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Carving a 1D Co^{II}-carboranylcarboxylate system by using organic solvents to create stable trinuclear molecular analogues: complete structural and magnetic studies†‡

Mònica Fontanet,^a Montserrat Rodríguez,^a Xavier Fontrodona,^a Isabel Romero,^{*a} Francesc Teixidor,^{*b} Clara Viñas^b and Núria Aliaga-Alcalde^{b,c}

This work presents a straightforward methodology to achieve small linear trinuclear molecules based on the Co^{II}-carboranylcarboxylate system obtained by carving a 1D polynuclear analogous system with the use of diethylether. The reaction of the carboranylcarboxylic ligand, 1-CH₃-2-CO₂H-1,2-*closo*-C₂B₁₀H₁₀ (**LH**) with different cobalt salts leads to the polynuclear compound [Co₂(μ-H₂O)(1-CH₃-2-CO₂-1,2-*closo*-C₂B₁₀H₁₀)₄(THF)₄], **1** and the polymeric [Co(μ-H₂O)(1-CH₃-2-CO₂-1,2-*closo*-C₂B₁₀H₁₀)₂]_n(H₂O)_n, **2**. This latter 1D chain has been obtained by an unprecedented synthetic strategy for the isolation of cobalt(II) compounds. [Co₃(μ-H₂O)₂(1-CH₃-2-CO₂-1,2-*closo*-C₂B₁₀H₁₀)₆(H₂O)₂(C₄H₁₀O)₂], **3** is formed by the dissociation of the polymeric structure that forms **2** when a mild coordinating solvent such as diethylether is added. These compounds have been characterized by analytical and spectroscopic techniques. X-ray analysis of **1** and **3** revealed that **1** presents a dinuclear structure whereas **3** is trinuclear; in both cases a six-coordinated Co^{II} compound with water molecules bridging each of the two Co^{II} centres has been observed. The magnetic properties of **1** and **3** show a weak antiferromagnetic behaviour, respectively, between the Co^{II} centres mediated by two carboxylate ligands and a molecule of water.

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Introduction

One major current challenge in materials science and nanotechnology is the association of different properties in the same material which can function either in independent or concerted (synergic) ways.¹ In this respect the study of coordination compounds provides effective solutions allowing a clear understanding of their functioning and improving the fundamental and potential applied research; also, these compounds could be used as building blocks in more elaborate structures.² The presence of a metal–ligand association in metal coordination compounds can lead to the generation of suitable materials with a particular usefulness or applicability;

the metal can provide redox and magnetic properties³ among others and the appropriate ligand may determine the functionality of the material. In this sense, coordination polymers are currently of great interest and represent an active area of coordination chemistry because of their special roles in fields such as ion exchange, gas storage, chemical separation, sensor technology, magnets, optoelectronics, energy conversion and storage and catalysis.⁴ Also, polynuclear metal compounds attract great interest owing to their relevance to many important naturally occurring processes. The cooperative action of closely coupled dinuclear or multinuclear centres is required for several enzymes to carry out their biological functions as for example in the photosystem II.⁵ The successful behaviour of these structures consists of the use of adequate building blocks that allow the synthesis of 1D linear or twisted chains, 2D squares and polygons and 3D cubes and polyhedra.⁶

The magnetic properties of 1D/polymetallic compounds derive from the cooperative exchange interaction between the paramagnetic metal ions through the bridging ligands. In this aspect the ligand design is crucial to efficiently transmit exhaustive interactions between the metal in a controller manner.⁷ Concerning organic spacers, carboxylic ligands are frequent choices for metal–organic networks, one reason

^aDepartament de Química and Serveis Tècnics de Recerca, Universitat de Girona, Campus de Montilivi, E-17071 Girona, Spain. E-mail: marisa.romero@udg.edu

^bInstitut de Ciència de Materials de Barcelona, ICMAB-CSIC, Campus UAB, E-08193 Bellaterra, Spain. E-mail: teixidor@icmab.es

^cICREA (Institució Catalana de Recerca i Estudis Avançats), Passeig Lluís Companys, 23, 08018 Barcelona, Spain

†In memory of our good friend and excellent boron chemist, Igor Chizhevsky, who passed away on May 23 2016.

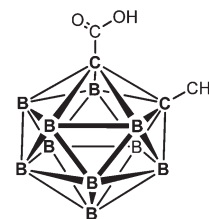
‡Electronic supplementary information (ESI) available. CCDC 1471844 and 1471845. For ESI and crystallographic data in CIF or other electronic format see DOI: 10.1039/c6dt01744d



being their rich modes of coordination.⁸ With reference to Co compounds containing bridging carboxylates and bridging O atoms (μ -aquo or μ -hydroxo), the structural details of Co core compounds, the nature of the role of coordinating water, the coordinating ligands and the hydrogen bridges are crucial to account for the magnetic properties of polynuclear Co compounds.⁹ Most of the studies on carboxylate based bi- and trinuclear cobalt metal compounds with bridging water molecules have been performed with trifluoro- or trichloroacetate bridging carboxylates¹⁰ and pivalates,¹¹ however the research on new carboxylate Co compounds with different electronic and structural parameters is necessary to develop new systems with potential applications in nanoscience.

Closo boron clusters are tridimensional aromatic compounds electronically related to the conventional Hückel $(4n + 2)\pi$ (ref. 12) planar aromatic compounds; therefore they are highly stable tridimensional structures very suitable for chemically harsh conditions.^{13,14} One of the most remarkable examples is $[B_{12}H_{12}]^{2-}$. This is a dianionic species, highly soluble in water. The substitution of one B^- by one carbon lowers the species' negative charge and hydrophilicity, e.g. $[CB_{11}H_{12}]^-$. If a further B^- substitution takes place the charge is further lowered producing a lipophilic species. Therefore the physicochemical properties of structurally very similar borane structures can be tuned according to the physicochemical needs.¹⁵ As is the case with aromatic heterocyclic compounds that are more reactive than their aromatic homocyclic analogs, *closo*-heteroboranes are more reactive than their *closo*-homoboranes. Therefore the extensive derivative chemistry of *ortho*- $C_2B_{10}H_{12}$ is not surprising.¹⁶ As a consequence of the electronic polarization originating in the two heteroatoms, in this particular case the adjacent carbon atoms, different substitutions are possible on C and B.¹⁷ Therefore the *ortho*- $C_2B_{10}H_{12}$ offers more possible derivatizations in a small volume than the vast majority of the available chemical scaffolds; further, it also offers very distinct environments in a very short distance, e.g. one lipophilic environment next to a hydrophilic one. Consequently, using carboranes in supramolecular chemistry is a field to be explored for their particular characteristics^{13b,18} that shall induce unconventional properties in the supramolecular structures in which they are inserted.¹⁹ Boron clusters can form supramolecular polymeric structures not only due to the bridging of polymetallic centres by various non-boron ligands, but also *via* hydrophobic interactions or halogen or dihydrogen bonding of polyhedral carboranes.²⁰

