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# The salts of chloronium ions $R\text{-Cl}^+\text{-R}$ ( $R = \text{CH}_3$ or $\text{CH}_2\text{Cl}$ ): formation, thermal stability, and interaction with chloromethanes†

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The interaction of  $\text{CH}_3\text{Cl}/\text{CD}_3\text{Cl}$  or  $\text{CH}_2\text{Cl}_2/\text{CD}_2\text{Cl}_2$  with the carborane acid  $\text{H}(\text{CHB}_{11}\text{Cl}_{11})$  (abbreviated as  $\text{H}\{\text{Cl}_{11}\}$ ) generates the salts of  $\text{CH}_3\text{-}\{\text{Cl}_{11}\}$  and  $\text{CH}_2\text{Cl-}\{\text{Cl}_{11}\}$  and their deuterio analogs, respectively, which are analogs of the salts of asymmetric chloronium cations. Next, salts of chloronium cations  $\text{CH}_3\text{-Cl}^+\text{-CH}_3$ ,  $\text{ClCH}_2\text{-Cl}^+\text{-CH}_2\text{Cl}$ , and  $\text{ClCH}_2\text{-Cl}^+\text{-CH}_3$  and their deuterio analogs were obtained from the above compounds. The asymmetric  $\text{ClCH}_2\text{-Cl}^+\text{-CH}_3$  cation was found to be unstable, and at ambient temperature, slowly disproportionated into symmetric cations  $(\text{CH}_3)_2\text{Cl}^+$  and  $(\text{CH}_2\text{Cl})_2\text{Cl}^+$ . At a high temperature (150 °C), disproportionation was completed within 5 minutes, and the resulting cations further decomposed into  $\text{CH}_3\text{-}\{\text{Cl}_{11}\}$  and  $\text{CH}_2\text{Cl-}\{\text{Cl}_{11}\}$ . The molecular fragment  $\text{ClCH}_2\text{-}(X)$  of the compounds ( $X = \{\text{Cl}_{11}\}$ ,  $\text{-Cl}^+\text{-CH}_2\text{Cl}$ , or  $\text{-Cl}^+\text{-CH}_3$ ) is involved in exchange reactions with  $\text{CH}_2\text{Cl}_2$  and  $\text{CHCl}_3$ , converting into  $\text{CH}_3\text{-}(X)$  with the formation of chloroform and  $\text{CCl}_4$ , respectively.

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Halonium ions ( $\text{R}_2\text{Hal}^+$ ) are well-recognized reactive intermediates in electrophilic chemistry.<sup>1,2</sup> Their stability increases in the order  $\text{Hal} = \text{F}, \text{Cl}, \text{Br}, \text{I}$ . Recently, the evidence of the formation of the symmetrical fluoronium ions in solutions was obtained,<sup>3,4</sup> and the nature of carbon-halogen bonds in the halonium ions was studied.<sup>5</sup> Mostly stable dimethylbromonium and -iodonium salts are presently commercialized and widely used in chemical ionization mass spectroscopy (gas phase chemistry) as effective methylating<sup>6–12</sup> and protonating agents<sup>13,14</sup> for a variety of compounds. Nevertheless, the chemistry of dialkylhalonium ions in condensed phases is virtually unknown. Recently, the salts of  $(\text{CH}_3)_2\text{Cl}^+$  and  $(\text{C}_2\text{H}_5)_2\text{Cl}^+$  with the exceptionally stable and inert toward reactive cations undecachlorocarborane ions,  $\text{CHB}_{11}\text{Cl}_{11}^-$ , were obtained and studied using X-ray and infrared (IR) spectroscopy.<sup>15</sup> The solid salt  $(\text{CH}_3)_2\text{Cl}^+(\text{CHB}_{11}\text{Cl}_{11}^-)$  is stable even at elevated temperatures and decomposes at 140 °C releasing  $\text{CH}_3(\text{CHB}_{11}\text{Cl}_{11})$  and  $\text{CH}_3\text{Cl}$ . This is an important method for the isolation of pure  $\text{CH}_3(\text{CHB}_{11}\text{Cl}_{11})$ , which can be viewed as a neutral analogue of an asymmetric chloronium ion related to the dimethylchloronium ylide,  $\text{CH}_3\text{ClCH}_2$ .<sup>16,17</sup>

In the present work we obtained the salts of symmetric and asymmetric chloronium cations,  $\text{CH}_3\text{-Cl}^+\text{-CH}_2\text{Cl}$  and

$\text{CH}_2\text{Cl-Cl}^+\text{-CH}_2\text{Cl}$ , both protio and deuterio analogs, with the  $\text{CHB}_{11}\text{Cl}_{11}^-$  counterion (hereafter abbreviated as  $\{\text{Cl}_{11}^-\}$ , Fig. S1, ESI†). We also explored their thermal stability and interaction with some simple chloromethanes. The carborane ion  $\{\text{Cl}_{11}^-\}$  was chosen as a counterion for chloronium salts because of its exceptionally low basicity and high thermal stability, which ensure the stability of chloronium salts at room temperature and above.<sup>15</sup>

## Experimental

Carborane acid  $\text{H}\{\text{Cl}_{11}\}$  was prepared as previously described.<sup>18</sup> The acid was sublimed at 150–160 °C under a pressure of  $10^{-5}$  Torr on cold Si windows of a specially designed IR cell-reactor as a very thin translucent layer.<sup>19</sup> The spectrum of the sublimed acid showed no traces of the  $\text{H}_3\text{O}^+$  cation.<sup>20</sup> Dry gaseous chloromethanes ( $\text{CH}_3\text{Cl}$ ,  $\text{CH}_2\text{Cl}_2$ , and  $\text{CHCl}_3$ ) were injected anaerobically into the IR cell. The values of their partial pressure were calculated as the ratio of absorption intensity to that of the standard spectrum recorded in the same cell filled with vapors at atmospheric pressure.

All procedures were performed in a Vacuum Atmospheres Corp. glovebox in the atmosphere of  $\text{N}_2$  ( $\text{O}_2$  and  $\text{H}_2\text{O} < 0.5$  ppm). The IR spectra were recorded on a PerkinElmer Spectrum-100 spectrometer inside a dry box in the transmission mode ( $400\text{--}4000\text{ cm}^{-1}$ ). The IR data were processed in the GRAMMS/A1 (7.00) software program from Thermo Fisher Scientific.

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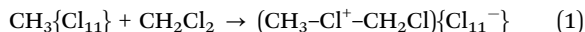


## Results and discussion

To explore the chemical processes involving chloronium ions by means of IR spectroscopy, their detailed IR spectra must be obtained and interpreted.

