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Proton coupled electron transfer from the excited state of a ruthenium(II) pyridylimidazole complex†

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Proton coupled electron transfer (PCET) from the excited state of $[\text{Ru}(\text{bpy})_2\text{pyimH}]^{2+}$ (bpy = 2,2'-bipyridine; pyimH = 2-(2'-pyridyl)imidazole) to *N*-methyl-4,4'-bipyridinium (monoquat, MQ⁺) was studied. While this complex has been investigated previously, our study is the first to show that the formal bond dissociation free energy (BDFE) of the imidazole-N–H bond decreases from (91 ± 1) kcal mol⁻¹ in the electronic ground state to (43 ± 5) kcal mol⁻¹ in the lowest-energetic ³MLCT excited state. This makes the $[\text{Ru}(\text{bpy})_2\text{pyimH}]^{2+}$ complex a very strong (formal) hydrogen atom donor even when compared to metal hydride complexes, and this is interesting for light-driven (formal) hydrogen atom transfer (HAT) reactions with a variety of different substrates. Mechanistically, formal HAT between ³MLCT excited $[\text{Ru}(\text{bpy})_2\text{pyimH}]^{2+}$ and monoquat in buffered 1:1 (v : v) CH₃CN/H₂O was found to occur *via* a sequence of reaction steps involving electron transfer from Ru(II) to MQ⁺ coupled to release of the N–H proton to buffer base, followed by protonation of reduced MQ⁺ by buffer acid. Our study is relevant in the larger contexts of photoredox catalysis and light-to-chemical energy conversion.

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1. Introduction

Hydrogen atom transfer (HAT) is important in enzymes and in synthetic organic chemistry, for example for hydrogenations of unsaturated compounds such as ketones and imines. It would be attractive to use visible light to perform HAT reactions under mild reaction conditions, and therefore we explored the (formal) HAT chemistry of photoexcited $[\text{Ru}(\text{bpy})_2\text{pyimH}]^{2+}$ (bpy = 2,2'-bipyridine; pyimH = 2-(2'-pyridyl)imidazole). In pure HAT the transferred electron and proton originate from the same donor orbital, whereas in unidirectional proton coupled electron transfer (PCET) the transfer of a net hydrogen atom occurs from different donor orbitals.^{1–5} This is in fact the case for excited $[\text{Ru}(\text{bpy})_2\text{pyimH}]^{2+}$ because the metal center acts as an electron donor, whereas the proton is released from the pyimH ligand. While early PCET studies have focused largely on reactions between molecules in their electronic ground states,^{6–11} photoinduced PCET is now receiving increasing attention.^{1,12–25} Formal HAT between a transition metal complex in its ³MLCT excited-state and various reaction partners either across a salt bridge,^{14,20–23} or *via* hydrogen bonding interactions have been explored.^{13,24,25} Even hydride transfer from the excited state of an iridium complex was reported recently.²⁶ In order to predict the reactivity of an excited state, thermodynamic quantities such

as its redox potential and acidity constant must be known.^{27,28} For HAT reactions, the determination of bond dissociation free energies (BDFEs) is useful, while for PCETs the calculation of formal BDFEs has proven meaningful.²⁸ This is possible for reactions in the ground state as well as for reactions with excited species. Based on this concept, photochemical conversions of ketones to ketyls could be rationalized.^{2,16,17}

The $[\text{Ru}(\text{bpy})_2\text{pyimH}]^{2+}$ complex (Fig. 1) has long been known,²⁹ in particular Haga and coworkers explored a variety of ruthenium and osmium complexes with pyimH and related (deprotonatable) ligands.^{30–35} Later, Gray and coworkers explored the acid–base and redox chemistry of $[\text{Ru}(\text{bpy})_2\text{pyimH}]^{2+}$ and related complexes in the ground and the lowest ³MLCT excited state.^{29,36} However, the formal BDFE of the peripheral N–H bond of $[\text{Ru}(\text{bpy})_2\text{pyimH}]^{2+}$ and related complexes has never been determined, and the excited-state PCET chemistry remained unexplored, except in the case of an Ir(III) complex with a

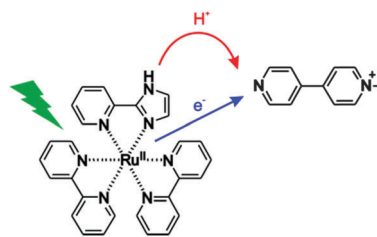


Fig. 1 The investigated process in this work: transfer of one electron and one proton from $[\text{Ru}(\text{bpy})_2\text{pyimH}]^{2+}$ to monoquat (MQ⁺) upon photoexcitation, corresponding to net transfer of a hydrogen atom.

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2,2'-biimidazole ligand.¹⁴ In structurally related complexes such as $[\text{Ru}(\text{acac})_2\text{pyimH}]^{2+}$ (acac = acetylacetonato), formal BDFEs were estimated for the electronic ground state, and values around 62 kcal mol^{-1} were found.^{9,37} We anticipated that $[\text{Ru}(\text{bpy})_2\text{pyimH}]^{2+}$ might exhibit an unusually low N–H BDFE in its long-lived ³MLCT excited state, making it potentially an equally potent (formal) hydrogen atom donor as previously investigated metal hydride complexes in their electronic ground states.^{4,38}

In the following we present the thermochemical characterization of ground and ³MLCT excited states of $[\text{Ru}(\text{bpy})_2\text{pyimH}]^{2+}$ in buffered 1:1 (v:v) $\text{CH}_3\text{CN}/\text{H}_2\text{O}$. In this solvent mixture, well-defined pH values can easily be obtained by a variety of buffers and the solubility of $[\text{Ru}(\text{bpy})_2\text{pyimH}]^{2+}$ as well as that of a variety of substrates is good. We find an N–H BDFE of only $(43 \pm 5) \text{ kcal mol}^{-1}$ in the emissive ³MLCT excited state based on thermodynamic cycles and on the photoinduced PCET chemistry with monoquat (MQ^+). The acceptor was chosen due to its ability to act as a combined electron–proton acceptor, the favourable spectroscopic properties of its radical form and the importance of pyridyl radicals for the reduction of CO_2 .^{39–41} The PCET reaction mechanism between photoexcited $[\text{Ru}(\text{bpy})_2\text{pyimH}]^{2+}$ and MQ^+ was explored in detail.

2 Results and discussion

Spectroscopy and thermodynamics of the ground state

pK_a of $[\text{Ru}(\text{bpy})_2\text{pyimH}]^{2+}$. For determination of the pK_a value of the electronic ground state, absorption spectra were recorded at different pH values between pH 3.7 and pH 10.2 (ESI† Fig. S1a) using suitable buffers. By plotting the absorbance at the MLCT absorption maxima at 460 nm (protonated form) and 491 nm (deprotonated form) and sigmoidal fitting, $pK_a = 8.1 \pm 0.1$ was found for 1:1 (v:v) $\text{CH}_3\text{CN}/\text{H}_2\text{O}$ (ESI† Fig. S1b), in agreement with a prior study that reported $pK_a = 7.9 \pm 0.1$ in H_2O containing 5% methanol.²⁹

pK_a of $[\text{Ru}(\text{bpy})_2\text{pyimH}]^{3+}$. Oxidative cyclic voltammetry sweeps probing metal oxidation were performed in the pH range between pH 1 and pH 11. Characteristic voltammograms for three pH regions of interest are shown in Fig. 2a. In the acidic pH range where $[\text{Ru}(\text{bpy})_2\text{pyimH}]^{2+}$ remains protonated after oxidation of Ru^{II} to Ru^{III} , the voltammograms show one reversible oxidation wave at $E_{\text{ox}}^{\text{prot}} = (1.00 \pm 0.05) \text{ V vs. SCE}$ with peak separations between 67 and 80 mV. This behaviour is observed up to $\text{pH} = 3.6 \pm 0.1$, corresponding to the acidity constant of $[\text{Ru}(\text{bpy})_2\text{pyimH}]^{3+}$ (pK_a^{ox}). At higher pH values oxidation is irreversible at a sweep rate of 100 mV s^{-1} (Fig. 2a and Fig. S5 in the ESI†) due to deprotonation of the pyimH ligand of the

