Triphos derivatives and diphosphines as ligands in the ruthenium-catalysed alcohol amination with NH₃†

N. Nakagawa, a E. J. Derrah, a M. Schelwies, b F. Rominger, c O. Trapp c and T. Schaub a,b

The ruthenium-triphos and diphosphine-catalysed amination of alcohols with ammonia is reported. Various types of triphos derivatives with electron-donating functional group were synthesized and used as ligands in the Ru-catalysed alcohol amination with NH₃. The triphos derivatives are effective for the formation of primary amines. On the other hand, if hemilabile diphosphines as tridentate ligands are used, mixtures of secondary-along with primary amines are obtained. It was found that even simple diphosphines can be used as ligands for the selective formation of the secondary amines. The diphosphine system allows a new entry to the Ru-catalysed formation of secondary amines.

Introduction

Amines are important building blocks in the manufacture of pharmaceuticals and agrochemicals.1 Intensive research has focused on the selective formation of primary, secondary and tertiary amines.2 One of the most straightforward and environmentally-friendly methods is the amination of alcohols using cheap and abundant ammonia (NH₃), which represents an atom efficient route towards this valuable class of organic compounds, along with water as the sole by-product.3 The amination of alcohols with ammonia is performed with heterogeneous catalysts on industrial scale. Using heterogeneous catalysts, it is in many cases difficult to control the selectivity as harsh reaction conditions are required.1c During the last decade, the amination of alcohols has been further developed with a series of homogeneous catalysts.4 In particular, ruthenium catalysts were identified to exhibit a high performance for the selective formation of primary amines5 and secondary amines.6 The combination of a ruthenium pre-catalyst and the tridentate phosphate ligand 1,1,1-tris(diphenylphosphinomethyl)ethane 1 (= triphos) was first published as one of the candidates in the alcohol amination by Beller and co-workers7a and investigated in detail by our laboratory (Scheme 1).7b We disclosed that the catalytic active species is a cationic ruthenium-triphos complex through experimental and computational investigations. Our previous report on this ruthenium catalysis described only the use of the commercial available triphos 1 and effects of other phosphate derivatives have still remained undisclosed and undeveloped. We herein report the synthesis and evaluation of triphos derivatives and the corresponding ruthenium complexes in order to gain an insight how variations of the triphos scaffold influence the ruthenium-catalysed amination of alcohols with NH₃. Our approach was to alter the coordination sphere of triphos-type Ru-complexes in order to change the selectivities in the corresponding amination reactions.

Results and discussion

Ligands

To evaluate the ligand influence on the catalytic performance of the ruthenium-catalysed amination of 1-octanol with NH₃,
the following triphos derivatives were used: triphos 1, the newly synthesised phenyl-p-tolyl-mixed triphos derivative 2, and the all-p-tolyl triphos derivative 3 (obtained via a modified procedure). We also examined diphenosphines with an electron-donating functional group such as a pyridyl substituted derivative 4, which is potentially hemilabile ligand, as well as oxygen-substituted ones 5 and 6 (Fig. 1).

**Synthesis of triphos derivatives via modified protocol**

C$_{iv}$ symmetrical triphos-type ligands such as 3 can be obtained using 1,1,1-tris(chloromethyl)ethane as a starting material. A route to C1 symmetric variants of the ligand with three different phosphine donors was reported by Huttner and Helmchen. However, since we were only interested altering one of the three donor groups of triphos, we chose 7 as key intermediate in our synthesis (Scheme 2). The substitution reaction of CH$_2$C(CH$_2$OMs)$_2$(CH$_2$Br) with an excess amount of lithium diphenylphosphide, followed by borane protection afforded a monomesyl diphosphine–borane compound 6-PG in 64% yield (Scheme 2a). After deprotection with DABCO, the intermediate 6 was obtained in 66% yield. We repeated this sequence using lithium di(p-tolyl)phosphide with monomesylate 6 and obtained the unsymmetrical triphos–borane 2-PG. The structure of 2-PG was confirmed by X-ray analysis (Fig. 2). Deprotection of 2-PG afforded the unsymmetrical triphos 2 in 27% overall yield. The same protocol was applied to the synthesis of all-p-tolyl triphos 3 (51% overall yield, Scheme 2b).

