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6, 4511An oxygen pool from YBaCo<sub>4</sub>O<sub>7-δ</sub>-based oxides for  
soot combustion†Guangkai Tian,<sup>a</sup> Hui Chen,<sup>a</sup> Chenxi Lu,<sup>a</sup> Ying Xin,<sup>a</sup> Qian Li,<sup>a</sup> James A. Anderson<sup>b</sup>  
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Soot, often referred to as black carbon emitted from diesel engines, is not only a particulate matter pollutant but also a light-absorbing agent that may affect global climate, but can be effectively controlled using a catalytic diesel particulate filter (DPF). A new YBaCo<sub>4</sub>O<sub>7-δ</sub>-type oxygen storage material is reported as an effective catalyst for soot combustion. Isotopic isothermal reactions demonstrate the activation of gaseous oxygen and subsequent oxygen storage and reaction/desorption during an oxidation process. High activity and structural stability are achieved by the substitution of Co with Al and Ga to form YBa(Co<sub>0.85</sub>Al<sub>0.075</sub>Ga<sub>0.075</sub>)<sub>4</sub>O<sub>7-δ</sub>. The specific rates at 300 °C of YBaCo<sub>4</sub>O<sub>7-δ</sub> and YBa(Co<sub>0.85</sub>Al<sub>0.075</sub>Ga<sub>0.075</sub>)<sub>4</sub>O<sub>7-δ</sub>, normalized by surface areas, are an order of magnitude higher than those of CeO<sub>2</sub>-based oxides. This kind of oxygen-storage material acts as an oxygen pool, which ensures that the accumulated soot on a DPF can be promptly combusted.

Received 15th December 2015,  
Accepted 5th February 2016

DOI: 10.1039/c5cy02181b

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## 1. Introduction

Soot, often referred to as black carbon, is not only a particulate matter pollutant but also a light-absorbing agent that may affect global climate.<sup>1</sup> Diesel engines are amongst the most abundant emission sources of soot that can be effectively controlled by a catalytic diesel particulate filter (DPF).<sup>2</sup> Commercial catalysts are composed of noble metals (Pt and Pd) supported on ceria-based oxides,<sup>3</sup> oxidizing NO into NO<sub>2</sub>, which is transported through a gas phase to soot aggregates where it oxidizes carbon while being reduced to NO.<sup>4</sup> A so-called continuously regenerated trap (CRT) overcomes the problem of poor contact between soot and a catalyst. However, the limited amounts of NO<sub>x</sub> present and the need to control NO<sub>x</sub> for Euro IV standards are not ideal situations to meet the requirements of the latest generation of diesel engines, leading to a drive to develop catalysts which can produce highly reactive oxygen species from gaseous O<sub>2</sub> molecules.<sup>5,6</sup>

Among this kind of catalysts, the extensively studied soot oxidation catalysts are ceria-based oxides due to the redox properties of the Ce<sup>3+</sup>/Ce<sup>4+</sup> couple and the capacity of ceria to exchange oxygen with a gas phase.<sup>7–12</sup> Unfortunately, the

effective utilization of such active oxygen species is limited by the extent of contact between soot and a catalyst. This limitation could be overcome by producing a sufficiently high amount of active oxygen to ensure that, at least, a proportion reaches the soot particle surface.<sup>5</sup>

In comparison with a conventional oxygen storage material, CeO<sub>2</sub>-ZrO<sub>2</sub>, YBaCo<sub>4</sub>O<sub>7</sub> (114 structure) shows a markedly larger oxygen storage capacity (OSC), particularly at low temperatures (200–400 °C), which is just within the range of temperatures reached at the exhaust of a diesel engine.<sup>13–16</sup> The extraordinary oxygen storage capability of the YBaCo<sub>4</sub>O<sub>7</sub>-based oxides is due to the variable valence of Co ions between Co<sup>2+</sup> and Co<sup>3+</sup>. The limitation of YBaCo<sub>4</sub>O<sub>7-δ</sub> is that it decomposes just above 600 °C in an oxygen-containing atmosphere, which limits its applications in catalytic combustion at elevated temperatures.<sup>13</sup> Recently, Karppinen and colleagues identified that an YBa(Co<sub>0.85</sub>Al<sub>0.075</sub>Ga<sub>0.075</sub>)<sub>4</sub>O<sub>7-δ</sub> phase, where Co is co-substituted by Al and Ga is stable up to high temperatures under oxidizing conditions,<sup>16</sup> which creates the potential for YBa(Co<sub>0.85</sub>Al<sub>0.075</sub>Ga<sub>0.075</sub>)<sub>4</sub>O<sub>7-δ</sub> to become a more effective catalyst for catalytic combustion.

