Silver complexation by metallacryptates†

Kevin Lamberts,** Matteo Tegoni,a Xiang Jiang,c Hui-Zhong Kou**c and Ulli Englerta,a

We report the first complete characterization of metallacryptates encapsulating Ag(I) cations: carboxylato ligands derived from L-proline and L-alanine chelate and bridge six Cu(II) centres arranged in a slightly distorted octahedral fashion. Eight oxygen atoms of these ligands are disposed in square-prismatic geometry and coordinate the monovalent cation. Two alternative metallacryptates based on alanine have been identified which differ with respect to aggregation: a solid in which pairs of encapsulating sites are formed competes with an infinite chain of M(i) coordinating sites. In contrast, the individual encrypting moieties are arranged as overall neutral and isolated molecular species in the proline-based metallacryptate. This proline derivative can accommodate Ag(I) and Na(I) cations and form a solid solution. Susceptibility measurements confirm ferromagnetic interactions between the Cu(II) within the hexanuclear proline cryptate and thus underline the similarity between solids accommodating Na(I) and Ag(I). Spectroscopic results suggest that these metallacryptates hardly dissociate in methanol solution.

1. Introduction

Organic crown ethers have proven useful since their discovery in the 1960s.1 These usually selective ligands are applied as cytotoxic agents,2 as building units for supramolecular polymers,3 as templates for crystal engineering4 and as stabilizers for selective catalysis.5 J. M. Lehn expanded the realm of polydentate ligands into the third dimension:6,7 cryptands literally "bury" cations, thus providing even higher stability constants and more pronounced selectivity than 2D crown ethers. Such functionalities are not restricted to organic molecules but may also be exerted by coordination complexes. These so-called metallacrowns8–12 and metallacryptands3–15 represent less popular but highly versatile chemical answers to the requirements of polydentate ligands. They may not only coordinate cations but also encapsulate anions;16,17 sulfate selectivity has been shown for a dinuclear Cu derivative.18 Quite obviously, the metal-containing polydentate ligands may feature certain extra properties in comparison to their classical organic counterparts, such as metal-based redox activity and the structural manifold of coordination centers with a potentially elevated number of binding sites. Rare earth cations were trapped in a trigonal prismatic hexanuclear Cu(u) scaffold and their magnetic interactions were studied by Sopasis et al.19

A class of octahedral hexanuclear metallacryptates from Cu(u) perchlorate and ω-amino acids shows the capability of binding Na(I) in its cavities: an octahedral arrangement of copper cations, connected by chelating and bridging amino acids, is stabilized by crosslinking perchlorates and the incorporated host cation. Hu et al.20 first reported the synthesis of Na[Na⊂Cu2{Cu(Gly)2}4(ClO4)4(OH)]·H2O and Na⊂Cu2–{Cu(Pro)2}4(ClO4)4(OH)]·H2O. Later, Wang et al.21 published an isomorphous proline metallacryptate and a hydroxyproline metallacryptate and reported their magnetic properties. Further occurrences of this structural motif were encountered for L-alanine,22 L-threonine23 and 1-amino-1-cyclopropane.24 The prototypic organic cryptand [2.2.2] with a cavity size radius of 1.4 Å may coordinate a number of different cations25 with a preference for K+. Expanding from a 6 fold coordinating cavity to 8, Xiang et al.26 found that the hexanuclear L-threonine cryptate coordinated Na+ exclusively, even in the presence of other cations.

In this contribution, we report the first complete structural characterizations of such hexanuclear copper(n)-metallacryptates which incorporate silver cations (see Fig. 1 for their general architecture): [Ag⊂Cu2{Cu(Pro)2}4(ClO4)4] (1a) is closely related to the structural outcome of a competition experiment, the solid solution [M⊂Cu2{Cu(Pro)2}4(ClO4)4], M = Ag, Na(i) (1b). Our investigation does not only cover the structural aspects of these compounds but also extends to their magnetic properties and their stability in solution. We compare these mono-cryptands to two alanine derivatives.
(2a & 2b) in which dimers or polymers of coordination sites aggregate via longer Cu⋯O Jahn–Teller contacts. As for monovalent Ag(I), no metallacryptates have been described to date, but a flexible metallacrown surrounding this cation has been reported; no structural characterization for this complex has been provided.

2. Experimental

Chemicals were used as purchased without further purification: l-alanine (99.8%, Evonik), l-proline (>99%, Evonik), CuClO4·6H2O (98%, Sigma-Aldrich), AgClO4 (97%, Sigma-Aldrich), NaClO4·H2O (99%, Merck), urea (99.5%, Grüssing).

2.1. Synthesis

The metallacryptates have been synthesised by self assembly via diffusion of an antisolvent into a stoichiometric solution of the constituents of the target products. X-ray powder diffraction (Fig. 2) confirms phase purity of 1a and 1b. The experimental powder patterns were obtained on flat samples at room temperature whereas the simulations are based on single crystal data collected at 100 K. Most solids exhibit positive thermal expansion and therefore larger lattice parameters at higher temperature, corresponding to a shift towards smaller diffraction angles in the experimentally observed patterns. A quantitative determination of the Na content in 1b is not possible by X-ray powder diffraction, since the diffractogram hardly differs from that of the isomorphous compound 1a. However, the reflection width indicates a homogeneous distribution of Na and Ag within the bulk sample. The powder diffractogram of the sample containing 2a and 2b shows the concomitant presence of both products (visible in the range up to 10° 2θ) together with impurities of l-alanine and at least one unidentified byproduct.