**Synthesis of ruthenium complexes**

Ruthenium complexes of the triphos derivatives 1–3 were prepared according to the reported protocol for Ru-triphos complex A (Scheme 3). When using 2 as a ligand, a mixture of three isomers is formed. $^{31}$P-NMR indicates that the three isomers differ in the position of the tolyl moiety. Elemental ana-
lysis confirms the composition of the mixture of B-1, B-2, and B-3. In case of the all-tolyl substituted ligand 3, only one complex (C) is formed like in the synthesis of A. The structure of C was determined by X-ray analysis (Fig. 3). The distance of Ru–P bonds in C is the range of 2.27–2.38 Å which is similar like in the reported structure for RuHCl(CO)(triphos) A.\textsuperscript{13}

We also conducted the preparation of the corresponding ruthenium complexes with pyridyl substituted diphosphine 4.\textsuperscript{9}

Using this ligand, a mixture of three complexes was obtained:

the two isomers D-1 and D-2 as well as the dichloro ruthenium species D-3 (Scheme 4). The structure of D-3 was confirmed by X-ray analysis (Fig. 4). Regarding the structure of D-3, the carbonyl ligand is located trans to the pyridyl moiety. Although the reaction pathway towards D-3 is not confirmed, we propose that the chlorine atom on D-3 was derived from the CH\textsubscript{2}Cl\textsubscript{2} solvent through the chlorine abstraction by ruthenium hydride species. Under these conditions, selective formation and complete separation of the three isomers turned out to be difficult.

In addition, we failed to obtain the ruthenium–pyridyl diphosphine complex selectively in toluene as solvent. Therefore, we used the pyridyl diphosphine 4 directly in combination with [Ru(PPh\textsubscript{3})\textsubscript{3}(CO)HCl] in the amination reaction to generate the desired complexes \textit{in situ} (\textit{vide infra}).

We were also successful to synthesise a novel hydroxyl-containing ruthenium complex E from 5\textsuperscript{10} (Scheme 5). The structure of E was confirmed by X-ray analysis (Fig. 5). One PPh\textsubscript{3} ligand from the precursor remains on ruthenium which is con-
Catalytic investigations

In order to investigate the ligand effects, the ruthenium-catalysed amination of 1-octanol with NH$_3$ was conducted for the above described ligands 1–6 (Table 1). As previously reported by our group,$^7$ the ruthenium-triphos complex A showed a high catalytic performance towards octylamine 8 along with small amounts of the secondary and tertiary amines 9 and 10 (entry 1). Note that it is unnecessary to use the isolated ruthenium catalyst A: the in situ prepared ruthenium-triphos catalyst showed the same activity and selectivity as A (entry 2). In the absence of any additional phosphine, the amination products 8 to 10 were not obtained (entry 3). When the p-tolyl-substituted triphos derivatives 2 and 3 were used instead of 1, it was found that the selectivity towards the primary amine is slightly shifted to dioctylamine 9 (entries 4 and 5). The use of the pyridyl diphosphine 4, a potentially hemilabile ligand, led to the formation of dioctylamine 9 (entry 6). The amination reactions using the hydroxyl–diphosphine complex E and its alkoxo complex F were sluggish and the dioctylamine 9 is formed in good selectivity (entries 7 and 8). These results can be interpreted as the complex E probably converts into the alkoxo complex F through deprotonation of the hydroxyl group under the basic conditions of the amination. When the mesyl diphosphine 6, which is the synthetic intermediate of the ligands 2 and 3, was used as a ligand, dioctylamine 9 is formed in good selectivity (entry 9). The mesyl group (OMs) will be displaced by the amino group (NH$_2$) via nucleophilic substitution under the ammonia atmosphere of the amination reaction.$^{11}$ To our surprise, even simpler diphosphines, such as dpppdmmp (1,3-bis(diphenylphosphino)-2,2-dimethylpropane, Ph$_2$PCH$_2$C(CH$_3$)$_2$CH$_2$PPh$_2$) and dppp (1,3-bis(diphenylphosphino)propane, Ph$_2$PCH$_2$CH$_2$CH$_2$PPh$_2$), were found to be good ligands for the preferable formation of dioctylamine 9 (entries 10 and 11).$^{14,15}$ Finally, the highly selective formation of dioctylamine 9 was achieved with dpppp at a higher substrate concentration (entries 12 and 13).

The reduced amount of toluene significantly affected the formation of the secondary amine. In order to get more insights in the reaction, we examined the reaction profile by analysing the reaction mixture at several reaction times (1–8 h; Fig. 6).$^{16}$ The obtained curves showed that there is no significant induction period, and that the amount of octylamine 8 and dioctylamine 9 increases as time elapses without decomposing or interconverting into the other products. We also confirmed the formation of small amount of octylcyclohexan-1-imine 11, which is assumed to be the intermediate before dioctylamine 9 is formed. The imine 11 is probably reactive to a Ru–H species to produce dioctylamine 9 through hydrogenation.

