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Methane trapped in the two distinct dodecahedral cages of the ionic clathrate hydrate of TBAB was studied by single crystal XRD and MD simulation. The relative CH₄ occupancies over the cage types were opposite to those of CO₂, which illustrates the interplay between the cage symmetry and guest shape and dynamics, and thus the gas selectivity.

Methane is the smallest hydrocarbon molecule which interacts with water hydrophobically. For the class of clathrate hydrates including the ionic or salt hydrates, methane is a suitable guest substance that fits well into the pentagonal dodecahedral (D) cages made of hydrogen-bonded water molecules.¹ The molecular size and the quasi-spherical shape of methane promote an isotropic distribution which is compatible with the void space of the D cages which can have a variety of local symmetries. Some recently obtained understanding of guest distributions in ionic clathrate hydrates suggests a novel method for modifying the preference of guest occupancies by controlling the D cage symmetries.²

The ionic clathrate hydrates form from water and ionic guests such as tetra-*n*-butylammonium bromide (TBAB)³ under relatively mild conditions, *e.g.*, 290 K for CH₄ + TBAB hydrates *versus* 275 K for the cubic sI CH₄ hydrate, at 3 MPa.⁴ Hence many potential applications of these materials are suggested, *e.g.*, gas separation and cool energy storage.⁵ Ionic hydrates can have the D cages occupied by small secondary guest gases, *e.g.*, CH₄, N₂ and CO₂.^{2,4c,6} In the hydrate structures formed with tetra-*n*-butylammonium and tetra-*n*-butylphosphonium salts,

Selective occupancy of methane by cage symmetry in TBAB ionic clathrate hydrate†

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there are two distinct types of D cages with significantly different shapes, particularly due to the water displacement by the large cation.⁷ Although they have almost the same volume, the distorted D cage (D_A) and relatively regular D cage (D_B) have very different cage occupancies by CO₂, *i.e.*, 0.867 *versus* 0.490, respectively, with each cage having the anisotropic angular distribution for CO₂.² So far, the D_A cage found in the TBAB hydrates has shown that it would provide different selectivity even for similarly sized molecules such as N₂ and CO₂.^{1b,6b} Here, we consider CH₄, which is a preferred guest molecule for the spherical D cage in most of the known hydrate structures, as a guest gas in the TBAB hydrate cages.

We determined the structure of the TBAB + CH₄ hydrate by single crystal X-ray diffraction (SCXRD) measurements. The crystallographic details are given in Notes‡ and the ESL† It has an *Pm* orthorhombic lattice with a 21.0329(15) × 12.5972(9) × 12.0333(8) Å³ cell size in agreement with Jeffrey's type IV hydrate structure as well as the TBAB + CO₂ hydrate.^{2,3a} The cell size is slightly shrunk from the pure TBAB hydrate (21.060(5) × 12.643(4) × 12.018(8) Å³ at 93.1 K)^{7a} likely due to the CH₄ occupancy and attractive CH₄–water interactions, however, the unit cell is not doubled along the *b*-axis in contrast to the *Im* lattice of the TBAB + CO₂ hydrate with unit cell vectors of 21.0197(7), 25.2728(8), and 12.0096(4) Å.² If we assume that the *Im* structure has the *b*-lattice length simply double that of the *Pm* structure, actually this is not true as shown later, the ratios of the axes based on the present *Pm* structure of the TBAB + CH₄ hydrate are 1.00, 2.01 and 1.00 for *a*, *b* and *c* axes, respectively. These values show that the unit cell size of the *Im* structure is scarcely affected by CO₂, and the symmetry lowering is due to the aforementioned local interactions in the D cages between the host water and the CO₂ guests. The chemical composition of the crystal unit cell was determined to be TBAB·38H₂O·2.16CH₄. Methods and results of SCXRD measurements have been detailed in our previous papers^{2,7b,c} and the ESL† As in the case of TBAB + CO₂ hydrate, the TBA cation was incorporated into a fused T²P² supercage, built out of two tetrakaidecahedral (T) cages having 12 pentagonal faces and two hexagonal faces, and two pentakaidecahedral (P) cages having 12 pentagonal faces and three hexagonal faces (see Fig. S1 in the ESL†).

