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We demonstrate two novel methods for the measurement of the temperatures of reaction spaces locally heated by microwaves, which have been applied here to two example systems, *i.e.*, BaTiO₃ particles covered with a SiO₂ shell (BaTiO₃–SiO₂) and layered tungstate particles. Photoluminescent (PL) probes showing the temperature-sensitivity in their PL lifetimes are located in the nanospaces of the above systems. In the case of BaTiO₃–SiO₂ core–shell particles, rhodamine B is loaded into the mesopores of the SiO₂ shell covering the BaTiO₃ core, which generates the heat through the dielectric loss of microwaves. The inner nanospace temperature of the SiO₂ shell is determined to be 28 °C higher than the bulk temperature under microwave irradiation at 24 W. On the other hand, Eu³⁺ is immobilized in the interlayer space of layered tungstate as the PL probe, showing that the nanospace temperature of the interlayer is only 4 °C higher than the bulk temperature. This method for temperature-measurement is powerful for controlling microwave heating and elucidates the ambiguous mechanisms of microwave special effects often observed in chemical reactions, contributing greatly to the practical application of microwaves in chemistry and materials sciences.

Introduction

Microwave irradiation has been attracting a lot of interest in chemistry and materials science as a useful tool for shortening the reaction and processing durations, and reducing temperatures. When microwaves pass through a solid, microwave energy is converted to thermal energy, *i.e.*, heat generation due to

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In situ temperature measurements of reaction spaces under microwave irradiation using photoluminescent probes†

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dielectric loss,¹ magnetic loss and Joule loss. Unlike conventional heating methods, microwaves heat solids directly, resulting in rapid heating, superheating, and substance-selective heating, which induce specific temperature distribution in the heated systems.^{2–12} However, it has also raised some controversial discussions on these advantages in using microwaves by often being called “microwave special effects” based on unusual experimental results which could not be explained fully by the heat generation due to microwaves.^{13–18} Recently, Wada’s group reported the acceleration of photo-induced electron transfer by microwaves observed for photoexcited CdS quantum dots.¹⁸ “Microwave special effects” also called “non-thermal effects” have caused a controversy among microwave chemists. It was pointed out that the reports showing the non-thermal effects were often based on inaccurate temperature measurements. Kappe *et al.* demonstrated that infrared sensors often detected the temperature of the outer surface of a reaction vessel which was different from the temperature of the inner reaction field in the reaction vessel.^{19,20}

We can find some challenging trials to estimate the accurate temperature of reaction fields heated by microwaves. Sumerlin *et al.* demonstrated the temperature distribution of the surroundings nearby Au nanoparticles after microwave irradiation at 100 W by using a PL probe and azo-bond.²¹ The temperature difference between the particle and the surrounding was estimated to be 70 °C at a distance of 0.5 nm from the particle surface, but it was 0 °C at that of 2 nm. Wada *et al.* observed non-equilibrium local heating of cobalt metal particles by measuring the Raman spectra of the surrounding molecules as a temperature probe under microwave irradiation.²² The Co particles were dispersed in dimethylsulfoxide (DMSO) solvent, and the temperature of the DMSO molecules in close proximity to the Co surface was determined to be higher by 70 °C compared to the bulk temperature at the maximum. However, an *in situ* method for the measurement of the accurate temperature has yet to be developed for understanding the “microwave special effects”.

We focused on the PL probes whose PL lifetimes are sensitive to temperature as the contactless molecular thermometers.^{23–27}



This study demonstrates *in situ* temperature-measurement methods based on the temperature dependence of the PL lifetimes of probe molecules under microwave irradiation for two submicroparticle systems: BaTiO₃–SiO₂ core–shell particles and layered tungstate particles intercalated with ammonium cations (C₁₆N⁺–W₂O₇^{2–}). The BaTiO₃ core is ferroelectric acting as a microwave absorber²⁸ and the SiO₂ shell provides BaTiO₃ with mesopores functioning as a reaction space accommodating the reaction molecules. The SiO₂ shell is transparent to microwaves and has low thermal conductivity, leading to the accumulation of heat generated in the BaTiO₃ core. We loaded RhB as the PL probe in the SiO₂ mesopores to measure the nanospace temperature of the inner SiO₂ shell (denoted as BaTiO₃–RhB@SiO₂ hereafter). The layered tungstate in C₁₆N⁺–W₂O₇^{2–}, reported recently as a new material, absorbs microwaves strongly.²⁹ Ammonium cations are intercalated into the interlayer for compensating the W₂O₇^{2–} anion layer. This interlayer with a distance of several nanometers acts as a reaction space for photocatalysis. We immobilized the Eu³⁺ cation as the PL probe in the interlayer *via* ion exchange of three C₁₆N⁺ cations with one Eu³⁺ (denoted as C₁₆N⁺–W₂O₇^{2–}–Eu³⁺ hereafter). BaTiO₃–RhB@SiO₂ and C₁₆N⁺–W₂O₇^{2–}–Eu³⁺ were dispersed in heptane respectively as in heterogeneous catalysis systems which can utilize the microwave energy very efficiently in their catalytic reactions. In these particles, non-equilibrium local heating can occur effectively, leading to the dramatic acceleration of chemical reactions under microwave irradiation. The purpose of this work is to demonstrate a trial of measuring the temperatures of the nanospaces expected to work as reaction fields under microwave irradiation.

