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Herein we report a dinickel azacryptand complex that enables fast, selective, and tight CO₂ binding from air. Exploiting the affinity of the cavitand towards azides, CO₂ release was observed. Despite the stability of the azido complex, UV irradiation under atmospheric conditions proved to be a suitable pathway for N₃[−] replacement by CO₂.

The continuous emission of carbon dioxide from combustion processes is a major reason for the increase of the global greenhouse effect. Thus CO₂ capture and sequestration technologies (CCS) are required to lower its concentration in our atmosphere and allow for storage and transportation.¹ In addition, reversible systems are needed that allow for its successive release and application as a C1 building block in industrial processes. In industry, aqueous ethanolic amine solutions are commonly applied for CCS but show high energy consumption and severe corrosion problems.² Furthermore, amino acids,³ metal oxides,⁴ and organic polymers⁵ were evaluated for this purpose. Additionally, metal–organic frameworks show good CO₂ uptake capabilities,^{6,7} and are also reported to allow for CO₂ reduction.^{8–10} However, although promising, they do not allow for a selective, non-competitive guest absorption/uptake of CO₂ *versus e.g.* H₂, N₂, or CH₄ and increasing the CO₂ affinity results in difficult desorption.¹¹ Furthermore, water, a significant component of flue gases, was found to predominantly decrease CO₂ capacities.¹²

In contrast, dinuclear metalloenzymes are known to selectively bind CO₂ and catalyze hydrolysis as well as redox reactions at ambient conditions. Among those, the nickel containing ureases¹³ as well as the iron and nickel containing

Bimetallic nickel complexes for selective CO₂ carbon capture and sequestration†

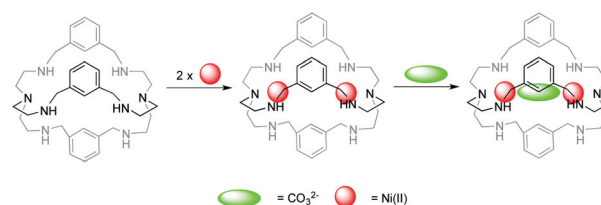
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CO dehydrogenases (CODH_{Ni})¹⁴ can bind CO₂ during the conversion of urea to CO₂/NH₃ or the reduction of CO₂ to CO, respectively. Commonly, such CO₂ fixation is facilitated by a Ni–OH moiety. Whereas complexes comprising a Ni–OH center are widely known to facilitate CO₂ fixation and Holm *et al.* showed their outstanding potential for fast CO₂ fixation,^{15–17} only few systems exist that selectively bind CO₂ between two nickel centers.^{18–22} General problems associated with such systems are either loose binding of the generated carbonate to nickel or formation of oligomers with bridging carbonates rendering the systems not applicable for reversible CCS and transportation. Contrary to those complexes, cryptands comprise well-defined metal binding cavities with tuneable metal–metal distances that allow for a strong binding of specific anions. Nelson and coworkers showed the uptake of carbonates into the cavity of dinickel azacryptands affording a stable dinickel complex with an *anti-anti* η₁,η₁-bound carbonate (Scheme 1).²³ Furthermore, Sullivan and coworkers showed that in presence of H₂, CO₂ can be reduced to afford CO and CH₄.²⁴ However, neither CO₂ uptake kinetics nor selectivity studies were reported in order to apply azacryptands for CCS. Additionally, Lu as well as Malthouse and coworkers showed evidence for a direct atmospheric CO₂ uptake by dinuclear copper²⁵ and zinc cryptands.²⁶ The uptake capability of these bimetallic cryptands and of dinuclear metalloenzymes as well as the high stability of the obtained carbonate complexes within those frameworks pointed our interest towards the application of dinickel azacryptands as highly selective CO₂

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†Electronic supplementary information (ESI) available: Synthesis and characterization of new compounds, X-ray crystallographic analysis, UV-Vis spectra, SQUID and kinetic data, detailed experimental conditions. CCDC 1418422–1418427. For ESI and crystallographic data in CIF or other electronic format see DOI: 10.1039/c5dt04267d



Scheme 1 Carbonate uptake by bimetallic complexes.

