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Path-dependent variational effects and multidimensional tunneling in multi-path variational transition state theory: rate constants calculated for the reactions of HO₂ with *tert*-butanol by including all 46 paths for abstraction at C and all six paths for abstraction at O†

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Multi-path variational transition state theory (MP-VTST) provides a conformationally complete framework for calculating gas-phase rate constants. For reactions in which the transition state has distinguishable torsional minima (which include most reactions), there are multiple possible reaction paths. In principle MP-VTST includes the contributions from all the reaction paths, and it should explicitly treat the variational and tunneling effects of each path, but in practice one may need to truncate the number of paths included in MP-VTST calculations in order to achieve a balance between computational cost and accuracy. In this work, we present calculations including all paths for two prototype combustion reactions, namely the two hydrogen abstraction reactions from *tert*-butanol by HO₂ radical. For both reactions we included all the reaction paths. Since abstraction at C has 46 paths, it provided a good opportunity to carry out a case study in which we investigated the errors introduced by truncating the number of paths. For the reaction studied, we found that the variational and multidimensional tunneling transmission coefficients are very different for different reaction paths, which provides new evidence that MP-VTST is necessary for treating path-dependent variational effects and multidimensional tunneling. We found that tunneling transmission coefficients can be much larger for higher-energy paths than for lower-energy ones. Interestingly, the simple hypothesis that higher barriers are narrower does not explain this finding in the present case; we found instead that the effect is due to higher-energy barriers having the possibility of tunneling at energies farther below the barrier top. We also show that a previously applied criterion for judging convergence with respect to the number of paths may not be reliable at low temperature.

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1. Introduction

The recently developed multi-path variational transition state theory^{1,2} (MP-VTST) provides a framework for taking account of all reactant and transition state conformations and all conformationally different reaction paths for calculating thermal rate constants and predicting experimentally unavailable branching ratios for complex reactants. It should be especially

useful in combustion chemistry, atmospheric chemistry, and all problems involving reactions of chain molecules. MP-VTST is an extension of multi-structural variational transition state theory³ (MS-VTST), which has now been used to study the chemical kinetics of a variety of chemically important systems, including hydrogen abstraction reactions in combustion of biofuels^{4–8} and silyl anion–silane polymerization reactions in nanodusty plasma chemistry.⁹

The internal rotations (torsions) of a molecule can lead to multiple conformational structures, *i.e.*, multiple local minima (for reactants and products) or multiple first-order saddle points (for transition structures) on the potential energy surface, and therefore multiple reaction paths. The contributions from all the reaction paths that correspond to the same chemical reaction are included in the MP-VTST calculations, and the variational effect and quantum tunneling contribution of each

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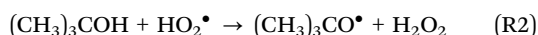
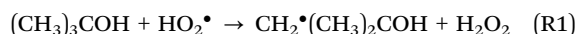
† Electronic supplementary information (ESI) available: Geometries optimized at M08-HX/MG3S level of all the conformers; classical barrier heights and energy of reaction computed by various methods; $I_{\text{act}}^{\text{MS-T}}$ factors; MP-VTST rate constants; fitting parameters for rate constants; activation energies and Gibbs free energies of activation. See DOI: 10.1039/c5cp05780a



path is treated explicitly in MP-VTST. In principle, all the reaction paths should be included in the calculations, and one should perform variational transition state theory calculations with multi-dimensional tunneling calculations for all the paths. Such a calculation can be based on minimum-energy paths (MEPs) generated by direct dynamics. Practically, however, the computational cost for such calculations, which involve calculating the energies, gradients, and generalized normal mode analyses at large numbers of points along each reaction coordinate, can be demanding. Although density functional theory can greatly accelerate such calculations as compared to reliable wave-function-based calculations, it is often unaffordable to accurately calculate all the reaction paths, which are generated from multiple conformers of the transition state, even for a medium-sized transition state (for instance, the transition state for hydrogen abstraction reaction from C-3 of 1-butanol by HO₂ radical consists of 262 distinguishable conformers¹⁰).

In our previous work,^{4,8} the number of paths included in MP-VTST was decided by ordering the paths in order of increasing barrier height and comparing calculations averaged over the first P paths and averaged over the first $P - 1$ paths (or P and $P - 2$ paths if the paths come in optically active pairs), and if the computed final rate constants have no significant differences (for instance, if the relative difference is smaller than 15%), then we stopped including more reaction paths. For a given reaction, it would be tempting to assume that various reaction paths are similar to each other except for the quantitative energetics, in which case this would be a reasonably reliable method. However, this is a dangerous assumption. In the current work, we will confirm that this kind of agreement does not necessarily mean that the MP-VTST rate constant has been converged using the first P paths because the transmission coefficients can be very different for different paths.

Hydrogen abstractions by small radicals (such as H, O, OH, HO₂, and CH₃) are a class of important reactions in the combustion mechanisms of common fuels.^{11,12} In Sarathy and co-workers' work,¹¹ rate constants for HO₂ + *tert*-butanol H-abstraction reaction are theoretically computed based on Evans-Polanyi approach, which not only ignores the temperature-dependence of the activation energy but also assumes a perfect correlation between bond energy and barrier height. *Ab initio* variational transition state calculations are needed in order to more accurately predict the thermal rate constants. In this work, we choose the *tert*-butanol + HO₂ reaction to study the effect of truncating the number of paths included in MP-VTST calculations, and we use the MP-VTST rate constants calculated based on all reaction paths as a reference. To be more specific, the reactions we consider here are:



Reaction (R1) has 46 reaction paths (23 pairs of mirror images), and (R2) has 6 paths (3 pairs of mirror images). (R1) is the reaction of major interest in the current work since it provides a

large number of reaction paths and also because it has a larger rate constant than reaction (R2).

2. Theory

The thermal rate constant in multi-path variational transition state theory is given by^{1,2}

$$k^{\text{MP-VTST}} = F_{\text{act}}^{\text{MS-T}} \langle \gamma \rangle_P k_1^{\text{ConTST}} \quad (1)$$

where k_1^{ConTST} is the rate constant computed by conventional transition state theory (ConTST) without tunneling using the lowest-energy transition structure, $F_{\text{act}}^{\text{MS-T}}$ is the multi-structural torsional anharmonicity factor of activation, which is defined as:

$$F_{\text{act}}^{\text{MS-T}} = \frac{Q_{\text{MS-ConTS}}^{\text{MS-T}} Q_{\text{R1,GM}}^{\text{SS-QH}} Q_{\text{R2,GM}}^{\text{SS-QH}}}{Q_{\text{MS-R1}}^{\text{MS-T}} Q_{\text{MS-R2}}^{\text{MS-T}} Q_{\text{ConTS,GM}}^{\text{SS-QH}}} \quad (2)$$

where MS-R1, MS-R2, and MS-ConTS denote the multi-structural reactant 1, the multi-structural reactant 2, and the multi-structural conventional transition state, and GM denotes "global minimum"; $Q^{\text{MS-T}}$ is the multi-structure (MS) conformational-rovibrational partition function with torsional anharmonicity (T) computed based on a coupled torsional potential;¹³ and $Q^{\text{SS-QH}}$ is the rovibrational partition function computed using the rigid-rotator-quasi-harmonic-oscillator approximation based on a global minimum structure. When we say GM for a transition state we mean the lowest-energy first-order saddle point that connects reactants to products, *i.e.*, the lowest-energy transition structure. Note also that a "transition state" is a dividing surface that passes through all the transition state structures, and the MS-ConTS is a dividing surface that passes through all the saddle points, which are also called transition structures. In a full calculation, $F_{\text{act}}^{\text{MS-T}}$ includes the contributions from all the conformational structures, *i.e.*, there is no truncation error introduced by limiting the sums of structures.

