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Ruthenium water oxidation catalysts containing the non-planar tetradentate ligand, biisoquinoline dicarboxylic acid (biqaH₂)[†]

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Two ruthenium complexes containing the tetradentate ligand [1,1'-biisoquinoline]-3,3'-dicarboxylic acid, and 4-picoline or 6-bromoisoquinoline as axial ligands have been prepared. The complexes have been fully characterised and initial studies on their potential to function as molecular water oxidation catalysts have been performed. Both complexes catalyse the oxidation of water in acidic media with Ce^{IV} as a stoichiometric chemical oxidant, although turnover numbers and turnover frequencies are modest when compared with the closely related Ru-bda and Ru-pda analogues. Barriers for the water nucleophilic attack and intermolecular coupling pathways were obtained from density functional theory calculations and the crucial influence of the ligand framework in determining the most favourable reaction pathway was elucidated from a combined analysis of the theoretical and experimental results.

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Introduction

As society strives to wean itself off its reliance on carbon-based fuels, alternative new technologies need to be discovered and developed to facilitate this revolution.^{1–5} Researchers in the field of *Artificial Photosynthesis* aim to achieve this by using sunlight to drive the production of so called *solar fuels* (e.g. H₂, MeOH) from the oxidation of water and the reduction of Earth-abundant substrates (e.g. H⁺, CO₂).^{6–8} Of the key steps involved in this process (light harvesting, charge separation, substrate reduction and water oxidation),⁹ water oxidation is considered to be a major bottle neck and new catalysts capable of facilitating this reaction are therefore constantly in high demand.^{10–13} The group of Licheng Sun has been particularly active in this area and has published a number of molecular water oxidation catalysts (WOCs), including the Ru-bda catalysts (bdaH₂ = bipyridine dicarboxylic acid),^{14–19} some of the fastest and most stable molecular WOCs reported to date. Much work was carried out on varying the axial ligands of these catalysts, which resulted in catalysts with increased turnover numbers (TONs) and turnover frequencies (TOFs).^{20–23} However, less work has been done with respect to studying equatorial ligand derivatives.^{24–28} In the particular example where the bdaH₂

ligand was exchanged for the pdaH₂ ligand (pdaH₂ = phenanthroline dicarboxylic acid), a profound effect on the water oxidation mechanism was observed, resulting in a switch from an intermolecular coupling mechanism (I2M) to water nucleophilic attack (WNA). The reasoning behind this switch was proposed to be caused by changes in the reorganisation energies throughout the catalytic cycle as a result of the increased rigidity of the ligand. Intrigued by these results, we decided to study the complementary effect of enforced non-planarity as opposed to enforced planarity and rigidity in the equatorial ligand plane.

Herein, we report a new variation of the Ru-bda-type WOCs, based on the non-planar ligand [1,1'-biisoquinoline]-3,3'-dicarboxylic acid (biqaH₂). The synthesis of the new equatorial ligand and two of its ruthenium complexes along with their characterisation and evaluation as WOCs is presented forthwith.

Results and discussion

Synthesis

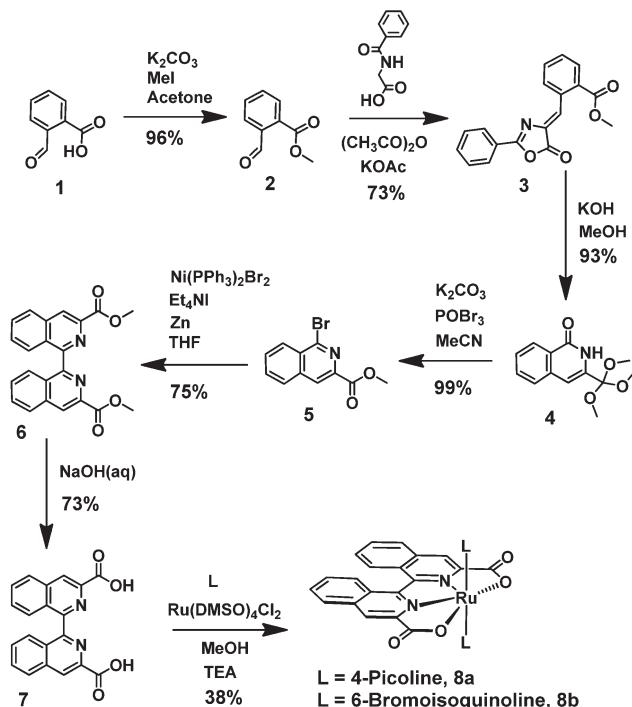
The biqaH₂ ligand **7** was prepared in good overall yield over six steps starting from cheap and readily available precursors (Scheme 1). The first four steps followed previously reported procedures,^{29–31} whilst the conditions for the final two steps were adapted from reported procedures of analogous compounds.^{32,33} An initial five step synthesis of the biqaH₂ ligand **7**, starting from 3-methylisoquinoline and forming the diacid

