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Carbon-rich “Click” 1,2,3-triazoles: hexaphenylbenzene and hexa-*peri*-hexabenzocoronene-based ligands for Suzuki–Miyaura catalysts†

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Hexaphenylbenzene (HPB) and hexa-*peri*-hexabenzocoronene (HBC) functionalised 1,2,3-triazoles have been synthesised using an optimised copper(i)-catalysed azide–alkyne cycloaddition (CuAAC) reaction. The coordination chemistry of these ligands was explored through the synthesis of the respective palladium(ii) complexes and their activity as catalysts in the Suzuki–Miyaura reaction assessed.

Polycyclic aromatic hydrocarbons (PAHs) are one of the most widely studied families of organic compounds, and present many interesting synthetic challenges. Breakthroughs in the area of PAH synthesis have led to a plethora of new materials with unique photophysical and electronic properties, and some of these materials have been incorporated into electronic devices.¹

Hexa-*peri*-hexabenzocoronene (HBC) is a graphitic, all benzenoid PAH that has remarkable thermal and chemical stability.² Traditionally the synthesis of HBC has been hampered by harsh conditions and as such could only be isolated in trace quantities.^{2,3} However, developments by the Müllen group have made HBCs more readily accessible, through the intramolecular oxidative aromatic coupling of hexaphenylbenzenes (HPBs).⁴ These methods now allow for the synthesis of HPBs and HBCs with a range of different substitution patterns and functionalities,⁵ allowing these units to be investigated in a variety of applications.⁶ Additionally, it has been shown that HBCs undergo facile self-assembly due to robust $\pi \cdots \pi$ interactions.⁷ Hexa(alkyl)-HBCs, in particular, self-assemble into highly ordered, columnar mesophases that have high charge-carrier mobility and as such act as molecular semiconductors.⁸ HPBs have also been shown to self-assemble⁹ and exploited for the fabrication of OLEDs.¹⁰ In addition to

organic optoelectronic materials based on HBC and HPB, there have been a few recent reports on the use of HBC and HPB ligands for the development of hybrid metal–organic materials with new optical properties.^{11–13} The photophysics of HBC-acetylide Pt(II) complexes and systems that incorporate nitrogen into the HBC framework have been examined by Draper and co-workers.¹⁴ We have recently reported an HBC(*t*-Bu)₅-bpy (bpy = 2,2'-bipyridine) ligand and its corresponding Re(I) chlorotricarbonyl complex; this complex exhibits a long-lived charge transfer excited state.¹¹ Likewise, HPB ligands have been used to develop photophysically active metal complexes.¹⁵

Despite the upsurge in interest in the development of HBC ligand systems, there are few general methods that enable the generation of soluble functionalised HBC architectures. One potential solution to the solubility and reactivity problems is through the use of the copper(i) catalysed azide–alkyne cycloaddition (CuAAC) “click” reaction.¹⁶ This is an efficient and functional group tolerant way of generating 1,4-disubstituted-1*H*-1,2,3-triazoles (tzls), from alkynes and azides, and the triazole units have been shown to be effective N-donor ligands.¹⁷ Additionally, the solubility of HPBs and HBCs can be tuned at a late stage of synthesis from alkyne-functionalised precursors, which in turn may give rise to novel liquid-crystalline materials. Recently we have demonstrated that triazoles act as electronic insulators¹⁸ and therefore could be used in the modulation of electronics in donor–acceptor systems. Herein we report the synthesis of directly appended HPB and HBC triazoles for the first time. The triazoles are employed as ligands to form [PdCl₂L₂] complexes which were subsequently investigated as precatalysts in the Suzuki–Miyaura reaction.