Experimental

BaTiO₃–SiO₂ core–shell particles were synthesized *via* a procedure adapted from the Stober method.³⁰ Barium titanate (Wako Pure Chemical Industries Ltd., 162 mg) was dispersed in a mixture of ultra-pure water (0.84 ml) and ethanol (4.32 ml) under ultrasonification for 15 minutes. After hexadecyl trimethyl ammonium bromide (Tokyo Chemical Industry Co., Ltd., 0.91 mg) was added into this dispersion, the ultrasonic dispersion was conducted again for 15 minutes. While the mixture was being stirred using a stirring bar, ammonia water (Wako Pure Chemical Industries, Ltd., 15.54 µl of 28%) and tetraethyl orthosilicate (Tokyo Chemical Industry Co., Ltd., 315 µl) were added. The mixture was heated at 40 °C in 12 hours under stirring. The resulting powder was washed with water and ethanol and dried under vacuum, giving BaTiO₃–SiO₂ core–shell particles. RhB was introduced into the mesopores of the SiO₂ shell using an immersion method for obtaining BaTiO₃–RhB@SiO₂. BaTiO₃–SiO₂ particles were immersed in RhB ($\geq 95\%$, Sigma-Aldrich Co.) aqueous solution (0.3 ml of 0.1 mM RhB) under ultrasonication for 30 minutes. The red-coloured powder was washed with water until the supernatant liquid became colourless. The powder was dried, giving BaTiO₃–RhB@SiO₂.

H₂W₂O₇ was synthesized *via* a procedure reported in ref. 29. A mixture of bismuth(III) oxide (Wako Pure Chemical Industries,

Ltd., 15.1 g) and tungsten(III) oxide (Kojundo Chemical Laboratory Co., Ltd., 15.0 g) was calcined at 800 °C for 20 hours, giving Bi₂W₂O₉, as a precursor of H₂W₂O₇. Then, Bi₂W₂O₉ (10 g) was dispersed in 6 M hydrochloric acid, and heated by microwaves using MicroSYNTH (Milestone Inc.) at 80 °C for one hour. The resulting powder was washed with water, dried for 6 hours under vacuum and then dried under air for 1 day, giving H₂W₂O₇. Hexadecylamine was intercalated into H₂W₂O₇, giving C₁₆N⁺–W₂O₇^{2–}, through the acid–base reaction of hexadecylamine with protons in the interlayer to form hexadecylammonium cations. A mixture of H₂W₂O₇ (1.0 g) and hexadecylamine (Tokyo Chemical Industry Co., Ltd., 5.0 g) in 50 ml of heptane was stirred for 6 hours at room temperature. The resulting powder was washed with acetone three times, dried for 6 hours under vacuum, and then dried under air for 1 day, giving C₁₆N⁺–W₂O₇^{2–}. C₁₆N⁺–W₂O₇^{2–}–Eu³⁺ was obtained *via* an ion exchange of three C₁₆N⁺ cations with one Eu³⁺. A mixture of C₁₆N⁺–W₂O₇^{2–} (0.50 g) and europium(III) chloride hexahydrate (Kanto Chemical Co. Inc., 1.27 g) in 50 ml of 50% ethanol water was stirred for 48 hours at room temperature. The resulting powder was washed with water and ethanol, giving C₁₆N⁺–W₂O₇^{2–}–Eu³⁺ after drying.