### Chloronium salts and IR spectra

The  $\text{CH}_3\text{-Cl}^+\text{-CH}_2\text{Cl}$  cation was obtained by introducing the  $\text{CH}_2\text{Cl}_2$  vapors at a partial pressure ( $P_{\text{CH}_2\text{Cl}_2}$ ) of 0.4 atm into an evacuated IR cell containing a film of  $\text{CH}_3\text{-}\{\text{Cl}_{11}\}$  sublimed on their Si-windows. Reaction (1) of  $\text{CH}_3\text{-Cl}^+\text{-CH}_2\text{Cl}$  formation proceeds very slowly, without a release of HCl.



After 40 minutes, the spectrum of the cation was isolated by subtracting the original spectrum of  $\text{CH}_3\text{-}\{\text{Cl}_{11}\}$  multiplied by the scaling factor  $f = 0.933$  (Fig. 1, red). This means that only 6.7% of the  $\text{CH}_3\text{-}\{\text{Cl}_{11}\}$  salt (100–93.3%) has reacted.

If the partial pressure of the injected  $\text{CH}_2\text{Cl}_2$  vapors was twofold higher (1 atm), then the reaction was accelerated significantly and completed within 6 minutes. The spectrum of the resultant  $\text{CH}_3\text{-Cl}^+\text{-CH}_2\text{Cl}$  cation slightly differs from that of the previous sample (Fig. 1), indicating that the cation is sensitive to changes in the environment.

The  $\text{CH}_3\text{-Cl}^+\text{-CD}_2\text{Cl}$  cation is formed when the vapors of  $\text{CD}_2\text{Cl}_2$  are introduced into the IR cell with sublimed  $\text{CH}_3\text{-}\{\text{Cl}_{11}\}$ .

The IR spectrum of the salt of this cation is shown in Fig. 2 (red; the gas phase was evacuated).

The  $\text{CD}_3\text{-Cl}^+\text{-CH}_2\text{Cl}$  and  $\text{CD}_3\text{-Cl}^+\text{-CD}_2\text{Cl}$  cations were formed when vapors of  $\text{CH}_2\text{Cl}_2$  or  $\text{CD}_2\text{Cl}_2$  respectively, were injected into the IR cell with a film of sublimed  $\text{CD}_3\text{-}\{\text{Cl}_{11}\}$  salt on the Si windows. Their IR spectra are presented in Fig. S2 and S3 (ESI<sup>†</sup>) and in Table 1.

### Interpretation of the IR spectra

IR spectra of the salts of the cations  $\text{CH}_3\text{-Cl}^+\text{-CH}_2\text{Cl}$  (red),  $\text{CD}_3\text{-Cl}^+\text{-CH}_2\text{Cl}$  (blue), and  $\text{CH}_3\text{-Cl}^+\text{-CD}_2\text{Cl}$  (green) that are normalized to the intensity of the anion are shown in Fig. 3. The figure shows that in the frequency range of CH stretch vibrations, the sum of the spectra of the fragments  $\text{ClH}_2\text{C-(Cl}^+\text{-)}$  (blue) and  $\text{CH}_3\text{-(Cl}^+\text{-)}$  (green) matches the spectrum of the  $\text{CH}_3\text{-Cl}^+\text{-CH}_2\text{Cl}$  cation. The same result was observed in the frequency range of the CH bend vibrations (Fig. 3, right). Interpretations of the IR spectra for all cations follow from the above data (Tables 1 and 2), taking into account that spectra of  $\text{CH}_3\{\text{Cl}_{11}\}$  and  $(\text{CH}_3\text{-Cl}^+\text{-CH}_3)\{\text{Cl}_{11}^-\}$  were interpreted earlier.<sup>15</sup> It was important to determine the existence of specific absorption bands for each compound under study; we will use these bands (shown in Tables 1 and 2) as markers for the identification of these compounds in the mixtures. Intensity of the marked bands allowed us to estimate relative amounts of the compounds.

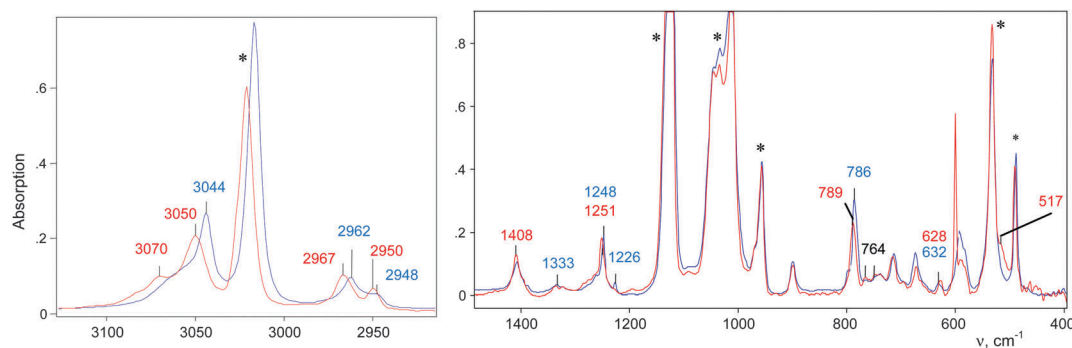


Fig. 1 Red: IR spectra of  $(\text{CH}_3\text{-Cl}^+\text{-CH}_2\text{Cl})\{\text{Cl}_{11}^-\}$  formed at low (red) and high (blue) partial pressure of  $\text{CH}_2\text{Cl}_2$  (0.4 and 1 atm respectively). Both spectra are normalized to intensity of the anion. The red spectrum was isolated after subtracting the spectrum of the unreacted  $\text{CH}_3\text{-}\{\text{Cl}_{11}\}$ . Intense bands (marked by an asterisk) belong to the  $\{\text{Cl}_{11}^-\}$  anion.

