

Cite this: *Dalton Trans.*, 2016, **45**, 13525Received 5th June 2016,
Accepted 26th July 2016
DOI: 10.1039/c6dt02242a

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Palladium nanoparticles supported on a nickel pyrazolate metal organic framework as a catalyst for Suzuki and carbonylative Suzuki couplings†

A. W. Augustyniak,^a W. Zawartka,^a J. A. R. Navarro^{*b} and A. M. Trzeciak^{*a}Methanolic reduction of [PdCl₂(CH₃CN)₂] on a [Ni(2,5-di(1*H*-pyrazol-4-yl)benzenesulfonate)₂] metal organic framework gives rise to Pd²⁺/Pd⁰ nanocomposites with Suzuki and carbonylative Suzuki heterogeneous catalytic activities.

One of the major challenges in modern catalysis is the efficient separation of a metal catalyst from organic products.^{1–7} This is particularly important in palladium catalysed cross-coupling reactions that lead to products which can be utilized in the pharmaceutical industry. For example, the Suzuki–Miyaura reaction presents an attractive pathway to substituted biaryls, precursors of anti-inflammatory drugs.^{8–12} The carbonylative Suzuki–Miyaura coupling offers easy access to diarylated ketones used in UV-screens, dyes, agrochemicals and pharmaceuticals.^{13–15}

Application of heterogeneous catalysts to these reactions enables improvement of the separation step and minimizes the palladium content in the reaction products. In this regard, metal organic frameworks, MOFs, have attracted an increasing attention as potential supports for palladium.^{16–22} Advantageously, MOFs possess large surface areas with well-defined pores of shapes that can be modulated. These highly porous materials can be easily modified by several functional groups facilitating the bonding of a variety of metal ions and metal nanoparticles (NPs).^{16–22}

One of the first reports dealing with the application of the MOF in the Suzuki–Miyaura reaction was the palladium(II) based system [Pd(2-pymo)₂]_n (2-pymo = 2-hydroxypyrimidinolate)^{23,24} which remained unaltered under the reaction conditions, at 150 °C, as proven by X-ray powder diffraction.²⁴ Other MOFs, mainly carboxylate systems of the MIL family, were used as supports for palladium catalysts of the Suzuki–Miyaura coupling.^{25–36} Thus, Pd NPs immobilized on MIL-101(Cr) formed efficient heterogeneous catalysts of the

water-mediated coupling of aryl chlorides at 80 °C with a TBAB additive.²⁵ Palladium supported on amino-functionalized MIL-53(Al)NH₂ catalyzed the Suzuki–Miyaura coupling with high activity and recyclability.²⁶ The effect of palladium loading on the catalytic performance was studied for Pd NPs immobilized on MIL-101(Cr)NH₂ and the best activity and recyclability were obtained for 8 wt% of Pd.²⁷ Further studies with this catalyst indicated that the type of the base used in the catalytic reaction has a decisive influence on the stability of the MOF structure.²⁸ Noteworthy, the MOF catalyst was degraded by the presence of carbonates, while it remained crystalline with fluorides.²⁸ Decomposition of Pd-MIL-101(Cr) and leaching of the active phase were observed.²⁹ A similar catalyst, Pd-MIL-101, containing well dispersed Pd NPs obtained by a “double solvent” method acted as the heterogeneous and recyclable catalyst of bromo- and iodobenzene coupling.³⁰

Pd-NH₂-MIL-125 catalyzed the Suzuki–Miyaura coupling of aryl chlorides at 100 °C in methanol with a yield of ca. 80%.³¹ It was also successfully recycled and used in four subsequent runs. The catalyst Pd-Uio-66, obtained by the microwave-assisted method, exhibited high catalytic activity in the Suzuki–Miyaura coupling of aryl bromides and chlorides at 30 °C.³² Pd NPs dispersed on the surface of DUT-67 micro-crystals were used for coupling of iodobenzene at 70 °C in water-ethanol.³³ Interestingly, the Pd(II) complex with a 2,2'-bipyridine-5,5'-dicarboxylate linker incorporated into the MOF (Uio-67) showed a remarkably higher activity than the homogeneous Pd counterparts.³⁴