In earlier studies, we have studied the coordination of the 1- CH_3 -2- CO_2H -1,2-*closo*- $C_2B_{10}H_{10}$ ligand to Cu^{II} and Mn^{II} ions together with their physical and chemical properties.²¹ In the case of Mn^{II} , the carboranycarboxylate ligands lead to highly uncommon water soluble inorganic polymers with water molecules bridging the Mn atoms; in contrast, for Cu^{II} the generated compounds do not differ substantially from the related organic based carboxylate ligands. We have also studied the molecular magnetism for these compounds containing carboranycarboxylate ligands; a strong antiferromagnetic interaction



1- CH_3 -2- $COOH$ -1,2-*closo*- $C_2B_{10}H_{10}$ (**LH**)

Chart 1 Drawing of the ligand used in this work.

has been observed in Cu^{II} compounds while there is a weak antiferromagnetic interaction in the corresponding Mn^{II} ones. In this work, we extend these studies to Co^{II} with the aim to learn whether the carboranyl fragment produces more uncommon coordination and nuclearity around the metal centre in the way it did with Mn^{II} or, contrastingly the performance is more classical as it is the case for Cu^{II} .

With all this in mind, in this work we describe the synthesis of new air-stable di-, tri- and polynuclear Co^{II} compounds using the carboranycarboxylate ligand 1- CH_3 -2- CO_2H -1,2-*closo*- $C_2B_{10}H_{10}$ (Chart 1); an exhaustive characterization of the obtained compounds as well as their magnetic properties is also presented in this work.

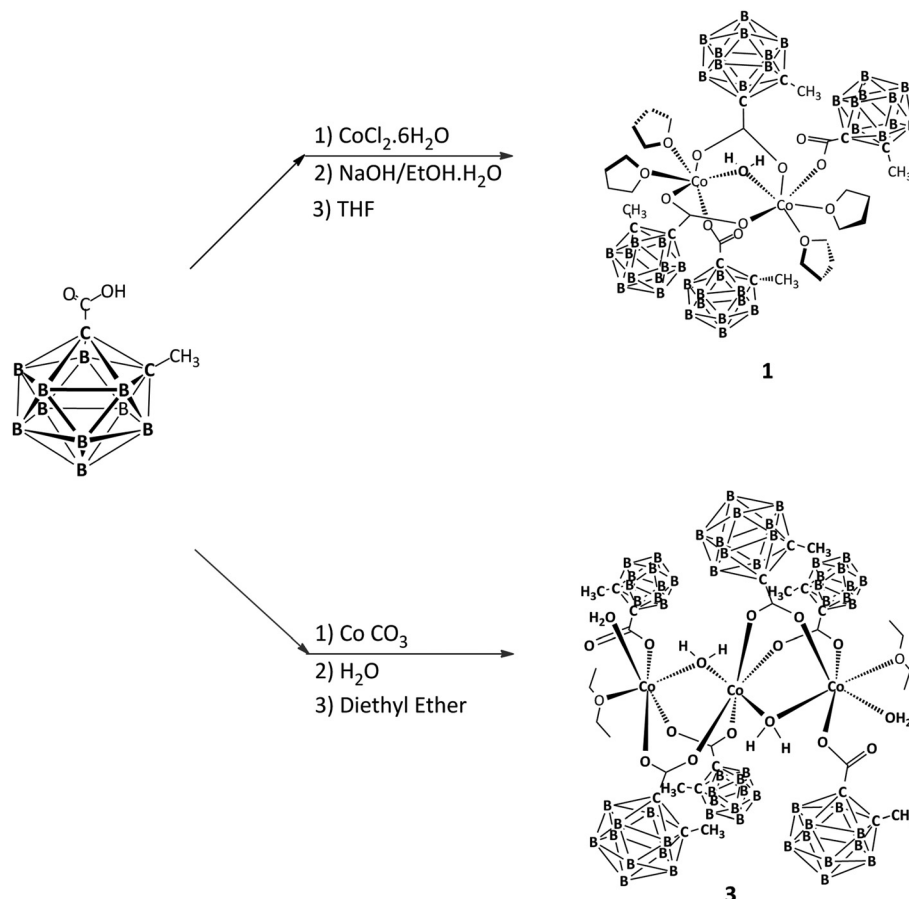
Results and discussion

Synthesis and structure

The synthetic strategy for the preparation of the Co^{II} compounds, **1–3**, containing the carboranycarboxylate ligand 1- CH_3 -2- CO_2H -1,2-*closo*- $C_2B_{10}H_{10}$ (**LH**) and THF, aqua or diethyl ether ligands is outlined in Schemes 1 and 2.

The dinuclear cobalt compound $[Co_2(\mu-H_2O)(1-CH_3-2-CO_2-1,2-*closo*- $C_2B_{10}H_{10}$)_4(THF)_4]$, **1**, was obtained by neutralization of carboranycarboxylic acid, **LH**, with NaOH followed by reaction with $CoCl_2 \cdot 6H_2O$ in ethanol/ H_2O and subsequent extraction in THF. Compound **2** was synthesized by reaction of a suspension of 1- CH_3 -2- CO_2H -1,2-*closo*- $C_2B_{10}H_{10}$, **LH**, with $CoCO_3$ in water. The trinuclear compound $[Co_3(\mu-H_2O)_2(1-CH_3-2-CO_2-1,2-*closo*- $C_2B_{10}H_{10}$)_6(H_2O)_2(C_4H_{10}O)_2]$, **3**, was obtained by recrystallization of **2** in diethyl ether. All attempts to obtain crystals of compound **2** in non-coordinating solvents or water were unsuccessful but its X-ray powder diffraction (XRD) was performed and compared with the calculated XRD for compound **3** (Fig. 1). Fig. 1 shows that both diagrams are different, proving that **2** and **3** display different structures. This led us to tentatively propose that complex **2** presents a polymeric structure (Scheme 2) where each Co^{II} atom is coordinated by four carboxylate oxygen atoms and two aqua oxygen atoms and is bridged to other Co atoms by two carboranycarboxylate ligands and by an aqua ligand, similar to the polymer manganese compound described by us,^{21b} which displays water molecules bridging every two Mn centres, an unusual feature in 1-D oligomer Mn^{II} compounds. In support of this, the elemental





Scheme 1 Synthetic strategy for the preparation of Co(II) compounds 1–3.

analysis of **2** is in agreement with the polymeric structure proposed (see the Experimental section). Compound **2** is broken in coordinating solvents such as diethylether, leading to a linear trinuclear Co^{II} compound **3**.