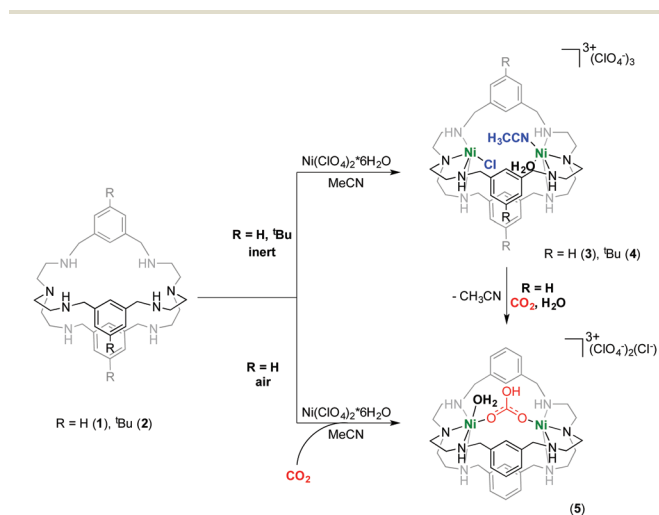
binding and transportation platform under ambient conditions.

Reaction of the azacryptands **1** or **2** with $\text{Ni}(\text{ClO}_4)_2$ comprising the non-coordinative perchlorate anion afforded complexes **3** and **4** under inert atmosphere as crystalline material in 56% and 52% yield, respectively (Scheme 2). The composition of both complexes was unequivocally confirmed by structure analysis (Fig. 1 and S1†). Both complexes **3** and **4** are isostructural and reveal, although possessing two identical metal coordination sites, two nickel centers with different coordination modes. One nickel atom is coordinated by four amines and an additional chloride ligand. The chlorido ligand herein most likely originates from the reduction of a perchlorate catalyzed by a Ni–O species,²⁷ and multiple synthetic attempts with independently generated starting materials gave similar results. In contrast, the second nickel center is coordinated by four amines as well as an additional acetonitrile and a water molecule. The Ni–O distances of 2.044 Å (in **3**) and 2.035 Å (in **4**) are in good agreement with literature data for nickel-bound water and are significantly longer than Ni–O distances observed for Ni–OH moieties.²⁸ Charge balance is provided by

three perchlorate anions. To exclude nickel oxidation due to reductive perchlorate decomposition within the complexes we performed SQUID measurements on **4**, which indicate two nickel centers with $S = 1$, consistent with two isolated Ni(II) centers (Fig. S2†).

Notably, whereas solutions of **4** were stable under atmospheric conditions, solutions of **3**, kept in acetonitrile in air rapidly changed color (Fig. 2). The resulting species was thus isolated and revealed a significantly altered IR spectrum (Fig. S3†). Crystals obtained by concentrating an acetonitrile solution of **3** under air revealed the molecular structure of the resulting compound **5** as depicted in Fig. 1. It is noteworthy, the cavity of the azacryptand is occupied by a bicarbonate bridging the two nickel-centers with a Ni–O_{CO₃} distance of 1.991 Å and 2.019 Å respectively. The Ni–Cl distance is significantly increased from 2.326 in **3** to 2.783 Å in **5** and the chloride is held in place by a hydrogen bond to the coordinated bicarbonate. As highlighted in Fig. 1, an internal hydrogen bond to an attached water molecule further stabilizes the bicarbonate ion between the two nickel(II) centers. Although a magnetic communication between both nickel centers can be expected, SQUID measurements unequivocally revealed two isolated nickel(II) centers in complex **5** with $S = 1$.

In order to prove that the formation of **5** is the result of a selective uptake of CO₂ from air, we performed binding experiments under ¹²CO₂ and ¹³CO₂ atmosphere. Treatment with either gas resulted in the formation of a bicarbonate-binding complex **5**. A comparison of both IR spectra did not allow for a clear assignment of the carbonate vibrations since the bicarbonate bands coincide with the azacryptand ligand vibrations. Mass spectrometry, however, confirmed the successful uptake of the respective isotopically-enriched substrates (Fig. S4†). Notably, no reaction of complex **3** with SO₂, NO, N₂O, NO₂, nor CO was observed by UV-vis spectroscopy.



Scheme 2 Synthesis and reactivity of dinickel azacryptands **3** and **4**.