Scanning electron microscopy (SEM) images and a cross-sectional SEM image were collected using an S-5500 (Hitachi High-Technologies Co.). The sample for the cross-sectional SEM was obtained by Ar⁺ sputtering with an Al substrate using an SM-09020CP (JEOL Ltd.), and modified using a conductive surface, 2 nm thickness of OsO₄. Cross-sectional transmission electron microscopy (TEM) images were collected using a JEM-2010 (JEOL Ltd.). The sample was sliced using UMT-UC7 ultramicrotomes (Leica Microsystems GmbH). Nitrogen adsorption was measured using a Belsorp-mini (MicrotracBEL Co.). The samples were purged by N₂ gas at 423 K. The UV-vis spectra were collected using a V-670 (JASCO Co.) equipped with a diffused reflex unit, ISN-470 (JASCO Co.). The emission spectra and excitation spectra were collected using a F-7000 (Hitachi High-Technologies Co.). X-ray diffraction spectra (XRD) were collected using a Miniflex (Rigaku Co.) and X'Pert MPD-OEC diffractometer (PANalytical B.V.). Elemental analysis was conducted using a CHN corder MT-6 (Yanaco Technical Science Co., Ltd.) and Prodigy ICP (ICP-OES) (Teledyne Leeman Labs). For ICP-OES, the sample was dissolved in a mixture of hydrofluoric acid, hydrochloric acid and nitric acid.

In situ temperature measurements of the reaction nanospaces were conducted as described below. BaTiO₃–RhB@SiO₂ (100 mg) or C₁₆N⁺–W₂O₇^{2–}–Eu³⁺ (20 mg) dispersed in heptane (4 ml) was placed in a rectangular four-sided quartz cell with an optical length of 1 cm.

For the measurement of PL under conventional heating, the above cell was placed in a spectroscopic cryostat, CoolSpek UV (Unisoku Co. Ltd.), with a temperature control unit, CS-AT-HT (Unisoku Co. Ltd.). This cryostat was set at the measuring position in a PL lifetime measuring instrument, Quantaurus-Tau (Hamamatsu Photonics K.K.). Solutions were stirred using a cuboid bar (1 mm × 1 mm × 8 mm) during the measurements. The temperatures of the bulk solutions were measured using a fiber-optic thermometer.



The PL lifetimes were derived by fitting the PL decays to a one component exponential function, taking the instrument response function obtained using a colloidal silica, Ludox[®] HS40 (Sigma-Aldrich Co.) dispersion in water. For BaTiO₃-SiO₂, the PL decay in a 3.6 to 10.1 ns range was used to determine the PL lifetime by fitting. For C₁₆N⁺-W₂O₇²⁻-Eu³⁺, the PL decay from 100 to 502 μ s was used. For the *in situ* temperature measurements of the PL decays under microwave irradiation, the same equipment as that previously reported by our group¹⁸ (Fig. S1, ESI[†]) was used. An ellipsoid microwave applicator (Chronix Ltd.) and a PL lifetime measuring instrument, Quantaurus-Tau, were assembled to form a system for *in situ* measurements of PL under microwave irradiation as shown in Fig. S1a and b (ESI[†]). The above cell was placed in the chamber of the microwave applicator. The chamber was drilled to permit a PL path. The fiber-optic thermometer was inserted into the cell and placed at the upper position of the cell where microwaves were not applied. The solutions were stirred using a cuboid bar (1 mm \times 1 mm \times 8 mm) during measurements.

The PL lifetimes under continuous microwave irradiation were measured many times continuously during the microwave irradiation. To minimize the variation of the temperatures, all measurements of the PL decays were conducted setting the cell at the same conditions, *i.e.*, position, stirring rate and PL counts. Especially, it should be emphasized that PL measurements were carried out 10 min after starting the stirring of the solutions for stabilizing the optical conditions of the solutions.

Results and discussion

The preparation scheme of BaTiO₃-RhB@SiO₂ core-shell particles is shown in Fig. 1a. The core size distribution of BaTiO₃-RhB@SiO₂ (Fig. 1b) was determined from the SEM images

of BaTiO₃ as the precursor (Fig. S4, ESI[†]). The shell thickness distribution of BaTiO₃-RhB@SiO₂ (Fig. 1c) was determined from the cross-section of the TEM image of BaTiO₃-SiO₂ (Fig. S5, ESI[†]). The cross-sectional SEM image is also shown in Fig. S6 (ESI[†]). These observations gave the averages of the core size and the shell thickness of BaTiO₃-SiO₂ as 447 nm and 10.4 nm, respectively. The nitrogen adsorption isotherms of BaTiO₃ and BaTiO₃-SiO₂ were shown to be of III type (Fig. S7, ESI[†]), essentially indicating the absence of pore structures in these particles. However, the hysteresis and the increased adsorption observed at high P/P_0 for the isotherm of BaTiO₃-SiO₂ showed that the SiO₂ shell of BaTiO₃-SiO₂ contained mesopores to a certain extent. The amount of RhB loaded in BaTiO₃-SiO₂ was calculated to be 183 mmol g⁻¹ BaTiO₃-SiO₂ from the RhB decrement in the immersion method using UV-vis spectroscopy. Since BaTiO₃-RhB@SiO₂ was washed with ultra-pure water many times, the RhB present on the outer surface of the SiO₂ shell was removed and only RhB in the SiO₂ mesopores remained. The excitation spectrum and the emission spectrum of BaTiO₃-RhB@SiO₂ were measured and are shown in Fig. 1d.