The variational effect and multidimensional tunneling (MT) contributions of each reaction path are explicitly treated using the path-averaged generalized transmission coefficient:

$$\langle \gamma \rangle_P = \frac{\sum_{p=1}^P k_p^{\text{MT}} \Gamma_p^{\text{CVT}} Q_p^{\ddagger\text{-SS-T}}}{\sum_{p=1}^P Q_p^{\ddagger\text{-SS-T}}} \quad (3)$$

where the subscript p denotes the index of a reaction path; P denotes the total number of reaction paths included in MP-VTST; $Q_p^{\ddagger\text{-SS-T}}$ is the single-structural rovibrational partition function with torsional anharmonicity of the transition structure of path p ; Γ_p^{CVT} is the variational recrossing transmission coefficient computed for path p ; and κ_p^{MT} is the multidimensional tunneling transmission coefficient of path p . The factor κ_p^{MT} is equal to the ratio $k_p^{\text{CVT/MT}}/k_p^{\text{CVT}}$, where CVT/MT refers to the CVT rate constant including multidimensional tunneling. We should emphasize that the multidimensional tunneling transmission coefficient not only includes the effect of quantum mechanical



tunneling (which increases the rate constant, especially at lower temperatures and for light particles), but also includes the contribution from nonclassical reflection (which decreases the rate constant) at energies above the barrier height. Thus a better technical name for κ_p^{MT} would be the transmission coefficient that accounts for quantum mechanical effects on reaction coordinate motion, including its coupling to other degrees of freedom, but we call it tunneling transmission coefficient for brevity.

The multidimensional tunneling transmission coefficient is evaluated with the small-curvature tunneling approximation¹⁴ (SCT) in the current work, but a few results are shown for the zero-curvature tunneling (ZCT) approximation^{15–17} for comparison. In the ZCT approximation, the system tunnels through the vibrationally adiabatic ground-state potential curve defined for path p by

$$V_{a,p}^{\text{G}}(s) = V_{\text{MEP},p}(s) + \varepsilon_{0,p}(s) \quad (4)$$

where s is the signed distance (from the saddle point) along the MEP through mass-scaled coordinates with reduced mass μ , $V_{\text{MEP},p}(s)$ is the potential energy along MEP p , and $\varepsilon_{0,p}(s)$ is the local zero-point vibrational energy of modes transverse to this MEP. The vibrationally adiabatic ground-state potential curve is sometimes called the effective barrier for tunneling. In the SCT approximation, the tunneling is again governed by $V_{a,p}^{\text{G}}(s)$, but $\sqrt{\mu}ds$ in the imaginary action integral (see next equation) is reduced to account for the shortening of the tunneling path by corner cutting in the small-curvature limit.^{14,17} In practice, this reduction is accomplished by replacing μ by $\mu_{\text{eff},p}(s) < \mu$, where $\mu_{\text{eff},p}(s)$ depends on the reaction-path curvature at s . Then the SCT tunneling probability at energy E is

$$\text{Prob}_p^{\text{SCT}}(E) = \left[1 + \exp\left(\frac{2}{\hbar} \int_{s_<}^{s_>} \sqrt{2\mu_{\text{eff},p}(s) [V_{a,p}^{\text{G}}(s) - E]} ds\right) \right]^{-1} \quad (5)$$

where \hbar is Planck's constant divided by 2π , $s_<$ is the turning point where $V_{a,p}^{\text{G}}(s) = E$ on the left side of the maximum, and $s_>$ is the turning point where $V_{a,p}^{\text{G}}(s) = E$ on the right side of the maximum.

In eqn (3), P is formally equal to the total number of distinguishable reaction paths (including paths that are mirror images), which is equal to the number N of distinguishable structures of the transition state (TS). Practically, due to the existence of a large number of reaction paths, we often have to limit the number of paths included in MP-VTST calculation by taking $P < N$. Notice that, if $P = N$, the denominator of eqn (3) is equal to $Q_{\text{TS}}^{\text{MS-T}}$; and if $P < N$, and this introduces truncation errors. The possible sources of errors introduced in the final calculated rate constant are complex to analyze. In the present work, we focus on the errors introduced by truncating the number of reaction paths in MP-VTST calculations, regardless of other factors such as the accuracy of electronic structure calculations or the accuracy with which anharmonicity is treated.

3. Computational details

An exhaustive conformational search was carried out for reactant, product, and transition structures by using M08-HX¹⁸/MG3S¹⁹ for the electronic structure and MSTor²⁰ for the vibrational anharmonicity and vibration-rotation coupling. In order to choose a reasonably accurate electronic structure method for carrying out VTST calculations, classical barrier heights for both forward and reverse reactions (V_{\ddagger}^{\ddagger} and V_{\ddagger}^{\ddagger}), and energy of reaction (ΔE) were computed based on the lowest-energy structures. (Classical barrier heights and energies of reaction are difference in Born–Oppenheimer energies for stationary structures; they do not include vibrational zero-point energy or thermal energy.) Four hybrid density functionals, M05,²¹ M06-2X,^{22,23} M08-HX,¹⁸ and M08-SO,¹⁸ were used, combined with MG3S¹⁹ and ma-TZVP²⁴ basis sets. The explicitly correlated coupled cluster theory CCSD(T)-F12a^{25–27} with the jun-cc-pVTZ basis²⁸ was chosen as the benchmark. All the DFT calculations were performed with integration grids of 99 radial shells and 974 angular points per shell using a locally modified *Gaussian 09*^{29,30} program; coupled cluster calculations were carried out with *Molpro2010*³¹ software.