The possible mechanism is described in Scheme 6. As we postulated in the previous report,$^7$ the selective mono-alkylation of ammonia using ruthenium-triphos complexes involves the formation of cationic ruthenium-triphos intermediates as active species (Scheme 6a). Among the triphos derivatives, the different ratio between the primary and secondary amines can originate from the difference of electron negativity of P atom. In the case of hemilabile diphosphines and simple diphosphines, the formation of the secondary amine is preferred under the same reaction conditions. This observation can be explained by considering a different mechanism. As Vogt and co-workers disclosed the mechanism on the amination of cyclohexanol using xantphos,$^{16}$ the dissociation of the remaining
PPh₃ can play an important role (Scheme 6b). Although the mechanism of dialkylation is still unclear, it is possible to consider that dissociation of PPh₃ and chloride can provide more vacant sites on the ruthenium than in the triphos systems. The transient ketone or imine can coordinate at the vacant site and incorporate in the further alkylation. When octylamine as a starting material is used instead of 1-octanol, the formation of dioctylamine was only observed in small amounts (eqn (1)). Therefore, it is conceivable that the ruthenium–diphosphate complex possesses a poor ability on dehydrogenation of the primary amine. Presumably, in situ generated ketone or imine (from the starting alcohol) are the reactants for the dilakylation.

Conclusions

A series of ruthenium-complexes bearing triphos derivatives and diphosphines with the electron-donating functional group were prepared and examined for the amination of primary alcohols with NH₃. Deviating the P-substituents on the triphos scaffold slightly affects the ratio of mono- and dialkylamines, but these triphos derivatives are basically effective for the formation of the primary amines. The reaction mechanism of mono-alkylation can involve the formation of the cationic ruthenium complex based on our previous investigations. On the other hand, the use of hemilabile diphosphines leads to a mixture of the primary and secondary amines. It seems to be clearer after these results, that the origin of high selectivities towards the primary amines is related to the very stable ligand spheres of Ru/triphos and the Milstein type Ru/pincer.
The selective formation of the secondary amine was finally realized by using simple dppp as the ligand. The reaction mechanism of the ruthenium – diphosphine system can be influenced by dissociation of attaching PPh3 which is different to the triphos-based systems. So far, the preparation of secondary amines from alcohols with NH3 using homogeneous catalysts has been achieved by Milstein and co-worker with Ru catalyst, or by Fujita, Yamaguchi, and co-workers as well as in previous work from our laboratory using Ir catalysts. Compared with these studies, our current diphosphine system offers a simple protocol for the preparation of diamines with commercially available catalysts.

**Experimental**

**General considerations**

All reactions were carried out under a positive pressure of argon in an MBraun glovebox or using standard Schlenk line techniques. All nondeuterated solvents were dried using an MBraun SPS-800 solvent purification system and degassed prior to use. 3-Methyl-3-oxetanemethanol (a precursor of 7), and 1-octanol were purchased from Aldrich and distilled prior to use. Di-p-tolyl and di-phenylphosphines were purchased from ABCR, RuHCl(CO)(PPh3)3 was supplied by BASF and used without further purification. All other products were purchased from Aldrich and used without further purification. Liquid reagents including deuterium solvents were distilled prior to use, and all others were used without further purification. 1H, 13C{1H}, and 31P{1H} NMR spectra were recorded on a Bruker Avance 200, 400, or 600 MHz spectrometer. 1H and 13C chemical shifts are reported relative to residual solvent signals of CD2Cl2 (5.32 and 54.0 ppm) and THF-d8 (5.38 and 67.21 ppm). 31P{1H} chemical shifts are referenced to an external 85% solution of phosphoric acid. The 13C NMR data were assigned by HSQC and HMBC spectra. FAB and HR mass spectrometry was measured at the Mass Spectrometry Facility (Institute of the Organic Chemistry, University Heidelberg). Gas chromatography was performed on an Agilent 6890N modular GC base equipped with a split-mode capillary injection system and a flame ionization detector using a BGB-5 capillary column (Agilent 122-1033; 30 m × 0.32 mm × 0.25 μm; He flow 1.0 mL min⁻¹, program: initial 50 °C for 2 min, ramp 6 °C min⁻¹, 300 °C for 10 min). Starting materials and products had the following retention times: octylamine (t_R = 14.54 min), 1-octanol (t_R = 15.16 min), diocetylamine (t_R = 31.34 min), and trioctylamine (t_R = 41.71 min). Elemental analysis were performed in the "Mikroanalytisches Laboratorium der Chemischen Institute der Universität Heidelberg". X-ray structures were solved by direct methods and refined against F² with a full-matrix least squares algorithm by using the SHELXTL (version 2014/7) software package. Intensities were corrected for Lorentz and polarisation effects. CCDC 1440322 (2-PG), 1440323 (C), 1440324 (D-3), and 1440325 (E) contain the supplementary crystallographic data for this paper.