In this paper, YBaCo<sub>4</sub>O<sub>7-δ</sub> and YBa(Co<sub>0.85</sub>Al<sub>0.075</sub>Ga<sub>0.075</sub>)<sub>4</sub>O<sub>7-δ</sub> were studied from the perspective of catalysis for diesel soot combustion. Their oxygen storage performance allows them to create an oxygen pool supply of active oxygen which is created from a gaseous phase.<sup>17</sup> Although YBaCo<sub>4</sub>O<sub>7-δ</sub> has been reported as a robust catalyst for H<sub>2</sub>O<sub>2</sub> oxidation of cyclohexene in the liquid phase,<sup>18</sup> the present result is the first report on high-temperature oxidation reactions for this kind of non-stoichiometric transition metal oxide materials.

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† Electronic supplementary information (ESI) available: Textural properties and XRD patterns of the samples. See DOI: 10.1039/c5cy02181b



## 2. Experimental section

The  $\text{YBaCo}_4\text{O}_{7+\delta}$ -type samples were synthesized from a precursor powder prepared by an EDTA (ethylenediaminetetraacetic acid) and CA (citric acid) complex gel method,<sup>13</sup> using  $\text{Y}(\text{NO}_3)_3 \cdot 6\text{H}_2\text{O}$ ,  $\text{Ba}(\text{NO}_3)_2$ ,  $\text{Co}(\text{NO}_3)_2 \cdot 6\text{H}_2\text{O}$ ,  $\text{Ga}(\text{NO}_3)_3 \cdot 6\text{H}_2\text{O}$  and  $\text{Al}(\text{NO}_3)_3 \cdot 9\text{H}_2\text{O}$  as the starting materials. The precursor powder was calcined in a muffle at 350 °C for 2 h, then pelletized and sintered in air at 1000 °C for 24 h, followed by rapid cooling to room temperature. Further thermal treatment at 300 °C under a pure  $\text{O}_2$  flow for 2 h obtained  $\text{YBaCo}_4\text{O}_{7+\delta}$  and  $\text{YBa}(\text{Co}_{0.85}\text{Al}_{0.075}\text{Ga}_{0.075})_4\text{O}_{7+\delta}$ .

X-ray powder diffraction (XRD) patterns were recorded on a Rigaku D/max-RC diffractometer employing  $\text{Cu K}\alpha$  radiation. Surface areas and pore size distributions were determined by  $\text{N}_2$  adsorption-desorption at 77 K using a Micromeritics ASAP 2020 instrument after outgassing at 300 °C for 5 h prior to analysis.

Temperature programmed desorption of  $\text{O}_2$  ( $\text{O}_2$ -TPD) experiments were conducted in a fixed-bed flow reactor. A 150 mg sample was heated under a flow of high purity  $\text{O}_2$  (30 ml  $\text{min}^{-1}$ ) at 300 °C for 1 h. After cooling to room temperature, high purity He was introduced. Desorption was started at a heating rate of 2 °C  $\text{min}^{-1}$  in He (30 ml  $\text{min}^{-1}$ ). The desorbed  $\text{O}_2$  was monitored by a quadruple mass spectrometer (MS, OminiStar 200, Balzers).

Temperature programmed reduction with  $\text{H}_2$  ( $\text{H}_2$ -TPR) experiments were performed in a quartz reactor with a thermal conductivity detector (TCD) to monitor  $\text{H}_2$  consumption. A 50 mg sample was pretreated *in situ* at 300 °C for 1 h under a flow of  $\text{O}_2$  and cooled to room temperature in the presence of  $\text{O}_2$ . After purging with  $\text{N}_2$ , TPR was conducted at 10 °C  $\text{min}^{-1}$  up to 700 °C under a 30 mL  $\text{min}^{-1}$  flow of 5 vol%  $\text{H}_2$  in  $\text{N}_2$ . To quantify the total amount of  $\text{H}_2$  consumption,  $\text{CuO}$  was used as a calibration reference.

“Dynamic” OSC (DOSOC) measurements with  $\text{CO-O}_2$  pulses were carried out at 200–500 °C.  $\text{CO}$  (4%  $\text{CO}/1\%$   $\text{Ar}/\text{He}$  at 300 mL  $\text{min}^{-1}$  for 10 s) and  $\text{O}_2$  (2%  $\text{O}_2/1\%$   $\text{Ar}/\text{He}$  at 300 mL  $\text{min}^{-1}$

for 10 s) streams were pulsed alternately with at a frequency of 0.05 Hz. A DOSOC value was obtained by integrating the  $\text{CO}_2$  formed during one  $\text{CO-O}_2$  cycle and was expressed as  $\mu\text{mol}$  of O per gram of catalyst ( $\mu\text{mol} [\text{O}] \text{g}^{-1}$ ). The concentration of  $\text{CO}_2$  was determined using a MS.

