Salt metathesis versus protonolysis routes for the synthesis of silylamide Hauser base (R$_2$NMgX; X = halogen) and amido-Grignard (R$_2$NMgR) complexes†

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The preparation of silylamide Hauser base (R$_2$NMgX, X = halide) and amido-Grignard (R$_2$NMgR) complexes from simple Grignard reagents using [K(N(SiMe$_2$)$_2$Bu)$_2$], [K(N(SiMe$_2$)$_2$Bu)[Si(Pr)$_3$]], and [K(N(SiPr$_3$)$_2$)]$_n$ and their parent silylamines, was explored. Both salt metathesis and protonolysis routes proved ineffective with allylmagnesium chloride as a starting material due to complex Schlenk equilibria, with [Mg(N(SiMe$_2$)$_2$Bu)- (THF)]$_2$ (N$_{RR}^-$ = {N(SiMe$_2$)$_2$Bu})$^{-2}$, 1; {N(SiBuMe$_2$)[Si(Pr)$_3$]}$^{-2}$, 2; {N(SiPr$_3$)$_2$}$^{-2}$, 3 and [Mg(N(SiPr$_3$)$_2$)[μ-C$_2$H$_4$]$_n$]$_4$ (4) identified as minor products. In contrast, salt metathesis protocols using potassium silylamides and methylmagnesium iodide gave [Mg(N$_{RR}^-$)[μ-CH$_3$]$_2$ (N$_{RR}^-$ = {N(SiMe$_2$)$_2$Bu})$^{-2}$, 7a; {N(SiBuMe$_2$)[Si(Pr)$_3$]}$^{-2}$, 8; (N/SiPr$_3$)$_2^{-2}$, 9 and [Mg(N(SiBuMe$_2$)$_2$)[CH$_3$](DME)]$_{7b}$ with [Mg(N(SiBuMe$_2$)$_2$)[μ-li(THF)]$_2$] (10) isolated as a side-product during the preparation of 7a. Unusually, methylmagnesium iodide, di-n-butylmagnesium and 7–9 did not react with HN$^{15}$ under the conditions we employed. The synthesis of [Na(N(SiBuMe$_2$)$_2$)- (THF)]$_2$ (5a) and [Na(N(SiBuMe$_2$)$_2$)(DME)]$_2$ (5b) from benzyl sodium and HN(SiBuMe$_2$)$_2$, and a solvent-free structure of [K(N(SiBuMe$_2$)$_2$)]$_2$ (6), are also reported. Complexes 1, 5b, 7a, 7b, 8, 9 and 10 are fully characterised by single crystal XRD, multinuclear NMR and IR spectroscopy and elemental analysis, whereas complexes 2–4, 5a and 6 were identified by XRD only.

Sterically demanding N-donor ligands are commonly employed in Ae solution chemistry as they impede oligomerisation, complex Schlenk equilibria and other unwanted degradation pathways in ethereal solvents, particularly for the heavier Ae metals.5,6 Bochmann and co-workers have shown that cationic magnesium complexes with bulky amido ligands such as [Mg(N’’’’(Et$_2$O)$_3$)][BAr$_7$] [N’’’’ = {N(SiMe$_3$)$_2$}]$^{2-}$, Bar$_7$ = {B(C$_6$F$_5$)$_3$}$^{1-}$) can act as potent ring opening polymerisation catalysts,7 an industrially significant process in which Chisholm and others have shown magnesium catalysts have great promise.8 In seminal work by Jones, bulky N-donor ligands have been used to stabilise the first examples of structurally characterised Mg(i) complexes,9 which have since proven their utility as selective one-electron reducing agents in a number of diverse transformations.10

Multidentate ligands such as guanidinates,9–11 amido-nitrides12 and β-diketimines6,9,11,13,14 dominate N-donor magnesium chemistry and there are relatively few examples of monodentate complexes.1b,c,9,15 In 1994 Power disclosed the first crystallographically authenticated Hauser base complex (R$_2$NMgX; X = halide),16 [Mg(N’’’’)(μ-Cl)(Et$_2$O)]$_2$,17 which was originally prepared in 1972 by Wannagat et al. via a proto-

