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Supramolecular metallogels with bulk self-healing properties prepared by *in situ* metal complexation

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In this feature article, we discuss a series of contributions dealing with the *in situ* fabrication of supramolecular metallogels (*i.e.* using low molecular weight ligands and metal ions) that show self-healing properties of the bulk gel phase after complete physical segregation. Most of the advances in this area have taken place during the last three years and are mainly represented by organogels, whereas examples of hydrogels and organic-aqueous gels are still a minority. *In situ* gelation via metal-coordination of low molecular weight compounds is conceptually different from the use of premade (*e.g.* in solution) coordination polymers and polymeric structures as gelators and ligands, respectively. In the case of *in situ* gelation, the cooperative effects of all components of the mixture (*i.e.* ligand, metal ion, counterions and solvent molecules) in an appropriate ratio under well-defined experimental conditions play a crucial role in the gelation phenomenon and self-healing properties of the material.

Introduction

Gels prepared in organic solvents (organogels), water (hydrogels) or aqueous organic mixtures represent a very productive research

field^{1–7} due to their large number of applications in many areas including regenerative medicine, drug-delivery, sensors, actuators, cosmetics, foods, environmental remediation, nanoelectronics ('bottom-up' approaches), and catalysis, among others.^{8–11} However, as it has observed with other materials with significant industrial potential, the development of high-tech applications usually takes place much faster than the acquisition of full knowledge of the mechanisms governing the key physical and/or chemical processes that define the properties of the materials.^{12,13}

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David Díaz Díaz

David Díaz Díaz received his PhD in Chemistry from the University of La Laguna (Spain). In 2002, he joined Prof. Finn's group as a postdoc at The Scripps Research Institute (USA). Since 2006, he has held various positions in academia and industry ('Ramón y Cajal' Researcher, UAM, Spain, 2006; Sr. Chemist, Dow Chemical, Switzerland, 2007; Tenured Scientist, CSIC, Spain, 2009; Alexander von Humboldt Experienced Researcher, University of Regensburg, Germany, 2010). In 2013, he was awarded the DFG Heisenberg Professorship and appointed as Associate Professor at the University of Regensburg. Among other distinctions, he has received the Young Investigator Award from the Polymer Network Group and is the Editor-in-Chief of Gels. His main research interest focuses on the development and understanding of functional soft materials.



Gel-based materials are not an exception to the rule because a large number of new gels are still nowadays discovered by serendipity. However, considerable efforts have been made during the last few decades to gain a deeper understanding of the gelation phenomenon, allowing the development of a general set of rules for the rational design of new molecules with the ability to form gels.¹⁴

Among different criteria, the classification of gels is usually based on the driving forces for molecular aggregation. This allows a fundamental distinction between chemical gels^{15,16} (*i.e.* gels based on covalent bonds, usually crosslinked polymers unable to redissolve) and physical gels^{17–25} (*i.e.* gels made of either low molecular weight (LMW)) molecules or polymers, which are based on non-covalent bonds, mainly hydrogen-bonding, van der Waals, charge-transfer, dipole–dipole, π – π stacking, and coordination interactions, which typically lead to reversible gel-to-sol phase transitions. Furthermore, systems based on both types of connections have also been described.^{26,27} The solid-like appearance of a gel results from the entrapment of the liquid (major component in the mixture) in the interstices of a solid 3D matrix with a large surface area (minor component of the mixture), frequently ascribed to surface tension and capillary forces.²⁸ The formation of the solid matrix is a consequence of the entanglement of 1D polymeric (or suprapolymeric) strands of macromolecules (or LMW compounds) – so-called gelators – typically with nanometer scale diameters and micrometer scale lengths through the above described forces. The immobilization of liquid molecules is governed by a metastable dissolution–crystallization equilibrium and usually increases the viscosity of the medium by a factor of 10.¹⁰ Overall, these viscoelastic materials may be considered hard or soft based on their rheological characteristics. All these characteristics provide many gels with the potential to respond to a variety of external stimuli^{29,30} including temperature, light, pH, mechanical action, electro/magnetic fields and/or chemical stimuli, which are manifested as remarkable changes from the molecular (*e.g.* hierarchical order, conformational state) to the macroscopic level (*e.g.* surface properties, shape).^{31–33}

On the other hand, one of the most impressive properties of biological materials is their ability to regenerate and heal after the infliction of physical damage. The origin of some natural self-healing processes (*e.g.*, merging of broken bones, cell-wall recovery upon deformation of wood) can be found in the presence of sacrificial bonds between molecules, which by breaking and reforming allow both a better adaptation of biomaterials to external conditions and the repair of damaged materials.^{34–37} This has inspired many scientists worldwide to develop a variety of materials,^{38–42} including gels, with self-healing properties through different strategies.^{43–50} Among different gels that exhibit self-healing properties in bulk,^{51–60} considerable attention has lately been drawn towards a specific type of gels (made either in organic solvents, water or aqueous organic mixtures) containing metal ions, so-called metallogels,^{61–67} due to the reversibility nature of metal coordination.⁶⁸ Besides the tunable coordination binding strength, the redox, electronic, magnetic and optical properties of metal ions are promising

features for numerous applications of metallogels in materials science.^{69–77} The majority of contributions in the field of self-healing metallogels deal with the use of polymeric ligands. During the preparation of this article, a comprehensive article covering the preparation and properties of coordination polymer gels^{78–88} was published by Sutar and Maji,⁸⁹ although it contains only two examples of self-healing supramolecular gels. In this context, during the last few years an increasing number of examples describing self-healing metallogels obtained by *in situ* gelation during metal complexation of LMW ligands have been reported. It is important to realize that *in situ* gelation using metal–LMW ligand coordination is significantly different from the use of premade (*e.g.* in solution) coordination polymers or polymeric structures as gelators or ligands, respectively. In the case of *in situ* gelation, the cooperative effects of all components of the mixture (*i.e.* ligand, metal and solvent molecules) in an appropriate ratio and under suitable experimental conditions play a crucial role in the gelation phenomenon and self-healing properties of the material.

Herein, we focus on this particular group of *in situ* formed metallogels, using LMW compounds, showing self-healing properties of the bulk gel phase after physical segregation (the term self-healing is used from now on under this description). The use of polymeric substrates or premade metal coordination polymers to form gels, including self-healing gels,^{90–104} is out of the scope of this contribution. In addition, those examples in which only rheological thixotropic loop tests are provided but not evidence of self-healing of the bulk material^{105–108} are not included in this article.

Metallohydrogels

In 2014, the group of A. Banerjee¹⁰⁹ described the first supramolecular *in situ* formed amino acid-based metallohydrogels with self-healing properties. These materials were prepared by combining a series of amphiphilic tyrosine derivatives, bearing alkyl chains of different lengths (1–3, Fig. 1) with NiCl₂. In particular, 1–3 (10 mM) were first dissolved in aqueous phosphate buffer (pH 7–8) by heating and subsequently mixed with a NiCl₂ solution, which turned the solution into an initial white suspension. Clear and stable metallohydrogels were obtained after sonication of the suspension and resting for a few minutes for 3, 3 h for 2 and 18 h for 1.

The authors found that a gelator:Ni²⁺ ion ratio between 1 : 1 and 1 : 0.3 was necessary in order to ensure complete gelation of the liquid phase and obtain gels with higher thermal stability. As a control experiment regarding the effect of the substitution of the amino acid on gelation, compound 4 was also synthesized and showed no gelation ability under similar conditions. It is also remarkable to note the apparent non-influence of the counterion of the metal salt on the gelation properties as well as the high metal-selectivity showed by 1–3, being unable to form any gel in the presence of other divalent metals such as Zn²⁺, Cu²⁺, Co²⁺, Mn²⁺, Fe²⁺, and Hg²⁺. The results of XRD studies indicated the formation of nickel hydroxide after addition



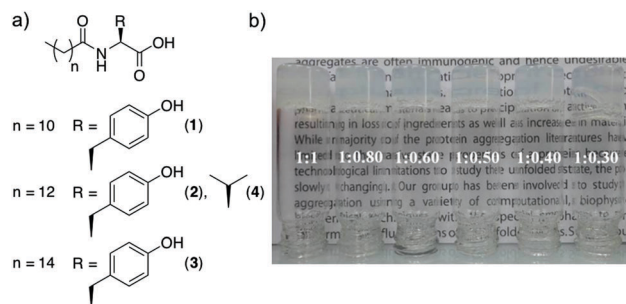


Fig. 1 (a) Chemical structures of various amphiphiles (**1–4**). (b) Pictures of glass vials containing metallohydrogels obtained from different proportions of **3** and NiCl₂. Critical gelation concentrations (CGC) = 10 mM (**1**), 9 mM (**2**), and 7 mM (**3**). Adapted with permission from ref. 109. Copyright©2014 The Royal Society of Chemistry.

of NiCl₂ into the phosphate buffer solution of **1–3** (pH 7.46). Moreover, comparative FT-IR analyses of the non-gel phase (in the absence of Ni²⁺) and in the gel phase (in the presence of Ni²⁺) suggested that the phenolic hydroxyl group of **1–3** strongly interacts with the hydrated nickel complex, whereas other functionalities (*i.e.* carboxylate and amide) of the ligand molecules are likely involved only in hydrogen bonding between the tape-like nanofibrillar assemblies observed by transmission electron microscopy (TEM), which was in accord with the results of small angle X-ray scattering (SAXS) experiments (Fig. 2).

Despite the small differences in the length of the alkyl chains in **1–3**, the rigidity of these metallohydrogels increased by increasing the side chain (**1** < **2** < **3**) as confirmed by oscillatory rheological measurements. Besides responding to multiple stimuli,¹¹⁰ *i.e.* pH, temperature, mechanical stress and metal chelating agents (Fig. 3a), the metallohydrogels exhibited remarkable self-healing properties. Interestingly, the kinetics of the self-healing process was correlated with the stiffness of

the gels (*i.e.* the gel made of **3** (*n* = 14) recovered immediately, whereas those made of **1** (*n* = 10) and **2** (*n* = 12) needed 8 and 5 min, respectively). However, further investigations are still necessary to explain this apparent relationship. Macroscopic self-healing of all metallohydrogels was demonstrated by cutting the gel blocks into two pieces and subsequently rejoining them with moderate pressing. It was found that after 25–30 min the two pieces merged into a continuous monolithic gel (Fig. 3b). The so-obtained self-healed monoliths displayed the same stability as uncut monoliths and could be stretched by pulling the ends, demonstrating the efficiency of the self-healing phenomenon. The recovery of the gel network phase (storage modulus $G' > \text{loss modulus } G''$) and its mechanical properties were also confirmed at the microscale by strain-recovery rheological experiments with alternating low and high strain cycles to ensure rupture of the gel network ($G' < G''$) (Fig. 3c) – note that while bulk self-healing of gels is usually accompanied by a thixotropic behavior at the microscale, the reverse is not necessarily true. The kinetics of the moduli recovery was again confirmed to increase with the increase in the length of the alkyl chain. Thus, the fine control achieved over several properties of these metallohydrogels by small changes in the alkyl chain length encourages us to keep in mind these results when investigating other amphiphilic LMW coordinating gelators with potential self-healing properties.