Temperature programmed oxidation (TPO) reactions were conducted in a fixed bed micro reactor consisting of a quartz tube (6 mm i. d.). Printex-U from Degussa was used as the model soot. The soot was mixed with the catalyst in a weight ratio of 1 : 9 in an agate mortar for 30 min, which resulted in a tight contact between the soot and the catalyst. A 50 mg sample of the soot/catalyst mixture was pre-treated under a flow of He (50 mL  $\text{min}^{-1}$ ) at 200 °C for 30 min to remove adsorbed species. After cooling to room temperature, a gas flow with 5 vol% oxygen in He was introduced and then TPO was initiated at a heating rate of 5 °C  $\text{min}^{-1}$  until 880 °C. For pure soot combustion (non-catalytic), the catalyst was substituted by silica.  $\text{CO}$  and  $\text{CO}_2$  concentrations in the effluent gas were monitored using an online gas chromatograph (GC) (SP-6890, Shandong Lunan Ruihong Chemical Instrument Corporation, China) fitted with a methanator. The ignition temperature for soot combustion was evaluated by the value of  $T_{10}$ , which is defined as the temperature at which 10% of the soot is converted. The selectivity to  $\text{CO}_2$  is defined as the percentage  $\text{CO}_2$  outlet concentration divided by the sum of the  $\text{CO}_2$  and  $\text{CO}$  outlet concentrations.

Isothermal reactions at 300 °C, at which a stable and low soot conversion (<15%) was achieved, were conducted within the kinetic regime. The reaction rate for soot combustion was obtained from the slope of the conversion lines with time. Specific rates normalized by BET surface areas and turnover frequency (TOF)<sup>19</sup> were used to characterize the activity for soot combustion.

An isotopic isothermal reaction was performed by switching the flowing gas from 1%  $^{16}\text{O}_2$  to 1%  $^{18}\text{O}_2$  diluted in Ar at 350 °C. 50 mg of a mixture of the soot and catalyst in a tight contact mode was employed. The effluent gas from the reactor was continuously monitored by a MS for all of the isotopic molecules of  $\text{CO}_2$  (at  $m/z = 44, 46$  and  $48$ ).

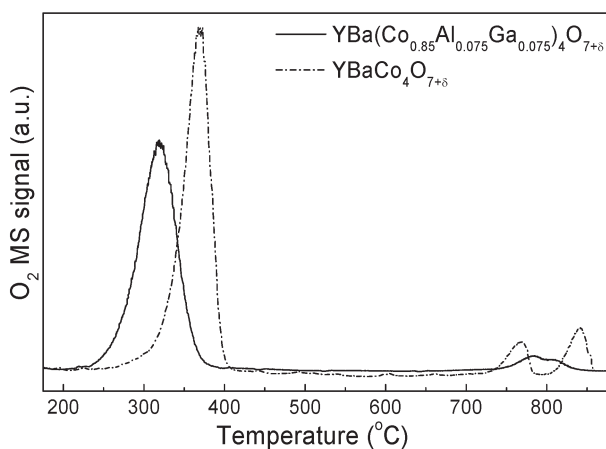


Fig. 1  $\text{O}_2$ -TPD spectra of  $\text{YBaCo}_4\text{O}_{7+\delta}$  and  $\text{YBa}(\text{Co}_{0.85}\text{Al}_{0.075}\text{Ga}_{0.075})_4\text{O}_{7+\delta}$  from room temperature to 900 °C.

## 3. Results and discussion

The fresh  $\text{YBaCo}_4\text{O}_{7+\delta}$  and  $\text{YBa}(\text{Co}_{0.85}\text{Al}_{0.075}\text{Ga}_{0.075})_4\text{O}_{7+\delta}$  were confirmed to be composed of a single phase (Fig. S1†), similar to that of hexagonal  $\text{LuBaAlZn}_3\text{O}_7$  (JCPDS 40-4099).<sup>13</sup> As the shifts in the XRD peaks for the separate substitution of Co by Al ( $r = 0.039$  nm) and Ga ( $r = 0.047$  nm) are opposed for each sample (Fig. S2†), the simultaneous substitution has a slight effect on the cell volume parameters. Furthermore, both ionic radii of Al and Ga are smaller than the high spin  $\text{Co}^{2+}$  ionic radius ( $r = 0.058$  nm), thus Maignan *et al.* suggested that Al and Ga are substituted for  $\text{Co}^{3+}$ .<sup>20</sup> The BET surface areas are fairly low, in accordance with the high-temperature sintering preparation (Table S1†) and the highly crystalline nature of the samples.