2.1.1 Synthesis of Cu2(l-Pro)4(H2O)3. Cu2(l-Pro)4(H2O)3 was synthesised similar to literature:27 CuSO4·5H2O (10 mmol, 2.496 g, 1 eq.) was dissolved in 10 mL H2O and l-proline (20 mmol, 2.303 g, 2 eq.) as well as NaOH (20 mmol, 0.8 g, 2 eq.) was added to the solution. Addition of 70 mL of acetone yielded a blue precipitate after 10 minutes of vigorous stirring. Since sodium sulfate precipitated as well, the solid was filtered off and suspended in 200 mL EtOH. Sodium sulfate remained as a solid and was removed by filtration. The filtrate was evaporated with a rotary evaporator and the resulting blue oil dissolved in 10 mL of H2O. Addition of 70 mL of acetone yielded blue crystalline powder of [Cu2(l-Pro)4(H2O)3]·2H2O. Drying in a dessicator yields Cu2(l-Pro)4(H2O)3 according to microanalysis. Anal. Caled for Cu2(l-Pro)4(H2O)3: C: 36.6%, H: 5.8%, N: 8.6%, Found: C: 36.6%, H: 5.8%, N: 8.4%.

2.1.2. Synthesis of 1a. While 1a can be synthesised in a one step reaction, larger quantities of 1a could be obtained by dissolving (Cu2(l-Pro)4(H2O)3)0.5 (2 mmol, 674 mg, 4 eq.), CuClO4·6H2O (1 mmol, 371 mg, 2 eq.) and AgClO4 (0.5 mmol, 112 mg, 1 eq.) in 50 mL MeOH and precipitating the product...
by adding 100 mL Et₂O. The solid was recovered by filtration and dried in a desiccator.

Anal. Calcld. for 1a: C: 26.0%, H: 3.7%, N: 6.0%, Found: C: 24.8%, H: 3.8%, N: 5.5%. Single crystals could be grown from a smaller scale reaction in the same stoichiometry, followed by gas phase diffusion of Et₂O into the solution. Crystals grew after 1–2 days as large square bipyramids. The crystals of 1a are a solvate with the formula [Ag₂(Cu-L-Pro)₂(H₂O)]₄(ClO₄)₅·(MeOH)₀.₅₈·(H₂O)₀.₄₄. The same product was obtained by replacing the Cu₂(L-Pro)₄(H₂O)₃ building block by CuClO₄·6H₂O and L-proline and performing the reaction with stoichiometric amounts of urea.

2.1.3. Synthesis of 1b. Mixed crystals containing Ag and Na were obtained by dissolving [Cu₂(L-Pro)₄(H₂O)]₀.₅(2 mmol, 674 mg, 4 eq.), CuClO₄·6H₂O (1 mmol, 371 mg, 2 eq.), AgClO₄ (0.25 mmol, 56 mg, 0.5 eq.) and NaClO₄·H₂O (0.25 mmol, 31 mg, 0.5 eq.) in 50 mL MeOH and precipitating the product by adding 100 mL Et₂O. The solid was recovered by filtration and dried in a desiccator. Single crystals could be grown from a smaller scale reaction in the same stoichiometry, followed by gas phase diffusion of Et₂O into the solution. Crystals grew after 1 day as large square bipyramids and are a solvate with the formula [Ag₀.₅₆Na₀.₄₄⊂Cu₂{Cu(L-Pro)₂}₄(ClO₄)₅·(MeOH)₀.₅₈·(H₂O)₀.₄₄]. The crystals are solvates with the formula [Ag₀.₅₆Na₀.₄₄⊂Cu₂{Cu(L-Pro)₂}₄(ClO₄)₅·(MeOH)₀.₅₈·(H₂O)₀.₄₄]. The crystals of 1b were confirmed as equal to those encountered in 1a. Difference Fourier maps suggested an end-to-end disorder for the axially coordinated moieties, one anion and a solvent molecule. Under these conditions, the additional water/methanol disorder for the latter could not be handled as simply as in the case of 1a. Only the coordinated oxygen atom was therefore refined with full site occupancy, and the many alternative sites with low fractional site occupancy for methanol C and H atoms were not taken into account.

X-ray powder diffraction was performed at the Institute of Inorganic Chemistry, RWTH Aachen using a Stoe Stadi P diffractometer with Guinier-geometry (Cu-Kα, λ = 1.54059 Å, Johansson germanium monochromator and Stoe Imageplate detector IP-PSD, 0.005° stepwidth in 2θ).

Microanalysis was performed at the Institute of Organic Chemistry, RWTH Aachen with a CHN-O-Rapid VarioEL from Heraeus.

3.3. Solution studies

The samples were prepared using reagent-grade methanol or doubly-distilled water. Dilution titration experiments were carried out as follows. Mother solutions of I(Na) (ca. 5.4 mM) or 1a (ca. 5.0 mM) were prepared by weight in methanol or water. Proper volumes of the solutions were diluted 1:1 to obtain samples diluted of a factor of 2. Successive dilutions allowed to obtain samples with dilution factors of 4, 8, 16, 32, 64, and 128. Analogous samples of I(Na) were prepared by diluting the methanolic or aqueous mother solution of the compound with a 0.1 M NaClO₄ solution in methanol or water, respectively. The most diluted solutions of these sample set had concentrations in the 32.6–40.2 μM range. Visible spectra of samples of Cu₂(L-Pro)₄(H₂O)₃ (ca. 6 mM) and CuClO₄₂·6H₂O (ca. 50 mM) prepared by weight in methanol were used to calculate the molar absorbance of the compounds in this solvent. Visible spectra of the samples were collected in the 400–900 nm interval using a Thermo Evolution Bio 260 UV-visible spectrophotometer provided with a Peltier device. Quartz cuvettes of 0.1, 1 and 5 cm path length were used.