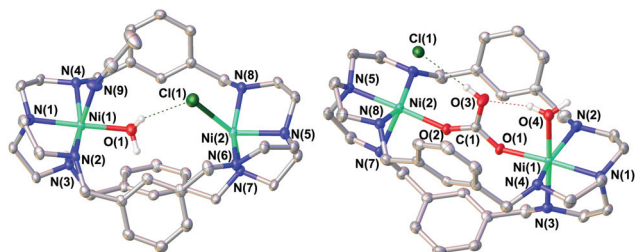


Fig. 1 Molecular structures of the apo-cryptand **3** (left) and its bicarbonate-binding equivalent **5** (right) (Counter ions and hydrogen atoms were omitted for clarity).

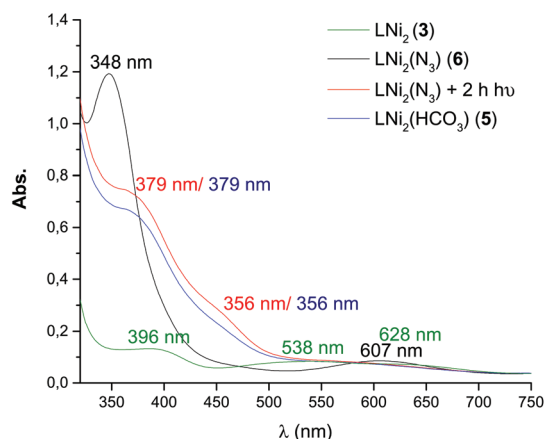


Fig. 2 Absorption spectra measured in acetonitrile at RT: green trace – complex **3** (LNi_2); blue trace – after bubbling CO₂ through the solution of **3** for 1 min ($\text{LNi}_2(\text{HCO}_3)$); black trace – after reaction of **6** ($\text{LNi}_2(\text{N}_3)$) with sodium azide for 6 h ($\text{LNi}_2(\text{N}_3)$); red trace – after UV-irradiation of **6** ($\text{LNi}_2(\text{N}_3)$) + 2 h $h\nu$).

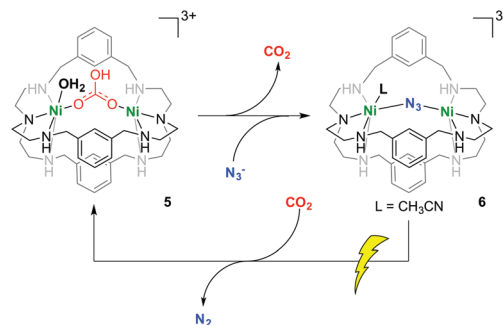


The kinetics of the CO₂ fixation process was furthermore followed by stopped-flow spectrophotometry at various temperatures and applying different concentrated CO₂ solutions in acetonitrile obtained from saturated stock solutions (Fig. S5†).²⁹ The reaction was monitored by the occurring absorbance changes at 470 nm. The reaction of **3** and CO₂ is of first-order in CO₂ concentration, yielding k_{obs} , and the apparent association rate constant $k_{2, 298 \text{ K}} = 0.067 \pm 0.005 \text{ M}^{-1} \text{ s}^{-1}$ is obtained from the slope of the plot of k_{obs} vs. CO₂ concentration (Fig. S5c†). Only a small increase of k_2 with increasing temperature is observed and an apparent enthalpy of the activation of $\Delta H^\ddagger = 7.55 \pm 1.07 \text{ kJ mol}^{-1}$ is obtained from the Eyring plot (Fig. S6†). This small value reflects that the association of CO₂ proceeds in more than one step. Therefore pre-equilibrium and activation enthalpies compose the apparent ΔH^\ddagger value.

Contrary to **3**, an uptake of CO₂ into complex **4** was not possible. This fact demonstrates the different uptake capacities of both species. As no significant structural differences are observed at the metal centers in **3** vs. **4**, we believe that the hydrophilic bicarbonate is repelled by the interaction with the hydrophobic *tert*-butyl group. Thus the CO₂ uptake can easily be adjusted by different bridging elements.