First we needed to obtain the relationship between the PL lifetime and the temperature in order to convert the PL lifetime to the temperature for BaTiO₃-RhB@SiO₂. PL lifetime measurements were carried out using the PL decay at 580 nm of RhB excited by 5 MHz pulse LED light at 405 nm. The PL decay was fitted using only one exponential component, giving the PL lifetime. Fig. 2a shows the PL decays of BaTiO₃-RhB@SiO₂ at 27, 40, 50, 60 and 70 °C under conventional heating. The PL lifetime value at each temperature was obtained as the average of the values repeated three times. The PL lifetimes at 27, 40, 50, 60 and 70 °C are plotted against the temperatures, giving a straight line. This relationship was used for obtaining the

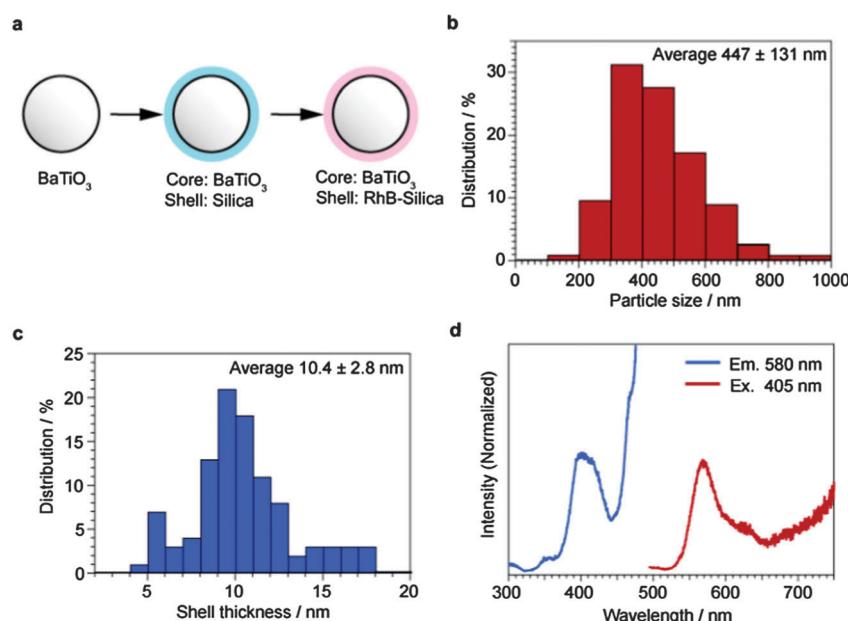


Fig. 1 Synthesis of BaTiO₃-RhB@SiO₂. (a) Synthesis scheme, (b) core size distribution, (c) shell thickness distribution, and (d) excitation spectrum and emission spectrum.