Based on the benchmark results discussed below, M08-HX/MG3S was used for VTST calculations. Canonical variational transition state theory calculations were carried out in non-redundant internal coordinates^{32,33} by using the *Gaussrate*³⁴ and *Polyrate*^{35,36} programs. Multidimensional tunneling transmission coefficients were computed based on the small-curvature tunneling approximation and reaction paths extending from $s = -2.0$ to $+2.0$ bohr, where s is the reaction coordinate taken as the distance from the saddle point along the isoinertial minimum-energy path scaled to a reduced mass of 1 amu. In calculating the rovibrational partition functions and performing generalized normal mode analysis, all the vibrational frequencies are scaled with a factor of 0.973.³⁷

4. Results and discussion

Benchmarking density functionals

We will judge the accuracy of density functionals for the present reactions by estimating errors as the differences of their predictions of classical barrier heights and classical energies of reaction at stationary-point geometries optimized by M08-HX/MG3S to the results obtained by CCSD(T)-F12a/jun-cc-pVTZ with the same geometries. A classical barrier height is the difference in potential energy between the lowest-energy saddle point (of the transition state) and the lowest-energy structure of reactants (forward barrier height) or the products (reverse barrier height); it is the main determinant of the magnitude of the rate constant. A classical energy of reaction is the difference in potential energy between the lowest-energy structure of the products and the lowest-energy structure of the reactants; it is the main component of the free energy of reaction, which is related to the chemical equilibrium constant, that connects the forward and reverse rate constants according to the principle of detailed balance. Classical barrier heights and energies of reaction computed by various density functionals



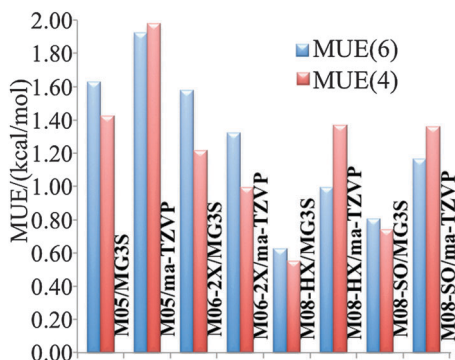


Fig. 1 Mean unsigned errors (MUEs) in kcal mol⁻¹ computed based on M08-HX/MG3S geometries using various methods, with respect to CCSD(T)-F12a/jun-cc-pVTZ as benchmark.

for reactions (R1) and (R2) are tabulated in Table S1 (tables labeled with “S” are in the ESI[†]). Forward barrier height V_{\ddagger}^{\ddagger} , reverse barrier height V_{\ddagger}^{\ddagger} , and energy of reaction ΔE are not independent because they are connected by the relation $V_{\ddagger}^{\ddagger} - V_{\ddagger}^{\ddagger} = \Delta E$. One may argue that for some purposes it is reasonable to compute mean unsigned errors (MUEs) based only on forward barrier heights and energies of reaction, rather than based on all three of them, whereas for other purposes it is appropriate to consider all three. Therefore we report two kinds of MUEs in Fig. 1: MUE(6) includes both forward barrier heights, both reverse barrier heights, and both energies of reaction; MUE(4) only includes forward barriers and energies of reaction. MUE(6) and MUE(4) differs less than 0.4 kcal mol⁻¹ in all the cases, and the use of either error criterion leads to similar conclusions about the relative accuracies of the method. The numerical results for the three most accurate of the eight methods tested are shown in Table 1. Among the model chemistries we tested, M08-HX/MG3S has the smallest MUEs (0.6 kcal mol⁻¹) and therefore is chosen for rate constant calculations.

Multi-structural anharmonicity factors, rate constants, and branching fractions

The $F_{\text{act}}^{\text{MS-T}}$ factors computed by M08-HX/MG3S are given for 12 temperatures in Table S2 (ESI[†]) and for six temperatures in Table 2. Computed MP-CVT/SCT rate constants for the forward and reverse reactions of (R1) and (R2) are given at 12 temperatures in Table S3 (ESI[†]) and for six temperatures in Table 3. The rate constants are fitted for 200–2400 K by the four-parameter functions proposed previously.^{2,38} The final values of the

Table 1 Classical barrier height for forward and reverse reaction and classical energy of reaction (kcal mol⁻¹) computed by various model chemistries. Single-point calculations were carried out based on the lowest-energy-structures optimized at M08-HX/MG3S level of theory

Model chemistries	R1			R2		
	V_{\ddagger}^{\ddagger}	V_{\ddagger}^{\ddagger}	ΔE	V_{\ddagger}^{\ddagger}	V_{\ddagger}^{\ddagger}	ΔE
M06-2X/ma-TZVP	19.9	2.5	17.3	21.2	1.0	20.2
M08-HX/MG3S	20.6	3.5	17.1	23.1	2.5	20.5
M08-SO/MG3S	21.0	4.2	16.7	22.5	1.8	20.7
CCSD(T)-F12a/jun-cc-pVTZ	20.2	4.1	16.2	23.4	3.4	20.0

Table 2 $F_{\text{act}}^{\text{MS-T}}$ factors computed by M08-HX/MG3S at various temperatures for (R1) and (R2) in the forward and reverse directions

T/K	R1		R2	
	$F_{\text{act, fwd}}^{\text{MS-T}}$	$F_{\text{act, rev}}^{\text{MS-T}}$	$F_{\text{act, fwd}}^{\text{MS-T}}$	$F_{\text{act, rev}}^{\text{MS-T}}$
200	1.8	0.43	1.8	0.95
298.15	2.5	0.47	1.9	1.00
600	6.0	0.96	2.3	1.1
1000	10.9	1.6	2.8	1.3
1500	15.4	2.4	3.0	1.4
2400	19.7	3.4	3.0	1.3

Table 3 MP-CVT/SCT rate constants (cm³ molecule⁻¹ s⁻¹) including all the paths for forward and reverse reactions of (R1) and (R2) at selected temperatures (K)

T/K	R1		R2	
	$k_{\text{fwd}}^{\text{MP-CVT/SCT}}$	$k_{\text{rev}}^{\text{MP-CVT/SCT}}$	$k_{\text{fwd}}^{\text{MP-CVT/SCT}}$	$k_{\text{rev}}^{\text{MP-CVT/SCT}}$
200	4.85×10^{-33}	1.11×10^{-16}	1.02×10^{-36}	2.50×10^{-15}
298.15	4.87×10^{-27}	1.34×10^{-16}	2.86×10^{-29}	3.88×10^{-15}
600	3.55×10^{-20}	5.48×10^{-16}	2.76×10^{-21}	1.17×10^{-14}
1000	6.97×10^{-17}	3.03×10^{-15}	1.06×10^{-17}	4.44×10^{-14}
1500	6.00×10^{-15}	1.41×10^{-14}	1.13×10^{-15}	1.50×10^{-13}
2400	2.96×10^{-13}	8.37×10^{-14}	6.38×10^{-14}	6.31×10^{-13}

forward rate constants (in units of cm³ molecule⁻¹ s⁻¹) are given by the fits as

$$k_{\text{R1, fwd}}^{\text{MP-CVT/SCT}} = 2.47 \times 10^{-17} \left(\frac{T}{300} \right)^{5.776} \times \exp \left[-\frac{11.120(T + 88.432)}{R(T^2 + 7820.238)} \right] \quad (6)$$

$$k_{\text{R2, fwd}}^{\text{MP-CVT/SCT}} = 8.04 \times 10^{-16} \left(\frac{T}{300} \right)^{3.950} \times \exp \left[-\frac{18.051(T + 4.954)}{R(T^2 + 24.539)} \right] \quad (7)$$

$$k_{\text{overall, fwd}}^{\text{MP-CVT/SCT}} = 1.09 \times 10^{-15} \left(\frac{T}{300} \right)^{4.440} \times \exp \left[-\frac{16.139(T - 11.694)}{R(T^2 + 136.747)} \right] \quad (8)$$

where R is the gas constant in kcal mol⁻¹, and T is temperature. The complete set of fitting parameters for both forward and reverse reactions are in Table S4 (ESI[†]).

Branching fractions into the two possible products are shown in Fig. 2. We see that the yield of product R2 increases from a very small number at low temperature to about 20% at high temperature.