Noteworthy, while the Pd–MOF catalysts found application in the Suzuki–Miyaura coupling, this work represents the first example of their use under a CO atmosphere. To the best of our knowledge, the only example until now was reported for the Sonogashira carbonylative coupling with a Pd(II)–Zr–MOF–BIPY system.³⁷

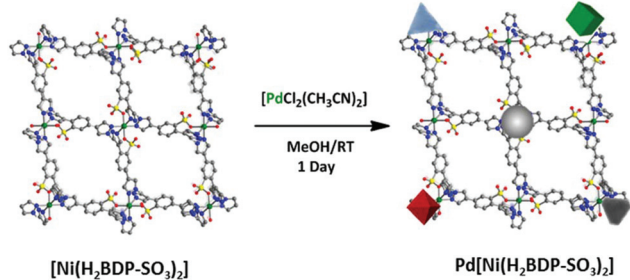
Most of the studies on the Pd–MOF systems in the Suzuki–Miyaura reaction are based on carboxylate MOFs of the MIL/

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† Electronic supplementary information (ESI) available: Experimental details, optimization of the reaction conditions, FT-IR spectra, XPS spectra, TGA curve, TEM image of Pd@[Ni(H₂BDP-SO₃)₂] after 4 uses. See DOI: 10.1039/c6dt02242a





Scheme 1 Green preparation of Pd@[Ni(H₂BDP-SO₃)₂] catalyst.

UiO families. In this regard, testing of other MOF materials of high chemical stability as supports for palladium could be beneficial. To this aim, MOFs containing strong donor azolate ligands like imidazoles or pyrazoles could be considered.^{38,39} Following this idea, we decided to choose pyrazolato based MOF containing sulfonate functionalities, [Ni(H₂BDP-SO₃)₂], for palladium immobilization.⁴⁰ We expected that the system, built from the assembly of Ni(II) cations and 2,5-di(1*H*-pyrazol-4-yl)benzenesulfonate linkers might be a suitable MOF support for Pd²⁺ and Pd NPs since both palladium forms are present in the Suzuki–Miyaura catalytic cycle (Scheme 1).^{4,6,7}

Noteworthy, the impregnation of the [Ni(H₂BDP-SO₃)₂] system with a methanolic solution of [PdCl₂(CH₃CN)₂] at room temperature gave rise to the straightforward formation of Pd@[Ni(H₂BDP-SO₃)₂] containing both Pd²⁺ ions and Pd NPs embedded in the MOF material. The uploaded amount of Pd, as determined by induced coupling plasma (ICP) analysis, is up to 5.6 wt%. TEM images indicated that the Pd NPs are homogeneously distributed on the MOF microparticles (Fig. 1) (for details see the ESI†). Remarkably, MeOH acted as both a solvent and a reductant.⁴¹ Indeed, we have observed a quick change of the colour of the [PdCl₂(CH₃CN)₂] solution from yellow to dark brown, while the MOF micropowder changed the colour from yellow to dark red, indicating that the palladium precursor was reduced to Pd NPs. The as-prepared Pd NPs were characterized by transmission electron microscopy (TEM) (see Fig. 1C and D) and elemental mapping images (Fig. 1E). The corresponding electron diffraction pattern (SAED) for the selected area is shown in Fig. 1B demonstrating that Pd NPs are crystalline. In the SAED pattern we observe the diffraction rings matched with (111), (200), (220), and (311) crystallographic planes of Pd nanocrystals, respectively. The TEM images also established that the palladium nanocrystals are in the 4–8 nm size distribution (Fig. 1A). It should also be noted that Pd NPs exhibit very-well defined geometrical shapes even though the synthesis conditions were very mild. This is unusual since the production of Pd NPs typically involves the addition of a reducing agent (*i.e.* N₂H₄, NaBH₄, H₂) leading to less defined shapes in most cases.^{42,43} In this regard, the use of methanol as a reducing agent, under the reported mild conditions, seems to be advantageous since it leads to a higher control over the shape and size of palladium nanocrystals com-

