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**We report for the first time the formation of a metal amide–hydride solid solution. The dissolution of KH into KNH<sub>2</sub> leads to an anionic substitution, which decreases the interaction among NH<sub>2</sub><sup>−</sup> ions. The rotational properties of the high temperature polymorphs of KNH<sub>2</sub> are thereby retained down to room temperature.**

The amides of alkaline and alkaline-earth metals were discovered and independently investigated by J. L. Gay-Lussac and H. Davy in the early 19th century.<sup>1,2</sup> Further studies on their properties were performed towards the end of the same century, most systematically by A. W. Titherley.<sup>3</sup> The determination of their crystal structures was possible only after the 1930s with the studies of Juza *et al.*,<sup>4–10</sup> later continued by Jacobs *et al.*<sup>11,12</sup>

At that time this class of compounds was mostly used for organic synthesis. However, more recently, metal amide–metal hydride mixtures have been proven to be suitable for reversible hydrogen storage.<sup>13</sup> Furthermore, metal amide–metal borohydride systems are regarded as potential solid state ionic conductors.<sup>14</sup>

The structural and thermal properties of light-weight amides prepared *via* the reaction of metal hydrides with ammonia have been systematically examined by *in situ* diffraction experiments.<sup>15</sup> While the reaction mechanism and products reported in previous studies were confirmed for the amides of lithium and sodium,<sup>16–18</sup> the formation of new K–N–H based intermediates was suggested for potassium amide. The same intermediates

## KNH<sub>2</sub>–KH: a metal amide–hydride solid solution†

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were proven to play a role in the desorption reactions of the K–Mg–N–H system.<sup>19</sup> These intermediates are isolated here in the KNH<sub>2</sub>–KH system and, to the best of our knowledge, identified as the first metal amide–hydride solid solution.

The crystal structures of pristine potassium amide and potassium hydride have already been investigated *via* X-ray and neutron diffraction.<sup>7,8,12,20–22</sup> Potassium hydride is known to crystallize in a cubic rock-salt type structure with space group (s.g.) *Fm* $\bar{3}$ *m* and no polymorphic changes are expected in the temperature range from room temperature (RT) to 390 °C.<sup>21</sup> For potassium amide, the stable polymorph at RT is monoclinic with the s.g. *P2*<sub>1</sub>*m*.<sup>7</sup> Upon heating this phase transforms into a tetragonal structure in *P4/nmm* at 54 °C.<sup>12</sup> The latter phase is stable only in a narrow temperature range; and already above 75 °C the stability of a cubic phase (*Fm* $\bar{3}$ *m*) prevails.<sup>12,20,22</sup> An increase of symmetry is therefore reached with the two phase transitions at a higher temperature. The explanation of the symmetry changes resides in the increasingly high orientational disorder of the amide anions, as proven by powder neutron diffraction (PND), quasielastic incoherent neutron scattering and orientation-dependent deuterium spin lattice relaxation.<sup>20,23,24</sup> If the maximum temperature upon heating is kept below the decomposition temperature of KNH<sub>2</sub> (*ca.* 340 °C), the two phase transitions and the associated rotational dynamics should be reversible upon cooling.

This is in agreement with the present *in situ* synchrotron radiation powder X-ray diffraction experiment (SR-PXD) on KNH<sub>2</sub> (ESI†). When a 0.5KNH<sub>2</sub> + 0.5KH mixture was investigated, the expected phase transformations of KNH<sub>2</sub> took place, but the interaction of KNH<sub>2</sub> and KH in the temperature range between 100 °C and 270 °C led to the formation of a new cubic structure (Fig. 1a).‡

PND at 270 °C of a potassium deuteramide–potassium deuteride mixture confirmed the presence of a single cubic phase at this temperature (Fig. 1b).§

The structure was resolved assuming an ionic crystal with unaltered positions for the potassium cations and partially occupied sites at the original positions of amide and hydride anions.