Almost at the same time, cooperation between the group of R. Banerjee and our group¹¹¹ resulted in the publication of the second example of a self-healing amino acid-derived metallohydrogel. Simply mixing two stock aqueous solutions containing a L-valine derivative (*i.e.* **V** = L-3-methyl-2-(pyridine-4-yl-methylamino)-butanoic acid), sodium perchlorate (**P**) and Zn(OAc)₂·2H₂O (**ZA**) leads to the instantaneous formation of an opaque hydrogel (**ZAVP**) with a CGC of 84 g L^{−1} (Fig. 4a–c). Similarly to the previous example, the ligand system **VP** showed specific affinity for metal cation (*i.e.*, Zn²⁺), whereas the effect of counter anions was negligible, with the possibility of obtaining similar gels with zinc perchlorate or zinc nitrate. As observed with other chiral gel systems,¹¹² the enantiomeric purity of the amino acid derivative had a major influence on the gelation ability. Specifically, complete gelation was observed when the enantiomeric excess (*ee*) was between 85 and 100%, whereas only partial gelation took place for *ee* between 60 and 80%.

In terms of the gelation mechanism, mass spectrometry studies of the corresponding xerogels suggested the direct involvement of a reversible octahedral Zn(II)-complex with two carboxylates of one acetate and one ligand molecule forming an octahedral complex bearing two water molecules in the axial positions. This complex may undergo consecutive self-assembly into fibrils (*ca.* 20–80 nm in diameter) through multiple hydrogen-bonding interactions, which entangle into a 3D fibrillar network as confirmed by scanning electron microscopy (SEM) imaging of the xerogels (Fig. 4d). Furthermore, rheological measurements confirmed the viscoelastic and brittle nature of the metallohydrogel. Strain-recovery experiments with alternating low and high strain cycles of the **ZAVP** hydrogel revealed full recovery of the gel network within 3 ± 1 min (Fig. 5). The cycles

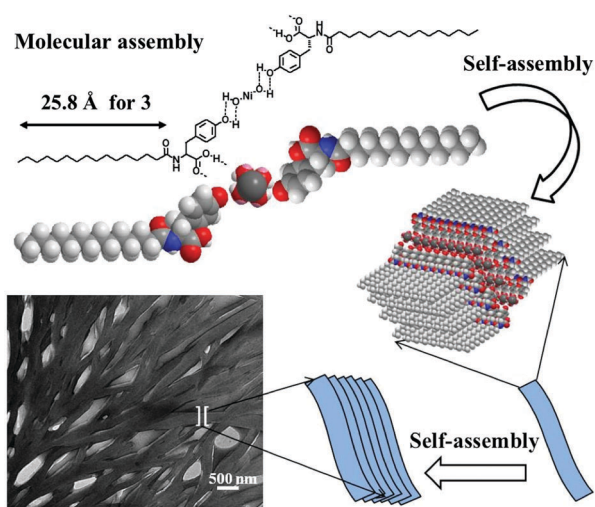


Fig. 2 Schematic representation illustrating the proposed stepwise assembly of the gelator molecules to form a tape-like network structure. The widths of the nanofibrils were in the range 60–80 nm, 60–70 nm and 120–150 nm for gels **1**, **2** and **3**, respectively. Adapted with permission from ref. 109. Copyright©2014 The Royal Society of Chemistry.



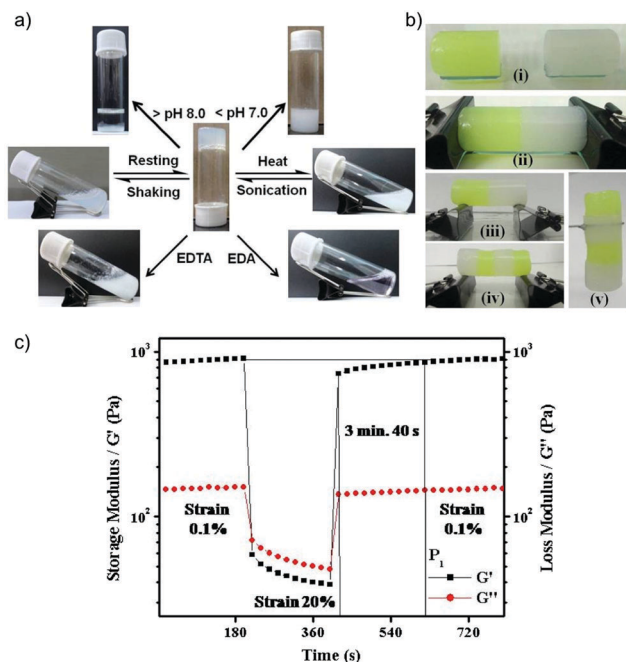


Fig. 3 (a) Multi-stimuli responsiveness shown by the hydrogel obtained from **3**. Above the gel-to-sol transition temperature (T_{gel}), all the gels turned into solutions and nickel hydroxide was precipitated out. These suspensions were transformed into metallohydrogels upon cooling them to room temperature followed by sonication. These metallohydrogels also showed sensitivity to metal chelating agents like ethylene diamine (EDA) or ethylenediaminetetraacetic acid (EDTA). (b) Illustration of self-healing behavior (i–v) shown by the metallohydrogel obtained from **3**. (c) Step strain experimental data obtained from the metallohydrogel made of **1**. The concentration was maintained at 10 mM at a 1:0.5 gelator and Ni^{2+} ratio. The angular frequency was constant at 1 rad s^{-1} in all steps. Adapted with permission from ref. 109. Copyright©2014 The Royal Society of Chemistry.

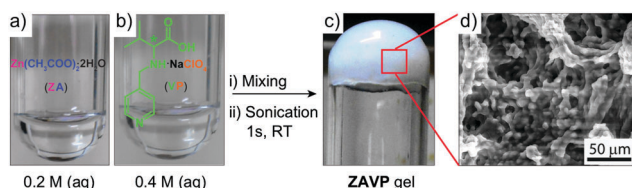


Fig. 4 (a–c) Synthesis of ZAVP at the CGC. Total volume of each solution = 0.5 mL. (d) Scanning electron microscopy (SEM) image of the corresponding xerogel. Adapted with permission from ref. 111. Copyright©2014 The Royal Society of Chemistry.

could be repeated at least 15 times without notable reduction of the average magnitude of G' and G'' , displaying a remarkable and rapid thixotropic behavior, which was in good agreement with the macroscopic thixotropy observed by shaking–resting cycles.

Also in this example, and in addition to a multi-responsive behavior towards external stimuli (*i.e.* temperature, pH, mechanical stress and metal complexing agents),¹¹¹ the obtained bulk metallohydrogel exhibited remarkable self-sustainability, load-bearing and self-healing properties. The gel was able to support up to 60-fold of its own weight without any visible change in its shape (Fig. 6a–c), which is not often found for LMW gels.¹¹³ Furthermore, bulk

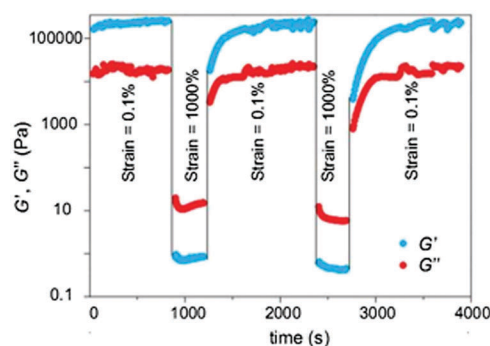


Fig. 5 Oscillatory rheology of the ZAVP hydrogel prepared at the CGC, thixotropy-loop test via continuous step-strain measurements at 0.1 Hz. Reprinted with permission from ref. 111. Copyright©2014 The Royal Society of Chemistry.

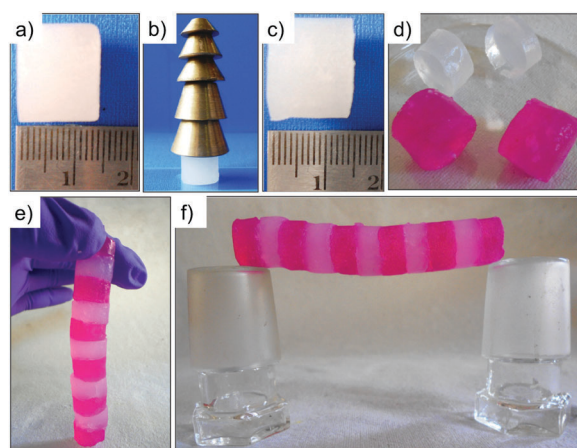


Fig. 6 Pictures of (a) freshly prepared hydrogel cylinder (thickness = 1.4 cm; diameter = 1.6 cm); (b) 2 g mass bearing a load up to 90 g; and (c) maintaining its shape after the load is released. (d) Moldable and shape-persistent hydrogel blocks cut from a larger cylinder. Pink objects correspond to hydrogel pieces obtained upon doping using Rose Bengal dye. (e and f) A bridge constructed by fusing 11 hydrogel blocks together. Light pink coloration of undoped gel blocks is due to the rapid diffusion of the dye through the self-healed hydrogel network. Adapted with permission from ref. 111. Copyright©2014 The Royal Society of Chemistry.

self-healing was demonstrated by cutting a gel monolith into small pieces and rejoining them by simple physical contact. Complete fusion of the blocks took place within 3 h allowing the reconstruction of long and stable self-healed gel bridges (Fig. 6d–f).

Metalloorganogels

The synthesis of macromolecules containing bipyridine and terpyridine metal complexes has gained considerable attention during the last two decades due to the large number of potential applications showed by these materials.¹¹⁴ Within the context of this structural moiety, Terech and co-workers¹¹⁵ used a tritopic ligand (BTC) (Fig. 7) exhibiting two terpyridine functionalities and a cyclan spacer to form weak (*i.e.* $G' \text{ ca. } 250 \pm 15 \text{ Pa}$ at 1 Hz) and self-healable metalloorganogels *in situ* when combined with

$\text{NiCl}_2 \cdot 6\text{H}_2\text{O}$ in DMF at 50°C for 24 h. The optimal stoichiometry was found to be $\text{Ni}^{2+}/\text{BTC} = 1.8$. Above a 2.0 molar ratio, the ligand was found to complex additional Ni^{2+} ions and showed fragmentation of the aggregates.

Visual inspection of the Ni_2BTC gel after macroscopic damage revealed its self-healing ability. The gel was damaged with a spatula and self-healed within 48 h (Fig. 8a), whereas a classical molecular gel made of 12-hydroxy stearic acid (HAS) with a similar elastic modulus (*ca.* 220 Pa) under the same treatment did not show any visual healing of the wound. Step strain tests were performed with high stresses up to the yield stress in the non-linear regime of deformations. The system was sheared for *ca.* 30 s and showed instantaneous recovery (albeit not completely) of the elasticity after removal of the applied stress. Moreover, the application of several cycles showed a progressive fatigue of the material, which was related to a decrease of its mean connectivity. Applying different protocols, a small and controlled deformation step in an oscillatory mode within the linear regime of deformations was followed by a shearing step where the gel was disrupted. Again, a significant part of the shear elasticity was recovered but not entirely (Fig. 8b). Moreover, impedance spectroscopy was performed in order to obtain an insight into the healing mechanism. The results confirmed the coexistence of two main dynamical processes: (1) a fast complexation reaction of Ni^{2+} with BTC (association constant $K = 5.0 \times 10^{10} \text{ M}^{-1}$) and (2) a slow repairing step involving optimization of the 3D interconnectivity of the transient gel network.