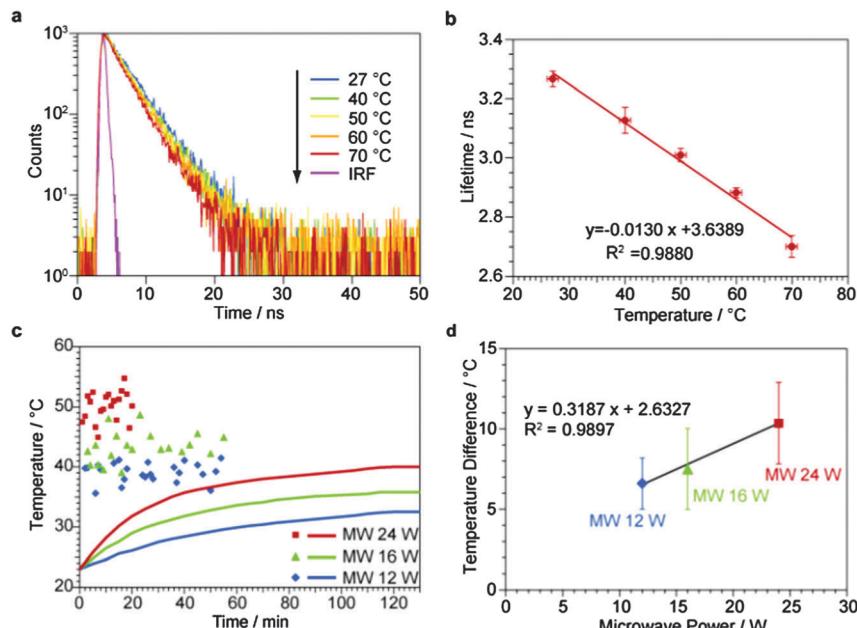


Fig. 2 *In situ* temperature measurements for $\text{BaTiO}_3\text{-RhB@SiO}_2$ under microwave irradiation. (a) PL decays at 27, 40, 50, 60 and 70 °C under conventional heating. (b) Calibration line between the lifetime and the temperature. (c) Nanospace temperatures in the SiO_2 mesopores plotted against the microwave irradiation time at 12, 16 and 24 W. Bulk temperatures shown as the curves. (d) Temperature differences between the nanospace and the bulk under microwave irradiation in the stable state plotted against the microwave power.

temperature experienced by RhB present in the SiO_2 shell of $\text{BaTiO}_3\text{-RhB@SiO}_2$ under microwave irradiation. The relationship between the temperature and the PL lifetime of RhB in ethanol solution is shown in Fig. S8 (ESI[†]) for comparison with the above measurements for $\text{BaTiO}_3\text{-RhB@SiO}_2$. It can be concluded that the sensitivity to temperature of the PL lifetime of RhB was reduced by adsorption in the mesopores of $\text{BaTiO}_3\text{-RhB@SiO}_2$, probably due to the restraint of the RhB molecule by the interaction with the wall of the mesopores.

Now, we demonstrate the temperature-measurement of the nanospace of the mesopores of the SiO_2 shell of $\text{BaTiO}_3\text{-RhB@SiO}_2$ under microwave irradiation by measuring the PL lifetime of RhB. When the solution of $\text{BaTiO}_3\text{-RhB@SiO}_2$ dispersed in heptane was irradiated by microwaves at 24 W, the bulk temperature of the solution measured by a fiber-optic thermometer was gradually increased from 23.0 to 40.0 °C during the 130 min irradiation as shown by the red line in Fig. 2c. The red dots indicate the temperature obtained by the PL lifetimes obtained by using the relationship shown in Fig. 2b against the microwave irradiation time. The raw data for these measurements of Fig. 2c are shown in Fig. S9 (ESI[†]). The first measurement gave a higher temperature (47.5 °C) than the bulk temperature (24.5 °C) after only 1 min of microwave irradiation, demonstrating the rapid increase of the nanospace temperature in the mesopores. The following measurements showed rather a large variation in the temperature values, but it should be mentioned that the temperature reached a stable state (50.4 °C), determined by taking the average of all plots considering the standard deviation of the temperatures (2.5 °C). This behaviour is explained by the generation of heat in $\text{BaTiO}_3\text{-RhB@SiO}_2$ balanced with the heat flow to the solution. The bulk

temperature also approached and reached the stable state (40.0 °C), demonstrating the temperature difference of 10.4 °C remained even in the stable state between the nanospace and the bulk, in which all the heat flows were balanced in the whole system. The same experiments when carried out at 12 and 16 W microwave irradiation gave the similar results, and the temperature differences were 6.6 °C and 7.5 °C, respectively. The temperature differences were plotted against the microwave powers (Fig. 2d). The linear relationship is shown between the temperature differences and microwave powers, displaying the non-equilibrium local heating in the nanospace of $\text{BaTiO}_3\text{-RhB@SiO}_2$. After the microwave irradiation, no RhB elution from $\text{BaTiO}_3\text{-RhB@SiO}_2$ to heptane was confirmed due to the absence of absorption peaks of RhB in the supernatant liquid (Fig. S10, ESI[†]), showing no changes of $\text{BaTiO}_3\text{-RhB@SiO}_2$ due to the microwave irradiation. We also conducted *in situ* temperature-measurement for a homogeneous RhB ethanol solution under microwave irradiation (Fig. S11, ESI[†]). The lifetimes under microwave irradiation showed no differences compared to the calibration curve, *i.e.*, the estimated temperature from the homogeneous RhB was the same as the bulk temperature measured by a fiber-optic thermometer.