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Scheme 1 Synthesis of the biqaH₂ ligand 7 and subsequent ruthenium complexes 8a and 8b.

via the chromic acid oxidation of 3,3'-dimethyl-1,1'-biisoquinoline, was originally planned,³⁴ however, this route proved to be unsuccessful due to the inability to selectively oxidise the methyl groups without affecting the isoquinoline moiety. The six step procedure without the need for the harsh oxidation step was consequently adopted (Scheme 1). Coordination of the BiqaH₂ ligand to ruthenium and substitution of the axial ligands with either 4-picoline (pic) or 6-bromoisoquinoline (Br-isoq) was accomplished in a single step to afford complexes 8a and 8b respectively. A mixture of compounds was formed under these conditions but the desired complexes were always the major product and were isolable *via* simple precipitation or column chromatography. Attempts to coordinate the ruthenium in a 2-step sequence and first isolate the Ru(biqa)(DMSO)₂ intermediate were unsuccessful and tended to yield mixtures of compounds, none of which were the desired intermediate compound as analysed by HPLC, MS and ¹H NMR. The isolated complexes 8a and 8b were fully characterised by standard analytical techniques (HPLC, 1D/2D NMR, MS, UV-vis, EA and voltammetry).

X-ray structure

Additionally for complex 8b, single crystals suitable for X-ray diffraction were grown and the absolute structural configuration was confirmed (see Fig. 1 and ESI[†]). Interestingly, the coordination environment of the Ru centre in complex 8b is very similar to analogous Ru-bda-type structures previously reported,^{15,16,35} despite the obvious distortion of the biisoquinoline moiety. The strain imposed by the inherent non-planarity

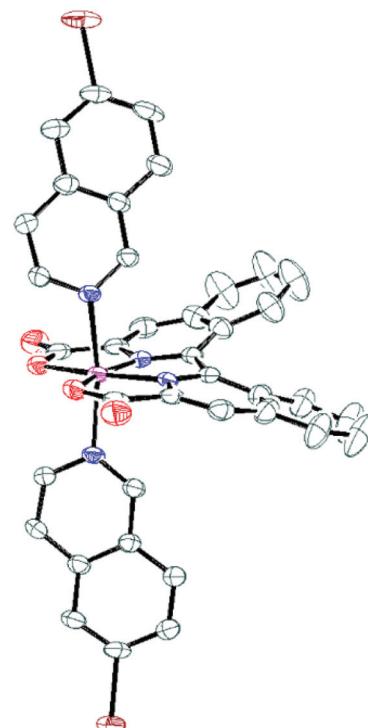


Fig. 1 ORTEP plot at 50% probability for compound 8b. Colour code: C, grey; Br, brown; N, blue; O, red; Ru, magenta. Hydrogens have been omitted for clarity.

of the biisoquinoline, combined with an almost planar coordination of the donor atoms to ruthenium, is therefore manifested in a bending of the isoquinoline carboxylates rather than a twisted coordination geometry. This is thought to occur because the octahedral geometry is already distorted due to the widening of the O–Ru–O angle from the ideal 90° to *ca.* 120°, similar to Ru-bda structures, and therefore does not permit any further distortion of the primary coordination sphere that may have been expected as a result of the non-planarity of the biqa ligand backbone.

Electrochemistry

Cyclic voltammetry (CV) and differential pulse voltammetry (DPV) in both organic and aqueous solutions were used to probe the electrochemical properties of compounds 8a and 8b. Not surprisingly, the electrochemical behaviour of the Ru-biqa complexes was found to be very similar to the analogous Ru-bda complexes. The CVs and DPVs in DCM (0.1 M TBA(PF₆)) + 20% trifluoroethanol (TFE) indicated a single reversible redox process for the Ru^{III–II} couple at 0.87 V and 0.92 V (*vs.* NHE)³⁶ for 8a and 8b respectively (Table 1, Fig. 2 and ESI[†]). In aqueous solvent (0.1 M HOTf_(aq) + TFE) three redox couples were observed due to the coordination of water permitting access to higher oxidation states through proton coupled electron transfer (PCET)^{37,38} and a large current increase above *ca.* 1.3 V was attributed to catalytic water oxidation. $E_{1/2}$ values for Ru^{III–II}, Ru^{IV–III}, and Ru^{V–IV} are given in Table 1 and their shifts



