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Development of type-I/type-II hybrid dye sensitizer with both pyridyl group and catechol unit as anchoring group for type-I/type-II dye-sensitized solar cell†

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A type-I/type-II hybrid dye sensitizer with a pyridyl group and a catechol unit as the anchoring group has been developed and its photovoltaic performance in dye-sensitized solar cells (DSSCs) is investigated. The sensitizer has the ability to adsorb on a TiO₂ electrode through both the coordination bond at Lewis acid sites and the bidentate binuclear bridging linkage at Brønsted acid sites on the TiO₂ surface, which makes it possible to inject an electron into the conduction band of the TiO₂ electrode by the intramolecular charge-transfer (ICT) excitation (type-I pathway) and by the photoexcitation of the dye-to-TiO₂ charge transfer (DTCT) band (type-II pathway). It was found that the type-I/type-II hybrid dye sensitizer adsorbed on TiO₂ film exhibits a broad photoabsorption band originating from ICT and DTCT characteristics. Here we reveal the photophysical and electrochemical properties of the type-I/type-II hybrid dye sensitizer bearing a pyridyl group and a catechol unit, along with its adsorption modes onto TiO₂ film, and its photovoltaic performance in type-I/type-II DSSC, based on optical (photoabsorption and fluorescence spectroscopy) and electrochemical measurements (cyclic voltammetry), density functional theory (DFT) calculation, FT-IR spectroscopy of the dyes adsorbed on TiO₂ film, photocurrent–voltage (*I*–*V*) curves, incident photon-to-current conversion efficiency (IPCE) spectra, and electrochemical impedance spectroscopy (EIS) for DSSC.

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Introduction

Since Grätzel and co-workers produced high-performance dye-sensitized solar cells (DSSCs) employing a Ru-complex dye-adsorbed TiO₂ electrode in 1991, which showed a solar energy-to-electricity conversion yield (η) of 11%,¹ DSSCs have received considerable attention as one of the most promising new renewable photovoltaic cells alternative to conventional solar cells based on silicon because they have interesting construction and operational principles, and a colorful and decorative nature.^{2–8} To further improve the photovoltaic performance based on the construction and operational principles of DSSC, much effort has been made towards the development of various types of organic dye sensitizers that possess the following characteristics: (i) good light-harvesting features over the wide spectral region of sunlight, (ii) high electron-injection efficiencies from photoexcited dyes to the conduction band (CB) of the

TiO₂ electrode, (iii) high dye loading and high surface coverage of the TiO₂ electrode, (iv) efficient re-reduction of the oxidized dyes by electron transfer from I[−] ions in the electrolyte containing the I₃[−]/I[−] redox couple, (v) reduction of charge recombination between the injected electrons in the CB of the TiO₂ electrode and the oxidized dyes or I₃[−] ions in the electrolyte, and (vi) reduction of dye aggregation on the TiO₂ surface.^{3,4}

TiO₂-based DSSCs can be classified into two types, I and II, depending on the electron-injection pathway from the dye to the CB of the TiO₂ electrode (Fig. 1).⁴ The pathway for type-I DSSCs is photoexcitation of the local band of the dye adsorbed on the TiO₂ surface followed by an electron injection from the photoexcited dye to the CB of the TiO₂ electrode (Fig. 1a: *i.e.*, an electron is excited from the HOMO to the LUMO level of the dye by sunlight illumination, followed by injection into the CB of TiO₂ electrode). Thus, this pathway can also be called the “two-step” or “indirect” electron-injection pathway. To achieve efficient electron injection from the excited dye to the CB of the TiO₂ electrode, the LUMO energy level of the dye for type-I DSSC must be higher (more negative) than the energy level (E_{cb}) of the CB (−0.5 V *versus* the normal hydrogen electrode (NHE)) of the TiO₂ electrode (some researchers have proposed that an energy

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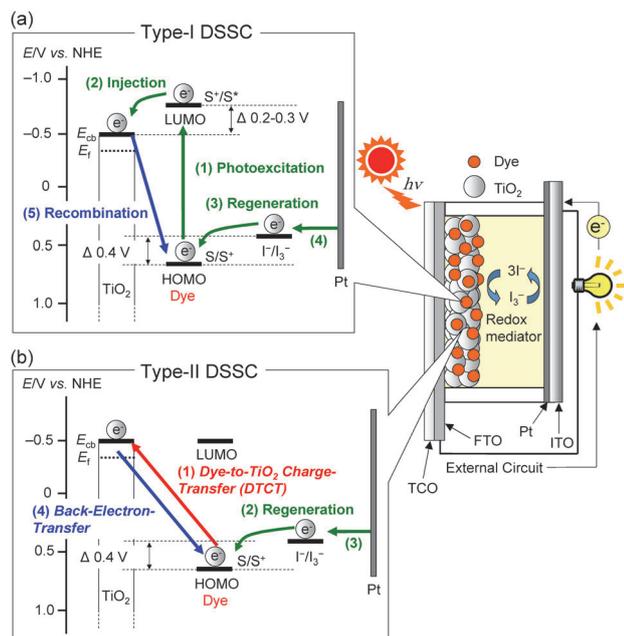


Fig. 1 Schematic representation of construction and operational principle of a typical TiO_2 -based DSSC (FTO: fluorine-doped SnO_2 , ITO: tin-doped indium oxide coated glass, TCO: transparent conductive oxide glass), and the pathway for electron injection from the dye to the CB of the TiO_2 electrode. (a) Type-I DSSC, in which an electron is excited from the HOMO to the LUMO level of the dye, followed by injection into the CB of the TiO_2 electrode. (b) Type-II DSSC, in which an electron is injected directly from the HOMO of the dye into the CB of the TiO_2 electrode upon photoexcitation.

gap of over 0.2–0.3 V is necessary for efficient electron injection).^{3–5} Almost all dye sensitizers developed so far for type-I DSSCs have carboxyl groups as anchoring groups, which are usually bound to the TiO_2 electrode through the bidentate bridging linkage between the carboxyl group of the dye and the Brønsted acid site (surface-bound hydroxyl groups, Ti-OH) on the TiO_2 surface (Fig. 2a). Ru complexes,^{1,9,10} porphyrin dyes,^{11–13} phthalocyanine dyes,¹⁴ and organic dyes^{3,4,6,15–17} bearing carboxyl groups have been known to inject electrons into the TiO_2 electrode, according to the type-I pathway. Among dye sensitizers bearing carboxyl groups used so far in type-I DSSCs, donor–acceptor– π -conjugated (D– π -A) dyes, which are constructed of a diphenyl- or dialkyl-amino group as the electron donor (D) and a carboxyl group as the electron acceptor (A) and anchor linked by π -conjugated bridges,

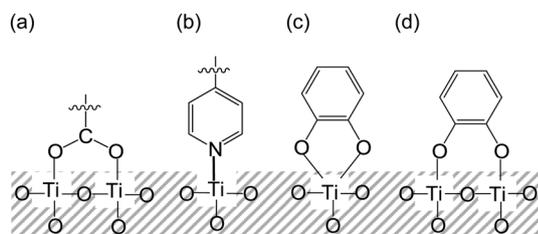


Fig. 2 Possible binding modes of carboxyl group, pyridyl group and catechol on TiO_2 surface. (a) A bidentate bridging linkage for carboxyl group, (b) a coordination bond for pyridyl group, (c) a bidentate mononuclear chelating linkage and (d) a bidentate dinuclear bridging linkage for catechol.

would be especially expected to be one of the most promising classes of organic dye sensitizers for type-I DSSCs because the wide variety of structures and the facile modification provide potential for molecular design, with the introduction of substituents onto the chromophore skeletons allowing easy control not only of their photophysical and electrochemical properties (the highest occupied molecular orbital (HOMO) and the lowest unoccupied molecular orbital (LUMO) energy levels), but also of their stereochemical structures. As a noteworthy structural feature of D– π -A dyes, the HOMO is localized over the π -conjugated system close to the D moiety, and the LUMO is localized over the A moiety.^{3,4,6} Therefore, in light of the D– π -A structure concept, the expansion of π -conjugation and the increase in the electron-donating and electron-accepting abilities of D moieties and A moieties, respectively, leads to the decrease in the energy gap between the HOMO and LUMO. Moreover, the photoabsorption properties of a D– π -A dye are associated with intramolecular charge transfer (ICT) excitation from the D to the A moiety of the dye, resulting in efficient electron transfer from the excited dye through the acceptor moiety (carboxyl group) into the CB of TiO_2 electrode. Moreover, D– π -A dyes can facilitate the ICT and the subsequent charge separation of the D and A moieties in the excited dye, and rapidly inject electrons into the CB of TiO_2 and effectively retard the charge recombination because the cationic charge on the D moiety of the excited dye is spatially separated by the π -conjugated bridge from the TiO_2 surface with the injected electrons.

On the other hand, we have reported that a new type of D– π -A dye sensitizers bearing a pyridyl group as an electron-withdrawing anchoring group are predominantly adsorbed on the TiO_2 electrode through the coordination bond between the pyridyl group of the dye and the Lewis acid site (exposed Ti^{IV} cations) on the TiO_2 surface (Fig. 2b).^{18,19} It was revealed that the new type of D– π -A dye sensitizers can inject electrons efficiently from the pyridyl group to the CB of the TiO_2 electrode through the coordination bond, compared to the bidentate bridging linkages of conventional D– π -A dye sensitizers bearing a carboxyl group. We demonstrated that the pyridyl group is a promising candidate as not only an electron-withdrawing anchoring group but also an electron-injecting group for type-I D– π -A dye sensitizers, compared to the carboxyl groups. Recently, some researchers reported that type-I DSSCs based on porphyrin dyes and D– π -A dyes bearing pyridyl groups.^{20,21} Consequently, much effort has been made on the development of various types of D– π -A organic dye sensitizers for type-I DSSCs and there has been a gradual accumulation of information about the relationship between the chemical structures and the photovoltaic performances of type-I DSSCs. As a result, type-I DSSCs have achieved an η value of up to 13%.^{12c,d} However, as mentioned above, to achieve electron injection from the excited dye to the CB of the TiO_2 electrode for type-I DSSCs, the LUMO levels of the type-I dye sensitizers must be higher than in the CB of the TiO_2 electrode. Thus, the restriction of the LUMO levels results in low light harvesting efficiency (LHE).