The $\text{C}_{16}\text{N}^+\text{-W}_2\text{O}_7^{2-}\text{-Eu}^{3+}$ layered structure was prepared by a scheme shown in Fig. 3a. The particle size distribution (Fig. 3b) was determined from the SEM image (Fig. S12, ESI[†]), and then the average particle size was estimated to be 0.887 μm . The XRD peaks of $\text{C}_{16}\text{N}^+\text{-W}_2\text{O}_7^{2-}\text{-Eu}^{3+}$ and its precursors are shown in Fig. 3c. The d value of the 001 peak is attributed to the center-to-center distance of the tungstate layers. Ammonium cations were intercalated into the $\text{C}_{16}\text{N}^+\text{-W}_2\text{O}_7^{2-}$ interlayer, resulting in the distance of 4.49 nm. $\text{C}_{16}\text{N}^+\text{-W}_2\text{O}_7^{2-}\text{-Eu}^{3+}$ was obtained



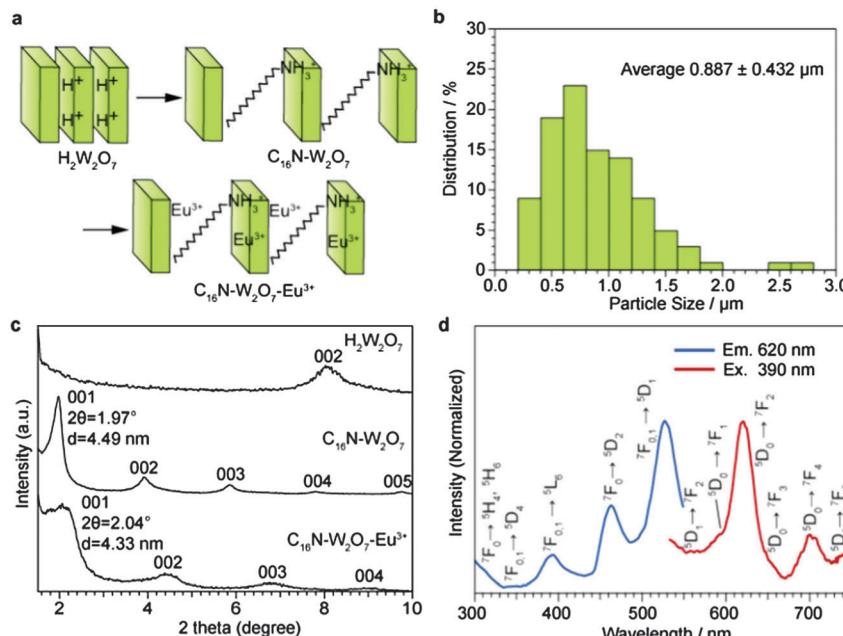


Fig. 3 Synthesis of $\text{C}_{16}\text{N}^+-\text{W}_2\text{O}_7^{2-}-\text{Eu}^{3+}$. (a) Synthesis scheme. (b) Particle size distribution. (c) XRD peaks. (d) Excitation spectrum and emission spectrum.

through the ion exchange of C_{16}N^+ with Eu^{3+} , reducing the center-to-center distance to 4.33 nm from the above value of $\text{C}_{16}\text{N}^+-\text{W}_2\text{O}_7^{2-}$. Since the thickness of a tungstate monolayer was estimated to be 0.96 nm in a previous report,³¹ the interlayer distance was determined to be 3.37 nm. Since the length of the alkyl chain of the C_{16}N^+ cation is estimated to be 2 nm and the ionic radius of Eu^{3+} cation is 1 nm, the interlayer space is wide enough to accommodate both C_{16}N^+ and Eu^{3+} . No change in the XRD peaks at angles higher than 20° (Fig. S13, ESI†) showed no structural change in each $\text{W}_2\text{O}_7^{2-}$ layer through the intercalation of ammonium cations and cation exchange with Eu^{3+} . The composition of $\text{C}_{16}\text{N}^+-\text{W}_2\text{O}_7^{2-}-\text{Eu}^{3+}$ was determined by the CHN analysis and ICP-OES (Table 1). The densities of C_{16}N^+ and Eu^{3+} were estimated to be 2.33 molecules nm^{-2} and 2.12 ions nm^{-2} , respectively. The excitation spectrum and the emission spectrum of $\text{C}_{16}\text{N}^+-\text{W}_2\text{O}_7^{2-}-\text{Eu}^{3+}$ were measured as shown in Fig. 4d.

Again, we obtained the relationship between the temperature and the PL lifetime for $\text{C}_{16}\text{N}^+-\text{W}_2\text{O}_7^{2-}-\text{Eu}^{3+}$. PL lifetime measurements were carried out using the PL decay at 620 nm of Eu^{3+} excited by 100 Hz pulse light at 390 nm from a xenon lamp through a 390 nm bandpass filter. The PL decays were fitted using only one exponential component, giving the PL lifetime. Fig. 4a shows the PL decays of $\text{C}_{16}\text{N}^+-\text{W}_2\text{O}_7^{2-}-\text{Eu}^{3+}$ at 22, 30, 40, 50, 60 and 70 °C under conventional heating. The PL lifetime value at each temperature was obtained as the average of the values repeated five times. The PL lifetimes at 22, 30, 40, 50, 60 and 70 °C are plotted against the temperature, giving a straight