Following their previous work on dynamic supramolecular polymers based on host-guest and metal-ligand interactions, Huang and co-workers¹¹⁶ described the use of heteroditopic monomer **5** with a bridging ligand and a metallic cross-linker, $[\text{PdCl}_2(\text{PhCN})_2]$ (0.5 equiv.), for the *in situ* preparation of a new self-healable metalloorganogel. Heteroditopic monomer **5** contained a 1,2,3-triazole group between the benzo-21-crown-7 host and its complementary guest dialkylammonium salt and the 1,2,3-triazole group acted as a ligand for coordination with Pd(II). With regard to the gel preparation, **5** was first dissolved in acetonitrile and then addition of 0.5 equiv. of $[\text{PdCl}_2(\text{PhCN})_2]$

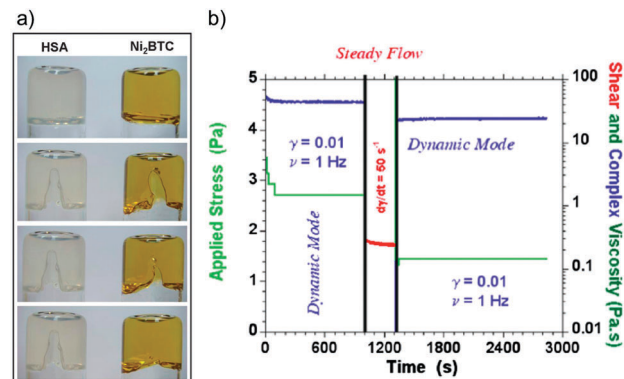


Fig. 8 (a) Photographs of upside-down vials containing the Ni_2BTC gel ($s = 1.8$, $C_{\text{BTC}} = 40 \text{ mM}$ in DMF) and the HAS gel used as reference ($C_{\text{HSA}} = 8.0 \text{ mM}$ in decane). From top to bottom: at rest; fracture; 24 hours after fracture; 48 hours after fracture, respectively. (b) Cycles of a small deformation (*ca.* 0.01, frequency 1 Hz) followed by a shearing protocol at $dg/dt = 50 \text{ s}^{-1}$ for 300 seconds during which the gel is fractured showing the partial self-healing of the system. Adapted with permission from ref. 115. Copyright©2014 The Royal Society of Chemistry.

and subsequent heating to 50°C led to a transparent solution, which spontaneously turned into a shape-persistent, free-standing gel upon cooling at 4°C (Fig. 9).

Self-aggregation of **5** to form linear or cross-linked supramolecular polymers by addition of Pd(II) was confirmed by 2D diffusion-ordered ^1H -NMR spectroscopy (DOSY) which proved an increase in the average aggregation size by addition of the metal salt. A sol-to-gel transition was induced by adding a competitive ligand (PPh_3) for Pd(II), which eliminated the cross-linking. SEM imaging of the corresponding xerogels showed an interconnected porous structure of entangled nanofibers (Fig. 10) where the pore size could be modulated by the amount of cross-linker.

At the CGC (30 mM), the gel exhibited reversible gel-to-sol transitions in response to pH, temperature, and addition of cations or metals, which form new complexes with the host. In addition, the gel showed good mechanical properties (*i.e.* plateau modulus = 10^1 – 10^4 Pa based on rheological analysis), self-sustainability, moldability, elastomeric behavior and shape-persistent properties. Moreover, freshly prepared gel blocks were able to adhere to each other to construct a stable gel bridge made up of up to four gel blocks (Fig. 11a–c). Although the authors did not provide direct data regarding the reversibility of this process (*i.e.* cutting and rejoining the pieces as usually done to prove self-healing), they found that fresh surfaces of the gel blocks were essential for adhesion. This is a common requisite to most self-healable gels, which suggests that the mechanical stress applied to freshly prepared gels cause significant alteration of the solvated structure due to the solid-liquid interface susceptibility to mechanical stress.¹¹⁷ Thus, we decided to include this interesting example in this feature article because it made us realize that the differences, and likely cooperative effects, between bulk self-adhesion and self-healing processes are still not specifically addressed in most publications.

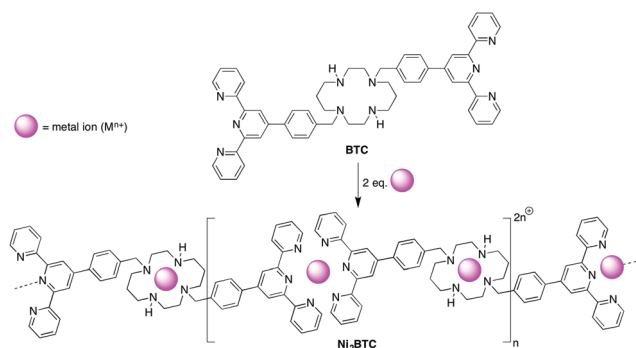


Fig. 7 Formation of Ni_2BTC suprapolymers through chelation of nickel ions (purple spheres) by BTC. The situation for stoichiometry $s = \text{Ni}^{2+}/\text{BTC} = 2$ is shown. Note that gelation takes place during the reaction without isolation of the proposed Ni_2BTC species. Adapted with permission from ref. 115. Copyright©2014 The Royal Society of Chemistry.



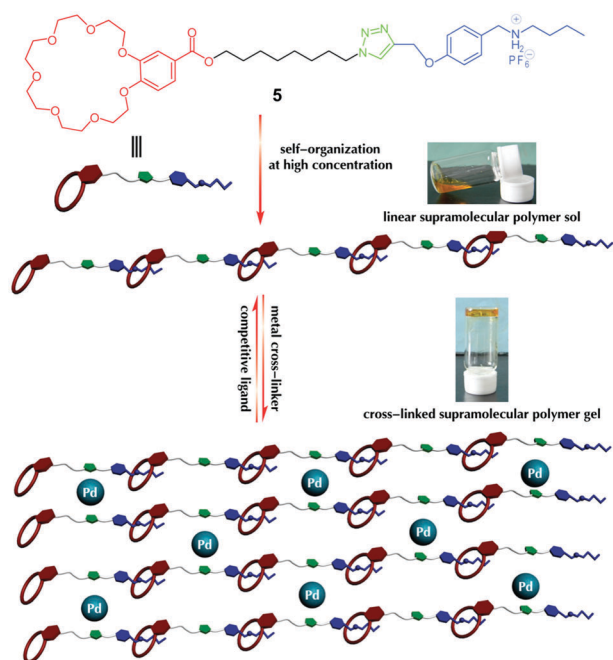


Fig. 9 Schematic representation of controlling the topology and rheology of the supramolecular gelator prepared from heteroditopic monomer **5**. Adapted with permission from ref. 116. Copyright©2012 Wiley-VCH Verlag GmbH & Co.

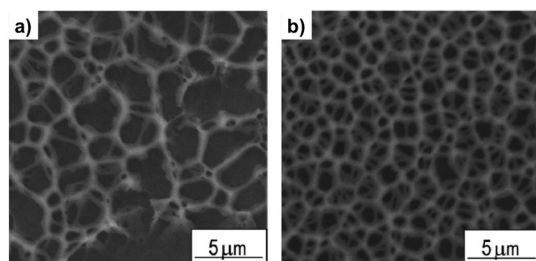


Fig. 10 SEM image of the supramolecular gel network with a cross-linking density of (a) 20% and (b) 60%. Adapted with permission from ref. 116. Copyright©2012 Wiley-VCH Verlag GmbH & Co.

Interestingly, the gel also showed excellent elastic properties, being able to rebound up to 20 cm after free-falling from a 25 cm height (Fig. 11d). This behavior can be attributed to the dynamically reversible complexation between the host and dialkylammonium salt and the good mechanical properties of the cross-linked 3D supramolecular network. In sharp contrast, the linear polymer (*i.e.* prepared in the absence of a cross-linker) did not show any elasticity, confirming the necessity of cross-linking to achieve elastic gels.

More recently, other N-containing heterocycles have been also explored as ligands for the preparation of supramolecular moldable and self-healing gels by *in situ* metal complexation. Thus, Yan and co-workers¹¹⁸ prepared a series of sticky metallogels by mixing two stock solutions of a series of tetrazole-based ligands (Fig. 12) and Pd(OAc)₂. Spontaneous gel formation occurred for ligands **6–9** at a 2:1 tetrazolyl unit: Pd²⁺

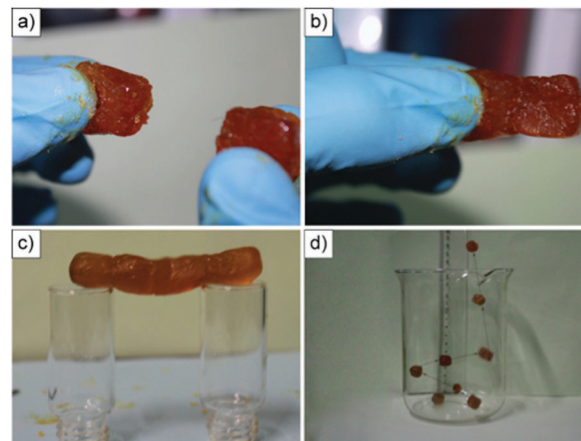


Fig. 11 (a) Two supramolecular gel blocks. (b) The two gel blocks adhered to each other to form a free-standing object. (c) A bridge constructed by connecting together four gel blocks suspended horizontally. (d) The moving trajectory of the gel blocks after rebounding from the bottom of the beaker (eight pictures overlaid together). Reprinted with permission from ref. 116. Copyright©2012 Wiley-VCH Verlag GmbH & Co.

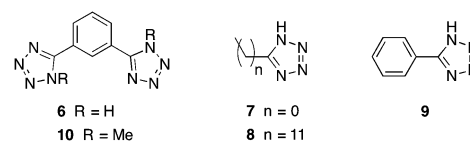


Fig. 12 Chemical structures of ligands **6–10**. Adapted with permission from ref. 118. Copyright©2014 The Royal Society of Chemistry.

molar ratio and a CGC of 3.0, 4.5, 3.6, and 3.9% (w/v), respectively. Methylation of **6** afforded compound **10** lacking H-donor ability, which inhibited the gelation process. TEM, atomic force microscopy (AFM) and dynamic light scattering (DLS) images revealed the initial formation of uniform spherical aggregates with an average diameter of around 30 nm, which further interconnected to form a more dense and entangled network.

The gel derived from ligand **6** was used as a model gel for further characterization. The viscoelastic and brittle nature of the gel was confirmed by typical rheological measurements. This metallogel displayed a remarkable stability towards a broad range of pH (1–14), heating (*ca.* 100 °C), shaking or sonication without a visible gel-to-sol transition, suggesting strong intermolecular interactions. The robustness of the gel allowed its molding into self-supporting blocks of any geometrical shape (Fig. 13).

