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# $^{125}\text{Te}$ NMR provides evidence of autoassociation of organo-ditellurides in solution†

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The frequency of the resonance of  $^{125}\text{Te}$  of two organo-ditellurides, R–Te–Te–R (R = 4- $\text{CH}_3\text{C}_6\text{H}_4$  and 2-( $\text{CH}_3$ ) $_2\text{NCH}_2\text{C}_6\text{H}_4$ ), in solution undergoes a low-field shift as the concentration of the sample increases. In sharp contrast, the resonance of a sterically hindered ditelluride (R = ( $\text{C}_6\text{H}_5(\text{CH}_3)_2\text{Si}$ ) $_3\text{C}$ ) and telluric acid display the opposite effect. While the negative concentration coefficients can be explained by the change in magnetic susceptibility, the positive coefficients are consistent with autoassociation of the molecules through tellurium-centred supramolecular interactions. Although the corresponding equilibrium constants are small, the process is shown to be exothermic. However, the influence of autoassociation is much smaller than the effects of solvent polarity and the conformation of the ditelluride bond.

## Introduction

Supramolecular association is a recurrent feature in the structural chemistry of organic derivatives of heavy p-block elements. The success of halogen bonding<sup>1–4</sup> in areas as diverse as crystal engineering,<sup>5–16</sup> photonic materials<sup>17,18</sup> organocatalysis<sup>19</sup> and biomimetic chemistry<sup>20</sup> – to name a few – has stimulated the resurgence of interest in the study of analogous interactions centred on other heavy main-group elements, including the chalcogens (S, Se, Te).<sup>21–23</sup>

While tellurium-centred supramolecular interactions (alternatively termed secondary bonding or chalcogen bonding interactions) are primarily characterized in the solid state through X-ray diffraction structural determination, non-crystalline systems must be studied by means of spectroscopic methods.

UV photoelectron spectroscopy (UVPES) has been unable to provide evidence of aggregation through these interactions in the gas phase. Even in the case of benzo-2,1,3-telluradiazole, which is strongly associated in the solid state, the UVPES spectrum corresponds to individual molecules only.<sup>24</sup> Electro-spray mass spectrometry of 1,2,5-telluradiazoles has detected protonated dimers<sup>25</sup> and the adduct of a neutral molecule with  $\text{Cl}^-$ .<sup>26</sup> However, the interpretation of such experiments is complicated by the charge of the ions and side reactions within the instrument. As organo-tellurium compounds are often coloured, UV-vis absorption spectroscopy is an appealing option in solution; this technique has been used to measure

association constants for the adducts of benzotelluradiazoles with Lewis bases<sup>26</sup> and monitor the association of anions with tellurophenes.<sup>27</sup> However, photolysis can be a concern with many systems.

Arguably the most practical spectroscopic method for the characterization of tellurium-centred supramolecular interactions in solution is NMR. Indeed  $^1\text{H}$  and  $^{19}\text{F}$  resonances have been used to monitor the association of 1,2,5-telluradiazoles with Lewis bases<sup>26</sup> and, very recently,  $^1\text{H}$  NMR was used to demonstrate the persistence in solution of the macrocyclic aggregates of isotellurazole *N*-oxides.<sup>28</sup> However, such measurements rely on effects secondary to the interactions of the chalcogen atom. Fortunately, tellurium can be directly observed because it has two NMR-active isotopes with spin 1/2 and no quadrupole moment:  $^{123}\text{Te}$  and  $^{125}\text{Te}$ . Observation of the latter is most convenient as it has relatively high receptivity ( $D^C = 13.1$ ) in spite of its low natural abundance (7.12%).<sup>29</sup> Its chemical shift range is wider than 4000 ppm:<sup>30</sup> from –1214 ppm for ( $\text{Me}_3\text{Sn}$ ) $_2\text{Te}$ <sup>31</sup> to 3102 ppm for the  $\text{Se}_2\text{Te}_2^{2+}$  dication in 30% oleum.<sup>32</sup> The frequency of resonance of  $^{125}\text{Te}$  is therefore very sensitive, so much that the effect of supramolecular association could possibly be masked by the influence of small structural changes or the dielectric and diamagnetic properties of the medium.

Based on these considerations, we have conducted a detailed  $^{125}\text{Te}$  NMR investigation of a selected set of tellurium compounds intended to probe the presence of tellurium-centred supramolecular interactions in solution, quantify their impact on the measured chemical shifts and compare this effect with the influence of other environmental factors. The molecules selected for this study are simple compounds that differ from each other in their relative ability to undergo autoassociation in solution though tellurium-centred supramolecular interactions.

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# Results and discussion

## System of choice

The compounds used for this study (Chart 1) are organo-ditellurides. The Te–Te bond in these heavy congeners of peroxides is analogous to the I–I linkage in molecular iodine in that the bond between the heavy atoms generates a pair of distal  $\sigma$  holes and a low-lying  $\sigma^*_{\text{Te-Te}}$  antibonding orbital. Fig. 1 portrays those features on the molecule of diphenyl ditelluride (**1**); being the prototypical organo-ditelluride, this compound was selected as a model for supporting DFT calculations. The combined electrostatic and covalent effects enable attractive interactions of the chalcogen atoms with Lewis bases (Scheme 1);<sup>33–35</sup> however, Symmetry-Adapted Perturbational Theory calculations strongly indicate that dispersion is an important component of such interactions.<sup>36</sup> Only a handful of ditelluride structures deposited in the Cambridge structural database exhibit no short contacts with the tellurium atoms, such cases are the result of steric hindrance or peculiar packing effects.