pared to traditional reducing agents (Fig. 1D). Moreover, analysis of X-ray photoelectron spectroscopy (XPS) spectra confirmed the presence of both Pd(0) (334.0, 339.3 eV) and Pd²⁺ (336.9, 342.6 eV) in a 15 : 85 ratio. On the other hand, XPS lines of Pd²⁺ are shifted towards lower binding energies, as compared to pristine [PdCl₂(CH₃CN)₂] indicating a change of the palladium environment. In fact, FT-IR analysis exhibited that the characteristic absorbance band for a C≡N bond at *ca.* 2340 cm⁻¹ was absent (Fig. S4†), suggesting that CH₃CN is removed from the coordination sphere of palladium during immobilization. The FT-IR region 600–100 cm⁻¹ was also studied. There is a strong absorption band centred at 352 cm⁻¹ which can unequivocally be assigned to ν(Pd–Cl).⁴⁴ Additionally, EDX analysis shows a *ca.* 1:1 atomic ratio of Pd:Cl. These results suggest that PdCl⁺ cations might be anchored to the outer surface of the Pd NPs and further stabilized in the [Ni(H₂BDP-SO₃)₂] structure by interactions with sulfonate oxygen atoms of the organic linkers. Indeed, two additional bands centred at 420 and 455 cm⁻¹ might be assigned to ν(Pd–O), supporting the plausible interaction of the core shell NPs with the sulfonate groups of the organic spacers. Alternatively, additional stabilizing interactions between Pd–Cl cations and NH groups of pyrazole could also be considered.

The low-angle XPRD pattern of the as-prepared [Ni(H₂BDP-SO₃)₂] (Fig. 2A) is in agreement with the simulated pattern reported by Colombo *et al.*⁴⁰ In addition, the structure of [Ni(H₂BDP-SO₃)₂] remains intact after palladium uploading and no changes in the XPRD patterns were detected, which supports that [Ni(H₂BDP-SO₃)₂] was chemically stable during the palladium reduction process and ulterior catalytic reaction. We have also evaluated the effect of palladium incorporation into the [Ni(H₂BDP-SO₃)₂] system by means of N₂ adsorption isotherms at 77 K (see Fig. 2B and Table 1). Noteworthy, the porous structure in Pd@[Ni(H₂BDP-SO₃)₂] is still highly accessible, although a diminution of approx. 45% in the specific surface area (SA) and pore volume is noticed. On the other hand, the TGA analysis reveals that the [Ni(H₂BDP-SO₃)₂] material possesses a notable thermal stability of up to 350 °C (Fig. S1†).

Once the physicochemical features of the Pd@[Ni(H₂BDP-SO₃)₂] composites were established we proceed to evaluate their plausible catalytic properties.

The C–C coupling reactions between aryl halides (–Br, –Cl) and phenylboronic acid were selected as model reactions for testing the catalytic activity of Pd@[Ni(H₂BDP-SO₃)₂]. With this aim, the reaction parameters were optimized using 4-bromoanisole as a screening substrate, and the results are summarized in Tables 2 and S1 (see the ESI†).

The reaction did not proceed either in the absence of a catalyst (Table S1, entry 1†) or in the presence of pristine [Ni(H₂BDP-SO₃)₂] (Table S1, entry 2†), which confirms the catalytic activity of Pd@[Ni(H₂BDP-SO₃)₂]. Indeed, Pd@[Ni(H₂BDP-SO₃)₂] was active, in 1:1 2-propanol and water solvent mixtures at 25 °C forming up to 91% of the coupling product. Under these conditions (25 °C) reactions of different aryl halides with phenylboronic acid showed conversions from



