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Rhodium catalyzed synthesis of difluoromethyl cyclopropanes†

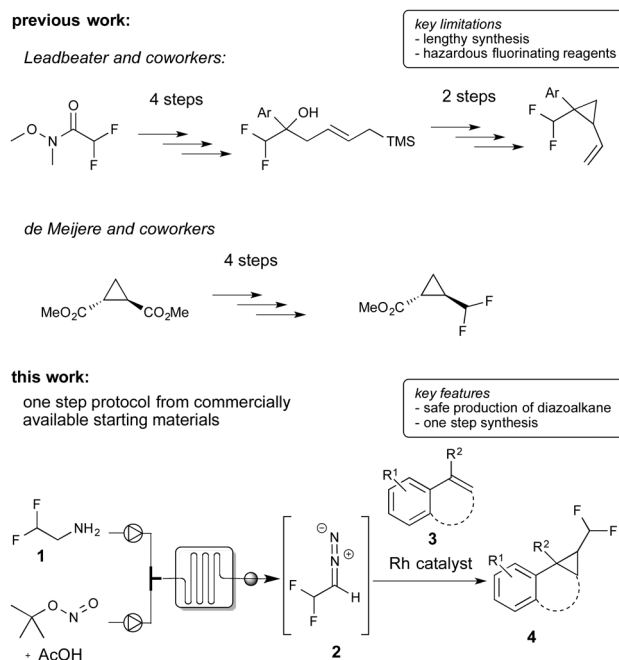
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Difluoromethyl-substituted cyclopropanes still remain one of the most challenging class of substrates. Despite significant progress in modern chemistry, an atom-economic and concise synthesis has not been described to date. Herein, we describe the first method for their catalytic, one-step synthesis using difluoromethyl diazomethane and a rhodium(II) catalyst.

In recent years, novel methods for the introduction of difluoromethyl groups have attracted the interest of organic chemists from all fields.^{1,2} Although the difluoromethyl group differs only in a single fluorine atom from the well-studied trifluoromethyl group, it possesses significantly altered properties making difluoromethylated compounds attractive synthetic targets. While trifluoromethyl groups are inert from a chemical and biological perspective, the difluoromethyl group possesses lower lipophilicity and increased polarity. Moreover, it is an alcohol or thiol bio-isostere and can act as a hydrogen bond donor and is thus highly attractive to drug discovery programs.^{1b}

Over the past decade, trifluoromethyl diazomethane found regular application in organic synthesis and led to a flourishing area of research.^{3,4} Interestingly, while trifluoromethyl diazomethane is known since 1943,⁵ the difluoromethyl-substituted analogue was first described only in 2015⁶ and applications of this novel reagent in organic synthesis are still scarce.^{6,7} Against this background, we became interested in the catalytic synthesis of difluoromethyl-substituted cyclopropanes, which should be accessible *via* this newly introduced reagent.⁸

To date, there are only two reports claiming 1-difluoromethyl-2-aryl cyclopropanes, yet without describing the synthesis.⁹ The synthesis of difluoromethyl cyclopropanes in general has been studied only to a very limited extent.¹⁰ Leadbeater and co-workers and Hanamoto and co-workers reported the synthesis of this class of substances applying ring closure reactions using tedious multi-step syntheses.^{10a,b} De Meijere and co-workers



Scheme 1 Synthetic approaches towards difluoromethyl cyclopropanes.

reported a multi-step synthesis of difluoromethyl cyclopropanes starting from cyclopropyl esters and by application of hazardous and highly reactive fluorinating reagents for the introduction of the difluoromethyl group.^{10c} To the best of our knowledge no catalytic methods for the efficient one-step preparation from readily available starting materials has been described yet (Scheme 1).

Difluoromethyl-substituted cyclopropanes find few applications,^{9,11} such as in the ROR-gamma modulator (6)^{9a} or as insecticides (7).^{9b,11b} The most prominent compound containing a difluoromethyl cyclopropane is voxilaprevir (5), which is currently in phase II clinical trials for treatment of hepatitis C virus infections (Fig. 1).^{11c,d}

We started our investigations towards difluoromethyl-substituted cyclopropanes by evaluating different batch transformations.