Captivatingly, the gel body was able to support up to 30-fold its own weight by unprecedented bottom-up load-bearing properties due to its sticky character (Fig. 14). This deserves attention because usually up-down load-bearing experiments are reported for supramolecular gels.¹¹⁹

Self-healing properties were also confirmed by the construction of a 3 cm self-supporting gel bridge after cutting gel monoliths and placing the pieces into contact for several minutes at room temperature (Fig. 15). The obtained bridge was stable enough to



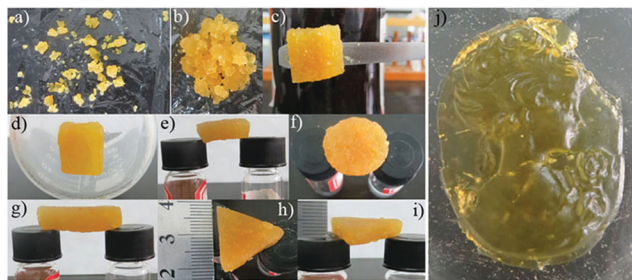


Fig. 13 (a–i) Shape-persistent, free-standing objects moulded from a gel made of **6** (4%). (j) The gel sculpture made from 4% DMF solution of the gel made of **6**. Adapted with permission from ref. 118. Copyright©2014 The Royal Society of Chemistry.

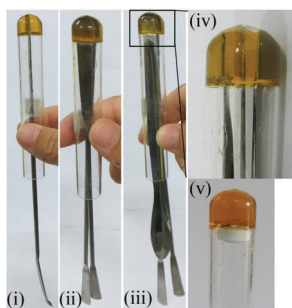


Fig. 14 Gel made of **1** (4%) showing bottom-up load-bearing properties. One to three spatulas (ca. 26 g) were inserted into the gel body and suspended just by holding the test tube. Adapted with permission from ref. 118. Copyright©2014 The Royal Society of Chemistry.

hold when suspended horizontally or vertically from a wire. The same self-healing capability was also found for the gels made from ligands **7–9**. In addition, the gels can be restored after any damage by adding DMF or H₂O and self-healing can be accelerated by sonication due to a faster permeating speed of the solvent. However, solvent-healed gels lost their viscoelastic response and only 30% can be recovered after aging for several days.

A year later, the same group¹²⁰ expanded their investigations using tetrazole-based ligands and reported for the first time a supramolecular self-healing heterometallic gel with synergistic

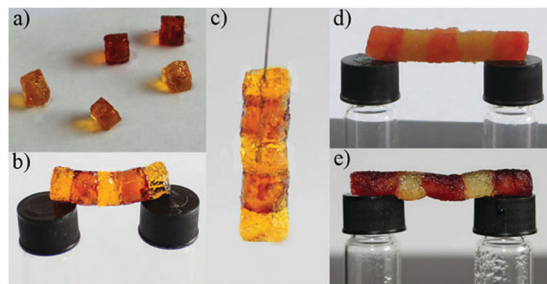


Fig. 15 (a and b) A bridge constructed by connecting freshly prepared blocks of gel **6** can be suspended horizontally and (c) held vertically, (d) a bridge constructed by connecting moulded gel blocks and (e) solvent-healed (DMF) gel blocks. Red objects correspond to metallogel blocks obtained upon doping with an aqueous solution of methyl red (ca. 0.1%). Reprinted with permission from ref. 118. Copyright©2014 The Royal Society of Chemistry.

properties. In this case, a DMF solution of simple ligand **11** was mixed with DMF solutions of NiCl₂·6H₂O and CoCl₂·6H₂O to spontaneously form homogeneous metallogels at Ni²⁺:ligand ratios from 1:1 to 3:1. AFM and TEM images confirmed that both monometallic gels (**11**-Ni and **11**-Co) self-assemble into spherical particles, whose interconnection creates a 3D entangled network (Fig. 16). In addition, the necessity of the –NH group in **11** for gelation was confirmed by the methylated derivative **12**, which was not capable of evolving into any gel in the presence of Ni²⁺ or Co²⁺.

Gel **11**-Ni with a CGC of 2.3 wt% exhibited high stability towards trifluoroacetic acid (TFA), heating (ca. 100 °C), shaking and sonication without a visible gel-to-sol transition. In contrast, Gel **11**-Co showed a reversible gel-to-sol transition by heating at 40 °C, which could be regulated by TFA/Et₃N. In agreement with rheological measurements **11**-Co exhibited an overall weaker stability than **11**-Ni (*i.e.* *G'* being ca. 20 times higher for **11**-Ni at a three-fold lower concentration than **11**-Co). The air stable and stiff **11**-Ni gel blocks were not able to adhere to each other, but exhibited an uncommon bouncing behavior, probably due to significant flexibility of the dynamically reversible non-covalent interactions. In sharp contrast, **11**-Co collapsed in air accompanied by a visible color change from blue to pink due to a transition from a tetrahedral complex (T_d) into an octahedral (O_h) complex induced by water, as supported by UV-vis spectroscopy (Fig. 17b and e). Interestingly, combination of both systems by adding 0.5 equivalents of NiCl₂·6H₂O into **11**-Co to prepare an air stable heterometallic gel **1**-Co-Ni(0.5) caused a 5-fold enhancement of the storage modulus compared to the initial value of the pristine gel **1**-Co. Strain-recovery tests of heterometallic gels prepared using different Co:Ni ratios showed that gels **11**-Co-Ni(0.2) and **11**-Co-Ni(0.5) were recovered within ca. 10–15 min after removal of the high oscillatory stress, whereas the gel **11**-Co-Ni(0.1) recovered almost immediately (Fig. 17c). Moreover, stable gel blocks allowed the preparation of 5 cm long self-supporting gel bridges by simply rejoining 12 blocks of **11**-Co-Ni(0.2) without any external stimuli (Fig. 17a). Due to the synergistic features of the constituent metal ions, visual self-healing was observed after several minutes and for multiple cycles with the wound formed in the same position.

Moving away from alkali and transition metals, Gunnlaugsson and co-workers¹²¹ have recently described the use of lanthanide

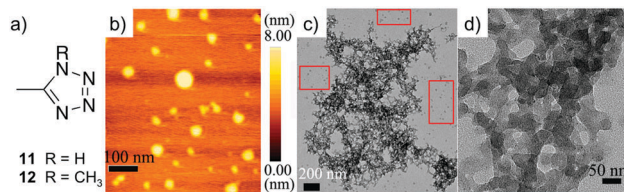


Fig. 16 (a) The chemical structures of ligands **11** and **12**. (b) Tapping-mode AFM height image of aggregates formed of metallogel **11**-Co (0.1 wt%). (c) TEM image of **11**-Co on a carbon coated copper grid (unstained, 0.3 wt%). (d) A partial zoomed-in image **11**-Co on a carbon coated copper grid (unstained, 0.3 wt%). Adapted with permission from ref. 120. Copyright©2014 The Royal Society of Chemistry.



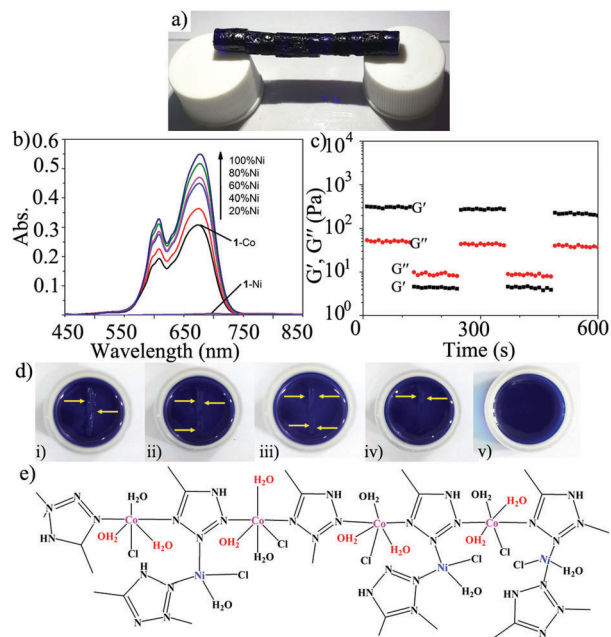


Fig. 17 (a) A self-healing bridge made of heterometallic gel **11**-Co-Ni(0.2). (b) The UV-vis spectra of **11**-Co with increased Ni^{2+} . (c) A loop test of **11**-Co-Ni(0.1) via continuous step-stress measurements at 1 Hz. **11**-Co-Ni(0.1) was subjected to 450 Pa stress for 2 min, and then again to 50 Pa for 2 min, and this process was repeated 3 times. (d) The third self-healing cycle of **11**-Co-Ni(0.2) (i) after damage, (ii–v) after free standing for 20 min, 40 min, 60 min and 2 h. (e) A schematic illustration of the Ni-supported network of heterometallic gel **11**-Co-Ni. Reprinted with permission from ref. 120. Copyright©2014 The Royal Society of Chemistry.

salts (*i.e.* $\text{Ln} = \text{Eu}^{3+}$, Tb^{3+}) to direct the self-assembly of a LMW ligand **H₂L** (*i.e.* **H₂L** = 4,40-(((pyridine-2,6-dicarbonyl)bis-(azanediyl))bis(methylene))-dibenzoic acid (Fig. 18a)) in MeOH affording different luminescent gels with self-healing properties. Although the authors carried out studies with the isolated metal complexes made in solution, they also found the exact conditions to achieve *in situ* gelation by adjusting the metal:ligand molar ratio, the nature of the counterion and the experimental conditions during the mixing of the reactants. $^1\text{H-NMR}$ titrations provided evidence of the involvement of terminal free carboxylic groups during the self-assembly process. The use of a **Ln(H₂L)₃:Ln(CH₃COO)₃** ratio of 1:0.5 afforded gels with a fibrillar morphology as observed by SEM imaging of the corresponding xerogels (Fig. 18b). However, these gels showed a weak nature as indicated by rheological measurements (*i.e.* G' below 1.2 kPa). The self-healing properties of the Ln-based gels were visualized by cutting the gels into two pieces, which reassembled without external stimuli (Fig. 18c–f). An excess number of free carboxylic acid groups from the ligands and a high coordination number of lanthanides were proposed to facilitate the self-healing process.