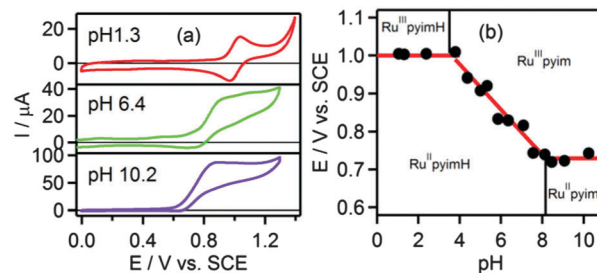


Fig. 2 (a) Voltammograms of $[\text{Ru}(\text{bpy})_2\text{pyimH}]^{2+}$ in 1:1 (v:v) $\text{CH}_3\text{CN}/\text{H}_2\text{O}$ with 0.05 M buffer at different pH values, (b) Pourbaix diagram of $[\text{Ru}(\text{bpy})_2\text{pyimH}]^{2+}$. The slope between pH 3.6 and pH 8.1 is $-(60 \pm 4) \text{ mV per pH}$. A comprehensive set of voltammograms is shown in Fig. S5 of the ESI.†

Table 2 Ground and excited state redox potentials (E_{ox} , $*E_{\text{ox}}$), ³MLCT energy (E_{0-0}), emission maxima at 25 °C (λ_{max}) and luminescence lifetimes at 25 °C under aerated and deaerated conditions of $[\text{Ru}(\text{bpy})_2\text{pyimH}]^{2+}$ in 1:1 (v:v) $\text{CH}_3\text{CN}/\text{H}_2\text{O}$ with 0.05 M buffer

	$[\text{Ru}(\text{bpy})_2\text{pyimH}]^{2+}$	$[\text{Ru}(\text{bpy})_2\text{pyim}]^+$
E_{ox} [V vs. SCE]	1.00 ± 0.05	0.73 ± 0.05
$*E_{\text{ox}}$ [V vs. SCE]	-1.1 ± 0.1	-1.2 ± 0.1
E_{0-0} [eV]	$2.1 \pm 0.1^{a,b}$	1.9 ± 0.1^b
λ_{max} [nm] 25 °C	625 ± 5	675 ± 5
τ [ns] aerated	110 ± 10	50 ± 5
τ [ns] deaerated	210 ± 20	70 ± 7

^a Taken from references.^{29,42} ^b Determined from emission in ethanol/methanol 4:1 (v:v) at 77 K.

oxidized complex. Oxidation potentials of the deprotonated complex were estimated by determining the relevant inflection points of the oxidation waves. Plotting the oxidation potentials in volt vs. pH gives the data points for the Pourbaix diagram in Fig. 2b. In the range between $pK_a^{\text{ox}} = 3.6$ and $pK_a = 8.1$ the $\text{Ru}^{\text{II/III}}$ oxidation wave shifts cathodically with a slope of $-(60 \pm 4) \text{ mV per pH unit}$, as expected for a 1-electron-1-proton-process. Oxidation of the deprotonated complex, $[\text{Ru}(\text{bpy})_2\text{pyim}]^+$, occurs at $E_{\text{ox}}^{\text{dep}} = (0.73 \pm 0.05) \text{ V vs. SCE}$ (Table 2). Thus, the total cathodic shift between oxidation potentials of $[\text{Ru}(\text{bpy})_2\text{pyimH}]^{2+}$ and $[\text{Ru}(\text{bpy})_2\text{pyim}]^+$ is 270 mV, which is smaller than the previously reported shift of 380 mV in neat acetonitrile.²⁹

Excited state properties

Excited state pK_a^* of $[\text{Ru}(\text{bpy})_2\text{pyimH}]^{2+}$. Luminescence spectra of $[\text{Ru}(\text{bpy})_2\text{pyimH}]^{2+}$ were measured between pH 3 and pH 10 in 1:1 (v:v) $\text{CH}_3\text{CN}/\text{H}_2\text{O}$ containing 0.05 M buffer to control the pH (Fig. 3a). All spectra were recorded at identical complex concentration and their intensity was normalized to the intensity of the most acidic sample; excitation occurred into the isosbestic point at 474 nm (Fig. S1a, ESI†). The decrease in intensity is due to different luminescence quantum yields of the protonated and deprotonated complex and due to deprotonation in the excited state. Emission decays were measured at 630 nm to determine the lifetimes of the protonated (τ^{prot}) and deprotonated ³MLCT state (τ^{dep}) in aerated and deaerated solution (Table 2). For these measurements the complex was excited at 532 nm with laser pulses of ca. 10 ns duration.

Table 1 Acidity constants of $[\text{Ru}(\text{bpy})_2\text{pyimH}]^{2+}$ in 1:1 (v:v) $\text{CH}_3\text{CN}/\text{H}_2\text{O}$ in the electronic ground state (pK_a), in the long-lived ³MLCT excited state (pK_a^*) and in the one-electron oxidized form (pK_a^{ox})

pK_a	8.1 ± 0.1
pK_a^*	5.6 ± 0.3
pK_a^{ox}	3.6 ± 0.1



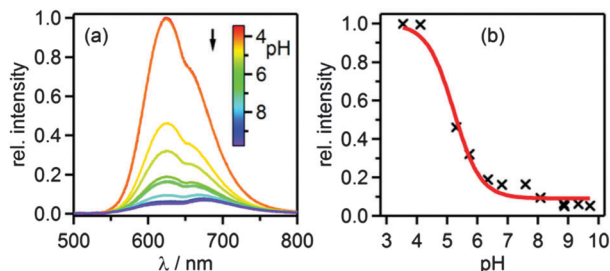


Fig. 3 (a) Luminescence of 40 μM $[\text{Ru}(\text{bpy})_2\text{pyimH}]^{2+}$ in aerated 1 : 1 (v : v) $\text{CH}_3\text{CN}/\text{H}_2\text{O}$ with 0.05 M buffer at different pH values following excitation into the isosbestic point at 474 nm. (b) Relative luminescence intensity at 625 nm vs. pH.

The acidity constant in the long-lived $^3\text{MLCT}$ -state ($\text{p}K_{\text{a}}^*$) was then determined from the inflection point of the steady-state emission titration curve ($\text{pH}_i = 5.2 \pm 0.2$) in Fig. 3b and the excited-state lifetimes of protonated (τ^{prot}) and deprotonated complex (τ^{dep}).

$$\text{p}K_{\text{a}}^* = \text{pH}_i + \log[\tau^{\text{prot}}/\tau^{\text{dep}}] \quad (1)$$

With eqn (1) one obtains $\text{p}K_{\text{a}}^* = 5.6 \pm 0.3$.²⁷ Based on the emission maxima of $[\text{Ru}(\text{bpy})_2(\text{pyimH})]^{2+}$ and $[\text{Ru}(\text{bpy})_2(\text{pyim})]^+$ at 25 $^\circ\text{C}$, the Förster equation yields $\text{p}K_{\text{a}}^* = 5.3 \pm 0.6$ (ESI,† page S6). This value is in good agreement with that determined from the luminescence titration, and also with the previously reported $\text{p}K_{\text{a}}^*$ in water with 5% methanol.²⁹ The increase of acidity in the excited state compared to the ground state indicates that the $^3\text{MLCT}$ state is localized on the bpy spectator ligands, as noted earlier.³⁶