Intramolecular interactions are common when substituents bearing atoms with lone pairs are present in the molecular structure. Intermolecular interactions in ditellurides are most commonly observed in the solid state with either Te atoms,<sup>37–40</sup> or aryl rings,<sup>41,42</sup> or through interactions with electron-donor functional groups on adjacent molecules.<sup>43</sup>

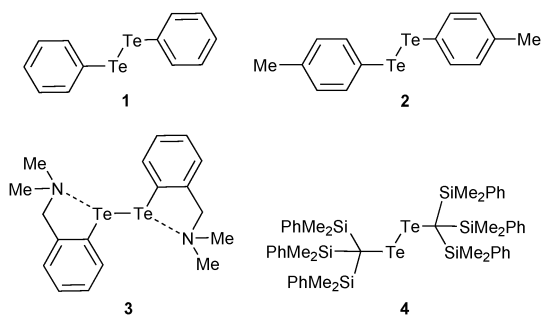


Chart 1

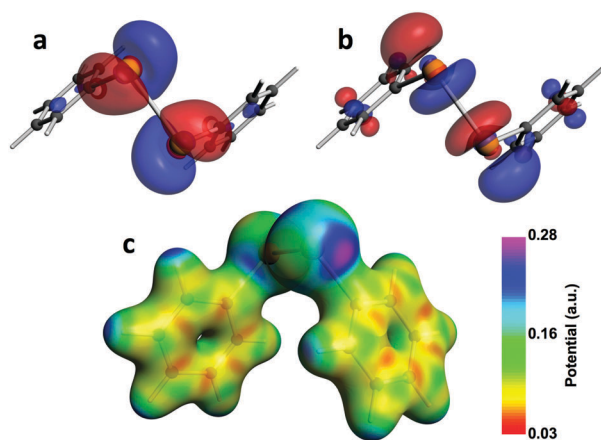
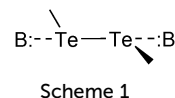


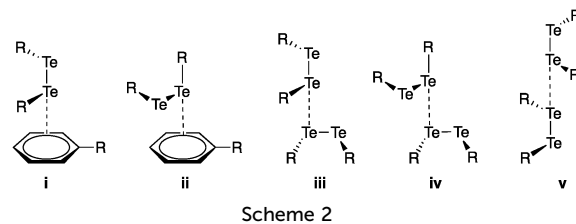
Fig. 1 Selected computational results of  $\text{Ph}_2\text{Te}_2$  (**1**). The (a) HOMO; (b) LUMO; and (c) electrostatic potential mapped onto the SCF electron density. All isosurfaces are plotted at 0.03 a.u.



Scheme 1

The crystal structures of organo-ditellurides which display short contacts with the chalcogen can be classified by the type of intermolecular interactions into two major groups: those displaying  $\text{Aryl} \cdots \text{Te}$  interactions and those with  $\text{Te} \cdots \text{Te}$  interactions. Within the first group, there are cases in which the aryl group is opposite to the Te–Te bond (Scheme 2 (i)) or a Te–C bond (Scheme 2 (ii)). Similarly, in the second group a Te atom is opposite to the Te–Te bond (Scheme 2 (iii)) or a Te–C bond (Scheme 2 (iv)), in a third category the lattice features linear  $(\text{Te}_2)_\infty$  chains (Scheme 2 (v)).

For our experimental investigations we compared the behavior of ditellurides **2–4**, in which the organic substituents were selected in order to influence the intermolecular interactions centred on tellurium. Table 1 presents a comparison of key properties of their molecules. Bis-4-methyl-phenyl ditelluride (**2**) displays a *gauche* conformation in the solid state.<sup>44</sup> Its tellurium atoms are readily accessible for interaction with other molecules. Because of the potential for photolysis and other degradation mechanisms, the *para*-methyl groups were used as a spectroscopic handle to readily monitor the integrity of the samples throughout all experiments; the effect of these substituents on autoassociation is assumed to be negligible. Bis(*ortho*-(dimethylamino-methyl)phenyl)-ditelluride (**3**) features two pendant dialkylamino groups that in the crystal engage in  $\text{Te} \cdots \text{N}$  intramolecular interactions with an average distance of 2.876(3) Å. Its Te–Te bond is somewhat longer than in **2**, the C–Te–Te–C dihedral angle<sup>45</sup> is wider and the compound is photosensitive. Should the  $\text{Te} \cdots \text{N}$  interactions in **3** persist in solution, they would compete with autoassociation and binding to solvent molecules. The bulky substituents of bis(tris-dimethylphenylsilyl)methyl ditelluride (**4**) prevent any intermolecular contacts with tellurium. The steric demands of the substituents force the C–Te–Te–C dihedral angle to 180° and elongate the Te–Te bond distance; the compound is stable in the solid state but decomposes slowly in solution, even in the dark.<sup>46</sup>



Scheme 2

Table 1 Selected properties of ditellurides **2–4**

Compound	$d_{\text{Te-Te}}$ (Å)	$\theta_{\text{C-Te-Te-C}}$ (°)	$\delta^{125}\text{Te}^a$ (ppm)	Ref.
<b>2</b>	2.6962(5)	85.7(4)	432.2	42 and 45
<b>3</b>	2.7480(7)	95.2(1)	355	43
<b>4</b>	2.7716(3)	180.0(1)	442	44

<sup>a</sup> Measurements reported for  $\text{CDCl}_3$  solution.



### Effect of concentration on the $^{125}\text{Te}$ chemical shift

If ditelluride **A** underwent autoassociation equilibrium in solution, dimerization in the simplest case (eqn (1)), the proportion of associated molecules ( $\alpha$ , eqn (2)) would increase with an increase of the total concentration:



$$\alpha = \frac{2C_{\text{A}_2}}{C} \quad (2)$$

Such equilibrium, with constant  $K$  (eqn (3)), would be fast enough for the NMR experiment to measure only an averaged resonance of  $^{125}\text{Te}$  nuclei, (eqn (4)):

$$K = \frac{\alpha/2}{(1-\alpha)^2 C} \quad (3)$$

$$\delta = \delta_{\text{A}} + (\delta_{\text{A}_2} - \delta_{\text{A}})\alpha = \delta_{\text{A}} + \Delta\delta\alpha \quad (4)$$

Therefore, the observed chemical shift would depend on the total concentration ( $C$ ) and temperature, reflecting changes in the position of the equilibrium. An early report<sup>31</sup> describes a change from 303 to 293 ppm in  $\delta^{125}\text{Te}$  when  $\text{Me}_2\text{Te}_2$  was diluted from 2 to 0.02 mol L<sup>-1</sup> in benzene.