Despite the large number of fascinating supramolecular self-healing gels reported up to 2016, none of them had been proved to impart self-healing properties to a material generated upon combination with other networks. Within this context, our group¹²² very recently reported the first supramolecular metalloorganogel with self-healing properties and the ability to

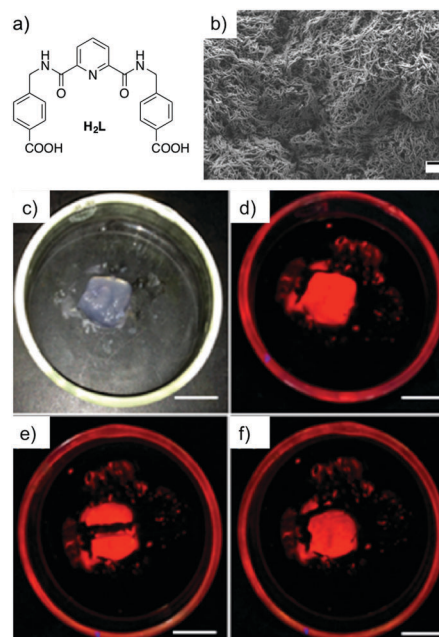


Fig. 18 (a) Chemical structure of ligand **H₂L**. (b) SEM image of the corresponding Eu(III)-based xerogel dried on a silicon plate (scale bar = 500 nm). Healing experiment of (c) Eu(III) gel in the day light, (d) the same gel under UV light, and (e) gel after being cut in half. (f) Self-healing properties of the gel (scale bars = 1 cm). Reprinted with permission from ref. 121. Copyright©2015 American Chemical Society.

induce self-healing in other organogel networks lacking this capacity. The metalloorganogel was prepared by mixing two equivalent DMF solutions of $\text{Cu}(\text{OAc})_2 \cdot \text{H}_2\text{O}$ (0.36 M) and oxalic acid dihydrate (0.66 M) at room temperature (Fig. 19a–d). The initially formed colloidal suspension was sonicated for 10 min affording a homogeneous gel (**CuA-Ox**) within 6 h. Previous experimental evidence demonstrated unprecedented proton conductivity for this material with a gelation mechanism involving the spontaneous formation of 1D coordination short oligomers (*i.e.* $n < 7$, $[\text{Ox}_n\text{Cu}_n + 1]]$), where Cu^{2+} is surrounded by four oxygen atoms of two oxalic acid molecules in a square-planar geometry.¹²³ Further intermolecular interactions between oxygen and copper atoms as well as hydrogen bonding between oxalic acid units may also play a key role in the overall entanglement trapping the solvent molecules in the interstices of the gel network. No gelation was achieved using commercially available copper oxalate hemihydrate, suggesting that all above-mentioned components must play a synergic role during the gelation phenomenon. The so-obtained gel monoliths displayed self-supporting properties and remarkable load-bearing capacity.¹²³ Moreover, the self-healing ability of the gel was demonstrated after cutting a gel monolith and reconnecting the pieces (Fig. 19e–j). The healing time could be reduced from 5 h to 10 min by applying 2 min sonication after reconnection of the pieces. It is important to emphasize that a self-healed **CuA-Ox** metalloorganogel was found to have essentially the same entwined nanofibrillar morphology (*i.e.* fibers width of *ca.* 20–50 nm), thermal stability (*i.e.* degradation of the coordination oligomer at *ca.* 300 °C), rheological properties (*i.e.* high G' values of *ca.* 10^5 Pa, low frequency dependence





Fig. 19 (a–d) Synthesis of 2 mL of monolithic **CuA-Ox** gel. (e–g) Self-supporting properties of monoliths. Rhodamine B was used as a dye to stain one monolith and for better visualization of the healing process. (h–j) Self-healing and load-bearing properties of a monolithic gel made by fusion of alternate dye doped and undoped block gels. Reprinted with permission from ref. 122. Copyright©2016 American Chemical Society.

$G' \approx \omega^{0.05}$; $\tan \delta \approx 0.12$), and proton conductivity (*i.e.* derived from the Nyquist plot, $\sigma \approx 1.4 \times 10^{-4} \text{ S cm}^{-1}$ at 27 °C under dehumidified conditions) as the freshly prepared monolithic sample.¹²³

Interestingly, this gel network showed the unprecedented ability to impart self-healing properties to other physical gel networks that inherently lack this capacity. A well-known diaminocyclohexane bis(amide)-based gelator (**DACBA**) was used to exemplify this behavior. Two strategies were possible, either by fusing alternating gel blocks of **CuA-Ox** and **DACBA** affording a stable gel bridge after 12 h or by preparing a (**DACBA**)-(CuA-Ox) hybrid gel through the simultaneous formation of both gel networks (Fig. 20). The lowest **DACBA**:**CuA-Ox** ratio necessary to prepare stable and self-healable hybrid monoliths was found to be 0.8:0.2 (v/v), namely hybrid-0.8. The effective interpenetration of both networks was suggested by comparing both mechanical and thermal properties of each network before and after hybridization. In brief, the average storage moduli G' of the gels **DACBA**, **CuA-Ox** and hybrid-0.8 were $2.1 \times 10^4 \text{ Pa}$, $7.5 \times 10^5 \text{ Pa}$ and $4.8 \times 10^4 \text{ Pa}$, respectively. Moreover, the T_{gel} of **DACBA** was established at 41 °C, whereas **CuA-Ox** remained stable until 150 °C and hybrid-0.8 until 120 °C. In terms of morphology, SEM imaging of the xerogel prepared from the hybrid gel revealed morphological features of both networks. It is worth mentioning that not only physical gels but also chemical gels could be hybridized with the **CuA-Ox** gel network to exhibit self-healing properties to some extent. This was verified using a polyelectrolyte gel based



Fig. 20 (a and b) Preparation of a self-supporting monolithic bridge by reconnection of alternate block gels made of **CuA-Ox** and **DACBA**. (c and d) Preparation of self-supporting and self-healing hybrid-0.8 (*i.e.* **DACBA**/**CuA-Ox** = 0.8:0.2 v/v) by the simultaneous formation of **CuA-Ox** and **DACBA** gel networks. (e) Hybrid-0.9. (f) Hybrid-0.95. (g) Spontaneous collapse of the gel made only of **DACBA**. Panels (d)–(f) correspond to the material obtained after cutting/reconnecting the gel (four pieces). Total volume in each case = 2 mL. Reprinted with permission from ref. 122. Copyright©2016 American Chemical Society.

on polyvinylidene difluoride (PVDF) and NaI in DMF instead of **DACBA**.¹²²

Furthermore, conductive composite black gels were also prepared by dispersing different carbon nanomaterials (*i.e.* graphite = 25 g L⁻¹, graphene oxide = 20 g L⁻¹, MWNT = 12.5 g L⁻¹, SWNT = 5 g L⁻¹) into an oxalic acid stock solution before mixing with a stock solution of copper acetate and sonicated for 2 min. The incorporation of the carbon nanomaterials within the gel network was also confirmed by electron microscopy imaging (Fig. 21a and b). The conductivity of the composite was demonstrated by connecting one monolithic composite to an electrical circuit where a LED was lit with a DC voltage of 3.0 V (Fig. 21c–f). Repeated cutting and reconnection of the composite blocks demonstrated the self-healing ability by restoration of the conductivity.

Metallogels in organic-aqueous solvent systems

In some cases, the solvent system used for the preparation of metallogels is neither pure water nor a pure organic solvent, but a combination of both. Thus, strictly speaking, these gels are neither hydrogels nor organogels regardless of the water: organic solvent ratio. In this section, we highlight self-healable metallogels formed *in situ* in water/organic solvent mixtures.



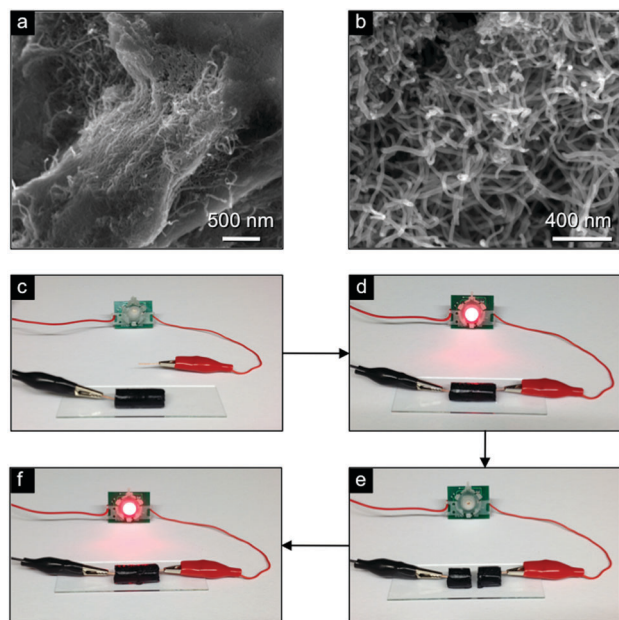


Fig. 21 (a and b) SEM images of MWNTs embedded in the **CuA-Ox** gel matrix. (c–f) Demonstration of the restoration of bulk conductivity of the **CuA-Ox**/MWNTs (10 g L^{-1}) composite gel used to bridge an electrical circuit. Adapted with permission from ref. 122. Copyright©2016 American Chemical Society.

Alcohols have been used as fuels with high octane rating for a long time. However, the liquid nature and high volatility are serious threats for expanding the use of these fuels. Therefore, strategies to formulate liquid alcohols into gel-like materials will be highly advantageous for easy handling, delivery and safety. Within this context, naturally occurring dipicolinic acid (*i.e.* it composes 5% to 15% of bacterial spores) is responsible for the heat resistance of bacterial endospores. This inspired Halder and Bera¹²⁴ to pursue research aiming to convert liquid alcohols into safe fuel gels aided by alkali metals. In particular, they described the instant formation of a strong self-healable transparent gel by mixing dimethyl dipicolinate (**13**) with 3 equiv. of KOH in (9:1) MeOH:H₂O or EtOH:H₂O solvent mixtures at room temperature (Fig. 22). The observed gelation was found to be highly selective towards K⁺ and OH[−] ions as well as the above-mentioned solvent systems. Very interestingly, due to its solvent selectivity the gel can be used to separate various miscible organic solvents in a green manner.¹²⁵ This was exemplified by adding dimethyl dipicolinate and aqueous KOH to a solvent mixture of EtOAc and MeOH. The gel was formed only with MeOH and precipitated out, which facilitated the separation of both solvents by decantation or simple filtration (MeOH can also be collected by distillation of the gel).

NMR, mass spectroscopy and powder X-ray diffraction (PXRD) measurements demonstrated the absence of ester hydrolysis of the dipicolinate during gelation and a sheet-like packing arrangement with KOH presumably acting as a crosslinker for the formation of monolayers. This simple and peculiar gel featured a helical tape-like nanofibrillar network as observed by field emission SEM (FE-SEM) imaging of the corresponding xerogel (Fig. 23a).

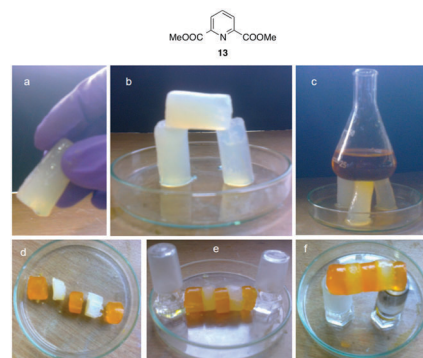


Fig. 22 Top: Chemical structure of ligand **13**. Bottom: (a) a gel cylinder taken in hand. (b) A gel gate made by fusing three stable gel cylinders. (c) Gel cylinders supporting the weight of several grams. (d) Alternate arrangement of methyl orange-doped and undoped gel cylinders. (e) Self-healing of the gel. (f) A cylinder made from five small gel discs suspended on two glass stoppers. Dye diffusion through the entire gel monolith demonstrated the dynamic exchange of dissolved molecules across the fusion supporting the self-healing process. Adapted with permission from ref. 124. Copyright©2015 The Royal Society of Chemistry.