In agreement with conclusions in the literature,<sup>13,16</sup> the O<sub>2</sub>-TPD profiles (Fig. 1) show that both YBaCo<sub>4</sub>O<sub>7+δ</sub> and YBa(Co<sub>0.85</sub>Al<sub>0.075</sub>Ga<sub>0.075</sub>)<sub>4</sub>O<sub>7+δ</sub> desorb a large amount of oxygen below 400 °C, corresponding to δ = 0.37 and δ = 0.34, respectively. The peak maximum of YBa(Co<sub>0.85</sub>Al<sub>0.075</sub>Ga<sub>0.075</sub>)<sub>4</sub>O<sub>7+δ</sub> is 50 °C lower than that of YBaCo<sub>4</sub>O<sub>7+δ</sub>, which suggests a promotion effect on O<sub>2</sub> desorption by doping, albeit with a slight decrease of the overall oxygen storage capability.<sup>16</sup> Above 700 °C, further desorption of O<sub>2</sub> is observed, corresponding to the possible decomposition of the 114 structure.<sup>13</sup> However, YBa(Co<sub>0.85</sub>Al<sub>0.075</sub>Ga<sub>0.075</sub>)<sub>4</sub>O<sub>7+δ</sub> shows less pronounced O<sub>2</sub> desorption than YBaCo<sub>4</sub>O<sub>7+δ</sub>. After O<sub>2</sub>-TPD at about 900 °C, the Y<sub>2</sub>O<sub>3</sub> and CoO phases are segregated from YBaCo<sub>4</sub>O<sub>7+δ</sub>, which is opposite to the stable structure of YBa(Co<sub>0.85</sub>Al<sub>0.075</sub>Ga<sub>0.075</sub>)<sub>4</sub>O<sub>7+δ</sub> (Fig. S3†).

A similar situation is observed in H<sub>2</sub>-TPR (Fig. S1†). The 114 structure of YBaCo<sub>4</sub>O<sub>7+δ</sub> is completely destroyed by H<sub>2</sub> in TPR to 700 °C (Fig. S1a and b†). In contrast, YBa(Co<sub>0.85</sub>Al<sub>0.075</sub>Ga<sub>0.075</sub>)<sub>4</sub>O<sub>7+δ</sub> was preserved with no formation of new oxide phases (Fig. S1c and d†). The structural stability of YBa(Co<sub>0.85</sub>Al<sub>0.075</sub>Ga<sub>0.075</sub>)<sub>4</sub>O<sub>7+δ</sub> is vital to high-temperature redox reactions. As shown in Fig. 2, two peaks were observed for both YBaCo<sub>4</sub>O<sub>7+δ</sub> and YBa(Co<sub>0.85</sub>Al<sub>0.075</sub>Ga<sub>0.075</sub>)<sub>4</sub>O<sub>7+δ</sub>. The low-temperature TPR peak can be assigned to the removal of non-stoichiometric excess O accommodated within the lattice, and the values of which are slightly larger than those consumed by δ (Table 1). The 114 structure of YBaCo<sub>4</sub>O<sub>7+δ</sub> is stable at this stage (Fig. S1a and b†). The second peak corresponds to the reduction of bulk and surface Co<sup>3+</sup> of YBaCo<sub>4</sub>O<sub>7+δ</sub> and YBa(Co<sub>0.85</sub>Al<sub>0.075</sub>Ga<sub>0.075</sub>)<sub>4</sub>O<sub>7+δ</sub>, respectively. The substitution of Co<sup>3+</sup> by Al and Ga protects the structure from decomposition under reducing atmospheres (H<sub>2</sub>-TPR and O<sub>2</sub>-TPD).<sup>20</sup> Furthermore, the lower TPR temperature of YBa(Co<sub>0.85</sub>Al<sub>0.075</sub>Ga<sub>0.075</sub>)<sub>4</sub>O<sub>7+δ</sub> in comparison with that of YBaCo<sub>4</sub>O<sub>7+δ</sub> coincides with the O<sub>2</sub>-TPD results.