The values of  $\delta^{125}\text{Te}$  of ditellurides 2–4 were measured in solution as a function of concentration and temperature. Because of their different solubilities, **3** had to be studied in hexane and toluene only, while **4** could only be dissolved in  $\text{CH}_2\text{Cl}_2$ ; **2** was soluble in all three solvents. The results are graphically presented in Fig. 2. In all cases there is an apparent linear dependence of the  $^{125}\text{Te}$  NMR chemical shift on concentration. Given the significant effect of concentration, comparable in magnitude to the effect of temperature in most instances, the chemical shift values were extrapolated to zero concentration for each temperature. Such chemical shifts at infinite dilution are hereafter denoted as  $\delta_0$ . In each case,  $\delta_0$  exhibits a linear dependence on temperature.

Concentration coefficients measured at selected temperatures are presented in Table 2. Notably, while  $\delta^{125}\text{Te}$  of ditellurides **2** and **3** increases with concentration, irrespective of the solvent, data from the sterically encumbered ditelluride **4** displays the opposite behaviour. Positive slopes (concentration coefficients) for **2** and **3** would be the result of supramolecular interactions if  $\delta_{\text{A}_2} > \delta_{\text{A}}$ . Although less data are available for  $^{125}\text{Te}$  NMR, supramolecular interactions usually cause downfield shifts of  $^{77}\text{Se}$  resonances. In fact, the  $\delta^{77}\text{Se}$  value of

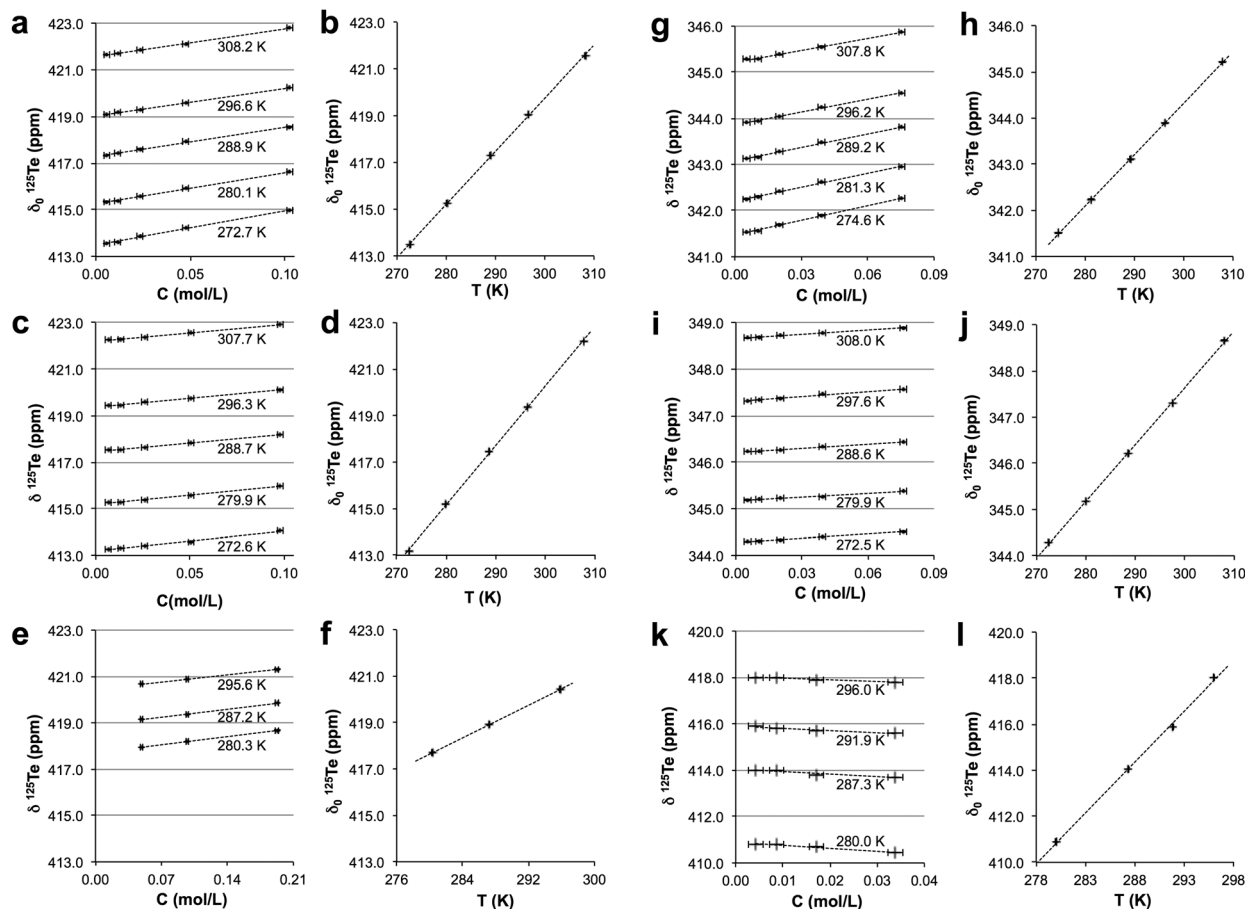


Fig. 2 Effects of concentration and temperature on the  $^{125}\text{Te}$  NMR chemical shift of organo-ditelluride solutions. (a and b) **2** in hexane; (c and d) **2** in toluene; (e and f) **2** in  $\text{CH}_2\text{Cl}_2$ ; (g and h) **3** in hexane; (i and j) **3** in toluene; and (k and l) **4** in  $\text{CH}_2\text{Cl}_2$ . Standard errors in  $\delta^{125}\text{Te}$  ( $< 0.005$  ppm) and  $T$  ( $< 0.05$  K) are not discernible at this scale.