Introduction

Grignard reagents are widely utilised due to their facile preparation and broad applicability in organic synthesis,1,3 yet in stark contrast the synthetic potential of Hauser bases, amido analogues with a N–Mg bond instead of a C–Mg bond, is only starting to be realised.3 This is remarkable as N-donor groups are harder than C-donors and a wider variety of synthetic routes are available to access homo- and heteroleptic N-donor alkaline earth (Ae) complexes4 than Ae organometallics.1,4,5 Furthermore, highly reactive heteroleptic magnesium complexes with N-donors, such as alkyls and hydrides, are desirable and useful reagents, as they can undergo σ-bond metathesis or protonolysis with a number of substrates, thereby providing access to various synthetic hetero-functionalisations.3,6

School of Chemistry, The University of Manchester, Oxford Road, Manchester, M13 9PL, UK. E-mail: david.mills@manchester.ac.uk 1 Electronic supplementary information (ESI) available. CCDC 1408447–1408447 for 1–3, 5a–b, 6, 7a–b, 8–10 and 4. For ESI and crystallographic data in CIF or other electronic format see DOI: 10.1039/c5dt02533d

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Reactions of allylmagnesium chloride with silylamines and potassium silylamides

We found that the reactions of both potassium silylamides, \([\text{K}[\text{N}^\text{RR}]]\), and their parent silylamines with allylmagnesium chloride reproducibly gave mixtures of products and poor yields. Given that propylmagnesium chloride reacts with \(\text{HN}^*\) in diethyl ether to give \([\text{Mg}[\text{N}^*]](\text{Et}_2\text{O})\), we added allylmagnesium chloride to THF solutions of \(\text{HN}(\text{Si}^\text{Bu})\text{Me}_2\), \(\text{HN}(\text{Si}^\text{Bu})\text{Me}_2(\text{Si}^\text{Pr}_3)\), and \(\text{HN}(\text{Si}^\text{Pr}_3)\) (Scheme 1). Work-up of the respective reaction mixtures after 8 hours reproducibly gave crystals of the chloride Hauser bases, \([\text{Mg}[\text{N}^\text{RR}]](\text{µ-C}_3\text{H}_5)(\text{THF})\), \(\text{N}^\text{RR} = \{\text{N}(\text{Si}^\text{Bu})\text{Me}_2\}^\text{−}\), \(\text{1}\); \(\{\text{N}(\text{Si}^\text{Bu})\text{Me}_2(\text{Si}^\text{Pr}_3)\}^\text{−}\), \(\text{2}\); \(\{\text{N}(\text{Si}^\text{Pr}_3)\}^\text{−}\), \(\text{3}\), which were minor components of intractable mixtures of products. Complexes 1–3 were structurally characterised (see below) but they could not be separated from the bulk material, which was shown to be a mixture of products by \(^1\text{H}\) NMR spectroscopy. By increasing the reaction time to 10 days, we were able to improve the crystalline yield of complex 1 to 19%. Improved yields of 2–3 were not achieved by this method so further analysis of these complexes could not be obtained. On the other hand, the attempted salt metathesis reactions between allylmagnesium chloride and the corresponding potassium amides, \([\text{K}[\text{N}^\text{RR}]]\), did not give the expected amido-Grignard reagents, \([\text{Mg}[\text{N}^\text{RR}]](\text{C}_3\text{H}_5)\), despite several attempts at optimising the reaction conditions. However, in the attempted preparation of \([\text{Mg}[\text{N}(\text{Si}^\text{Pr}_3)\text{Me}_2]\)–(\(\text{C}_3\text{H}_5)\), several crystals of complex 3 were identified together with a crystal of \([\text{Mg}[\text{N}(\text{Si}^\text{Pr}_3)\text{Me}_2]\)(\(\text{µ-C}_3\text{H}_5\))\) by single crystal XRD (see below). These observations suggested that complex Schlenk equilibria processes are taking place in these reaction mixtures.1,31