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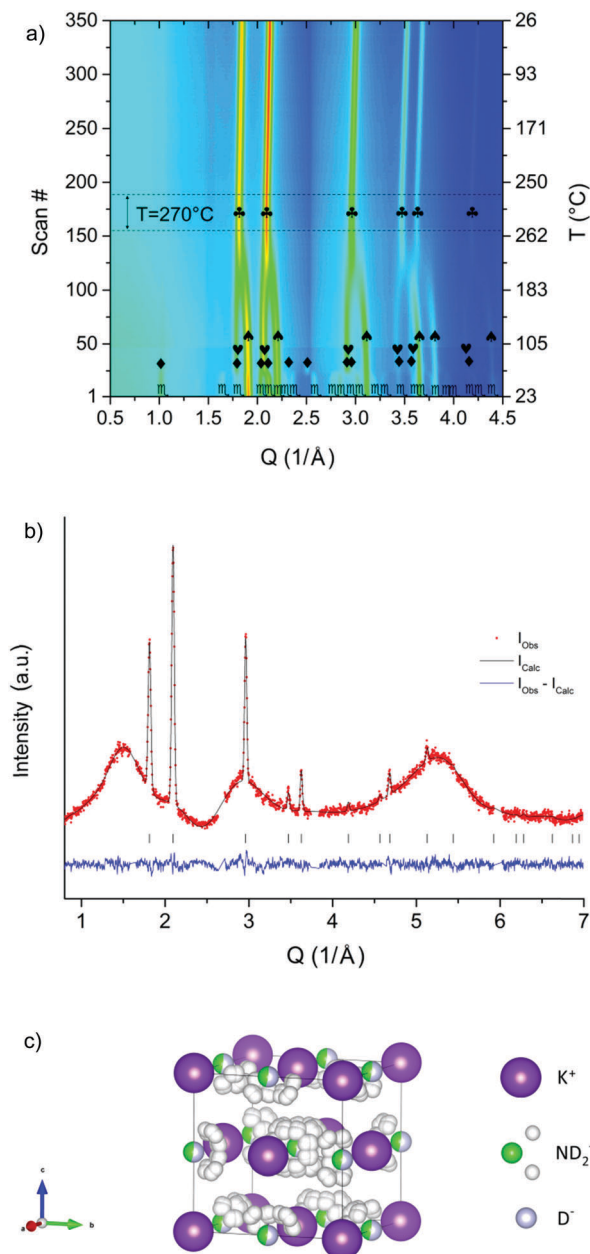
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† Electronic supplementary information (ESI) available: Powder X-ray and neutron diffractograms, Rietveld refinements, thermal decomposition, and experimental details. See DOI: 10.1039/c6cc05777b





**Fig. 1** (a) *In situ* SR-PXD experiment on the 0.5KNH<sub>2</sub> + 0.5KH sample.  $\square$  = KNH<sub>2</sub> (*P2<sub>1</sub>/m*),  $\blacklozenge$  = KNH<sub>2</sub> (*P4/nmm*),  $\blacktriangledown$  = KNH<sub>2</sub> (*Fm-3m*),  $\blacklozenge$  = KH (*Fm-3m*),  $\clubsuit$  = new phase (*Fm-3m*). (b) PND pattern of the nominal 0.5KNH<sub>2</sub> + 0.5KH mixture annealed up to 270 °C and kept under isothermal conditions during the data collection.  $R_{wp}$  (%) = 2.8 (corrected for background). The wavy background was originated from the quartz sample holder (see the ESI†). (c) Structural model of the cubic phase (s.g. *Fm-3m*) of composition K(ND<sub>2</sub>)<sub>0.46</sub>D<sub>0.54</sub> obtained after the Rietveld refinement, taking into account the anionic substitution at the position (0.5 0.5 0.5) and the orientational disorder of amide anions.