Although H<sub>2</sub>-TPR and O<sub>2</sub>-TPD data may be useful in rapidly evaluating the potential OSC of the candidate materials, DOSC provides better simulation of instantaneous oscillations between lean (oxidizing) and rich (reducing) exhaust

conditions during real operation and is therefore much more useful in the evaluation of the activity of OSC materials.<sup>21</sup> Fig. 3(a) shows the collected DOSC data and the corresponding transition curves at 320 °C of YBa(Co<sub>0.85</sub>Al<sub>0.075</sub>Ga<sub>0.075</sub>)<sub>4</sub>O<sub>7+δ</sub> as an example, with alternate dynamic pulses of 4% CO/1% Ar/He (10 s) and 2% O<sub>2</sub>/1% Ar/He (10 s) under 0.05 Hz given in Fig. 3(b). In comparison with CeO<sub>2</sub>-ZrO<sub>2</sub>, the normalized DOSC values by the BET surface areas of YBaCo<sub>4</sub>O<sub>7+δ</sub> and YBa(Co<sub>0.85</sub>Al<sub>0.075</sub>Ga<sub>0.075</sub>)<sub>4</sub>O<sub>7+δ</sub> are more than thirty times larger (DOSC values of YBaCo<sub>4</sub>O<sub>7+δ</sub>, YBa(Co<sub>0.85</sub>Al<sub>0.075</sub>Ga<sub>0.075</sub>)<sub>4</sub>O<sub>7+δ</sub> and Ce<sub>0.43</sub>Zr<sub>0.57</sub>O<sub>2</sub> at 400 °C are 40.9, 41.6 and 1.3 μmol [O] m<sup>-2</sup>, respectively),<sup>22</sup> confirming the fast responses between reduction and oxidation environments. Furthermore, the higher DOSC of YBa(Co<sub>0.85</sub>Al<sub>0.075</sub>Ga<sub>0.075</sub>)<sub>4</sub>O<sub>7+δ</sub> in comparison with that of YBaCo<sub>4</sub>O<sub>7+δ</sub> is consistent with the O<sub>2</sub>-TPD and H<sub>2</sub>-TPR results.

The catalytic activity for soot combustion was first checked by TPO (Fig. 4a). YBaCo<sub>4</sub>O<sub>7+δ</sub> and YBa(Co<sub>0.85</sub>Al<sub>0.075</sub>Ga<sub>0.075</sub>)<sub>4</sub>O<sub>7+δ</sub> decrease from a T<sub>10</sub> value of 530 °C for non-catalytic combustion to 387 and 379 °C, respectively, confirming the catalytic effect of the YBaCo<sub>4</sub>O<sub>7+δ</sub>-type material and the higher activity of the latter than the former. In terms of the selectivity towards CO<sub>2</sub> formation, the non-catalytic combustion is only 43.3%, while YBaCo<sub>4</sub>O<sub>7+δ</sub> and YBa(Co<sub>0.85</sub>Al<sub>0.075</sub>Ga<sub>0.075</sub>)<sub>4</sub>O<sub>7+δ</sub> yield nearly 100% CO<sub>2</sub>. After the TPO reactions, no phase decomposition occurs even for YBaCo<sub>4</sub>O<sub>7+δ</sub> (Fig. S4†), probably due to the high heating rate in 5 vol% oxygen in He.<sup>23</sup> Furthermore, the XRD peaks of YBaCo<sub>4</sub>O<sub>7+δ</sub> and YBa(Co<sub>0.85</sub>Al<sub>0.075</sub>Ga<sub>0.075</sub>)<sub>4</sub>O<sub>7+δ</sub> after TPO shift to higher angles, suggesting a lattice shrinkage, which confirms the participation of bulk oxygen.

The intrinsic activity was further demonstrated by kinetic rates at 300 °C, which can be obtained from the slope of the lines shown in Fig. 4b. As observed in Table 1, the specific rates of YBaCo<sub>4</sub>O<sub>7+δ</sub> and YBa(Co<sub>0.85</sub>Al<sub>0.075</sub>Ga<sub>0.075</sub>)<sub>4</sub>O<sub>7+δ</sub>, normalized by BET surface areas, are an order of magnitude larger than that of Ce<sub>0.43</sub>Zr<sub>0.57</sub>O<sub>2</sub>.<sup>19–22</sup> This is significant because 300 °C is a relevant temperature for light diesel engines. This particularly high reaction rate can ensure that the accumulated soot on the DPF can be readily combusted, leading to a lower balance point temperature (BPT) at which the rate of soot oxidation is matched with the rate of soot accumulation.<sup>24</sup> Furthermore, both the specific rate and TOF (Table 1) of YBa(Co<sub>0.85</sub>Al<sub>0.075</sub>Ga<sub>0.075</sub>)<sub>4</sub>O<sub>7+δ</sub> are a little higher than that of