**Synthesis of H3CC(CH2OMs)2(CH2Br) (7)**

A 100 mL round-bottom Schlenk flask was charged with 25 mL dioxane and 3-methyl-3-oxetanemethanol (5.45 g, 53 mmol). Aqueous HBr (48%, 7.2 mL, 64 mmol) was slowly added over a 5–7 min period to give a light yellow solution that was slightly warm to the touch. The solution was refluxed for 3.5 h and the solvent was removed under vacuum (60 °C) to give H3CC(CH2OH)2(CH2Br) (9.38 g, 51 mmol, 96% yield) as an orange-brown solid. The product was used in the next step without further purification.

1H NMR (200.1 MHz, CD2Cl2, δ): 3.63 (s, 4H, CH2OH), 3.54 (s, 2H, CH2Br), 2.95 (s, 2H, OH), 0.91 (s, 3H, CH3).

A 500 mL round-bottom Schlenk flask was charged with H3CC(CH2OH)2(CH2Br) (4.08 g, 22.3 mmol), triethylamine (4.96 g, 0.49 mol) and 200 mL CH2Cl2. The colourless solution was cooled in an ice bath and neat methanesulfonyl chloride (5.36 g, 47 mmol) was added dropwise to give a colourless solution and white precipitate. The solution was stirred for 2 h
at 0 °C. The solution was concentrated to ~100 mL under vacuum and water (100 mL) was added. The organic layer was extracted and washed with another 100 mL of water, dried over MgSO₄, and the organic solvent was removed under vacuum to obtain H₃CC(CH₂OMes)₂(CH₂Br) 7 (6.14 g, 18 mmol, 81% yield) as a thick yellowish oil.

**Synthesis of H₃CC(CH₂PPh₂-BH₃)₂(CH₂OMes) (6-PG)**

A 500 mL round-bottom Schlenk flask was charged with HPP₃H (7.99 g, 43 mmol) and 60 mL THF. The colourless solution was cooled in an ice bath and 2.5 M n-BuLi in hexane (19 mL, 47 mmol) was added drop wise via syringe to give a red/orange solution. The solution was stirred for 30 min before being transferred to a 100 mL addition funnel affixed to a 500 mL round-bottom Schlenk flask charged with H₃CC(CH₂OMes)₂(CH₂Br) (7.28 g, 21 mmol) in 150 mL THF. The solution was cooled to −40 °C and the LiPP₃-solution was added slowly over a 2 h period. The temperature was maintained at −40 °C for 3 h before being slowly warmed up to room temperature and the orange mixture was stirred overnight. A solution of BH₃-SMe₂ (35 mL, 71 mmol) in THF (2.0 M) was slowly added to the light yellow reaction mixture. The resulting mixture was stirred for 2 h and the solvent was removed under vacuum. 250 mL diethyl ether (250 mL) and 250 mL water were added to the white solid. The organic layer was separated and washed with 250 mL water, dried over MgSO₄, and the organic solvent was removed under vacuum. The white solid was loaded onto a silica column (10 × 4 cm) and eluted with a 70:30 mixture of ether and petroleum ether. The first fraction was filtered through a small plug of silica (3 cm) on a glass filter frit and washed with toluene (3 × 100 mL). The solvent was removed from the supernatant to give H₃CC(CH₂PPh₂-BH₃)₂(CH₂OMes) 6 (1.57 g, 2.9 mmol, 52% yield) as a thick white oil.