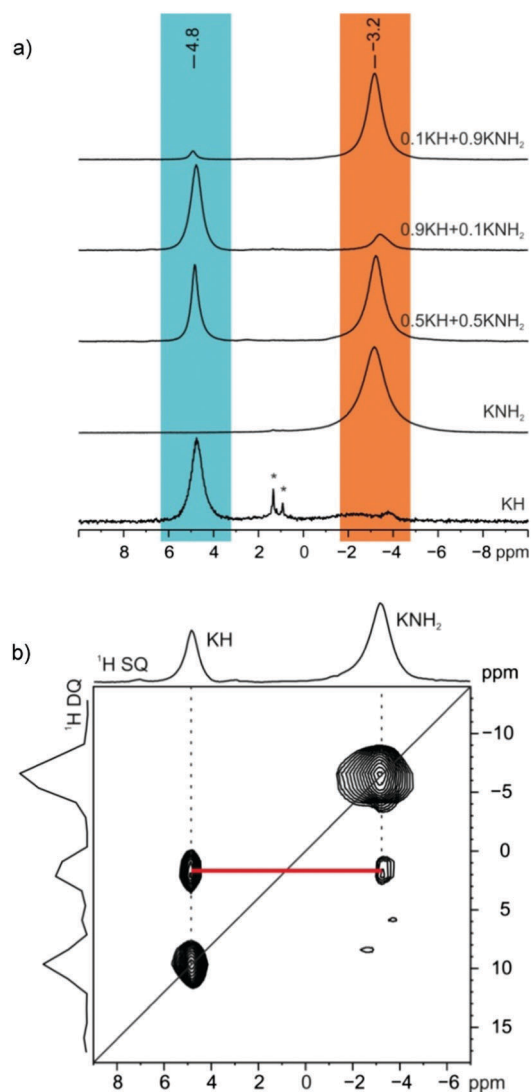
Indeed the formation of a solid solution is highly favorable due to the structural similarities between the two cubic polymorphs of KNH<sub>2</sub> and KH (same space group, same cation, same charge for the anions, similar lattice constants).

The local symmetry of the deuteramide groups is compatible with the s.g. *Fm-3m* only assuming rotational dynamics and

orientational disorder. In this case the restrictions imposed from both the s.g. and the rigid amide groups resulted in partially occupied sites (multiplicity = 192) for the deuterium atoms of each amide group (Fig. 1c).

The final Rietveld refinement<sup>25,26</sup> of the PND pattern (Fig. 1b) confirmed the structural model, *i.e.* the formation of a potassium amide–hydride solid solution. To the best of our knowledge, no similar cases have been reported so far for other alkaline metal amide–metal hydride mixtures.

The similarity of the chemical environment between the starting materials and the  $x\text{KNH}_2 + (1 - x)\text{KH}$  samples at different compositions ( $x = 0.1, 0.5, 0.9$ ) after annealing was verified by <sup>1</sup>H magic angle spinning solid-state NMR (MAS SSNMR), see Fig. 2.†



**Fig. 2** (a) <sup>1</sup>H (400.23 MHz) MAS SSNMR spectra of the starting reagents and  $x\text{KNH}_2 + (1 - x)\text{KH}$  samples at different compositions ( $x = 0.1, 0.5, 0.9$ ) after annealing recorded with a spinning speed of 32 kHz. Asterisks denote impurities. (b) 2D <sup>1</sup>H (400.23 MHz) DQ MAS SSNMR spectrum of the KNH<sub>2</sub> + KH ( $x = 0.5$ ) after annealing recorded with a spinning speed of 32 kHz. The red line highlights the DQ correlation between the KH and KNH<sub>2</sub> signals.



No significant shifts are observed while the integral values perfectly reflect the composition of the solid solution. Furthermore, the same  $T_1$   $^1\text{H}$  value (46 s) for both signals supports the formation of a solid solution since it indicates that spin diffusion is active. This is only possible if they belong to the same phase or in the case of homogeneous samples on a nanometer scale.<sup>27</sup> Direct evidence of the solid solution formation is provided by the  $^1\text{H}$  double-quantum (DQ) MAS SSNMR experiment (Fig. 2b). Indeed, the observed DQ correlation between the  $\text{KNH}_2$  (−3.2 ppm) and KH (4.8 ppm) signals implies that they are in close spatial proximity to each other (less than 5 Å). This is only possible if they are intimately related as in a solid solution.<sup>28,29</sup> A similar correlation, although much weaker, is observed for the sample before annealing (ESI†), which can be explained with the formation of a small fraction of solid solution due to the fast rotation and slightly increased temperature during the NMR experiment.

The effect of the starting composition of the mixture on the final structure was studied by means of *in situ* SR-PXD. A linear relationship was found between the unit cell parameter of the cubic phase at  $T = 270$  °C and the molar fraction of amide anions (Fig. 3 and Table 1) as expected from Vegard's law.<sup>30</sup> Therefore an almost ideal behavior is expected for the solid solution ( $\Delta V_{\text{mix}} \sim 0$  and  $\Delta H_{\text{mix}} \sim 0$ ).