Table 1 Summary of electrochemical data for complexes **8a** and **8b**^a

| Complex | $E_{1/2}$ Ru ^{III-II} (ΔE_p) | $E_{1/2}$ Ru ^{IV-III} (ΔE_p) | $E_{1/2}$ Ru ₂ ^{IV-V} |
|------------------------|---|---|---|
| 8a ^b | 0.87 (80) | — | — |
| 8a ^c | 0.67 (50) | 1.11 (60) | 1.30 ^e |
| 8b ^b | 0.92 (80) | — | — |
| 8b ^d | 0.73 (50) | 1.15 (40) | 1.31 ^e |

^a $E_{1/2}$ values in V vs. NHE (ref. 36) and ΔE in mV. ^b 0.5 mM in 0.1 M TBA(PF₆) DCM with 20% TFE. ^c 0.3 mM in 0.1 M HOTf_(aq) with 20% TFE. ^d 0.3 mM in 0.1 M HOTf_(aq) with 33% TFE. ^e Values taken from DPV.

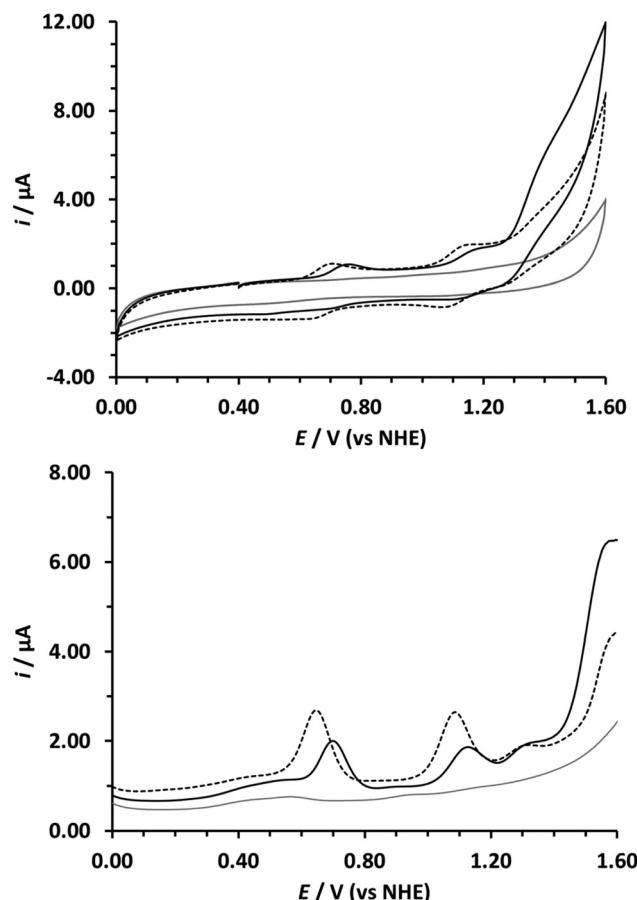


Fig. 2 Cyclic voltammograms (top) and differential pulse voltammograms (bottom) of **8a** (dashed black) and **8b** (solid black). $[8a] = 0.3$ mM in 0.1 M HOTf_(aq) with 20% TFE, $[8b] = 0.3$ mM in 0.1 M HOTf_(aq) with 33% TFE. Electrolyte only blank scan (solid grey). CV scan rate = 100 mV s⁻¹. DPV sweep rate = 20 mV s⁻¹. GC disc as working electrode, Pt rod as counter electrode, Ag/AgCl as reference electrode (values converted to NHE using $E_{1/2}$ Ag/AgCl = 0.2 V vs. NHE).

as a function of changes in pH are presented in the corresponding Pourbaix diagrams in the ESI (Fig. S7 and S8†). Previous work by Sun *et al.* showed that small quantities of TFE had little or no effect on the electrochemistry of Ru-bda complexes in aqueous solutions at pH 1 but at higher concentrations the catalytic currents were suppressed.¹⁶ TFE coordination to ruthenium has also been observed for the

Ru(Mebimpy)(bpy)(OH₂)²⁺ WOC previously reported by Meyer *et al.*³⁹ In our electrochemical studies we also found that high TFE quantities could significantly influence the electrochemical behaviour of the Ru-biqa complexes. Due to the poor solubility of **8b** in 0.1 M aqueous HOTf, 33% TFE was required to obtain a 0.3 mM solution but no effect on the electrochemical response was observed. In the case of **8a** only 20% of TFE was required to fully dissolve the complex, again without disturbing the electrochemical response, however, the Ru^{IV-III} and Ru^{V-IV} waves that were initially observed for catalyst **8a** were greatly diminished at TFE concentrations of 33% and above (ESI Fig. S6†). These results indicated an equilibrium between H₂O and TFE coordination for the Ru-biqa complexes but as the initial Ru-biqa complexes are coordinatively saturated this would require prior decoordination of one of the biqa ligand donor atoms. A potentially related observation made during the electrochemical experiments was a change in colour of the acidic aqueous solutions of the catalysts from blue to yellow. UV-vis spectroscopy was used to probe this further.