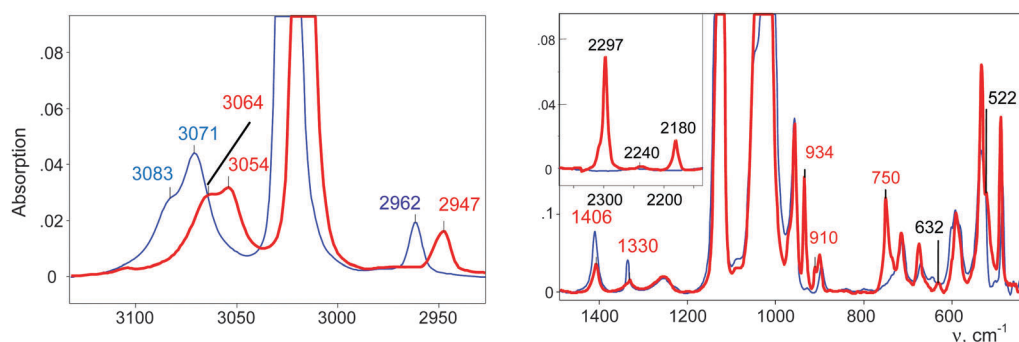
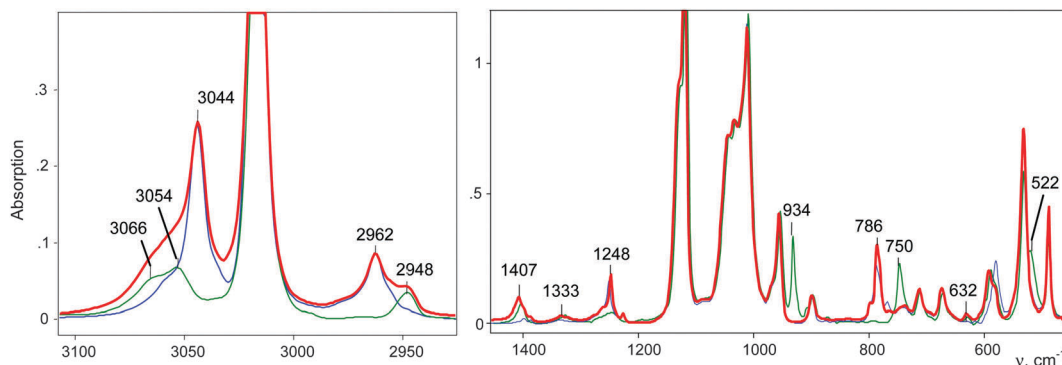


Fig. 2 IR spectra of the initial  $\text{CH}_3\text{-}\{\text{Cl}_{11}\}$  salt (blue) and the salt of the  $\text{CH}_3\text{-Cl}^+\text{-CD}_2\text{Cl}$  cation (red).



**Table 1** IR frequencies of the methyl and bridged C–Cl<sup>+</sup>–C groups of the compounds analyzed, in comparison with the IR spectrum of CH<sub>3</sub>Cl (the most characteristic or intense bands, used as markers to identify the cations in their mixtures, are underlined)

Compound	$\nu_{\text{as}}\text{CH}_3$	$\nu_{\text{s}}\text{CH}_3$	$\delta_{\text{as}}\text{CH}_3$	$\delta_{\text{s}}\text{CH}_3$	CH <sub>3</sub> rock	$\nu_{\text{as}}(\text{CClC})$	$\nu_{\text{s}}(\text{CClC})$
CH <sub>3</sub> Cl (gas) <sup>21</sup>	3043	<u>3039</u>	2968	2879	1452	1355	732 ( $\nu_{\text{C}}\text{Cl}$ )
CH <sub>3</sub> {Cl <sub>11</sub> } <sup>15</sup>	3083	<u>3071</u>	2962	1409	1335	—	—
CH <sub>3</sub> –Cl <sup>+</sup> –CH <sub>3</sub> <sup>15</sup>		<u>3068</u>	2957	1417	1324	<u>636</u>	
H <sub>3</sub> C–Cl <sup>+</sup> –CD <sub>2</sub> Cl	3064	<u>3054</u>	2947	1406	1330	934	632

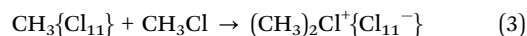
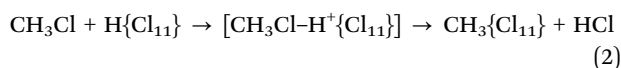
**Fig. 3** IR spectra of salts of cations: CH<sub>3</sub>–Cl<sup>+</sup>–CH<sub>2</sub>Cl (red), CD<sub>3</sub>–Cl<sup>+</sup>–CH<sub>2</sub>Cl (blue) and CH<sub>3</sub>–Cl<sup>+</sup>–CD<sub>2</sub>Cl (green).**Table 2** IR frequencies of the methylene chloride and bridged C–Cl<sup>+</sup>–C groups of the compounds analyzed, in comparison with the IR spectrum of CH<sub>2</sub>Cl<sub>2</sub> (the most characteristic or intense bands, used as markers to identify the cations in their mixtures, are underlined)

Compound	$\nu_{\text{as}}\text{CH}_2$	$\nu_{\text{s}}\text{CH}_2$	$\delta\text{CH}_2$ scissor	$\delta\text{CH}_2$ waggle	$\nu_{\text{as}}\text{CCl}_2/\nu_{\text{C}}\text{–Cl}$	$\nu_{\text{as}}(\text{CClC})$					
CH <sub>2</sub> Cl <sub>2</sub> (liquid) <sup>22</sup>	3045	2990	1424	1265vs	739	—					
ClH <sub>2</sub> C–{Cl <sub>11</sub> }	3079	<u>3063</u>	2978	1391	<u>1243vs</u>	793					
ClH <sub>2</sub> C–Cl <sup>+</sup> –CH <sub>2</sub> Cl <b>IIa</b>	3073	<u>3064</u>	3048	2980	1416	1324	<u>1284vs</u>	1248	798	<u>783</u>	635
ClH <sub>2</sub> C–Cl <sup>+</sup> –CH <sub>2</sub> Cl <b>IIb</b>		<u>3044</u>	2964br	1406	1338	1283vs	1248	798	<u>783</u>	672	593
ClH <sub>2</sub> C–Cl <sup>+</sup> –CH <sub>3</sub>		3044	2962	<u>1407</u>	1386	1261	<u>1248s</u>	1226	<u>786</u>	<sup>a</sup>	
ClH <sub>2</sub> C–Cl <sup>+</sup> –CD <sub>3</sub>		3044	2962	<u>1398</u>	1387	1263	<u>1251s</u>	1226	<u>787</u>	<sup>a</sup>	

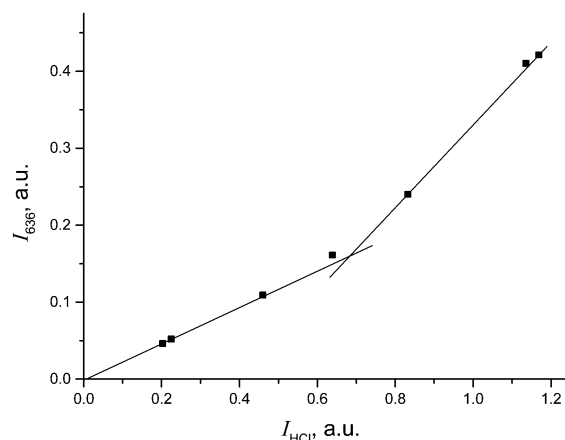
<sup>a</sup> Not determined.