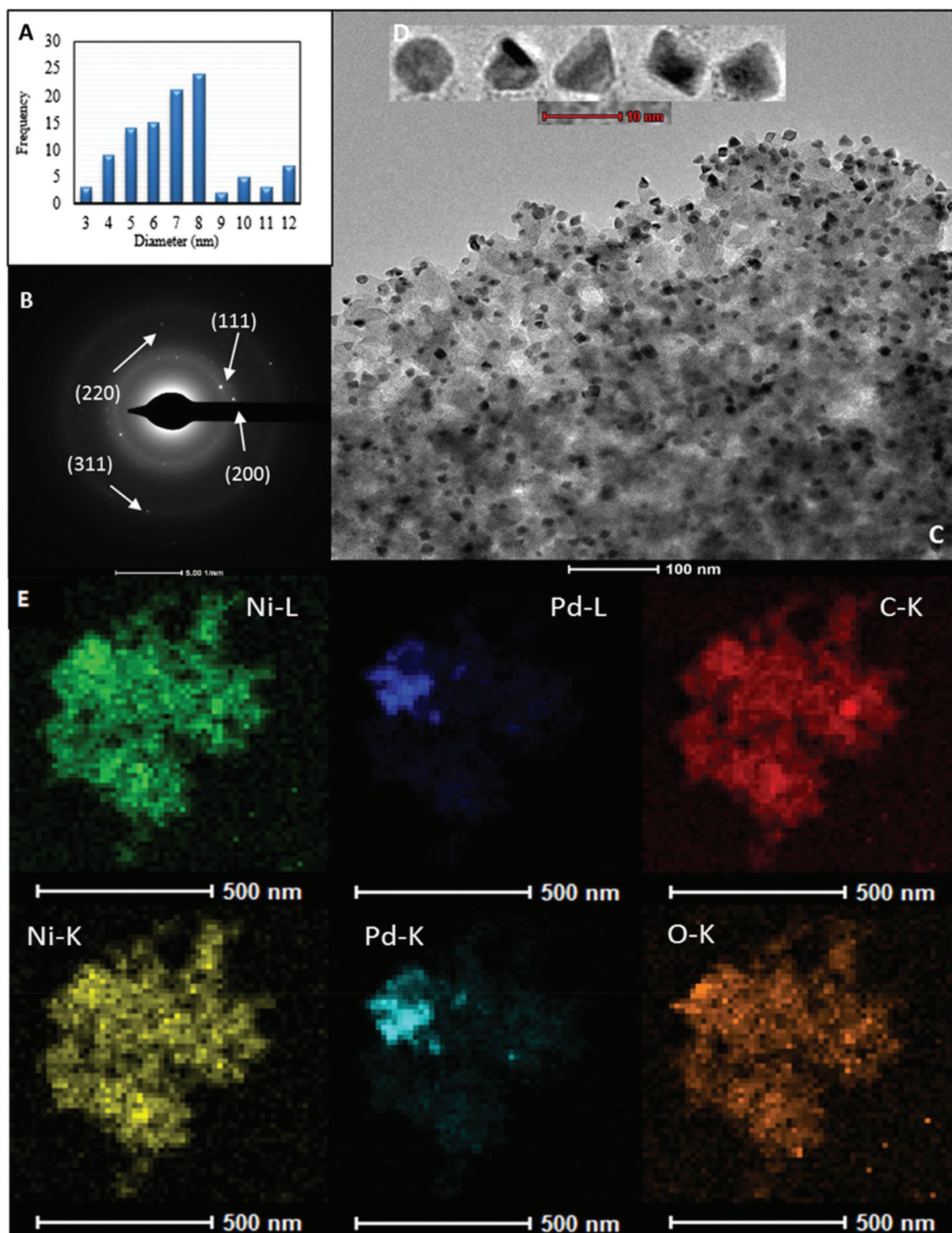


Fig. 1 (A) Histogram of Pd particle size, (B) SAED pattern of Pd nanocrystals, (C, D) TEM images of Pd NPs with different shapes found in Pd@[Ni(H₂BDP-SO₃)₂] nanocrystals, (E) EDX elemental mappings of Pd@[Ni(H₂BDP-SO₃)₂].