In contrast, the electron-injection pathway for type-II DSSCs is direct, that is, “one-step” electron injection from the ground



state of the dye to the CB of TiO₂ by photoexcitation of the dye-to-TiO₂ charge transfer (DTCT) bands (Fig. 1b: *i.e.*, an electron is injected directly from the HOMO of the dye into the CB of TiO₂ electrode upon photoexcitation).^{4b} Catechol (Cat) dyes, such as dopamine and fluorine, numerous natural pigments, such as bromopyrogallol red, and anthocyanins containing Cat moieties, which are classified as type-II dye sensitizers, show strong new absorption bands in the longer wavelength region corresponding to the DTCT upon binding to the TiO₂ surface.^{22–29} The Cat dyes have been known to bind to the TiO₂ surface through a bidentate mononuclear chelating linkage and/or a bidentate binuclear bridging linkage at the Brønsted acid site on TiO₂ (Fig. 2c and d). The most advantageous aspect of type-II DSSCs over type-I DSSCs is their light-harvesting capabilities over the wide spectral range of sunlight because the direct electron injection in type-II DSSCs can lead to the creation of a broad DTCT band and the easing of restrictions on the LUMO levels of the dye sensitizers relative to type-I DSSCs with an indirect electron-injection pathway, although the absorption coefficient of the DTCT band for type-II dye sensitizers is much lower than in the absorption band based on the $\pi \rightarrow \pi^*$ transition for conventional dye sensitizers due to lower overlap of the ground (molecular-localized) and the excited state (Ti-localized) wave functions.^{30a} However, the photovoltaic performance of type-II DSSCs based on Cat dyes is significantly lower than in type-I DSSCs based on D- π -A organic dye sensitizers. Fundamental studies for the interfacial electron transfer dynamics of Cat dyes adsorbed on TiO₂ nanoparticles have revealed that the back-electron transfer (charge recombination) of electrons injected into the TiO₂ electrode for the type-II pathway takes place in a sub-picosecond timescale, which was significantly faster than in the type-I pathway (μ s–ms). Thus, type-II Cat dyes have rarely been employed as sensitizers for DSSCs and thus there have been few attempts to develop type-II Cat dye sensitizers. In fact, the η values of type-II DSSCs based on Cat dyes had never exceeded 2.5%.^{30b} In an attempt to overcome this problem, Yoon and co-workers showed that attaching electron-donating moieties to the Cat unit can lead to retardation of the back-electron-transfer rate and a dramatic increase in electron-injection efficiency, and a red-shift of the DTCT band caused by the increase in the donor strength of the dye.^{27a} Their work inspired us to conceive the application of the molecular design of D- π -Cat dye sensitizers with strong electron donors, such as triphenylamine, indoline, and carbazole, which is one of the most promising strategies not only to inhibit back electron transfer, but also to enhance direct electron-injection efficiency in type-II DSSCs based on Cat dyes. Thus, to provide a direction in molecular design of Cat dye sensitizers for type-II DSSCs, in our previous work, we designed and synthesized a D- π -Cat dye, which shows a broad absorption band corresponding to the DTCT upon binding to TiO₂ film.^{29a} It was found that the photocurrent for the DSSC based on the D- π -Cat dye is mainly generated by a direct electron-injection pathway from the D- π -Cat dye to the CB of the TiO₂ electrode, that is, type-II DSSCs. Moreover, we have demonstrated that the stabilization

of the LUMO level and the expansion of the π -conjugated system by the introduction of a long π -bridge such as terthiophene on the Cat moiety can lead to an increase in the ICT excitation based on the π - π^* transition with a decrease in the DTCT characteristics, resulting in enhancement of an indirect electron injection pathway from the excited dye to the CB of the TiO₂ electrode by photoexcitation of the local band of the adsorbed dye on the TiO₂ surface. Interestingly, Nagata *et al.* reported that a DSSC based on 2-anthroic acid exhibits the type-II pathway with an η value of 2.2%.^{30b} Thus, the combination of the two injection (direct and indirect) mechanisms with some dye sensitizers has been investigated.^{24,30} To study the electronic injection mechanism in more depth and to generalize the main differences between the type-I and type-II pathways, Sanz and co-workers studied five organic dyes of increasing size, Cat, 2,3-naphthalenediol, alizarin, coumarin C343 having a carboxylic acid group, and coumarin derivative NKX-2311 having a cyanoacrylic acid group, and the electronic structures and optical response of the five dyes attached to the TiO₂ electrode were analyzed and compared by using time-dependent density functional theory (TDDFT).^{24b} From their analysis, NKX-2311 and Cat show purely indirect (type-I pathway) and direct (type-II pathway) electron-injection behavior, respectively. In the sequence from Cat to 2,3-naphthalenediol, C343, alizarin, and NKX-2311, the electron injection changes from a purely direct mechanism to a purely indirect mechanism. Alizarin and C343 represent intermediate behavior in which both injection regimes are present. Consequently, the electron injection mechanisms are related to the relative position of the LUMO energy of the dye to the edge of the CB of TiO₂ electrode.

Therefore, on the basis of the useful knowledge about the molecular design of D- π -A dye sensitizers with a pyridyl group for type-I DSSC and D- π -Cat dye sensitizers with a Cat unit for type-II DSSC obtained from our previous study,^{18,19,29} in this work, we have designed and developed a type-I/type-II hybrid dye sensitizer **OF-Py-Cat** with a pyridyl group and a catechol unit as the anchoring group possessing the ability to adsorb on the TiO₂ electrode through both a coordination bond at the Lewis acid sites and a bidentate binuclear bridging linkage at the Brønsted acid sites on the TiO₂ surface (Fig. 3). It would be expected that the type-I/type-II hybrid dye sensitizer makes it possible to inject electrons into the CB of the TiO₂ electrode by the ICT excitation (type-I pathway) and by the photoexcitation of the DTCT band (type-II pathway). Furthermore, not only to gain insight into the influence of the molecular structure of type-I/type-II hybrid dye sensitizer on the appearance of ICT and DTCT bands and the electron-injection mechanism, but also to investigate the impacts of the ICT and DTCT characteristics of the type-I/type-II hybrid dye sensitizer on the photovoltaic performances of DSSCs, we also synthesized a type-I dye sensitizer **OF-Car-Car** with two carboxyl groups, a type-I dye sensitizer **OF-Py-Car** with a pyridyl group and a carboxyl group, a type-II dye sensitizer **OF-Cat-Cat** with two catechol units, as well as a type-I dye sensitizer **OF-Py-Py** (**YNI-2**) with two pyridyl groups.^{18d} It was found that the type-I/type-II hybrid dye sensitizer adsorbed on TiO₂ film exhibits a broad photoabsorption



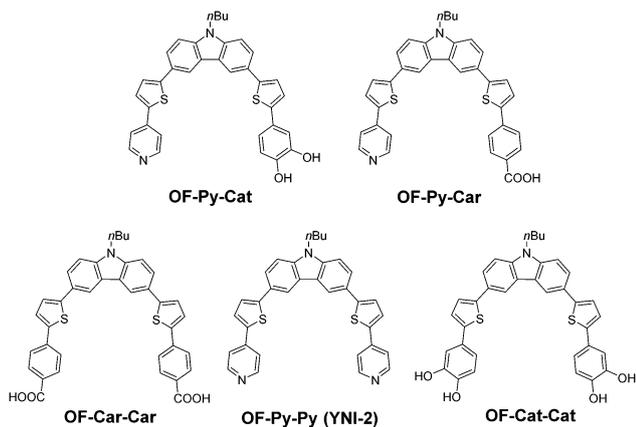


Fig. 3 Chemical structures of **OF-Py-Cat**, **OF-Py-Car**, **OF-Car-Car**, **OF-Py-Py (YNI-2)** and **OF-Cat-Cat**.

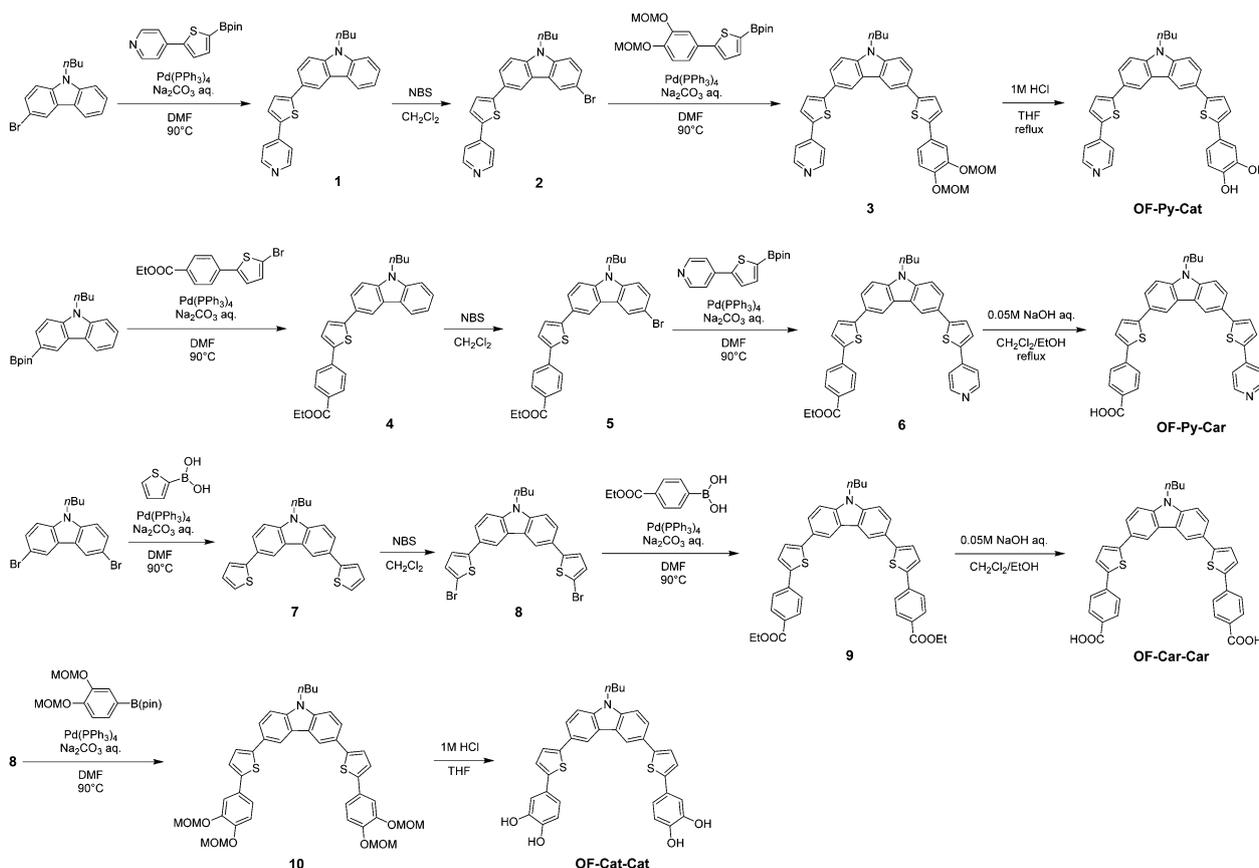
band originating from the ICT and DTCT characteristics. Here we reveal the photophysical and electrochemical properties of the type-I/type-II hybrid dye sensitizers bearing a pyridyl group and a Cat unit, along with their adsorption mode onto TiO_2 film, and their photovoltaic performances in type-I/type-II DSSC, based on optical (photoabsorption and fluorescence spectroscopy) and electrochemical measurements (cyclic voltammetry), density functional theory (DFT) calculation, FT-IR spectroscopy of the dyes adsorbed on TiO_2 film, photocurrent–voltage (I – V)

curves, incident photon-to-current conversion efficiency (IPCE) spectra, and electrochemical impedance spectroscopy (EIS) for DSSC.