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† Electronic supplementary information (ESI) available: Experimental and computational details. CCDC 1431843. For ESI and crystallographic data in CIF or other electronic format see DOI: 10.1039/c6cc00264a



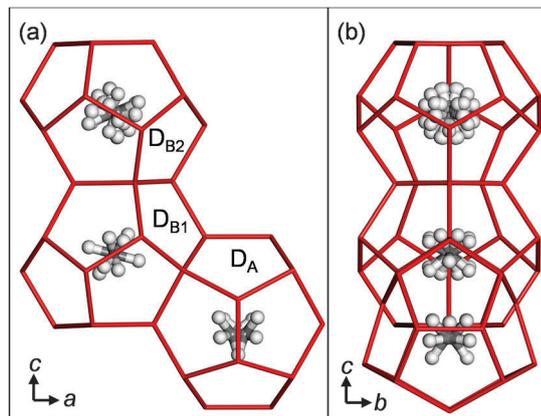


Fig. 1 Three D cages in the TBAB + CH₄ hydrate. (a) A view from the *b* axis. (b) A view from the *a* axis.

The bromide anion replaces a water molecule in the lattice of the T²P² cage and some water molecules are missing at the vertices where the cages join in order to accommodate the large cations. The cation and anion positions were the same as for the TBAB + CO₂ hydrate and the simple TBAB hydrate in the absence of guest gas.^{2,7a} Three distinct D cages were found in the structure as shown in Fig. 1: a strongly distorted cage (D_A) and two relatively regular cages (D_{B1} and D_{B2}), thus the cage framework can be described as T²P²·D_A·D_{B1}·D_{B2}. Symmetries of both D_{B1} and D_{B2} cages were 2/*m*, therefore, these should be identical for CH₄. Cage occupancies by CH₄ were 0.174 for D_A, 0.989 for D_{B1} and 0.993 for D_{B2}. Table 1 compares the D cages in the TBAB + CH₄ and TBAB + CO₂ hydrates. The relative CH₄ occupancies of the D_A and D_B cages were opposite to those of CO₂.²

The results show that CH₄ molecules prefer the regular quasi-spherical D_B cages rather than the distorted aspherical D_A cage, whereas CO₂ prefers the elongated D_A cage. The averaged occupancies over all cage types by CH₄ and CO₂ are close to each other, *i.e.*, 0.719 and 0.616, respectively. Table S2 in the ESI† further compares the D cage symmetries between ionic and canonical clathrate hydrates. This relationship between D cage occupancies gives an explanation for the formation conditions of TBAB hydrate formed with CH₄ and/or CO₂^{4c,6a,8} (see Fig. 2). Phase equilibrium data of the TBAB + CH₄ and TBAB + CO₂ hydrates with TBAB mole fractions in the aqueous phase $x_{\text{TBAB}} = 0.013$ and 0.026 are available.^{4c,6a} They are fairly close as shown in Fig. 2. With $x_{\text{TBAB}} \sim 0.013$, equilibrium conditions of TBAB hydrates formed with CH₄ + CO₂ mixed gas were also reported.⁸ With this solution, there is about 0.3 K of equilibrium temperature difference between TBAB + CH₄ and TBAB + CO₂ hydrates. However, TBAB + CH₄ + CO₂

Table 1 The D cage symmetries and occupancies in the TBAB hydrate structure

Gas	D _A cage		D _B cage	
	Symmetry	Occupancy	Symmetry	Occupancy
CH ₄ ^b	<i>mm</i> 2	0.174	2/ <i>m</i>	0.991 ^a
CO ₂ ^b	2/ <i>m</i>	0.867	2	0.490

^a An averaged value for D_{B1} and D_{B2} cages. ^b Reported in ref. 2.

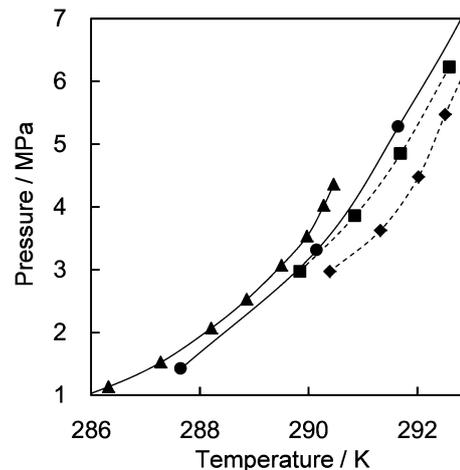


Fig. 2 Phase diagram for TBAB + CH₄ and/or CO₂ hydrates. TBAB + CO₂ hydrate:^{6a} Δ , $x_{\text{TBAB}} = 0.026$; \blacktriangle , $x_{\text{TBAB}} = 0.013$. TBAB + CH₄ hydrates:^{4c} \circ , $x_{\text{TBAB}} = 0.023$; \bullet , $x_{\text{TBAB}} = 0.014$. TBAB + CH₄ + CO₂:⁸ \blacklozenge , $x_{\text{TBAB}} = 0.014$, $y_{\text{CO}_2} = \sim 0.55$; \blacksquare , $x_{\text{TBAB}} = 0.014$, $y_{\text{CO}_2} = \sim 0.35$, where x_{TBAB} and y_{CO_2} denote the mole fraction of TBAB in an aqueous phase and the mole fraction of CO₂ in a gas phase, respectively.