However, in contrast to most chiral gelators, both left and right-handed helices were observed in this case. Moreover, rheological experiments demonstrated the stability and viscoelastic properties of the gel (*i.e.* the evolution of G' and G'' as a function of shear stress and strain at a frequency of 10 rad s^{-1} showed that G' was one order of magnitude larger than G'' with no crossover point). It is remarkable that gels prepared at 1 wt% were stable to handle with tweezers, could be shaped into any self-supporting geometry by using appropriate molds, and could support a weight of several grams (Fig. 22c). The self-healing ability of the gels was also demonstrated by fusing several gel blocks to each other after cutting the monoliths into pieces (Fig. 22d–f). Although a thixotropic behavior was also shown during loop strain-recovery rheological experiments, the storage modulus gradually decreased after each cycle under the tested conditions (Fig. 23b).

Due to high stability at high temperatures and the lack of thermo-reversibility the gel was tested as a potential candidate for a safer fuel. The gel made using EtOH/H₂O displayed 1300 kJ mol^{-1} heat of combustion compared to 726 kJ mol^{-1} for the gel made using MeOH/H₂O. Interestingly, after burning, the gels can be reused up to four times by addition of (9:1) MeOH:H₂O. Unfortunately, after that the gelation efficiency decreased with each cycle due to basic hydrolysis of the ester groups at high temperatures (Fig. 24). However, as the gels are not thermoreversible, they remain in a gel phase at the time of burning and take longer times than other standard heating-cooling alcoholic gels to be consumed.

Very recently, the group of Bhattacharjee¹¹⁶ described self-healing metallogels with excellent selectivity for dye adsorption and separation. Opaque white metallogels were prepared by reacting the disodium salt (**Na₂HL**) of a multidentate ligand (*i.e.* **H₃L** = carboxymethyl-(3,5-di-*tert*-butyl-2-hydroxybenzyl)amino acetic acid), as previously described by the same group,^{127,128} with Cd(II) and Zn(II) halides in an aqueous methanolic solution at room temperature for 1 min. Depending on the metal salt,



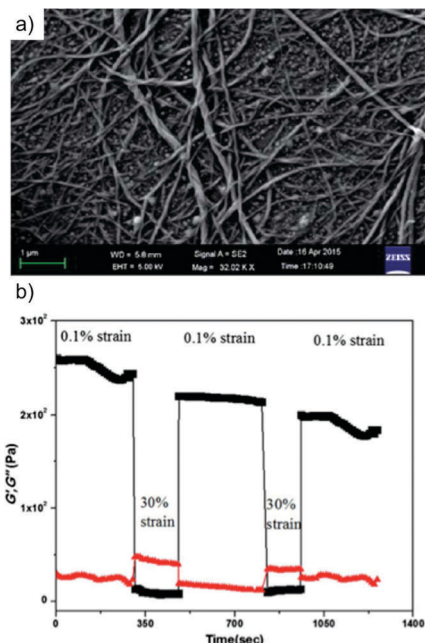


Fig. 23 (a) FE-SEM image of the xerogel prepared from the corresponding gel made of KOH responsive compound **13** in 1:9 H₂O:MeOH at 0.5% w/v (CGC). (b) Loop step strain experiment of the previous gel at a frequency of 10 rad s⁻¹. Adapted with permission from ref. 124. Copyright©2015 The Royal Society of Chemistry.

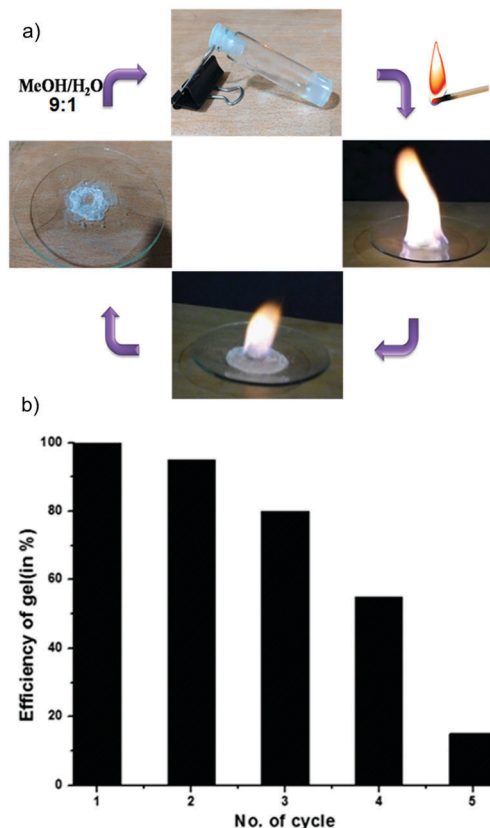


Fig. 24 (a) The burning methanol gel and recharge by 9:1 MeOH:H₂O. (b) Efficiency of the gel after various cycles. Adapted with permission from ref. 124. Copyright©2015 The Royal Society of Chemistry.

4 different metallogels (**MOG**) were obtained (Fig. 25a). The effective gelation ratio of the gelator and metal to ensure complete gelation was found to be in the range of 1:0.6 to 1:1, with CGCs ranging from 0.8 to 1.27 (w/v). These metallogels showed thixotropic behavior as demonstrated by rheological data (Fig. 25b), and gel-to-sol transitions in response to either temperature, pH outside of the range 6–8 or addition of competitive chelating agents. With regard to the morphological features, FE-SEM imaging of the corresponding xerogels revealed 3D aggregated multisheet arrangements for **MOG-1**, **MOG-2** and **MOG-3** and a spongy-like structure for **MOG-4** (Fig. 26a and b).

Although all metallogels showed good self-sustainability, **MOG-1** was found to be the best allowing the construction of free-standing gel blocks in any shape with very good load-bearing capacity (*i.e.* 18 g cm⁻² without compression). Furthermore, the self-healing ability of **MOG-1** was demonstrated as usual by cutting the gel blocks into pieces and subsequently reconnection of the pieces by physical contact. In this manner, the reconstruction of a stable 2.5 cm gel bridge was achieved within 10 min (Fig. 26c–f). It is worth mentioning that besides the self-healing behavior of the gels, the corresponding porous xerogels of **MOG-1** and **MOG-2** were able to absorb N₂ gas at 77 K which showed a type-III adsorption isotherm and exhibited a saturated sorption amount of 32.09 and 46.41 cm³ g⁻¹, comprising a total pore volume of 4.96 × 10⁻² and 7.17 × 10⁻² cm³ g⁻¹, respectively. In addition, these metallogels showed the ability to quickly and selectively remove cationic dyes from mixtures of cationic and anionic dyes

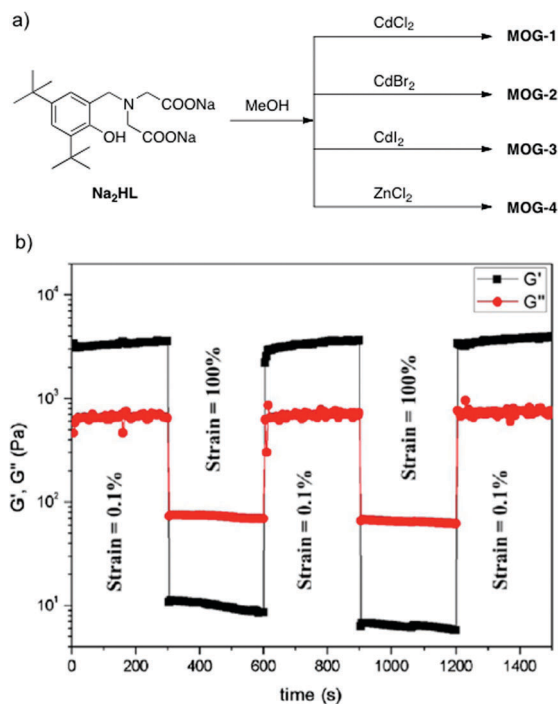


Fig. 25 (a) Synthesis of **MOG**. CGC (w/v) = 0.89 (**MOG-1**); 1.02 (**MOG-2**); 1.27 (**MOG-3**); 0.8 (**MOG-4**). (b) Loop step strain experiment of **MOG-1** at a frequency of 10 rad s⁻¹ showing complete recovery after each cycle. Adapted with permission from ref. 126. Copyright©2016 American Chemical Society.

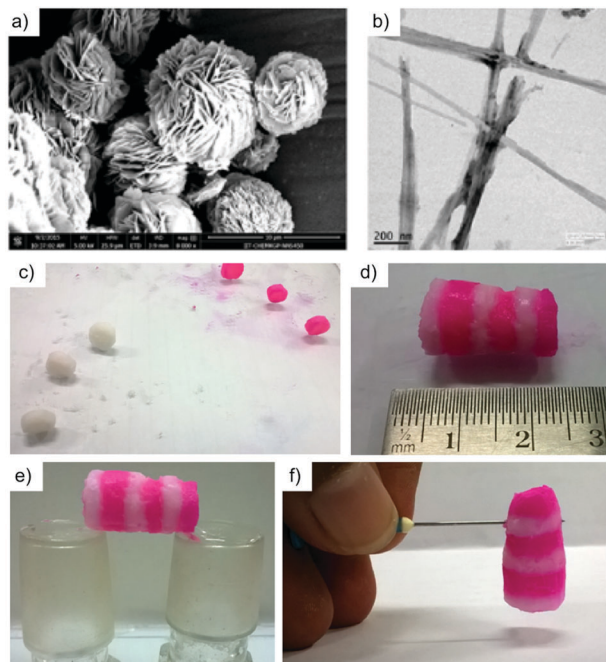


Fig. 26 (a) FESEM of **MOG-1**. (b) TEM of **MOG-1**. (c) Dye doped and undoped gel blocks. (d–f) A bridge constructed by alternate dye doped and undoped 6 **MOG-1** blocks. Diffusion of the dye from doped blocks to the undoped blocks was clearly observed during the formation of the bridge. Adapted with permission from ref. 126. Copyright©2016 American Chemical Society.

(thereby the sodium ions of the ligand are replaced by the cationic dyes).

Conclusions and outlook

In this feature article, we have discussed a series of contributions dealing with the *in situ* fabrication of supramolecular metallogels, using LMW compounds, which show self-healing properties of the bulk gel phase after physical segregation. Most of the advances in this area within the last three years are represented by organogels, whereas examples of hydrogels and organic-aqueous gels are still a minority. The overview of these contributions should stimulate the scientific community interested in gels to look closer at the possibility of achieving gel materials with multi-stimuli responsiveness and self-healing properties by *in situ* metal complexation of LMW ligands. Amino acid-based ligands seem to be promising for the preparation of hydrogels, while heterocycle-containing ligands and other non-amino acid ligands with several coordination centers are more favorable for the formation of organogels. The synthetic versatility of LMW ligands may constitute also a major advantage for tuning the moldability, load-bearing and self-healing properties of *in situ* formed gels. The research carried out with these gels has exposed their potential as safe fuels, conductive materials and materials for environmental remediation. The highly dynamic nature of these gels has also opened an interesting avenue for imparting self-healing properties to other soft gel networks that intrinsically

lack the ability to self-repair upon damage. Moreover, the *in situ* preparation of both monometallic and heterometallic soft gels with self-repair capacity may boost the applications of supramolecular metallogels in areas such as catalysis and drug delivery and overcome some difficulties found in more crystalline structures like metal organic frameworks (MOFs) with lower solubility and processability.^{129,130}

We foresee intense work in this emerging area to gain a deeper understanding of both chemical and physical events that govern the self-healing ability of these materials as well as a coherent design of LMW ligands with suitable structures and solubility properties in order to achieve self-healing metallogels under well-defined conditions that are otherwise difficult to prepare. In this sense, a precise description of the experimental conditions in which the gels are formed is crucial for future advances in the field, where researchers should also use or develop new characterization methods to quantify the contribution from both self-adhesion and self-healing processes to the final properties of the bulk material.