Results and discussion

Synthesis

The type-I dye sensitizer **OF-Py-Py (YNI-2)**^{18d} with two pyridyl groups was prepared by Stille coupling of 3,6-dibromo-9-butyl-9H-carbazole with 4-(5-tributylstannanyl-thiophen-2-yl)-pyridine. In this work, to develop a new synthetic route to D- π -A dye sensitizers with two anchoring groups, we adopted a synthetic strategy utilizing Suzuki coupling for the synthesis of the D- π -A dye sensitizers employed (Scheme 1). Indeed, the type-I/type-II hybrid dye sensitizer **OF-Py-Cat** and the type-I dye sensitizer **OF-Py-Car** with two different anchoring groups were successfully prepared by using Suzuki coupling twice to introduce the anchoring (electron acceptor) moiety, where 3-bromo-9-butyl-9H-carbazole³¹ and (9-butyl-9H-carbazol-3-yl)boronic acid pinacol ester,³¹ respectively, as starting materials were employed. Moreover, 3,6-bis(5-bromothiophen-2-yl)-9-butyl-9H-carbazole **8** was synthesized as a useful intermediate capable of introducing anchoring groups, that is, the corresponding anchoring unit was introduced by Suzuki coupling of **8** with 4-ethoxycarbonylphenylboronic acid and (3,4-bis(methoxymethoxy)phenyl)boronic acid pinacol



Scheme 1 Synthesis of **OF-Py-Cat**, **OF-Py-Car**, **OF-Car-Car** and **OF-Cat-Cat**.



ester to give **9** and **10**, respectively, which is then treated with the appropriate alkali and acid, respectively, to provide the type-I dye sensitizer **OF-Car-Car** with two carboxyl groups and the type-II dye sensitizer **OF-Cat-Cat** with two catechol units.

Photophysical properties in solution

The UV/vis absorption and fluorescence spectra of **OF-Car-Car**, **OF-Py-Py**, **OF-Cat-Cat**, **OF-Py-Car** and **OF-Py-Cat** in THF are shown in Fig. 4 and their spectral data are summarized in Table 1. The five dyes show a bimodal absorption band with peaks at around 350 and 380 nm (Fig. 4a). Interestingly, for **OF-Car-Car**, **OF-Py-Py** and **OF-Py-Car** with two electron-withdrawing anchoring groups the molar extinction coefficient (ϵ) of the absorption band at around 380 nm is higher than that of the absorption band at around 350 nm, but for **OF-Cat-Cat** and **OF-Py-Cat** with a catechol unit as the relatively electron-donating anchoring group, the ϵ of the absorption band at around 380 nm is lower than in the absorption band at around 350 nm. Moreover, the absorption band at around 380 nm for **OF-Cat-Cat** is blue-shifted by *ca.* 10 nm compared with that of **OF-Car-Car**, **OF-Py-Py** and **OF-Py-Car**. Thus, for the four dyes **OF-Car-Car**, **OF-Py-Py**, **OF-Py-Car** and **OF-Py-Cat** with an electron-withdrawing anchoring group except **OF-Cat-Cat**, this result shows that the intense absorption band at the longest wavelength corresponds to the intramolecular charge transfer (ICT) excitation from the carbazole moiety for **OF-Car-Car**, **OF-Py-Py** and **OF-Py-Car** and the carbazole containing a thienylcatechol moiety for **OF-Py-Cat**, respectively, to the electron-withdrawing

anchoring group. The corresponding fluorescence spectra (Fig. 4b) show that only the dye **OF-Cat-Cat** exhibits a vibronic-structured fluorescence band ($\lambda_{\text{em}} = 412$ nm), which occurs at a shorter wavelength by *ca.* 35–55 nm than in the other four dyes ($\lambda_{\text{fl}} = \text{ca. } 445\text{--}465$ nm). Moreover, the dye **OF-Cat-Cat** ($\Phi_{\text{f}} = 0.16$) exhibits a lower fluorescence quantum yield (Φ_{f}) than those of the other four dyes ($\Phi_{\text{f}} = \text{ca. } 0.5\text{--}0.7$).

Photophysical properties of dye-adsorbed TiO₂ film

The photoabsorption spectra of the dyes adsorbed on the TiO₂ film are shown in Fig. 5a. It is worth noting that in addition to the ICT band observed in THF, **OF-Cat-Cat** and **OF-Py-Cat** adsorbed on the TiO₂ film show a broad absorption band corresponding to the DTCT upon binding to the TiO₂ film, that is, the DTCT band appears in the region of 430 to 650 nm for **OF-Cat-Cat** and 450 to 650 nm for **OF-Py-Cat**. The red-shift of the absorption peak wavelengths for dyes with a catechol unit upon binding to the TiO₂ film can be attributed to the stabilization of the LUMO level of the Ti-Cat dye complexes, which has been demonstrated by DFT calculations in our previous work.^{29a} As shown in Fig. 5b, the colors of the dyes in THF are light yellow or yellow for **OF-Car-Car**, **OF-Py-Py**, **OF-Py-Car** and **OF-Py-Cat**, and light purple for **OF-Cat-Cat**. For **OF-Car-Car**, **OF-Py-Py** and **OF-Py-Car**, the colors of dyes adsorbed on the TiO₂ film are similar to those in THF. In contrast, the colors of **OF-Cat-Cat** and **OF-Py-Cat** adsorbed on the TiO₂ film are almost black and dark-orange, respectively, which indicates a decrease in the transition energy by the appearance of the DTCT band associated with the formation of the surface-bound Ti-Cat complex. Thus, this result reveals that the dye **OF-Py-Cat** adsorbed on the TiO₂ film exhibits a broad photoabsorption band originating from the ICT and DTCT characteristics, that is, the dye **OF-Py-Cat** can act as a type-I/type-II hybrid dye sensitizer.

Electrochemical properties

The electrochemical properties of the five dyes were determined by cyclic voltammetry (CV) in DMF containing 0.1 M tetrabutylammonium perchlorate (Bu₄NClO₄) for **OF-Car-Car**, **OF-Cat-Cat**, **OF-Py-Car** and **OF-Py-Cat** and THF containing 0.1 M Bu₄NClO₄ for **OF-Py-Py**^{18d} (Fig. 6). The potentials were referenced to ferrocene/ferrocenium (Fc/Fc⁺) as the internal reference. The oxidation waves for **OF-Car-Car**, **OF-Py-Py** and

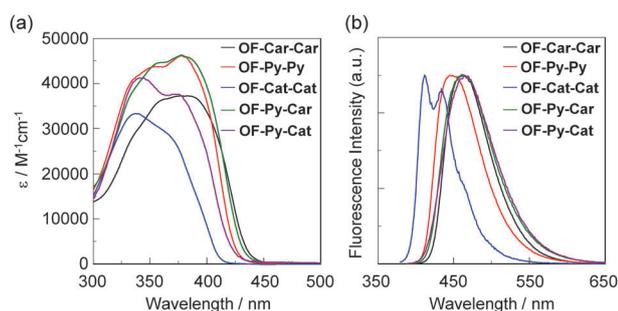


Fig. 4 (a) UV/vis absorption and (b) fluorescence spectra of **OF-Car-Car**, **OF-Py-Py**, **OF-Cat-Cat**, **OF-Py-Car** and **OF-Py-Cat** in THF.

Table 1 Optical and electrochemical data, HOMO and LUMO energy levels, and DSSC performance parameters of **OF** dye sensitizers

Dye	$\lambda_{\text{max}}^{\text{abs}}/\text{nm}$ ($\epsilon/\text{M}^{-1}\text{cm}^{-1}$) ^a	$\lambda_{\text{fl}}/\text{nm}$ (Φ_{f}) ^b	$E_{1/2}^{\text{ox}}$ or $E_{\text{onset}}^{\text{ox}}/\text{V}^{\text{c}}$	HOMO/V ^d	LUMO/V ^d	Molecules/cm ⁻² ^e	$J_{\text{sc}}/\text{mA cm}^{-2}$ ^f	$V_{\text{oc}}/\text{mV}^{\text{f}}$	ff ^f	η (%) ^f
OF-Car-Car	383 (37300)	461 (0.56)	0.55	1.27	-1.63	9.1×10^{16}	4.60	480	0.62	1.37
OF-Py-Py	378 (46100)	448 (0.67)	0.53	1.22	-1.75	1.0×10^{17}	5.88	542	0.64	2.04
OF-Cat-Cat	370 (27700)	412 (0.16)	0.28	1.00	-2.12	9.2×10^{16}	0.47	304	0.49	0.07
OF-Py-Car	378 (46200)	464 (0.69)	0.55	1.27	-1.65	1.1×10^{17}	5.28	485	0.62	1.51
OF-Py-Cat	373 (37600)	465 (0.51)	0.15	0.87	-2.05	1.2×10^{17}	0.40	299	0.48	0.06

^a In 1,4-dioxane. ^b Fluorescence quantum yields (Φ_{f}) were determined by using a calibrated integrating sphere system ($\lambda_{\text{ex}} = 383, 378, 370, 378$ and 373 nm for **OF-Car-Car**, **OF-Py-Py**, **OF-Cat-Cat**, **OF-Py-Car**, and **OF-Py-Cat**, respectively). ^c Half-wave potentials versus Fc/Fc⁺ of oxidation ($E_{1/2}^{\text{ox}}$) for **OF-Py-Py** were recorded in THF/Bu₄NClO₄ (0.1 M) solution. Onset ($E_{\text{onset}}^{\text{ox}}$) versus Fc/Fc⁺ of the oxidation potential for **OF-Car-Car**, **OF-Cat-Cat**, **OF-Py-Car**, and **OF-Py-Cat** were recorded in DMF/Bu₄NClO₄ (0.1 M) solution. ^d Versus normal hydrogen electrode (NHE). ^e Adsorption amount of dye molecules per unit area of the TiO₂ electrode, when the 9 μm thick TiO₂ electrode was immersed into 0.1 mM dye solution in THF. ^f Under a simulated solar light (AM 1.5, 100 mW cm⁻²).



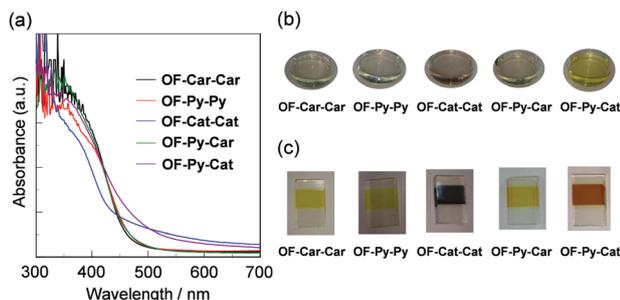


Fig. 5 (a) Photoabsorption spectra of **OF-Car-Car**, **OF-Py-Py**, **OF-Cat-Cat**, **OF-Py-Car** and **OF-Py-Cat** adsorbed on TiO_2 film. Photographs of the dyes (b) in THF and (c) adsorbed on TiO_2 film.