The \(^1\text{H}\) and \(^{13}\text{C}\{^1\text{H}\}\) NMR spectra of 1 in \(d_6\)-benzene are unremarkable and not shifted considerably from those reported for \([\text{K}[\text{N}(\text{Si}^\text{Me}_2\text{Bu})\text{Me}_2]\)], although the \(^{29}\text{Si}\{^1\text{H}\}\) NMR spectrum of 1 exhibits a single resonance (δ: -0.40 ppm) that is shielded in comparison to the related potassium complex (δ: -15.72 ppm). Microanalysis was performed multiple times on discrete crystalline samples of complex 1 and consistently low C values were obtained despite good agreement between measured and expected H and N content. This phenomenon was also encountered for complexes 5b, 7b and 9 (see below), even though the \(^1\text{H}, ^{13}\text{C}\{^1\text{H}\}\) and \(^{29}\text{Si}\{^1\text{H}\}\) NMR spectra of these complexes exhibit only minor impurity signals.
(see ESI Fig. S1–S9†). Similar amounts of trace impurities were detected in the NMR spectra of other structurally characterised complexes reported herein, yet these have excellent agreement between predicted and observed microanalysis values. Therefore, we attribute these discrepancies to carbide formation, which has been cited previously as a recurring phenomenon in elemental analyses of silicon-rich complexes.29a,32

The identities of 1–4 were determined by single crystal XRD experiments (the molecular structure of 1 is depicted in Fig. 2; see ESI Fig. S10–S12† for the structures of 2–4. Selected bond lengths and angles are compiled in Table 1). The dataset obtained for 4 is of poor quality as the crystals were weakly diffracting. Therefore no discussion of the geometric parameters of 4 is given here, but as the connectivity is clear-cut the structure is included for completeness. The magnesium centres in 4 are 3-coordinate, which is rare for magnesium amide structures and even more so for monodentate ligands.21g,33 They are each bridged by two allyls, which generates an infinite 1D polymer comprised of two parallel-bridged chains. Complexes 1–3 exhibit similar motifs in the solid state, therefore only the structure of 1 is discussed for brevity. Complex 1 is dimeric with a central Mg2Cl2 rhomboid [Mg–Cl–Mg 92.15(6)°; Cl–Mg–Cl 87.85(6)°] and two 4-coordinate magnesium centres that are each bound by THF. The Mg–Cl [2.418(2) and 2.411(2) Å] and Mg–O [2.018(4) Å] distances are longer than those in [Mg(NH2)2(μ-Cl)(Et2O)2] [Mg–Cl 2.401(1) and 2.405(1) Å; Mg–O 2.000(3) Å],17 which we attribute to the increased steric demands of {N(Si′BuMe2)2}2−. Each magnesium coordination sphere of 1–3 is completed by two short Mg⋯C–H contacts [mean Mg⋯C 3.3485(8) Å (1); 3.192(4) Å (2); 3.267(6) Å (3)] and several short Mg⋯H distances [range Mg⋯H 2.71–2.75 Å (1); 2.63–3.02 Å (2); 2.56–2.67 Å (3)], where only one close Mg⋯C–H contact is present in [Mg(NAr)(μ-Cl)(Et2O)2] [Mg⋯C 3.207(5) Å; mean Mg⋯H 3.02 Å].17

Reactions of allyl-magnesium chloride with sodium silylamides

The reaction of sodium silylamides, [Na(NRR′)]30, with allyl-magnesium chloride did not give improved yields of the expected amido-Grignard complexes. During the course of these studies, we considered that sodium silylamides could be employed as alternative ligand transfer agents that may be less susceptible to byproduct formation. HN(Si′BuMe2)2 was treated with an excess of sodium hydride in THF, but this deprotonation strategy was laborious and only 30% conversion was observed by 1H NMR spectroscopy after 72 h reflux. In contrast, the deprotonation of HN(Si′BuMe2)2 with benzyl sodium34 in THF proceeded at room temperature to give [Na(N(Si′BuMe2)2)(THF)]3 as a dark oil following work-up (Scheme 2). Attempted crystallisation from pentane gave only several crystals of [Na(μ-N(Si′BuMe2)2)(THF)]3 (5a); therefore the residue was treated with DME and recrystallised from pentane to afford [Na(N(Si′BuMe2)2)(DME)]3 (5b) as a beige solid in fair yield (65%). The NMR spectra of 5b exhibit similar features to the potassium salt [K{N(SiMe3)2}]35 although 5b contains two additional resonances for coordinated DME in both the 1H (δ = 2.80 and 2.88 ppm) and 13C (δ = 59.21 and 71.30 ppm) NMR spectra.25Si{1H} NMR spectroscopy shows a single resonance for 5b (δ = −13.89 ppm; cf. [K{N(SiMe3)2}]3, δ = −15.73 ppm).29a