### Interaction of mono- and dichloromethane with H{Cl<sub>11</sub>} and chloronium cations

CH<sub>3</sub>Cl interacts with the H{Cl<sub>11</sub>} acid in two stages. At first, it is protonated with the release of HCl and the formation of CH<sub>3</sub>{Cl<sub>11</sub>} (eqn (2)):

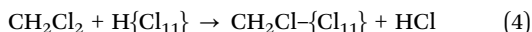


Then, CH<sub>3</sub>Cl binds to CH<sub>3</sub>{Cl<sub>11</sub>}, thus forming (CH<sub>3</sub>)<sub>2</sub>Cl<sup>+</sup>{Cl<sub>11</sub><sup>–</sup>} (eqn (3)). The formation of the resultant CH<sub>3</sub>{Cl<sub>11</sub>} can be observed by monitoring the intensity of the IR absorption of the released HCl ( $I_{\text{HCl}}$ ). The amount of the resulting ion (CH<sub>3</sub>)<sub>2</sub>Cl<sup>+</sup> can be traced from the intensity of its band  $\nu_{\text{as}}(\text{CCl}^+\text{C})$  at 636 cm<sup>–1</sup> ( $I_{636}$ ; Fig. S4, ESI<sup>†</sup>). Dependence of  $I_{636}$  on  $I_{\text{HCl}}$  was initially proportional (Fig. 4), which means that both reactions proceed at constant speeds. When the thickness of the layer of the formed surface products increases sufficiently to slow down the diffusion of CH<sub>3</sub>Cl molecules to the acid, the rate of

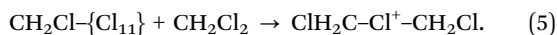
**Fig. 4** Dependence of the intensity of band  $\nu_{\text{as}}(\text{CClC})$  at 636 cm<sup>–1</sup> of the (CH<sub>3</sub>)<sub>2</sub>Cl<sup>+</sup> cation on the intensity of HCl absorption [reflects the dependence of the (CH<sub>3</sub>)<sub>2</sub>Cl<sup>+</sup> formation on the CH<sub>3</sub>{Cl<sub>11</sub>} formation].

$\text{CH}_3^+\{\text{Cl}_{11}^-\}$  formation decreases and the slope increases (Fig. 4). The cation of the ultimate salt  $(\text{CH}_3)_2\text{Cl}^+\{\text{Cl}_{11}^-\}$  does not show any further interaction with  $\text{CH}_3\text{Cl}$ .

Dichloromethane vapors (at a partial pressure of 0.65 atm) interacted with sublimed acid in the same way as chloromethane did. At the first stage, IR spectra registered the emergence of absorbance of the gaseous HCl and the surface  $\text{CH}_2\text{Cl}-\{\text{Cl}_{11}\}$  compound, which are formed in accordance with eqn (4):



The IR spectrum of  $\text{CH}_2\text{Cl}-\{\text{Cl}_{11}\}$  showed the characteristic band of the C–Cl stretch at  $793\text{ cm}^{-1}$ , and the absence of the bands of C–Cl<sup>+</sup>–C group vibrations in the frequency range  $650\text{--}500\text{ cm}^{-1}$  (Fig. 5 and Table 2). Intensity of HCl absorption ( $I_{\text{HCl}}$ ) was used to quantify the total amount of the formed  $\text{CH}_2\text{Cl}-\{\text{Cl}_{11}\}$ , whereas intensity of the band of the terminal C–Cl stretch ( $I_{\text{CCl}}$ ) reflects the current amount of this compound. Dependence of  $I_{\text{CCl}}$  on  $I_{\text{HCl}}$  was linear for the first 5 hours of the reaction (Fig. 6); this result means that only  $\text{CH}_2\text{Cl}-\{\text{Cl}_{11}\}$  was formed. Then, the dependence started to drop (point 22 in Fig. 6), indicating the second stage of the reaction: involvement of  $\text{CH}_2\text{Cl}-\{\text{Cl}_{11}\}$  in the formation of the  $\text{ClH}_2\text{C}-\text{Cl}^+-\text{CH}_2\text{Cl}$  cation (eqn (5)).



Extrapolation of  $I_{\text{CCl}}$  from  $I_{\text{HCl}}$  to the moment of the reaction stoppage (at the 1380th minute, when  $\text{CH}_2\text{Cl}_2$  was pumped out) yields  $I_{\text{CCl}} = 0.084$  (arbitrary units; Fig. 6). This value corresponds to the amount of  $\text{CH}_2\text{Cl}-\{\text{Cl}_{11}\}$  that was formed in accordance with the amount of HCl production (eqn (4)). Nevertheless, because  $\text{CH}_2\text{Cl}-\{\text{Cl}_{11}\}$  is further consumed (reaction (5)),  $I_{\text{CCl}}$  decreases to 0.040 arbitrary units (Fig. 6), meaning that 48% ( $0.040/0.084$ ) of this compound survived, and 52% was converted to the chloronium ion.

An IR spectrum of  $(\text{CH}_2\text{Cl})_2\text{Cl}^+$  was obtained by subtracting the spectrum of  $\text{CH}_2\text{Cl}-\{\text{Cl}_{11}\}$  from the spectrum of the mixture of  $(\text{CH}_2\text{Cl})_2\text{Cl}^+$  with the  $\text{CH}_2\text{Cl}-\{\text{Cl}_{11}\}$  up to complete compensation

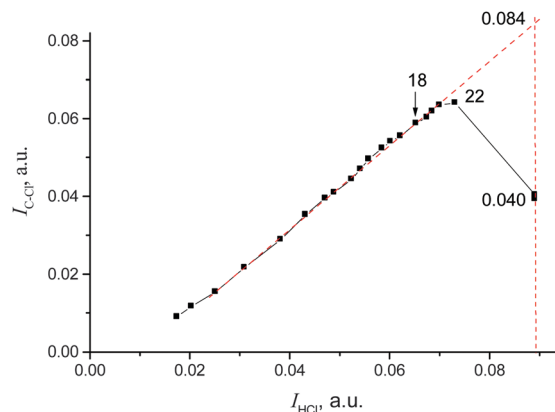


Fig. 6 Dependence of  $I_{\text{CCl}}$  on  $I_{\text{HCl}}$ , pointing to  $\text{ClH}_2\text{C}-\{\text{Cl}_{11}\}$  formation and its further consumption on the formation of the  $\text{ClH}_2\text{C}-\text{Cl}^+-\text{CH}_2\text{Cl}$  cation.