Table 2 Fitted concentration coefficients for  $\delta^{175}\text{Te}$  NMR (ppm) at selected temperatures

Ditelluride	Solvent	$T$ (K)	Coefficient (ppm L mol <sup>-1</sup> )
2	Hexane	296.6	11.8 ± 0.1
2	Toluene	296.3	7.6 ± 0.1
2	CH <sub>2</sub> Cl <sub>2</sub>	295.6	5.04 ± 0.04
3	Hexane	296.2	9.3 ± 0.2
3	Toluene	297.6	3.5 ± 0.2
4	CH <sub>2</sub> Cl <sub>2</sub>	296.0	-7.2 ± 0.9

dialkyl-2-oxazolonyl-diaryldiselenides correlates with the crystallographic Se...N distance.<sup>9</sup>

In contrast, the behaviour of the <sup>125</sup>Te resonance of **4** parallels observations made for telluric acid. Fig. 3 summarizes the dependency on concentration and temperature of the <sup>125</sup>Te chemical shift of Te(OH)<sub>6</sub> in D<sub>2</sub>O. The octahedral arrangement of oxygen atoms in the molecule of telluric acid shields the chalcogen from interactions with other atoms. Although intermolecular hydrogen-bonding between hydroxyl groups is possible, in dilute aqueous solutions the interaction with the solvent is dominant. Here the negative concentration coefficients should result from the increase of magnitude of the diamagnetic (negative) susceptibility of the solution. In dilute solutions the concentration coefficient should be constant (eqn (5), where  $\chi_v$  are the volume magnetic susceptibilities,  $\text{FW}_{\text{sample}}$  and  $\rho_{\text{sample}}$ , respectively, are the formula weight and density of the pure solute). The concentration coefficient estimated in this way from literature<sup>47,48</sup> data for water and telluric acid is -0.5 ppm L mol<sup>-1</sup> at 293 K, which compares well to the -0.4 ± 0.1 ppm L mol<sup>-1</sup> experimentally measured at 296.0 K.

$$\frac{d\delta}{dC} = 4\pi(\chi_{v,\text{sample}} - \chi_{v,\text{solvent}}) \frac{\text{FW}_{\text{sample}}}{\rho_{\text{sample}}} \quad (5)$$

If the preceding arguments hold in the case of sterically encumbered **4**, the steeper negative slope would reflect the much larger formula weight and smaller density than those of telluric acid. From the crystallographic density, the volume susceptibility of **4** is estimated as  $\chi_v = -1.5 \times 10^{-6}$  esu at 296.0 K.

The augmentation of the diamagnetic response upon an increase of concentration must occur as well in solutions of **2** and **3**, but the

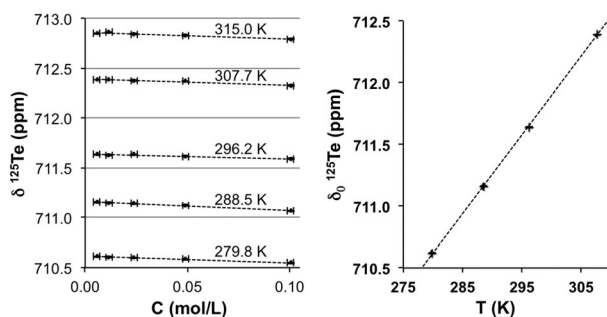


Fig. 3 Effects of concentration (left) and temperature (right) on  $\delta^{125}\text{Te}$  NMR (ppm) of Te(OH)<sub>6</sub> in water. Standard errors in  $\delta^{125}\text{Te}$  (<0.005 ppm) and  $T$  (<0.05 K) are not discernible at this scale.

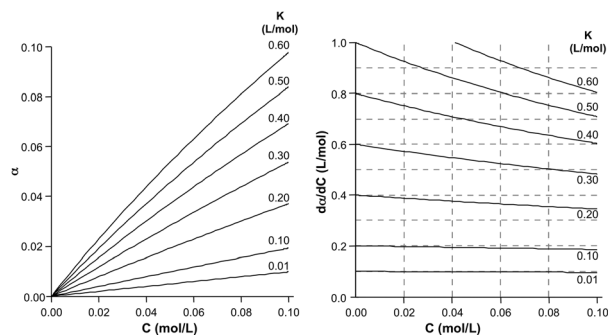


Fig. 4 Calculated fraction of associated molecules ( $\alpha$ ) and the corresponding curve slopes as a function of total concentration ( $C$ ) and magnitude of the dimerization equilibrium constant ( $K$ ).

effect is clearly overshadowed by the high-field shift of the <sup>125</sup>Te resonance. Beyond qualitative arguments, the positive concentration coefficients should be examined in terms of the equilibrium of autoassociation. While the fraction of associated molecules  $\alpha$  does not depend linearly on concentration (eqn (6)), simulations (Fig. 4) show that at low concentrations and small equilibrium constants, the  $\alpha(C)$  curves approach linear behaviour. Indeed it can be shown that the slope (eqn (7)) at low concentrations is nearly constant for  $K \ll 1$  and approaches  $2K$  (eqn (8)), thus the concentration coefficient is  $\approx 2\Delta\delta K$ ,

$$\alpha = 1 + \frac{1 - (1 + 8CK)^{1/2}}{4CK} \quad (6)$$

$$\frac{d\alpha}{dC} = -\frac{1}{C(1 + 8CK)^{1/2}} - \frac{1 - (1 + 8CK)^{1/2}}{4C^2K} \quad (7)$$

$$\lim_{C \rightarrow 0} \frac{d\alpha}{dC} = 2K. \quad (8)$$

The concentration coefficients of **3** in both hexane and toluene are clearly smaller than those of **2** but a direct comparison is not possible because the difference in chemical shifts is unknown. However, it is clear that although the intramolecular Te...N interaction is likely in competition, it does not prevent autoassociation. It is indeed possible for the molecule of **3** in solution to exist in a dynamic equilibrium between conformations with and without intramolecular Te...N interactions; the latter would allow autoassociation. Only the steric hindrance in **4** is effective in this regard.