ESI mass spectra were recorded using a WatersAcquity SQ Detector with ESI interface and a direct infusion device. Data were processed by using the spectrometer software (MassLinx 4.1). The measurements were performed on solutions of I(Na) or 1a in methanol, water, or methanol:water mixtures 9:1 (v:v). The sample concentration was ca. 10⁻⁴ M. Direct infusion analyses were always performed at 10 μL min⁻¹. Experimental conditions were as follows: ES capillary 3.0 kV; cone 30–80 V; extractor 4 V; source block temperature 80 °C; desolvation temperature 150 °C; cone and desolvation gas (N₂)
### Table 1  Crystallographic details of the structures \[\text{M} \subset \text{Cu}_2\text{Cu}(\text{Cu-amino-carboxylate})_2\text{Cl}_4\text{(ClO}_4\text{)}_5\text{solvent}\]

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<th>Compound/CCDC#</th>
<th>1a/1048757</th>
<th>1b/1048758</th>
<th>2a/1427077</th>
<th>2b/1427078</th>
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<td>Crystal data</td>
<td></td>
<td></td>
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<td>Chemical formula</td>
<td>(\text{C}<em>{40.58}\text{H}</em>{68}\text{N}_8\text{AgCl}_5\text{Cu}<em>6\text{O}</em>{37.42})</td>
<td>(\text{C}<em>{40.58}\text{H}</em>{68}\text{N}_8\text{AgCl}_5\text{Cu}_6\text{N}<em>4\text{O}</em>{41.17})</td>
<td>(\text{C}<em>{62}\text{H}</em>{72}\text{N}_8\text{AgCl}_6\text{Cu}<em>6\text{O}</em>{53})</td>
<td>(\text{C}<em>{62}\text{H}</em>{72}\text{N}_8\text{AgCl}_6\text{Cu}_6\text{N}<em>4\text{O}</em>{53})</td>
</tr>
<tr>
<td>M, amino-carboxylate</td>
<td>Ag(l), (\text{l-Pro})</td>
<td>Ag(l), (\text{l-Pro})</td>
<td>Ag(l), (\text{l-Ala})</td>
<td>Ag(l), (\text{l-Ala})</td>
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<tr>
<td>Solvent</td>
<td>0.58 MeOH 0.84 H(_2)O</td>
<td>1.6 and 8 L min(^{-1}), respectively. Scanning was performed at m/z 100–2000.</td>
<td>0.58 MeOH 1.5 H(_2)O</td>
<td>0.58 MeOH 1.5 H(_2)O</td>
</tr>
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<td>Crystal system, space group</td>
<td>Tetragonal, (\text{I}_{4})</td>
<td>Tetragonal, (\text{I}_{4})</td>
<td>Triclinic, (\text{P})</td>
<td>Monoclinic, (\text{P}_2_1)</td>
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<td>(a) (Å)</td>
<td>13.512(3)</td>
<td>13.4554(16)</td>
<td>12.7360(7)</td>
<td>9.8711(12)</td>
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<td>(b) (Å)</td>
<td>18.097(6)</td>
<td>17.931(2)</td>
<td>19.6150(12)</td>
<td>5.596(3)</td>
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<td>(c) (Å)</td>
<td>12.8490(8)</td>
<td>24.766(3)</td>
<td>13.512(3)</td>
<td>13.4554(16)</td>
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<td>(\beta) (°)</td>
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<td>100.770(2)</td>
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<td>101.270(2)</td>
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<td>(\gamma) (°)</td>
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<td>(V) (Å(^3))</td>
<td>3304.2(18)</td>
<td>3246.4(9)</td>
<td>3050.4(3)</td>
<td>6147.2(13)</td>
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<td>(\mu) (mm(^{-1}))</td>
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<td>2.41</td>
<td>2.70</td>
<td>2.67</td>
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<td>0.14 × 0.09 × 0.07</td>
<td>0.16 × 0.14 × 0.08</td>
<td>0.22 × 0.14 × 0.04</td>
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<td>Data collection</td>
<td>(I &gt; 2\sigma(I)) reflections</td>
<td>0.533, 0.745</td>
<td>0.609, 0.745</td>
<td>0.648, 0.746</td>
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<td>(\sin \theta/\lambda) (Å(^{-1}))</td>
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<td>0.611</td>
<td>0.714</td>
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<tr>
<td>(R_{int})</td>
<td>0.083</td>
<td>0.068</td>
<td>0.035</td>
<td>0.080</td>
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<td>Refinement</td>
<td>(R[F^2 &gt; 2\sigma(F^2)], wR[F^2], S)</td>
<td>0.054, 0.150, 1.04</td>
<td>0.042, 0.115, 1.04</td>
<td>0.055, 0.135, 1.03</td>
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<td>No. of reflections</td>
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<td>3116</td>
<td>34212, 27196</td>
<td>47383, 34212, 27196</td>
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<td>241</td>
<td>47383, 34212, 27196</td>
<td>93433, 35809, 22713</td>
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<tr>
<td>No. of restraints</td>
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<td>149</td>
<td>1516</td>
<td>1504</td>
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<tr>
<td>(\Delta\alpha_{max} - \Delta\alpha_{min}) (e Å(^{-3}))</td>
<td>0.93, −0.89</td>
<td>1.20, −0.42</td>
<td>3.02, −3.12</td>
<td>1.67, −1.92</td>
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<td>Flack parameter</td>
<td>0.04(2)</td>
<td>−0.019(14)</td>
<td>0.017(9)</td>
<td>0.074(7)</td>
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</table>

1.6 and 8 L min\(^{-1}\), respectively. Scanning was performed at m/z 100–2000.