of the bands  $\nu\text{CCl} = 792\text{ cm}^{-1}$  and  $\delta\text{CH}_2 = 1243\text{ cm}^{-1}$ , which are specific to  $\text{CH}_2\text{Cl}-\{\text{Cl}_{11}\}$  (Fig. 5, blue; Table 2). The frequencies of the  $(\text{CH}_2\text{Cl})_2\text{Cl}^+$  cation are very close to those of  $\text{CH}_2\text{Cl}-\{\text{Cl}_{11}\}$  except for one intense band at  $1284\text{ cm}^{-1}$ , which can be used as a marker of this cation (Table 2). Furthermore, we will denote  $\text{CH}_2\text{Cl}-\{\text{Cl}_{11}\}$  as compound **I** and the  $(\text{CH}_2\text{Cl})_2\text{Cl}^+$  salt as compound **II**.

After stoppage of the reaction at minute 1380, an IR spectrum of the sample retained a strong absorption pattern of the unreacted acid, which constituted 70% of the original spectrum of the acid.

To continue reactions (4) and (5) at a higher speed, the  $\text{CH}_2\text{Cl}_2$  vapors were reintroduced into our IR cell at higher partial pressure (1 atm). Under these conditions, the formation of compound **I** was completed after 8 minutes with full utilization of the acid and termination of the HCl release (point 30 in Fig. 7). An IR spectrum of this sample is shown in Fig. 8 (black).

Further interaction of compound **I** with  $\text{CH}_2\text{Cl}_2$  and the formation of **II** is manifested in IR spectra as a decrease in the intensity of compound **I** and the upregulation of compound **II** with the appearance of isosbestic points (Fig. 8 and Fig. S5, ESI<sup>†</sup>).

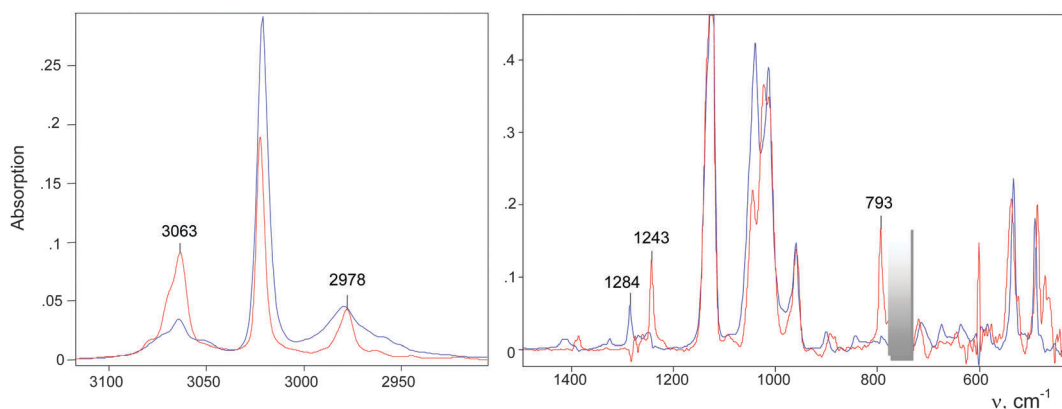


Fig. 5 IR spectra of the  $\text{CH}_2\text{Cl}-\{\text{Cl}_{11}\}$  compound (red) and the salt of  $\text{CH}_2\text{Cl}-\text{Cl}^+-\text{CH}_2\text{Cl}$  cation (blue). The spectra of unreacted  $\text{H}\{\text{Cl}_{11}\}$  acid, gaseous HCl and  $\text{CH}_2\text{Cl}_2$  were subtracted. The region of strong absorption of gaseous  $\text{CH}_2\text{Cl}_2$  (C–Cl stretches), where compensation does not work, is shaded.



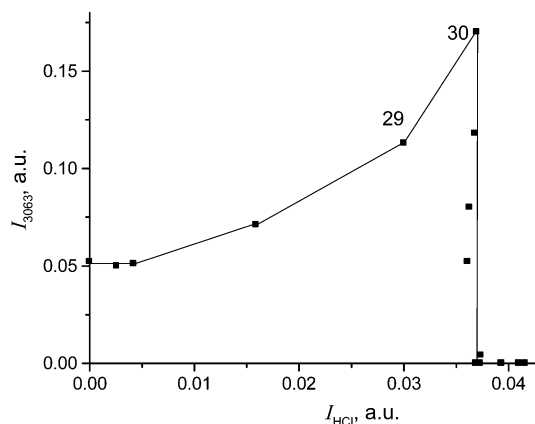


Fig. 7 Dependence of  $\text{CH}_2\text{Cl}-(\text{Cl}_{11})$  formation (determined by means of intensity of its band  $\nu_{\text{as}}\text{CH}_2$  at  $3063\text{ cm}^{-1}$ ) on the amount of released HCl (determined by means of intensity of the  $\nu\text{HCl}$  band).

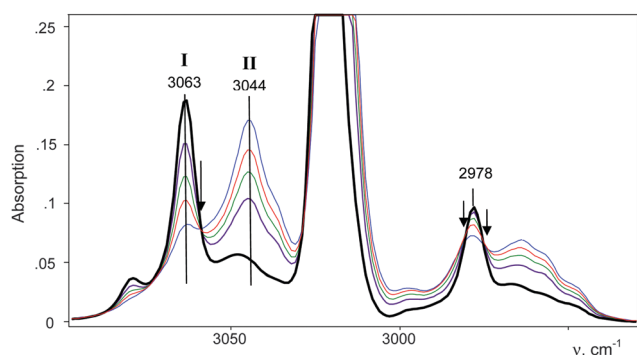


Fig. 8 A change in intensity of the IR spectra of compounds  $\text{CH}_2\text{Cl}-(\text{Cl}_{11})$  and  $(\text{CH}_2\text{Cl}-\text{Cl}^+-\text{CH}_2\text{Cl})(\text{Cl}_{11}^-)$  as reaction (5) proceeds. The arrows indicate isosbestic points. The spectrum of gaseous  $\text{CH}_2\text{Cl}_2$  was subtracted.