Transient absorption spectra were measured using pulsed laser excitation at 532 nm. Under acidic conditions (Fig. 4a), $[\text{Ru}(\text{bpy})_3]^{2+}$ -like transient spectra were observed, exhibiting a bleach around 450 nm and increased intensity around 370 nm.⁴³ At basic pH the MLCT band is red-shifted (ESI,† Fig. S1), and therefore the MLCT-bleach in transient absorption is red-shifted as well (Fig. 4b). Transients that were recorded in the pH range between pH 5.0 and 8.1 exhibit a prominent feature at 500 nm (Fig. 4c) which can be explained by deprotonation in the excited state and the accumulation of deprotonated complex in the ground state. This interpretation is confirmed by a subtraction of the ground-state UV-vis spectra of the protonated and deprotonated complex (Fig. 4d) which also exhibits the prominent positive feature at 500 nm. The temporal evolution of this signal at pH 6.4 in presence of 0.05 M acetic acid/sodium acetate buffer and $ca. 10^{-5}$ M complex concentration is shown below. In the absence of any reaction partner the formation of $[\text{Ru}(\text{bpy})_2\text{pyim}]^+$, the deprotonated complex in its ground state, occurs with a time constant of $\tau = (60 \pm 10)$ ns. Re-protonation then occurs with a time constant of $\tau = (106 \pm 10)$ ns as discussed later. The kinetics of these deprotonation and protonation events are dictated by the buffer concentration. At pH 6.3 the quenching constant k_q is $(8.3 \pm 0.5) \times 10^8 \text{ L mol}^{-1} \text{ s}^{-1}$ based on a Stern–Volmer luminescence quenching experiment (ESI,† Fig. S7).

At any given pH and buffer concentration, the amount of accumulated $[\text{Ru}(\text{bpy})_2\text{pyim}]^+$ in the ground state correlates

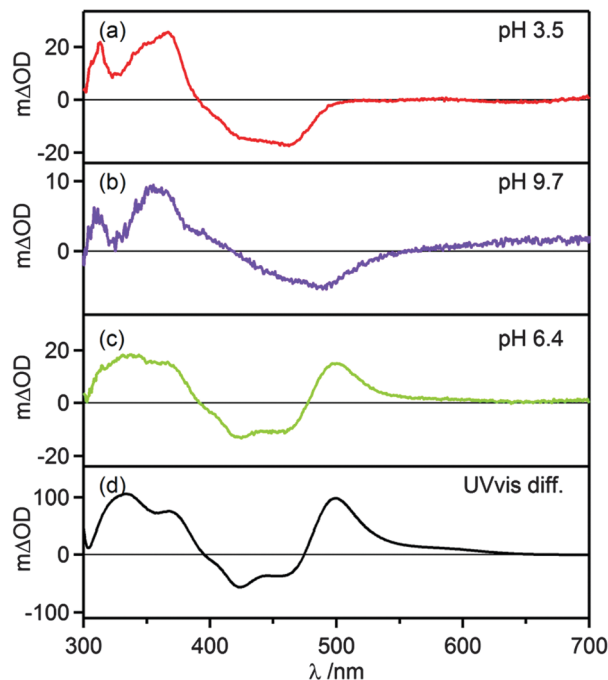


Fig. 4 (a–c) Transient absorption spectra of 20 μM $[\text{Ru}(\text{bpy})_2\text{pyimH}]^{2+}$ in 1 : 1 (v : v) $\text{CH}_3\text{CN}/\text{H}_2\text{O}$ with 0.05 M buffer at different pH values. Excitation occurred at 532 nm with laser pulses of $ca. 10$ ns duration, the spectra were recorded without time delay over a period of 200 ns. (d) Difference of ground state UV-vis spectra of protonated and deprotonated complex.

with the acid–base equilibration in the $^3\text{MLCT}$ excited state, and this reflects directly in the intensity of the transient band at 500 nm. A plot of the change in optical density at 500 nm vs. pH yields an inflection point at $\text{pH} 5.6 \pm 0.2$ (ESI,† Fig. S4), in line with the $\text{p}K_{\text{a}}^*$ value determined by luminescence titration and the Förster equation.

Thus, at sufficiently high pH, $[\text{Ru}(\text{bpy})_2\text{pyimH}]^{2+}$ exhibits ordinary photoacid behavior similar to hydroxyarenes such as naphthols and hydroxypyrenetrisulfonate (“pyranine”).⁴⁴

Excited state redox chemistry of $[\text{Ru}(\text{bpy})_2\text{pyimH}]^{2+}$. The excited state oxidation potentials (E_{ox}^*) were estimated based on the ground state redox potentials (E_{ox}) and the $^3\text{MLCT}$ energy (E_{0-0}) using eqn (2).⁴²

$$E_{\text{ox}}^* = E_{\text{ox}} - E_{0-0} \quad (2)$$

E_{0-0} was determined from low-temperature luminescence spectroscopy (Fig. S2, ESI†). For $[\text{Ru}(\text{bpy})_2\text{pyimH}]^{2+}$ we determined $E_{0-0}^{\text{prot}} = (2.1 \pm 0.1) \text{ eV}$ ^{29,42} and for $[\text{Ru}(\text{bpy})_2\text{pyim}]^+$ we found $E_{0-0}^{\text{dep}} = (1.9 \pm 0.1) \text{ eV}$. Based on these E_{0-0} values (see also Table 2), excited state redox potentials of $E^{\text{prot}} = -(1.1 \pm 0.1) \text{ V}$ vs. SCE and $E^{\text{dep}} = -(1.2 \pm 0.1) \text{ V}$ vs. SCE were estimated for the protonated and deprotonated complex, respectively. Expectedly, oxidation is far easier in the $^3\text{MLCT}$ excited state than in the electronic ground state.

BDFEs and “cube” scheme

A graphical summary of all relevant thermodynamic parameters for $[\text{Ru}(\text{bpy})_2\text{pyimH}]^{2+}$ is provided in Scheme 1. This so-called

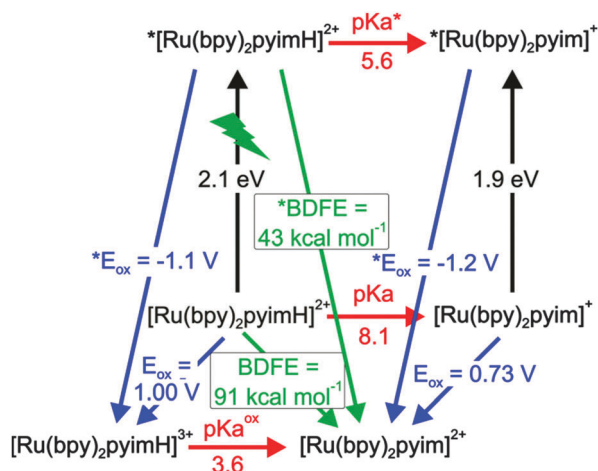


"cube" scheme is a three dimensional illustration displaying ground state redox potentials (blue) and acidity constants (red) on the bottom and excited state potentials (blue) and pK_a values (red) on the top. ³MLCT excitation energies are represented by vertical black arrows. N–H bond dissociation free energies (BDFEs) can be estimated using eqn (3) and the experimentally determined acidity constants and oxidation potentials.^{2,45,46}

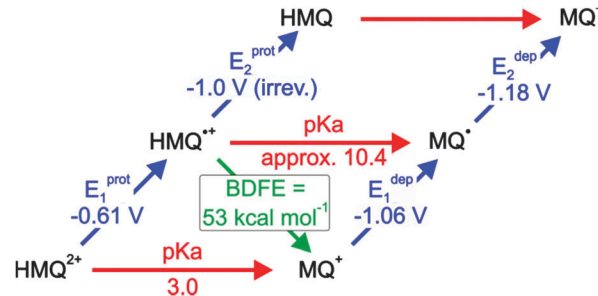
$$\text{BDFE (N-H)} = 1.37 \text{ pK}_a + 23.06 E^\circ + 57.6 \text{ kcal mol}^{-1} \quad (3)$$