The solid state structures of 5a and 5b are depicted in Fig. 3 and 4 respectively, and selected bond lengths and angles are shown in Table 1. Both complexes exhibit bond metrics typical of silylamide group 1 salts.26,35 Complex 5a is dimeric, with a central Na2N2 solvent-capped core and a mean Na–N distance of 2.4903(6) Å, which is longer than the mean Na–silylamide bond lengths in [Na(μ-N(SiMe3)2)(THF)]3 [2.399(2) Å],35a [Na(μ-N(SiMe3)2)(μ-TMEDA)]3 (TMEDA = N,N,N′,N′-tetramethylethylenediamine) [2.4433(10) Å]15b and [Na(μ-N(SiMe3)2)(Ph)]3 (THF)]3 [2.440(4) Å].15c This can be attributed to the increased steric demands of the silylamide ligands in 5a. The sodium coordination spheres in 5a are completed by multiple short Na⋯C–H/Na⋯H distances [Na⋯C range 2.827 (4)–3.028(4) Å; Na⋯H range 2.48–2.88 Å], in common with the potassium salt [K{N(SiMe3)2}2(C–H)]3, which exhibits analogous K⋯C–H/K⋯H contacts.28 Complex 5b is monomeric in the solid state, with a 5-coordinate sodium centre and exhibits shorter Na–N distances [2.350(3) Å] than those in 5a, as would be expected for a terminally bound silylamide. The Na–O distances of 5b [range 2.425(3)–2.567(4) Å] are within the range of those for Na–DME interactions in the literature [2.402(9)–2.658(7) Å].16 To the best of our knowledge, [Sm{[N(Ar)(C(CH3)2)]2CHCNC(Pr)2Na(DME)}2κ3,N–N′,N′κ1{N(SiMe3)2}2] (Ar = C6H11, HfPr2,2,6)36d is the only other literature example containing a terminal, DME-capped N–Na bond. As with 5a, the coordination sphere of sodium in 5b is also completed by short Na⋯H contacts [range Na⋯H 2.84–3.06 Å].

The reaction of 5b with allylmagnesium chloride in THF was performed in an attempt to synthesise [Mg{N(SiMe3)2}]−.
Table 1  Selected bond lengths (Å) and angles (°) for 1–3 and 5–10

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<th>N(1)–Mg(1)</th>
<th>Mg(1)–Cl(1)</th>
<th>Mg(1)–Cl(2)</th>
<th>N(1)–Si(2)</th>
<th>Si(1)–N(1)–Si(2)</th>
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<td>1.998(4)</td>
<td>2.174(2)</td>
<td>2.018(4)</td>
<td>1.717(2)</td>
<td>126.8(2)</td>
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<tr>
<td>2</td>
<td>2.005(2)</td>
<td>2.207(2)</td>
<td>1.722(2)</td>
<td>1.909(2)</td>
<td>117.9(7)</td>
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<tr>
<td>3</td>
<td>2.028(3)</td>
<td>2.237(2)</td>
<td>1.736(3)</td>
<td>1.909(2)</td>
<td>117.9(7)</td>
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<tr>
<td>4</td>
<td>2.050(3)</td>
<td>2.243(2)</td>
<td>1.725(3)</td>
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[C(CH₃)₂]. However, we were not able to isolate the amidogrignard reagent from this reaction mixture, even after 10 days of stirring, and ¹H NMR spectroscopy revealed a complex mixture of products. We concluded from these observations that allylmagnesium chloride is unsuitable for facile and reproducible synthetic routes to NRR-containing Hauser bases, at least in the reaction conditions employed, and therefore we switched our attention to a different organomagnesium starting material (see below).