The spectrum of the resultant cation **II** differed from that of the same cation formed during a slow reaction of  $\text{CH}_2\text{Cl}-(\text{Cl}_{11})$  with  $\text{CH}_2\text{Cl}_2$  (Fig. 9 and Table 2). Thus, two isomers of  $\text{CH}_2\text{Cl}-\text{Cl}^+-\text{CH}_2\text{Cl}$

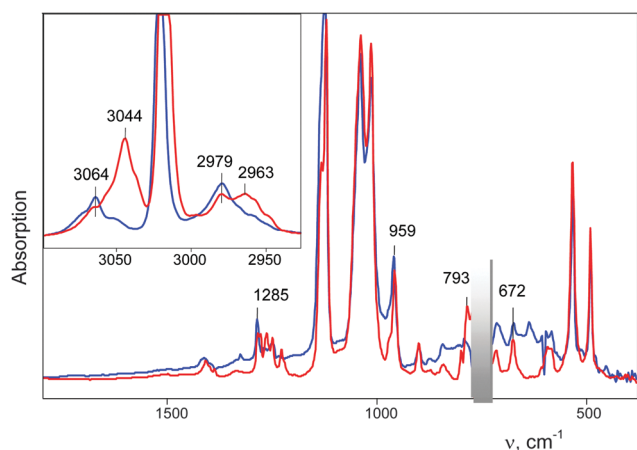


Fig. 9 IR spectra of  $(\text{CH}_2\text{Cl})_2\text{Cl}^+(\text{Cl}_{11})$ , formed by slow (**IIa**, some hours, blue) and quick interaction of  $\text{CH}_2\text{Cl}-(\text{Cl}_{11})$  with  $\text{CH}_2\text{Cl}_2$  (**IIb**, some minutes, red). The region of strong absorption of gaseous  $\text{CH}_2\text{Cl}_2$  (C–Cl stretches), where compensation does not work, is shaded.

are formed: compound **IIa** from the slowly proceeding reaction (5) and compound **IIb** from the rapid reaction (5). They differ in frequencies of stretches CH and CCl. Nonetheless, their bend CH vibrations are very similar and have one specific band  $\delta\text{CH}_2$  ( $1284\text{ cm}^{-1}$ ), which does not overlap with the bands of other types of cations. Therefore, the intensity of this band ( $I_{1284}$ ) can be used for the estimation of the amount of the **IIa** + **IIb** mixture.

Fig. 10 shows the time dependence of the intensity of absorption bands of compounds **I** ( $I_{3063}$ ) and **II** ( $I_{1284}$ ) that reflects the formation of the corresponding cations. One can see that the formation of compound **II** passes through a maximum at point 34, when compound **I** disappears, and then decreases.

Simultaneously, the bands indicative of the cation  $\text{CH}_3-\text{Cl}^+-\text{CH}_2\text{Cl}$  at  $1261$  and  $1248\text{ cm}^{-1}$  made an appearance (Table 2) and increased in intensity. We will denote this cation as compound **III**. Its narrow band  $\delta\text{CH}_2 = 1248\text{ cm}^{-1}$  ( $I_{1248}$ ) does not overlap with the bands of other cations and can be used for the assessment of the relative amount of compound **III** (Fig. 10, green). Along with the spectrum of compound **III**, the absorption bands of the cation  $(\text{CH}_3)_2\text{Cl}^+$  also develop (hereafter denoted as compound **IV**). Intensity of its single band at  $1324\text{ cm}^{-1}$  ( $I_{1324}$ ) was used to assess the formation of compound **IV**. Fig. 10 summarizes the sequence of the formation of compounds **I**, **II**, **III**, and **IV** and the relative amounts of each compound, but does not describe the quantitative relations among them.

Cations **III** and **IV** can be formed only if the reaction of **I** or **II** with  $\text{CH}_2\text{Cl}_2$  is accompanied by the formation of gaseous products. The spectra of the gas phase revealed the band of H–C–Cl bend vibration of chloroform at  $1219\text{ cm}^{-1}$ . The time dependence of its intensity shows that the formation of chloroform is associated with the formation of compounds **I** and **II** (Fig. 10).

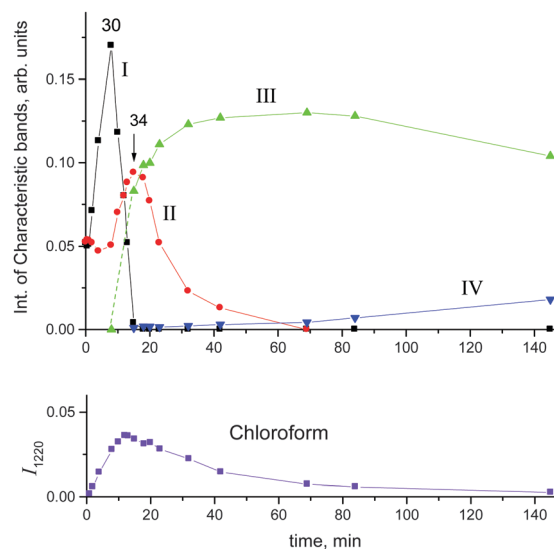


Fig. 10 Kinetic curves of the formation of compounds **I–IV** and chloroform; these curves were constructed by means of the intensity values of their bands at  $3068$  (compound **I**),  $1284$  (**II**),  $1248$  (**III**),  $1322$  (**IV**), and  $1220\text{ cm}^{-1}$  (chloroform). Curves **I–IV** do not quantitatively indicate the proportion of a cation in the mixture of compounds **I–IV**.





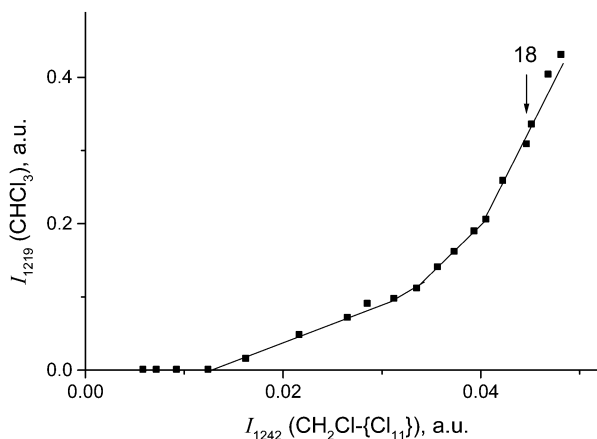
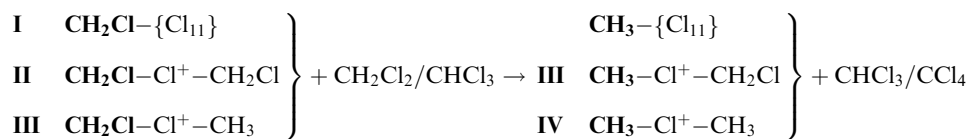
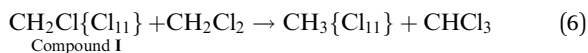


Fig. 11 Dependence of  $\text{CHCl}_3$  formation (shown as intensity of the band at  $1219\text{ cm}^{-1}$ ) on  $\text{CH}_2\text{Cl}-\{\text{Cl}_{11}\}$  formation (indicated by intensity of its band at  $1242\text{ cm}^{-1}$ ) under the conditions when only compound **I** is formed. Cation **II** starts to form after point 18.