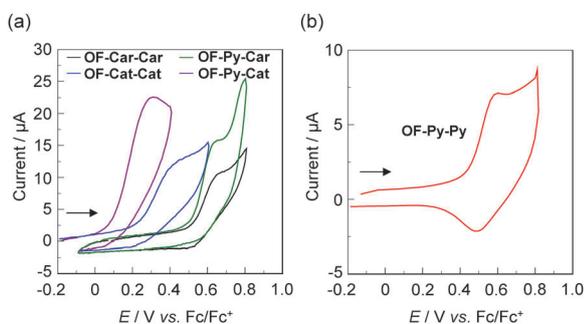


Fig. 6 Cyclic voltammograms of (a) **OF-Car-Car**, **OF-Cat-Cat**, **OF-Py-Car** and **OF-Py-Cat** in DMF containing 0.1 M Bu_4NClO_4 and (b) **OF-Py-Py** in THF containing 0.1 M Bu_4NClO_4 . The arrow denotes the direction of the potential scan.

OF-Py-Car were observed at around 0.6 V *versus* Fc/Fc^+ , whereas the oxidation waves for **OF-Cat-Cat** and **OF-Py-Cat** were observed at around 0.4 V and 0.3 V, respectively, which are cathodically shifted by 0.2 V and 0.3 V, respectively, compared with those of **OF-Car-Car**, **OF-Py-Py** and **OF-Py-Car**. The corresponding reduction wave was clearly observed at 0.48 V for **OF-Py-Py**, whereas a clear reduction wave was not observed for **OF-Car-Car**, **OF-Cat-Cat**, **OF-Py-Car** and **OF-Py-Cat**. These results show that the redox process of **OF-Py-Py** is electrochemically reversible, but those of **OF-Car-Car**, **OF-Cat-Cat**, **OF-Py-Car** and **OF-Py-Cat** are electrochemically irreversible. The HOMO energy levels of the dyes were estimated from the onset ($E_{\text{onset}}^{\text{ox}}$) of the oxidation potential for **OF-Car-Car** (0.55 V), **OF-Cat-Cat** (0.28 V), **OF-Py-Car** (0.55 V) and **OF-Py-Cat** (0.15 V), and the half-wave potential ($E_{1/2}^{\text{ox}} = 0.53$ V) for **OF-Py-Py** (Table 1). The HOMO energy levels (*versus* NHE) of **OF-Cat-Cat** (1.00 V) and **OF-Py-Cat** (0.87 V) are higher than in **OF-Car-Car**, **OF-Py-Py** and **OF-Py-Car** (*ca.* 1.25 V). This result shows that the HOMO energy levels are more positive than the I_3^-/I^- redox potential (0.4 V), and thus this ensures an efficient regeneration of the oxidized dyes by electron transfer from the I_3^-/I^- redox couple in the electrolyte. The LUMO energy levels of the dyes were estimated from the $E_{\text{onset}}^{\text{ox}}$ or the $E_{1/2}^{\text{ox}}$ and an intersection of the UV/vis absorption and fluorescence spectra (428 nm; 2.90 eV for **OF-Car-Car**, 417 nm; 2.97 eV for **OF-Py-Py**, 398 nm; 3.12 eV for **OF-Cat-Cat**, 424 nm; 2.92 eV for **OF-Py-Car** and 418 nm; 2.97 eV for **OF-Py-Cat**).

The LUMO energy levels of **OF-Cat-Cat** (−2.12 V) and **OF-Py-Cat** (−2.05 V) are also higher than in **OF-Car-Car**, **OF-Py-Py** and **OF-Py-Car** (*ca.* −1.70 V). Therefore, this result shows that from **OF-Car-Car**, **OF-Py-Py** and **OF-Py-Car** to **OF-Cat-Cat**, that is, by replacing the electron-withdrawing carboxyl and pyridyl groups with the electron-donating catechol unit as the anchoring group, the elevation of the LUMO energy level is larger than that of the HOMO energy level. Consequently, it was revealed that the blue-shift of the ICT absorption band for **OF-Cat-Cat** relative to **OF-Car-Car**, **OF-Py-Py** and **OF-Py-Car** is attributed to the destabilization of the LUMO energy level by the introduction of two catechol units with electron-donating ability, resulting in an increase in the HOMO–LUMO band gap. Evidently, the LUMO energy levels of the five dyes are higher than the E_{cb} of the CB of TiO_2 (−0.5 V) electrode, suggesting that an electron injection from the photoexcited dye to the CB of TiO_2 electrode for type-I DSSCs is thermodynamically feasible.

Theoretical calculations

To examine the electronic structures of **OF-Car-Car**, **OF-Py-Py**, **OF-Cat-Cat**, **OF-Py-Car** and **OF-Py-Cat**, the molecular orbitals of these dyes were calculated using density functional theory (DFT) at the B3LYP/6-31G(d,p) level.³³ The DFT calculations indicate that for **OF-Car-Car**, **OF-Py-Py** and **OF-Py-Car** the HOMOs are mostly localized on the carbazole moiety containing a thiophene ring (Fig. 7). The LUMOs for **OF-Car-Car** and **OF-Py-Py** are mostly localized on the two thienylbenzoic acid moieties and the two thienylpyridine moieties, respectively. On the other hand, for **OF-Py-Car** the LUMO is mostly localized on the thienylbenzoic acid moiety, and the LUMO+1 is mostly localized on the thienylpyridine moiety. Interestingly, for **OF-Py-Cat** the HOMO is mostly localized on the carbazole containing a thienylcatechol moiety and a thiophene ring, and the LUMO is mostly localized on the thienylpyridine moiety. Moreover, for **OF-Cat-Cat** the HOMO is delocalized over the whole molecule but the LUMO is mostly localized on the carbazole moiety containing the two thiophene rings. Accordingly, the DFT calculations reveal that for the four dyes with an electron-withdrawing anchoring group, except **OF-Cat-Cat**, the dye excitations upon light irradiation induce a strong ICT from the carbazole moiety to the electron-withdrawing anchoring groups.

FTIR spectra

To elucidate the adsorption states of the five dyes on TiO_2 nanoparticles, we measured the FTIR spectra of the dye powders and the dyes adsorbed on TiO_2 film. For the dye powders of **OF-Car-Car** and **OF-Py-Car** (Fig. 8a and d), the C=O stretching band of the carboxyl group was observed at 1678 cm^{-1} for **OF-Car-Car** and 1695 cm^{-1} for **OF-Py-Car**. When the two dyes were adsorbed on the TiO_2 surface, the C=O stretching bands of the carboxyl group disappeared completely; this indicates the formation of a bidentate bridging linkage between the carboxyl group of the dye and the Brønsted acid site on the TiO_2 surface. In addition, for the dye powder of **OF-Py-Car** the C=N stretching band of the pyridyl group was clearly observed at 1601 cm^{-1} . In the FTIR spectrum of **OF-Py-Car** adsorbed on the TiO_2 film, a new band appeared at around 1612 cm^{-1} , which can be



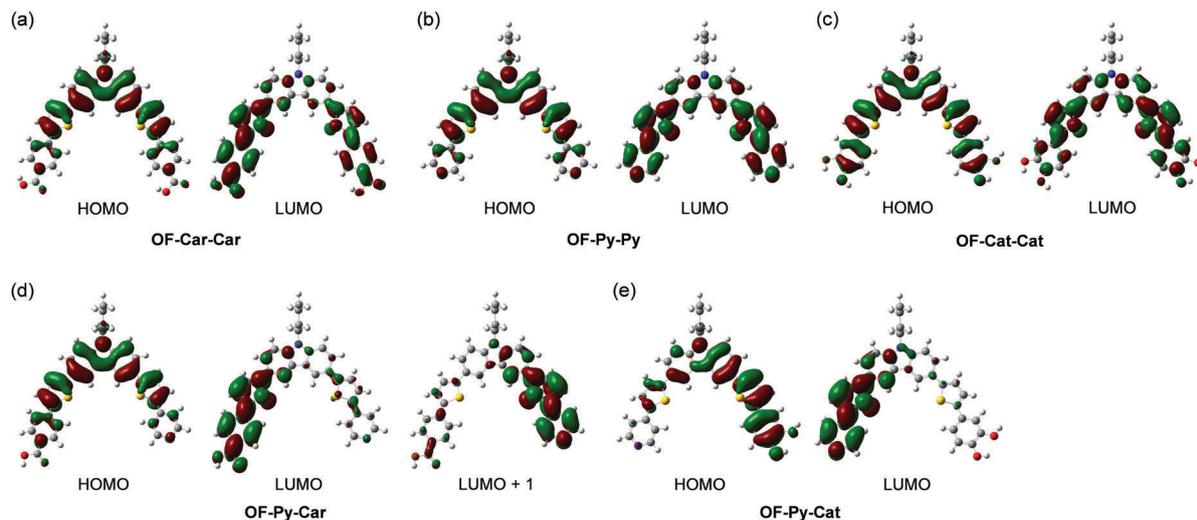


Fig. 7 HOMO and LUMO of (a) **OF-Car-Car**, (b) **OF-Py-Py**, (c) **OF-Cat-Cat**, (d) **OF-Py-Car** and (e) **OF-Py-Cat** derived from the DFT calculations at the B3LYP/6-31G(d,p) level.