**Synthesis of H₃CC(CH₂PPh₂-BH₃)₂(CH₂P(p-tol)-BH₃) (2-PG)**

A 100 mL round-bottom Schlenk flask was charged with HP(p-tol)₂ (0.344 g, 2.1 mmol) and 25 mL THF. The colourless solution was cooled in an ice bath and 2.5 M n-BuLi in hexane (0.89 mL, 2.2 mmol) was slowly added via syringe to give a red/orange solution. The solution was stirred for 30 min before drop wise addition to a 25 mL THF solution of H₃CC(CH₂PPh₂-BH₃)₂(CH₂OMes) (845 mg, 1.58 mmol) in a 100 mL round-bottom Schlenk flask cooled to −40 °C. After the addition the solution was slowly warmed to room temperature and stirred overnight. The reaction mixture was additionally heated to reflux for 2 h. A solution of BH₃-SMe₂ (2.84 mL, 5.69 mmol) in THF (2.0 M) was slowly added to the reaction mixture. The resulting mixture was stirred overnight and the volatile compounds were removed under vacuum. 100 mL Et₂O and 100 mL water were added to the white solid. The organic layer was separated and the aqueous layer was extracted with Et₂O for 2 times. The combined organic layer was washed with 100 mL water for 2 times, and then dried over MgSO₄. After the removal of Et₂O under vacuum, the white residue in 20 mL Et₂O was diffused into hexane to give a white powder. The obtained powder was filtered off and washed 2 times with 10 mL hexane. After drying in vacuum, H₃CC(CH₂PPh₂-BH₃)₂(CH₂P(p-tol)-BH₃) 2-PG (0.710 g, 1.02 mmol, 65% yield) was obtained as a white powder.
Synthesis of RuHCl(CO)(H₂CC(CH₂PPh₂)₂(CH₂P(p-tol))₃) (B)

A round-bottom Schlenk flask was charged with H₂CC(CH₂PPh₂)₂(CH₂P(p-tol))₃ (0.663 g, 1.0 mmol), RuHCl(CO)(PPh₃)₃ (0.881 g, 0.92 mmol) and 40 mL toluene. The colorless solution with an off white precipitate was heated at reflux for 2 h to give a light yellow precipitate. The solution was filtered through a filter frit, the yellow residue washed three times with 10 mL toluene and dried in vacuum to give RuHCl(CO)(H₂CC(CH₂PPh₂)₂(CH₂P(p-tol))₃) (B) (0.544 g, 0.66 mmol, 72% yield) as a yellowish white powder.

Improved synthesis of H₂CC(CH₂P(p-tol))₃ (3)

A 500 mL round-bottom Schlenk flask was charged with HP(p-tol)₂ (3.06 g, 14 mmol) and 80 mL THF. The colorless solution was cooled in an ice bath and 2.5 M n-ButLi in hexanes (6 mL, 15 mmol) was slowly added via syringe to give a red/orange solution. The solution was stirred for 30 min before drop wise addition to a THF (100 mL) solution of H₂CC(CH₂OMes)(CH₂Br) (7.12 g, 3.6 mmol) in a 500 mL round-bottom Schlenk flask cooled to −40 °C. The resulting mixture was slowly warmed to room temperature, then heated at reflux overnight. A solution of BH₃·SMe₂ (8.0 mL, 16 mmol) in THF (2.0 M) was slowly added to the colorless solution. The resulting mixture was stirred for 2 h and the solvent was removed under vacuum. 100 mL Ether and 100 mL water were added to the white solid. The organic layer was separated and washed with 150 mL water, dried over MgSO₄ and the organic solvent was removed under vacuum. The residue was suspended in 20 mL diethyl ether, 20 mL hexane was added, the product filtered off and washed twice with 20 mL hexane. After drying in vacuum H₂CC(CH₂P(p-tol))₂·BH₃ (3-PG, 1.92 g, 2.6 mmol, 72% yield) was obtained as a white powder. A round-bottom Schlenk flask was charged with H₂CC(CH₂PPh₂)₂(2-pyridyl) (4 (160 mg, 0.31 mmol, 1.2 equiv.), PPh₃)RuHCl(CO) (250 mg, 0.26 mmol, 1 equiv.) and dichloromethane (40 mL). The light yellow solution was stirred at room temperature for 5 days to give yellow solution. The solution was filtered to remove a small amount of white precipitate and the volume was reduced to 2 mL. A fine yellow precipitate was obtained with the addition of hexane (∼20 mL). The yellow solid contained a mixture of D-1, D-2, and D-3. Recrystallization by slow solvent diffusion of pentane (20 mL) into a concentrated DCM solution (2 mL) gave a small amount of 21 (∼20 mg) in reasonable purity (∼95%). A second recrystallization attempt of the same sample gave a mixture of D-1, D-2, and D-3. Single crystal of D-3 was isolated from the sample and used for X-ray diffraction analysis.