In the situation when only **I** is formed ( $P_{\text{CH}_2\text{Cl}_2} = 0.65\text{ atm}$ ), the dependence of the chloroform formation ( $I_{1219}$ ) on the formation of **I** ( $I_{1242}$ ) increases both with an increase in the amount of compound **I** and with an increase in the contact time of **I** with dichloromethane (Fig. 11). Subtraction of the spectrum of compound **I** from the spectrum of the products (at point 18 in Fig. 11) leads to the manifestation of the spectrum of compound  $\text{CH}_3\{\text{Cl}_{11}\}$  (Fig. S6, ESI<sup>†</sup>). Hence,  $\text{CH}_2\text{Cl}_2$  reacts with  $\text{CH}_2\text{Cl}\{\text{Cl}_{11}\}$  according to eqn (6):



With the rapid formation of  $\text{CH}_2\text{Cl}\{\text{Cl}_{11}\}$  in reaction (4) at  $P_{\text{CH}_2\text{Cl}_2} = 1\text{ atm}$ , chloroform is formed symbatically (Fig. 12),

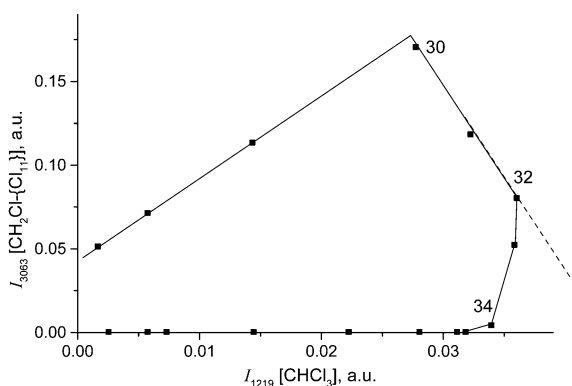
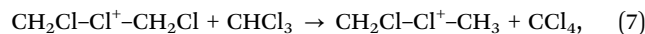


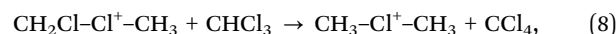
Fig. 12 The link between formation of  $\text{CH}_2\text{Cl}-\{\text{Cl}_{11}\}$  (indicated by intensity of its band at  $3063\text{ cm}^{-1}$ ) and  $\text{CHCl}_3$  (indicated by intensity of its band at  $1219\text{ cm}^{-1}$ ).

confirming reaction (6). In contrast, starting from point 32, chloroform formation stopped, whereas from point 34, it begins to decrease, when the amount of **II** passes through a maximum, and compound **I** is exhausted (Fig. 10). The subsequent expenditure of chloroform occurs simultaneously with the consumption of compound **II** and the increasing amount of **III**. This finding points to reaction (7):



(anions are omitted), which is suggestive of  $\text{CCl}_4$  formation.

In the same way, compound **IV** can form:



The detection of  $\text{CCl}_4$  was carried out as follows. After the completion of the reaction, the gaseous phase and all surface-adsorbed molecules were removed by pumping. The difference in IR spectra before and after the evacuation represented the spectrum of removed molecules. It consists of a strong absorption pattern of the original dichloromethane, the characteristic band at  $1219\text{ cm}^{-1}$  of chloroform, and a weak band at  $790\text{ cm}^{-1}$ , which may belong to the C-Cl stretch of  $\text{CCl}_4$ . The latter frequency is lower than that of gaseous  $\text{CCl}_4$  ( $795\text{ cm}^{-1}$ ), but equals that of  $\text{CCl}_4$  solvated with dichloromethane in its solutions ( $789\text{ cm}^{-1}$ ). Therefore, the traces of the formed  $\text{CCl}_4$  are adsorbed by superficial chloronium salts.

To sum up, we can conclude that  $\text{CH}_2\text{Cl}_2$  and  $\text{CHCl}_3$  interact with the molecular fragment  $\text{CH}_2\text{Cl}-$ , whose reactivity is highest in  $\text{CH}_2\text{Cl}\{\text{Cl}_{11}\}$  and is consistently reduced in cations **II** and **III**. In general, these interactions can be expressed as:

Decreasing the reactivity of the molecular fragment  $\text{CH}_2\text{Cl}-$  in compounds **I**, **II**, and **III** is correlated with its decreasing polarizability, which in turn is determined by the change in ionicity of the bonds in chloronium bridge  $\text{C}-\text{Cl}^+-\text{C}$ .

### Ionicity/covalency of the bonds in the $\text{C}-\text{Cl}^+-\text{C}$ bridge of chloronium ions

The stretch frequencies of the methyl group, especially  $\nu_{\text{as}}\text{CH}_3$ , are sensitive to  $\text{CH}_3$  polarization: the higher these frequencies, the stronger the  $\text{CH}_3$  group is polarized and the higher is ionicity of the  $\text{CH}_3-\text{Cl}(\text{X})$  bond. In compounds  $\text{CH}_3-\{\text{Cl}_{11}\}$  and  $\text{CH}_3-\text{Cl}^+-\text{CH}_3$ , the CH stretches differ insignificantly (Table 1), pointing to almost the same ionicity of  $\text{CH}_3^+$  bonding to  $\{\text{Cl}^-\}$  and  $\text{Cl}-\text{CH}_3$ , respectively. Nonetheless, one would expect weaker ionicity (stronger covalency) of the bonds in  $\text{CH}_3-\text{Cl}^+-\text{CH}_3$  than in  $\text{CH}_3-\{\text{Cl}_{11}\}$ . The crystal structure of the  $(\text{CH}_3-\text{Cl}^+-\text{CH}_3)\{\text{Cl}_{11}^-\}$  salt shows (Fig. S7 in ESI<sup>†</sup>) that the chloronium Cl-atom forms six ionic bonds with Cl-atoms of the four  $\{\text{Cl}_{11}^-\}$  anions of its