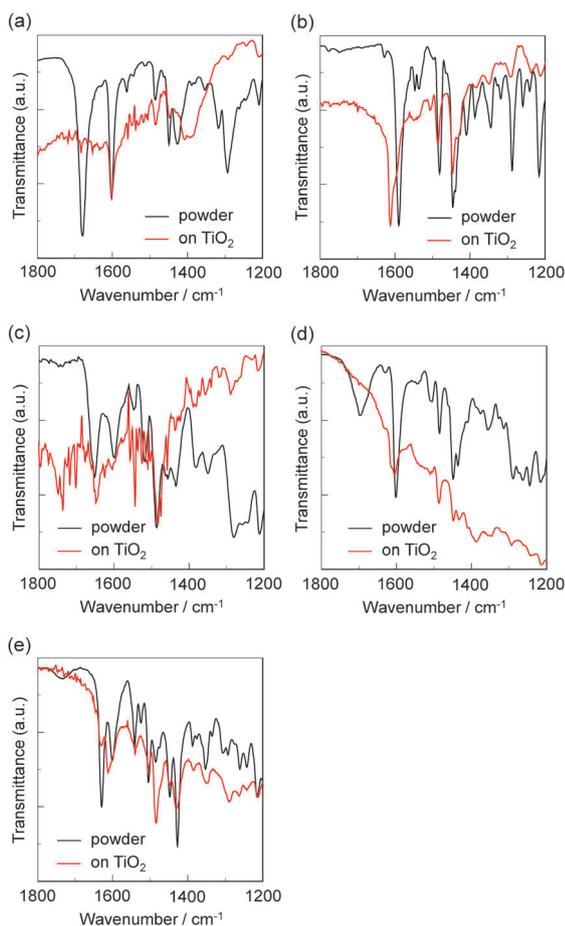


Fig. 8 FTIR spectra of the powders and the dyes adsorbed on TiO_2 nanoparticles. (a) **OF-Car-Car**, (b) **OF-Py-Py**, (c) **OF-Cat-Cat**, (d) **OF-Py-Car** and (e) **OF-Py-Cat**.

assigned to the pyridyl group coordinated to the Lewis acid sites of the TiO_2 surface. As reported previously, the FTIR spectrum of **OF-Py-Py** adsorbed on TiO_2 films demonstrated

that the dye **OF-Py-Py** is predominantly adsorbed on the TiO_2 surface through the coordination bond between the pyridyl group of the dye and the Lewis acid site on the TiO_2 surface, that is, the fact is based on a new band at around 1614 cm^{-1} in the FTIR spectrum of **OF-Py-Py** adsorbed on the TiO_2 film (Fig. 8b).^{18d} Consequently, the dye **OF-Py-Car** with a pyridyl group and a carboxyl group was adsorbed on the TiO_2 surface through both the bidentate bridging linkage between the carboxyl group of the dye and the Brønsted acid site of the TiO_2 surface and the coordination bond between the pyridyl group of the dye and the Lewis acid site of the TiO_2 surface. On the other hand, for the dye powders of **OF-Cat-Cat** and **OF-Py-Cat** (Fig. 8c and e), the stretching vibrations of the phenolic group (C–OH) for the catechol unit was observed at 1280 , 1256 and 1242 cm^{-1} for **OF-Cat-Cat** and 1306 , 1292 , 1260 and 1242 cm^{-1} for **OF-Py-Cat**. In the FTIR spectra of dyes adsorbed on TiO_2 film, the stretching vibrations merge to one prominent signal or lose their hyperfine structure. Thus, these observations indicate that the two dyes are adsorbed on the TiO_2 surface through the formation of a bidentate mononuclear chelating linkage and/or a bidentate binuclear bridging linkage between the Cat unit of dye and the TiO_2 surface.²⁸ In addition, for the FTIR spectrum of **OF-Py-Cat** adsorbed on the TiO_2 film, a new band appeared at around 1612 cm^{-1} , which can be assigned to the pyridyl group coordinated to the Lewis acid sites of the TiO_2 surface. Therefore, the dye **OF-Py-Cat** with a pyridyl group and a catechol unit was adsorbed on the TiO_2 surface through both the bidentate linkage between the catechol unit of the dye and the Brønsted acid site of the TiO_2 surface and the coordination bond between the pyridyl group of the dye and the Lewis acid site of the TiO_2 surface.

Photovoltaic performances

The DSSCs based on **OF-Car-Car**, **OF-Py-Py**, **OF-Cat-Cat**, **OF-Py-Car** and **OF-Py-Cat** were fabricated by using the dye-adsorbed TiO_2 electrode (9 mm), Pt-coated glass as a counter electrode, and an acetonitrile solution with iodine (0.05 M), lithium iodide (0.1 M),



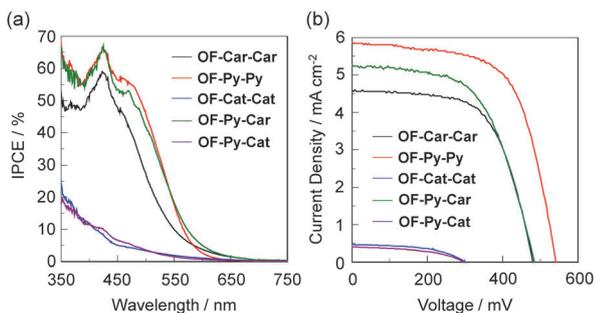


Fig. 9 (a) IPCE spectra and (b) I - V curves of DSSCs based on **OF-Car-Car**, **OF-Py-Py**, **OF-Cat-Cat**, **OF-Py-Car** and **OF-Py-Cat**.

and 1,2-dimethyl-3-propylimidazolium iodide (0.6 M) as the electrolyte. The photocurrent-voltage (I - V) characteristics were measured under simulated solar light (AM 1.5, 100 mW cm⁻²). The IPCE spectra and the I - V curves are shown in Fig. 9. The photovoltaic performance parameters are collected in Table 1. It is worth mentioning here that the adsorption amount of all the five dyes bearing two anchoring groups adsorbed on the TiO₂ electrode is *ca.* 1.0×10^{17} molecules per cm², which is comparable to that of conventional dye sensitizers bearing one anchoring group. The IPCE spectra of the DSSCs based on these dyes demonstrated that the IPCE of **OF-Car-Car**, **OF-Py-Py** and **OF-Py-Car** are much higher than those of **OF-Cat-Cat** and **OF-Py-Cat** (Fig. 9a). Interestingly, the maximum IPCE values of **OF-Py-Py** (66% at 426 nm) and **OF-Py-Car** (66% at 428 nm) are somewhat higher than in **OF-Car-Car** (59% at 423 nm). For the DSSCs based on **OF-Cat-Cat** and **OF-Py-Cat** the DTCT bands make surprisingly little contribution to the IPCE spectra, that is, the IPCE originating from the DTCT band, which is below 10% in the region of 450 to 600 nm, is observed. The I - V curves show that the short circuit photocurrent density (J_{sc}) and η increase in the order of **OF-Py-Cat** (0.40 mA cm⁻², 0.06%) \approx **OF-Cat-Cat** (0.47 mA cm⁻², 0.07%) < **OF-Car-Car** (4.60 mA cm⁻², 1.37%) < **OF-Py-Car** (5.28 mA cm⁻², 1.51%) < **OF-Py-Py** (5.88 mA cm⁻², 2.04%) (Fig. 9b). Thus, this result demonstrates that the photovoltaic performances of DSSC based on **OF-Py-Car** and **OF-Py-Py** with at least a pyridyl group are higher than in **OF-Car-Car** with only carboxyl group, which is attributed to efficient electron injection from the pyridyl group to the CB of the TiO₂ electrode through the coordination bond at the Lewis acid sites on the TiO₂ surface, compared to the carboxyl group through the bidentate bridging linkages at the Brønsted acid sites on the TiO₂ surface. On the other hand, the significantly low photovoltaic performance of DSSC based on **OF-Cat-Cat** with the two catechol units is due to the facilitation of the back-electron transfer from the electrons injected into TiO₂ to the oxidized dye. Moreover, the most important finding to be emphasized is that the totally unexpected result for DSSC based on **OF-Py-Cat** with both a pyridyl group and a catechol unit may be due to significantly faster back-electron transfer, in spite of the efficient electron injection from the pyridyl group to the CB of the TiO₂ electrode.

On the other hand, it is also worth mentioning here that the open-circuit photovoltage (V_{oc}) values of DSSC based on **OF-Cat-Cat**

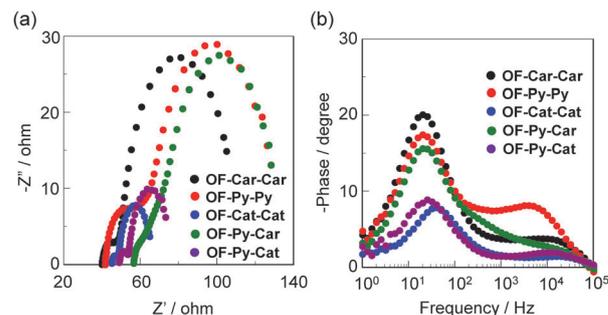


Fig. 10 (a) Nyquist plots and (b) Bode phase plots of DSSCs based on **OF-Car-Car**, **OF-Py-Py**, **OF-Cat-Cat**, **OF-Py-Car** and **OF-Py-Cat**.

(304 mV) and **OF-Py-Cat** (299 mV) with a catechol unit are much lower than in **OF-Car-Car** (480 mV), **OF-Py-Py** (542 mV) and **OF-Py-Car** (485 mV). Thus, electrochemical impedance spectroscopy (EIS) analysis was performed to study the electron recombination process in DSSCs based on the dyes in the dark under a forward bias of -0.35 or -0.55 V with a frequency range of 1 Hz to 100 kHz. The large semicircle in the Nyquist plot (Fig. 10a), which corresponds to the midfrequency peaks in the Bode phase plots, represents the charge recombination between the injected electrons into the TiO₂ electrode and I₃⁻ ions in the electrolyte, that is, the charge-transfer resistances at the TiO₂/dye/electrolyte interface. The Nyquist plots show that the electron recombination (R_{rec}) value of the large semicircle for **OF-Car-Car** (64 Ω), **OF-Py-Py** (68 Ω) and **OF-Py-Car** (50 Ω) is higher than in **OF-Cat-Cat** (16 Ω) and **OF-Py-Cat** (20 Ω), indicating that the DSSCs with high V_{oc} values have a tendency to exhibit a large R_{rec} value. The electron recombination lifetimes (τ_e), which are an expression of the electron recombination between the injected electrons into TiO₂ and I₃⁻ ions in the electrolyte, are extracted from the angular frequency (ω_{re}) at the mid-frequency peak in the Bode phase plot (Fig. 10b) using $\tau_e = 1/\omega_{re}$. The τ_e value (8 ms) for **OF-Car-Car**, **OF-Py-Py** and **OF-Py-Car** is higher than in **OF-Cat-Cat** (4 ms) and **OF-Py-Cat** (6 ms), that is, the τ_e values have the same tendency as the R_{rec} values. This result revealed that the charge recombination between the injected electrons into the TiO₂ electrode and I₃⁻ ions in the electrolyte may be one of the major reasons for the difference in V_{oc} value between the DSSC based on the five dyes, that is, the dyes **OF-Car-Car**, **OF-Py-Py** and **OF-Py-Car** can effectively retard the charge recombination compared with **OF-Cat-Cat** and **OF-Py-Cat**, but the decrease in the number of electrons injected into the CB of the TiO₂ electrode results in a lower V_{oc} value for the DSSC based on **OF-Cat-Cat** and **OF-Py-Cat**.^{3,4} Consequently, on the basis of the photophysical and electrochemical properties, the dye-adsorption modes onto TiO₂ film and the photovoltaic performances for **OF-Car-Car**, **OF-Py-Py**, **OF-Cat-Cat**, **OF-Py-Car** and **OF-Py-Cat**, this work develops a new concept of the molecular design and synthesis of a type-I/type-II hybrid dye sensitizer capable of efficiently injecting electrons and with good light-harvesting features, although the photovoltaic performance of the DSSC based on the type-I/type-II hybrid dye sensitizer with a pyridyl group and a catechol unit that has been developed in



this current stage is much lower than in DSSCs based on type-I dye sensitizers with two carboxyl groups, two pyridyl groups, or both a pyridyl group and a carboxyl group.