Synthesis of RuHCl(CO)(H₂CC(CH₂P(p-tol))₃) (C)

A round-bottom Schlenk flask was charged with H₂CC(CH₂P(p-tol))₃ (500 mg, 0.71 mmol), RuHCl(CO)(PPh₃)₃ (612 mg, 0.64 mmol) and 40 mL toluene. The colorless solution with an off white precipitate was heated at reflux for 2 h to give a light yellow precipitate. 10 mL pentane was added to the reaction mixture, the precipitate was filtered off via a sinter frit and washed three times with 10 mL pentane. After drying in vacuum, RuHCl(CO)(H₂CC(CH₂P(p-tol))₃) C (499 mg, 0.57 mmol, 80% yield) was obtained as a yellow powder.

Improved synthesis of H₂CC(CH₂P(p-tol))₃ (3)

A 500 mL round-bottom Schlenk flask was charged with HP(p-tol)₂ (3.06 g, 14 mmol) and 80 mL THF. The colorless solution was cooled in an ice bath and 2.5 M n-ButLi in hexanes (6 mL, 15 mmol) was slowly added via syringe to give a red/orange solution. The solution was stirred for 30 min before drop wise addition to a THF (100 mL) solution of H₂CC(CH₂OMes)(CH₂Br) (7.12 g, 3.6 mmol) in a 500 mL round-bottom Schlenk flask cooled to −40 °C. The resulting mixture was slowly warmed to room temperature, then heated at reflux overnight. A solution of BH₃·SMe₂ (8.0 mL, 16 mmol) in THF (2.0 M) was slowly added to the colorless solution. The resulting mixture was stirred for 2 h and the solvent was removed under vacuum. 100 mL Ether and 100 mL water were added to the white solid. The organic layer was separated and washed with 150 mL water, dried over MgSO₄ and the organic solvent was removed under vacuum. The residue was suspended in 20 mL diethyl ether, 20 mL hexane was added, the product filtered off and washed twice with 20 mL hexane. After drying in vacuum H₂CC(CH₂P(p-tol))₂·BH₃ (3-PG, 1.92 g, 2.6 mmol, 72% yield) was obtained as a white powder. A round-bottom Schlenk flask was charged with H₂CC(CH₂PPh₂)₂(2-pyridyl) (4 (160 mg, 0.31 mmol, 1.2 equiv.), PPh₃)RuHCl(CO) (250 mg, 0.26 mmol, 1 equiv.) and dichloromethane (40 mL). The light yellow solution was stirred at room temperature for 5 days to give yellow solution. The solution was filtered to remove a small amount of white precipitate and the volume was reduced to 2 mL. A fine yellow precipitate was obtained with the addition of hexane (∼20 mL). The yellow solid contained a mixture of D-1, D-2, and D-3. Recrystallization by slow solvent diffusion of pentane (20 mL) into a concentrated DCM solution (2 mL) gave a small amount of 21 (∼20 mg) in reasonable purity (∼95%). A second recrystallization attempt of the same sample gave a mixture of D-1, D-2, and D-3. Single crystal of D-3 was isolated from the sample and used for X-ray diffraction analysis.

Reaction of RuHCl(CO)(PPh₃)₃ with H₂CC(CH₂PPh₂)₂(2-pyridyl)

A round-bottom Schlenk flask was charged with H₂CC(CH₂PPh₂)₂(2-pyridyl) (4 (160 mg, 0.31 mmol, 1.2 equiv.), PPh₃)RuHCl(CO) (250 mg, 0.26 mmol, 1 equiv.) and dichloromethane (40 mL). The light yellow solution was stirred at room temperature for 5 days to give yellow solution. The solution was filtered to remove a small amount of white precipitate and the volume was reduced to 2 mL. A fine yellow precipitate was obtained with the addition of hexane (∼20 mL). The yellow solid contained a mixture of D-1, D-2, and D-3. Recrystallization by slow solvent diffusion of pentane (20 mL) into a concentrated DCM solution (2 mL) gave a small amount of 21 (∼20 mg) in reasonable purity (∼95%). A second recrystallization attempt of the same sample gave a mixture of D-1, D-2, and D-3. Single crystal of D-3 was isolated from the sample and used for X-ray diffraction analysis.

Elemental Analysis Calc. (Found): C 64.58% (64.16%), H 5.42% (5.32%).
D-3: $^{31}$P($^1$H) NMR (121.4 MHz, CD$_2$Cl$_2$, $\delta$): 37.9 ppm ($s$); $^1$H NMR (300.1 MHz, CD$_2$Cl$_2$, $\delta$): all peaks are overlapping with those of D-1 and D-2.