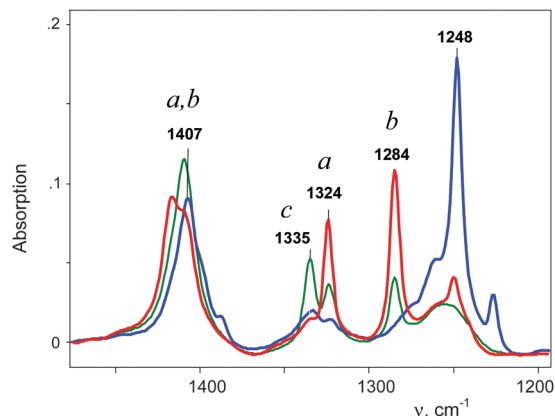
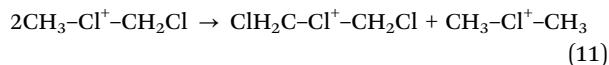


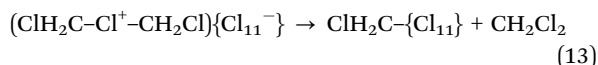
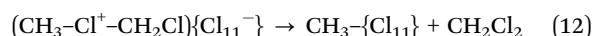
Fig. 14 IR spectra in the frequency region of CH bend vibrations of the salt  $(\text{CH}_3\text{-Cl}^+\text{-CH}_2\text{Cl})\{\text{Cl}_{11}^-\}$  before (blue) and after heating for 5 minutes at 100 °C (red) and 150 °C (green). The most characteristic bands of the formed compounds are marked with (a) for  $(\text{CH}_3)_2\text{Cl}^+$ , (b) for  $(\text{CH}_2\text{Cl})_2\text{Cl}^+$  (IIa) and (c) for  $\text{CH}_3\text{-Cl}_{11}$ .

bands a and b, respectively, Fig. 14 and Fig. S8, ESI,† red). The absorption bands of gaseous HCl did not appear.

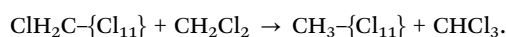
Therefore, at 100 °C, the asymmetric cation  $\text{CH}_3\text{-Cl}^+\text{-CH}_2\text{Cl}$  quickly disproportionates into the more stable symmetric cations:



Further heating of the sample for 5 minutes at 150 °C led to an equal (threefold) reduction in intensity of the bands at 1324 and 1284  $\text{cm}^{-1}$  of the cations  $(\text{CH}_3)_2\text{Cl}^+$  and  $(\text{CH}_2\text{Cl})_2\text{Cl}^+$ , respectively, and appearance and an increase in intensity of the bands of  $\text{CH}_3\text{-Cl}_{11}$  (1335  $\text{cm}^{-1}$ ; Fig. 14, green) and  $\text{ClH}_2\text{C-Cl}_{11}$  ( $\nu_{\text{CCl}_{\text{term}}} = 793 \text{ cm}^{-1}$ ). Additionally, bands with a rotational structure at 1270 and 761  $\text{cm}^{-1}$  of gaseous  $\text{CH}_2\text{Cl}_2$  were observed. Consequently, an increase in temperature facilitates the decomposition of chloronium ions according to eqn (12) and (13); this change should lead to increasing intensity of the IR spectrum of released dichloromethane.



Longer heating of the sample at 150 °C did not enhance the spectrum of  $\text{CH}_2\text{Cl}_2$ . It remained virtually unchanged. In contrast, the bands of  $\text{CHCl}_3$  appeared (1219 and 772  $\text{cm}^{-1}$ ) and were enhanced with heating and a very weak absorption pattern of gaseous HCl developed. The solid-phase spectrum shows increased intensity of the  $\text{CH}_3\text{-Cl}_{11}$  band at 1335  $\text{cm}^{-1}$  and disappearance of the adsorption of  $\text{ClH}_2\text{C-Cl}_{11}$ . Taken together, our data indicate that the  $\text{CHCl}_3$  formation is caused by the interaction of  $\text{ClH}_2\text{C-Cl}_{11}$  with dichloromethane:

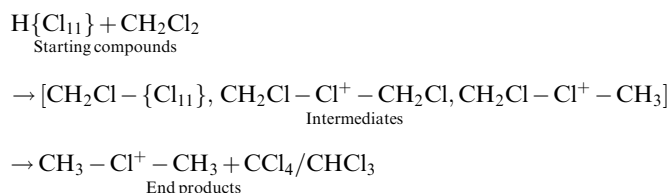


To determine the reason for the formation of trace amounts of HCl, additional studies are needed.

## Conclusions

The salts of symmetric methylchloronium cations,  $(\text{CH}_3)_2\text{Cl}^+\{\text{Cl}_{11}^-\}$  and  $(\text{CH}_2\text{Cl})_2\text{Cl}^+\{\text{Cl}_{11}^-\}$ , and their neutral analogs  $\text{CH}_3\text{-Cl}_{11}$  and  $\text{CH}_2\text{Cl-Cl}_{11}$ , are stable at ambient and increased temperatures. Nevertheless, the asymmetric cation  $\text{ClCH}_2\text{-Cl}^+\text{-CH}_3$  even at ambient temperature disproportionates into symmetrical  $(\text{CH}_3)_2\text{Cl}^+$  and  $(\text{CH}_2\text{Cl})_2\text{Cl}^+$ . The molecular fragment  $\text{CH}_2\text{Cl-}$  of the chloronium ions enters exchange reactions with  $\text{CH}_2\text{Cl}_2$  and  $\text{CHCl}_3$  with increasing reactivity in the order  $\text{ClCH}_2\text{-Cl}^+\text{-CH}_3$ ,  $\text{ClCH}_2\text{-Cl}^+\text{-CH}_2\text{Cl}$ , and  $\text{CH}_2\text{Cl-Cl}_{11}$ , yielding more stable and less reactive  $(\text{CH}_3)_2\text{Cl}^+\{\text{Cl}_{11}^-\}$  and  $\text{CH}_3\text{-Cl}_{11}$  compounds.

If we take into account the reactivity of  $\text{H}\{\text{Cl}_{11}\}$  acid or its chloronium salts with  $\text{CH}_2\text{Cl}_2$ , and the simultaneous disproportionation of the  $\text{CH}_3\text{-Cl}^+\text{-CH}_2\text{Cl}$  cation (eqn (11); initiated by the elevated temperature), then the end products of interactions at ambient, or slightly elevated temperature, are the chemically inert  $\text{CH}_3\text{-Cl}^+\text{-CH}_3$  and  $\text{CCl}_4$ :



Reactions (9)–(13) (as determined in this study) proceed with the formation of a trace amount of HCl, thus indicating the negligible presence of parallel reactions. Under the other conditions, they may appear to be significant and interesting for future research.

## Acknowledgements

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