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Notes and references

- 1 P. Terech and R. G. Weiss, *Chem. Rev.*, 1997, **97**, 3133–3260.
- 2 R. G. Weiss and P. Terech, *Molecular Gels: Materials with Self-Assembled Fibrillar Networks*, Springer, New York, 2006.
- 3 M. George and R. G. Weiss, *Acc. Chem. Res.*, 2006, **39**, 489–497.
- 4 J. H. van Esch and B. L. Feringa, *Angew. Chem., Int. Ed.*, 2000, **39**, 2263–2266.
- 5 O. Gronwald and S. Shinkai, *Chem. – Eur. J.*, 2001, **7**, 4328–4334.
- 6 D. J. Abdallah and R. G. Weiss, *Adv. Mater.*, 2000, **12**, 1237–1247.
- 7 B.-K. An, J. Gierschner and S. Y. Park, *Acc. Chem. Res.*, 2012, **44**, 544–554.
- 8 N. Fujita, P. Mukhopadhyay and S. Shinkai, *Annu. Rev. Nano Res.*, 2006, **1**, 385.
- 9 A. R. Hirst, B. Escuder, J. F. Miravet and D. K. Smith, *Angew. Chem., Int. Ed.*, 2008, **47**, 8002–8018.
- 10 D. D. Diaz, D. Kühbeck and R. J. Koopmans, *Chem. Soc. Rev.*, 2011, **40**, 427–448.
- 11 Y. Osada and K. Kajiura, in *Gels handbook*, ed. T. Fushimi, O. Hirasa, Y. Hirokawa, T. Matsunaga, T. Shimomura and L. Wang and H. Ishida, Academic Press, San Diego, 2001, vol. 1–3.
- 12 J. H. Jung, J. H. Lee, J. R. Silverman and G. John, *Chem. Soc. Rev.*, 2013, **42**, 924–936.
- 13 M. D. Segarra-Maset, V. J. Nebot, J. F. Miravet and B. Escuder, *Chem. Soc. Rev.*, 2013, **42**, 7086–7098.
- 14 J. H. van Esch, *Langmuir*, 2009, **25**, 8392–8394.
- 15 T. Tanaka, *Sci. Am.*, 1981, **244**, 110–116.
- 16 D. Derossi, R. Kajiura and Y. Osada, *Polymer Gels: Fundamentals and Biomedical Applications*, Plenum Press, New York, 1991.
- 17 L. A. Estroff and A. D. Hamilton, *Chem. Rev.*, 2004, **104**, 1201–1218.
- 18 N. M. Sangeetha and U. Maitra, *Chem. Soc. Rev.*, 2005, **34**, 821–836.
- 19 A. Ajayaghosh, V. K. Praveen and C. Vijayakumar, *Chem. Soc. Rev.*, 2008, **37**, 109–122.
- 20 P. Xie and R. Zhang, *J. Mater. Chem.*, 2005, **15**, 2529–2550.
- 21 D. K. Smith, *Chem. Commun.*, 2006, 34–44.
- 22 M. George, R. Mathew and R. G. Weiss, *Mol. Gels*, 2006, 449.
- 23 T. Yi, K. Sada, K. Sugiyasu, T. Hatano and S. Shinkai, *Chem. Commun.*, 2003, 344–345.



- 24 J. H. Jung, Y. Ono and S. Shinkai, *Angew. Chem., Int. Ed.*, 2000, **39**, 1862–1865.
- 25 E. Gazit, *Chem. Soc. Rev.*, 2007, **36**, 1263–1269.
- 26 S. M. Aharoni, in *Synthesis, Characterization, and Theory of Polymeric Networks and Gels*, ed. S. M. Aharoni, Plenum, New York, 1992.
- 27 E. R. Zubarev, M. U. Pralle, E. D. Sone and S. I. Stupp, *Adv. Mater.*, 2002, **14**, 198–203.
- 28 X. Huang, P. Terech, S. R. Raghavan and R. G. Weiss, *J. Am. Chem. Soc.*, 2005, **127**, 4336–4344.
- 29 F. Ilmain, T. Tanaka and E. Kokufuta, *Nature*, 1991, **349**, 400–401.
- 30 Y. Osada and A. R. Khokhlov, *Polymer Gels and Networks*, Marcel Dekker, New York, 2002.
- 31 P. Bawa, V. Pillay, Y. E. Choonara and L. C. du Toit, *Biomed. Mater.*, 2009, **4**, 15.
- 32 M. A. C. Stuart, W. T. S. Huck, J. Genzer, M. Müller, C. Ober, M. Stamm, G. B. Sukhorukov, I. Szleifer, V. V. Tsukruk, M. Urban, F. Winnik, S. Zauscher, I. Luzinov and S. Minko, *Nat. Mater.*, 2010, **9**, 101–113.
- 33 M. W. Urban, *Handbook of Stimuli-Responsive Materials*, Wiley-VCH, Weinheim, 2011.
- 34 V. Amendola and M. Meneghetti, *Nanoscale*, 2009, **1**, 74–88.
- 35 P. Fratzl, *J. R. Soc., Interface*, 2007, **4**, 637–642.
- 36 G. Fantner, E. Oroudjev, G. Schitter, L. Golde, P. Thurner, M. Finch, P. Turner, T. Gutsmann, D. Morse, H. Hansma and P. Hansma, *Biophys. J.*, 2006, **90**, 1411–1418.
- 37 J. Thompson, J. Kindt, B. Drake, H. Hansma, D. Morse and P. Hansma, *Nature*, 2001, **414**, 773–776.
- 38 F. Herbst, D. Döhler, P. Michael and W. H. Binder, *Macromol. Rapid Commun.*, 2013, **34**, 203–220.
- 39 M. Gragert, M. Schunack and W. H. Binder, *Macromol. Rapid Commun.*, 2011, **32**, 419–425.
- 40 Y. Chen, A. M. Kushner, G. A. Williams and Z. Guan, *Nat. Chem.*, 2012, **4**, 467–472.
- 41 Q. Wei, J. Wang, X. Shen, X. A. Zhang, J. Zhi Sun, A. Qin and B. Z. Tang, *Sci. Rep.*, 2013, **3**, 1093, DOI: 10.1038/srep01093.
- 42 S. Bode, R. K. Bose, S. Matthes, M. Ehrhardt, A. Seifert, F. H. Schacher, R. M. Paulus, S. Stumpf, B. Sandmann, J. Vitz, A. Winter, S. Hoeppe, S. J. Garcia, S. Spange, S. van der Zwaag, M. D. Hager and U. S. Schubert, *Polym. Chem.*, 2013, **4**, 4966–4973.
- 43 Y. Jin, C. Yu, R. Denman and W. Zhang, *Chem. Soc. Rev.*, 2013, **42**, 6634–6654.
- 44 N. Roy, B. Bruchmann and J. Lehn, *Chem. Soc. Rev.*, 2015, **44**, 3786–3807.
- 45 Z. Wei, J. Yang, J. Zhou, F. Xu, M. Zrínyi, P. Dussault, Y. Osada and Y. Chen, *Chem. Soc. Rev.*, 2014, **43**, 8114–8131.
- 46 X. Yu, L. Chen, M. Zhang and T. Yi, *Chem. Soc. Rev.*, 2014, **43**, 5346–5371.
- 47 J. Cui and A. del Campo, *Chem. Commun.*, 2012, **48**, 9302–9304.
- 48 P. Froimowicz, D. Klinger and K. Landfester, *Chem. – Eur. J.*, 2011, **17**, 12465–12475.
- 49 Y. Zhang, L. Tao, S. Li and Y. Wei, *Biomacromolecules*, 2011, **12**, 2894–2901.
- 50 M. Nakahata, Y. Takashima, H. Yamaguchi and A. Harada, *Nat. Commun.*, 2011, **2**, 511, DOI: 10.1038/ncomms1521.
- 51 S. Roy, A. Baral and A. Banerjee, *Chem. – Eur. J.*, 2013, **19**, 14950–14957.
- 52 U. K. Das, S. Banerjee and P. Dastidar, *Chem. – Asian J.*, 2014, **9**, 2475–2482.
- 53 A. Vidyasagar, K. Handore and K. M. Sureshan, *Angew. Chem., Int. Ed.*, 2011, **50**, 8021–8024.
- 54 M. Zhang, D. Xu, X. Yan, J. Chen, S. Dong, B. Zheng and F. Huang, *Angew. Chem.*, 2012, **124**, 7117–7121.
- 55 D. Liu, D. Wang, M. Wang, Y. Zheng, K. Koynov, G. K. Auernhammer, H.-J. Butt and T. I. Ikeda, *Macromolecules*, 2013, **46**, 4617–4625.
- 56 G. Akay, A. Hassan-Raeisi, D. C. Tuncaboylu, N. Orakdogan, S. Abdurrahmanoglu, W. Oppermann and O. Okay, *Soft Matter*, 2013, **9**, 2254–2261.
- 57 G. Deng, F. Li, H. Yu, F. Liu, C. Liu, W. Sun, H. Jiang and Y. Chen, *ACS Macro Lett.*, 2012, **1**, 275–279.
- 58 J. M. Garcia, G. O. Jones, K. Virwani, B. D. McCloskey, D. J. Boday, G. M. ter Huurne, H. W. Horn, D. J. Coady, A. M. Bintaleb, A. M. S. Alabdulrahman, F. Alsewailam, H. A. A. Almegren and J. L. Hedrick, *Science*, 2014, **344**, 732–735.
- 59 A. B. South and L. A. Lyon, *Angew. Chem.*, 2010, **122**, 779–783.
- 60 T. Kakuta, Y. Takashima, M. Nakahata, M. Otsubo, H. Yamaguchi and A. Harada, *Adv. Mater.*, 2013, **25**, 2849–2853.
- 61 M. H. Filby and J. W. Steed, *Coord. Chem. Rev.*, 2006, **250**, 3200–3218.
- 62 J. Zhang and C. Y. Su, *Coord. Chem. Rev.*, 2013, **257**, 1373–1408.
- 63 F. Fages, *Angew. Chem., Int. Ed.*, 2006, **45**, 1680–1682.
- 64 M.-O. M. Piepenbrock, G. O. Lloyd, N. Clarke and J. W. Steed, *Chem. Rev.*, 2010, **110**, 1960–2004.
- 65 K. Murata, M. Aoki, T. Nishi, A. Ikeda and S. Shinkai, *Chem. Commun.*, 1991, 1715–1718.
- 66 A. Y.-Y. Tam and V. W.-W. Yam, *Chem. Soc. Rev.*, 2013, **42**, 1540–1567.
- 67 K. Murata, M. Aoki, T. Suzuki, T. Harada, H. Kawabata, T. Komori, F. Ohseto, K. Ueda and S. Shinkai, *J. Am. Chem. Soc.*, 1994, **116**, 6664–6676.
- 68 Q. Lin, T.-T. Lu, X. Zhu, T.-B. Wei, H. Li and Y.-M. Zhang, *Chem. Sci.*, 2016, **7**, 5341–5346.
- 69 Q. Lin, T.-T. Lu, X. Zhu, B. Sun, Q.-P. Yang, T.-B. Wei and Y.-M. Zhang, *Chem. Commun.*, 2015, **51**, 1635–1638.
- 70 Q. Lin, B. Sun, Q.-P. Yang, Y.-P. Fu, X. Zhu, T.-B. Wei and Y.-M. Zhang, *Chem. – Eur. J.*, 2014, **20**, 11457–11462.
- 71 A. Kishimura, T. Yamashita and T. Aida, *J. Am. Chem. Soc.*, 2005, **127**, 179–183.
- 72 T. Aida, E. W. Meijer and S. I. Stupp, *Science*, 2012, **335**, 813–817.
- 73 J. H. Lee, S. Kang, J. Y. Lee, J. Jaworski and J. H. Jung, *Chem. – Eur. J.*, 2013, **19**, 16665–16671.
- 74 R. Afrasiabi and H. B. Kraatz, *Chem. – Eur. J.*, 2013, **19**, 17296–17300.
- 75 A. Y.-Y. Tam, K. M.-C. Wong and V. W.-W. Yam, *J. Am. Chem. Soc.*, 2009, **131**, 6253–6260.
- 76 S. Bhowmik, B. N. Ghosh, V. Marjomki and K. Rissanen, *J. Am. Chem. Soc.*, 2014, **136**, 5543–5546.
- 77 W. Bai and D. A. Spivak, *Angew. Chem., Int. Ed.*, 2014, **53**, 2095–2098.
- 78 H. B. Aiyappa, S. Saha, P. Wadge, R. Banerjee and S. Kurungot, *Chem. Sci.*, 2015, **6**, 603–607.
- 79 S. Saha, G. Das, J. Thote and R. Banerjee, *J. Am. Chem. Soc.*, 2014, **136**, 14845–14851.
- 80 R. Banerjee, *Cryst. Growth Des.*, 2014, **14**, 3434–3437.
- 81 S. Sarkar, S. Dutta, S. Chakrabarti, P. Bairi and T. Pal, *ACS Appl. Mater. Interfaces*, 2014, **6**, 6308–6316.
- 82 A. E. Rowan, M. C. Feiters and R. J. M. Nolte, *Chem. Commun.*, 1997, 545–546.
- 83 Q. Wei and S. L. James, *Chem. Commun.*, 2005, 1555–1556.
- 84 M. Shirakawa, S. Kawano, N. Fujita, K. Sada and S. Shinkai, *J. Org. Chem.*, 2003, **68**, 5037–5044.
- 85 M. George, G. P. Funkhouser, P. Terech and R. G. Weiss, *Langmuir*, 2006, **22**, 7885–7893.
- 86 B. Xing, M.-F. Choi and B. Xu, *Chem. – Eur. J.*, 2002, **8**, 5028–5032.
- 87 P. Terech, G. Gebel and R. Ramasseul, *Langmuir*, 1996, **12**, 4321–4323.
- 88 N. Fujita, Y. Sakamoto, M. Shirakawa, M. Ojima, A. Fujii, M. Ozaki and S. Shinkai, *J. Am. Chem. Soc.*, 2007, **129**, 4134–4135.
- 89 P. Sutar and T. K. Maji, *Chem. Commun.*, 2016, **52**, 8055–8074.
- 90 A. Phadke, C. Zhang, B. Arman, C.-C. Hsu, R. A. Mashelkar, A. K. Lele, M. J. Tauber, G. Arya and S. Varghese, *Proc. Natl. Acad. Sci. U. S. A.*, 2012, **109**, 4383–4388.
- 91 Y. Shi, M. Wang, C. Ma, Y. Wang, X. Li and G. Yu, *Nano Lett.*, 2015, **15**, 6276–6281.
- 92 Z. Shafiq, J. Cui, L. Pastor-Pérez, V. San Miguel, R. A. Gropeanu, C. Serrano and A. del Campo, *Angew. Chem., Int. Ed.*, 2012, **51**, 4332–4335.
- 93 M. Krogsgaard, M. A. Behrens, J. S. Pedersen and H. Birkedal, *Biomacromolecules*, 2013, **14**, 297–301.
- 94 M. Guvendiren, P. B. Messersmith and K. R. Shull, *Biomacromolecules*, 2008, **9**, 122–128.
- 95 N. Holten-Andersen, M. J. Harrington, H. Birkedal, B. P. Lee, P. B. Messersmith, K. Y. C. Lee and J. H. Waite, *Proc. Natl. Acad. Sci. U. S. A.*, 2011, **108**, 2651–2655.
- 96 T.-A. Asoh, H. Yoshitake, Y. Takano and A. Kikuchi, *Macromol. Chem. Phys.*, 2013, **214**, 2534–2539.
- 97 F. R. Kersey, D. M. Loveless and S. L. Craig, *J. R. Soc., Interface*, 2007, **4**, 373–380.
- 98 J. Yuan, X. Fang, L. Zhang, G. Hong, Y. Lin, Q. Zheng, Y. Xu, Y. Ruan, W. Weng, H. Xia and G. Chen, *J. Mater. Chem.*, 2012, **22**, 11515–11522.
- 99 B. Yang, H. Zhang, H. Peng, Y. Xu, B. Wu, W. Weng and L. Lia, *Polym. Chem.*, 2014, **5**, 1945–1953.
- 100 Y. Zhang, B. Yang, X. Zhang, L. Xu, L. Tao, S. Li and Y. Wei, *Chem. Commun.*, 2012, **48**, 9305–9307.