UV-vis

Changes in the absorbance spectra and a number of other observations were recorded after dissolving catalysts **8a** and **8b** in TFE-0.1 M HOTf_(aq) mixed solvent (see ESI†): the rate of decay at 600 nm was identical for both catalysts, the spectra of the resulting pale yellow solutions contained very similar absorbance features for both compounds, no change in absorbance was observed for either compound when the TFE-0.1 M HOTf_(aq) mixed solvent was replaced with a TFE-water mix, and finally, the process could be reversed by the addition of a reducing agent like sodium ascorbate. The conclusions that can be drawn from these results are; (i) the process involves a reversible oxidation event, (ii) the oxidation is pH dependent, and (iii) the axial ligands have little effect on the conversion. The final proof was obtained by carrying out redox titrations of the catalyst solutions with the strong oxidant, ceric ammonium nitrate (CAN). The same spectra as those observed after one hour standing in air were obtained after the addition of 1.0 eq. of CAN, thus confirming the 1-electron oxidation of the catalysts as the reason for the observed colour changes.

Chemical water oxidation

Given the structural similarities between the Ru-biqa WOCs and the Ru-bda WOCs, conditions for assessing the chemical water oxidation were chosen to give the closest possible comparison. Catalysts **8a** and **8b** were therefore added to a solution of excess CAN oxidant in 0.1 M HOTf_(aq) and the oxygen concentration was measured in the aqueous phase using a fluorescence-based oxygen probe. Fig. 3 shows the oxygen evolution profiles at various catalyst concentrations of **8a** and **8b**. What is evident from these plots is that catalyst **8a** and catalyst **8b** behave quite differently under the same reaction conditions: catalyst **8a** has a fairly smooth oxygen evolution profile

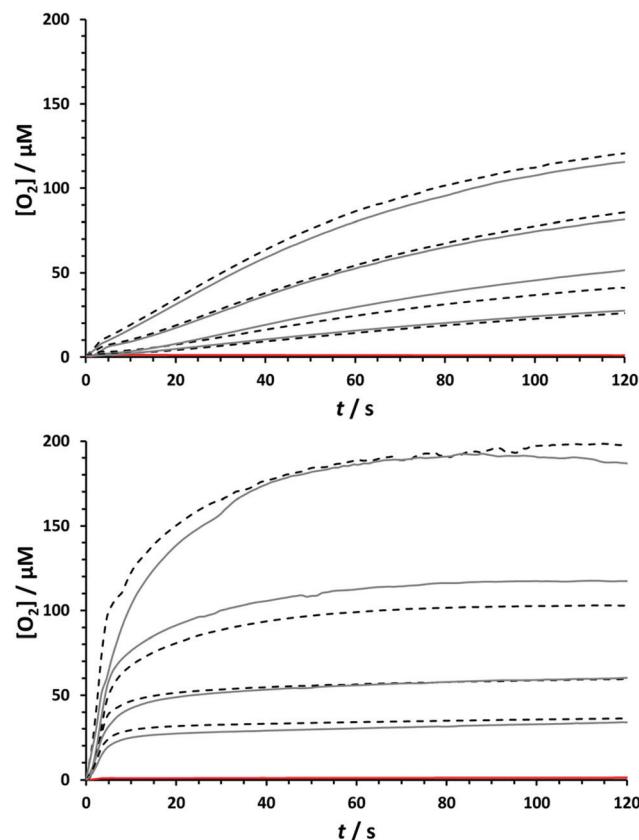


Fig. 3 Oxygen evolution profiles after addition of **8a** (top) and **8b** (bottom) to CAN measured with fluorescence probe. $[\text{CAN}] = 1.5 \text{ mM}$ in $0.1 \text{ M } \text{HOTf}_{(\text{aq})}$ ($\text{pH} = 1$). $[\mathbf{8a}] = [\mathbf{8b}] = 4.0, 2.0, 1.0, 0.5 \mu\text{M}$. $T = 25^\circ\text{C}$. Reaction volume = 4.0 mL. Experiments for each [cat.] were carried out in duplicate (1st experiment, solid grey; 2nd experiment, dashed black). The red line is the blank experiment for the addition of solvent with no catalyst to the CAN solution.