Activation energies and phenomenological Gibbs free energies of activation

Fig. 3 shows temperature-dependent Tolman activation energies, defined by³⁹

$$E_a = -RT \frac{d \ln k}{d(1/T)} \quad (9)$$



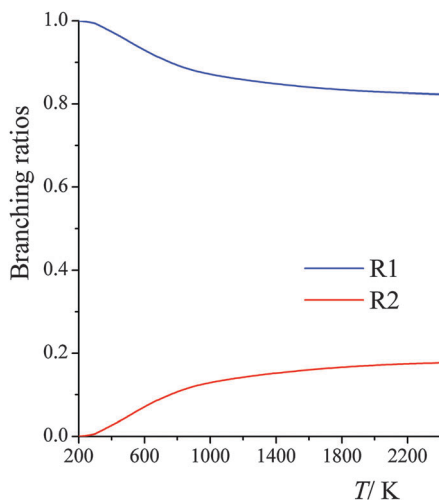


Fig. 2 Branching fractions at various temperatures for forward reactions of (R1) and (R2).

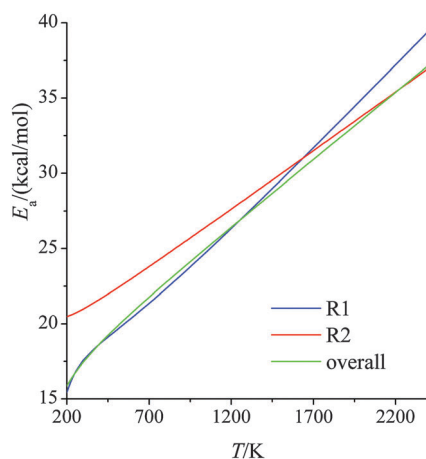


Fig. 3 Activation energies (kcal mol^{-1}) for forward reactions (R1) and (R2) and the overall reaction at various temperatures.

The derivatives were computed analytically from the fits (see Table S4, ESI[†]). We see that, over the temperature range studied, the activation energy increases by $\sim 25 \text{ kcal mol}^{-1}$ for reaction (R1) and by $\sim 15 \text{ kcal mol}^{-1}$ for reaction (R2). This temperature-dependence of activation energy is significant and should not be ignored in theoretical combustion modeling. Since reaction (R1) dominates the overall reaction, the results for the overall reaction are very close to those for reaction (R1).

Fig. S1 (ESI[†]) shows phenomenological Gibbs free energies of activation, which are given for bimolecular gas-phase reactions by^{40,41}

$$\Delta G_{\text{act, fwd}}^{\circ} = RT \ln \left(\frac{(k_{\text{B}} T)^2}{h k_{\text{fwd}}^{\text{MP-CVT/SCT}} p^0} \right) \quad (10)$$

where k_{B} and h are the Boltzmann and Planck constants, $k_{\text{fwd}}^{\text{MP-CVT/SCT}}$ is in the units of $\text{m}^3 \text{ molecule}^{-1} \text{ s}^{-1}$, and p^0 is the

standard-state pressure (10^5 Pa). As for the activation energy, the result is a rapidly increasing function of temperature.

Path-dependent variational effects

The variational transition states are slightly away from the saddle points. For example, numbering the paths in order of increasing classical barrier height, we find that at 200 K, the canonical variational generalized transition state is 0.002, 0.012, 0.009 Å away from the saddle point for paths 1, 41, and 45, respectively; at 600 K, they are 0.003, 0.023, 0.013 Å away from the saddle points for these paths; and at 2400 K, they are 0.008, 0.081, and 0.039 Å away from the saddle points for these paths.

Fig. 4 shows the computed CVT variational recrossing transmission coefficients of reaction (R1) for temperatures from 200 K to 2400 K for paths 1, 7, and 41 (the results for paths 2, 8, and 42 are the same as for 1, 7, and 41 because the paths come in enantiomeric pairs). Over the entire temperature range, the variational recrossing effect is negligible for path 1, for which the variational recrossing transmission coefficient varies from 0.998 at 200 K to 0.996 at 2400 K. For reaction path 7, which emanates from a transition structure $0.97 \text{ kcal mol}^{-1}$ higher than that generating path 1, the behavior is significantly different from that for path 1. Variational recrossing transmission coefficients of path 7 increase from 0.913 at 200 K to a maximum 0.941 at 600 K and then gradually decrease to 0.930 at 1500 K, followed by a rapid decrease to 0.750 at 2400 K. For reaction path 41, there is an almost linear decrease of the variational recrossing transmission coefficient with temperature from 250 to 2400 K. At 2400 K, the variational transmission coefficient for the lowest-energy path is larger than the one for path 41 by a factor of 1.4.

Multidimensional tunneling

Fig. 5 shows the SCT tunneling transmission coefficients for all the reaction paths of R1 at 200 K and 1000 K. At 200 K the

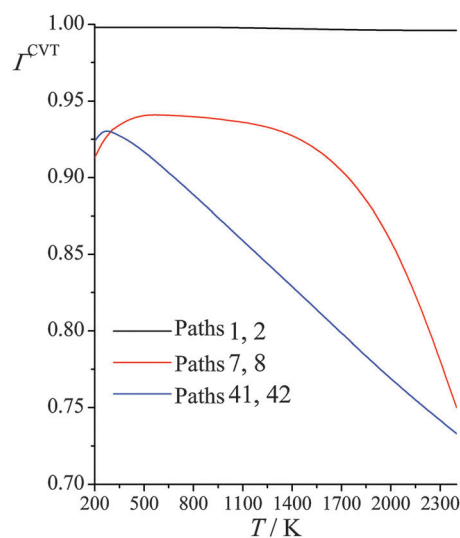


Fig. 4 Calculated CVT variational transmission coefficients at various temperatures (K) for different paths for reaction (R1): $(\text{CH}_3)_3\text{COH} + \text{HO}_2^{\bullet} \rightarrow \text{CH}_2^{\bullet}(\text{CH}_3)_2\text{COH} + \text{H}_2\text{O}_2$.



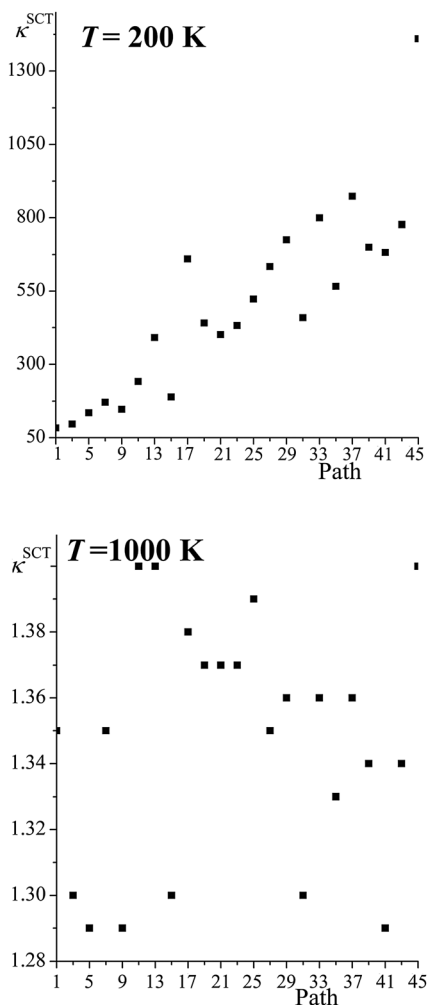


Fig. 5 SCT transmission coefficient of all the paths for reaction (R1) computed at 200 K and 1000 K. The abscissa is the ordinal number of the path. The paths are numbered 1–46 in order of increasing classical barrier height, but because the paths come in enantiomeric pairs, we only give the results for odd-numbered paths.