Reactions of methyldimagnesium iodide with silylamines and potassium silylamides

Methyldimagnesium iodide reacts with potassium silylamides, [K(NRR)]₂, to give amido-grignard complexes in fair yields, but protonolysis reactions with the parent silylamines were unsuccessful. As the reaction of [K(N₂BuMe₂)]₂ or 5b with allylmagnesium chloride is slow and gives poor yield of 1 (see above), we selected methyldimagnesium iodide as an alternative starting material, which has been extensively used in the literature as a reagent for the preparation of L–Mg and L–Mg–R complexes. In the course of synthesising potassium silylamide precursors for this study, crystals of [K(μ-N₂BuMe₂)]₂ [C₃H₇] were obtained and identified by single crystal XRD (Fig. 5, selected bond lengths and angles are shown in Table 1). We previously reported that desolvation of [K(μ-N₂BuMe₂)]₂ [C₃H₇] under vacuum yielded 6, but this was not structurally characterised. In the solid state, complex 6 is dimeric with a K₂N₂ core. K⋯C–H agostic interactions [K⋯C 3.188(4) Å] between the dimers form a 1D chain and numerous other K⋯C–H/K⋯C−H agostic and agostic interactions are also present. Each potassium centre in 6 is close to two nitrogen atoms, with N–K distances [2.856(3) Å mean] and K–N–K [74.59(6)°] and N–K–N [105.41(6)°] angles that are comparable to those in [K(μ-N₂BuMe₂)]₂ [C₃H₇] and [K(μ-N₂BuMe₂)]₂ [C₃H₇] where K–N 2.874(2) Å mean; K–N–K 78.37(5)°; N–K–N 101.63(5)° and [K(μ-N₂)]₂ [K–N 2.787(3) Å mean; K–N–K 85.53(9)°; N–K–N 94.47(9)°].

The amido-grignard complexes [Mg(NRR)₂(μ-CH₃)]₂ [NRR = N₂BuMe₂] were evaluated for use as diethyl ether, and treatment of 7a with DME gave [Mg[N₂BuMe₂]₂] (Scheme 3). Workup and crystallisation from
hot toluene afforded colourless crystals in poor (7a and 7b), fair (9) and excellent (8) yields, respectively. On one occasion a small crop of the Hauser base [Mg{N(Si\text{t}BuMe2)2}(\mu-I)(THF)]2 (10) was obtained (4%) when attempting to isolate 7a, which has presumably formed via complex Schlenk equilibria.\textsuperscript{1} Treatment of methylmagnesium iodide with HNRR′ in diethyl ether gave no reaction even after extended reaction times (>3 days at room temperature) or heating the reaction mixture under reflux for three hours. Furthermore, we found that 7a does not react with HN(Si\text{t}BuMe2)2 under similar forcing conditions in hexanes. Germane to this, no reaction was observed between di-\text{t}butylmagnesium and HN(Si\text{t}BuMe2)2 in a mixture of heptane and hexanes. Refluxing this reaction mixture for extended periods gave an intractable mixture of products. Together, these experiments illustrate the sluggishness of protolysis reactions between alkylmagnesium complexes and HNRR′. Complexes 7–10 were characterised by single crystal XRD studies (see below), elemental analysis and NMR and IR spectroscopies. To the best of our knowledge, 7a, 8 and 9 are the first examples of structurally characterised dimeric...
CH$_3$-bridged Hauser bases with monodentate N-donor ligands. Complex 7b is unusual as it is monomeric with a terminal Mg–CH$_3$ group, and there are few examples of this motif in the literature.

The $^1$H and $^{13}$C($^1$H) NMR spectra of 7–10 contain unremarkable silylamide ligand resonances, but the shielded methyl resonances in 7–9 [δ$_H$: −0.71 ppm (7a); −0.99 ppm (7b); −0.81 ppm (8); −0.80 ppm (9); δ$_C$: −10.04 ppm (7a), −13.65 ppm (7b), −13.02 ppm (8), −12.19 ppm (9)] are in agreement with those of structurally characterised examples in the literature (δ$_H$ range −2.00 ppm to −0.67 ppm). The similarity of the methyl group chemical shifts in 7–9 with monodentate N-donor complexes in the literature suggests that the spectator ligand does not greatly influence these values.6,19,40

One resonance is observed in the $^{29}$Si($^1$H) NMR spectra of 7a, 7b, 8, 9 and 10 and two resonances were found for 8 [δ$_H$: −1.93 ppm (7a); −2.43 ppm (7b); −2.90 and −1.76 ppm (8); −3.15 ppm (9); −1.19 ppm (10)], correlating with the number of unique silicon environments.