line. This relationship was used for obtaining the temperature experienced by Eu^{3+} of $\text{C}_{16}\text{N}^+-\text{W}_2\text{O}_7^{2-}-\text{Eu}^{3+}$ under microwave irradiation. The relationship between the PL lifetime and the temperature of Eu^{3+} in 80% ethanol water is shown in Fig. S14 (ESI†) for comparing with the above measurements for $\text{C}_{16}\text{N}^+-\text{W}_2\text{O}_7^{2-}-\text{Eu}^{3+}$. It can be concluded that the sensitivity to temperature of the PL lifetime of Eu^{3+} was significantly increased by the introduction of Eu^{3+} into the $\text{C}_{16}\text{N}^+-\text{W}_2\text{O}_7^{2-}-\text{Eu}^{3+}$ interlayer, probably due to some interactions of the Eu^{3+} molecule with the $\text{W}_2\text{O}_7^{2-}$ layer.

Then, we demonstrate the temperature-measurement of the interlayer nanospace of $\text{C}_{16}\text{N}^+-\text{W}_2\text{O}_7^{2-}-\text{Eu}^{3+}$ under microwave irradiation by measuring the PL lifetime of Eu^{3+} . When the solution of $\text{C}_{16}\text{N}^+-\text{W}_2\text{O}_7^{2-}-\text{Eu}^{3+}$ dispersed in heptane was irradiated by microwaves at 24 W, the bulk temperature of the solution measured by a fiber-optic thermometer was gradually increased from 26.0 to 54.0 °C during the 70 min irradiation as shown by the red line in Fig. 4c. The red dots indicate the temperatures obtained at the PL lifetimes obtained using the relationship shown in Fig. 4b against the microwave irradiation time. The raw data of the above measurements in Fig. 4c are shown in Fig. S15 (ESI†). The second measurement gave a higher temperature (31.5 °C) than that for the solution (28.6 °C) after only 4 min of microwave irradiation, demonstrating the increase in the nanospace temperature of the interlayer. The following measurements showed a gradual increase in the interlayer temperature, showing rather a large dispersion in the temperature

Table 1 Chemical composition of $\text{C}_{16}\text{N}^+-\text{W}_2\text{O}_7^{2-}-\text{Eu}^{3+}$ from CHN analysis and ICP-OES

	C/wt%	H/wt%	N/wt%	O/wt%	W/wt%	Eu/wt%	C_{16}N^+ density/ nm^{-2}	Eu^{3+} density/ nm^{-2}
$\text{C}_{16}\text{N}^+-\text{W}_2\text{O}_7^{2-}-\text{Eu}^{3+}$	10.7	2.46	0.76	18.6	60	7.75	2.33	2.12

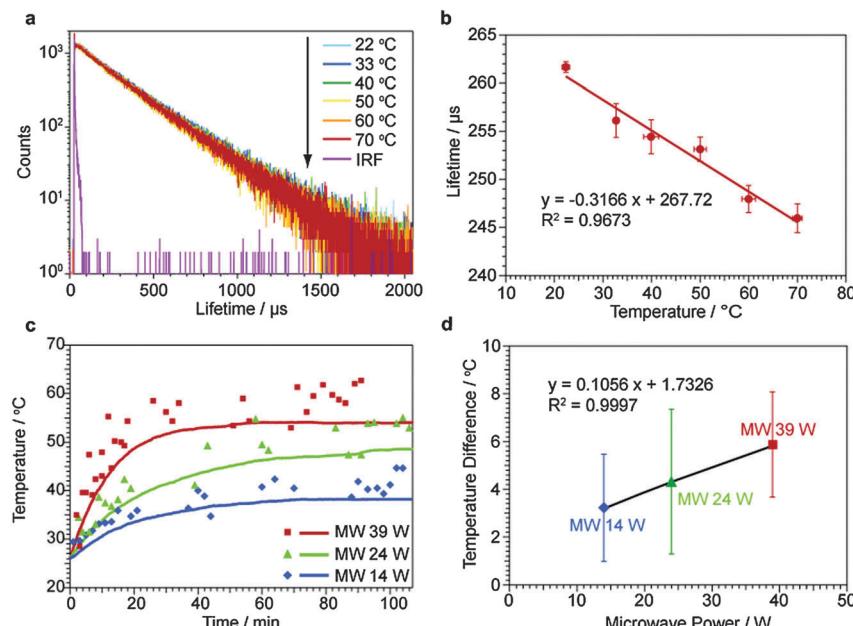


Fig. 4 *In situ* temperature measurements for $\text{C}_{16}\text{N}^+-\text{W}_2\text{O}_7^{2-}-\text{Eu}^{3+}$ under microwave irradiation. (a) PL decays at 22, 33, 40, 50, 60 and 70 °C under conventional heating, (b) calibration line between the lifetime and the temperature, and (c) nanospace temperatures of the $\text{W}_2\text{O}_7^{2-}$ interlayer plotted against the microwave irradiation time at 14, 24 and 39 W. Bulk temperatures shown as the curves. (d) Temperature differences between the nanospace and the bulk under microwave irradiation in the stable state plotted against microwave power.