tunneling transmission coefficients increase dramatically as the barrier heights increase; the tunneling transmission coefficients of the highest-energy paths are very much larger than the ones of the lowest-energy path, with the largest ratio being 17 and several other ratios being in the 9–12 range. As expected, the tunneling transmission coefficients are much closer to unity at 1000 K, because the Boltzmann factor for overbarrier reaction is much larger, but perhaps not so expected is the finding that the tunneling transmission coefficients cluster close to one another – with an average value of 1.35 and a standard deviation of only 0.04.

We may now ask: why are the tunneling transmission coefficients larger for the paths with higher classical heights? The answer to this question will illuminate an interesting feature of the way that tunneling effects enter a multi-path calculation. In order to give this answer we need to consider some aspects of the tunneling in more detail, and we now turn to this.

Fig. 6 shows both the ZCT and SCT tunneling transmission probabilities of various paths of reaction (R1) as functions of tunneling energy; the figure includes the lowest- and highest-energy paths and some intermediate-energy paths. SCT tunneling transmission probabilities are greater than ZCT ones because reaction-path curvature increases the tunneling probability by allowing corner cutting. The zero of energy for the tunneling calculations is the potential energy of the equilibrium structure of reactants. For each path the semiclassical tunneling probability rises approximately exponentially with energy until it reaches 0.5 at the energy of the maximum of $V_{a,p}^G(s)$ for that path; this maximum is called $V_{a,p}^{AG}$, and it is usually higher for paths with a higher ordinal number, *i.e.*, for paths with a higher classical barrier height.

The $V_{a,p}^G(s)$ curves for reaction for path 1 (which, along with its mirror image, path 2, is the lowest-energy path), for path 41, and for path 45 (which, along with its mirror image, path 46, is the highest-energy path) are shown in Fig. 7, and the potential energies along the minimum energy paths (MEPs through the iso-inertial coordinate system, which are sometimes called IRCs (intrinsic reaction coordinates)) for these three reaction paths are shown in Fig. 8. Fig. 9 shows the forming O–H bond distance as a function of the breaking C–H bond distance along the MEPs for paths 1, 41, and 45. The MEPs are very similar for these two critical bond distances, but the shapes of the vibrationally adiabatic potential energy curves are different because of the variations of the zero-point vibrational energy along the reaction paths. We see that $V_{a,p}^G(s)$ peaks close to, but not precisely at $s = 0$, which, by convention, is placed at the saddle point, which is the maximum of $V_{\text{MEP}}(s)$.

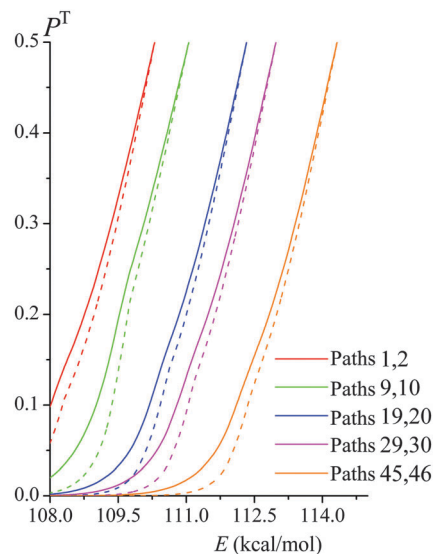


Fig. 6 SCT transmission probabilities $P^{\text{SCT}}(E)$ (solid lines) and ZCT transmission probabilities $P^{\text{ZCT}}(E)$ (dashed lines) in the tunneling regime of various paths of reaction (R1). Note that each conformer of the transition state generates its own reaction path, and so each curve in this figure corresponds to a different reaction path. They are labeled with two path numbers because transition state structures that are mirror images generate paths that are mirror images and therefore have the same transmission coefficient.



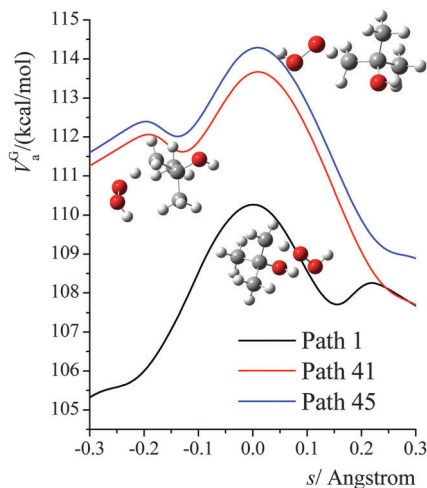


Fig. 7 Vibrationally adiabatic potential energy (in kcal mol⁻¹) along the minimum energy paths (MEP) for paths 1, 41, 45 of reaction (R1), where s is the signed distance from the saddle point (*i.e.*, transition structure) along each curved MEP in isoinertial coordinates scaled to a reduced mass of 1 amu. Note that each conformer of the transition state generates its own reaction path, and so each curve in this figure corresponds to a different reaction path, where the paths are shown in Fig. 9. The transition structures are shown next to the vibrationally adiabatic ground-state potential energy curve for the path they generate.

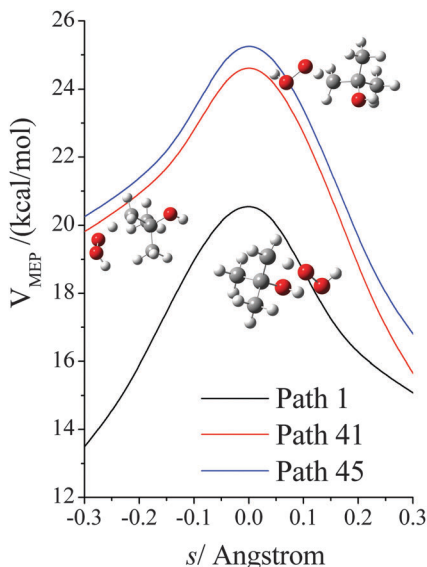


Fig. 8 Potential energy (in kcal mol⁻¹) along the minimum energy path (MEP) for paths 1, 41, 45 of reaction (R1), where s is the signed distance from the saddle point (*i.e.*, transition structure) along the curved MEP in isoinertial coordinates scaled to a reduced mass of 1 amu. Note that each conformer of the transition state generates its own reaction path, and so each curve in this figure corresponds to a different reaction path, where the paths are shown in Fig. 9. The transition structures are shown next to the vibrationally adiabatic ground-state potential energy curve for the path they generate.