In eqn (3), E° must be entered in units of V vs. NHE; the last summand is a solvent-characteristic parameter describing solvation of hydrogen atoms. The resulting BDFEs for [Ru(bpy)₂pyimH]²⁺ in the electronic ground and excited states are BDFE = (91 ± 1) kcal mol⁻¹ and *BDFE = (43 ± 5) kcal mol⁻¹ (green arrows in Scheme 1). The ground-state BDFE is comparable to primary and secondary amines^{2,47,48} but somewhat higher than related ruthenium pyridylimidazole complexes which were characterized in acetonitrile.^{9,37} Upon photoexcitation, the BDFE drops by 48 kcal mol⁻¹, which is essentially equal to the energy of the absorbed visible photon. Interestingly, the resulting excited-state BDFE is comparable to metal hydride catalysts for hydrogenation reactions for which M–H BDFEs ranging from 50 to 55 kcal mol⁻¹ for M = vanadium, 58 kcal mol⁻¹ for M = chromium, and 68 kcal mol⁻¹ for M = tungsten have been reported.^{4,38} In principle this drop in BDFE is expected to occur for other related metal complexes in the course of photoexcitation, but prior studies have not explicitly reported on this effect. Presumably this is due to the fact that in many cases the necessary redox potentials and acidity constants were not always determined in the same solvent, which complicates the application of eqn (3). Estimations based on prior work yields for the first N–H BDFE of [Ru(bpy)₂(2,2'-biimidazole)]²⁺ a decrease from 86 to 40 kcal mol⁻¹ between the electronic ground state and the long-lived ³MLCT state.^{29,49}

In order to test whether the N–H BDFE is really that low, we set out to react photoexcited [Ru(bpy)₂pyimH]²⁺ in PCET



Scheme 1 Thermodynamic "cube" scheme for [Ru(bpy)₂pyimH]²⁺ in 1:1 (v:v) CH₃CN/H₂O based on the data in Tables 1 and 2. Horizontal/red: pK_a values, orthogonal in black: triplet energy E_{0-0} , pointing towards the reader in blue: oxidation potentials in V vs. SCE, diagonal in green: BDFEs.



Scheme 2 Thermodynamic "square" scheme of monoquat (MQ⁺) in 1:1 (v:v) CH₃CN/H₂O. Horizontal with red arrows: pK_a values, upward right with blue arrows: redox potentials in V vs. SCE, downward left with green arrow: BDFEs. The pK_a of HMQ^{•+} was taken from the literature.⁴⁶

chemistry with a suitable formal H atom acceptor. We identified *N*-methyl-4,4'-bipyridinium (monoquat, MQ⁺) as a promising candidate. Meyer and coworkers already proposed the use of MQ⁺ for detecting PCET photoproducts because its one-electron reduced and protonated congener (MQH^{•+}) exhibits absorption features that can be identified unambiguously.⁵⁰

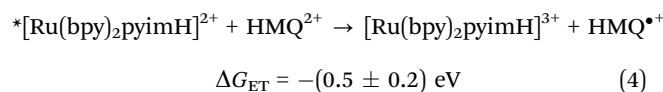
Before performing actual photochemical experiments between [Ru(bpy)₂pyimH]²⁺ and MQ⁺, the thermodynamic properties of MQ⁺ in 1:1 (v:v) CH₃CN/H₂O were determined. The results from acid–base titration and electrochemical experiments are in the ESI[†] (Fig. S8 and S9). Here we merely report the final results in a thermodynamic "square" scheme (Scheme 2).

The key finding is that the HMQ^{•+} radical has an N–H BDFE of (53 ± 1) kcal mol⁻¹. Consequently, photoexcitation of [Ru(bpy)₂pyimH]²⁺ in presence of MQ⁺ is expected to lead to formal HAT, resulting in [Ru(bpy)₂pyim]⁺ and HMQ^{•+}. Based on an N–H *BDFE of (43 ± 5) kcal mol⁻¹ for the photoexcited complex (Scheme 1), the driving-force for this reaction should be $-(10 \pm 6)$ kcal mol⁻¹, which corresponds to $-(0.4 \pm 0.3)$ eV. In the following we report on the photochemistry between [Ru(bpy)₂pyimH]²⁺ and MQ⁺ as a function of pH.

From simple photoinduced ET to formal HAT

In the electronic ground state, electron transfer (ET) and proton transfer reactions (PT) between [Ru(bpy)₂pyimH]²⁺ and MQ⁺ are strongly endergonic ($\Delta G_{\text{ET}} = +(2.1 \pm 0.1)$ eV, $\Delta G_{\text{PT}} = +(0.30 \pm 0.02)$ eV) and therefore no ground-state chemistry occurs. Regarding excited-state chemistry, there are in fact three different pH domains which are discussed individually in the following 3 sub-sections.

Acidic pH – photoinduced electron transfer. In the acidic range, both the complex and the acceptor are protonated in the ground and excited state. Under these conditions, the expected reaction is photoinduced electron transfer (eqn (4)).



The transient absorption spectrum recorded at pH = 2 (Fig. 5a) confirms this expectation. Directly after the laser pulse one observes a bleach around 450 nm which is compatible with



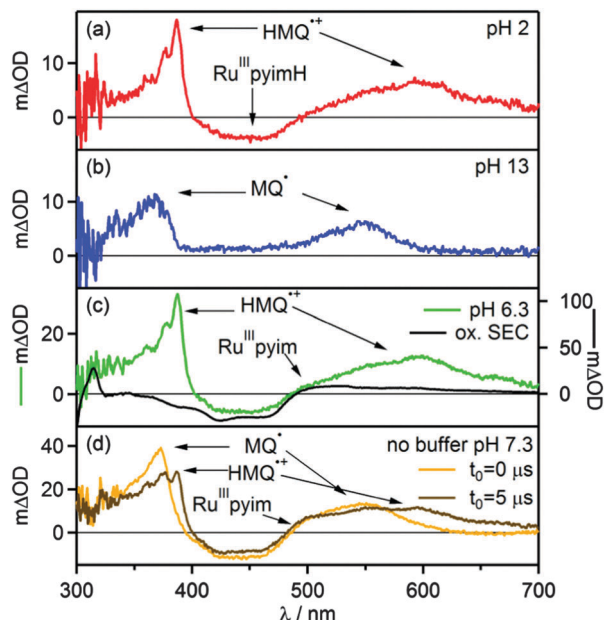
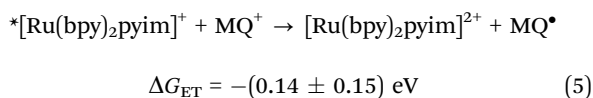


Fig. 5 Transient absorption spectra of $[\text{Ru}(\text{bpy})_2\text{pyimH}]^{2+}$ in 1:1 (v:v) $\text{CH}_3\text{CN}/\text{H}_2\text{O}$ recorded in presence of 60 mM MQ^+ following excitation at 532 nm with laser pulses of ca. 10 ns duration. Detection occurred by integration over 200 ns. (a) 30 μM $[\text{Ru}(\text{bpy})_2\text{pyimH}]^{2+}$ at pH 2 with 0.05 M buffer, data recorded with a delay time (t_0) of 5 μs , (b) 30 μM $[\text{Ru}(\text{bpy})_2\text{pyimH}]^{2+}$ at pH 13, data recorded with a delay time (t_0) of 5 μs , (c) 50 μM $[\text{Ru}(\text{bpy})_2\text{pyimH}]^{2+}$ at pH 6.3 with 0.05 M acetate buffer, data recorded with a delay time (t_0) of 1 μs . (d) 50 μM $[\text{Ru}(\text{bpy})_2\text{pyimH}]^{2+}$ in unbuffered in 1:1 (v:v) $\text{CH}_3\text{CN}/\text{H}_2\text{O}$ at pH 7.3 recorded without delay time (orange trace) and with a time delay of 5 μs (brown trace).

metal oxidation, and the signatures of the $\text{HMQ}^{\bullet+}$ cation radical appear at 387 nm and 610 nm.⁴⁶ The latter closely resemble the well-known radical of methyl viologen (MV^{2+}).⁵¹ The photoinduced ET reaction from eqn (4) is associated with $\Delta G_{\text{ET}} = -(0.5 \pm 0.2)$ eV. A Stern–Volmer experiment under acidic conditions reveals a quenching constant of $k_q = (6.8 \pm 0.1) \times 10^8 \text{ L mol}^{-1} \text{ s}^{-1}$ (ESI,† Fig. S14), which is comparable to what was found for the reaction of $^*\text{[Ru}(\text{bpy})_3]^{2+}$ with MV^{2+} in water ($k_q = 5.9 \times 10^8 \text{ L mol}^{-1} \text{ s}^{-1}$).^{52,53} The thermal reverse ET from $\text{HMQ}^{\bullet+}$ to $[\text{Ru}(\text{bpy})_2\text{pyimH}]^{3+}$ in the electronic ground state then occurs on a time scale of approximately 100 μs , as determined by monitoring the $\text{HMQ}^{\bullet+}$ signal at 610 nm (ESI,† Fig. S10).