It is noteworthy that, for the compositions  $x = 0.1, 0.3, 0.5, 0.7$ , the structure did not change during the cooling process down to RT, except for the thermal contraction of the unit cell volume (more details in Table 1 and in the ESI†).

It appears that the addition of potassium hydride can stabilize the cubic geometry, which retains the rotation of the amide anions even at RT. A similar behavior was previously reported for the sodium borohydride–sodium chloride systems, at different temperatures ( $T = -81$  °C).<sup>31</sup> *Ex situ* PXD of annealed samples ( $x = 0.3, 0.5, 0.7, 0.9$ ) collected with a Bragg Brentano diffractometer proves the coexistence of at least two different

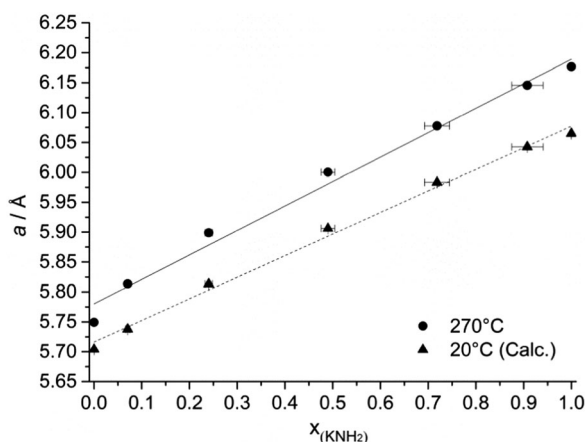


Fig. 3 Unit cell parameters (dots) obtained by *in situ* SR-PXD experiments under isothermal conditions ( $T = 270$  °C) as a function of the potassium amide content and the corresponding linear fit (continuous line). The values for  $T = 20$  °C (triangles) were calculated using the thermal expansion coefficient of each phase and then fitted (dotted line). The error bars for molar fractions and cell parameters were calculated considering the errors of the Rietveld process.

Table 1 Structural details of the  $\text{K}(\text{NH}_2)_x\text{H}_{(1-x)}$  phase at different compositions

$x^a$	$x^b$	$a/\text{Å}$	$\alpha/\text{K}^{-1}$	$a/\text{Å}$
0	0	5.74934(14)	$3.19(2) \times 10^{-5}$	5.70390(14)
0.1	0.070(5)	5.81372(13)	$5.30(1) \times 10^{-5}$	5.73769(13)
0.3	0.240(8)	5.8989(2)	$5.88(2) \times 10^{-5}$	5.8134(2)
0.5	0.490(14)	6.0003 (2)	$6.32(1) \times 10^{-5}$	5.9060(2)
0.7	0.72(3)	6.07768(12)	$6.83(1) \times 10^{-5}$	5.98320(12)
0.9	0.91(3)	6.14553(14)	$6.83(2) \times 10^{-5}$	6.04230(14)
1	1	6.17667(8)	$7.38(1) \times 10^{-5}$	6.06473 (8)

<sup>a</sup> Nominal composition. <sup>b</sup> Refined composition (molar fraction) of  $\text{KNH}_2$ . <sup>c</sup> Refined cell parameters at 270 °C. <sup>d</sup> Calculated linear thermal expansion coefficients. <sup>e</sup> Calculated cell parameters at 20 °C. Estimated standard deviations are given in parentheses.

cubic structures. The composition  $x = 0.1$ , however, presented unchanged cubic phase (see the ESI†). At the composition  $x = 0.9$ , coexistence of the cubic and monoclinic phases was found. These results suggest the presence of a two phase field at RT. Nevertheless, even in the cases where phase segregation occurred, both cubic structures were proven to be unchanged even after several months, hence a complete transformation back to the pure monoclinic phase of  $\text{KNH}_2$  and cubic phase of KH did not occur.

It is noteworthy that the same structures are formed simply by mechanochemical treatments of the starting reactants. In some cases ( $x = 0.1, 0.3$  and  $0.9$ ) manual grinding is enough to promote the formation of detectable amounts of  $\text{K}(\text{NH}_2)_x\text{H}_{(1-x)}$  solid solution. These species are therefore easily formed and are expected to be possibly identified as reaction products or intermediates in future studies of amide-based systems containing potassium.