The distribution diagrams of the complex species for the Cu\(^{2+}\)/proline system was studied with the Hyss 2006 program.\(^{32}\)

### 3. Results and discussion

Four new metallacryptates are presented. They share a very similar architecture with respect to primary cation coordination: four \(\text{cis}\)-\(\text{bis}\)-\(\text{l}\)-aminocarboxylato-copper(II) units are connected \(\text{via}\) their \(\text{exo}\)-carboxylato-O atoms towards two copper(II) cations.

An octahedral arrangement of 6 copper cations is formed, leaving a coordination pocket in the middle that offers an eightfold binding site by the \(\text{endo}\)-carboxylato-O atoms. This site is occupied by Ag(II) (or statistically occupied by Ag(I) and Na(I) in the case of the solid solution 1b). Four perchlorate anions bridge the equatorial Cu-units so that an overall monoplatonic metallacryptate motif is formed.

The metallacryptates 1a-2b differ with respect to the arrangement of the primary cation coordination sites. Structures 1a and 1b are synthesised from \(\text{l}\)-proline. 1a is iso- morphous to its Na-analogue\(^{31}\) (here referred to as 1(Na)) and very closely related to the solid solution of a mixture of both (1b). Since these compounds are monomeric structures (axial copper cations are coordinated by solvent or the remaining perchlorate anion, see Fig. 3) 1a was suited for investigations on its magnetic properties and solution behaviour. Moreover, studies on the stability of 1(Na) and 1a in aqueous and methanolic solutions were carried out.

Structures 2a and 2b, synthesised from \(\text{l}\)-alanine, remarkably are not isomorphous to their related Na compounds.\(^{32}\) Both exhibit the aforementioned general architecture. 2a however, is a dimer of two connected metallacryptate units, 2b is a polymer with two crystallographically independent strands.

Structural key points and distinctive features will be discussed for all compounds. The magnetic properties and solution behaviour of 1a will be presented separately.

#### 3.1. Structural properties of 1a and 1b

The metallacryptate in 1a consists of four neutral \(\text{cis}\)-\(\text{bis}\)-\(\text{l}\)-prolinato-copper(II) units. They connect the two axial Cu(II) cations and the central Ag(I), thus resulting in a formally pentacatonic aggregate. The charge is compensated by one perchlorate anion, see Fig. 3) 1a was suited for investigations on its magnetic properties and solution behaviour. Moreover, studies on the stability of 1(Na) and 1a in aqueous and methanolic solutions were carried out.

1a forms as a phase pure solid in almost quantitative yield by self assembly of its constituents in the stoichiometry required for the formation of the target product: 1a can be obtained by evaporation of methanolic solution of the reactants Cu(ClO\(_4\))\(_2\), \(\text{l}\)-proline and AgClO\(_4\) in a 6 : 8 : 1 ratio.
The encrypted metal is coordinated by one oxygen atom of each proline and thus enclosed in a slightly distorted square prismatic coordination sphere, with a torsion angle between the top and bottom square planes of the oxygen prism of ca. 10°. The six Cu(n) cations, two associated with the base planes and four with the lateral edges of the prism, adopt a slightly clinched octahedral geometry (Cu⋯Cuaxial = 6.960(10) Å, Cu⋯Cuaxial = 7.131(3) Å) about the central Ag(i). A more detailed inspection of the coordinative bonds in the cation reveals that axial and equatorial Cu(n) differ: the former are in a typical Jahn–Teller type square pyramidal coordination, with short Cu–O bonds (ca. 2 Å) to the amino acid ligands and significantly longer contacts to either a perchlorate counter anion (ca. 2.5 Å) or a coordinated solvent molecule (ca. 2.3 Å). The cations in the equatorial cis-bis-l-prolinato-copper(n) units are O,N chelated by short bonds to the amino acid molecules. In addition to this square planar coordination, they interact with much more distant (2.7–2.9 Å) bridging perchlorate anions. Selected bond distances for all structures are given in Table 3. Precise geometry data and their standard uncertainties have been compiled in the ESI.† The pentacationic complex and the bridging or terminal perchlorate anions aggregate via the above-mentioned longer Cu⋯O distances, and hence 1a can, despite its formal composition as a salt, be described as a solid built from discrete “molecules”. Fig. 4 shows two projections of a [Ag ⊘ Cu2(Pro)2]5[(ClO4)6 aggregate in the [100] and [001] view directions.

A view along [100] also illustrates the arrangement of the aggregates in the body-centered unit cell (Fig. 5). The axial residues are disordered about the crystallographic fourfold axis and fill the cavities A (terminal perchlorate) and B (methanol/water solvent molecule).