Penta(*tert*-butyl)-substituted hexaphenylbenzene (HPB)- and hexa-*peri*-hexabenzocoronene (HBC)-acetylenes were synthesised using established coupling conditions (ESI†).^{11,13} Due to the limited solubility of these species in conventional polar solvents, an optimised synthesis of the triazoles was needed (ESI†). The triazoles **1a** and **1b** were optimally synthesised in dichloromethane using benzyl azide, [Cu(CH₃CN)₄]BF₄/tris(benzyltriazolylmethyl)amine (TBTA), and triethylamine at

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† Electronic supplementary information (ESI) available: Experimental procedures,

¹H and ¹³C NMR, MALDI-TOF-MS, crystallographic data, ORTEP and packing diagrams, molecular modelling. CCDC 1502527–1502529. For ESI and crystallographic data in CIF or other electronic format see DOI: 10.1039/c6cc07413h



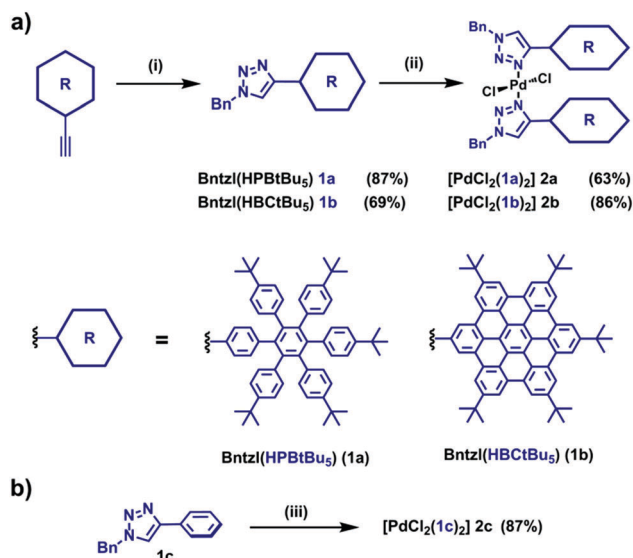


Fig. 1 (a) Synthesis of HPB and HBC triazoles and [PdCl₂L₂] complexes. Conditions: (i) R-CCH (1 equiv.), benzyl azide (1.2 equiv.), NEt₃ (1.2 equiv.), [Cu(MeCN)₄]BF₄ (0.10 equiv.), TBTA (0.10 equiv.), CH₂Cl₂, rt, 18 h; (ii) [PdCl₂(MeCN)₂] (0.5 equiv.), acetone (**2a**) or CHCl₃ (**2b**), rt, 1 h; (b) synthesis of model complex. Conditions: (iii) [PdCl₂(MeCN)₂] (0.5 equiv.), acetone, rt, 1 h.

room temperature in good yields (**1a**: 87%, **1b**: 69%; Fig. 1). ¹H and ¹³C NMR spectroscopy supported the formation of **1a** and **1b**, showing signals consistent with the presence of the triazole ring, benzylic methylenes, as well as the absence of the alkynyl proton signals of the precursors. Additionally, significant downfield shifts are observed within the polyaromatic cores for the nuclei proximal to the electron-poor triazole ring (ESI[†]). MALDI-TOF-MS data show signals corresponding to the molecular ion as well as the loss of dinitrogen ([M-N₂]⁺). Triazole **1b** was also synthesised from **1a** by oxidative aromatic cyclodehydrogenation in quantitative yield; the ¹H NMR spectrum and MALDI-TOF-MS data are identical to those of **1b** formed *via* the CuAAC method (ESI[†]).

With **1a** and **1b** in hand, we examined the ability of the new triazoles to act as ligands. The [PdCl₂(**1a-b**)₂] (**2a-b**) complexes were formed by reaction of the ligands with [PdCl₂(MeCN)₂] (0.5 equiv.) in acetone (**2a**) or chloroform (**2b**) solution and were isolated as yellow solids (**2a**: 63%, **2b**: 86%; Fig. 1). Elemental analyses were consistent with a 2 : 1 ligand/metal ratio. ¹H and ¹³C NMR spectra of **2a** and **2b** show similar profiles to their ligand precursors. The triazole ring protons of each complex shift upfield upon complexation, presumably due to increased shielding by the benzyl ring. Conversely, the benzylic methylene protons shift downfield because of metal-induced deshielding. With respect to the HBC proton signals of **2b**, the protons proximal to the triazole shift downfield upon complexation and broaden significantly due to slow rotation in solution; the remaining proton signals shift upfield consistent with the mutual face-to-face overlap of the two HBC units. MALDI-TOF-MS spectra of the complexes show weak signals associated with the complex (**2a**: [M-2Cl-H]⁺; **2b**: [M]⁺) and stronger signals associated with the fragmentation of the respective ligands.