Measurements performed for **2** in donor solvents like DMSO and pyridine also yielded positive concentration coefficients (Fig. 5),  $18.8 \pm 1.4$  and  $7.9 \pm 0.2$  ppm L mol<sup>-1</sup>, respectively, at 303.0 K; the corresponding  $\delta_0$  values are 403.9 and 422.1 ppm. This observation suggests that autoassociation is still present in these media. However, these experiments were complicated by a chemical reaction. Changes are noticeable in the <sup>1</sup>H NMR spectrum and a white solid is separated from the solutions over time. This process could be slowed by shielding samples from light and storing them in sealed containers to exclude oxygen. Although similar observations have been reported,<sup>49,50</sup> the products are not conclusively identified. The proposed products include a variety of oxygen-containing species (R<sub>2</sub>Te<sub>2</sub>O, R<sub>2</sub>Te<sub>2</sub>O<sub>2</sub>, and



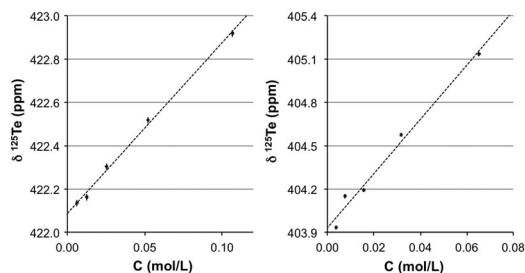


Fig. 5 Effect of the concentration on  $\delta^{125}\text{Te}$  NMR (ppm) of **2** at 303.0 K in solution with: (a) pyridine and (b) DMSO. Standard errors in  $\delta^{125}\text{Te}$  ( $<0.005$  ppm) and  $T$  ( $<0.05$  K) are not discernible at this scale.

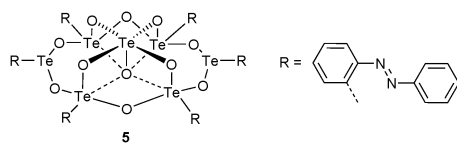


Chart 2

$\text{R}_2\text{Te}_2\text{O}_3$ ) based on infrared spectroscopy and elemental analysis, but it is likely that these are species related to  $5^{51}$  (Chart 2).

### Effect of temperature

Temperature coefficients for  $\delta_0^{125}\text{Te}$  of ditellurides **2–4** in the non-donor solvents and telluric acid in water are compiled in Table 3. The sensitivity of  $\delta_0^{125}\text{Te}$  to temperature correlates with the size of the molecule; compound **4** displays the largest coefficient as expected from the number of vibrational degrees of freedom and molecular weight.

Although in the absence of values of the differences in chemical shift of the individual molecules and their dimers ( $\Delta\delta$ ) in each solvent it is not possible to extract the actual equilibrium constants, a van't Hoff plot based on the observed concentration coefficients would provide an estimate of  $\Delta H$ . The fitted value for **3** in hexane solution (Fig. 6) of  $-4.1 \pm 0.2 \text{ J mol}^{-1}$  is small and is likely biased by solvation and the lack of a diamagnetic correction. The relative errors for the  $\Delta H$  values fitted from the other measurement sets are too large for a comparison but it is clear that the association of the molecules of **2** is exothermic in all three solvents.

### Aggregation vs. other effects

Compared to the range of possible values of  $^{125}\text{Te}$  NMR chemical shift, the influence of concentration on  $\delta$  is small

Table 3 Fitted temperature coefficients for the infinite dilution  $^{125}\text{Te}$  NMR chemical shifts ( $\delta_0$ ) and interpolated values at 298.0 K

Compound	Solvent	Coefficient $\times 10^2$ (ppm $\text{K}^{-1}$ )	$\delta_0$ at 298.0 K (ppm)
<b>2</b>	Hexane	$22.7 \pm 0.2$	419.3
<b>2</b>	Toluene	$25.5 \pm 0.3$	419.8
<b>2</b>	$\text{CH}_2\text{Cl}_2$	$17.9 \pm 0.3$	420.8
<b>3</b>	Hexane	$11.2 \pm 0.1$	344.1
<b>3</b>	Toluene	$12.3 \pm 0.2$	347.4
<b>4</b>	$\text{CH}_2\text{Cl}_2$	$44 \pm 1$	418.8
$\text{Te}(\text{OH})_6$	Water	$6.38 \pm 0.04$	711.8

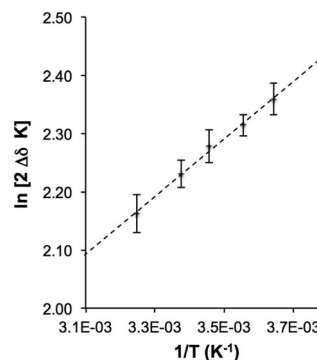


Fig. 6 van't Hoff plot for the proposed equilibrium of autoassociation of **3** in hexane. The standard error in  $T$  ( $<0.05$  K) is not discernible at this scale.

and should be compared with other effects that originate in the local environment of the molecule.

The first is the structure of the molecule itself. Table 1 illustrates the wide range of values the C–Te–Te–C dihedral angle can adopt. This internal coordinate is especially important because of its effect on the energy ( $\Delta E$ ) of the first electronic excitation, LUMO  $\leftarrow$  HOMO, which in turn impacts the shielding ( $\sigma$ ) according to Ramsey's equation (eqn (9)). The HOMO is particularly sensitive to the torsion angle because it consists of a combination of the lone-pair orbitals on the tellurium atoms (Fig. 1). TD-DFT/DFT-GIAO calculations for compound **1** estimate  $\Delta E$  values from 2.9 to 1.7 eV with concomitant changes in  $\delta^{125}\text{Te}$  within the range 350–750 ppm for C–Te–Te–C torsion angles of 0–180°. The same conformational variation would result in molecular dipole moments ranging from 0 to 3.0 D,

$$\sigma = \frac{1}{3} \frac{e^2}{2mc^2} \left[ \begin{array}{c} \left\langle 0 \left| \frac{y^2 + z^2}{r^3} \right| 0 \right\rangle + \\ \left\langle 0 \left| \frac{x^2 + z^2}{r^3} \right| 0 \right\rangle + \\ \left\langle 0 \left| \frac{x^2 + y^2}{r^3} \right| 0 \right\rangle \end{array} \right] + \left[ \frac{2}{3} \left( \frac{e\hbar}{mc} \right) \left\langle \frac{1}{r^3} \right\rangle \frac{1}{\Delta E} \right] \quad (9)$$

In spite of the dependence of the molecular dipole moment on the conformation of ditelluride, COSMO-DFT calculations predict changes less than 4.2° from low (hexane) to high (DMSO) polarity solvents. Even so, these would translate into  $\delta^{125}\text{Te}$  changes of up to 40 ppm. Experimentally, we found variations within a 50 ppm-wide window. The experimental solvatochromic effect is not larger than 0.20 eV, while COSMO TD-DFT calculations predict less than 0.1 eV.