The framework around the silver cation can be seen as [2.2.2.2.2] metallacryptate. The diagonal length of the square prismatic cage amounts to ca. 5.1 Å; when a certain degree of flexibility is taken into account, the coordination cage may enclose a range of cations. Selectivity of cryptands does, however, not only depend on radii criteria but also on additional parameters such as Pearson hardness33 or the bite angle of the ligands. The cation selectivity in metallacryptates has been addressed by Wang et al.21 These authors shortly mention the selectivity of their l-proline cryptate towards Na(i) in presence of Li(i) or K(i) and add that an analogous reaction with Ag(i) is possible. Additionally, they comment on a possible existence of metallacryptates in solution. However, no sophisticated evidence is given.

For the case of our Cu2(Cu(Pro)2)5[(ClO4)6] metallacryptate, we have obtained structural evidence for the inclusion of both silver and sodium cations under competitive conditions. When the templated assembly of the cryptate is conducted as above but from an equimolar mixture of AgClO4 and NaClO4, a solid solution is obtained in which both monovalent cations are incorporated in roughly 1 : 1 ratio. This solid 1b has also been structurally characterized. The most relevant result concerns selectivity: The significantly different electron density of both alternative M(i) cations allows to refine their occupancy with good accuracy. 1b contains 56.3(6)% Ag versus 43.7(6)% Na,
indicating a slight preference for silver coordination. In X-ray powder diffraction, reflections of 1b are of similar width as those of 1a, thus indicating a homogenous solid solution.

The most prominent effects of partial Na substitution are observed in the geometry of the coordination cage and changes in the geometry of the Cu octahedron. For eightfold coordination the crystal radius of Na(i) is 0.1 Å shorter than that of Ag(i).34 As a consequence, the M(i)-O distances shorten (M = O = 2.553(6) and 2.544(7) Å) and the overall geometry of the cryptate adopts a more symmetrical arrangement: the differences in length between the edges within the Cu6 octahedron are reduced by an order of magnitude. Table 2 summarizes the structural differences between 1a, 1b and the Na(i) analogue 1(Na)22 (measured at room temperature) and the two polymorphs from alanine 2a and 2b.

### Table 2

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<thead>
<tr>
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<th>1a</th>
<th>1b</th>
<th>1(Na)22</th>
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<tr>
<td>σoCu</td>
<td>0.262</td>
<td>0.057</td>
<td>0.029</td>
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<td>ΔoCu × 10^4 (Cu)</td>
<td>0.155</td>
<td>0.047</td>
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<tr>
<td>Cu1...Cu3 [Å]</td>
<td>6.960</td>
<td>6.972</td>
<td>7.011</td>
</tr>
<tr>
<td>Cu2...Cu'2 [Å]</td>
<td>7.131</td>
<td>7.031</td>
<td>7.058</td>
</tr>
<tr>
<td></td>
<td>2a</td>
<td>2b</td>
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<tr>
<td>σoCu</td>
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<td>5.232, 3.701</td>
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<tr>
<td>ΔoCu × 10^4 (Cu)</td>
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<td>Cu1_6...Cu1_8 [Å]</td>
<td>7.107(2)</td>
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<tr>
<td>σoCu = 1/11 \sum_{i=1}^{12} (θ_{i} - 90°)^2; ΔoCu = 6/11 \sum_{i=1}^{6}</td>
<td>I_{i} - \bar{I}</td>
<td>^2;</td>
<td>I_{i}</td>
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Symmetry code: 'y, 1 - x, z'ii; '1 - x, 1 - y, z'ii; '1 + x, y, z'.

### Table 3

<table>
<thead>
<tr>
<th></th>
<th>1a [Å]</th>
<th>1b [Å]</th>
<th>2a [Å]</th>
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<tbody>
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<td>2.553(7)</td>
<td>2.544(7)</td>
</tr>
<tr>
<td>Cu1-O9</td>
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<td>2.45(2)</td>
<td>2.380(17)</td>
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<tr>
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<td>2.709(12)</td>
<td>2.703(7)</td>
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<tr>
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<td>2.879(15)</td>
<td>2.738(8)</td>
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<tr>
<td>Cu3-O13</td>
<td>2.223(19)</td>
<td>2.30(2)</td>
<td>2.30(2)</td>
</tr>
</tbody>
</table>

### 3.2. Magnetic properties of 1a

The presence of magnetically active Cu(i) cations within the metallacryptates suggests that susceptibility measurements can contribute additional information to a comparison between closely related Na(i) and Ag(i) coordinating species. The former information has been provided by Wang et al.21
and later summarised for other Na(i) containing cryptates by Xiang et al.23

Temperature-dependent magnetic susceptibility measurement for 1a was carried out on a Quantum Design SQUID magnetometer. The experimental susceptibility was corrected for the diamagnetism of the constituent atoms (Pascal’s Tables). Temperature-dependent magnetic susceptibilities of 1a were measured in the temperature range of 2–300 K under an applied magnetic field of 2000 Oe.