Attempts to release the bicarbonate by purging solutions of **5** with nitrogen gas or heating up the same solutions did not result in the re-generation of **3** (Fig. S7†). Azacryptands are well known to support the binding of different anions with varying binding strength.^{30,31} To reverse the strong CO₂ binding of **5**, we therefore added different anions to complex **5** to test whether we would be able to release the tightly-bound bicarbonate. Whilst addition of Cl[−] or Br[−], as well as SO₂, NO, N₂O, NO₂, CO, or CS₂ did not reveal any dissociation of the bicarbonate, attempts of reactivation through treatment with NaOH, HOAc, or HBF₄ resulted in precipitation of Ni(OH)₂, formation of [Ni(OAc)₂(H₂O)₄] (Fig. S8†),³² or other, unidentifiable, decomposition products. However, when N₃[−] or SCN[−] was allowed to react with **5**, the formation of a single well-defined product was observed *via* UV-vis spectroscopy (Fig. 2 and S9†). The identity of the N₃[−] reaction product was elucidated as compound **6** by single-crystal X-ray diffractometry and confirmed the exchange of the bicarbonate by an azide ligand (Scheme 3 and Fig. S10†). Furthermore, ESI-MS experiments as well as infrared spectra of this solution are in good agreement with a solution of **6**, independently generated by reaction of **1** with NaN₃ and Ni(ClO₄)₂. The release of CO₂ from complex **6** is slow ($k_{\text{obs}, 303.15 \text{ K}} = 0.00223 \text{ min}^{-1}$) (Fig. S11†) and reveals the high stability of compound **5**. Likewise, reaction of **5** and SCN[−] lead to exchange of HCO₃[−] by SCN[−] visible in ESI-MS. Notably, although **4** is not capable to incorporate the bicarbonate anion, [Ni₂(2)(N₃)](ClO₄)₃ **7** can be easily obtained by reaction of **4** with NaN₃ (Fig. S10†) and shows that the cavity is not intrinsically blocked by the *tert*-butyl groups.

We wondered about the general nature of the released “CO₂-species”. We therefore treated ¹³CO₂-**5** with the different substrates and investigated changes by ¹³C NMR spectroscopy, measured in an inert tube in deuterated acetonitrile, and using



Scheme 3 CO₂ binding and release utilizing dinickel azacryptands **5** and **6**.

gas chromatography. This treatment unequivocally showed the release of CO₂ as the only C1-species in all cases (Fig. S12†).

Similar reactivity was observed upon treatment of complex **5** with NaHCO₃ and provides evidence for a reversible CO₂ binding within the cryptand. Unfortunately, complex **6** did not allow for further ligand substitution reactions. It is, however, well known that Ni–N₃ complexes decompose under UV irradiation.³³ We therefore irradiated an acetonitrile solution of **6** with a 150 W Xe-lamp under air. This treatment revealed a change of the initial UV-vis bands and complete conversion was achieved within 2 hours (Fig. 2). Inspection of the UV-vis spectrum of the obtained compound showed close resemblance to the spectrum of complex **5** (Fig. 2).

Furthermore, the obtained IR and MS spectra are in good agreement with the formation of compound **5** through UV-irradiation of **6** (Fig. S13†). Although we are at this point not able to provide an exact mechanism for the uptake of CO₂, its expelling by N₃[−] and the subsequent cleavage of the resulting free azide, an overall pathway is presented in Scheme 3.

In conclusion, we have shown for the first time that nickel-azacryptands are superior compounds for highly selective CO₂ capture under ambient conditions. Due to their defined CO₂-hosting cavity, typical problems associated with gas selectivity and low binding strength in the presence of moisture can be avoided as compound **5** is thermodynamically stable in the presence of excess SO₂, NO, N₂O, NO₂, CO, or CS₂. As shown with complex **4**, the CO₂ binding affinity can be easily adjusted by manipulation of the bridging elements connecting both coordination cavities. Complex **5** showed high stability even at higher temperatures. We were able to show a selective release of CO₂ upon treatment with pseudo-halogenides. Especially interesting is the reaction of **5** with NaN₃ to give **6**. This reaction proved to be very useful since irradiation with a standard UV Xe-lamp emptied the cryptands' cavity and allowed for subsequent binding of CO₂.

The well-defined cavity of azacryptands provides a highly sophisticated system for carbon capture and sequestration in a reversible manner. In addition, it shows that azacryptands can serve as molecular reactors allowing specific reactions to proceed in an elaborate, protected environment. Further investigations concerning the mechanism of N₃[−] uptake,



photocleavage, as well as manipulation of the substitution kinetics and reduction experiments are currently under way.

Acknowledgements

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