values. The temperature difference of 3.7 °C was constant with the standard deviation of 2.8 °C in the whole range of the microwave irradiation time. We regarded the last eight plots under microwave irradiation at 24 W as the measurements in the stable state, and obtained the temperature difference as 4.3 °C with the standard deviation of 3.0 °C. The same experiments carried out at 14 and 39 W gave the similar results, and the temperature differences in the stable state were 3.2 °C and 5.8 °C, respectively. These temperature differences were plotted against the microwave powers (Fig. 4d). A linear relationship between the temperature difference and the microwave power was observed here, showing the non-equilibrium local heating in the interlayer nanospace of $\text{C}_{16}\text{N}^+-\text{W}_2\text{O}_7^{2-}-\text{Eu}^{3+}$.

Compared with $\text{BaTiO}_3-\text{RhB}@\text{SiO}_2$ under microwave irradiation at 24 W, the temperature differences observed for $\text{C}_{16}\text{N}^+-\text{W}_2\text{O}_7^{2-}-\text{Eu}^{3+}$ were not large. On the other hand, the increase of the bulk solution temperature under microwave irradiation was faster for $\text{C}_{16}\text{N}^+-\text{W}_2\text{O}_7^{2-}-\text{Eu}^{3+}$ than $\text{BaTiO}_3-\text{RhB}@\text{SiO}_2$. We attributed this result to the insufficient ability for insulating heat conduction in $\text{C}_{16}\text{N}^+-\text{W}_2\text{O}_7^{2-}-\text{Eu}^{3+}$ causing much heat flow to the solution easily. It can be concluded that the ability of heat insulation is critical for creating non-equilibrium heating, for promoting catalysis under microwave irradiation effectively. After the microwave irradiation, no Eu^{3+} elution from $\text{C}_{16}\text{N}^+-\text{W}_2\text{O}_7^{2-}-\text{Eu}^{3+}$ to heptane was confirmed due to the absence of absorption peaks of Eu^{3+} in the supernatant liquid (Fig. S16, ESI†). We also measured the SEM images (Fig. S17, ESI†) to determine the particle size distribution (Fig. S18, ESI†), XRD peaks (Fig. S19, ESI†), and chemical composition (Table S1, ESI†) of $\text{C}_{16}\text{N}^+-\text{W}_2\text{O}_7^{2-}-\text{Eu}^{3+}$ after the microwave irradiation. These results showed no changes in the $\text{C}_{16}\text{N}^+-\text{W}_2\text{O}_7^{2-}-\text{Eu}^{3+}$ structure

and the chemical composition by the microwave irradiation. Additionally, the lifetimes of $\text{C}_{16}\text{N}^+-\text{W}_2\text{O}_7^{2-}-\text{Eu}^{3+}$ before and after the microwave irradiation were the same as the calibration line (Fig. S20, ESI†), showing that the lifetimes were not changed by the microwave irradiated time. We also conducted the *in situ* temperature-measurement for a homogeneous mixture of Eu^{3+} 80% in ethanol water (Fig. S21, ESI†). The lifetimes under microwave irradiation showed no differences from the calibration line, in which the estimated temperature from the homogeneous Eu^{3+} was the same as the bulk temperature measured using a fiber-optic thermometer.

Conclusions

In summary, we demonstrate the *in situ* temperature-measurement of local temperatures under microwave irradiation. This method has been applied to two systems: $\text{BaTiO}_3-\text{RhB}@\text{SiO}_2$ and $\text{C}_{16}\text{N}^+-\text{W}_2\text{O}_7^{2-}-\text{Eu}^{3+}$ dispersed in heptane, respectively, in which PLs of RhB and Eu^{3+} are sensitive to temperature, and are used as a molecular thermometer. We have succeeded in determining the temperature difference occurring in nano-hybrid systems under microwave irradiation. We conclude that microwaves heat particles locally, leading to “non-equilibrium local heating”. The heating extents are significantly varied with different particles, which can be attributed to the ability as the heat insulator contained in the nano-hybrid systems. The temperature-measuring method is very important in controlling microwave heating in chemical reactions. Furthermore it should be a strong tool for elucidating the mechanism of microwave special effects.



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