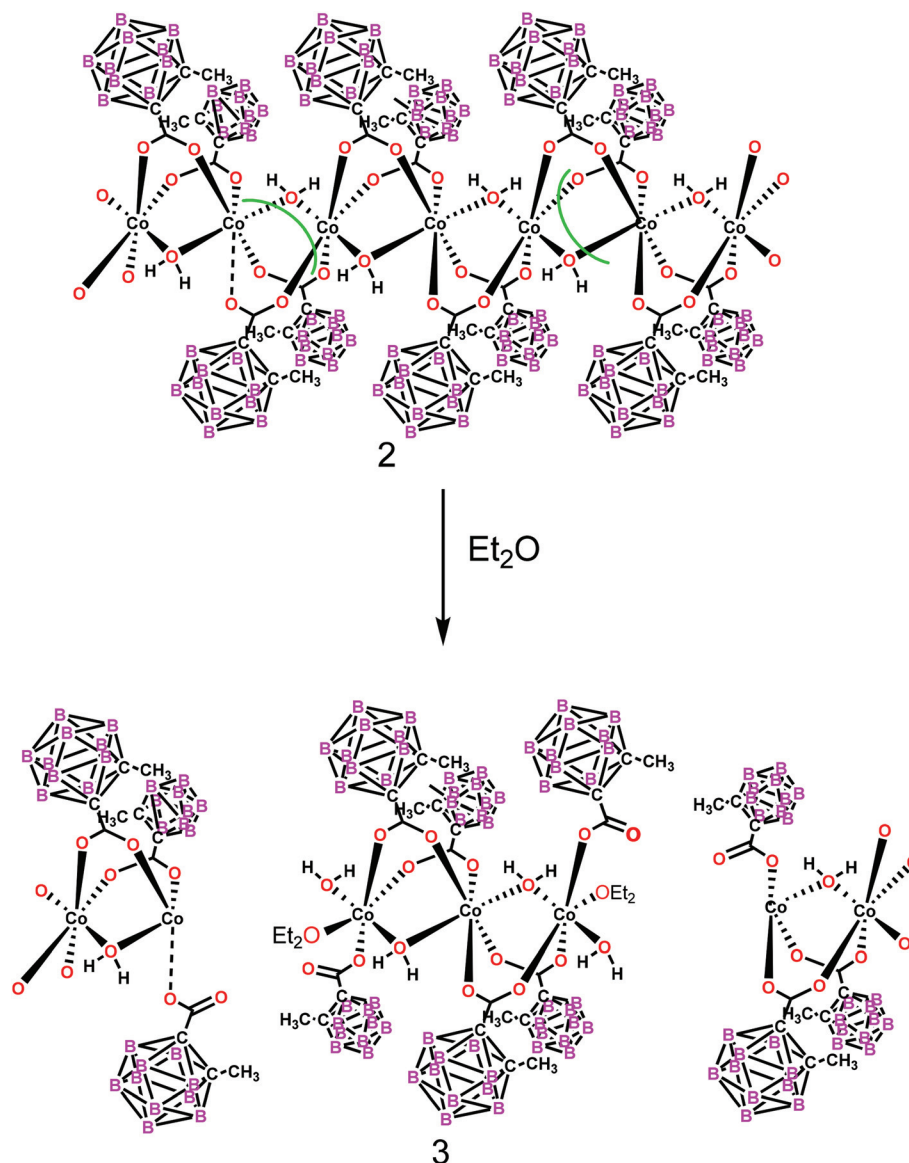
The crystal structures of **1** and **3** have been solved by X-ray diffraction analysis. Crystallographic data and selected bond distances and angles for **1** and **3** are presented in Tables 1 and 2, respectively. ORTEP plots with the corresponding atom labels for the X-ray structures of both compounds are presented in Fig. 2.

Dinuclear compound **1** displays a structure with the two cobalt(II) atoms holding together through two *syn,syn* η¹:η¹:μ₂-carboxylate bridges and one bridged aqua molecule as shown in Fig. 2. The distorted octahedral geometry around each metal ion is completed by one additional monodentate carboranylcarboxylate ligand and two oxygen donor atoms from two THF ligands leading to a local Co(O_{THF})₂(O_{carboxy})₃(O_{H₂O}) coordination. Although there are many reported examples with a CoO₆ environment^{10b,22} in which the metal ion is Co^{II}, compound **1** is the first one that contains the carboranyl fragment with structural and electronic parameters different from those previously reported in the literature. The molecule is located in a crystallographically imposed two-fold axis which passes through the oxygen O(10) belonging to the μ₂-bridged

aqua molecule. The hydrogen atoms of the bridged aqua molecule are involved in intramolecular O(10)–H...O–C hydrogen bonds with the uncoordinated O atoms of the monodentate carboranylcarboxylate ligands (O(10)–H...O(2)–C(10), 1.607 Å) (see Fig. 3). The bond length Co(1)–O(10)_w distance is 2.083(17) Å while the corresponding Co(1)–O(10)–Co(1)#1 angle is 117.82(14)°. This bond distance is shorter than the distance observed for similar carboxylate compounds containing pyridine as the terminal ligand in spite of THF.^{22c–e} The Co–O_{THF} bond length corresponding to the THF molecule in the *trans* position to the bridged aqua oxygen atom is significantly shorter (2.083 Å) than the same Co–O_{THF} distance for the THF molecule *trans* to the bridged carboranylcarboxylate oxygen atom (2.092 Å). The structural packing of this compound (Fig. 4a) along the *c* axis shows the formation of channels in which the carboranylcarboxylate ligands are orientated to these channels and opposite to the water bridged molecules, leading to the existence of a hydrophobic environment in these cavities. Interactions B–H...H–B of 2.284 Å between neighbouring molecules and B–H...H–C_{THF} (2.3–2.4 Å) between carboxylate ligands and THF molecules along the channels have been observed.

The X-ray diffraction of **3** discloses a trinuclear cobalt compound with aqua bridged entities that is made up of a linear





Scheme 2 Cleavage of **2** by Et_2O and formation of the trinuclear complex **3**.

array of three hexacoordinated Co^{II} ions with a crystallographic inversion centre located on the central $\text{Co}(2)$. This central $\text{Co}(2)$ is coordinated by four carboxylate and two aqua oxygen atoms and is bonded to both terminal $\text{Co}(1)$ ions by two carboranylcarboxylate ligands and by an aqua ligand; the distorted octahedral geometry of the central Co atom is probably similar to the one that is most probably presented by Co^{II} atoms in the polymeric structure of **2**. The two outer Co atoms fulfill their hexacoordination with one oxygen atom from one diethyl ether ligand, one terminal H_2O ligand, and one monodentate carboranylcarboxylate ligand that probably originally could be shared with another Co^{II} atom in the polymer structure (Scheme 2).