30% for 2-chlorotoluene to 91% for 4-bromoanisole and 4-bromotoluene. The highest conversions, up to 97–99%, were obtained for 2-bromotoluene, 4-bromotoluene and 4-bromoanisole at 60 °C after a 4 h reaction (Table 2). Kinetic profiles

of Suzuki–Miyaura reactions of 4-bromoanisole, illustrating a positive effect of a higher temperature on the reaction course, are shown in Fig. 3. By contrast, in the reaction of 2-chlorotoluene at 100 °C only 24% of the corresponding coupling



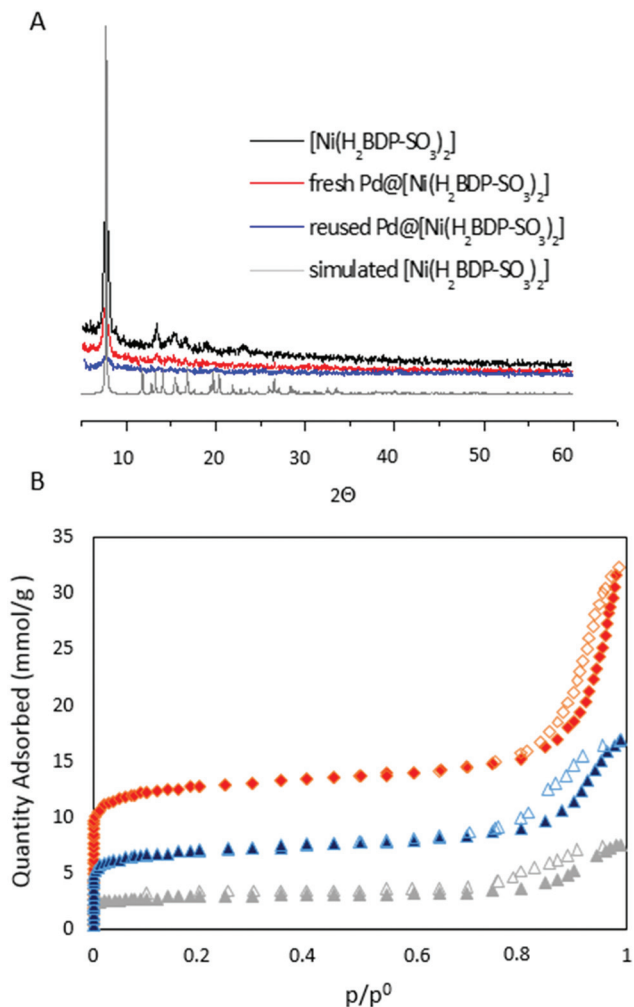


Fig. 2 (A) XPRD patterns of $[\text{Ni}(\text{H}_2\text{BDP-SO}_3)_2]$ as synthesized (black), $\text{Pd}@[\text{Ni}(\text{H}_2\text{BDP-SO}_3)_2]$ (red), 4-use catalytic cycle (blue), simulated (grey);⁴⁰ (B) N_2 adsorption isotherms measured at 77 K for $[\text{Ni}(\text{H}_2\text{BDP-SO}_3)_2]$ (red), $\text{Pd}@[\text{Ni}(\text{H}_2\text{BDP-SO}_3)_2]$ (blue) and reused $\text{Pd}@[\text{Ni}(\text{H}_2\text{BDP-SO}_3)_2]$ (grey). Filled and empty symbols represent adsorption and desorption, respectively.