The semiclassical tunneling transmission coefficient for path p is given by

$$\kappa_p = \frac{\int_{E_0}^{+\infty} \text{Prob}_p(E) \exp(-E/RT) dE}{\int_{E_0}^{+\infty} \text{Prob}_p^C(E) \exp(-E/RT) dE} \quad (11)$$

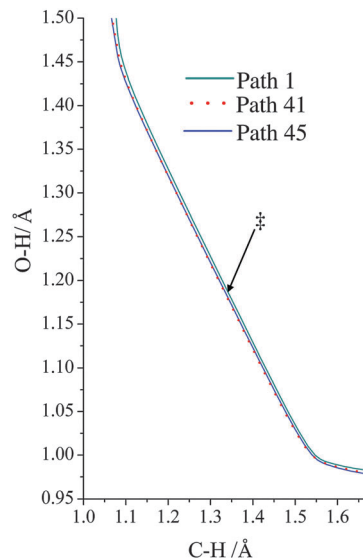


Fig. 9 Minimum energy path for paths 1, 41, and 45, where we show the breaking C-H bond distance and the forming H-O bond distance. Note that all three paths are shown, but as far as the forming and breaking bond distances are concerned, the paths are so similar as to be hard to distinguish on the plot. The locations of the transition structures along the paths are similar for the three paths and are indicated by ‡.

where E_0 is the maximum of the ground-state vibrational energy of reactants and the ground-state vibrational energy of products (this is the lowest energy at which reaction can occur), $\text{Prob}_p(E)$ is the transmission probability of eqn (5), and $\text{Prob}_p^C(E)$ is the classical transmission probability along path p at energy E . Since the classical transmission probability is a unit step function at $E = V_a^{\text{AG}}$, we can carry out the integral in the denominator, which yields

$$\kappa_p = \frac{1}{RT} \int_{E_0}^{+\infty} \text{Prob}_p(E) \exp\left[-\left(E - V_{a,p}^{\text{AG}}\right)/RT\right] dE \quad (12)$$

the combination of eqn (5) and (12) shows that the different κ_p for various paths may result from different effective reduced masses or different $V_{a,p}^{\text{G}}(s)$. However, since the higher-energy paths also have larger transmission coefficients in the ZCT approximation, and the reduced masses are all the same in the SCT approximation, we look to the shape of the $V_{a,p}^{\text{G}}(s)$ for an explanation.

The energy at which the integrand of eqn (12) peaks is called the representative tunneling energy, $E_{\text{rep},p}(T)$; some examples (in the SCT approximation) are shown in Table 4. Table 4 shows that representative tunneling energies are 2.6–5.2 kcal mol⁻¹ below the effective barrier maxima. A simple hypothesis might be that the higher-energy paths show larger tunneling transmission coefficients because the effective barriers are thinner near their tops. However, examination of the $V_{a,p}^{\text{G}}(s)$ curves shows that, in this energy range, these curves have about the same widths (see Table S7, ESI[†]). Consistent with this, Table 4 shows that the tunneling probabilities are about the same for all paths at energies 2 to 3 kcal mol⁻¹ below the maximum of the effective potential energy curve for tunneling. So this hypothesis



Table 4 SCT representative tunneling energies at 200 K, energies of the maxima of the vibrationally adiabatic ground-state potential curves, and SCT transmission probabilities at energies 2 and 3 kcal/mol lower than $V_{a,p}^{AG}$ for various paths of reaction (R1)

Path	$E_{rep,p}$ (200 K) (kcal mol ⁻¹)	$V_{a,p}^{AG}$ (kcal mol ⁻¹)	Prob _p ^{SCT a}	Prob _p ^{SCT b}
1	107.67	110.31	0.14	—
9	107.67	111.06	0.11	0.022
19	107.67	112.33	0.12	0.024
29	108.11	112.98	0.12	0.029
41	109.03	113.72	0.12	0.022
45	109.14	114.33	0.13	0.030

^a Evaluated at an energy 2 kcal mol⁻¹ below $V_{a,p}^G(s)$. ^b Evaluated at an energy 3 kcal mol⁻¹ below $V_{a,p}^G(s)$.

does not explain the observation of larger tunneling transmission coefficients for the higher-energy paths.

We plot the integrand of eqn (12) for the SCT tunneling approximation at 200 K in Fig. 10 for various paths of reaction (R1). The curves in this figure start at an energy of 107.7 kcal mol⁻¹ because that is the energy of the lowest-energy state of the products (*i.e.* CH₂^{*}(CH₃)₂COH radical and H₂O₂) of this endothermic reaction (the zero of energy is the potential energy of separated reactants at their equilibrium geometries; on this scale the potential energy of products is 17.1 kcal mol⁻¹, and their zero-point energy is 90.6 kcal mol⁻¹). The argument of the exponential suggests that it would be instructive to replot the results in Fig. 10 as functions of $E - V_{a,p}^{AG}$ and this is done in Fig. 11. Fig. 11 shows that the curves are very similar as functions of this variable, but curves extend lower for the higher-energy curves, giving more area under the curve. The reason for this is that the curves stop at 107.1 kcal mol⁻¹, as explained in the previous paragraph. Since eqn (12) shows that the area under the curve is proportional to the tunneling

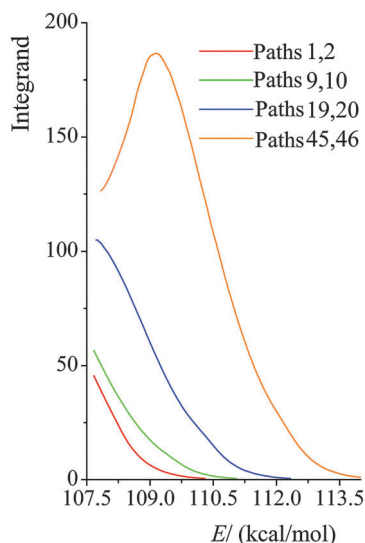


Fig. 10 Values of the integrand in eqn (12) at 200 K for various paths of reaction (R1) as functions of total energy E , with the zero of energy at the overall zero of energy, which is taken as the energy of the equilibrium structure of reactants.

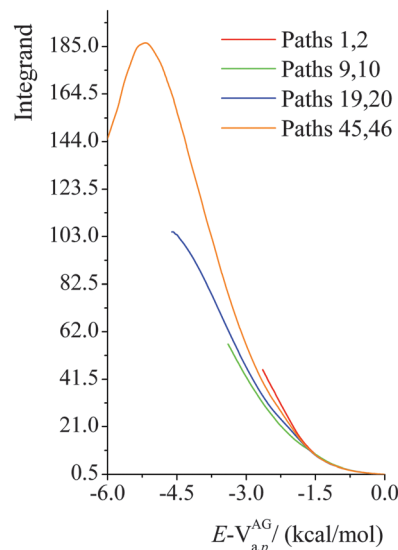


Fig. 11 Values of the integrand in eqn (12) as functions of $E - V_{a,p}^{AG}$ at 200 K for various paths of reaction (R1).

transmission coefficient, we see that the reason that higher-energy paths have larger tunneling transmission coefficients is that they can tunnel at energies farther below the maximum of the effective barrier for tunneling. Thus, for an endothermic reaction, the higher-energy paths have more to gain by tunneling because they can tunnel at energies farther below the barrier top.