The identities of 7a, 7b, 8, 9 and 10 were determined by single crystal XRD and are depicted herein (7a: Fig. 6; 7b: Fig. 7; 10: Fig. 8; see ESI Fig. S13 and S14† for the structures of 8 and 9), with selected bond lengths and angles compiled in Table 1. The structures of 7a, 8 and 9 are broadly similar, with 3-coordinate Mg centres and bridging CH$_3$ moieties to form central Mg$_3$C$_2$ rhomboids with centres of inversion. The Mg–N distances [7a: 1.956(3) Å, 8: 1.9658(18) Å; 9: 1.977(2) Å] are slightly shorter than in 1, 2, and 3 respectively, which is a reflection of the absence of coordinated solvent molecules in 7–9. The mean Mg–C bond lengths in 7a, 8 and 9 are comparable [7a: 2.217(6) Å; 8: 2.236(3) Å; 9: 2.232(5) Å] and are typical of methylmagnesium amides [previously reported range Mg–C 1.977(3)–2.434(13) Å].33 The MgNSi$_2$ fragments of all three ligands are roughly planar in 7a, 8 and 9; these planes are twisted relative to the central Mg$_3$C$_2$ plane by differing amounts [7a: 74.38(11)°; 8: 63.95(8)°; 9: 67.77(10)°] with no clear trend. In the dimeric unit, the two ligand Si–N–Si fragments are also co-planar with each other (to within 0.1°) for all three structures. The coordination spheres of the magnesium centres in 7a, 8 and 9 are completed by a number of short Mg⋯C–H/Mg⋯H distances. Complex 7b is monomeric due to the coordinated DME molecule and the Mg–N [2.0293(15) Å] and Mg–C [2.118(2) Å] bonds are correspondingly shorter than those in 7a. However, the coordination sphere of the magnesium centre is again completed by short Mg⋯C–H/Mg⋯H distances. Complex 10 exhibits a near-square Mg$_2$I$_2$ central motif, with the sum of the four internal angles calculated at 359.99(12)°. The two Mg–I bond lengths in 10 [2.8280(18) and 2.8187(19) Å] are similar to those observed for literature examples that contain N–Mg–I moieties and bridging iodide ligands [range 2.7766(12)–2.901(3) Å].6,19,41 Finally, in common with 1–3, the coordination sphere of Mg is completed by a THF molecule [Mg–O 2.010(4) Å] and multiple short Mg⋯C–H/Mg⋯H distances.

**Conclusions**

We have structurally characterised a series of silylamide Hauser bases during our exploration of straightforward syn-
thetic routes to these complexes. Protonolysis routes to Hauser base complexes by treating allylmagnesium chloride, di-\(n\)-butylmagnesium or methylmagnesium iodide with a series of silylamines were found to be slow and prone to complex Schlenk equilibria. A fully characterised isolated product, 1, was achieved on only one occasion by using extended reaction times. The salt metathesis reactions of allylmagnesium chloride with potassium silylamides, and the novel sodium silylamine, 5b, gave intractable mixtures. We concluded from these studies that, using these ligand systems and methodologies, allylmagnesium chloride is unsuitable as a starting material for reproducible syntheses of silylamine Hauser base and amido-Grignard complexes.

Salt metathesis reactions of potassium silylamides with methylmagnesium iodide were found to be a far more successful strategy, giving modest to excellent yields of the amido-Grignard complexes 7–9. The isolation of a small amount of the iodide Hauser base complex 10 indicated that complex Schlenk equilibria also operate in these reaction mixtures, slightly reducing the yields of the target complexes. Despite these side-reactions, facile synthetic routes to amido-Grignard complexes of three different silylamine ligands have been achieved. We envisage that such amido-Grignard complexes could be useful reagents for heterofunctionalisation reactions. Additionally, such compounds could pave the way for the preparation of novel bimetallic systems with alkali metals, which could potentially be employed in synergic metalation reactions.

Experimental

Materials and methods

All manipulations were carried out using standard Schlenk and glove box techniques under an atmosphere of dry argon. Solvents were dried by refluxing over potassium and degassed before use. All solvents were stored over potassium mirrors (with the exception of THF and DME which were stored over activated 4 Å molecular sieves). Deuterated solvents were distilled from potassium, degassed by three freeze–pump–thaw cycles and stored under argon. \([\text{K}[\text{N(Si}^\text{tBuMe}_2)_2]]_{29}^\text{a} \) \([\text{K}[\text{N(Si}^\text{tBuMe}_2)(\text{Si}^\text{Pr}_3)_2]]_{28}^\text{b} \) \([\text{K}[\text{N(Si}^\text{Pr}_3)_2]]_{29}^\text{b} \) were prepared according to published procedures. \([\text{Na(CH}_2\text{C}_6\text{H}_5}]\) was prepared via a modification of published procedures.\(^{34}\) All other chemicals were used as purchased and stored appropriately. Most solid reagents were dried under vacuum for four hours and most liquid reagents were dried over 4 Å molecular sieves and distilled before use. \(^1\text{H}, \(^{13}\text{C}\{^1\text{H}\} \) and \(^{29}\text{Si}\{^1\text{H}\} \) NMR spectra were recorded on a spectrometer operating at 400.2, 100.6 and 79.5 MHz, respectively; chemical shifts are quoted in ppm and are relative to TMS. FTIR spectra were recorded as Nujol mulls in KBr discs on a Perkin Elmer Spectrum RX1 spectrometer. Elemental microanalyses were carried out by Mr Stephen Boyer at the Microanalysis Service, London Metropolitan University, UK.