Table 1 N_2 sorption isotherm data of $[\text{Ni}(\text{H}_2\text{BDP-SO}_3)_2]$ and $\text{Pd}@[\text{Ni}(\text{H}_2\text{BDP-SO}_3)_2]$

Sample ^a	S_{BET} ($\text{m}^2 \text{g}^{-1}$)	V_{pore} ($\text{cm}^3 \text{g}^{-1}$)
$[\text{Ni}(\text{H}_2\text{BDP-SO}_3)_2]$	1082(2)	1.11
$\text{Pd}@[\text{Ni}(\text{H}_2\text{BDP-SO}_3)_2]$	583(2)	0.58
$\text{Pd}@[\text{Ni}(\text{H}_2\text{BDP-SO}_3)_2]^b$	246(1)	0.26

^a Desolvated under reduced pressure at 180 °C for 12 h. ^b Reused $\text{Pd}@[\text{Ni}(\text{H}_2\text{BDP-SO}_3)_2]$.

product was formed as a result of higher strength of the C-Cl bond compared to C-Br (Table 2). The recyclability of the $\text{Pd}@[\text{Ni}(\text{H}_2\text{BDP-SO}_3)_2]$ catalyst was tested over the Suzuki-Miyaura reaction (Fig. 4). After each run, the solid catalyst was easily collected by centrifugation, washed with MeOH, and used for a further run up to four consecutive runs under the same conditions.

Table 2 $\text{Pd}@[\text{Ni}(\text{H}_2\text{BDP-SO}_3)_2]$ catalyzed Suzuki-Miyaura coupling showing the effect of substituents in aryl halides^a

Entry	Aryl halide	Yield ^a (%) [*]	Yield ^a (%) ^{**}
1		85	99
2		58	75
3		88	94
4		69	76
5		91	98
6		91	97
7		—	24 ^{***}

^a Reaction conditions: aryl halide (1 mmol), phenylboronic acid (1.5 mmol), K_2CO_3 (1.5 mmol), $i\text{PrOH}$ (2.5 mL), H_2O (2.5 mL), $\text{Pd}@[\text{Ni}(\text{H}_2\text{BDP-SO}_3)_2]$ (0.5 mol% of Pd), * 8 h at 25 °C, ** 4 h at 60 °C, *** 20 h at 100 °C.

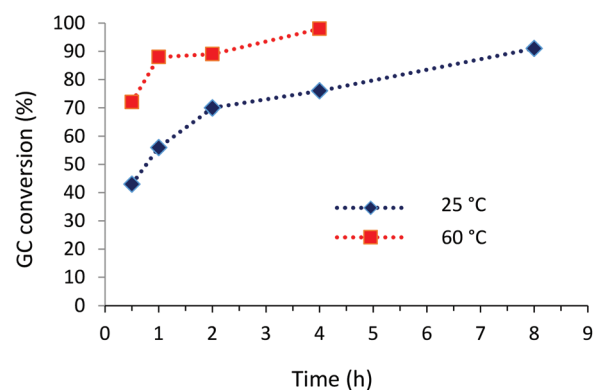


Fig. 3 Time conversion plot for the Suzuki-Miyaura coupling of 4-bromoanisole with phenylboronic acid.



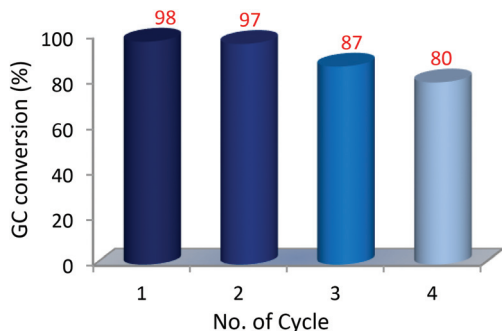


Fig. 4 Recycling catalytic test of Pd@[Ni(H₂BDP-SO₃)₂]. Reaction conditions: 4-bromoanisole (1 mmol), phenylboronic acid (1.2 mmol), K₂CO₃ (1.5 mmol), Pd@[Ni(H₂BDP-SO₃)₂] (0.5 mol% of Pd) H₂O/PrOH (5 mL), 60 °C, 4 h.