**Synthesis of RuHCl(CO)(PPh$_3$)$_3$[H$_2$CC(CH$_2$PPh$_2$)$_2$]CH$_2$OH] (E)**

In a glovebox, a round-bottom Schlenk flask was charged with H$_2$CC(CH$_2$PPh$_2$)$_2$CH$_2$OH 5 (0.502 g, 1.1 mmol), RuHCl(CO)(PPh$_3$)$_3$ (952 mg, 1.0 mmol) and 40 mL toluene. The colorless solution with an off white precipitate heated at reflux for 2 h, then the volume of toluene was reduced under vacuum to ca. 10 mL. In the glovebox, 10 mL pentane was added to the reaction mixture, forming a pale yellow powder. The precipitate was filtered off via a sinter frit and washed three times with 10 mL pentane. After drying in vacuum, RuHCl(CO)(PPh$_3$)$_3$[H$_2$CC(CH$_2$PPh$_2$)$_2$]CH$_2$OH] (E) (0.610 g, 0.69 mmol, 69% yield) was obtained as a pale yellow powder. The $^{31}$P NMR indicated that the isolated compounds were the mixture of regioisomers of E.

IR (KBr): 517, 695, 742, 806, 839, 1054, 1091, 1158, 1188, 1312, 1434, 1481, 1586, 1924 (vs, CO and tBu-1H), 2928, 3053 cm$^{-1}$.

Elemental Analysis Calc. (Found): C 65.19% (62.58%), H 6.73% (6.73%), N 5.24% (5.15%).

$^{31}$P($^1$H) NMR (162.1 MHz, CD$_2$Cl$_2$, $\delta$): 44.2 ($dd$, $^2$J$_{PP} = 282$ Hz, $^2$J$_{HP} = 20$ Hz), 43.6 ($dd$, $^2$J$_{PP} = 282$ Hz, $^2$J$_{HP} = 20$ Hz), 27.4 ($dd$, $^2$J$_{PP} = 282$ Hz, $^2$J$_{HP} = 20$ Hz), 26.9 ($dd$, $^2$J$_{PP} = 282$ Hz, $^2$J$_{HP} = 20$ Hz), 3.0 ($t$, $^3$J$_{HP} = 20$ Hz), 0.4 ($t$, $^3$J$_{HP} = 20$ Hz). All peaks are derived from the regioisomers of E; $^1$H NMR (400.3 MHz, CD$_2$Cl$_2$, $\delta$): 8.05–8.60 (m, 35H, CH), 3.30–0.18 (m, 9H, CH$_2$ and CH$_3$), $–$6.73 ($ddd$, $^3$J$_{HP} = 109$ Hz, $^2$J$_{HP} = 23$ Hz, $^2$J$_{HP} = 17$ Hz, 1H, RuH), $–$6.83 ($ddd$, $^3$J$_{HP} = 109$ Hz, $^2$J$_{HP} = 23$ Hz, $^2$J$_{HP} = 17$ Hz, 1H, RuH). All peaks are derived from the regioisomers of E.

**Reaction and RuHCl(CO)(PPh$_3$)$_3$[H$_2$CC(CH$_2$PPh$_2$)$_2$]CH$_2$OH] (E) with t-BuOK**

In a glovebox, a round-bottom Schlenk flask was charged with RuHCl(CO)(PPh$_3$)$_3$[H$_2$CC(CH$_2$PPh$_2$)$_2$]CH$_2$OH] E (0.114 g, 0.13 mmol), t-BuOK (14.4 mg, 0.13 mmol) and 5 mL toluene, forming a red solution at room temperature in 2 h. All volatile compounds were removed under vacuum, then 0.5 mL toluene and 5 mL pentane was added to a red powder. The resulting suspension was passed through a pad of celite, then a red filtrate was dried under vacuum to give alkoxo ruthenium complex F (81.3 mg, 0.096 mmol, 75% yield) as a deep red powder. The $^1$H and $^{31}$P NMR spectra of the obtained powder suggested that it was constituted of mainly alkoxo ruthenium complex F with some impurities including small amounts of the unreacted E.

Characteristic peaks of F are described below:

$^{31}$P($^1$H) NMR (243.0 MHz, d$_6$-THF, $\delta$): 35.7 ($dd$, $^2$J$_{PP} = 276$ Hz, $^2$J$_{HP} = 20$ Hz), 33.5 ($dd$, $^2$J$_{PP} = 276$ Hz, $^2$J$_{HP} = 20$ Hz), 13.7 ($dd$, $^2$J$_{PP} = 20$ Hz); $^1$H NMR (600.2 MHz, d$_6$-THF, $\delta$): $–$4.65 ($ddd$, $^2$J$_{HP} = 113$ Hz, $^3$J$_{HP} = 26$ Hz, $^3$J$_{HP} = 13$ Hz, 1H, RuH). CH$_3$ peaks from t-Bu group were not observed in $^1$H NMR spectra.