## Experimental section

### Materials and methods

The manipulation of air-sensitive materials was carried out in a glove box or using standard Schlenk techniques under an atmosphere of nitrogen. Solvents were purified by using an Innovative Technologies purification system or by distillation



over appropriate dehydrating agents. Elemental tellurium (Cerac) and deuterated solvents (Aldrich) were used as received without further purification. Telluric acid (Aldrich) was purified by recrystallization. The laboratory sample of dimethyl telluride was handled within a sealed quartz NMR tube at all times.

Ditellurides **2**,<sup>52</sup> **3**,<sup>33</sup> and **4**<sup>46</sup> were prepared by methods adapted from the literature. In general, elemental tellurium was treated with organolithium or Grignard reagents to give a telluroate which was oxidized by I<sub>2</sub> or O<sub>2</sub>. The final products were extracted in an aqueous work up and rigorously purified by recrystallization, sublimation or chromatography.

NMR spectra were acquired on a Bruker Avance 500 (500 MHz). A Bruker 5 mm broad band inverse probe was used to acquire <sup>1</sup>H spectra and a 5 mm broad band probe was used to acquire <sup>13</sup>C{<sup>1</sup>H} and <sup>125</sup>Te spectra. Chemical shifts are reported in  $\delta$  units and are referenced to the residual resonances of the solvent in the cases of <sup>1</sup>H and <sup>13</sup>C. To ensure the highest accuracy possible,  $\delta_0$  <sup>125</sup>Te is referenced to the absolute frequency of resonance interpolated to 298.0 K from measurements of a sample of pure Me<sub>2</sub>Te between 270.0 and 305.0 K; within this range the resonance depends linearly ( $1.03 \pm 0.01 \times 10^{-1}$  ppm K<sup>-1</sup>) on temperature, in agreement with the literature.<sup>53</sup> Corrections for the diamagnetic susceptibility of each solvent were performed using eqn (10). The volume susceptibility of dimethyl telluride ( $\chi_{v,\text{reference}}$ ) was approximated using literature methods,<sup>54</sup>

$$\delta_{\text{corr}} = \delta - \frac{4\pi}{3}(\chi_v - \chi_{v,\text{reference}}). \quad (10)$$

Variable temperature spectra were acquired under either a cold or an ambient temperature gas flow with a BV-T 2000 variable temperature controller. Temperature at the probe was measured by monitoring the chemical shift of a 4% solution of methanol in methanol-d<sub>4</sub> between 178 and 300 K and 100% ethylene glycol between 300 and 380 K.<sup>55</sup>

### Ultraviolet-visible spectroscopy

Absorption spectra were measured in quartz cuvettes on a Varian Cary 300 spectrometer with a temperature-controlled stage in dual beam mode from 200 to 800 nm with a scan rate of 10 nm s<sup>-1</sup>. The solvent background was subtracted using external blank samples.

### Computational details

All calculations were carried out using the Amsterdam Density Functional Package.<sup>56</sup> For computational expediency all calculations were carried out within the general gradient approximation using the PW91<sup>57</sup> functional, while relativistic effects were approached using the Zeroth Order Regular Approximation.<sup>58-62</sup> In preliminary geometry optimizations it was established that while most optimized bond distances, bond angles and torsion angles satisfactorily reproduced the experimental values the C-Te-Te-C torsion angle required a basis set of quadruple- $\zeta$  quality with four polarization sets (QZ4P) to be calculated within 0.5 degrees of the experimental value. Solvent properties were modeled using the Conductor like Screening Model<sup>63-65</sup> (COSMO). Time dependent calculations were used to evaluate the first 100 excitations in the electronic spectrum using the

Davidson diagonalization method and the statistical average of model potentials<sup>66</sup> (SAOP). GIAO NMR shielding constants were calculated using the NMR ADF module.<sup>67-69</sup>

## Conclusions

The solution <sup>125</sup>Te NMR chemical shift of all the compounds included in this study is dependent on concentration. While negative concentration coefficients are explicable by changes in the diamagnetism of the sample solutions, positive coefficients are consistent with independent observations of the effect of chalcogen-centred supramolecular interactions. Such intermolecular interactions could only be prevented when the tellurium atom is inaccessible because of steric protection. The influence of concentration on the measured  $\delta$  <sup>125</sup>Te, however, is dwarfed by the effects of conformational changes or the dielectric properties of the medium. This observation has important implications for spectroscopic studies of supramolecular interactions of tellurium compounds. For example, studying the association of Lewis bases with these or other tellurium-containing molecules in solution would likely require a significant excess of base that could easily alter the dielectric constant of the medium and mask the effect of association on  $\delta$  <sup>125</sup>Te.

Albeit small, the effect of concentration is enough to cause variations that can complicate the comparison and interpretation of <sup>125</sup>Te NMR data, reporting precise chemical shifts requires specifying the concentration in addition to temperature or stating the value at infinite dilution. The current IUPAC standard for <sup>125</sup>Te NMR is "a solution of less than 10% benzene-d<sub>6</sub> in dimethyl telluride".<sup>70</sup> Aside from its high toxicity, (CH<sub>3</sub>)<sub>2</sub>Te (**1**) poses a fundamental challenge: its molecules are very likely to undergo autoassociation equilibria because of the two  $\sigma$  holes on tellurium,<sup>71</sup> the orientation of the  $\sigma^*_{\text{Te-C}}$  orbitals<sup>72</sup> and the high polarizability of the chalcogen.<sup>36</sup> Although the use of an absolute scale, in which the resonances are referenced to the frequency of TMS protons, has been suggested,<sup>73</sup> a simpler alternative would be to specify a precise composition and temperature for the standard.