The results show a relatively small conversion decrease, from 98 to 80%, highlighting the good stability of the Pd@[Ni(H₂BDP-SO₃)₂] catalyst. After the 4th use, the catalyst was examined by XPRD (Fig. 2A) and TEM (Fig. S3†) and showed a certain degree of MOF crystallinity loss. Noteworthily, the amount of Pd(0) increased from the original 15% to 39% similarly to other heterogenized palladium systems in Suzuki–Miyaura reactions^{45–47} (Fig. S2†). To estimate the role of soluble palladium forms leached from the solid support, in the catalytic process, the activity of the liquid phase was investigated. However, the filtrate obtained after separation of the Pd@[Ni(H₂BDP-SO₃)₂] catalyst showed no activity (Table S1, entry 7†). According to ICP analysis, the palladium content in the liquid phase was 0.035 ppm evidencing insignificant palladium leaching. Consequently, the heterogeneous reaction pathway can be proposed as the main one, with catalytic activity taking place on the Pd NP surface.

Furthermore, TEM images of the recovered catalyst indicate that the Pd NPs are homogeneously dispersed even after the catalytic reactions (Fig. 5A and B).

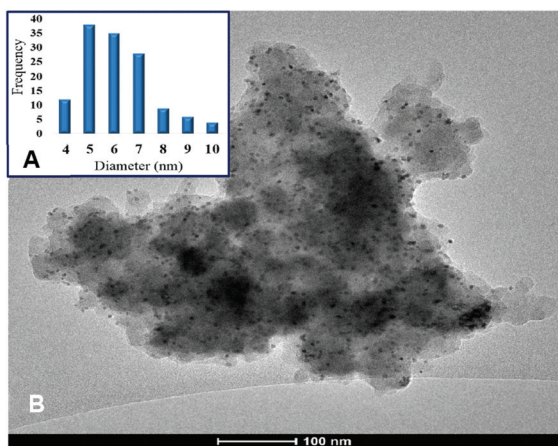


Fig. 5 (A) Histogram of Pd NPs; (B) TEM images of Pd@[Ni(H₂BDP-SO₃)₂] catalyst after reaction.

Table 3 Pd@[Ni(H₂BDP-SO₃)₂] catalyzed carbonylative Suzuki–Miyaura coupling of aryl iodides with arylboronic acid^a

Entry	Aryl halides	Boronic acid	Conv. (%)	A (%)	B (%)
1			80 ^{*,b}	21	59
2			66 ^{***,b}	51	15
3			50 ^{***,b}	50	0
4			58	58	0
5			33	33	0
6			54	54	0
7			71	71	0
8			59	59	0
9			53	53	0
10			35	35	0

^a Reaction conditions: aryl halide (1 mmol), phenylboronic acid (1.2 mmol), Pd@[Ni(H₂BDP-SO₃)₂] (1.0 mol% of Pd), K₂CO₃ (3 mmol), solvent (anisole) 5 mL, CO (10 bar) at 105 °C. *CO balloon, **CO (5 bar), ***CO (10 bar). ^b 0.5 mol% of Pd at 105 °C.



Moreover, agglomeration of Pd NPs was not appreciable but some decrease of an average size was noted. The morphology was also changed and Pd NPs became more round. These observations support the suggestion that Pd NPs do actively participate in the catalytic reaction taking place on their surface. The Pd amount found in the reused Pd@[Ni(H₂BDP-SO₃)₂], as determined by ICP analysis, was 5.3 wt%. The specific area determined by N₂ adsorption decreased to 246 m² g⁻¹, probably as a result of incorporation of organic products into pores (Table 1). As the leaching of palladium was not significant, it can be assumed that the decrease of the catalytic activity during recycling is caused by lower availability of active palladium centres blocked by organic products.