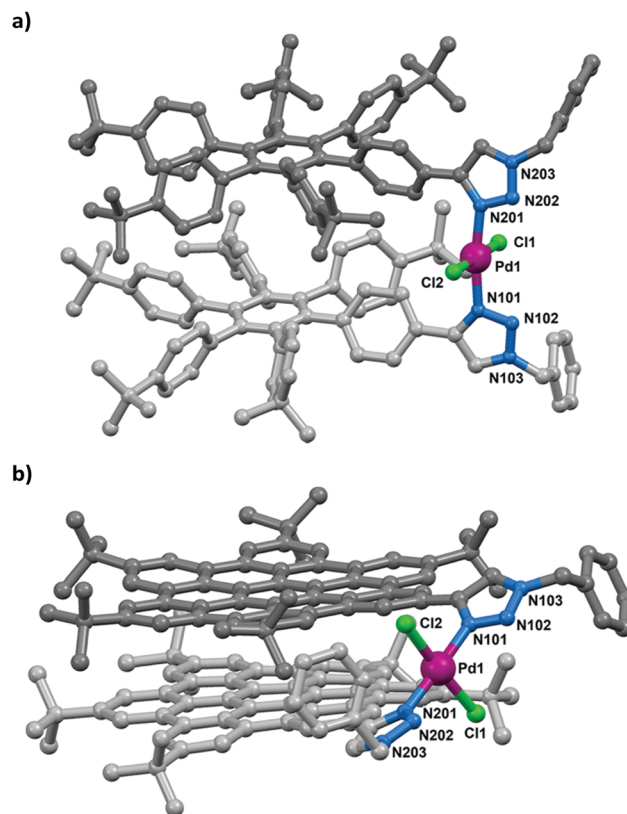


Fig. 2 Ball-and-stick representations of the X-ray crystal structures of (a) **2a**·7CH₂Cl₂ and (b) **2b**. Hydrogen atoms and solvate molecules have been omitted for clarity. Key bond lengths (Å) and angles (°) for **2a**: Pd1-N101 = 2.029(5) Å, Pd1-N201 = 1.994(5) Å, Pd1-Cl1 = 2.297(2) Å, Pd1-Cl2 = 2.279(2) Å, N101-Pd1-N201 = 171.3(4)°, Cl1-Pd1-Cl2 = 178.95(7)°. **2b**: Pd1-N101 = 1.993(4) Å, Pd1-N201 = 1.996(4) Å, Pd1-Cl1 = 2.320(2) Å, Pd1-Cl2 = 2.316(2) Å, N101-Pd1-N201 = 177.9(2)°, Cl1-Pd1-Cl2 = 174.18(5)°.

X-ray quality crystals of both complexes were obtained by either a layered diffusion of methanol into a solution (**2a**) or the slow evaporation of a CH₂Cl₂ solution (**2b**). Both complexes show the expected connectivity of *trans*-N-donor complexes and a geometry about the palladium(II) centres typical of palladium-triazole complexes.¹⁹ The Pd-N(tzl) distances range from 1.993(4)–2.029(5) Å and sit within the lower range of reported Pd-N(tzl) distances (1.983–2.166 Å). Although the N-Pd-N bond angle of **2b** at 177.9(2)° is close to the mean of reported complexes *trans*-bis(tzl) palladium complexes (176.4°), that of **2a** (171.3(2)°) is rather acute (Fig. 2), presumably because of increased ligand size and resultant intramolecular interactions. The complexes are isostructural in the sense that the polyaromatic cores of each ligand in the respective complexes are intramolecularly stacked and eclipsed. The stacking in **2a** appears to be stabilised by multiple edge-to-face C-H...π interactions. Each molecule of **2a** co-crystallised with seven CH₂Cl₂ solvate molecules, five of which are situated within clefts of HPB moieties and stabilised by two-fold C-H...π interactions (ESI[†]). For **2b**, however, the HBC moieties are stabilised through offset face-to-face π-interactions facilitated by the π-electron-rich nature of the HBC unit (central ring inter-centroid distance = 3.486 Å),



thus supporting the NMR evidence of stacking in solution. The chloro ligands of **2b** are positioned within bay positions of the HBC units and are stabilised by C–H...Cl interactions (ESI†).