that gradually slows down over time due to catalyst deactivation, whereas for catalyst **8b** there appears to be a fast initial phase, abruptly followed by a slow second phase. Despite these differences, the TONs of both catalysts were very similar; 87 for **8a** and 80 for **8b** (see ESI† for all TON and TOF values). The absorbance decay of Ce^{IV} , which was proportional to oxygen production, was measured under the same reaction conditions and used to indirectly estimate the reaction rates (see ESI†). Linear fitting of the data from 4–20 s gave almost identical first order rate constants of 2.51 s^{-1} for **8a** and 2.27 s^{-1} for **8b**, corresponding to TOFs⁴⁰ of 0.63 s^{-1} and 0.57 s^{-1} respectively. This data fits well with the previous studies of Sun *et al.*, where they compared the Ru-bda catalyst framework to the more rigid phenanthroline analogue, Ru-pda, and found that the increased rigidity shifted the preference from the I2M mechanism for the Ru-bda's to WNA for the Ru-pda's.²⁴ Considering the reaction is first order in Ru-biqa catalyst and the TOFs are closer to the $(\text{pic})_2\text{Ru}(\text{pda})$ analogue than the $(\text{pic})_2\text{Ru}(\text{bda})$ analogue (0.1 s^{-1} and 3.0 s^{-1} respectively) measured under the same conditions, a similar shift towards WNA was indicated here.^{15,24} The rates recorded for

catalyst **8b** over the initial period (0–4 s) after catalysts injection were visibly faster than the data recorded after 4 s but this range is not as accurate because it includes the errors for manual injection of the catalyst solution and so no reliable correlations to rate order could be extracted from this data. It is possible that the energy barriers for the two possible pathways (I2M vs. WNA) are perhaps so close for the Ru-biqa compounds that both pathways are energetically feasible and that compound **8b** begins catalysis *via* I2M but quickly switches to WNA after only a few cycles. Density functional theory (DFT) calculations were performed on Ru-biqa catalyst models to determine the energy barriers of the two pathways and the results are presented in the following computational section.

Another possible explanation is that the initial complex is converted to another active species during the initial phase of catalysis, thus resulting in the more complex behaviour. HPLC analysis of the catalysis reaction solutions for both compounds **8a** and **8b** supported this hypothesis to some extent (see ESI†): for catalysis solutions of **8b** (16 μM) and CAN (1.5 mM, *ca.* 94 eq.), both the starting complex and free biqa ligand were identified in the reaction mixture after catalysis. For catalysis solutions of **8b** (8 μM) and CAN (1.5 mM, *ca.* 188 eq.), both the starting complex and free biqa ligand were again present in the reaction mixture but the relative amount of free ligand had increased. No starting complex was observed at the end of catalysis reactions with **8a** and CAN, free biqa ligand was the only species identified at both catalyst concentrations. Hence, both catalysts **8a** and **8b** are not stable under the reaction conditions and the majority of the original catalyst has been converted after only 23–47 cycles. What is not entirely clear, is what the other degradation products are, however, the fact that Ce^{IV} consumption does not continue after the original catalysts are gone would indicate that their activity is negligible in comparison to the original molecular species. Work to determine the fate of the axial ligands and the ruthenium and to identify the exact nature of other reaction by-products is still ongoing in our labs.

Computational studies

Assuming that the molecular catalysts are responsible for catalysis in the initial phases of the reaction with CAN and that the O–O bond formation is the rate limiting step, as was previously observed for the analogous Ru-bda compounds,^{20,41} the activation barriers of the two well-known I2M and WNA mechanisms^{24,42} were compared for model complexes of the new Ru-biqa compounds **8a** and **8b** in an exhaustive DFT study (see ESI† for more details). For both model complexes, Ru-biqa-py and Ru-biqa-isoq (py = pyridine, isoq = isoquinoline as axial ligands), the barriers for the WNA pathway were found to be lower than those of the I2M pathway (see Table 2 and Fig. 4). The flexibility of the bda ligand allows it to adapt to the various transition state and intermediate structures of the I2M pathway.²⁴ Both the pda and biqa ligands are restrained in their movement and are therefore less adaptable, thus increasing the I2M barriers beyond those of WNA. The axial ligand identity did not appear to play an important role in determi-



Table 2 Activation energies for the I2M and WNA mechanisms

| Ligands | $E_a(\text{I2M-HS})$ (kcal mol ⁻¹) | $E_a(\text{WNA-H}_2\text{O})$ (kcal mol ⁻¹) |
|--------------|--|---|
| Ru-bda-py | 22.5 ^a (15.0 ^b) | 26.9 ^a (18.0 ^b /32.7 ^c) |
| Ru-biqa-py | 30.7 ^a | 23.9 ^a |
| Ru-biqa-isoq | 25.1 ^a | 23.6 ^a |

^a Present work. ^b Ref. 24. ^c Ref. 43.