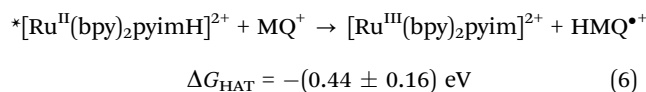
Basic pH – ET. In the basic range, the complex and the acceptor are both deprotonated in the ground and excited state, and consequently they are expected to undergo photoinduced electron transfer according to eqn (5).



In transient absorption spectroscopy, the neutral monoquat radical (MQ^{\bullet}) with characteristic absorptions at 365 nm and 545 nm is observed.⁵¹ The expected MLCT bleach of the ruthenium complex overlaps with a positive contribution of MQ^{\bullet} hence the flat region in the spectrum between 390 and 480 nm.

The Stern–Volmer luminescence quenching experiment yields a quenching constant of $k_q = (20.7 \pm 0.3) \times 10^8 \text{ L mol}^{-1} \text{ s}^{-1}$ (ESI,† Fig. S15), which is close to the diffusion limit. The thermal reverse ET from MQ^{\bullet} to $[\text{Ru}(\text{bpy})_2\text{pyim}]^{2+}$ takes place on a time scale of 10 μs (ESI,† Fig. S11).

Middle pH range – formal HAT. In the middle pH range the most interesting photochemistry is expected. At pH 6.3 the complex is protonated in its ground state but becomes deprotonated upon excitation to the $^3\text{MLCT}$ state as well as upon oxidation of Ru^{II} to Ru^{III} (Scheme 1). On the other hand, $\text{MQ}^{\bullet+}$ is not protonated, but it is expected to be protonated upon one-electron reduction (Scheme 2). According to the formal N–H BDFEs determined above for photoexcited $[\text{Ru}(\text{bpy})_2\text{pyimH}]^{2+}$ ($(43 \pm 5) \text{ kcal mol}^{-1}$, Scheme 1) and for $\text{HMQ}^{\bullet+}$ ($(53 \pm 1) \text{ kcal mol}^{-1}$, Scheme 2) the formal hydrogen atom transfer reaction in eqn (6) should be associated with a reaction free energy of $-(10 \pm 6) \text{ kcal mol}^{-1}$.

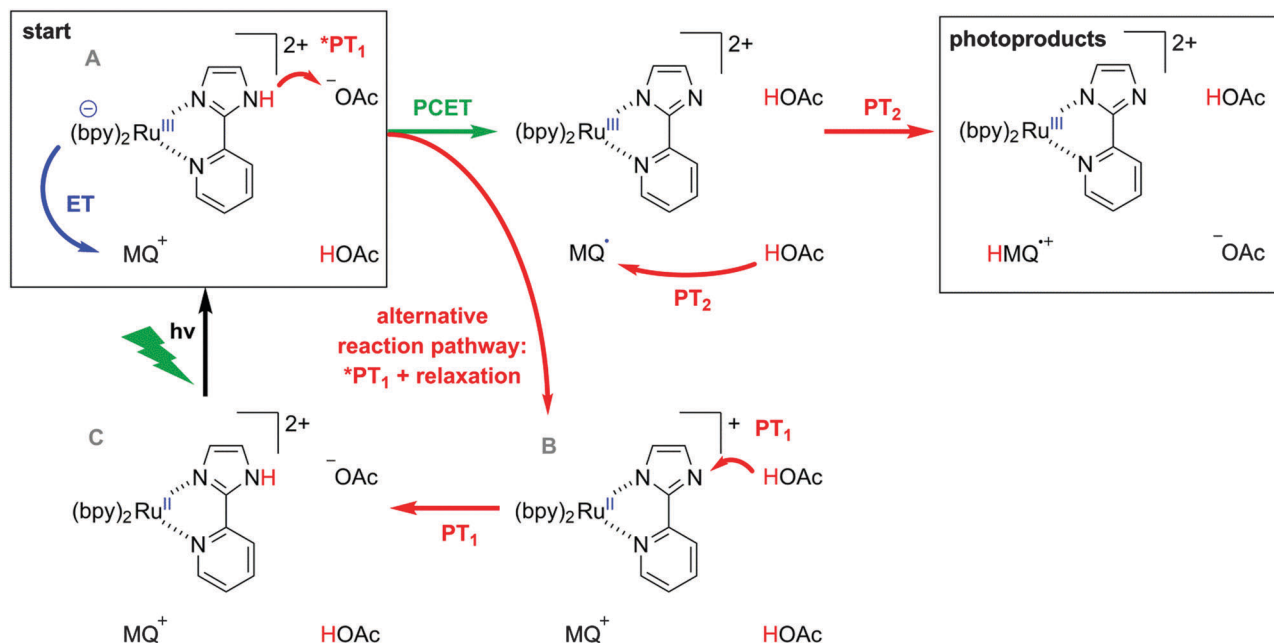


In transient absorption spectroscopy the two photoproducts from eqn (6) are indeed observed (Fig. 5c). When recording transient absorption spectra with a time delay (t_0) of 1 μs there is clear evidence for $\text{HMQ}^{\bullet+}$ (signals at 387 and 610 nm) and for $[\text{Ru}^{\text{III}}(\text{bpy})_2\text{pyim}]^{2+}$ (bleach around 450 nm), as confirmed by spectro-electrochemical (SEC) studies (black trace in Fig. 5c). However, mechanistically direct HAT between $^*\text{[Ru}^{\text{II}}(\text{bpy})_2\text{pyimH}]^{2+}$ and MQ^+ in presence of aqueous buffer is highly improbable, particularly in view of the positive charges on both reactants. Moreover, the lowest-energetic MLCT excitation in $[\text{Ru}^{\text{II}}(\text{bpy})_2\text{pyimH}]^{2+}$ involves promotion of an electron into a bpy-localized orbital rather than a pyimH orbital (see above).³⁶ Consequently, formal HAT between photoexcited $[\text{Ru}^{\text{II}}(\text{bpy})_2\text{pyimH}]^{2+}$ and MQ^+ most likely involves a sequence of electron and proton transfer steps as illustrated in Scheme 3: following excitation of $[\text{Ru}^{\text{II}}(\text{bpy})_2\text{pyimH}]^{2+}$, photoinduced electron transfer to MQ^+ is coupled to release of the pyimH N–H proton to buffer base (acetate anion (AcO^-), $^*\text{PT}_1$ in Scheme 3). This PCET process can occur either in stepwise or concerted fashion, and it results in $[\text{Ru}^{\text{III}}(\text{bpy})_2\text{pyim}]^{2+}$, MQ^{\bullet} , and HOAc (PCET arrow in Scheme 3). Subsequent proton transfer (PT_2 step in Scheme 3) between buffer acid (HOAc) and MQ^{\bullet} then leads to the photoproducts detected in Fig. 5c. In fact, these two reaction steps can be temporally resolved (see below).