As palladium complexes are used as catalysts in a wide range of reactions,²⁰ we investigated the catalytic potential of **2a** and **2b**. Triazoles feature prominently as supporting ligands for a variety of metal-based catalysts.^{21,22} Although triazoles often comprise part of bidentate ligands, few catalytically-active complexes employ triazoles as monodentate ligands.²² Palladium(II) triazole dendrimers have shown good catalytic activity and high turnover in the Suzuki–Miyaura reaction;²³ accordingly we examined **2a** and **2b**, along with phenyl-substituted model **2c**¹⁹ (Fig. 1b) and commercially available [PdCl₂(PPh₃)₂] (**3**), as catalysts in the Suzuki–Miyaura cross-coupling reaction (Table 1 and ESI†). Upon addition of model **2c** (5 mol%) to phenylboronic acid and 4-bromoacetophenone at 80 °C for 4 hours, an 84% yield of 4-acetylbiphenyl was obtained confirming this bis(monodentate)triazole complex is catalytically active. Repeating the reaction under identical conditions with **2a** and **2b** resulted in slightly improved yields of 94% and 92%, respectively. Commercially available [PdCl₂(PPh₃)₂] (**3**) afforded the product in 90% yield.

To further probe whether the incorporation of polyaromatic groups in ligands **1a** and **1b** has an effect on the catalytic activity, the reaction was repeated with the less activated 3,5-dimethoxybromobenzene substrate. Utilising the same reaction conditions (but a longer reaction time) complexes **2a** and **2b**, along with model **2c** and **3**, were all found to be active. Surprisingly, employing **2a** and **2c** resulted in a significantly lower yield of the biphenyl product (30–39%), while HBC-containing **2b** and **3** produced good yields of 69% and 76%, respectively. The reactions with **2b** and **3** were also monitored *via* ¹H NMR spectroscopy and revealed a small difference in activity. It is noteworthy that for this system, conversion ceased after *ca.* 5 hours, indicative of catalyst decomposition for both species (ESI†). This outcome was supported by an increase in

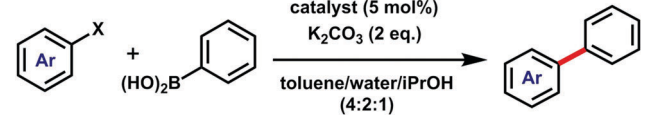
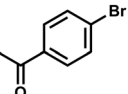
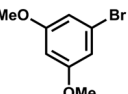
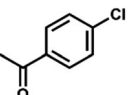
the intensity of free ligand signals in the ¹H NMR spectra. To test the limits of the triazole catalysts **2a–c**, in comparison with **3**, the haloarene substrate 4'-chloroacetophenone was reacted under similar conditions. Conversions dropped markedly for all catalysts with low yields of 4-acetylbiphenyl (0–14%). The HBC-based complex **2b** was catalytically active while the HPB analogue **2a** was inactive; remarkably, the Ph-bearing model **2c** showed comparable activity to **2b**, albeit low in both cases (13–14%). Subsequent monitoring by ¹H NMR spectroscopy indicated rapid catalyst decomposition (*ca.* 2 hours) for both **2b** and **3** (ESI†), rendering the longer reaction times unnecessary.

The activity of Werner complexes **2a–c** (*cf.* phosphine complex **3**) is intriguing because it is expected that N-donor ligands have poor donor properties. To investigate this, the donor strengths of all ligands were assessed using a ¹³C NMR technique developed by Huynh and coworkers which involved the synthesis of [PdBr₂(Pr₂-bimy)(L)] complexes (**4a–c,5**; ESI†).²⁴ The data revealed the following: The phosphine ligand has a greater donor strength than all triazole-based ligands, and the triazole ligand with the more electron-rich HBC core (**2b**) has a greater donor strength than the phenyl-appended model analogue **2c**. Thus, the catalytic activity cannot be explained by electronic arguments alone.