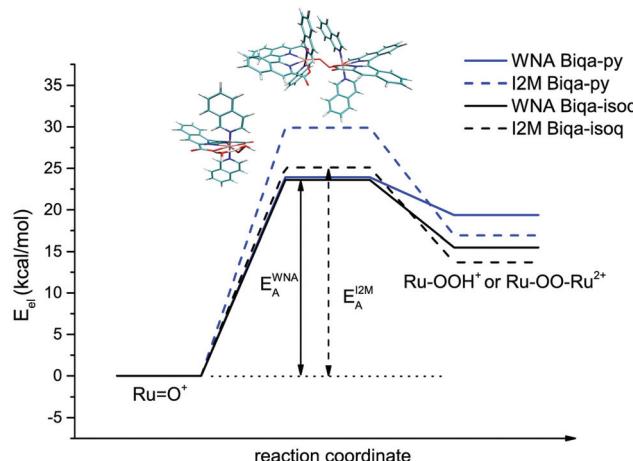


Fig. 4 Energy diagram comparing the energy barriers of oxygen bond formation for the WNA and I2M mechanisms for Ru-biqa model complexes.

ning the transition state geometries or the activation energies for the WNA mechanism (see Table 2, Fig. 5 and S33†). Transition state (TS) structures calculated using either one or two explicit water molecules highlighted the importance of an accurate description of the solvation shell. In particular, the hydrogen bonding network which shuttles the proton off the attacking water molecule to the carboxylate of the biqa ligand plays a crucial role. Overall, our findings are in good agreement with previous studies by Tong *et al.*, where the necessity of an additional water molecule was discussed in detail.²⁴

For the I2M mechanism, we observed a significant contribution of π - π stacking that leads to a lower barrier for Ru-biqa-isoq compared to Ru-biqa-py. This π -stacking effect is in line with observations made by Sun and co-workers for Ru-bda complexes with pyridine and isoquinoline axial ligands.²⁰ However, this effect would only become relevant in situations where the I2M was favoured over WNA.

The barriers obtained from the DFT calculations help to explain the observations made during the catalysis experiments, which showed regular oxygen evolution curves and first order kinetics for the Ru-biqa-pic complex (**8a**) but a more complex behaviour for the Ru-biqa-Br-isoq complex (**8b**): while the activation energy difference between the two mechanisms for Ru-biqa-py is 6.8 kcal mol⁻¹, for Ru-biqa-isoq it is only 1.5 kcal mol⁻¹. This clear preference for the WNA mechanism for Ru-biqa-py agrees well with the experimentally observed

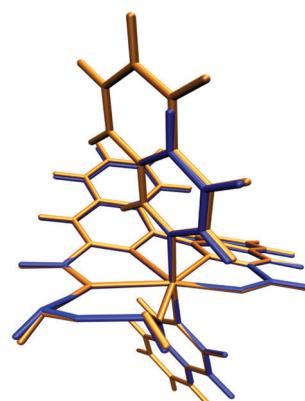
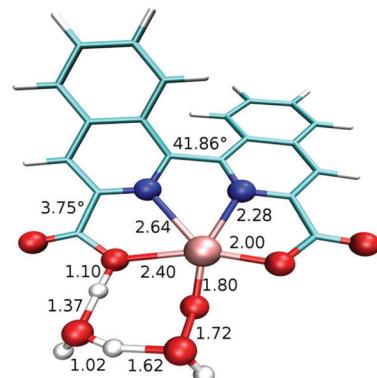


Fig. 5 Top: WNA transition state for Ru-biqa-py calculated with two explicit water molecules. Selected bond lengths (in Å) and dihedral angles are also given. Axial ligands are not shown for clarity. Bottom: Overlay of TS structure for Ru-biqa-py (blue) and Ru-biqa-isoq (orange) highlighting the fact that the axial ligand does not influence the WNA-TS structure.

first order kinetics (in catalyst) for compound **8a**, whereas for the Ru-biqa-isoq compound the energy difference of the two pathways is negligible, which indicates that both pathways are equally energetically feasible. Thus, for compound **8b** both mechanisms could potentially operate, resulting in a more complex kinetic behaviour as the mechanism switches or the compound begins to degrade. It is also worth pointing out that other mechanisms for O–O bond formation with Ru-bda complexes have recently been proposed,⁴⁴ which could offer an alternative theoretical explanation as well and would be worth considering in future studies.