Cobalt centers display a tetragonal distortion that means that the average of the equatorial $\text{Co}-\text{O}$ distances for terminal Cobalt ions is 2.06 Å and the average of the axial distances is

2.13 Å. The distances and angles observed in this compound are similar and comparable to other similar trinuclear compounds described in the literature most of which form in part either linear chains^{23a} or coordination network systems.^{23c,d} It is worth mentioning that our structure is stabilized by terminal ligands as diethyl ether, which is never observed in cobalt trimer compounds containing carboxylate bridged ligands. It is well known that carboranes induce unconventional properties in the structures in which they are inserted.¹⁹ This effect is caused by their volume and their rigid geometry that provide a higher atomic efficiency if compared to aromatic phenyl ligands. These steric properties together with the higher electron-withdrawing character of the carboranylcarboxylate cluster stabilize the unique carboranyl structures presented in this paper. A similar situation could occur in the case of compound **1** where two terminal THF ligands are



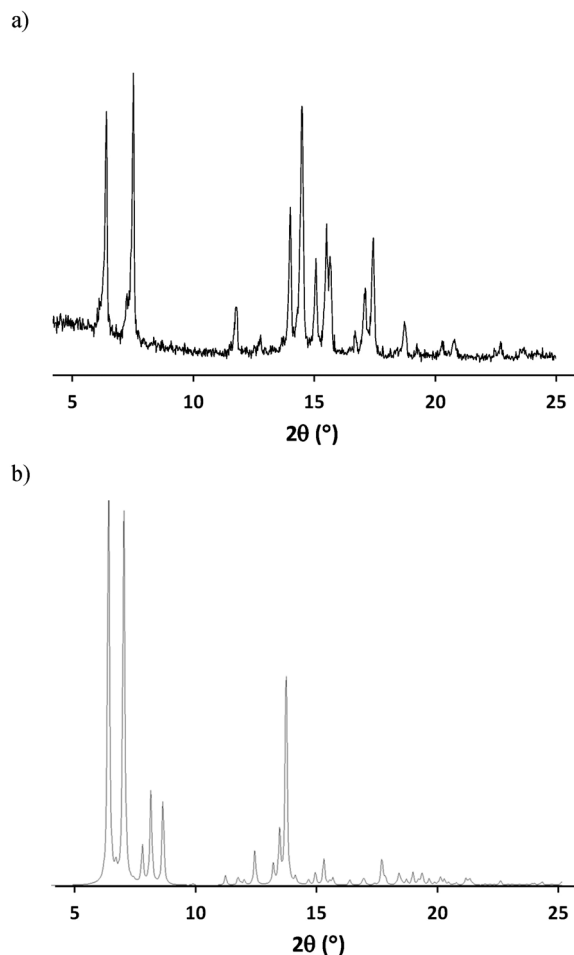


Fig. 1 a) X-ray powder diffraction (XRPD) for 2 and b) calculated XRPD for 3.

Table 1 Crystal data for X-ray structures of 1 and 3

	1	3
Empirical formula	C ₃₂ H ₈₆ B ₄₀ O ₁₃ Co ₂	C ₄₈ H ₁₄₆ B ₆₀ O ₂₂ Co ₃
Formula weight	1229.27	1901.04
Crystal system	Orthorhombic	Triclinic
Space group	<i>Pccn</i>	<i>P</i> $\bar{1}$
<i>a</i> [Å]	12.392(5)	14.420(9)
<i>b</i> [Å]	22.830(10)	14.806(10)
<i>c</i> [Å]	23.802(10)	15.656(10)
α [°]	90	63.006(12)
β [°]	90	73.533(13)
γ [°]	90	67.443(11)
<i>V</i> [Å ³]	6734(5)	2726(3)
Formula units/cell	4	1
ρ_{calc} [g cm ⁻³]	1.212	1.158
μ [mm ⁻¹]	0.542	0.506
R_1^a , [<i>I</i> > 2 σ (<i>I</i>)]	0.0636	0.0979
wR_2^b [all data]	0.2183	0.2603

^a $R_1 = \sum \|F_o - F_c\| / \sum F_o$. ^b $wR_2 = [\sum \{w(F_o^2 - F_c^2)^2\} / \sum \{w(F_o^2)^2\}]^{1/2}$, where $w = 1/[\sigma^2(F_o^2) + (0.0042P)^2]$ and $P = (F_o^2 + 2F_c^2)/3$.

present around each cobalt atom. The existence of two intra-molecular hydrogen bonds between the uncoordinated oxygen atom of the monodentate carboxylate (O(2)) and H8A and H8B of the bridged aqua (H(8A)–O(2), 1.666 Å; H(8B)–O(2), 2.955 Å) is also noticeable. Another intramolecular weak hydrogen bond is observed between H7B of the terminal aqua ligand and O(9) of the terminal diethyl ether molecule, H(9B)–O(7), 2.891 Å (Fig. 3b).

The packing arrangement of compound 3 along the *c* axis shows parallel lineal chains where the terminal diethyl ether ligands of adjacent molecules from the same chain are facing each other, Fig. 4b.

Spectroscopic properties

The IR spectra of the compounds described display typical $\nu(\text{B–H})$ absorption at frequencies above 2590 cm⁻¹ (Fig. S1†), characteristic of *closo* carborane derivatives.²⁴ Differences between the frequencies of the symmetric and antisymmetric stretches of the carboxylate ligands lie within the ranges quoted for bidentate bridged ligands.^{24c} The ¹H{¹¹B}-, ¹¹B-, ¹¹B{¹H}- and ¹³C{¹H}-NMR spectra of these compounds were recorded and are presented in the ESI (Fig. S2†). It can be observed that the ¹H{¹¹B}- spectra exhibit resonances around $\delta = 2$ ppm attributed to the C_c–CH₃ protons and the resonances of the protons bonded to the B atoms appear as broad singlets over a wide chemical-shift range in the region from $\delta = 0$ to +5 ppm. The ¹¹B{¹H}-NMR resonances for all the compounds featured similar patterns in the range from $\delta = 2$ to –15 ppm that agrees with a *closo* cluster.²⁵

The electronic UV-vis spectra of compounds 1 and 2 are shown in Fig. S3† and display two characteristic bands, one in the range 300–400 nm and the other in the range 420–570 nm. The pink binuclear compound 1 revealed two bands at 353 nm and 534 nm, whereas that the spectrum for the solid 2 presents one band at 510 nm, and two shoulders at 356 and 472 nm. The higher energy bands observed in these compounds can be assigned to ligand to metal charge transfer (LMCT) transitions. Absorptions in the visible region at 534 nm for 1 and at 472 and 510 nm for 3 are assigned to the two spin-allowed d–d transitions ⁴T_{1g}(F) → ⁴A_{2g}(F) and ⁴T_{1g}(F) → ⁴T_{1g}(P).^{22b,e,26}

Magnetic properties

Magnetic structural correlations were performed for compounds 1 and 3, where in both cases the cobalt centres are linked to each other through [1-CH₃-2-CO₂-1,2-*closo*-C₂B₁₀H₁₀]⁻ ligands and a single molecule of H₂O. The two compounds contain hexacoordinated Co^{II} centres with distorted octahedral geometries and in both (1, 3) the bulky nature of the ligand prevents intermolecular interactions among neighbouring molecules (distances Co^{II}...Co^{II} above 9 Å). Altogether, 3 can be seen as the extension of 1 by the addition of one extra Co^{II} centre and therefore, it can be predicted that magnetically both systems will behave in a similar manner. Variable-temperature magnetic susceptibility