Synthetic procedures

\([\text{Mg}[\text{N(Si}^\text{tBuMe}_2)_2](\mu-\text{Cl})(\text{THF})]_2\) (1). Allylmagnesium chloride (1.5 mL, 2.0 M in THF, 3 mmol) was added dropwise to a precooled solution –78 °C) of \([\text{HN(Si}^\text{tBuMe}_2)_2]]_{29}^\text{a} \) in THF (20 mL). The reaction mixture was allowed to warm to room temperature, forming a white precipitate which redissolved to give a colourless solution. After 10 days stirring at room temperature volatiles were removed \textit{in vacuo} to leave a
white solid which was washed with hexanes (4 mL). The solid residue was dried in vacuo and extracted with hot toluene (15 mL). Storage of the toluene solution at −20 °C for 24 hours gave 1 as colourless crystals (0.44 g, 19%). Anal. Caled for C20H50NNaO4Si2: C, 53.68; H, 11.26; N, 3.13. Found C, 40.48; H, 8.60; N, 3.38 (low C value attributed to 1 being a silicon-rich molecule, as has been observed previously).

1H NMR (6-benzene, 298 K): δ = 0.38 (s, 24H, Si(CH3)2), 1.18 (s, 36H, Si(CH3)2), 1.24 (m, 8H, THF–CH2), 2.84 (m, 8H, THF–OCH2). 13C{1H} NMR (6-benzene, 298 K): δ = 2.72 (Si(CH3)2), 20.98 (Si(CH3)2), 25.16 (THF–CH2), 32.93 (Si(CH3)2), 71.19 (THF–OCH2). 29Si{1H} NMR (6-benzene, 298 K): δ = −0.40 (Si(BuMe2)), FTIR (Nujol, cm−1): ν = 1259 (s), 1246 (s), 1019 (s), 983 (s), 844 (m), 829 (s), 795 (s), 722 (m), 660 (m).

[Mg(N[Si(ButMe2)2][S(Pr)3])] ([μ-Cl]THF)] (7a) and [Mg[N(Si(ButMe2)2)]([CH3]2) (DME)] (7b) and [Mg[N(Si(ButMe2)2)]([μ-I]THF)] (10). Methylmagnesium iodide (2 mL, 3 M in diethyl ether, 6 mmol) was added dropwise to a pre-cooled −78 °C solution of [KI[N(Si(ButMe2)2)] (1.70 g, 6 mmol) in diethyl ether (20 mL). The reaction was allowed to warm to room temperature and stirred for 8 hours, forming a white precipitate. Filtration and removal of volatiles in vacuo gave a colourless oil. Crystallisation from toluene (20 mL at room temperature yielded 7a (0.91 g, 32%). During the synthesis and attempted crystallisation of 7a from THF, we obtained a small crop of 10 (0.10 g, 4%). Following isolation and characterisation of 7a and 10, the extracts were recombined and volatiles were removed in vacuo. Dimethoxymethane (5 mL) was added and the resultant light yellow solution was stirred for 2 hours. The volatiles were removed in vacuo to leave a white crystalline solid that was dissolved in toluene (3 mL). The reaction was transferred to a −20 °C freezer for 24 hours to form colourless crystals of 7b (1.03 g, 46%). Data for 7a: Anal. Caled for C25H32MgN2O5Si2: C, 55.00; H, 11.72; N, 4.93. Found C, 54.96; H, 11.79; N, 4.90. 1H NMR (6-benzene, 298 K): δ = −0.64 (s, 6H, Mg–CH3), 0.16 (s, 24H, Si(CH3)2), 1.01 (s, 36H, Si(CH3)2). 13C{1H} NMR (6-benzene, 298 K): δ = −10.04 (Mg–CH3), 0.86 (Si(CH3)2), 19.92 (Si(CH3)2), 27.79 (Si(CH3)2). 29Si{1H} NMR (6-benzene, 298 K): δ = −1.93 (Si[BuMe2]). FTIR (Nujol, cm−1): ν = 1250 (s), 1211 (s), 1077 (m), 1030 (s), 936 (m), 753 (m), 642 (m).