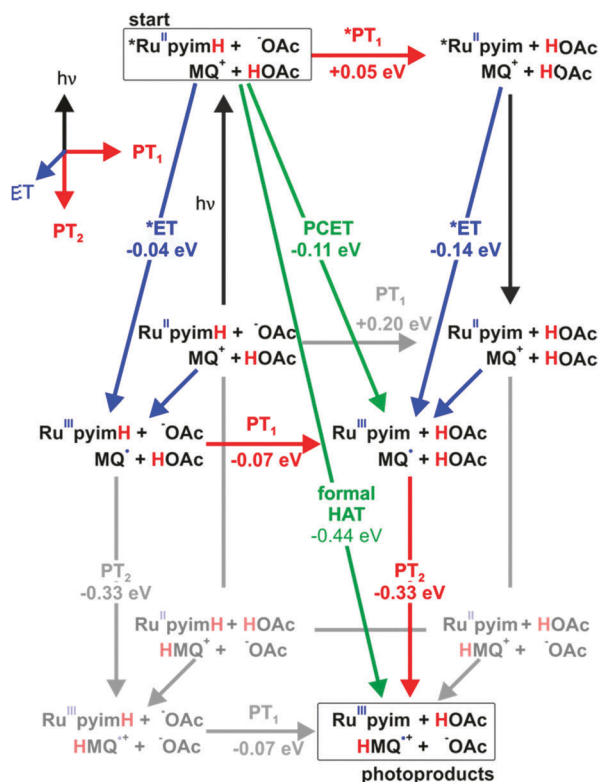
The reaction free energies for all conceivable reaction pathways are summarized in Scheme 4. The driving-forces in Scheme 4 emerge directly from the thermodynamic parameters of the two reaction partners in Schemes 1 and 2. The overall formal HAT process is dissected into photoexcitation (black arrows), electron transfer from the metal complex to monoquat (blue arrows), proton transfer from the metal complex to buffer base (PT_1 , red arrows), and proton transfer from buffer acid to monoquat (PT_2 , red arrows) as discussed above on the basis of Scheme 3.

From the starting point at the top left corner of Scheme 4, concerted proton-electron transfer (CPET) to form $[\text{Ru}^{\text{III}}(\text{bpy})_2\text{pyim}]^{2+}$, MQ^{\bullet} , and protonated buffer base (HOAc) is a plausible initial





Scheme 3 Possible pathways for reaction of $^*[\text{Ru}(\text{bpy})_2\text{pyimH}]^{2+}$ with MQ^+ in acetate-buffered solution.



Scheme 4 Extended "cube"-scheme illustrating all reaction pathways for formal HAT between $^*[\text{Ru}(\text{bpy})_2\text{pyimH}]^{2+}$ and MQ^+ in acetate-buffered solution (based on a combination of Schemes 1 and 2). Excited state chemistry is shown in the upper level, ET processes are pointing towards the reader, PT_1 (complex to buffer base) is shown in horizontal direction and PT_2 (buffer acid to monoquat) is shown in the two lower levels. Driving forces for the individual reaction steps are based on experimentally determined redox potentials and acidity constants.

reaction pathway since $\Delta G_{\text{PCET}} = -(0.11 \pm 0.16)$ eV for this process (green arrow). However, a sequence of electron and proton transfer events cannot be excluded on thermodynamic grounds and would be equally compatible with our experimental data.

Classical Stern-Volmer luminescence quenching experiments could not be performed for determination of the kinetics of the initial PCET process (green arrow in Scheme 3) because the luminescence of $[\text{Ru}(\text{bpy})_2\text{pyimH}]^{2+}$ is strongly quenched in presence of buffer molecules ($k_q = (8.3 \pm 0.5) \times 10^8 \text{ L mol}^{-1} \text{ s}^{-1}$ as described earlier). Therefore, we performed experiments at 5 mM buffer concentration for which we used a different spectroscopic observable to monitor the kinetics of the relevant PCET process: when measuring the transient absorption spectrum immediately after laser excitation, an additional band at 500 nm becomes observable. Based on the data in Fig. 4c and d, this band can be attributed unambiguously to $[\text{Ru}^{\text{II}}(\text{bpy})_2\text{pyim}]^+$, *i.e.*, to the deprotonated Ru^{II} complex in the electronic ground state. This species accumulates in a side-reaction to PCET as illustrated in the lower line, right side of Scheme 3; a subset of all excited complexes undergoes PCET chemistry to the photoproducts shown in the top right corner of Scheme 3 whereas another subsets merely acts as a photoacid (lower line, right side). Thus, the intensity of the transient absorption signal at 500 nm is a measure for the amount of ruthenium complexes that have been photoexcited but that have not undergone PCET chemistry. The temporal evolution of the transient signal at 500 nm shows classical $A \rightarrow B \rightarrow C$ reaction kinetics (Fig. 6a), with species A corresponding to $^*[\text{Ru}^{\text{II}}(\text{bpy})_2\text{pyimH}]^{2+}$, species B being $[\text{Ru}^{\text{II}}(\text{bpy})_2\text{pyim}]^+$, and species C corresponding to the protonated ground state as shown in Scheme 3. In absence of MQ^+ the signal at 500 nm rises with $\tau^{A \rightarrow B} = (105 \pm 10)$ ns and decays with $\tau^{B \rightarrow C} = (1.0 \pm 0.1)$ μs . With increasing MQ^+ concentration $\tau^{A \rightarrow B}$ decreased, $\tau^{B \rightarrow C}$ remained constant and



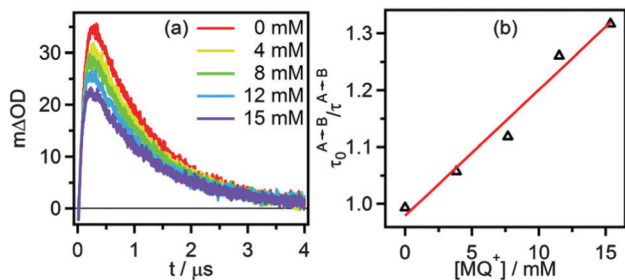


Fig. 6 (a) Change in optical density at 500 nm as a function of time after 532 nm excitation of 40 μM $[\text{Ru}(\text{bpy})_2\text{pyimH}]^{2+}$ in 1:1 (v:v) $\text{CH}_3\text{CN}/\text{H}_2\text{O}$ containing 5 mM acetate buffer and increasing concentrations of MQ^+ . (b) Pseudo Stern–Volmer plot based on the kinetics for step A \rightarrow B with constant B \rightarrow C lifetime, see text for details.

the maximum intensity of the signal decreased. A plot of $\tau_0^{A \rightarrow B}/\tau^{A \rightarrow B}$ vs. $[\text{MQ}^+]$ is shown in Fig. 6b. A linear regression fit yields a pseudo-Stern–Volmer constant (K_{SV}') of $(22 \pm 2) \text{ L mol}^{-1}$ and a quenching constant (k_q') of $(2.1 \pm 0.4) \times 10^8 \text{ L mol}^{-1} \text{ s}^{-1}$. This PCET-quenching constant is on the same order of magnitude as the $^*\text{PT}_1$ -quenching constant and therefore both processes are competitive.