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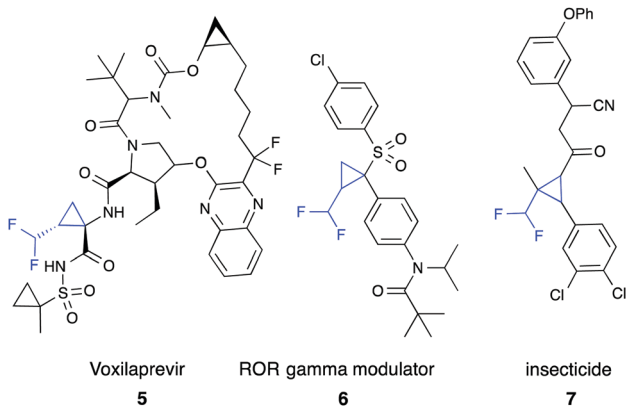
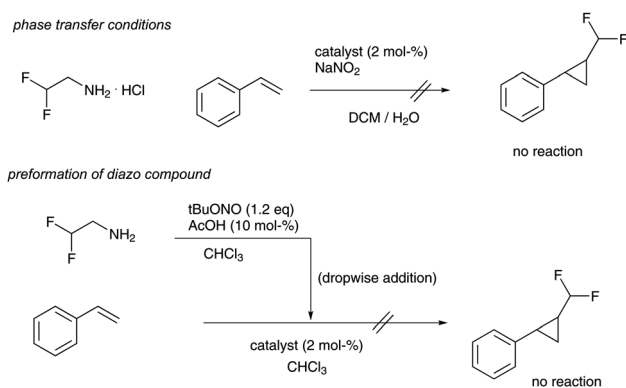


Fig. 1 Difluoromethyl cyclopropanes in medicinal and agrochemical research.



Scheme 2 Attempts to difluoromethyl cyclopropane synthesis in batch reactions (catalysts investigated = Rh_2OAc_4 , $\text{Ru}(\text{TPP})\text{CO}$, $\text{Fe}(\text{TPP})\text{Cl}$, Rh_2Piv_4).

First, we investigated phase transfer protocols,^{4c,i} that find regular application in reactions using trifluoromethyl diazomethane, yet the desired product could not be observed using different cyclopropanation catalysts. Similarly, preformation of difluoromethyl diazomethane and subsequent dropwise addition of the reagent to styrene and different cyclopropanation catalysts failed completely (Scheme 2). This observation is in sharp contrast to trifluoromethyl diazomethane and the corresponding cyclopropanation reactions.⁴

Recently, we reported on the continuous-flow synthesis of fluoroalkyl-substituted diazoalkanes and their application in pyrazole and pyrazoline synthesis.^{7,12,13} As conventional batch reactions did not provide the desired product, we became interested in evaluating the flow protocol in this cyclopropanation reaction.

For the continuous-flow generation of difluoromethyl diazomethane, we used a micromixer (Little Things Factory, MR-LAB Type MST, mixer volume: 200 μL) to mix a stream of difluoroethylamine with a mixture of *t*BuONO and AcOH. We then examined a range of different conditions, *e.g.* temperature, amount of acetic acid and residence time in the microreactor, by monitoring aliquots of the outlet by NMR spectroscopy. 40 mol% of acetic acid and a residence time of 2 minutes in the

Table 1 Optimization of reaction conditions

Entry ^a	Catalyst	Loading (mol%)	Solvent	Yield ^b (%)
1	$\text{Fe}(\text{TPP})\text{Cl}$	2	CHCl_3	Traces
2	$\text{Ru}(\text{TPP})\text{CO}$	2	CHCl_3	No rct.
3	$\text{Co}(\text{salen})$	2	CHCl_3	No rct.
4	CuOTf	2	CHCl_3	No rct.
5	$[\text{Cp}^*\text{IrCl}_2]_2$	2	CHCl_3	No rct.
6	$[\text{Cp}^*\text{RhCl}_2]_2$	2	CHCl_3	No rct.
7	$\text{Pd}(\text{acac})_2$	2	CHCl_3	No rct.
8	Rh_2OAc_4	2	CHCl_3	12
9	Rh_2Piv_4	2	CHCl_3	6
10	Rh_2TPA_4	2	CHCl_3	No rct.
11	Rh_2esp_2	2	CHCl_3	24
12	Rh_2esp_2	5	CHCl_3	67 ^c
13	Rh_2esp_2	5	CHCl_3	13 ^d

^a A solution of difluoroethylamine (0.2 M, 4 eq.) in CHCl_3 and a mixture of AcOH (1.6 eq.) and *t*BuONO (0.24 M, 4.8 eq.) in CHCl_3 were transferred by syringe pump into the microreactor (Little Things Factory, MR Lab, Type MST, inner volume 200 μL) at a constant flow rate of 50 $\mu\text{L min}^{-1}$ for each syringe. The microreactor was heated to 75 °C and the outlet of the microreactor was connected to a back pressure regulator (40 psi). The reaction solution from the microreactor was added to a standard reaction flask at room temperature under Ar atmosphere containing styrene (0.4 mmol, 1 eq.) and the appropriate catalyst in 4 mL CHCl_3 .