Table 2 Selected bond lengths (Å) and angles (°) for compounds 1 and 3

	1		3		3
Co(1)–O(3)	2.039(2)	Co(1)–O(1)	2.099(5)	Co(2)–O(6)	2.052(4)
Co(1)–O(6)	2.083(2)	Co(1)–O(3)	2.059(5)	Co(2)–O(4)	2.072(4)
Co(1)–O(10)	2.0832(17)	Co(1)–O(5)	2.027(5)	Co(2)–O(8)	2.134(4)
Co(1)–O(5)	2.092(2)	Co(1)–O(7)	2.021(5)	Co(2)–O(6)#1	2.052(4)
Co(1)–O(4)	2.092(2)	Co(1)–O(8)	2.217(4)	Co(2)–O(4)#1	2.072(4)
Co(1)–O(1)	2.121(2)	Co(1)–O(9)	2.111(5)	Co(2)–O(8)#1	2.134(4)
O(3)–Co(1)–O(6)	90.64(10)	O(7)–Co(1)–O(8)	88.09(16)	O(4)–Co(2)–O(4)#1	180.0(2)
O(3)–Co(1)–O(10)	91.76(8)	O(1)–Co(1)–O(9)	90.96(19)	O(4)–Co(2)–O(6)	91.91(19)
O(6)–Co(1)–O(10)	176.61(9)	O(3)–Co(1)–O(9)	89.45(19)	O(4)#1–Co(2)–O(6)	88.09(19)
O(3)–Co(1)–O(5)	178.25(10)	O(5)–Co(1)–O(9)	88.78(18)	O(4)–Co(2)–O(6)#1	88.09(19)
O(6)–Co(1)–O(5)	88.23(10)	O(7)–Co(1)–O(9)	92.40(19)	O(4)#1–Co(2)–O(6)#1	91.91(19)
O(10)–Co(1)–O(5)	89.42(8)	O(8)–Co(1)–O(9)	178.67(19)	O(6)–Co(2)–O(6)#1	180.0(1)
O(3)–Co(1)–O(4)	90.10(10)	O(1)–Co(1)–O(3)	175.49(19)	O(4)–Co(2)–O(8)#1	90.43(17)
O(6)–Co(1)–O(4)	89.01(9)	O(1)–Co(1)–O(5)	92.1(2)	O(4)#1–Co(2)–O(8)#1	89.57(17)
O(10)–Co(1)–O(4)	93.39(8)	O(3)–Co(1)–O(5)	92.42(19)	O(6)–Co(2)–O(8)#1	84.84(16)
O(5)–Co(1)–O(4)	88.54(9)	O(1)–Co(1)–O(7)	87.08(19)	O(6)#1–Co(2)–O(8)#1	95.16(16)
O(3)–Co(1)–O(1)	90.37(10)	O(3)–Co(1)–O(7)	88.4(2)	O(4)–Co(2)–O(8)	89.57(17)
O(6)–Co(1)–O(1)	88.95(9)	O(5)–Co(1)–O(7)	178.56(18)	O(4)#1–Co(2)–O(8)	90.43(17)
O(10)–Co(1)–O(1)	88.63(8)	O(1)–Co(1)–O(8)	87.83(16)	O(6)–Co(2)–O(8)	95.16(16)
O(5)–Co(1)–O(1)	90.95(9)	O(3)–Co(1)–O(8)	91.79(16)	O(6)#1–Co(2)–O(8)	84.84(16)
O(4)–Co(1)–O(1)	177.91(8)	O(5)–Co(1)–O(8)	90.71(16)	O(8)#1–Co(2)–O(8)	180.0(1)

measurements for 1 and 3 were performed in the 2–300 K temperature range applying a 0.5 T dc field. Fig. 5 shows together the two plots of $\chi_M T$ vs. T for 1 and 3. The $\chi_M T$ values at 300 K for compounds 1 and 3 are 7.26 and 13.26 cm³ mol^{−1} K, respectively. Both values are much higher than the spin-only value expected for non-interacting $S = 3/2$ centres (3.75 cm³ mol^{−1} K and 5.62 cm³ mol^{−1} K, in that order, when $g = 2$). This fact is a clear indication of the significant orbital contribution of the metallic ions, well-known for high-spin octahedral Co^{II} (HS *Oh* Co^{II}) centres.²⁷ The continuous decrease observed in $\chi_M T$ for 1 and 3 with decreasing temperature may be due to weak antiferromagnetic interactions and/or to the depopulation of spin–orbit split levels arising from the ⁴T_{1g} state in the *Oh* Co^{II} centres (the ⁴T_{1g} term in HS *Oh* Co^{II} ions splits into a sextet: a quartet and a Kramers' doublet by spin–orbit coupling).²⁸ At 2 K, the $\chi_M T$ of 1 arrives at a value of 2.41 cm³ mol^{−1} K whereas the $\chi_M T$ of 3 reaches a value of 4.10 cm³ mol^{−1} K at the same temperature. At such temperatures, only the Kramers' doublets are populated and each Co^{II} centre could be considered to have an effective spin $S' = 1/2$. The $M/N\mu_B$ values at 2 K and 5 T are of 4.7, and 7.7 for compounds 1 and 3, respectively, in the range of what should be expected for two and three $S' = \frac{1}{2}$ systems with relevant spin orbital contributions (Fig. S4†).

M vs. H/T data were also recorded at different applied external fields (0.5–5 T) in the 1.8–6.8 K temperature range (Fig. S5†). The magnetization sharply increases upon cooling the sample to 1.8 K and sweeping field up to 5 T, for the two systems, and clear saturation was observed in the reduced magnetization plots of both (1–3) at the highest achievable field and lowest temperature with values of 4.71 and 7.30 $N\mu_B$, respectively. No superposition of the isofield lines was observed, suggesting the existence of relevant anisotropy.²⁹

Our goal in the present work was the estimation of the J parameter using the whole temperature range. This has only been possible in recent years through sophisticated computer programs due to the magnetic complexity of cobalt species.³⁰ Here, the data were fitted by diagonalization of the spin Hamiltonian matrix, using the program PHI,³¹ which allows the correlation of experimental magnetic data of orbitally degenerate systems using multiple sources; in this case, $\chi_M T$ vs. T data together with M vs. H/T results were used simultaneously. This way, the $\chi_M T$ vs. T and M vs. H/T data fittings in the whole 2–300 K range (Fig. 1) afforded the parameters: $J = -0.30$ cm^{−1}, $g = 2.18$, $TIP = 490 \times 10^{-6}$ cm³ mol^{−1} and $\rho = 0.6\%$ for 1 and $J_1 = J_2 = -0.40$ cm^{−1}, $g = 2.40$, $TIP = 633 \times 10^{-6}$ cm³ mol^{−1} and $\rho = 0.5\%$ for 3. As Fig. 5 depicts, a more reliable fitting was obtained in the case of 3, thus the final data achieved for 1 should be taken as illustrative although, due to the structural similarities among the two compounds the resulting magnetic parameters for the latter may be close to reality.