In buffered solution at pH 6.3 the final PCET photoproducts (*i.e.* $[\text{Ru}^{\text{III}}(\text{bpy})_2\text{pyim}]^{2+}$ and HMQ^{*+}) are formed within the first microsecond at buffer concentrations between 5 and 50 mM (Fig. 5c). When monitoring the transient absorption signals at 387 nm and 610 nm (Fig. 7b and c) it becomes evident that the formation of HMQ^{*+} occurs more slowly than deactivation of the $^3\text{MLCT}$ excited state of the $[\text{Ru}^{\text{II}}(\text{bpy})_2\text{pyimH}]^{2+}$ complex (Fig. 7a). Based on the emission decay at 630 nm, the $^3\text{MLCT}$

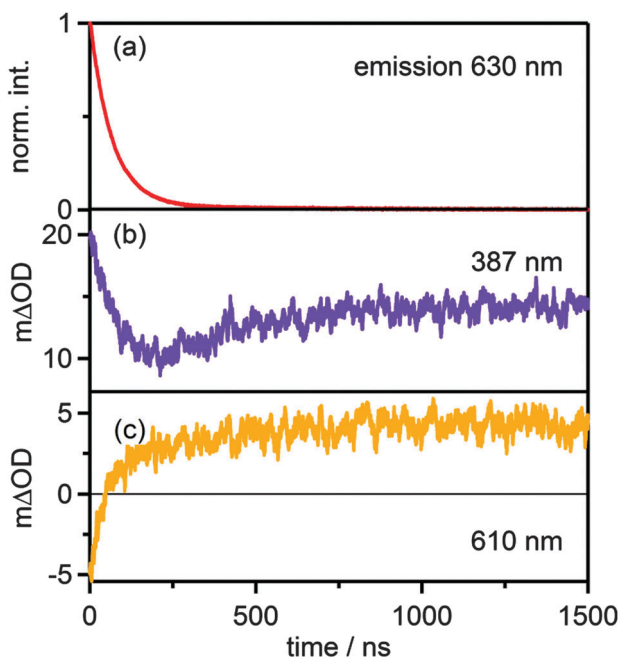


Fig. 7 (a) Decay of the emission signal at 630 nm and temporal evolution of the transient absorption signal at 387 nm (b) and at 610 nm (c) in the reaction of 40 μM $[\text{Ru}(\text{bpy})_2\text{pyimH}]^{2+}$ with 15 mM MQ^+ in 1:1 (v:v) $\text{CH}_3\text{CN}/\text{H}_2\text{O}$ at pH 6.3 with 5 mM acetate buffer.

Table 3 Stern–Volmer constants (K_{SV}) and bimolecular luminescence quenching constants (k_q) for $^*\text{[Ru}(\text{bpy})_2\text{pyimH}]^{2+}$ in presence of MQ^+ in deaerated 1:1 (v:v) $\text{CH}_3\text{CN}/\text{H}_2\text{O}$ at different pH values in absence and presence of buffer

Conditions	$K_{\text{SV}} [\text{L mol}^{-1}]$	$k_q [10^8 \text{ L mol}^{-1} \text{ s}^{-1}]$	Process
Buffered pH 2	143 ± 13	6.8 ± 0.1	ET
Unbuffered pH 13	124 ± 12	20.7 ± 0.3	ET
Buffered pH 6.3	34 ± 2	5.7 ± 0.3	PCET
Unbuffered H_2O pH 7.3	70 ± 5	3.4 ± 0.5	PCET
Unbuffered D_2O pD 7.3	58 ± 6	2.5 ± 0.8	PCET

lifetime under these conditions is $(70 \pm 7) \text{ ns}$. Biexponential fits to the transients at 387 nm and 610 nm yield time constants of $(70 \pm 7) \text{ ns}$ and $(250 \pm 25) \text{ ns}$ with different signs of amplitude, corresponding to the $^3\text{MLCT}$ lifetime and the time constant for formation of HMQ^{*+} , respectively. Thus, photoexcited $[\text{Ru}^{\text{II}}(\text{bpy})_2\text{pyimH}]^{2+}$ disappears more rapidly than HMQ^{*+} forms, and this can be explained by rapid PCET (green arrow in Scheme 3) followed by slow PT_2 (red arrow in Scheme 3). By increasing the buffer concentration the kinetics of these processes is accelerated as described in the ESI† (Fig. S12).

In absence of buffer the reaction kinetics are different. The absence of buffer base leaves MQ^+ and water molecules as potential proton acceptors. The luminescence quenching experiment performed with acetate buffer (Fig. S7, ESI†) clearly shows that buffer base is a far better proton acceptor than water vis-à-vis photoexcited $[\text{Ru}^{\text{II}}(\text{bpy})_2\text{pyimH}]^{2+}$. Transient absorption spectra recorded in 1:1 (v:v) $\text{CH}_3\text{CN}/\text{H}_2\text{O}$ in absence of buffer are shown in Fig. 5d. The spectrum recorded without time delay (orange trace in Fig. 5d) is compatible with the formation of neutral MQ^{\bullet} radical as a primary photoproduct, exhibiting a characteristic broad absorption centered on 545 nm. After 5 μs the population of MQ^{\bullet} has decreased and HMQ^{*+} can be detected (brown trace in Fig. 5d). Thus, the overall photochemistry is the same as in presence of buffer (PCET followed by PT_2), but the kinetics are much different. The time constant for protonation of MQ^{\bullet} is $(2.5 \pm 0.3) \mu\text{s}$ based on the temporal evolution of the transient signal at 610 nm (ESI† Fig. S13a) which is a factor of 10 slower than in the presence of 5 mM acetate buffer. Stern–Volmer luminescence quenching studies yielded $k_q = (3.4 \pm 0.5) \times 10^8 \text{ L mol}^{-1} \text{ s}^{-1}$ for the initial excited-state quenching process. When going from $\text{CH}_3\text{CN}/\text{H}_2\text{O}$ to $\text{CH}_3\text{CN}/\text{D}_2\text{O}$, the protonation of MQ^{\bullet} (PT_2) is slowed down by a factor of 2.6, and the initial excited state quenching process becomes a factor of (1.4 ± 0.6) slower (Table 3). In principle, an H/D kinetic isotope effect of 1.4 would be compatible with concerted electron–proton transfer (CPET), but definitive assignment of the PCET step in Scheme 3 (green arrow) either to concerted or consecutive electron and proton transfer steps is currently not possible.

3 Summary and conclusion

The thermodynamic properties of $[\text{Ru}(\text{bpy})_2\text{pyimH}]^{2+}$ in its electronic ground state and in the long-lived $^3\text{MLCT}$ excited state in 1:1 (v:v) $\text{CH}_3\text{CN}/\text{H}_2\text{O}$ were explored in detail. In the



ground state the formal BDFE of the N–H bond is in the range of primary and secondary amines ((91 ± 1) kcal mol⁻¹),^{2,47,48} but upon excitation with a visible photon the BDFE drops by roughly 50 kcal mol⁻¹ to only (43 ± 5) kcal mol⁻¹. Thus, photoexcitation leads to a formal N–H BDFE in the range of metal hydride complexes which are used as hydrogenation catalysts in their electronic ground states.^{4,38} Transient absorption spectroscopy demonstrates that photoexcited [Ru(bpy)₂pyimH]²⁺ and (*N*-methyl-4,4'-bipyridinium, MQ⁺) undergo a formal HAT reaction, thereby confirming the finding of a very low formal N–H BDFE in the ruthenium complex; in HMQ⁺ the N–H BDFE is (53 ± 1) kcal mol⁻¹. Mechanistically, formal HAT between these two reactants is found to proceed *via* a sequence of PCET and proton transfer reaction steps involving buffer or solvent molecules. More generally, our study demonstrates that aromatic imines can be reduced from the excited state of [Ru(bpy)₂pyimH]²⁺.