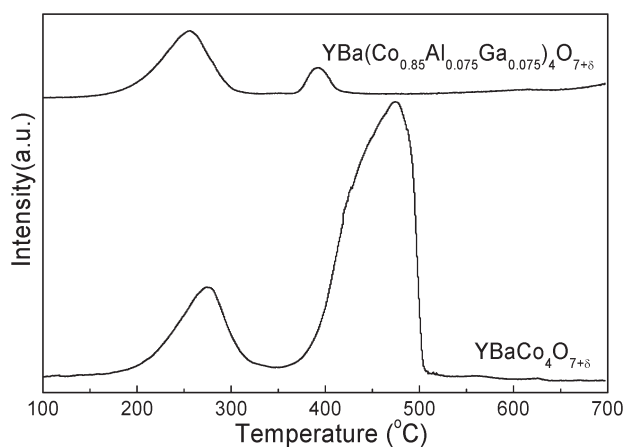


Fig. 2 H<sub>2</sub>-TPR spectra of YBaCo<sub>4</sub>O<sub>7+δ</sub> and YBa(Co<sub>0.85</sub>Al<sub>0.075</sub>Ga<sub>0.075</sub>)<sub>4</sub>O<sub>7+δ</sub> from room temperature to 700 °C.

Table 1 H<sub>2</sub>-TPR peak temperatures (T, °C), H<sub>2</sub> consumption (C, μmol g<sup>-1</sup>), specific rates at 300 °C normalized by BET surface areas (mol s<sup>-1</sup> m<sup>-2</sup> × 10<sup>-8</sup>), and TOF (s<sup>-1</sup> × 10<sup>-3</sup>) at 290 °C

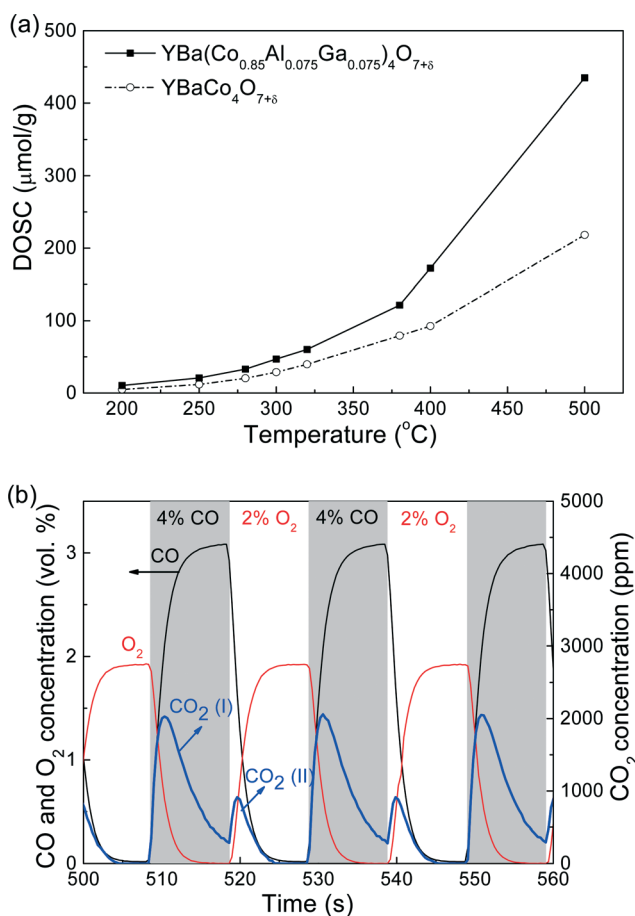
Sample	Peak 1		Peak 2		Specific rates <sup>a</sup>	TOF
	T	C	T	C		
YBaCo <sub>4</sub> O <sub>7+δ</sub>	275	74.9	475	298.5	2.11	1.34
YBa(Co <sub>0.85</sub> Al <sub>0.075</sub> Ga <sub>0.075</sub> ) <sub>4</sub> O <sub>7+δ</sub>	255	57.8	390	10.0	2.17	1.57

<sup>a</sup> The specific rate at 300 °C normalized by BET surface areas (80.4 m<sup>2</sup> g<sup>-1</sup>) for Ce<sub>0.43</sub>Zr<sub>0.57</sub>O<sub>2</sub> is 0.21 mol s<sup>-1</sup> m<sup>-2</sup> × 10<sup>-8</sup>.