As mentioned above, magnetic measurements for bridged Co^{II} compounds and more precisely, for μ -OH₂, μ -O₂CR dinuclear/trinuclear Co^{II} systems, are scarce in the literature. However, some worth noting publications have appeared in the last few years.^{10a,32} Dinuclear Co^{II} systems, described with a molecule of H₂O and two carboxylates as bridging ligands between the two metallic centres, are more abundant than trinuclear species. Magnetically, the characterization in former times was mostly based on the oxidation state and description of the nature of the exchange (always antiferro) but, excellent studies have already provided exchange parameter values between -1 and -3 cm^{−1}.^{10b,22b,33} In the case of Co₃ species, only two systems have been well-characterized magnetically: Calvo-Pérez *et al.*^{10a} performed exhaustive analyses of a system related to 3, finding that the J value was -0.4 cm^{−1} and similar



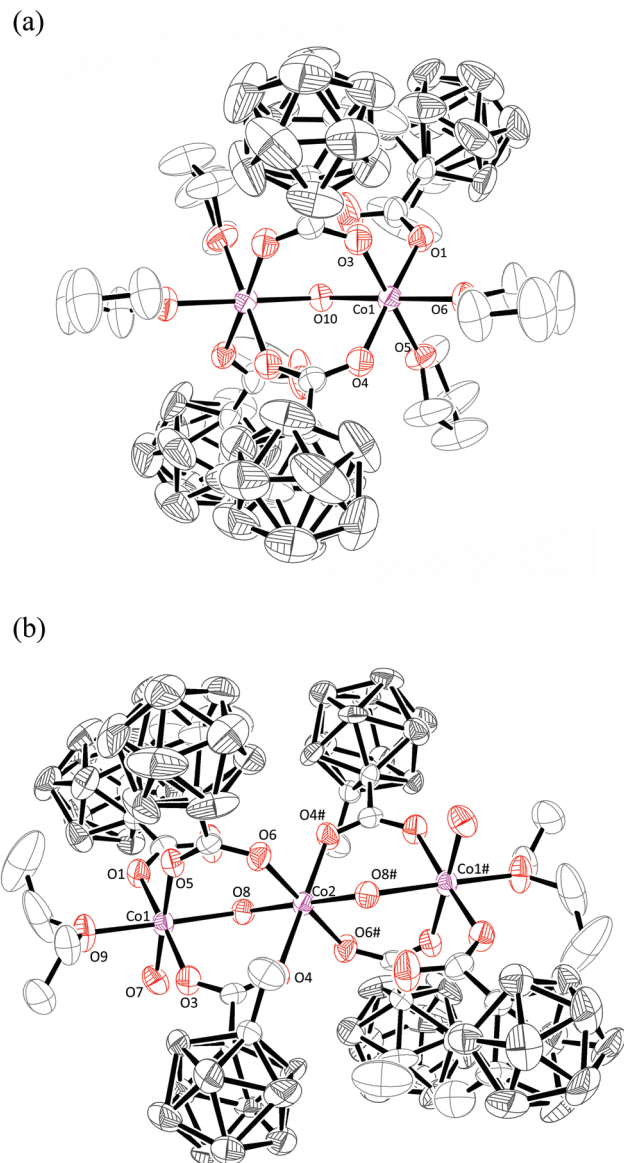


Fig. 2 X-ray structures (ORTEP plots with ellipsoids at 40% probability level) and the labelling scheme for compounds a) **1** and b) **3**.

values where described in additional Co_3 systems by some of the same authors some years later.³² The comparison of the data accomplished in this work and those previously published allows some preliminary information. This way, our results confirm what has been described in the past: structures containing strictly two carboxylate ligands and one H_2O as bridging units among Co^{II} centres provide small and negative exchange couplings (related to the *syn-syn* conformation of the carboxylate ligands and the existence of one molecule of H_2O as a linker). On the other hand, even though the range of J values is very small (from -0.4 to -3 cm^{-1}) and the number of systems is not that extensive, it is worth mentioning that the lowest values (-0.4 cm^{-1}) have been accomplished with carboxylates of an electron-withdrawing nature ($-\text{O}_2\text{CCF}_3$)^{10a} and ours, the carboranylcarboxylate ligand $[\text{1-CH}_3\text{-2-CO}_2\text{-1,2-closo-}$

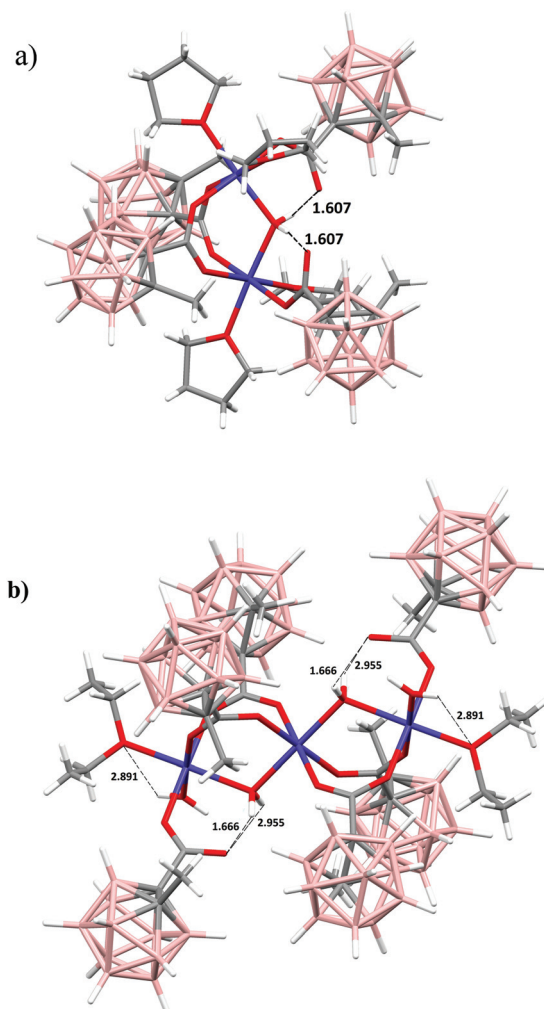


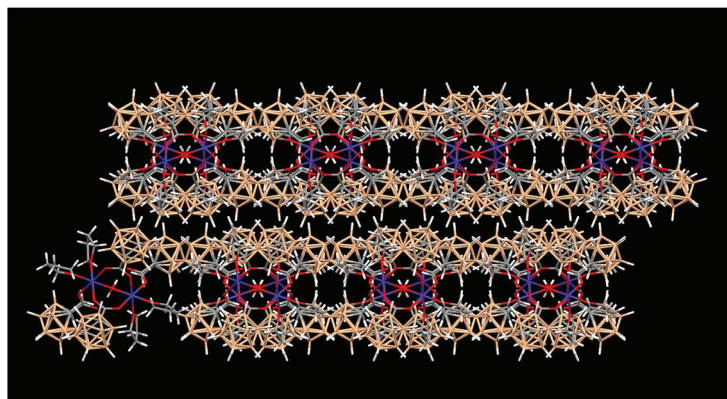
Fig. 3 X-ray structures showing the intramolecular hydrogen bonds, a) for **1** and b) for **3**.

