*, 1995, 2407–2415; (c) T. Beringhelli, G. D'Alfonso and M. G. Garavaglia, *J. Chem. Soc., Dalton Trans.*, 1996, 1771–1773.
- 31 The reduction potential is -1.34 V for the parent complex $[\text{Re}_2(\mu\text{-Cl})_2(\text{CO})_6(\mu\text{-pyridazine})]$ and -1.58 V for $[\text{Re}_2(\mu\text{-Cl})_2(\text{CO})_6(\mu\text{-4,5-Mez-pyridazine})]$ containing an electron-rich pyridazine ligand.
- 32 The same behaviour has been observed in the complex containing, as an ancillary ligand, the carboxylate derivative of cyclopentadithiophene. E. Quartapelle Procopio, V. Bonometti, M. Panigati, P. Mercandelli, P. Mussini, T. Benincori, G. D'Alfonso and F. Sannicolò, *Inorg. Chem.*, 2014, **53**, 11242–11251.