- 101 J. Yuan, X. Fang, L. Zhang, G. Hong, Y. Lin, Q. Zheng, Y. Xu, Y. Ruan, W. Weng, H. Xia and G. Chen, *J. Mater. Chem.*, 2012, **22**, 11515–11522.
- 102 M. Zhang, D. Xu, X. Yan, J. Chen, S. Dong, B. Zheng and F. Huang, *Angew. Chem., Int. Ed.*, 2012, **124**, 7117–7121.
- 103 S. Varguese, A. Lele and R. Mahelkar, *J. Polym. Sci., Part A: Polym. Chem.*, 2006, **44**, 666–670.
- 104 T. Sato, M. Ebara, S. Tanaka, T.-A. Asoh, A. Kikuchi and T. Aoyagi, *Phys. Chem. Chem. Phys.*, 2013, **15**, 10628–10635.
- 105 W. Fang, X. Liu, Z. Lu and T. Tu, *Chem. Commun.*, 2014, **50**, 3313–3316.
- 106 W. Fang, Z. Sun and T. Tu, *J. Phys. Chem. C*, 2013, **117**, 25185–25194.
- 107 A. Dey, S. K. Mandal and K. Biradha, *CrystEngComm*, 2013, **15**, 9769–9778.
- 108 T. D. Hamilton, D.-K. Bučar, J. Baltrusaitis, D. R. Flanagan, Y. Li, S. Ghorai, A. V. Tivanski and L. R. MacGillivray, *J. Am. Chem. Soc.*, 2011, **133**, 3365–3371.
- 109 S. Basak, J. Nanda and A. Banerjee, *Chem. Commun.*, 2014, **50**, 2356–2359.
- 110 Z. Sun, Q. Huang, T. He, Z. Li, Y. Zhang and L. Yi, *ChemPhysChem*, 2014, **15**, 2421–2430.
- 111 S. Saha, J. Bachl, T. Kundu, D. D. Díaz and R. Banerjee, *Chem. Commun.*, 2014, **50**, 3004–3006.
- 112 J.-S. Shen, G.-J. Mao, Y.-H. Zhou, Y.-B. Jiang and H.-W. Zhang, *Dalton Trans.*, 2010, **39**, 7054–7058.
- 113 S. Samai and K. Biradha, *Chem. Mater.*, 2012, **24**, 1165–1173.
- 114 U. S. Schubert and C. Eschbaumer, *Angew. Chem., Int. Ed.*, 2002, **41**, 2892–2926.
- 115 P. Terech, M. Yan, M. Maréchal, G. Royal, J. Galvez and S. K. P. Velu, *Phys. Chem. Chem. Phys.*, 2013, **15**, 7338–7344.
- 116 X. Yan, D. Xu, X. Chi, J. Chen, S. Dong, X. Ding, Y. Yu and F. Huang, *Adv. Mater.*, 2012, **24**, 362–369.
- 117 H. Ibach, *Surf. Sci. Rep.*, 1997, **29**, 195–263.
- 118 L. Yan, S. Gou, Z. Ye, S. Zhang and L. Ma, *Chem. Commun.*, 2014, **50**, 12847–12850.
- 119 P. Sahoo, R. Sankolli, H.-Y. Lee, S. R. Raghavan and P. Dastidar, *Chem. – Eur. J.*, 2012, **18**, 8057–8063.
- 120 L. Yan, L. Shen, M. Lv, W. Yu, J. Chen, S. Wang, X. Fu and Z. Ye, *Chem. Commun.*, 2015, **51**, 17627–17629.
- 121 M. Martínez-Calvo, O. Kotova, M. E. Möbius, A. P. Bell, T. McCabe, J. J. Boland and T. Gunnlaugsson, *J. Am. Chem. Soc.*, 2015, **137**, 1983–1992.
- 122 T. Feldner, M. Haring, S. Saha, J. Esquena, R. Banerjee and D. D. Díaz, *Chem. Mater.*, 2016, **28**, 3210–3217.
- 123 S. Saha, E.-M. Schön, C. Cativiela, D. D. Díaz and R. Banerjee, *Chem. – Eur. J.*, 2013, **19**, 9562–9568.
- 124 S. Bera and D. Haldar, *J. Mater. Chem. A*, 2016, **4**, 6933–6939.
- 125 S. Song, A. Song, L. Feng, G. Wei, S. Dong and J. Hao, *ACS Appl. Mater. Interfaces*, 2014, **6**, 18319–18328.
- 126 C. K. Karan and M. Bhattacharjee, *ACS Appl. Mater. Interfaces*, 2016, **8**, 5526–5535.
- 127 S. Khatua, H. Stoeckli-Evans, T. Harada, R. Kuroda and M. Bhattacharjee, *Inorg. Chem.*, 2006, **45**, 9619–9621.
- 128 S. Khatua, T. Harada, R. Kuroda and M. Bhattacharjee, *Chem. Commun.*, 2007, 3927–3929.
- 129 Themed Issue: Metal–organic Frameworks, *Chem. Rev.*, 2012, **112**, 673–674.
- 130 Themed Issue: Metal–organic Frameworks, *Chem. Soc. Rev.*, 2014, **43**, 5415–5418.