$\text{C}_2\text{B}_{10}\text{H}_{10}]^-$. In the past, we observed similar effects regarding this feature in dinuclear paddle-wheel Cu^{II} structures.^{21a} Summarizing, in this work, compounds **1** and **3** exhibit weak anti-ferromagnetic behaviours among the Co^{II} centres as it has been reported for analogous systems of Mn^{II} published in the past by the group in ref. 21c.

Conclusions

An unprecedented strategy for the synthesis of a 1D Co^{II} carboranycarboxylate polymer together with a straightforward methodology to achieve trinuclear molecules based on this 1D system has been presented in this work. The reaction of the carboranycarboxylate ligand, 1-CH₃-2-CO₂H-1,2-*closo*-C₂B₁₀H₁₀ (LH) with CoCO₃ in water leads easily to the isolation of the polymer. Its crystallization in slightly coordinating solvents, such as diethylether, leads to the trinuclear compound with diethyl ether acting as the ancillary ligand on the two terminal

a)



b)

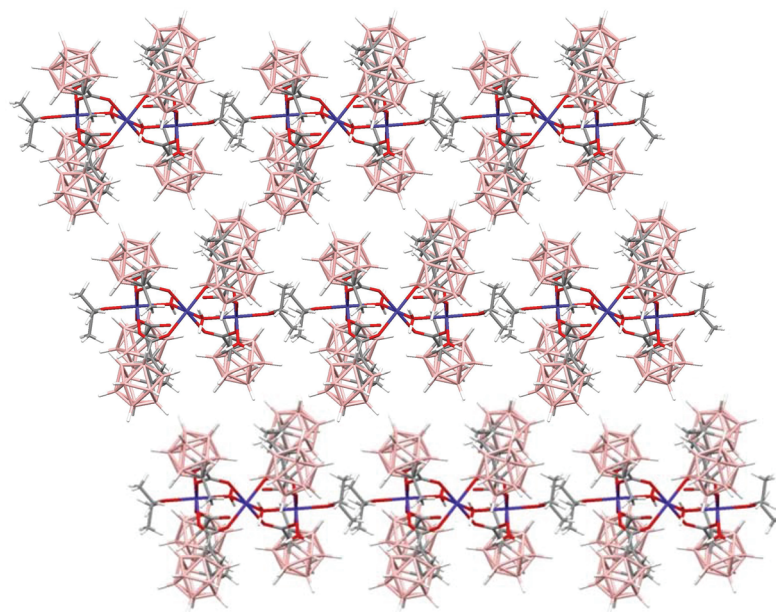


Fig. 4 Packing diagram of compounds a) **1** along the *c* axis b) **3** along the *c* axis.

Co^{II} ions. This structure containing diethyl ether as terminal ligands has never been observed in cobalt trimer compounds containing carboxylate bridged ligands; this effect is caused by the unconventional properties of the carboranyl ligand such as its volume, and rigid geometry together with its electron-withdrawing character that stabilizes the unique structure presented in this paper. This crystal shows the inner Co...Co pairs with the bridging water unit and the two bridging carboxylate ligands while the terminal Co centres contain a monodentate carboranylcarboxylate ligand, one terminal water molecule and one terminal fragmenting solvent, Et₂O, evidencing their origin as a result of polymer fragmentation. A new dinuclear structure has also been presented in this work, where Co atoms held together through two carboxylate bridges and one bridged aqua molecule and one additional monodentate car-

boranylcarboxylate ligand and two oxygen donor atoms from two THF ligands complete the octahedral geometry. Interestingly, the packing of this compound shows the formation of hydrophobic cavities due to the orientation of the carboranylcarboxylate ligands in the crystal. Magnetic measurements of polynuclear compounds were carried out, and showed, in all cases, weak antiferromagnetic interactions between the cobalt atoms. This methodology to achieve small molecules by carving 1D Co^{II} carboranylcarboxylate systems with organic solvents could be an important challenge in the development of new molecular building blocks with different ancillary solvent ligands. These molecules tend to impart a dynamic behaviour to the potential resultant coordination networks and consequently improve their properties. More studies in this sense are being developed in our laboratory.



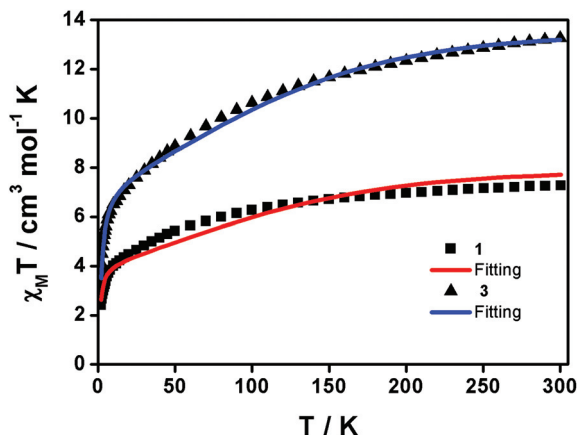


Fig. 5 Experimental data including $\chi_M T$ vs. T of **1** (black squares) and **3** (black triangles) between 2.0 and 300.0 K. The red and blue lines correspond to the theoretical values from the fitting, also in that order.

Experimental section

Materials

All reagents used in the present work were obtained from Aldrich Chemical Co and were used without further purification. Reagent grade organic solvents were obtained from SDS and high purity de-ionized water was obtained by passing distilled water through a nano-pure Mili-Q water purification system. 1-CH₃-1,2-*closo*-C₂B₁₀H₁₁ was purchased from Katchem.