^b Yields determined by ^{19}F -NMR of the crude reaction mixture using α,α,α -trifluorotoluene as internal standard. ^c Isolated yield after column chromatography. ^d Reaction temperature of the cyclopropanation reaction 0 °C.

microreactor at 75 °C proved to yield the desired difluoromethyl diazomethane in about 40% yield. With the optimal conditions for the preparation of difluoromethyl diazomethane in hand, we next assessed the cyclopropanation reaction of styrene with difluoromethyl diazomethane, prepared in continuous-flow, using a broad variety of different cyclopropanation catalysts (Table 1).

Initially, we examined iron and ruthenium porphyrine complexes, which were already demonstrated to be versatile catalysts for the preparation of trifluoromethyl substituted cyclopropanes. To our surprise, only trace amounts of the desired difluoromethyl-substituted cyclopropane could be observed using $\text{Fe}(\text{TPP})\text{Cl}$. Other typical cyclopropanation catalysts, such as $\text{Ru}(\text{TPP})\text{CO}$, $\text{Co}(\text{salen})$ complexes or $\text{Cu}(\text{I})$, or $\text{Ru}(\text{III})$ and $\text{Ir}(\text{III})$ complexes failed completely and no product was observed.

Interestingly, both Rh_2OAc_4 and Rh_2Piv_4 gave minor amounts of the desired cyclopropanation product, while Rh_2TPA_4 failed to yield the cyclopropane. These results are in sharp contrast to previously reported cyclopropanation reactions using trifluoromethyl diazomethane, which we assume to result from (a) the highly instable nature of difluoromethyl diazomethane and (b) from facile β -hydride elimination of the intermediate metal-carbene complex when using difluoromethyl diazomethane. We thus hypothesized, that sterically more demanding catalysts would reduce the rate of β -hydride elimination and result in higher yield of the desired cyclopropane.



To our delight, Rh₂esp₂¹⁴ proved significantly better and we could observe increased yields of the cyclopropanation product. Encouraged by these results we increased the catalyst loading to 5 mol% and we could finally isolate the desired difluoromethylated cyclopropane **4a** for the first time in 67% yield.

In further experiments we investigated the substrate scope of this transformation (Table 2). Using our continuous-flow protocol for the preparation of difluoromethyl diazomethane, we could show that a range of different styrenes can be converted to the corresponding difluoromethyl-substituted cyclopropanes in

moderate to good yield. This transformation is particularly useful, as our newly developed protocol allows for the first time an efficient synthesis of difluoromethyl-substituted cyclopropanes in an operationally simple one-step protocol from cheap, commercially available olefins and amines.

In general electron donating and electron-withdrawing substituents are well tolerated and the corresponding cyclopropanes can be isolated in moderate to good yield. Moreover, different substitution patterns, such as meta-substituents are compatible under these reaction conditions and the desired reaction products can be obtained in good yield.

Further investigations concentrated on substituted styrenes. We probed α -methyl styrene and *trans*- β methyl styrene under these reaction conditions and were delighted to observe that the corresponding cyclopropanes (**4l** and **4m**) can be isolated in good yield. Similarly, indene and dihydronaphthalene gave the desired cyclopropanation products (**4n** and **4o**) in moderate yield. Heterocyclic compounds (**4p**) proved to give only unsatisfactory yield of the desired cyclopropane. Allylbenzene, as a non-aromatic olefinic substrate, did not provide the desired product. The diastereomeric ratio of all cyclopropanes obtained was in the range of 1.2–1.7 : 1. Only in the case of **4o** we observed a slightly higher diastereomeric ratio of 3 : 1.

In summary, we have developed a new protocol that allows for the first time a catalytic one-step synthesis of difluoromethyl cyclopropanes. A key element for the development of this protocol is the efficient preparation of difluoromethyl diazomethane in continuous-flow. Subsequent application in rhodium(II) catalyzed cyclopropanation reactions provide for the first time difluoromethyl substituted cyclopropanes in a one-step protocol and good isolated yields.

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Table 2 Scope of the cyclopropanation reaction

Cyclopropane ^a	Yield ^b (%)	Cyclopropane ^a	Yield ^b (%)
	67		52
	68		34
	57		49
	42		32
	56		63
	59		45
	60		50
			49
			< 10 ^c

^a A solution of difluoroethylamine (0.2 M, 4 eq.) in CHCl₃ and a mixture of AcOH (1.6 eq.) and *t*BuONO (0.24 M, 4.8 eq.) in CHCl₃ were transferred by syringe pump into the microreactor (Little Things Factory, MR Lab, Type MST, inner volume 200 µL) at a constant flow rate of 50 µL min⁻¹ for each syringe. The microreactor was heated to 75 °C and the outlet of the microreactor was connected to a back pressure regulator (40 psi). The reaction solution the microreactor was added to a standard reaction flask under Ar atmosphere containing olefine (0.4 mmol, 1 eq.) and Rh₂esp₂ (5 mol%) in 4 mL CHCl₃. ^b Isolated yield after column chromatography. ^c Determined by NMR.



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