There are several postulates to explain the activity of **2b**, each involving the influence of the HBC core; the interaction of the two HBC units within the complex transforms the triazoles into one pseudo-bidentate ligand and/or that the HBC cores help to stabilise catalytically-active palladium nanoparticles that form during the reaction. In the former, intramolecular π -stacking may confer greater stability, however, molecular modelling (ESI†) suggests that while the *cis*-arrangement required for reductive elimination is accessible, it is higher in energy than the *trans*-isomer observed crystallographically. More likely here, is that nanoparticles are playing a significant role in the catalysis. A repeat catalytic run under the conditions of entry 6 (Table 1), but in the presence of a mercury drop, led to a fall in yield from 69% to 22%. It is generally thought that N-donor ligands do not effectively stabilise monoatomic Pd(0) resulting in Pd nanoparticle formation,²⁵ and polyaromatic groups can help slow aggregation of the particles to Pd black precipitates.²⁶ Further work will concentrate on determining the exact nature of the catalytically-active species and the strength of the HBC–HBC interaction in solution.

1,4-Disubstituted-1,2,3-triazoles functionalised with carbon-rich hexaphenylbenzen and hexa-*peri*-hexabenzocoronene substituents have been synthesised, using the CuAAC methodology. Furthermore, the triazole group was found to be tolerant of the oxidative dehydrogenation conditions used for converting HPBs to HBCs. These HPB and HBC compounds have been used as ligands to generate *trans*-[PdCl₂L₂] complexes which show π -interactions between each ligand in the solid state, and in solution for complex **2b**. The triazole-palladium complexes are catalytically active in the Suzuki–Miyaura cross-coupling reaction, with the HBC-bearing **2b** appearing to be the best across the substrates surveyed. Both catalysts have limited lifetimes (2–5 hours) as made evident through ¹H NMR studies.

Table 1 Suzuki–Miyaura catalysis with complexes **2a–c** and complex **3**

				
Ar-X	Entry	Catalyst ^a	Time (h)	Yield ^b (%)
	1	2a	4	94
	2	2b	4	92
	3	2c	4	84
	4	3	4	90
	5	2a	24	30
	6	2b	24	69
	7	2c	24	39
	8	3	24	76
	9	2a	30	0
	10	2b	30	13
	11	2c	30	14
	12	3	30	3

^a [PdCl₂L₂] complex. ^b Isolated yield; average of at least two runs ($\pm 5\%$).



The versatile and functional group tolerant CuAAC “click” approach developed herein should enable the synthesis of a wide range of HBC and HPB 1,2,3-triazole ligands which could be exploited to generate new catalysts and hybrid metal-organic materials with tunable optical properties.