Instrumentation and measurements

FT-IR spectra were recorded on a Mattson-Galaxy Satellite FT-IR spectrophotometer containing a MKII Golden Gate Single Reflection ATR System. Elemental analyses were performed using a CHNS-O Elemental Analyser EA-1108 from Fisons. UV-Vis spectroscopy was performed on a Cary 50 Scan (Varian) UV-Vis spectrophotometer with 1 cm quartz cells or with an immersion probe of 5 mm path length. NMR spectra have been recorded with a Bruker ARX 300 or a DPX 400 instrument equipped with the appropriate decoupling accessories. ¹H and ¹H{¹¹B} NMR (300.13/400.13 MHz), ¹³C{¹H} NMR (75.47/100.62 MHz) and ¹¹B and ¹¹B{¹H} NMR (96.29/128.37 MHz) spectra were recorded in d₆-acetone and D₂O. Chemical shift values for ¹¹B NMR spectra were referenced to external BF₃ ← OEt₂ and those for ¹H, ¹H{¹¹B} and ¹³C{¹H} NMR spectra were referenced to SiMe₄. Chemical shifts are reported in units of parts per million downfield from the reference, and all coupling constants are in Hz. The deconvolution of ¹¹B{¹H} spectra has been performed with the software OriginPro 8 SR0, v. 8.0724.

X-ray structure determination

Measurement of the crystals was performed on a Bruker Smart Apex CCD diffractometer using graphite-monochromated Mo

K α radiation ($\lambda = 0.71073$ Å) from an X-ray tube. Data collection, Smart V. 5.631 (BrukerAXS 1997-02); data reduction, Saint+ Version 6.36A (Bruker AXS 2001); absorption correction, SADABS version 2.10 (Bruker AXS 2001) and structure solution and refinement, SHELXL-2013 (Sheldrick, 2013). For structure **3**, two disordered ethyl ether solvent molecules per asymmetric unit were removed using the SQUEEZE option in PLATON.³⁴ The crystallographic data as well as details of the structure solution and refinement procedures are reported in Table 1. CCDC 1471844 (**1**) and 1471845 (**3**) contain the supplementary crystallographic data for this paper.

Magnetic susceptibility studies

Magnetic susceptibility measurements were carried out on polycrystalline samples with a DSM5 Quantum Design susceptometer working in the range 2–300 K under a magnetic field of 0.5 T. TIP and ρ stand for the temperature-independent paramagnetism and impurities (in %), respectively. Diamagnetic corrections were estimated from Pascal Tables.

Preparations

1-CH₃-2-CO₂H-1,2-*closo*-C₂B₁₀H₁₀ was prepared according to the literature procedures.³⁵ The synthetic manipulations of cobalt compounds were routinely performed under ambient conditions.

Synthesis of [Co(μ -H₂O)(1-CH₃-2-CO₂-1,2-*closo*-C₂B₁₀H₁₀)₂] \cdot (THF)₄, **1.** 1-CH₃-2-CO₂H-1,2-*closo*-C₂B₁₀H₁₀, **LH**, (0.200 g, 0.990 mmol) in ethanol (1.5 mL) was neutralized with a 0.1 M aqueous NaOH solution with phenolphthalein as an indicator at room temperature and immediately mixed with a solution of CoCl₂·6H₂O (0.120 g, 0.490 mmol) in water (1 mL). The temperature of the pink solution gradually rose to 40 °C. After 2 h the solvent was removed under vacuum, and the solid residue was redissolved in THF (30 mL) and again dried under vacuum. The last process was repeated three times. The obtained residue was once more dissolved in THF, the solution was filtered to remove solid Na₂SO₄ and the solvent was again removed under vacuum. The resulting pink solid was recrystallized by the slow diffusion of pentane into a diethylether solution to afford air-stable pink crystals of **1** suitable for X-ray diffraction analysis. Yield: 0.28 g (93%). ¹H{¹¹B} NMR (400.13 MHz, d₆-acetone, 25 °C): δ = 4.70 (s, μ -H₂O), 2.30 (br s, B-H), 2.23 (br s, B-H), 2.13 (s, CH₃). ¹¹B{¹H} NMR (128.37 MHz, d₆-acetone, 25 °C): δ = −1.4 (1B), −2.5 (1B), −3.7 (1B), −6.9 (2B), −9.6 (1B), −10.5 (1B), −11.1 (1B), −12.8 (2B) ppm. IR: ν = 2985, 2894, 2581, 1677, 1443, 1361, 916, 871, 761, 725 cm^{−1}. UV-Vis (CH₂Cl₂, 1.10^{−3} M) $\lambda_{\max}(\epsilon)$ = 353 nm (134 M^{−1} cm^{−1}), 534 nm (97 M^{−1} cm^{−1}). Elemental analysis calcd (%) for C₃₂H₈₆B₄₀O₁₃Co₂·0.3THF: C 31.88, H 7.12; found: C 31.77, H 7.34.

Synthesis of [Co(μ -H₂O)(1-CH₃-2-CO₂-1,2-*closo*-C₂B₁₀H₁₀)₂] \cdot (H₂O)_m, **2.** To a suspension of 1-CH₃-2-CO₂H-1,2-*closo*-C₂B₁₀H₁₀, **LH**, (0.093 g, 0.458 mmol) in water (20 mL) was added CoCO₃ (0.055 g, 0.458 mmol) in water (5 mL). The solution was stirred and heated to 40 °C for 22 h. Afterward, the solution was filtered and the solvent was removed under



vacuum to obtain a pale pink solid. The resulting product was recrystallized in a dichloromethane solution. Yield: 0.103 g (94%). $^1\text{H}\{^{11}\text{B}\}$ NMR (400.13 MHz, D_2O , 25 °C): δ = 2.19 (br s, B–H), 2.10 (br s, B–H), 1.96 (br s, B–H), 1.90 (br s, B–H), 1.79 (s, CH_3). $^{11}\text{B}\{^1\text{H}\}$ NMR (128.37, D_2O , 25 °C): δ = –4.0 ($^1\text{J}(\text{B},\text{H})$ = 148.9, 1B), –6.9 ($^1\text{J}(\text{B},\text{H})$ = 151.5, (1B), –10.09 3B), –10.91 ($^1\text{J}(\text{B},\text{H})$ = 147.6, 5B). IR: ν = 3585, 3556, 2583, 1376, 1149, 1015, 846, 772, 725 cm^{-1} . UV-Vis (H_2O , 1.10^{-3} M) $\lambda_{\text{max}}(\epsilon)$ = 356 nm (5 $\text{M}^{-1} \text{cm}^{-1}$), 472 nm (4 $\text{M}^{-1} \text{cm}^{-1}$), 510 nm (5 $\text{M}^{-1} \text{cm}^{-1}$). Anal. Found (calcd) (%) for $\text{C}_8\text{H}_{28}\text{O}_5\text{B}_{20}\text{Co} \cdot \text{H}_2\text{O}$: C 19.52(19.32); H 6.17(6.08).

Synthesis of $[\text{Co}_3(\mu\text{-H}_2\text{O})_4(1\text{-CH}_3\text{-2-CO}_2\text{-1,2-closo-C}_2\text{B}_{10}\text{H}_{10})_6(\text{C}_4\text{H}_{10}\text{O})_2]$, 3. By recrystallization of 2 in a diethyl ether solution and then slow diffusion of pentane into this solution, colourless needles suitable for X-ray diffraction were obtained corresponding to complex 3.

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