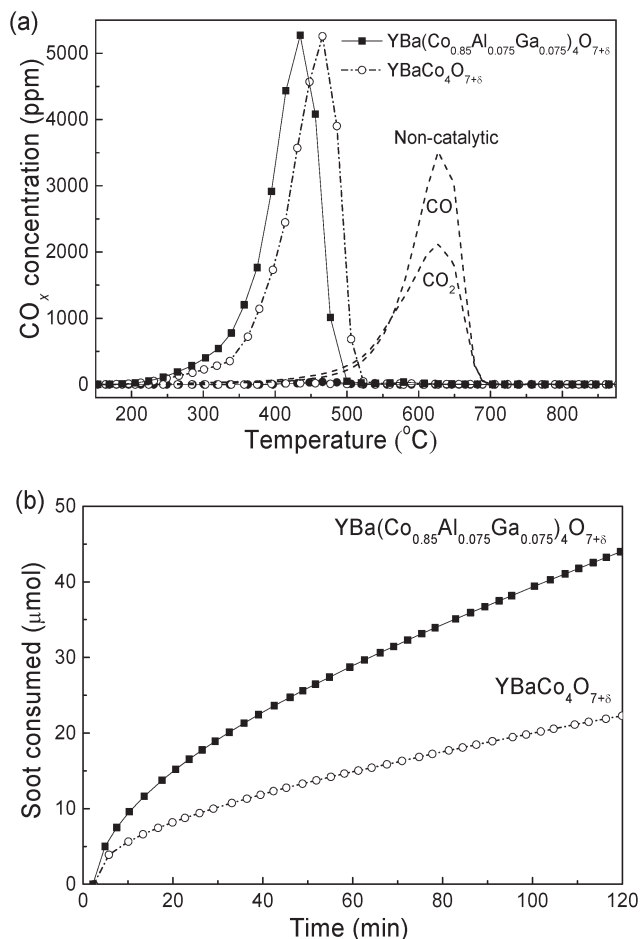


$\text{YBaCo}_4\text{O}_{7+\delta}$ , confirming the higher intrinsic activity of  $\text{YBa}(\text{Co}_{0.85}\text{Al}_{0.075}\text{Ga}_{0.075})_4\text{O}_{7+\delta}$ .

In order to explore the origin of the active oxygen, isotopic isothermal oxidation at 350 °C was performed (Fig. 5). Before switching from  $^{16}\text{O}_2$  to  $^{18}\text{O}_2$  (to the left of the dashed line), the main product was  $\text{C}^{16}\text{O}_2$ , confirming that the soot oxidation occurs with the bulk oxygen species. The concentration of  $\text{C}^{16}\text{O}_2$  for  $\text{YBa}(\text{Co}_{0.85}\text{Al}_{0.075}\text{Ga}_{0.075})_4\text{O}_{7+\delta}$  is much larger than that for  $\text{YBaCo}_4\text{O}_{7+\delta}$ , which is again consistent with the  $\text{O}_2$ -TPD,  $\text{H}_2$ -TPR, DOSC,  $T_{10}$  and specific rates. After switching from  $^{16}\text{O}_2$  to  $^{18}\text{O}_2$  (to the right of the dashed line), the sum of the products of  $\text{C}^{16}\text{O}^{18}\text{O}$ ,  $\text{C}^{18}\text{O}_2$  and  $^{16}\text{O}^{18}\text{O}$  still possesses higher concentrations of  $\text{YBa}(\text{Co}_{0.85}\text{Al}_{0.075}\text{Ga}_{0.075})_4\text{O}_{7+\delta}$  than that of  $\text{YBaCo}_4\text{O}_{7+\delta}$ . However, the concentration of  $\text{C}^{16}\text{O}_2$  decreased gradually to a very low value due to the depletion of  $^{16}\text{O}_2$  in the gaseous phase. Comparatively, for both  $\text{YBaCo}_4\text{O}_{7+\delta}$  and  $\text{YBa}(\text{Co}_{0.85}\text{Al}_{0.075}\text{Ga}_{0.075})_4\text{O}_{7+\delta}$ , the production of  $\text{C}^{16}\text{O}^{18}\text{O}$  increases and then reaches a maximum, while the  $\text{C}^{18}\text{O}_2$  production monotonically increases. This indicates that only the gaseous oxygen which has been activated by  $\text{YBaCo}_4\text{O}_{7+\delta}$  and  $\text{YBa}(\text{Co}_{0.85}\text{Al}_{0.075}\text{Ga}_{0.075})_4\text{O}_{7+\delta}$  can be used to



**Fig. 3** DOSC of  $\text{YBaCo}_4\text{O}_{7+\delta}$  and  $\text{YBa}(\text{Co}_{0.85}\text{Al}_{0.075}\text{Ga}_{0.075})_4\text{O}_{7+\delta}$  (a) and enlarged transition curves of  $\text{YBa}(\text{Co}_{0.85}\text{Al}_{0.075}\text{Ga}_{0.075})_4\text{O}_{7+\delta}$  (b) at 320 °C with dynamic pulses of 4% CO/1% Ar/He and 2%  $\text{O}_2$ /1% Ar/He under 0.05 Hz.  $\text{CO}_2(\text{I})$  was produced by CO and surface oxygen;  $\text{CO}_2(\text{II})$  was attributed to the reaction of absorbed CO and oxygen gas.<sup>22</sup>