To the best of our knowledge, amide/hydride solid solutions were not reported so far. In this sense, the chemistry of the K-based amide/hydride system is interestingly peculiar and differs substantially from that of the other alkaline metal amides/hydrides.

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## Notes and references

‡ The *in situ* SR-PXD experiments were performed at the diffraction beamline I711, MAX Lab (Lund, Sweden) and at the diffraction beamline P02, DESY (Hamburg, Germany) (monochromatic beams of  $\lambda \approx 0.99$  Å and  $\approx 0.2$  Å were employed respectively). The *in situ* cells and the procedure used are described elsewhere.<sup>19,32</sup>

§ PND was performed at the PUS instrument at the JEEP II reactor at IFE, Norway.<sup>33</sup> Neutrons with  $\lambda = 1.5539$  Å were provided by a focussing Ge(511) monochromator at 90° take-off angle. The data was collected in the range  $2\theta = 10$ – $130^\circ$  ( $\Delta 2\theta = 0.05^\circ$ ) by 2 detector banks each with 7



vertically stacked position sensitive detectors. The sample was contained in an argon-filled quartz tube with 6 mm diameter placed inside an in-house built furnace. The structure solution was carried out using the software "FOX".<sup>34</sup> The Rietveld refinement was performed by means of the software GSAS<sup>35</sup> and the EXPGUI graphic interface.<sup>36</sup>

¶ Solid-state NMR experiments were run on a Bruker AVANCE II 400 instrument operating at 400.23 MHz for <sup>1</sup>H and equipped with a 2.5 mm probe. The <sup>1</sup>H MAS spectra were recorded at the spinning speed of 32 kHz with the DEPTH sequence ( $\pi/2 - \pi - \pi$ ) for the suppression of the probe background signal (<sup>1</sup>H 90° = 2.5  $\mu$ s; scans = 16; relaxation delay = 53 s). The 2D <sup>1</sup>H DQ MAS experiments were performed at 32 kHz with the back-to-back (BABA) recoupling pulse sequence with excitation time durations of one rotor period (<sup>1</sup>H 90° = 2.5  $\mu$ s; 32 scans;  $t_1$  increments = 46; relaxation delay = 53 s). <sup>1</sup>H scale was calibrated with adamantane (<sup>1</sup>H signal at 1.87 ppm) as external standards.