Conclusions

Working toward the creation of a type-I/type-II hybrid dye sensitizer for dye-sensitized solar cells (DSSCs) that makes it possible to inject electron into the conduction band (CB) of the TiO₂ electrode by intramolecular charge-transfer (ICT) excitation (type-I pathway) and by the photoexcitation of the dye-to-TiO₂ charge transfer (DTCT) band (type-II pathway), we have designed and developed the dye sensitizer **OF-Py-Cat** with a pyridyl group and a catechol unit as the anchoring group possessing the ability to adsorb on the TiO₂ electrode through both the coordination bond at the Lewis acid sites and the bidentate binuclear bridging linkage at the Brønsted acid sites on the TiO₂ surface. Furthermore, not only to gain insight into the influence of the molecular structure of the type-I/type-II hybrid dye sensitizer on the appearance of the ICT and DTCT bands and the electron-injection mechanism, but also to investigate the impacts of the ICT and DTCT characteristics of type-I/type-II hybrid dye sensitizer on the photovoltaic performances of DSSCs, type-I dye sensitizer **OF-Car-Car** with two carboxyl groups, type-I dye sensitizer **OF-Py-Py** with two pyridyl groups, type-I dye sensitizer **OF-Py-Car** with a pyridyl group and a carboxyl group, and type-II dye sensitizer **OF-Cat-Cat** with two catechol units have been also synthesized. It was found that the dye **OF-Py-Cat** adsorbed on the TiO₂ film exhibits a broad photoabsorption band originating from the ICT and DTCT characteristics, that is, the dye **OF-Py-Cat** can act as a type-I/type-II hybrid dye sensitizer. However, the photovoltaic performance of the type-I/type-II DSSC based on **OF-Py-Cat** is lower than in the type-I DSSC based on the dye sensitizers (**OF-Car-Car**, **OF-Py-Py**, and **OF-Py-Car**) with a carboxyl group or a pyridyl group, but is equivalent to that of type-II DSSC based on the dye sensitizer (**OF-Cat-Cat**) with a catechol unit. On the basis of the photophysical and electrochemical properties, the dye-adsorption modes onto TiO₂ film and the photovoltaic performances for **OF-Car-Car**, **OF-Py-Py**, **OF-Cat-Cat**, **OF-Py-Car** and **OF-Py-Cat**, we revealed that the low photovoltaic performance of the DSSC based on **OF-Py-Cat** may be due to significantly faster back-electron transfer from the electrons injected into TiO₂ electrode to the oxidized dye, in spite of the efficient electron injection from the pyridyl group to the CB of TiO₂ electrode, which is the most important finding to be emphasized in this work. Consequently, this work develops a new concept of the molecular design and synthesis of a type-I/type-II hybrid dye sensitizer capable of efficiently injecting electrons and good light-harvesting feature, and thus in order to provide a direction in the molecular design of the high performance type-I/type-II hybrid dye sensitizer, further study to gain greater insight into the effects of the molecular structure of type-I/type-II hybrid dye sensitizer on the appearance of the ICT and DTCT characteristics, and the photovoltaic performances is now in

progress by developing dye sensitizers bearing the substituted-pyridyl group and -catechol unit.

Experimental

General

Melting points were measured with a Yanaco micro melting point apparatus MP model. IR spectra were recorded on a PerkinElmer Spectrum One FT-IR spectrometer by ATR method. High-resolution mass spectral data were acquired on a Thermo Fisher Scientific LTQ Orbitrap XL. High-resolution gas chromatography-time of flight mass spectral data (GC-TOFMS) were acquired on a JMS-T100 GCV 4G (JEOL). ¹H NMR and ¹³C NMR spectra were recorded on a Varian-400 (400 MHz) or a Varian-500 (500 MHz) FT NMR spectrometer. Absorption spectra were observed with a Shimadzu UV-3150 spectrophotometer and fluorescence spectra were measured with a HORIBA FluoroMax-4 spectrofluorometer. The fluorescence quantum yields in solution were determined by a HORIBA FluoroMax-4 spectrofluorometer by using a calibrated integrating sphere system (λ_{ex} = 383, 378, 370, 378 and 373 nm for **OF-Car-Car**, **OF-Py-Py**, **OF-Cat-Cat**, **OF-Py-Car**, and **OF-Py-Cat**, respectively). Cyclic voltammetry (CV) curves were recorded in THF/Bu₄NClO₄ (0.1 M) solution for **OF-Py-Py** and DMF/Bu₄NClO₄ (0.1 M) solution for **OF-Car-Car** and **OF-Cat-Cat**, **OF-Py-Car**, and **OF-Py-Cat** with a three-electrode system consisting of Ag/Ag⁺ as the reference electrode, a Pt plate as the working electrode, and Pt wire as the counter electrode by using an Electrochemical Measurement System HZ-7000 (HOKUTO DENKO). The HOMO and LUMO energy levels of **OF-Car-Car** and **OF-Py-Py**, **OF-Cat-Cat**, **OF-Py-Car**, and **OF-Py-Cat** were evaluated from the spectral analyses and the CV data (the HOMO energy level was evaluated from the $E_{\text{onset}}^{\text{ox}}$ of the oxidation potential or the $E_{1/2}^{\text{ox}}$). The LUMO energy level was estimated from the $E_{\text{onset}}^{\text{ox}}$ or the $E_{1/2}^{\text{ox}}$ and an intersection of absorption and fluorescence spectra (428 nm; 2.90 eV for **OF-Car-Car**, 417 nm; 2.97 eV for **OF-Py-Py**, 398 nm; 3.12 eV for **OF-Cat-Cat**, 424 nm; 2.92 eV for **OF-Py-Car** and 418 nm; 2.97 eV for **OF-Py-Cat**) in THF, that is, the LUMO energy level was obtained through the equation $-[E_{0-0} - \text{HOMO}]$, where the E_{0-0} transition energy is an intersection of the absorption and fluorescence spectra corresponding to the energy gap between the HOMO and the LUMO. Electrochemical impedance spectroscopy (EIS) for DSSCs in the dark under a forward bias of -0.35 or -0.55 V with a frequency range of 1 Hz to 100 kHz was measured with an AMETEK Versa STAT 3.

Synthesis

9-Butyl-3-(5-(pyridin-4-yl)thiophen-2-yl)-9H-carbazole (1). To a DMF solution (30 mL) of 5-(pyridin-4-yl)-2-thienylboronic acid pinacol ester (1.0 g, 3.5 mmol) and 3-bromo-9-butyl-9H-carbazole³¹ (1.0 g, 3.3 mmol) under an argon atmosphere was added aqueous Na₂CO₃ (0.74 g, 6.6 mmol) and Pd(PPh₃)₄ (0.09 g, 0.1 mmol), which was then stirred for 6 h at 90 °C. After concentrating under reduced pressure, the resulting residue was dissolved in chloroform and washed with water. The chloroform extract was evaporated under reduced pressure. The residue was chromatographed on



silica gel (chloroform as eluent) to give **1** (1.0 g, yield 78%) as a yellow solid; m.p. 150–152 °C; IR (ATR): $\tilde{\nu}$ = 1590, 1445, 1206 cm⁻¹; ¹H NMR (400 MHz, acetone-d₆) δ = 0.94 (t, *J* = 7.6 Hz, 3H), 1.40–1.43 (m, 2H), 1.86–1.90 (m, 2H), 4.47 (t, *J* = 6.8 Hz, 2H), 7.22–7.27 (m, 1H), 7.47–7.51 (m, 1H), 7.56 (d, *J* = 4.0 Hz, 1H), 7.60 (d, *J* = 8.2 Hz, 1H), 7.64–7.67 (m, 3H), 7.76 (d, *J* = 3.8 Hz, 1H), 7.85 (dd, *J* = 1.9 and 8.6 Hz, 1H), 8.25 (d, *J* = 7.4 Hz, 1H), 8.54 (d, *J* = 1.3 Hz, 1H), 8.58 (dd, *J* = 1.7 and 4.4 Hz, 2H) ppm; HRMS (ESI): *m/z* (%): [M + H⁺] calcd for C₂₅H₂₃N₂S, 383.15765; found 383.15845.

3-Bromo-9-butyl-6-(5-(pyridin-4-yl)thiophen-2-yl)-9H-carbazole (2). A solution of **1** (0.06 g, 0.16 mmol) and *N*-bromosuccinimide (0.034 g, 0.19 mmol) in dichloromethane (5 ml) was stirred at 0 °C in dark conditions. After 12 h, Na₂S₂O₃ aq. was added to the reaction mixture. The organic extract was concentrated under reduced pressure. The residue was dissolved in toluene, and recycling gel permeation chromatography (GPC) was performed to give **2** (0.024 g, yield 33%) as a yellow solid; m.p. 171–172 °C; IR (ATR): $\tilde{\nu}$ = 1590, 1453, 1214 cm⁻¹; ¹H NMR (400 MHz, acetone-d₆) δ = 0.93 (t, *J* = 7.6 Hz, 3H), 1.34–1.40 (m, 2H), 1.79–1.85 (m, 2H), 4.27 (t, *J* = 7.2 Hz, 2H), 7.27 (d, *J* = 8.6 Hz, 1H), 7.33 (d, *J* = 3.8 Hz, 1H), 7.39 (d, *J* = 8.6 Hz, 1H), 7.47–7.51 (m, 3H), 7.54 (dd, *J* = 1.9 and 8.6 Hz, 1H), 7.74 (dd, *J* = 1.9 and 8.6 Hz, 1H), 8.23 (d, *J* = 1.7 Hz, 1H), 8.26 (d, *J* = 1.4 Hz, 1H), 8.58 (d, *J* = 4.4 Hz, 2H) ppm; HRMS (ESI): *m/z* (%): [M + H⁺] calcd for C₂₅H₃₂N₂BrS, 461.06816; found 461.06839.

3-(5-(3,4-Bis(methoxymethoxy)phenyl)thiophen-2-yl)-9-butyl-6-(5-(pyridin-4-yl)thiophen-2-yl)-9H-carbazole (3). To a DMF solution (10 mL) of 5-(3,4-bis(methoxymethoxy)phenyl)thiophen-2-yl)boronic acid pinacol ester (0.45 g, 1.3 mmol) and **2** (0.55 g, 1.2 mmol) under an argon atmosphere was added aqueous Na₂CO₃ (0.26 g, 2.4 mmol) and Pd(PPh₃)₄ (0.28 g, 0.24 mmol), followed by stirring for 6 h at 90 °C. After concentrating under reduced pressure, the resulting residue was dissolved in chloroform and washed with water. The chloroform extract was evaporated under reduced pressure. The residue was chromatographed on silica gel (chloroform as eluent) and then recycling GPC (toluene as eluent) was performed to give **3** (0.25 g, yield 32%) as a reddish brown viscous solid; m.p. 66–67 °C; IR (ATR): $\tilde{\nu}$ = 1590, 1483, 1241, 1071 cm⁻¹; ¹H NMR (400 MHz, acetone-d₆) δ = 0.95 (t, *J* = 7.6 Hz, 3H), 1.39–1.48 (m, 2H), 1.87–1.95 (m, 2H), 3.49 (s, 3H), 3.53 (s, 3H), 4.50 (t, *J* = 7.5 Hz, 2H), 5.24 (s, 2H), 5.30 (s, 2H), 7.19 (d, *J* = 8.4 Hz, 1H), 7.33 (dd, *J* = 2.2 and 8.4 Hz, 1H), 7.38 (d, *J* = 4.0 Hz, 1H), 7.47–7.49 (m, 2H), 7.60 (d, *J* = 3.8 Hz, 1H), 7.65–7.69 (m, 4H), 7.78 (d, *J* = 3.8 Hz, 1H), 7.84–7.89 (m, 2H), 8.58 (dd, *J* = 1.6 and 4.4 Hz, 2H), 8.61 (d, *J* = 1.2 Hz, 1H), 8.68 (d, *J* = 1.4 Hz, 1H) ppm; HRMS (ESI): *m/z* (%): [M + H⁺] calcd for C₃₉H₃₇O₄N₂S₂, 661.21893; found 661.21930.