As shown in Fig. 6, the $\chi_m T$ value at 300 K is 2.5 cm$^3$ mol$^{-1}$ K, which is slightly larger than the theoretical value of 2.25 cm$^3$ mol$^{-1}$ K for six spin-only Cu(II) ions ($S = 1/2, g = 2$). Furthermore, the $\chi_m T$ value increases upon temperature cooling, indicating the presence of overall ferromagnetic interaction between adjacent Cu(II) centers. The $\chi_m$ versus T data above 25 K were fitted by the Curie-Weiss law, giving a Curie constant of 2.52 cm$^3$ mol$^{-1}$ K and a Weiss constant of 1.55 K, which also evidences the presence of ferromagnetic interaction within the Cu$_6$. In this octahedral Cu$_6$ cluster, the magnetic coupling between the equatorial Cu(II) atoms and axial Cu(II) atoms via the syn–anti carboxylate groups dominates and can be reasonably considered as identical. The magnetic coupling between the equatorial Cu(II) atoms via the central diamagnetic Ag$^+$/pathway is negligible. Thus, the exchange coupling constant $J$ is evaluated to be $+0.77(1)$ cm$^{-1}$ with a $g$ value of 2.12(1) by using the equation reported before.23 The positive $J$ value further confirms the ferromagnetic interaction within Cu$_6$ cluster.

For each Cu(n) ion with a distorted square pyramidal or elongated octahedral configuration, the unpaired electrons occupy the magnetic orbital $d_{x^2-y^2}$. For a syn–anti bridging mode of carboxylate, the 2p orbitals of two oxygen atoms respectively linked to the two Cu(II) ions are oriented to unfavourably give a strong orbital overlap for antiferromagnetic coupling. Therefore, most syn–anti carboxylate-bridged Cu(n) complexes display ferromagnetic interaction.23,35–44

For complex 1a, the adjacent $d_{x^2-y^2}$ orbitals from equatorial and axial position of the octahedral Cu$_6$ cluster are nearly perpendicular to each other with the dihedral angle between the N$_2$O$_2$ and O$_4$ coordination planes of 75.86°. According to the strict orthogonality of the magnetic orbitals of Cu(II) ions, such perpendicular arrangement of magnetic orbitals of adjacent Cu(n) atoms would lead to the overall ferromagnetic interaction.37 Moreover, the nonplanarity of the Cu-O–C–O–Cu linkages further weakens the orbital overlap to invalidate the antiferromagnetic interaction,41 which contributes to the overall ferromagnetic coupling in complex 1a. DFT calculation results also corroborate the ferromagnetic nature of Cu(n)–Cu(n) magnetic coupling via the syn–anti carboxylate bridge.44

3.3. Results of solution studies of 1a, 1b and 1(Na)

In contrast to organic cryptands, the metallacryptates are formed by self assembly of linkers and labile coordination centers; the question arises to which extend this process will be reversible. Do the complex architectures exist in solution or rather undergo disassembly to generate the Cu-prolinate and Cu(n) ions?

We decided to get insights on the integrity in solution of the metallacryptates containing Na$^+$ (1(Na)) and Ag$^+$ (1a). We therefore started with collecting the ESI-MS spectra of compounds 1(Na) and 1a in methanol (f$_{1(Na)}$,1a ca. 100 μM) by direct infusion of the solution. Signals corresponding to copper complexes were observed in positive-ion mode. In the 1000–2000 m/z range the spectrum of 1(Na) presents only one significant multiplet at $m/z = 1715$ (base peak) consistent with the metallacryptate species [NaCu$_2$[Cu(Pro)$_2$]$_4$ClO$_4$]$_n$ (Fig. 7 and S4†). Surprisingly, the same peak was observed also in the spectrum of 1a together with peaks at $m/z = 1037, 1315$ (base peak) and 1733 (Fig. 8 and S4†).

These latter three peaks, less intense, are present also in the spectrum of 1(Na). Possibly more surprisingly, none of the
peak in the spectrum of 1a is associated to Ag⁺ species, but rather to the aquated Na⁺ metallacryptate (m/z = 1733, NaCu₂[Cu(Pro)]₄(ClO₄)₄(H₂O)) or to a fragment of the metallacryptate construct (m/z = 1037, Cu(a)[Cu(Pro)]₂(ClO₄)⁴). The base peak at m/z = 1315 has been interpreted as a fragment of the metallacryptate interacting with one sodium and one potassium ion ([NaKCu[Cu(Pro)]₃(ClO₄)₄(H₂O)]⁺, the latter being possibly an impurity in the solution. These results show that the metallacryptate scaffold does not result into a complete disaggregation in pure methanolic solution, although they also suggest that the Ag-containing compound 1a is less stable than its sodium analogue 1(Na) under ionization conditions. Importantly, no signals above m/z = 1000 were observed for solutions of 1(Na) or 1a in water (Fig. S5†), while spectra similar to those obtained in pure methanol were observed for samples prepared in methanol/water 9:1 (v/v) mixtures (Fig. S6 and S7†). These data confirm the metallacryptates are completely disaggregated in aqueous samples, likely as the result of favourable solvation effects of the components of the construct. However, our data suggest that the compounds do not undergo a complete dissociation when water in methanol amounts to up to 10% in volume.

If we consider Cu²⁺ and [Cu(Pro)]₂ as fragments (components) of the metallacryptate scaffold, then we can describe the assembly process of 1(Na) and 1a through the equilibrium

\[ Na^+(Ag^+) + 2Cu^{2+} + 4[Cu(Pro)]_2 + 4ClO_4^- = [Na(Ag)Cu_2[Cu(Pro)]_4(ClO_4)_4] \]

The related stoichiometric equilibrium constant \( K = [\text{metallacryptate}]^{[\text{Na}]} \cdot [\text{Cu}^{2+}]^{[\text{Cu}(\text{Pro})_2]}^{[\text{ClO}_4^-]} \) has the dimension of a M⁻⁸. Therefore this equilibrium is concentration dependent, and favoured at higher concentrations.