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## Notes and references

- G. R. Desiraju, P. S. Ho, L. Kloo, A. C. Legon, R. Marquardt, P. Metrangolo, P. Politzer, G. Resnati and K. Rissanen, *Pure Appl. Chem.*, 2013, **85**, 1711-1713.
- A. Priimagi, G. Cavallo, P. Metrangolo and G. Resnati, *Acc. Chem. Res.*, 2013, **46**, 2686-2695.



- 3 P. Metrangolo and G. Resnati, *Cryst. Growth Des.*, 2012, **12**, 5835–5838.
- 4 V. R. Pedireddi, D. S. Reddy, B. S. Goud, D. C. Craig, A. D. Rae and G. R. Desiraju, *J. Chem. Soc., Perkin Trans. 2*, 1994, 2353–2360.
- 5 *Halogen Bonding*, ed. P. Metrangolo and G. Resnati, Springer Berlin Heidelberg, Berlin, Heidelberg, 2008, vol. 126.
- 6 K. Rissanen, *CrystEngComm*, 2008, **10**, 1107–1113.
- 7 P. Metrangolo and G. Resnati, *Chem. – Eur. J.*, 2001, **7**, 2511–2519.
- 8 P. Metrangolo, H. Neukirch, T. Pilati and G. Resnati, *Acc. Chem. Res.*, 2005, **38**, 386–395.
- 9 M. Fourmigué, *Curr. Opin. Solid State Mater. Sci.*, 2009, **13**, 36–45.
- 10 P. Metrangolo, G. Resnati, T. Pilati, R. Liantonio and F. Meyer, *J. Polym. Sci., Part A: Polym. Chem.*, 2006, **45**, 1–15.
- 11 M. Erdelyi, *Chem. Soc. Rev.*, 2012, **41**, 3547–3557.
- 12 P. Metrangolo, *Curr. Opin. Colloid Interface Sci.*, 2003, **8**, 215–222.
- 13 P. Metrangolo, F. Meyer, T. Pilati, G. Resnati and G. Terraneo, *Angew. Chem., Int. Ed.*, 2008, **47**, 6114–6127.
- 14 R. Bertani, P. Sgarbossa, A. Venzo, F. Lelj, M. Amati, G. Resnati, T. Pilati, P. Metrangolo and G. Terraneo, *Coord. Chem. Rev.*, 2010, **254**, 677–695.
- 15 P. Metrangolo and G. Resnati, *Science*, 2008, **321**, 918–919.
- 16 K. Raatikainen and K. Rissanen, *Chem. Sci.*, 2012, **3**, 1235–1239.
- 17 O. Bolton, K. Lee, H.-J. Kim, K. Y. Lin and J. Kim, *Nat. Chem.*, 2011, **3**, 205–210.
- 18 M. Virkki, O. Tuominen, A. Forni, M. Saccone, P. Metrangolo, G. Resnati, M. Kauranen and A. Priimagi, *J. Mater. Chem. C*, 2015, **3**, 3003–3006.
- 19 F. Kniep, S. H. Jungbauer, Q. Zhang, S. M. Walter, S. Schindler, I. Schnapperelle, E. Herdtweck and S. M. Huber, *Angew. Chem., Int. Ed.*, 2013, **52**, 7028–7032.
- 20 D. Manna and G. Mugesh, *J. Am. Chem. Soc.*, 2012, **134**, 4269–4279.
- 21 T. Chivers and R. S. Laitinen, *Chem. Soc. Rev.*, 2015, **44**, 1725–1739.
- 22 P. Politzer, J. Murray, G. Janjić and S. Zarić, *Crystals*, 2014, **4**, 12–31.
- 23 E. R. T. Tiekink and J. Zukerman-Schpector, *Coord. Chem. Rev.*, 2010, **254**, 46–76.
- 24 A. F. Cozzolino, N. E. Gruhn, D. L. Lichtenberger and I. Vargas-Baca, *Inorg. Chem.*, 2008, **47**, 6220–6226.
- 25 A. F. Cozzolino, G. Dimopoulos-Italiano, L. M. Lee and I. Vargas-Baca, *Eur. J. Inorg. Chem.*, 2013, 2751–2756.
- 26 G. E. Garrett, G. L. Gibson, R. N. Straus, D. S. Seferos and M. S. Taylor, *J. Am. Chem. Soc.*, 2015, **137**, 4126–4133.
- 27 G. E. Garrett, E. I. Carrera, D. S. Seferos and M. S. Taylor, *Chem. Commun.*, 2016, **52**, 9881–9884.
- 28 P. C. Ho, P. Szydłowski, J. Sinclair, P. J. W. Elder, J. Kübel, C. Gendy, L. M. Lee, H. Jenkins, J. F. Britten, D. R. Morim and I. Vargas-Baca, *Nat. Commun.*, 2016, **7**, 11299.
- 29 P. Enghag, *Encyclopedia of the Elements*, Wiley, Weinheim, 2008.
- 30 P. E. Lazlo, *NMR of newly accessible nuclei*, Academic Press, New York, 1983.
- 31 H. C. E. McFarlane and W. McFarlane, *J. Chem. Soc., Dalton Trans.*, 1973, 2416–2418.
- 32 G. J. Schrobilgen, R. C. Burns and P. Granger, *J. Chem. Soc., Chem. Commun.*, 1978, 957–960.
- 33 G. Mugesh, A. Panda, S. Kumar, S. D. Apte, H. B. Singh and R. J. Butcher, *Organometallics*, 2002, **21**, 884–892.
- 34 F. H. Kruse, R. E. Marsh and J. D. McCullough, *Acta Crystallogr.*, 1957, **10**, 201–209.
- 35 K. Kandasamy, S. Kumar, H. B. Singh and G. Wolmershäuser, *Organometallics*, 2003, **22**, 5069–5078.
- 36 C. Bleiholder, D. B. Werz, H. Köppel and R. Gleiter, *J. Am. Chem. Soc.*, 2006, **128**, 2666–2674.
- 37 S. Ludlow and A. E. McCarthy, *J. Organomet. Chem.*, 1981, **219**, 169–176.
- 38 S. K. Kumar, H. B. Singh, K. Das and U. C. Sinha, *J. Organomet. Chem.*, 1990, **397**, 161–167.
- 39 Y. Aso, K. Yui, T. Miyoshi, T. Otsubo, F. Ogura and J. Tanaka, *Bull. Chem. Soc. Jpn.*, 1988, **61**, 2013–2018.
- 40 K. K. Bhasin, V. Arora, T. M. Klapötke and M. J. Crawford, *Eur. J. Inorg. Chem.*, 2004, 4781–4788.
- 41 A. Edelmann, S. Brooker, N. Bertel, M. Noltemeyer, H. W. Roesky, G. M. Sheldrick and F. T. Edelmann, *Z. Naturforsch., B: Chem. Sci.*, 1992, **47**, 305–309.
- 42 S. K. Tripathi, U. Patel, D. Roy, R. B. Sunoj, H. B. Singh, G. Wolmershäuser and R. J. Butcher, *J. Org. Chem.*, 2005, **70**, 9237–9247.
- 43 R. A. Zingaro, C. Herrera and E. A. Meyers, *J. Organomet. Chem.*, 1986, **306**, C36–C40.
- 44 W.-W. du Mont, A. Martens-von Salzen, F. Ruthe, E. Seppälä, G. Mugesh, F. A. Devillanova, V. Lippolis and N. Kuhn, *J. Organomet. Chem.*, 2001, **623**, 14–28.
- 45 R. Kaur, H. B. Singh and R. J. Butcher, *Organometallics*, 1995, **14**, 4755–4763.
- 46 T. M. Klapötke, B. Krumm, H. Nöth, J. C. Gálvez-Ruiz, K. Polborn, I. Schwab and M. Suter, *Inorg. Chem.*, 2005, **44**, 5254–5265.
- 47 E. W. Washburn, *International Critical Tables of Numerical Data, Physics, Chemistry and Technology*, Knovel, 1st electronic edn, 2003, app.knovel.com.
- 48 O. Lindqvist, *Acta Chem. Scand.*, 1970, **24**, 3178–3188.
- 49 M. Sharma, K. K. Bhasin, S. K. Mehta, N. Singh and S. Kumar, *Radiat. Phys. Chem.*, 2006, **75**, 2029–2038.
- 50 R. T. Mehdi, D. Miller and T. A. Tahir, *Inorg. Chim. Acta*, 1984, **90**, 85–89.
- 51 K. Srivastava, S. Sharma, H. B. Singh, U. P. Singh and R. J. Butcher, *Chem. Commun.*, 2010, **46**, 1130–1132.
- 52 G. Llabres, O. Dideberg and L. Dupont, *Acta Crystallogr., Sect. B: Struct. Crystallogr. Cryst. Chem.*, 1972, **28**, 2438–2444.
- 53 P. Granger and S. Chapelle, *J. Magn. Reson.*, 1980, **39**, 329–334.
- 54 G. A. Bain and J. F. Berry, *J. Chem. Educ.*, 2008, **85**, 532.
- 55 C. Ammann, P. Meier and A. Merbach, *J. Magn. Reson.*, 1982, **46**, 319–321.
- 56 G. te Velde, F. M. Bickelhaupt, E. J. Baerends, C. Fonseca Guerra, S. J. A. van Gisbergen, J. G. Snijders and T. Ziegler, *J. Comput. Chem.*, 2001, **22**, 931–967.