- 1 J. L. Gay and L. J. Thenard, *Ann. Phys.*, 1809, **32**, 23–39.
- 2 H. Davy, *Philos. Trans. R. Soc. London*, 1808, **98**, 333–370.
- 3 A. W. Titherley, *J. Chem. Soc., Trans.*, 1894, **65**, 504–522.
- 4 R. Juza, *Angew. Chem.*, 1964, **76**, 290–300.
- 5 R. Juza, *Z. Anorg. Allg. Chem.*, 1937, **231**, 121–135.
- 6 R. Juza, K. Fasold and C. Haeberle, *Z. Anorg. Allg. Chem.*, 1937, **234**, 75–85.
- 7 R. Juza, H. Jacobs and W. Klose, *Z. Anorg. Allg. Chem.*, 1965, **338**, 171–178.
- 8 R. Juza and H. Liedtke, *Z. Anorg. Allg. Chem.*, 1957, **290**, 205–208.
- 9 R. Juza and K. Opp, *Z. Anorg. Allg. Chem.*, 1951, **266**, 313–324.
- 10 R. Juza, H. H. Weber and K. Opp, *Z. Anorg. Allg. Chem.*, 1956, **284**, 73–82.
- 11 H. Jacobs, *Z. Anorg. Allg. Chem.*, 1971, **382**, 97–109.
- 12 H. Jacobs and E. Von Osten, *Z. Naturforsch.*, 1976, **31**, 385–386.
- 13 P. Chen, Z. Xiong, J. Luo, J. Lin and K. L. Tan, *Nature*, 2002, **420**, 302–304.
- 14 M. Matsuo, A. Remhof, P. Martelli, R. Caputo, M. Ernst, Y. Miura, T. Sato, H. Oguchi, H. Maekawa, H. Takamura, A. Borgschulte, A. Züttel and S.-i. Orimo, *J. Am. Chem. Soc.*, 2009, **131**, 16389–16391.
- 15 C. Pistidda, A. Santoru, S. Garroni, N. Bergemann, A. Rzeszutek, C. Horstmann, D. Thomas, T. Klassen and M. Dornheim, *J. Phys. Chem. C*, 2015, **119**, 934–943.
- 16 W. I. F. David, M. O. Jones, D. H. Gregory, C. M. Jewell, S. R. Johnson, A. Walton and P. P. Edwards, *J. Am. Chem. Soc.*, 2007, **129**, 1594–1601.
- 17 J. W. Makepeace, M. O. Jones, S. K. Callear, P. P. Edwards and W. I. F. David, *Phys. Chem. Chem. Phys.*, 2014, **16**, 4061–4070.
- 18 H. Yamamoto, H. Miyaoka, S. Hino, H. Nakanishi, T. Ichikawa and Y. Kojima, *Int. J. Hydrogen Energy*, 2009, **34**, 9760–9764.
- 19 A. Santoru, S. Garroni, C. Pistidda, C. Milanese, A. Girella, A. Marini, E. Masolo, A. Valentoni, N. Bergemann, T. T. Le, H. Cao, D. Haase, O. Balmes, K. Taube, G. Mulas, S. Enzo, T. Klassen and M. Dornheim, *Phys. Chem. Chem. Phys.*, 2016, **18**, 3910–3920.
- 20 M. Müller, J. Senker, B. Asmussen, W. Press, H. Jacobs, W. Kockelmann, H. M. Mayer and R. M. Ibberson, *J. Chem. Phys.*, 1997, **107**, 2363–2373.
- 21 V. G. Kuznetsov and M. M. Shkrabkina, *J. Struct. Chem.*, 1962, **3**, 532–537.
- 22 M. Müller, B. Asmussen, W. Press, J. Senker, H. Jacobs, H. Büttner, W. Kockelmann and R. M. Ibberson, *Phys. B*, 1997, **234–236**, 45–47.
- 23 M. Müller, B. Asmussen, W. Press, J. Senker, H. Jacobs, H. Büttner and H. Schober, *J. Chem. Phys.*, 1998, **109**, 3559–3567.
- 24 J. Senker, *Solid State Nucl. Magn. Reson.*, 2004, **26**, 22–35.
- 25 H. M. Rietveld, *Acta Crystallogr.*, 1967, **22**, 151–152.
- 26 H. M. Rietveld, *J. Appl. Crystallogr.*, 1969, **2**, 65–71.
- 27 K. Gaglioti, M. R. Chierotti, F. Grifasi, R. Gobetto, U. J. Griesser, D. Hasa and D. Voinovich, *CrystEngComm*, 2014, **16**, 8252–8262.
- 28 D. Braga, L. Chelazzi, F. Grepioni, E. Dichiarante, M. R. Chierotti and R. Gobetto, *Cryst. Growth Des.*, 2013, **13**, 2564–2572.
- 29 M. R. Chierotti and R. Gobetto, *CrystEngComm*, 2013, **15**, 8599–8612.
- 30 L. Vegard, *Z. Phys.*, 1921, **5**, 17–26.
- 31 J. E. Olsen, P. Karen, M. H. Sørby and B. C. Hauback, *J. Alloys Compd.*, 2014, **587**, 374–379.
- 32 U. Bösenberg, C. Pistidda, M. Tolkiehn, N. Busch, I. Saldan, K. Suarez-Alcantara, A. Arendarska, T. Klassen and M. Dornheim, *Int. J. Hydrogen Energy*, 2014, **39**, 9899–9903.
- 33 B. C. Hauback, H. Fjellvåg, O. Steinsvoll, K. Johansson, O. T. Buset and J. Jørgensen, *J. Neutron Res.*, 2000, **8**, 215–232.
- 34 V. Favre-Nicolin and R. Cerny, *J. Appl. Crystallogr.*, 2002, **35**, 734–743.
- 35 A. C. L. A. R. B. V. Dreele, Los Alamos National Laboratory Report LAUR 86–748, 2000.
- 36 B. Toby, *J. Appl. Crystallogr.*, 2001, **34**, 210–213.