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Notes and references

- D. Braga and G. Horowitz, *Adv. Mater.*, 2009, **21**, 1473; H. Dong, X. Fu, J. Liu, Z. Wang and W. Hu, *Adv. Mater.*, 2013, **25**, 6158; Y. Zhao, Y. Guo and Y. Liu, *Adv. Mater.*, 2013, **25**, 5372; W. Wu, Y. Liu and D. Zhu, *Chem. Soc. Rev.*, 2010, **39**, 1489; S. Varghese and S. Das, *J. Phys. Chem. Lett.*, 2011, **2**, 863; K. Okamoto, J. Zhang, J. B. Housekeeper, S. R. Marder and C. K. Luscombe, *Macromolecules*, 2013, **46**, 8059; X. Guo, M. Baumgarten and K. Müllen, *Prog. Polym. Sci.*, 2013, **38**, 1832; A. Hagfeldt, G. Boschloo, L. Sun, L. Kloo and H. Pettersson, *Chem. Rev.*, 2010, **110**, 6595.
- E. Clar, C. T. Ironside and M. Zander, *J. Chem. Soc.*, 1959, 142; A. Halleux, R. H. Martin and G. S. D. King, *Helv. Chim. Acta*, 1958, **41**, 1177.
- W. Hendel, Z. H. Khan and W. Schmidt, *Tetrahedron*, 1986, **42**, 1127.
- A. Stabel, P. Herwig, K. Müllen and J. P. Rabe, *Angew. Chem., Int. Ed. Engl.*, 1995, **34**, 1609; V. S. Iyer, M. Wehmeier, J. D. Brand, M. A. Keegstra and K. Müllen, *Angew. Chem., Int. Ed. Engl.*, 1997, **36**, 1604.
- R. Yamaguchi, S. Ito, B. S. Lee, S. Hiroto, D. Kim and H. Shinokubo, *Chem. – Asian J.*, 2013, **8**, 178.
- J. Wu, W. Pisula and K. Müllen, *Chem. Rev.*, 2007, **107**, 718.
- J. P. Hill, W. Jin, A. Kosaka, T. Fukushima, H. Ichihara, T. Shimomura, K. Ito, T. Hashizume, N. Ishii and T. Aida, *Science*, 2004, **304**, 1481; Y. Yamamoto, G. Zhang, W. Jin, T. Fukushima, N. Ishii, A. Saeki, S. Seki, S. Tagawa, T. Minari, K. Tsukagoshi and T. Aida, *Proc. Natl. Acad. Sci. U. S. A.*, 2009, **106**, 21051.
- W. Pisula, X. Feng and K. Müllen, *Chem. Mater.*, 2011, **23**, 554.
- S. Hiraoka, K. Harano, T. Nakamura, M. Shiro and M. Shionoya, *Angew. Chem., Int. Ed.*, 2009, **48**, 7006; S. Hiraoka, K. Harano, M. Shiro and M. Shionoya, *J. Am. Chem. Soc.*, 2008, **130**, 14368.
- J.-s. Y. T. Ma, S.-l. Lou, X.-q. Tang and Y.-d. Jiang, *Chem. Res. Chin. Univ.*, 2009, **25**, 590; K. R. J. Thomas, M. Velusamy, J. T. Lin, C. H. Chuen and Y.-T. Tao, *J. Mater. Chem.*, 2005, **15**, 4453.
- A. B. S. Elliott, R. Horvath, X.-Z. Sun, M. G. Gardiner, K. Müllen, N. T. Lucas, M. W. George and K. C. Gordon, *Inorg. Chem.*, 2016, **55**, 4710.
- D. Nolan, B. Gil, L. Wang, J. Zhao and S. M. Draper, *Chem. – Eur. J.*, 2013, **19**, 15615; F. A. Murphy and S. M. Draper, *J. Org. Chem.*, 2010, **75**, 1862; N. T. Lucas, H. M. Zareie and A. M. McDonagh, *ACS Nano*, 2007, **1**, 348; Y. Fogel, M. Kastler, Z. Wang, D. Andrienko, G. J. Bodwell and K. Müllen, *J. Am. Chem. Soc.*, 2007, **129**, 11743; B. E. Hamaoui, L. Zhi, J. Wu, J. Li, N. T. Lucas, Ž. Tomović, U. Kolb and K. Müllen, *Adv. Funct. Mater.*, 2007, **17**, 1179; K.-Y. Kim, S. Liu, M. E. Köse and K. S. Schanze, *Inorg. Chem.*, 2006, **45**, 2509; B. E. Hamaoui, F. Laquai, S. Baluschev, J. Wu and K. Müllen, *Synth. Met.*, 2006, **156**, 1182; P. T. Herwig, V. Enkelmann, O. Schmelz and K. Müllen, *Chem. – Eur. J.*, 2000, **6**, 1834.