Conclusions

In summary, the novel biqaH₂ ligand framework was successfully prepared and complexed with ruthenium to give a new family of molecular water oxidation catalysts, the Ru-biqa's. The ruthenium complexes were shown to be active for chemically induced water oxidation catalysis when added to CAN oxidant at pH 1, however, their activity was greatly inferior to



the analogous Ru-bda compounds tested under similar experimental conditions.^{15,20,22,35} DFT studies revealed that for Ru-biq₄ complexes, in contrast to Ru-bda analogues, the oxygen–oxygen bond formation is more favourable *via* the WNA pathway than the I2M pathway. The slow catalytic rates and relatively lower barriers for WNA pathway were also in better agreement with previous studies of ruthenium complexes containing the rigid pda ligand framework, which were proposed to catalyse water oxidation *via* the WNA pathway.²⁴ It therefore appears that inducing non-planarity in the equatorial coordinating ligand framework has the same negative effect as inducing rigidity; resulting in a switch from the I2M mechanism to the WNA mechanism and consequently slower catalysis. Furthermore, the stability (indicated by TON) of the Ru-biq₄ complexes was lower than both the Ru-bda's and the Ru-pda's. Replacement of the bda ligand by biq₄ therefore offers no improvement with regards to catalytic performance, however, these findings help to further consolidate our understanding of the mechanisms surrounding catalytic water oxidation and will be invaluable in the design of future molecular catalysts. For example, future work that may look to investigate substitution effects on the bda ligand should avoid substituents positioned at the 3 and 3' positions as they may also lead to similarly induced non-planarity and a consequent drop in catalytic performance.

Acknowledgements

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References

- 1 N. S. Lewis, *Science*, 2016, **351**(6271), DOI: 10.1126/science. aad1920.
- 2 S. H. Jensen, C. Graves, M. Mogensen, C. Wendel, R. Braun, G. Hughes, Z. Gao and S. A. Barnett, *Energy Environ. Sci.*, 2015, **8**, 2471–2479.
- 3 Q. Yi, W. Li, J. Feng and K. Xie, *Chem. Soc. Rev.*, 2015, **44**, 5409–5445.
- 4 M. Z. Jacobson, M. A. Delucchi, G. Bazouin, Z. A. F. Bauer, C. C. Heavey, E. Fisher, S. B. Morris, D. J. Y. Piekutowski, T. A. Vencilla and T. W. Yeskooa, *Energy Environ. Sci.*, 2015, **8**, 2093–2117.
- 5 N. Armaroli and V. Balzani, *Chem. – Eur. J.*, 2015, **22**, 32–57.
- 6 D. Kim, K. K. Sakimoto, D. Hong and P. Yang, *Angew. Chem., Int. Ed.*, 2015, **54**, 3259–3266.
- 7 J. A. Herron, J. Kim, A. A. Upadhye, G. W. Huber and C. T. Maravelias, *Energy Environ. Sci.*, 2015, **8**, 126–157.
- 8 S. A. Bonke, M. Wiechen, D. R. MacFarlane and L. Spiccia, *Energy Environ. Sci.*, 2015, **8**, 2791–2796.
- 9 S. Berardi, S. Drouet, L. Francàs, C. Gimbert-Suriñach, M. Guttentag, C. Richmond, T. Stoll and A. Llobet, *Chem. Soc. Rev.*, 2014, **43**, 7501–7519.
- 10 L. Francàs, R. Bofill, J. García-Antón, L. Escriche, X. Sala and A. Llobet, *Molecular Water Oxidation Catalysis: A Key Topic for New Sustainable Energy Conversion Schemes*, John Wiley & Sons Ltd, Chichester, UK, 2014.
- 11 J. D. Blakemore, R. H. Crabtree and G. W. Brudvig, *Chem. Rev.*, 2015, **115**, 12974–13005.
- 12 R. Cao, W. Lai and P. Du, *Energy Environ. Sci.*, 2012, **5**, 8134–8157.
- 13 H. Inoue, T. Shimada, Y. Kou, Y. Nabetani, D. Masui, S. Takagi and H. Tachibana, *ChemSusChem*, 2011, **4**, 173–179.
- 14 L. Duan, L. Wang, F. Li, F. Li and L. Sun, *Acc. Chem. Res.*, 2015, **48**, 2084–2096.
- 15 L. Duan, A. Fischer, Y. Xu and L. Sun, *J. Am. Chem. Soc.*, 2009, **131**, 10397–10399.
- 16 L. Duan, L. Wang, A. K. Inge, A. Fischer, X. Zou and L. Sun, *Inorg. Chem.*, 2013, **52**, 7844–7852.
- 17 L. Wang, L. Duan, B. Stewart, M. Pu, J. Liu, T. Privalov and L. Sun, *J. Am. Chem. Soc.*, 2012, **134**, 18868–18880.
- 18 L. Duan, Y. Xu, P. Zhang, M. Wang and Li. Sun, *Inorg. Chem.*, 2010, **49**, 209–215.
- 19 Y. Jiang, F. Li, B. Zhang, X. Li, X. Wang, F. Huang and L. Sun, *Angew. Chem., Int. Ed.*, 2013, **52**, 3398–3401.
- 20 L. Duan, F. Bozoglian, S. Mandal, B. Stewart, T. Privalov, A. Llobet and L. Sun, *Nat. Chem.*, 2012, **4**, 418–423.
- 21 L. Duan, C. M. Araujo, M. S. G. Ahlquist and L. Sun, *Proc. Natl. Acad. Sci. U. S. A.*, 2012, **109**, 15584–15588.
- 22 L. Wang, L. Duan, Y. Wang, M. S. G. Ahlquist and L. Sun, *Chem. Commun.*, 2014, **50**, 12947–12950.
- 23 C. J. Richmond, R. Matheu, A. Poater, L. Falivene, J. Benet-Buchholz, X. Sala, L. Cavallo and A. Llobet, *Chem. – Eur. J.*, 2014, **20**, 17282–17286.
- 24 L. Tong, L. Duan, Y. Xu, T. Privalov and L. Sun, *Angew. Chem., Int. Ed.*, 2011, **50**, 445–449.
- 25 T.-T. Li, W.-L. Zhao, Y. Chen, F.-M. Li, C.-J. Wang, Y.-H. Tian and W.-F. Fu, *Chem. – Eur. J.*, 2014, **20**, 13957–13964.
- 26 Z. Liu, Y. Gao, Z. Yu, M. Zhang and J. Liu, *Chin. J. Cat.*, 2015, **36**, 1742–1749.
- 27 R. Matheu, M. Z. Ertem, J. Benet-Buchholz, E. Coronado, V. S. Batista, X. Sala and A. Llobet, *J. Am. Chem. Soc.*, 2015, **137**, 10786–10795.
- 28 D. W. Shaffer, Y. Xie, D. J. Szalda and J. J. Concepcion, *Inorg. Chem.*, 2016, DOI: 10.1021/acs.inorgchem.6b02193, asap.
- 29 E. F. Cotter, *Patent WO 2010/125343A1*, 2010.
- 30 E. T. Stiller, *J. Chem. Soc.*, 1937, 473–476.
- 31 Y. L. Janin, E. Roulland, A. Beurdeley-Thomas, D. Decaudin, C. Monneret and M.-F. Poupon, *J. Chem. Soc., Perkin Trans. 1*, 2002, 529–532.