- 57 J. P. Perdew, J. A. Chevary, S. H. Vosko, K. A. Jackson, M. R. Pederson, D. J. Singh and C. Fiolhais, *Phys. Rev. B: Condens. Matter Mater. Phys.*, 1992, **46**, 6671–6687.
- 58 E. van Lenthe, A. Ehlers and E.-J. Baerends, *J. Chem. Phys.*, 1999, **110**, 8943–8953.
- 59 E. Van Lenthe, E. J. Baerends and J. G. Snijders, *J. Chem. Phys.*, 1993, **99**, 4597–4610.
- 60 E. Van Lenthe, E. J. Baerends and J. G. Snijders, *J. Chem. Phys.*, 1994, **101**, 9783–9792.
- 61 E. Van Lenthe, J. G. Snijders and E. J. Baerends, *J. Chem. Phys.*, 1996, **105**, 6505–6516.
- 62 E. Van Lenthe, R. van Leeuwen, E. J. Baerends and J. G. Snijders, *Int. J. Quantum Chem.*, 1996, **57**, 281–293.
- 63 A. Klamt and G. Schüürmann, *J. Chem. Soc., Perkin Trans. 2*, 1993, 799–805.
- 64 A. Klamt, *J. Phys. Chem.*, 1995, **99**, 2224–2235.
- 65 A. Klamt and V. Jonas, *J. Chem. Phys.*, 1996, **105**, 9972–9981.
- 66 O. V. Gritsenko, P. R. T. Schipper and E. J. Baerends, *Chem. Phys. Lett.*, 1999, **302**, 199–207.
- 67 G. Schreckenbach and T. Ziegler, *J. Phys. Chem.*, 1995, **99**, 606–611.
- 68 G. Schreckenbach and T. Ziegler, *Int. J. Quantum Chem.*, 1996, **60**, 753–766.
- 69 G. Schreckenbach and T. Ziegler, *Int. J. Quantum Chem.*, 1997, **61**, 899–918.
- 70 R. K. Harris, E. D. Becker, S. M. C. De Menezes, P. Granger, R. E. Hoffman and K. W. Zilm, *Magn. Reson. Chem.*, 2008, **46**, 582–598.
- 71 P. Politzer, J. S. Murray and T. Clark, *Phys. Chem. Chem. Phys.*, 2013, **15**, 11178–11189.
- 72 A. F. Cozzolino, P. J. W. Elder and I. Vargas-Baca, *Coord. Chem. Rev.*, 2011, **255**, 1426–1438.
- 73 R. K. Harris, E. D. Becker, S. M. Cabral de Menezes, R. Goodfellow and P. Granger, *Pure Appl. Chem.*, 2001, **73**, 1795–1818.