Truncation of number of paths included in MP-VTST

The number of paths included in MP-VTST affects the values of the generalized transmission coefficients $\langle \gamma \rangle_P$. Unsigned percentage errors as compared to using all N paths and unsigned percentage deviations between two successive approximations of the generalized transmission coefficients for reaction R1 are shown in Fig. 12. The unsigned percentage errors (UPEs) are computed as:

$$\text{UPE}_P = \frac{|\langle \gamma \rangle_P - \langle \gamma \rangle_N|}{\langle \gamma \rangle_N} \times 100\% \quad (13)$$

where UPE_P is the unsigned error of generalized transmission coefficient that includes P distinguishable paths (including paths that are mirror images), and $\langle \gamma \rangle_N$ is the exact generalized transmission coefficient that includes all 46 distinguishable paths. Unsigned percentage deviations (UPDs) are computed as:

$$\text{UPD}_P = \frac{|\langle \gamma \rangle_{P-2} - \langle \gamma \rangle_P|}{\langle \gamma \rangle_P} \times 100\% \quad (14)$$

where UPD_P is the unsigned deviation between the generalized transmission coefficient based on an average over P paths and the one that is based on $P-2$ paths. (We use $P-2$ rather than $P-1$ because the paths occur in enantiomeric pairs.) Since the generalized transmission coefficient is the only P -dependent quantity in eqn (1), UPD_P defined in eqn (14) is equivalent to the unsigned percentage deviation of the MP-VTST rate constants computed including P and $P-2$ paths.



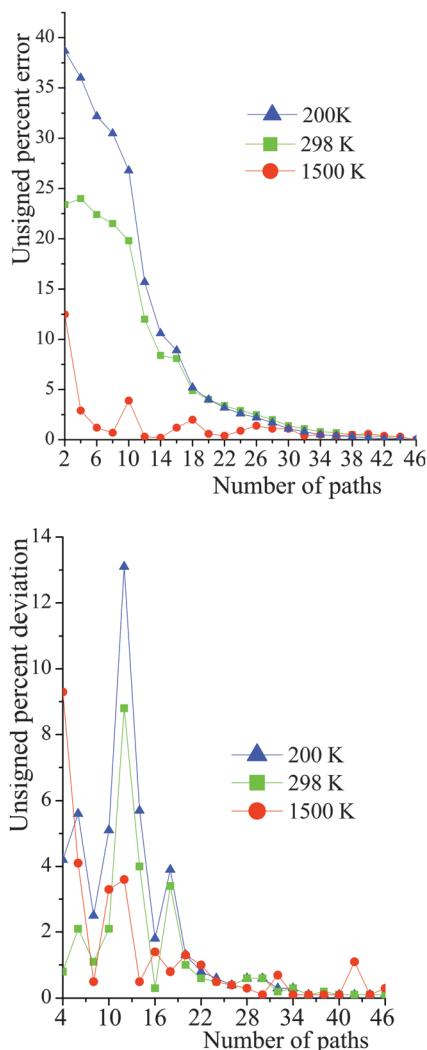


Fig. 12 Unsigned percent deviations of the generalized transmission coefficients for reaction (R1). The number of paths included (P) is the abscissa. The paths are indexed 1–46, with path numbering in order of increasing classical barrier height. The top figure is the deviation between including P paths and including all N paths. The bottom figure is the deviation between including P paths and including $P - 2$ paths. See eqn (13) and (14).

For small values of P , UPD_P is fluctuating rather than converging as P increases; thus, even if UPD_P is very small ($<5\%$), the truncation error could still be quite large. For instance, at 200 K, UPD_8 is only 2.5%, however UPE_8 is 30%. Nevertheless, at high temperatures (for instance, 1500 K), including a few paths is accurate enough for the final generalized transmission coefficients.

5. Summary

In the current work we discussed the errors introduced in truncating the number of paths in MP-VTST calculations. For hydrogen abstraction from *tert*-butanol by HO_2 radical, we found that the variational effects and the contributions of multidimensional tunneling are very different among the

various reaction paths. As a consequence, including only one or a few paths does not yield accurate rate constants at low temperature. Over the entire temperature range, the variational effect is negligible for the lowest-energy path; but for some of the higher-energy paths, the behavior is significantly different from that for the lowest-energy path; we gave an example where the variational recrossing transmission coefficients for different paths vary by up to a factor of 1.4. The differences are even greater – much greater – for the tunneling transmission coefficients. At 200 K, the highest-energy path has a small-curvature tunneling transmission coefficient 17 times larger than the one for lowest-energy path. The higher-energy paths can have much larger tunneling transmission coefficients because the effective barrier for tunneling extends higher above the reactants and products, giving a greater energy range in which tunneling can operate. We conclude that the multi-path methodology is necessary for better understanding of the detailed kinetics and predicting the thermal rate constants.

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References

- 1 T. Yu, J. Zheng and D. G. Truhlar, *J. Phys. Chem. A*, 2012, **116**, 297–308.
- 2 J. Zheng and D. G. Truhlar, *Faraday Discuss.*, 2012, **157**, 59–88.
- 3 T. Yu, J. Zheng and D. G. Truhlar, *Chem. Sci.*, 2011, **2**, 2199–2213.
- 4 J. Zheng, R. Meana-Pañeda and D. G. Truhlar, *J. Am. Chem. Soc.*, 2014, **136**, 5150–5160.
- 5 X. Xu, T. Yu, E. Papajak and D. G. Truhlar, *J. Phys. Chem. A*, 2012, **116**, 10480–10487.
- 6 I. M. Alecu, J. Zheng, E. Papajak, T. Yu and D. G. Truhlar, *J. Phys. Chem. A*, 2012, **116**, 12206–12213.
- 7 P. Seal, G. Oyedepo and D. G. Truhlar, *J. Phys. Chem. A*, 2013, **117**, 275–282.
- 8 J. L. Bao, R. Meana-Pañeda and D. G. Truhlar, *Chem. Sci.*, 2015, **6**, 5866–5881.
- 9 J. L. Bao, P. Seal and D. G. Truhlar, *Phys. Chem. Chem. Phys.*, 2015, **17**, 15928–15935.
- 10 P. Seal, E. Papajak and D. G. Truhlar, *J. Phys. Chem. Lett.*, 2012, **3**, 264–271.
- 11 S. M. Sarathy, S. Vranckx, K. Yasunaga, M. Mehl, P. Obwald, W. K. Metcalfe, C. K. Westbrook, W. J. Pitz, K. Kohse-Hoinghaus, R. X. Fernandes and H. J. Curran, *Combust. Flame*, 2012, **159**, 2028–2055.
- 12 H.-H. Carstensen, A. M. Dean and O. Deutschmann, *Proc. Combust. Inst.*, 2007, **31**, 149–157.
- 13 J. Zheng and D. G. Truhlar, *J. Chem. Theory Comput.*, 2013, **9**, 1356–1367.