32 M. Hapke, H. Staats, I. Wallmann and A. Lützen, *Synthesis*, 2007, 2711–2719.

33 M. Slangy and P. Stang, *Synthesis*, 1996, 1019–1028.

34 G. Bozoklu, C. Marchal, C. Gateau, J. Pécaut, D. Imbert and M. Mazzanti, *Chem. – Eur. J.*, 2010, **16**, 6159–6163.

35 C. J. Richmond, R. Matheu, A. Poater, L. Falivene, J. Benet-Buchholz, X. Sala, L. Cavallo and A. Llobet, *Chem. – Eur. J.*, 2014, **20**, 17282–17286.

36 Values recorded *vs.* Fc/Fc^+ internal standard and converted to *vs.* NHE by addition of 630 mV.

37 M. H. V. Huynh and T. J. Meyer, *Chem. Rev.*, 2007, **107**, 5004–5064.

38 C. J. Gagliardi, B. C. Westlake, C. A. Kent, J. J. Paul, J. M. Papanikolas and T. J. Meyer, *Coord. Chem. Rev.*, 2010, **254**, 2459–2471.

39 Z. Chen, J. J. Concepcion, H. Luo, J. F. Hull, A. Paul and T. J. Meyer, *J. Am. Chem. Soc.*, 2010, **132**, 17670–17673.

40 TOF values calculated by dividing the rate of oxygen production ($\mu\text{M s}^{-1}$) by the ruthenium catalyst concentration (μM).

41 R. Staehle, L. Tong, L. Wang, L. Duan, A. Fischer, M. S. G. Ahlquist, L. Sun and S. Rau, *Inorg. Chem.*, 2014, **53**, 1307–1319.

42 S. Romain, L. Vigara and A. Llobet, *Acc. Chem. Res.*, 2009, **42**, 1944–1953.

43 R. Kang, K. Chen, J. Yao, S. Shaik and H. Chen, *Inorg. Chem.*, 2014, **53**, 7130–7136.

44 J. J. Concepcion, D. K. Zhong, D. J. Szalda, J. T. Muckerman and E. Fujita, *Chem. Commun.*, 2015, **51**, 4105–4108.