4-(5-(9-Butyl-6-(5-(pyridin-4-yl)thiophen-2-yl)-9H-carbazol-3-yl)thiophen-2-yl)benzene-1,2-diol (OF-Py-Cat). To compound **3** (0.10 g, 0.15 mmol) was added THF (3 ml) and 1 M HCl (2 ml), followed by stirring at 80 °C. After 3 h, aqueous Na₂CO₃ was added to the reaction mixture to neutralize the solution. The resulting solid was filtered to collect **OF-Py-Cat** (0.025 g, yield 29%) as a red solid; m.p. 183–185 °C; IR (ATR): $\tilde{\nu}$ = 1629, 1600, 1427 cm⁻¹; ¹H NMR (400 MHz, acetone-d₆) δ = 0.95 (t, *J* = 7.4 Hz, 3H), 1.39–1.48 (m, 2H), 1.87–1.89 (m, 2H, overlapping peak of residual proton in acetone-d₆), 4.50 (t, *J* = 7.8 Hz, 2H),

6.89 (d, *J* = 8.2 Hz, 1H), 7.07 (dd, *J* = 2.2 and 8.2 Hz, 1H), 7.19 (d, *J* = 2.2 Hz, 1H), 7.26 (d, *J* = 3.8 Hz, 1H), 7.43 (d, *J* = 3.8 Hz, 1H), 7.60 (d, *J* = 3.8 Hz, 1H), 7.64–7.69 (m, 4H), 7.78 (d, *J* = 3.8 Hz, 1H), 7.83 (dd, *J* = 1.9 and 8.6 Hz, 1H), 7.87 (dd, *J* = 1.9 and 8.6 Hz, 1H), 8.58–8.59 (m, 3H), 8.67 (d, *J* = 1.6 Hz, 1H) ppm; HRMS (ESI): *m/z* (%): [M + H⁺] calcd for C₃₅H₂₉O₂N₂S₂, 573.16650; found 573.16656.

Ethyl 4-(5-(9-butyl-9H-carbazol-3-yl)thiophen-2-yl)benzoate (4). To a DMF solution (20 mL) of ethyl 4-(5-bromothiophen-2-yl)benzoate³⁴ (0.85 g, 2.7 mmol) and (9-butyl-9H-carbazol-3-yl)boronic acid pinacol ester³² (0.95 g, 2.7 mmol) under an argon atmosphere was added aqueous Na₂CO₃ (0.58 g, 5.4 mmol) and Pd(PPh₃)₄ (0.09 g, 0.1 mmol), followed by stirring for 12 h at 90 °C. After concentrating under reduced pressure, the resulting residue was dissolved in chloroform and then washed with water. The chloroform extract was evaporated under reduced pressure. The residue was chromatographed on silica gel (dichloromethane as eluent) to give **4** (1.07 g, yield 87%) as a yellow solid; m.p. 116–118 °C; IR (ATR): $\tilde{\nu}$ = 1710, 1595, 1443, 1275 cm⁻¹; ¹H NMR (400 MHz, acetone-d₆) δ = 0.94 (t, *J* = 7.2 Hz, 3H), 1.36–1.45 (m, 5H), 1.86–1.90 (m, 2H), 4.36 (q, *J* = 7.2 Hz, 2H), 4.47 (t, *J* = 7.2 Hz, 2H), 7.22–7.26 (m, 1H), 7.47–7.51 (m, 1H), 7.54 (d, *J* = 3.8 Hz, 1H), 7.61 (d, *J* = 8.3 Hz, 1H), 7.64–7.66 (m, 2H), 7.83–7.86 (m, 3H), 8.05 (dd, *J* = 2.0 and 6.7 Hz, 2H), 8.25 (d, *J* = 7.5 Hz, 1H), 8.53 (d, *J* = 1.3 Hz, 1H) ppm; HRMS (ESI): *m/z* (%): [M + H⁺] calcd for C₂₉H₂₈O₂NS, 454.18353; found 454.18335.

Ethyl 4-(5-(6-bromo-9-butyl-9H-carbazol-3-yl)thiophen-2-yl)benzoate (5). A solution of **4** (0.5 g, 1.1 mmol) and *N*-bromosuccinimide (0.195 g, 1.1 mmol) in dichloromethane (10 ml) was stirred at 0 °C in dark conditions. After 12 h, Na₂S₂O₃ aq. was added to the reaction mixture. The organic extract was concentrated under reduced pressure. The residue was dissolved in toluene, and recycling GPC was performed to give **5** (0.12 g, yield 20%) as a yellow solid; m.p. 158–159 °C; IR (ATR): $\tilde{\nu}$ = 1701, 1604, 1454, 1274 cm⁻¹; ¹H NMR (400 MHz, dichloromethane-d₂) δ = 0.95 (t, *J* = 7.2 Hz, 3H), 1.35–1.42 (m, 5H), 1.81–1.88 (m, 2H), 4.29 (t, *J* = 7.2 Hz, 2H), 4.37 (q, *J* = 7.2 Hz, 2H), 7.33 (d, *J* = 8.6 Hz, 1H), 7.37 (d, *J* = 3.8 Hz, 1H), 7.44–7.47 (m, 2H), 7.57 (dd, *J* = 2.0 and 8.6 Hz, 1H), 7.71–7.74 (m, 2H), 7.79 (dd, *J* = 1.9 and 8.4 Hz, 1H), 8.03–8.06 (m, 2H), 8.25 (d, *J* = 2.4 Hz, 1H), 8.31 (d, *J* = 1.8 Hz, 1H) ppm; HRMS (APCI): *m/z* (%): [M⁺] calcd for C₂₉H₂₆O₂NBrS, 531.08621; found 531.08643.

Ethyl 4-(5-(9-butyl-6-(5-(pyridin-4-yl)thiophen-2-yl)-9H-carbazol-3-yl)thiophen-2-yl)benzoate (6). To a DMF solution (3 mL) of 5-(pyridin-4-yl)-2-thienylboronic acid pinacol ester (0.065 g, 0.23 mmol) and **5** (0.11 g, 0.21 mmol) under an argon atmosphere was added aqueous Na₂CO₃ (0.044 g, 0.41 mmol) and Pd(PPh₃)₄ (0.007 g, 0.006 mmol), followed by stirring for 6 h at 90 °C. After concentrating under reduced pressure, the resulting residue was dissolved in chloroform and washed with water. The chloroform extract was evaporated under reduced pressure. The residue was chromatographed on silica gel (chloroform as eluent) to give **6** (0.1 g, yield 82%) as an orange yellow solid; m.p. 207–208 °C; IR (ATR): $\tilde{\nu}$ = 1710, 1590, 1440, 1271 cm⁻¹; ¹H NMR (400 MHz, dichloromethane-d₂) δ = 0.95 (t, *J* = 7.3 Hz, 3H),



1.38–1.45 (m, 5H), 1.86–1.93 (m, 2H), 4.33–4.40 (m, 4H), 7.42 (d, $J = 3.7$ Hz, 1H), 7.43 (d, $J = 3.8$ Hz, 1H), 7.47–7.50 (m, 2H), 7.53–7.54 (m, 2H), 7.58 (d, $J = 3.8$ Hz, 1H), 7.71–7.76 (m, 3H), 7.80–7.83 (m, 2H), 8.05–8.07 (m, 2H), 8.43–8.44 (m, 2H), 8.58 (d, $J = 4.6$ Hz, 2H) ppm; HRMS (ESI): m/z (%): $[M + H^+]$ calcd for $C_{38}H_{33}O_2N_2S_2$, 613.19780; found 613.19824.

4-(5-(9-Butyl-6-(5-(pyridin-4-yl)thiophen-2-yl)-9H-carbazol-3-yl)thiophen-2-yl)benzoic acid (OF-Py-Car). To a solution of **6** (0.035 g, 0.06 mmol) in a mixed solvent of ethanol (14 ml) and dichloromethane (5 ml) was added dropwise aqueous NaOH (5 ml, 0.05 M) with stirring at 75 °C. After further stirring for 12 h, the solution was concentrated under reduced pressure. The residue was dissolved in dichloromethane, and washed with water, in which the aqueous layer was acidified to pH 2 with 2 M HCl. The resulting solid was filtered to collect **OF-Py-Car** (0.027 g, yield 81%) as an orange solid; m.p. over 300 °C; IR (ATR): $\tilde{\nu} = 1695, 1601, 1448, 1288$ cm^{-1} ; 1H NMR (500 MHz, tetrahydrofuran- d_8) $\delta = 0.96$ (t, $J = 7.0$ Hz, 3H), 1.39–1.44 (m, 2H), 1.55–1.88 (m, 2H, overlapping peak of residual proton in tetrahydrofuran- d_8), 4.44 (t, $J = 7.0$ Hz, 2H), 7.48 (d, $J = 3.9$ Hz, 1H), 7.51 (d, $J = 3.9$ Hz, 1H), 7.56–7.58 (m, 4H), 7.68 (d, $J = 4.0$ Hz, 1H), 7.74–7.79 (m, 3H), 7.81–7.84 (m, 2H), 8.03–8.05 (m, 2H), 8.54–8.56 (m, 4H) ppm; HRMS (ESI): m/z (%): $[M + H^+]$ calcd for $C_{36}H_{29}O_2N_2S_2$, 585.16650; found 585.16685.