- 14 Y. P. Liu, G. C. Lynch, T. N. Truong, D.-H. Lu and D. G. Truhlar, *J. Am. Chem. Soc.*, 1993, **115**, 2408–2415.
- 15 D. G. Truhlar and A. Kuppermann, *J. Am. Chem. Soc.*, 1971, **93**, 1840–1851.
- 16 B. C. Garrett, D. G. Truhlar, R. S. Grev and A. W. Magnuson, *J. Phys. Chem.*, 1980, **84**, 1730–1748.
- 17 A. Fernandez-Ramos, B. A. Ellingson, B. C. Garrett and D. G. Truhlar, in *Reviews in Computational Chemistry*, ed. K. B. Lipkowitz and T. R. Cundari, Wiley-VCH, Hoboken, NJ, 2007, vol. 23, pp. 125–232.
- 18 Y. Zhao and D. G. Truhlar, *J. Chem. Theory Comput.*, 2008, **4**, 1849–1868.
- 19 B. J. Lynch, Y. Zhao and D. G. Truhlar, *J. Phys. Chem. A*, 2003, **107**, 1384–1388.
- 20 J. Zheng, S. L. Mielke, K. L. Clarkson and D. G. Truhlar, *MSTor computer program, version 2011*, University of Minnesota, Minneapolis, MN, 2011.
- 21 Y. Zhao, N. E. Schultz and D. G. Truhlar, *J. Chem. Phys.*, 2005, **123**, 161103.
- 22 Y. Zhao and D. G. Truhlar, *Theor. Chem. Acc.*, 2008, **120**, 215–241.
- 23 Y. Zhao and D. G. Truhlar, *Acc. Chem. Res.*, 2008, **41**, 157–167.
- 24 J. Zheng, X. Xu and D. G. Truhlar, *Theor. Chem. Acc.*, 2011, **128**, 295–305.
- 25 T. B. Adler, G. Knizia and H. J. Werner, *J. Chem. Phys.*, 2007, **127**, 221106.
- 26 G. Knizia, T. B. Adler and H. J. Werner, *J. Chem. Phys.*, 2009, **130**, 054104.
- 27 F. R. Manby, *J. Chem. Phys.*, 2003, **119**, 4607–4613.
- 28 E. Papajak and D. G. Truhlar, *J. Chem. Theory Comput.*, 2011, **7**, 10–18.
- 29 M. J. Frisch, G. W. Trucks, H. B. Schlegel, G. E. Scuseria, M. A. Robb, J. R. Cheeseman, G. Scalmani, V. Barone, B. Mennucci, G. A. Petersson, H. Nakatsuji, M. Caricato, X. Li, H. P. Hratchian, A. F. Izmaylov, J. Bloino, G. Zheng, J. L. Sonnenberg, M. Hada, M. Ehara, K. Toyota, R. Fukuda, J. Hasegawa, M. Ishida, T. Nakajima, Y. Honda, O. Kitao, H. Nakai, T. Vreven, J. A. Montgomery, Jr., J. E. Peralta, F. Ogliaro, M. Bearpark, J. J. Heyd, E. Brothers, K. N. Kudin, V. N. Staroverov, R. Kobayashi, J. Normand, K. Raghavachari, A. Rendell, J. C. Burant, S. S. Iyengar, J. Tomasi, M. Cossi, N. Rega, J. M. Millam, M. Klene, J. E. Knox, J. B. Cross, V. Bakken, C. Adamo, J. Jaramillo, R. Gomperts, R. E. Stratmann, O. Yazyev, A. J. Austin, R. Cammi, C. Pomelli, J. Ochterski, R. L. Martin, K. Morokuma, V. G. Zakrzewski, G. A. Voth, P. Salvador, J. J. Dannenberg, S. Dapprich, A. D. Daniels, O. Farkas, J. B. Foresman, J. V. Ortiz, J. Cioslowski and D. J. Fox, *GAUSSIAN 09 (Revision D. 01)*, Gaussian, Inc., Wallingford, CT, 2009.
- 30 Y. Zhao, R. Peverati, K. Yang and D. G. Truhlar, *MN-GFM version 5.2 computer program module*, University of Minnesota, Minneapolis, 2011.
- 31 H.-J. Werner, P. J. Knowles, F. R. Manby, M. Schütz, P. Celani, G. Knizia, T. Korona, R. Lindh, A. Mitrushenkov, G. Rauhut, T. B. Adler, R. D. Amos, A. Bernhardsson, A. Berning, D. L. Cooper, M. J. O. Deegan, A. J. Dobbyn, F. Eckert, E. Goll, C. Hampel, A. Hesselmann, G. Hetzer, T. Hrenar, G. Jansen, C. Küoppl, Y. Liu, A. W. Lloyd, R. A. Mata, A. J. May, S. J. McNicholas, W. Meyer, M. E. Mura, A. Nicklaß, P. Palmieri, K. Pflüger, R. Pitzer, M. Reiher, T. Shiozaki, H. Stoll, A. J. Stone, R. Tarroni, T. Thorsteinsson, M. Wang and A. Wolf, *MOLPRO, version 2012.1*, University of Birmingham, Birmingham, 2012.
- 32 K. A. Nguyen, C. F. Jackels and D. G. Truhlar, *J. Chem. Phys.*, 1996, **104**, 6491–6496.
- 33 C. F. Jackels, Z. Gu and D. G. Truhlar, *J. Chem. Phys.*, 1995, **102**, 3188–3201.
- 34 J. Zheng, S. Zhang, J. C. Corchado, Y. Y. Chuang, E. L. Coitiño, B. A. Ellingson and D. G. Truhlar, *GAUSSRATE, version 2009-A*, University of Minnesota, Minneapolis, 2010.
- 35 J. Zheng, S. Zhang, B. J. Lynch, J. C. Corchado, Y.-Y. Chuang, P. L. Fast, W.-P. Hu, Y.-P. Liu, G. C. Lynch, K. A. Nguyen, C. F. Jackels, A. F. Ramos, B. A. Ellingson, V. S. Melissas, J. Vill_a, I. Rossi, E. L. Coitino, J. Pu, T. V. Albu, R. Steckler, B. C. Garrett, A. D. Isaacson and D. G. Truhlar, *POLYRATE, version 2010-A*, University of Minnesota, Minneapolis, 2010.
- 36 D.-H. Lu, T. N. Truong, V. Melissas, G. C. Lynch, Y.-P. Liu, B. C. Garrett, R. Steckler, A. D. Isaacson, S. N. Rai and G. C. Hancock, *et al.*, *Comput. Phys. Commun.*, 1992, **71**, 235–262.
- 37 I. M. Alecu, J. Zheng, Y. Zhao and D. G. Truhlar, *J. Chem. Theory Comput.*, 2010, **6**, 2872–2887.
- 38 J. Zheng and D. G. Truhlar, *Phys. Chem. Chem. Phys.*, 2010, **12**, 7782–7793.
- 39 D. G. Truhlar, *J. Chem. Educ.*, 1978, **55**, 309–311.
- 40 M. M. Kreevoy and D. G. Truhlar, *Transition State Theory, in Investigation of Rates and Mechanisms of Reactions*, ed. C. F. Bernasconi, Techniques of Chemistry Series, John Wiley and Sons, New York, 4th edn, 1986, vol. 6, Part 1, pp. 13–95.
- 41 M. Garcia-Viloca, J. Gao, M. Karplus and D. G. Truhlar, *Science*, 2004, **303**, 186–195.