Among these components, [Cu(Pro)]₂ and Cu²⁺ (as Cu(ClO₄)₂) have a \( \lambda_{\text{max}} \) of absorption in the visible range of 606 and 829 nm in methanol, respectively (\( \lambda_{\text{max}} = 619 \) and 810 nm in water, Fig. S8 and S9†). The molar absorption of [Cu(Pro)]₂ is significantly higher than that of the Cu²⁺ ion in both methanol (\( \varepsilon_{\text{max}} = 68 \) and 11 M⁻¹ cm⁻¹, respectively) and in water (\( \varepsilon_{\text{max}} = 51 \) and 12 M⁻¹ cm⁻¹, respectively). On the other hand, the metallacryptates 1(Na) and 1a absorb at \( \lambda_{\text{max}} \) ca. 660 nm in methanol, respectively (\( \varepsilon = \text{ca.} 350 \) M⁻¹ cm⁻¹, Fig. S8†), while in water their absorption is at slightly lower wavelengths (\( \lambda_{\text{max}} = 651 \) and 652 nm, respectively, \( \varepsilon = \text{ca.} 250 \) M⁻¹ cm⁻¹, Fig. S9†).

In both solvents, the absorption maxima of 1(Na) and 1a are intermediate between those of [Cu(Pro)]₂ and Cu²⁺.

With these data in our hands, we decided to study the disaggregation processes of the metallacryptates in solution through dilution experiments monitored by visible absorption spectrophotometry. In the hypothesis that disaggregation of the metallacryptate construct occurs upon dilution, then a blue shift from \( \lambda_{\text{max}} \) ca. 655 toward 610 nm should be observed, as the consequence of the appearance of [Cu(Pro)]₂ which is largely more absorbing than Cu²⁺. On the contrary, if the metallacryptates are already fully dissociated even in the most concentrated solution, then the \( \lambda_{\text{max}} \) should not vary significantly upon dilution.

Indeed, by dilution of solutions of both 1(Na) and 1a (ca. 5 mM to 33–40 μM; 128-fold dilution), we observed a significant 18–19 nm blue shift of the \( \lambda_{\text{max}} \) (Fig. 9 and S10†). On the contrary, when the same experiments were performed in water the \( \lambda_{\text{max}} \) remained constant up to 8-fold dilution, moving toward higher wavelengths (red shift) at lower concentrations (Fig. 10 and S11†).

The blue shift observed in methanol, and not observed for aqueous solutions, indicates that a progressive dissociation of 1(Na) and 1a is obtained in this medium upon dilution. This result is consistent with ESI MS data which suggested the presence of undissociated metallacryptate at \( 10^{-4} \) M concentration. On the other hand, the red shift observed for higher dilutions in water could be very well interpreted taking into account the dissociation of [Cu(Pro)]₂ which occurs at lower concentrations following the \([\text{Cu}(\text{Pro})_2] = [\text{Cu}(\text{Pro})]^+ + \text{Pro}^-\) equilibrium. Because the formation constants of copper(n) and proline complexes in water are known⁴⁵–⁴⁷ we could quantitatively calculate the concentration of the species at the equilibrium. The speciation of the system is reported in Fig. 11A and it refers to a Cu²⁺/proline = 6:8 corresponding to a fully dissociated metallacryptate in unbuffered solution. In this system the most
abundant species is the $[\text{Cu(Pro)}]^+$ complex (ca. 58%). As $[\text{Cu(Pro)}]^2$ absorbs at 610 nm while $[\text{Cu(Pro)}]^+$ at ca. 710 nm, the mixture of the two species should result into an intermediate $\lambda_{\text{max}}$ consistent with the observed 655 nm in water. The concentration of $[\text{Cu(Pro)}]^2$ remains constant up 8-fold dilution, and then it starts to decrease with a concomitant increase of the amount of $[\text{Cu(Pro)}]^+$. This results into the red shift, as represented in Fig. 11B: the trend in shift as higher wavelengths overlaps very well with the increase in concentration of $[\text{Cu(Pro)}]^+$.

Very interestingly, similar experiments were carried out also on $1(\text{Na})$ diluted with either 0.1 M methanolic or aqueous NaClO$_4$ solution (Fig. S12 and S13,$^\dagger$ experiments using $1\text{a}$ resulted into precipitation issues). The results are similar to those obtained with pure solvents, but the presence of sodium perchlorate in methanol the blue shift amounts to only 11 nm. Although from this experiments we cannot discriminate the effect of the sodium and perchlorate ions separately, the results suggest that the presence of a high concentration of these ions prevents, as expected, the dissociation of the metallacryptate. On the contrary, the high concentration of sodium perchlorate has no effects in aqueous solutions where the metallacryptate $1(\text{Na})$ remains fully dissociated.

Overall, our results suggest that the dissociation of $1(\text{Na})$ and $1\text{a}$ is complete in water regardless of the total metallacryptate concentration and the presence of perchlorate ions in the medium. On the contrary, in methanol the dissociation occurs at a minor extent. The absence of polymetallic species in aqueous solutions is fully consistent with the speciation of the Cu$^{2+}$/amino acid (L) systems in aqueous solution: where the side chain of the amino acid is non-coordinating, $[\text{Cu(L)}]^+$ and $[\text{Cu(L)}]^2$ are the major species with no reported formation of polymeric complexes.$^{45-47}$ Moreover, our results explain how polymetallic species containing amino acids can be conveniently crystallized from methanolic solutions. Finally, as the behaviour of $1(\text{Na})$ and $1\text{a}$ in the dilution experiments is not significantly different, we should conclude that the stability of the two proline metallacryptates is not very different.