Data for 7b: Anal. Caled for C25H32MgN2O5Si2·0.1C6H6·C6H4: C, 55.47; H, 11.52; N, 3.66. Found C, 49.79; H, 11.44; N, 3.75 (low C value attributed to 7b being a silicon-rich molecule, as has been observed previously).

1H NMR (6-benzene, 298 K): δ = −0.99 (s, 3H, Mg–CH3), 0.34 (s, 12H, Si(CH3)2), 1.18 (s, 18H, Si(CH3)2), 2.72 (s, 4H, OCH3), 2.79 (s, 6H, OCH3). 13C{1H} NMR (6-benzene, 298 K): δ = −13.65 (Mg–CH3), 2.19 (Si(CH3)2), 21.10 (Si(CH3)2), 29.00 (Si(CH3)2), 58.79 (OCH3), 68.48 (OCH3). 29Si{1H} NMR (6-benzene, 298 K): δ = −2.43 (Si[BuMe2]), FTIR (Nujol, cm−1): ν = 1246 (s), 1047 (s), 1015 (s), 822 (s), 791 (s), 649 (m). Data for 10: Anal. Caled for C32H71Mg2N4O12Si2·C, 41.56; H, 8.19; N, 2.96. Found C, 40.95; H, 8.24; N, 3.02. 1H NMR (6-benzene, 298 K): δ = −0.36 (s, 24H, Si(CH3)2), 1.24 (s, 36H, Si(CH3)2), 1.40 (m, 8H, THF–CH2), 3.59 (m, 8H, THF–OCH2), 13C{1H} NMR (6-benzene, 298 K): δ = 3.44 (Si(CH3)2), 21.29 (Si(CH3)2), 25.93 (THF–CH2), 29.75 (Si(CH3)2), 68.60 (THF–OCH2), 29Si{1H} NMR (6-benzene, 298 K): δ = −1.19 (Si[BuMe2]). FTIR (Nujol, cm−1): ν = 1233 (s), 1016 (s), 981 (s), 823 (s), 795 (s), 645 (m).

[Mg[N(Si(ButMe2)2)]([μ-CH3]) (8). Methylmagnesium iodide (1 mL, 3 M in diethyl ether, 3 mmol) was added dropwise to a pre-cooled −78 °C solution of [KI[N(Si(ButMe2)2)] (1.95 g, 3 mmol) in diethyl ether (15 mL). The reaction was allowed to warm to room temperature and stirred for 8 hours,
forming a white precipitate. Filtration and removal of volatiles in vacuo yielded a white precipitate. This was dissolved in hot toluene (15 mL) and colourless crystals of 8 formed at room temperature (1.46 g, 93%). Anal. Calcd For C_{26}H_{57}Mg_2N_2Si_4·0.75C_7H_8: C, 62.05; H, 11.74; N, 3.89. Found C, 62.05; H, 11.74; N, 3.89. 1H NMR (298 K): δ = −0.81 (s, 6H, Mg–CH_3), 3.12 (s, 12H, Si(CH_3)_2), 1.12 (m, 6H, Si(CH_3)_2), 1.15 (s, 18H, Si(CH_3)_2), 1.31 (d, 18H, J_{HH} = 4 Hz, SiCH(CH_3)_2), 13C{1H} NMR (298 K): δ = −13.02 (Mg–CH_3), 1.82 (Si(CH_3)_2), 17.00 (Si(CH(CH_3)_2)), 20.14 (Si(CH(CH_3)_2)), 20.73 (Si(CH(CH_3)_2)), 28.55 (Si(CH(CH_3)_2)), 29Si{1H} NMR (298 K): δ = −2.90 (Si^3Pr_3), −1.76 (Si^3BuMe_2). FTIR (Nujol, cm\(^{-1}\)): ν = 1251 (s), 1015 (s), 820 (s), 780 (s), 722 (s), 641 (m), 627 (m).

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**References**