**Fig. 4** TPO patterns of  $\text{CO}_x$  for soot combustion with  $\text{O}_2$  over  $\text{YBaCo}_4\text{O}_{7+\delta}$  and  $\text{YBa}(\text{Co}_{0.85}\text{Al}_{0.075}\text{Ga}_{0.075})_4\text{O}_{7+\delta}$  (a); isothermal reactions for soot combustion at 300 °C within the kinetic regime (b) under tight contact conditions between the soot and catalysts.

oxidize the soot. In addition, the desorption of  $^{16}\text{O}^{18}\text{O}$  is detected, which suggests that the desorbed  $^{16}\text{O}^{18}\text{O}$  species do not interact with the soot to produce the product containing carbon, demonstrating that the intimate contact between the soot and catalysts is essential. Since the high DOSC can make up for the missing oxygen, the intrinsic activity of  $\text{YBaCo}_4\text{O}_7$ -based oxides is much more active than that of  $\text{CeO}_2$ -based oxides.

## Conclusions

In conclusion, a new  $\text{YBaCo}_4\text{O}_{7+\delta}$ -type oxygen storage material was used to catalyze soot combustion. Isotopic isothermal reactions demonstrate the activation of gaseous oxygen and subsequent oxygen storage and reaction/desorption during the oxidation process. Higher activity and structural stability are achieved by the substitution of Co with Al and Ga to form  $\text{YBa}(\text{Co}_{0.85}\text{Al}_{0.075}\text{Ga}_{0.075})_4\text{O}_{7+\delta}$ . The specific rates at 300 °C of  $\text{YBaCo}_4\text{O}_{7+\delta}$  and  $\text{YBa}(\text{Co}_{0.85}\text{Al}_{0.075}\text{Ga}_{0.075})_4\text{O}_{7+\delta}$ , normalized by BET surface areas, are an order of magnitude larger than that of  $\text{CeO}_2$ -based oxides. This type of oxygen-storage material



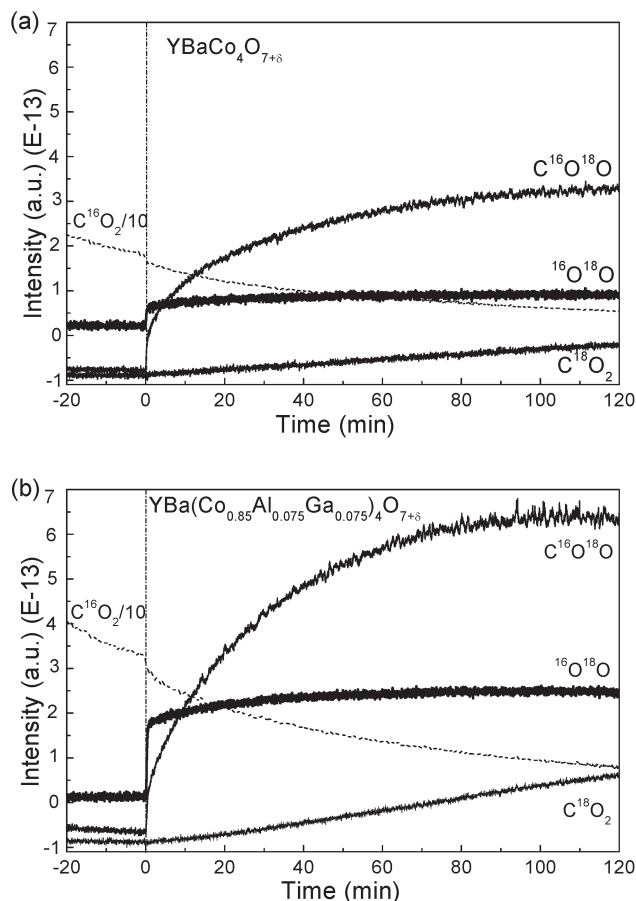


Fig. 5 Isothermal reactions for soot combustion at 350 °C after 1%  $^{16}\text{O}_2$  was switched to 1%  $^{18}\text{O}_2$  in He on  $\text{YBaCo}_4\text{O}_{7+\delta}$  (a) and  $\text{YBa}(\text{Co}_{0.85}\text{Al}_{0.075}\text{Ga}_{0.075})_4\text{O}_{7+\delta}$  (b).

provides an oxygen pool, which ensures that the accumulated soot on a DPF can be readily combusted.

## Acknowledgements

This work is financially supported by the National Natural Science Foundation of China (no. 21477046, 21277060 and 21547007).

## Notes and references

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