IR (KBr): 515, 695, 742, 836, 999, 1092, 1186, 1434, 1480, 1586, 1896 (m, tBu-H), 1910 (vs, CO), 2674, 2754, 2921, 2951, 3051 cm$^{-1}$.

MS (LIFDI): m/z: 848.01 (100%, F$^+$).

**A representative procedure for catalyst screening described in Table 1**

Entry 1: In an argon-filled glovebox, to a premix autoclave (60 mL, stainless steel) equipped with a magnetically coupled propeller blade stirrer was added RuHCl(CO)(triphos) A (36 mg, 0.047 mmol), toluene (17 mL), and 1-octanol (3.0 g, 24 mmol), then the autoclave was sealed. After introducing NH$_3$ gas (7–8 bar) at room temperature, the autoclave was heated to 165 °C for 15 h with vigorous stirring (700–800 rpm) without reintroducing NH$_3$. After cooling to the room temperature, the pressurized NH$_3$ gas in the autoclave was released in a fume hood. An aliquot of the reaction mixture was taken for the GC analysis, and the conversion of 1-octanol and the yield of amine products were calculated based on GC area %.

Entry 2: RuHCl(CO)(PPh$_3$)$_3$ (45 mg, 0.047 mmol) and triphos (33 mg, 0.052 mmol) were used.

Entry 3: RuHCl(CO)(PPh$_3$)$_3$ (45 mg, 0.047 mmol) was used.

Entry 4: RuHCl(CO)[H$_2$CC(Ch$_2$PPh$_2$)$_2$]CH$_2$F]$_3$ (38 mg, 0.047 mmol) was used.

Entry 5: RuHCl(CO)[H$_2$CC(Ch$_2$P$_2$tol)$_2$]$_3$ (40 mg, 0.047 mmol) was used.

Entry 6: RuHCl(CO)(PPh$_3$)$_3$ (45 mg, 0.047 mmol) and CH$_2$C[H$_2$PPh$_2$]$_2$-2-pyridyl 4 (26 mg, 0.046 mmol) were used.

Entry 7: RuHCl(CO)(PPh$_3$)[H$_2$CC(Ch$_2$P$_2$tol)$_2$]CH$_2$OH] (E) (41 mg, 0.047 mmol) was used.

Entry 8: The proposed alkoxo-ruthenium complex F (20 mg, 0.023 mmol), toluene (8.5 mL), and 1-octanol (1.5 g, 12 mmol) were used.

Entry 9: RuHCl(CO)(PPh$_3$)$_3$ (45 mg, 0.047 mmol) and CH$_2$C[H$_2$PPh$_2$]$_2$-2-pyridyl 4 (26 mg, 0.046 mmol) were used.

Entry 10: RuHCl(CO)(PPh$_3$)$_3$ (45 mg, 0.047 mmol) and dppp (23 mg, 0.052 mmol) were used.

Entry 11: RuHCl(CO)(PPh$_3$)$_3$ (45 mg, 0.047 mmol) and dppp (21 mg, 0.052 mmol) were used.

Entry 12: RuHCl(CO)(PPh$_3$)$_3$ (45 mg, 0.047 mmol), dppp (21 mg, 0.052 mmol), and toluene (6 mL) were used.

Entry 13: RuHCl(CO)(PPh$_3$)$_3$ (45 mg, 0.047 mmol), dppp (21 mg, 0.052 mmol), toluene (6 mL), and p(NH$_3$) = 4 bar (53 mmol of NH$_3$) were used.

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**Notes and references**


11 Attempts to install other diorganophosphino moiety or heteroatom such as P(o-tolyl), P(p-F-C₆H₄)₂, PCy₂, PPr₃, and NH₂, were unsuccessful by using our approach.


14 Preparation and characterization of RuHCl(CO)(dppp) (PPh₃) from RuHCl(CO)(PPh₃)₃ and dppp was reported, see: (a) A. Santos, J. Lópe, J. Montoya, P. Noheda, A. Romero and A. M. Echavarren, *Organometallics*, 1994, 13, 3605; (b) S. Huh, Y. Cho and J.-M. Jun, *Polyhedron*, 1994, 13, 1887.

15 Cyclohexanol and ammonia undergo the selective nomo-alkylation with dppp ligand. See ref. 5e.

16 In Fig. 6, P(NH₃) = 4.1–4.5 bar was used, which corresponded to approximately 53 mmol of NH₃ in the autoclave.