9-Butyl-3,6-di(thiophen-2-yl)-9H-carbazole (7)³². A solution of 3,6-dibromo-9-butyl-9H-carbazole (3.8 g, 10.0 mmol), 2-thienylboronic acid (2.5 g, 25.0 mmol), aqueous Na_2CO_3 (2.5 g, 20.0 mmol) and $Pd(PPh_3)_4$ (0.61 g, 0.5 mmol) in DMF (40 ml) was stirred for 7 h at 80 °C. After concentrating under reduced pressure, the resulting residue was dissolved in dichloromethane and washed with water. The dichloromethane extract was evaporated under reduced pressure. The residue was chromatographed on silica gel (dichloromethane:hexane 1:1 as the eluent) to give **7** (1.6 g, yield 40%) as a white solid; 1H NMR (400 MHz, acetone- d_6) $\delta = 0.94$ (t, $J = 7.2$ Hz, 3H), 1.39–1.45 (m, 2H), 1.85–1.93 (m, 2H), 4.47 (t, $J = 7.2$ Hz, 2H), 7.13 (dd, $J = 3.5$ and 5.1 Hz, 2H), 7.39 (dd, $J = 1.2$ and 5.1 Hz, 2H), 7.48 (d, $J = 1.2$ and 3.5 Hz, 2H), 7.62 (d, $J = 8.6$ Hz, 2H), 7.79 (dd, $J = 1.9$ and 8.6 Hz, 2H), 8.56 (d, $J = 1.4$ Hz, 2H) ppm.

3,6-Bis(5-bromothiophen-2-yl)-9-butyl-9H-carbazole (8). A solution of **7** (3.8 g, 10.0 mmol) and *N*-bromosuccinimide (3.7 g, 20.0 mmol) in chloroform (200 ml) was stirred at 0 °C in dark conditions. After 2 h, $Na_2S_2O_3$ aq. was added to the reaction mixture. The organic extract was concentrated under reduced pressure. The residue was compound **8** (5.3 g, yield 95%) as a yellow solid; m.p. 150–152 °C; IR (ATR): $\tilde{\nu} = 1601, 1482, 1430$ cm^{-1} ; 1H NMR (400 MHz, acetone- d_6) $\delta = 0.93$ (t, $J = 7.3$ Hz, 3H), 1.37–1.44 (m, 2H), 1.84–1.90 (m, 2H), 4.47 (t, $J = 7.2$ Hz, 2H), 7.18 (d, $J = 3.9$ Hz, 2H), 7.30 (d, $J = 3.9$ Hz, 2H), 7.64 (d, $J = 8.4$ Hz, 2H), 7.74 (dd, $J = 1.8$ and 8.4 Hz, 2H), 8.52 (d, $J = 1.3$ Hz, 2H) ppm; HRMS (APCI): m/z (%): $[M + H^+]$ calcd for $C_{24}H_{20}NBr_2S_2$, 545.93780; found 545.93811.

Diethyl 4,4'-((9-butyl-9H-carbazole-3,6-diyl)bis(thiophene-5,2-diyl))dibenzoate (9). A solution of **8** (1.0 g, 1.8 mmol), 4-ethoxy-carbonylphenylboronic acid (0.9 g, 4.5 mmol), aqueous Na_2CO_3 (0.4 g, 3.6 mmol) and $Pd(PPh_3)_4$ (0.11 g, 0.09 mmol) in DMF

(10 ml) was stirred for 4 h at 90 °C. After concentrating under reduced pressure, the resulting residue was dissolved in dichloromethane and washed with water. The dichloromethane extract was evaporated under reduced pressure. The residue was chromatographed on silica gel (dichloromethane:hexane 3:1 as eluent) to give **9** (0.65 g, yield 50%) as a yellow solid; m.p. 172–174 °C; IR (ATR): $\tilde{\nu} = 1705, 1601, 1444, 1270$ cm^{-1} ; 1H NMR (400 MHz, acetone- d_6) $\delta = 0.96$ (t, $J = 7.6$ Hz, 3H), 1.36–1.47 (m, 8H), 1.88–1.93 (m, 2H), 4.36 (q, $J = 7.2$ Hz, 4H), 4.51 (t, $J = 7.2$ Hz, 2H), 7.58 (d, $J = 3.8$ Hz, 2H), 7.67–7.70 (m, 4H), 7.85–7.90 (m, 6H), 8.05–8.08 (m, 4H), 8.67 (d, $J = 1.3$ Hz, 2H) ppm; HRMS (APCI): m/z (%): $[M + H^+]$ calcd for $C_{42}H_{38}O_4NS_2$, 684.22368; found 684.22290.

4,4'-((9-Butyl-9H-carbazole-3,6-diyl)bis(thiophene-5,2-diyl))-dibenzoic acid (OF-Car-Car). To a solution of **9** (0.1 g, 0.15 mmol) in a mixed solvent of ethanol (100 ml) and dichloromethane (30 ml) was added dropwise aqueous NaOH (32 ml, 0.05 M) with stirring at 75 °C. After further stirring for 12 h, the solution was concentrated under reduced pressure. The residue was dissolved in dichloromethane and washed with water, in which the aqueous layer was acidified to pH 1 with 2 M HCl. The resulting solid was filtered to collect **OF-Car-Car** (0.08 g, yield 89%) as a yellow solid; m.p. over 300 °C; IR (ATR): $\tilde{\nu} = 1678, 1601, 1425, 1292$ cm^{-1} ; 1H NMR (400 MHz, tetrahydrofuran- d_8) $\delta = 0.96$ (t, $J = 7.5$ Hz, 3H), 1.36–1.48 (m, 2H), 1.56–1.89 (m, 2H, overlapping peak of residual proton in tetrahydrofuran- d_8), 4.43 (t, $J = 7.0$ Hz, 2H), 7.48 (d, $J = 3.5$ Hz, 2H), 7.56–7.58 (m, 4H), 7.78–7.83 (m, 6H), 8.04 (d, $J = 8.5$ Hz, 4H), 8.55 (d, $J = 1.5$ Hz, 2H) ppm; ^{13}C NMR (125 MHz, tetrahydrofuran- d_8) $\delta = 14.03, 21.09, 31.96, 43.01, 110.29, 118.26, 123.89, 124.13, 124.76, 125.29, 126.42, 127.65, 130.07, 130.90, 131.07, 139.24, 141.59, 147.19, 167.01$ ppm; TOFMS (FD): m/z (%): $[M^+]$ calcd for $C_{38}H_{29}O_4NS_2$, 627.15380; found 627.15430.

3,6-Bis(5-(3,4-bis(methoxymethoxy)phenyl)thiophen-2-yl)-9-butyl-9H-carbazole (10). A solution of **8** (0.2 g, 0.37 mmol), (3,4-bis(methoxymethoxy)phenyl)boronic acid pinacol ester (0.3 g, 0.9 mmol), aqueous Na_2CO_3 (0.08 g, 0.7 mmol) and $Pd(PPh_3)_4$ (0.02 g, 0.02 mmol) in DMF (3 ml) was stirred for 4 h at 90 °C. After concentrating under reduced pressure, the resulting residue was dissolved in dichloromethane and washed with water. The dichloromethane extract was evaporated under reduced pressure. The residue was chromatographed on silica gel (ethyl acetate:hexane 1:1 as eluent) and then recycling GPC (toluene as eluent) was performed to give **10** (0.15 g, yield 52%) as a yellow solid; m.p. 41–43 °C; IR (ATR): $\tilde{\nu} = 1600, 1484, 1241, 1151, 1071$ cm^{-1} ; 1H NMR (400 MHz, acetone- d_6) $\delta = 0.82$ (t, $J = 7.2$ Hz, 3H), 1.23–1.29 (m, 2H), 1.70–1.75 (m, 2H), 3.48 (s, 6H), 3.51 (s, 6H), 4.23 (t, $J = 7.8$ Hz, 2H), 5.21 (s, 4H), 5.28 (s, 4H), 7.14 (d, $J = 8.4$ Hz, 2H), 7.25–7.29 (m, 4H), 7.37 (d, $J = 3.7$ Hz, 2H), 7.41 (d, $J = 8.6$ Hz, 2H), 7.48 (d, $J = 2.2$ Hz, 2H), 7.70 (dd, $J = 1.8$ and 8.6 Hz, 2H), 8.53 (d, $J = 1.8$ Hz, 2H) ppm; HRMS (ESI): m/z (%): $[M + Na^+]$ calcd for $C_{44}H_{45}O_8NS_2Na$, 802.24788; found 802.24664.

4,4'-((9-Butyl-9H-carbazole-3,6-diyl)bis(thiophene-5,2-diyl))-bis(benzene-1,2-diol) (OF-Cat-Cat). To compound **10** (0.15 g, 0.19 mmol) was added THF (3 ml) and 1 M HCl (2 ml), followed



by stirring at 80 °C for 12 h. After concentrating under reduced pressure, the resulting residue was subjected to reprecipitation to afford **OF-Cat-Cat** (0.1 g, yield 91%) as a dark green solid; m.p. over 300 °C; IR (ATR): $\tilde{\nu}$ = 1651, 1599, 1486 cm⁻¹; ¹H NMR (500 MHz, acetonitrile-d₃) δ = 0.94 (t, *J* = 7.5 Hz, 3H), 1.36–1.41 (m, 2H), 1.79–2.00 (m, 2H, overlapping peak of residual proton in acetonitrile-d₃), 4.39 (t, *J* = 7.0 Hz, 2H), 6.87 (d, *J* = 8.0 Hz, 2H), 7.09 (dd, *J* = 2.5 and 8.0 Hz, 2H), 7.17 (d, *J* = 2.5 Hz, 2H), 7.26 (d, *J* = 3.5 Hz, 2H), 7.39 (d, *J* = 3.5 Hz, 2H), 7.57 (d, *J* = 8.5 Hz, 2H), 7.80 (dd, *J* = 1.5 and 8.5 Hz, 2H), 8.49 (d, *J* = 2.0 Hz, 2H) ppm; TOFMS (FD): *m/z* (%): [*M*⁺] calcd for C₃₆H₂₉O₄NS₂, 603.15380; found 603.15521.

Preparation of DSSCs

The TiO₂ paste (JGC Catalysts and Chemicals Ltd, PST-18NR) was deposited on a fluorine-doped-tin-oxide (FTO) substrate by doctor-blading, and sintered for 50 min at 450 °C. The 9 μm thick TiO₂ electrode was immersed into 0.1 mM dye solution in THF for 15 hours, which was enough to adsorb the dye sensitizers. The DSSCs were fabricated by using the TiO₂ electrode (0.5 × 0.5 cm² in the photoactive area) thus prepared, Pt-coated glass as the counter electrode, and a solution of 0.05 M iodine, 0.1 M lithium iodide, and 0.6 M 1,2-dimethyl-3-propylimidazolium iodide in acetonitrile as the electrolyte. The photocurrent–voltage characteristics were measured using a potentiostat under simulated solar light (AM 1.5, 100 mW cm⁻²). IPCE spectra were measured under monochromatic irradiation with a tungsten-halogen lamp and a monochromator. The amount of dye molecules adsorbed on TiO₂ film was determined from the calibration curve by absorption spectral measurement of the concentration change of the dye solution before and after adsorption. The quantification of dye was made based on the molar extinction coefficient for $\lambda_{\text{max}}^{\text{abs}}$ of dye in the above solution. Absorption spectra of the dyes adsorbed on TiO₂ film were recorded on the dyes-adsorbed TiO₂ film in the transmission mode with a calibrated integrating sphere system.

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