Once we proved that the Pd@[Ni(H₂BDP-SO₃)₂] system was catalytically active in the Suzuki–Miyaura coupling, we decided to test it in a more demanding reaction, like carbonylative Suzuki–Miyaura coupling. It was interesting to check whether the catalytic activity can be switched from standard coupling to the carbonylative coupling by a simple change in the reaction conditions. The carbonylative coupling of iodoanisole with phenylboronic acid was selected as the model reaction, and the results are summarized in Tables 3 and S2 (see the ESI†). Anisole, toluene and DMF were chosen as solvents, while K₂CO₃, KHCO₃ and KOH were selected as bases. At 1 bar of CO only 21% of diarylketone was formed, however, its yield increased to 51% after pressure was increased to 5 bar. At 10 bar of CO diarylketone was obtained as the only product. Consequently, this pressure was used in further experiments performed in anisole as a solvent with 1 mol% of Pd@[Ni(H₂BDP-SO₃)₂]. Reactions of different aryl iodides with phenylboronic acid show conversions from 33% to 71% (Table 3). The highest conversion, up to 71%, was obtained for 3-iodoanisole. Typically, a correlation between the conversion and position of the substituent in the aryl ring is observed.

The much higher reactivity of 3-iodoanisole, as compared to 4-iodoanisole and 2-iodoanisole, may be attributed to the reduced pore availability of Pd@[Ni(H₂BDP-SO₃)₂] for a 3-iodoanisole molecule facilitating the reaction on the surface of Pd@[Ni(H₂BDP-SO₃)₂]. Reactions with substituted boronic acids showed conversions ranging from 59% for 3,5-dimethoxyboronic acid to 35% for 4-acetylphenylboronic acid. Interestingly the best yield was observed for bulky 3,5-dimethoxyboronic acid. In all reactions at 10 bar of CO the Pd@[Ni(H₂BDP-SO₃)₂] catalyst gave exceptionally 100% selectivity.

Conclusions

A new green synthetic method leading to Pd@[Ni(H₂BDP-SO₃)₂] containing Pd²⁺ and Pd NPs in a Ni(II) pyrazolate MOF has been presented. Noteworthy, the obtained crystalline Pd NPs supported on the [Ni(H₂BDP-SO₃)₂] MOF material can be regarded as a highly active heterogeneous catalyst for the Suzuki–Miyaura coupling in an environmentally friendly mixture of solvents (2-propanol/water). The catalyst can be recovered and reused for at least four runs without a significant loss in its activity.

At 25 °C our catalyst converted 91% of 4-bromotoluene and 4-bromoanisole to the coupling products. As far as we know, only for a Zn-free MOF-5-NPC-Pd catalyst very good results were obtained at 25 °C.³⁶ In other MOF systems a higher temperature (40–200 °C) was applied. The catalytic activity reported by us at 60 °C with 0.5 mol% of Pd is similar to that found for Pd/MIL-53-NH₂³⁰ and Pd-MCM-41.⁵¹ On the other hand, a higher activity in the Suzuki–Miyaura reaction was shown by Pd/MOF which produced 92% of the coupling product at 40 °C in 0.5 h.³⁵ Under the same conditions this catalyst converted other bromoarenes to appropriate products in 87–95%.³⁵

The activity of the Pd@[Ni(H₂BDP-SO₃)₂] catalyst can be easily switched to carbonylative coupling by introduction of CO and the excellent selectivity to diarylketone was noted at 10 bar of CO. According to the best of our knowledge we have reported for the first time the results of carbonylative coupling with a MOF supported Pd catalyst. At 1 bar of CO our catalyst produced 21% of ketone, showing lower activity than Pd nanoparticles supported on dopamine-functionalized magnetite.⁴⁹ Palladium nanoparticles immobilized onto SILLP formed diarylketones with the yield of 60–87% in 12 h at 30 atm of CO at 120 °C.⁵⁰ These conditions are significantly harsher than used by us. The Pd catalyst supported on amino-modified silica microspheres used at 5 atm of CO in an amount of 0.05 mol% gave better results, up to 90% conversion, in 5 h.⁴⁸

Further work on exploring new catalytic reactions with such a Pd@MOF system is currently underway.

Acknowledgements

We are grateful for financial support from National Science Centre (NCN, Poland) with grant 2014/15/B/ST5/02101 (A. W. A., W. Z., A. M. T.), Spanish MINECO (CTQ2014-53486-R) and EU FEDER Funding (JARN).

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