### 3.4. Structural properties of $2\text{a}$ and $2\text{b}$

The coordinating amino acid (l-proline in the case of $1\text{a}$ and $1\text{b}$) plays an important role for the formation of metallacryptates. Ag($\text{i}$) coordination can also be achieved with l-alanine: in this case, the concomitant products $2\text{a}$ and $2\text{b}$ have been obtained. With respect to primary cation coordination, they closely correspond to the proline derivatives: four cis-bis-alanine-copper($\text{ii}$) connect two axial Cu($\text{ii}$)-cations and enclose an Ag($\text{i}$) in their centre. However, the alanine-based cryptates are considerably more distorted from an ideal fourfold symmetry, and additional coordinative bonds link primary metallacryptate sites to overall more complex solids.

$2\text{a}$ crystallises in P1 as a dimer. The two cryptate molecules are connected via two Cu–O contacts of axial copper atoms (Cu–O ca. 2.39 Å). The alternative product $2\text{b}$ on the other hand crystallises in P2$_1$ with a slightly longer connection between metallacryptate subunits (Cu–O ca. 2.45 Å). However, the two symmetry independent molecules are linked on both sides towards its neighbour and hence build two independent polymer chains extending along $b$ (Fig. 12).

While the connection of the metallacryptate molecules to more extended aggregates is the most distinctive difference between the products from proline and alanine, the local symmetry of the hexanuclear units of $2\text{a}$ and $2\text{b}$ is also significantly different. With all atoms on general positions the $C_{\text{i}}$ (or approximate $D_3$) symmetry is broken. This is reflected in the coordination environment around the silver cation: Ag–O distances range from ca. 2.45 to 2.77 Å in $2\text{a}$ and 2.54 to 2.71 Å in $2\text{b}$. The base planes of the distorted square prismatic coordination polyhedron are almost not tilted towards each other although this geometry is hard to parametrise since all Ag–O distances are independent. The overall octahedral arrangement of the copper cations is also much more variable (see...
Table 2). While Cu⋯Ag⋯Cu angles have a much broader distribution than in the metallacryptates from proline, their Cu⋯Ag distance distribution is comparable. Distances across the octahedron are around 7.1 Å in 2a. In 2b they are shorter in chain direction (7.04 Å) than perpendicular to the chain (7.18 Å). Cu⋯Cu distances along the edges range from 4.657(2) to 5.338(2) Å in 2a and from 4.89(1) to 5.30(1) Å in 2b.

Most notably, the bridging of the cis-bis-alaninato-copper(II) units with perchlorate anions shows a larger variety: in both structures O–Cl–O bridging similar to that in 1a or 1b can be observed. Additionally, direct bridging with only one oxygen atom of the perchlorate anion occurs. Each compound features one symmetrically independent metallacryptate site in which only one pair of neighbouring Cu(II) centers is directly bridged (Fig. 13, right) and a second residue in which two of these Cu–O–Cu connections occur (Fig. 13, left).

While we showed that for the case of proline both cations, Na(I) and Ag(I), yield isomorphous or closely related molecular solids, the structures from alanine show polymorphism in the case of both cations and all structures are actually quite different: three structures are known for the analogue Na(I) chemistry. While one of the structures is not a metallacryptate with a full shell as presented here, the other two polymorphs are a molecular structure similar to 1a and 1b, and a dimeric structure comparable to 2a but with additional NaClO₄ bridging. To our knowledge 2b is therefore the first axially connected polymer of hexanuclear metallacryptates of its kind. However, a polymeric connection has been reported for a metallacryptate from hydroxyproline, though its connection is via the peripheral hydroxy groups.

4. Conclusions

What can be learned from Ag(I) coordination by metallacryptates, and perhaps more general, how do metallacryptates compare to their organic counterparts?

The proline-based metallacryptands 1a and 1b suggest that Na(I) and Ag(I) complexation will not lead to very different results: both cations may form a solid solution with the same metallacryptand, and the pure Ag(I) and Na(I) species are closely related solids with similar magnetic properties and similar stabilities in solution. One might be tempted to consider these results for the proline derivatives as predictable, but the corresponding alanine compounds convey a very different message. Ag(I) coordination by alanine-based metallacryptates is entirely different from Na(I) complexation. Three different Na(I) cryptates form based on the applied stoichiometry. One adopts a structure analogous to 1a; the second arranges in a dimeric fashion like 2a but contains additional NaClO₄ connecting the molecules. The third structure is a partial metallacryptate that is connected to an infinite polymer chain. Here, one bis-alaninato-copper unit is missing for the full hexanuclear cryptate arrangement.

Metallacryptates and organic cryptands differ in a very obvious aspect: the organic multidentate ligands might adopt alternative conformations when non-coordinating but they will surely not completely dissociate and be present as possible major constituent in solutions of specific solvents such as methanol. The characterization of metallacryptates to date mostly relies on solid state methods, but the results concerning their existence in solution, never reported previously, are encouraging. In our future work we will attempt to further bridge the gap between solid state characterization and
dynamic behaviour in solution and gain an overall better understanding of the versatile class of metallacryptates.

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References

47. R. M. Smith, A. E. Martell and R. J. Motekaitis, NIST Critically Selected Stability Constants of Metal Complexes, Database 46, 7.0, Gaithersburg, MD, USA, 2003.