- D. Nolan, B. Gil, F. A. Murphy and S. M. Draper, *Eur. J. Inorg. Chem.*, 2011, 3248.
- C. Delaney, G. M. Ó. Máille, B. Twamley and S. M. Draper, *Org. Lett.*, 2016, **18**, 88; L. P. Wijesinghe, B. S. Lankage, G. M. O. Maille, S. D. Perera, D. Nolan, L. Wang and S. M. Draper, *Chem. Commun.*, 2014, **50**, 10637; D. J. Gregg, C. M. A. Ollagnier, C. M. Fitchett and S. M. Draper, *Chem. – Eur. J.*, 2006, **12**, 3043; D. J. Gregg, C. M. Fitchett and S. M. Draper, *Chem. Commun.*, 2006, 3090; D. J. Gregg, E. Bothe, P. Höfer, P. Passaniti and S. M. Draper, *Inorg. Chem.*, 2005, **44**, 5654; S. M. Draper, D. J. Gregg, E. R. Schofield, W. R. Browne, M. Duati, J. G. Vos and P. Passaniti, *J. Am. Chem. Soc.*, 2004, **126**, 8694; S. M. Draper, D. J. Gregg and R. Madathil, *J. Am. Chem. Soc.*, 2002, **124**, 3486.
- S. Hiraoka, Y. Yamauchi, R. Arakane and M. Shionoya, *J. Am. Chem. Soc.*, 2009, **131**, 11646; K. Harano, S. Hiraoka and M. Shionoya, *J. Am. Chem. Soc.*, 2007, **129**, 5300; S. Hiraoka, K. Harano, M. Shiro, Y. Ozawa, N. Yasuda, K. Toriumi and M. Shionoya, *Angew. Chem., Int. Ed.*, 2006, **45**, 6488.
- V. V. Rostovtsev, L. G. Green, V. V. Fokin and K. B. Sharpless, *Angew. Chem., Int. Ed.*, 2002, **41**, 2596; C. W. Tornøe, C. Christensen and M. Meldal, *J. Org. Chem.*, 2002, **67**, 3057.
- B. Schulze and U. S. Schubert, *Chem. Soc. Rev.*, 2014, **43**, 2522; J. D. Crowley and D. A. McMorran, in *Click Triazoles*, ed. J. Košmrlj, Springer Berlin Heidelberg, Berlin, Heidelberg, 2012, p. 31; H. Struthers, T. L. Mindt and R. Schibli, *Dalton Trans.*, 2010, **39**, 675; D. Schweinfurth, N. Deibel, F. Weisser and B. Sarkar, *Nachr. Chem.*, 2011, **59**, 937.
- T. Y. Kim, A. B. S. Elliott, K. J. Shaffer, C. J. McAdam, K. C. Gordon and J. D. Crowley, *Polyhedron*, 2013, **52**, 1391.
- J. D. Crowley and E. L. Gavey, *Dalton Trans.*, 2010, **39**, 4035.
- Palladium in Organic Synthesis*, ed. J. Tsuji, Springer Berlin Heidelberg, 2005.
- D. Huang, P. Zhao and D. Astruc, *Coord. Chem. Rev.*, 2014, **272**, 145.
- D. Wang, L. N. S. Gautam, C. Bollinger, A. Harris, M. Li and X. Shi, *Org. Lett.*, 2011, **13**, 2618; W. Yan, X. Ye, N. G. Akhmedov, J. L. Petersen and X. Shi, *Org. Lett.*, 2012, **14**, 2358.
- L. Wang, D. J. Kiemle, C. J. Boyle, E. L. Connors and I. Gitsov, *Macromolecules*, 2014, **47**, 2199.
- H. V. Huynh, Y. Han, R. Jothibasu and J. A. Yang, *Organometallics*, 2009, **28**, 5395.
- T. Iwasawa, M. Tokunaga, Y. Obora and T. Tsuji, *J. Am. Chem. Soc.*, 2004, **126**, 6554; J. Cookson, *Platinum Met. Rev.*, 2012, **56**, 83; L. Li, T. Wu, J. Wang and R. Wang, *ChemPlusChem*, 2014, **79**, 257.
- C. Deraedt, L. Salmon and D. Astruc, *Adv. Synth. Catal.*, 2014, **356**, 2525; M. Gómez-Martínez, E. Buxaderas, I. M. Pastor and D. A. Alonso, *J. Mol. Catal. A: Chem.*, 2015, **404–405**, 1; S. S. Kotha, N. Sharma and G. Sekar, *Tetrahedron Lett.*, 2016, **57**, 1410; B. F. Machado and P. Serp, *Catal. Sci. Technol.*, 2012, **2**, 54; M. Pérez-Lorenzo, *J. Phys. Chem. Lett.*, 2012, **3**, 167.