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Notes and references

- D. R. Weinberg, C. J. Gagliardi, J. F. Hull, C. F. Murphy, C. A. Kent, B. C. Westlake, A. Paul, D. H. Ess, D. Granville and T. J. Meyer, *Chem. Rev.*, 2012, **112**, 4016–4093.
- J. J. Warren, T. A. Tronic and J. M. Mayer, *Chem. Rev.*, 2010, **110**, 6961–7001.
- S. Hammes-Schiffer, *J. Am. Chem. Soc.*, 2015, **137**, 8860–8871.
- M. Bourrez, R. Steinmetz, S. Ott, F. Gloaguen and L. Hammarström, *Nat. Chem.*, 2015, **7**, 140–145.
- J. M. Mayer, D. A. Hrovat, J. L. Thomas and W. T. Borden, *J. Am. Chem. Soc.*, 2002, **124**, 11142–11147.
- T. Irebo, S. Y. Reece, M. Sjödin, D. G. Nocera and L. Hammarström, *J. Am. Chem. Soc.*, 2007, **129**, 15462–15464.
- T. Irebo, M.-T. Zhang, T. F. Markle, A. M. Scott and L. Hammarström, *J. Am. Chem. Soc.*, 2012, **134**, 16247–16254.
- V. W. Manner, A. D. Lindsay, E. A. Mader, J. N. Harvey and J. M. Mayer, *Chem. Sci.*, 2012, **3**, 230.
- A. Wu, J. Masland, R. D. Swartz, W. Kaminsky and J. M. Mayer, *Inorg. Chem.*, 2007, **46**, 11190–11201.
- J. P. Roth, J. C. Yoder, T. Won and J. M. Mayer, *Science*, 2001, **294**, 2524–2526.
- J. P. Roth and J. M. Mayer, *Inorg. Chem.*, 1999, **38**, 2760–2761.
- O. S. Wenger, *Coord. Chem. Rev.*, 2015, **46**, 150–158.
- J. Nomrowski and O. S. Wenger, *Inorg. Chem.*, 2015, **54**, 3680–3687.
- J. C. Freys, G. C. Bernardinelli and O. S. Wenger, *Chem. Commun.*, 2008, 4267–4269.
- T. T. Eisenhart and J. L. Dempsey, *J. Am. Chem. Soc.*, 2014, **136**, 12221–12224.
- H. G. Yayla and R. Knowles, *Synlett*, 2014, 2819–2826.
- K. T. Tarantino, P. Liu and R. R. Knowles, *J. Am. Chem. Soc.*, 2013, **135**, 10022–10025.
- J. J. Concepcion, M. K. Brennaman, J. R. Deyton, N. V. Lebedeva, M. D. E. Forbes, J. M. Papanikolas and T. J. Meyer, *J. Am. Chem. Soc.*, 2007, **129**, 6968–6969.
- N. V. Lebedeva, R. D. Schmidt, J. J. Concepcion, M. K. Brennaman, I. N. Stanton, M. J. Therien, T. J. Meyer and M. D. E. Forbes, *J. Phys. Chem. A*, 2011, **115**, 3346–3356.
- Y. Deng, J. A. Roberts, S.-M. Peng, C. K. Chang and D. G. Nocera, *Angew. Chem., Int. Ed.*, 1997, **36**, 2124–2127.
- J. P. Kirby, J. A. Roberts and D. G. Nocera, *J. Am. Chem. Soc.*, 1997, **119**, 9230–9236.
- J. Rosenthal, J. M. Hodgkiss, E. R. Young and D. G. Nocera, *J. Am. Chem. Soc.*, 2006, **128**, 10474–10483.
- E. R. Young, J. Rosenthal, J. M. Hodgkiss and D. G. Nocera, *J. Am. Chem. Soc.*, 2009, **131**, 7678–7684.
- C. Bronner and O. S. Wenger, *J. Phys. Chem. Lett.*, 2012, **3**, 70–74.
- C. Bronner and O. S. Wenger, *Inorg. Chem.*, 2012, **51**, 8275–8283.
- S. M. Barrett, C. L. Pitman, A. G. Walden and A. J. M. Miller, *J. Am. Chem. Soc.*, 2014, **136**, 14718–14721.
- J. F. Ireland and P. A. H. Wyatt, *Adv. Phys. Org. Chem.*, Elsevier, 1976, vol. 12, pp. 131–221.
- C. R. Waidmann, A. J. M. Miller, C.-W. A. Ng, M. L. Scheuermann, T. R. Porter, T. A. Tronic and J. M. Mayer, *Energy Environ. Sci.*, 2012, **5**, 7771–7780.
- M.-A. Haga, *Inorg. Chim. Acta*, 1983, **75**, 29–35.
- M.-A. Haga and A. Tsunemitsu, *Inorg. Chim. Acta*, 1989, **164**, 137–142.
- A. M. Bond and M. Haga, *Inorg. Chem.*, 1986, **25**, 4507–4514.
- M. Haga, M. M. Ali, S. Koseki, K. Fujimoto, A. Yoshimura, K. Nozaki, T. Ohno, K. Nakajima and D. J. Stufkens, *Inorg. Chem.*, 1996, **35**, 3335–3347.
- M.-A. Haga, *Inorg. Chim. Acta*, 1980, **45**, L183–L184.
- M. Haga, T. Ano, T. Ishizaki, K. Kano, K. Nozaki and T. Ohno, *J. Chem. Soc., Dalton Trans.*, 1994, 263.
- X. Xiaoming, M. Haga, T. Matsumura-Inoue, Y. Ru, A. W. Addison and K. Kano, *J. Chem. Soc., Dalton Trans.*, 1993, 2477.
- K. M. Lancaster, J. B. Gerken, A. C. Durrell, J. H. Palmer and H. B. Gray, *Coord. Chem. Rev.*, 2010, **254**, 1803–1811.
- A. Wu and J. M. Mayer, *J. Am. Chem. Soc.*, 2008, **130**, 14745–14754.
- J. Choi, M. E. Pulling, D. M. Smith and J. R. Norton, *J. Am. Chem. Soc.*, 2008, **130**, 4250–4252.
- A. J. Morris, R. T. McGibbon and A. B. Bocarsly, *ChemSusChem*, 2011, **4**, 191–196.
- J. Chen, K. Wu, B. Rudshteyn, Y. Jia, W. Ding, Z.-X. Xie, V. S. Batista and T. Lian, *J. Am. Chem. Soc.*, 2016, **138**, 884–892.
- D. J. Boston, C. Xu, D. W. Armstrong and F. M. MacDonnell, *J. Am. Chem. Soc.*, 2013, **135**, 16252–16255.
- A. A. Vlcek, E. S. Dodsworth, W. J. Pietro and A. B. P. Lever, *Inorg. Chem.*, 1995, **34**, 1906–1913.
- A. Yoshimura, M. Z. Hoffman and H. Sun, *J. Photochem. Photobiol., A*, 1993, **70**, 29–33.
- L. M. Tolbert and K. M. Solntsev, *Acc. Chem. Res.*, 2002, **35**, 19–27.



- 45 J. W. Park, Y. Kim and J. M. Kim, *Bull. Korean Chem. Soc.*, 1993, 174–175.
- 46 A. Harriman, G. R. Millward, P. Neta, M. C. Richoux and J. M. Thomas, *J. Phys. Chem.*, 1988, **92**, 1286–1290.
- 47 F. G. Bordwell, X. Zhang and J. P. Cheng, *J. Org. Chem.*, 1991, **56**, 3216–3219.
- 48 F. G. Bordwell, X. M. Zhang and J. P. Cheng, *J. Org. Chem.*, 1993, **58**, 6410–6416.
- 49 D. P. Rillema, R. Sahai, P. Matthews, A. K. Edwards, R. J. Shaver and L. Morgan, *Inorg. Chem.*, 1990, **29**, 167.
- 50 C. J. Gagliardi, B. C. Westlake, C. A. Kent, J. J. Paul, J. M. Papanikolas and T. J. Meyer, *Coord. Chem. Rev.*, 2010, **254**, 2459–2471.
- 51 P. S. Braterman and J. I. Song, *J. Org. Chem.*, 1991, **56**, 4678–4682.
- 52 D. M. Roundhill, *Photochemistry and Photophysics of Metal Complexes*, Springer US, Boston, MA, 1994.
- 53 A. Juris, V. Balzani, F. Barigelletti, S. Campagna, P. Belser and A. von Zelewsky, *Coord. Chem. Rev.*, 1988, **84**, 85–277.