hydrates show higher equilibrium temperatures, ~ 1 K higher than the TBAB + CH₄ or CO₂ hydrate. In addition, the mixed CH₄ + CO₂ gas hydrate in the incongruent solution ($x_{\text{TBAB}} = 0.014$) is as stable as the TBAB + CO₂ and TBAB + CH₄ hydrates in the congruent solution ($x_{\text{TBAB}} = 0.026$ with the hydration number: 38). This is likely due to the fact that the distorted D_A cages are occupied by CO₂ and the regular D_B cages are occupied by CH₄. Consequently, the TBAB hydrate structure may be stabilized with the mixed gas.

Fig. 3 shows the cages with different shapes and symmetries occupied by CH₄ and CO₂. In the D_A cage, CH₄ is displaced from the central position along the *c* axis. In contrast, the carbon atom of CO₂ is located at the center, and the CO₂ guests have rotational motion in a plane which is almost perpendicular to the *b* axis and along the *c* axis.

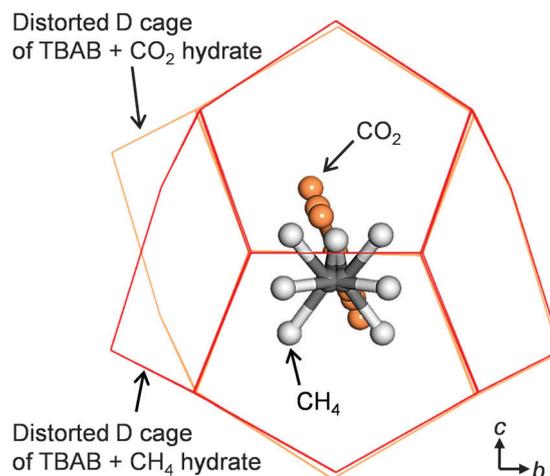


Fig. 3 Comparison of guest gas positions between CH₄ and CO₂ in the distorted D cage (D_A) of TBAB hydrate from a view of the *a*-axis.



We performed molecular dynamics (MD) simulations on this system with CH₄ cage occupancies corresponding to experimental values. The goal of the simulations is to detect the differences of structural and dynamical aspects of the methane guest motions in the two types of D cages. The details of the MD simulations are given in Notes[‡] and the ESI.[†] In Fig. 4 we show the calculated radial distribution function (RDF) for cage water oxygen atoms with the methane carbon atoms at 150 and 250 K.

The D_A cages are distorted and elongated along the direction which allows the methane guest molecules to approach some cage water molecules at distances between 3 and 3.2 Å. There is much smaller probability of seeing methane molecules in the D_B cages within this distance range of the cage waters. The elongation of the D_A cage also leads to the first peak in the C...OW RDF in the D_A cage being broader than the peaks in the D_B cages. As the ionic clathrate hydrate does not dissociate over the temperature range of the simulations (150 K to 250 K), there will not be a change in the coordination number of the methane guests in the D_A and D_B cages, with respect to water oxygens, associated with the changes in temperature. The changes in the RDF include broadening of the first peaks at higher temperature which is related to the increase of amplitude of the methane motions in the cages.

The dynamics of the methane molecules in the D cages is studied using the velocity autocorrelation function (VACF), $\psi(t) = \langle v_i(t) \cdot v_i(0) \rangle / \langle v_i(0) \cdot v_i(0) \rangle$ where $v_i(t)$ is the velocity of methane i at time t and the brackets represent an average over all methane molecules in a particular type of D cage. The velocity autocorrelation functions for methane in the D_A, D_{B1}, and D_{B2} cages at 150 and 250 K are shown in Fig. 5.

The motions of the methane molecules in the D_{B1} and D_{B2} cages are more regular and a greater number of rattling vibration cycles are observed in these cage types before the motion is randomized. The motions of the methane molecules in the D_A cages have a smaller initial period and the VACF of the methane guests in

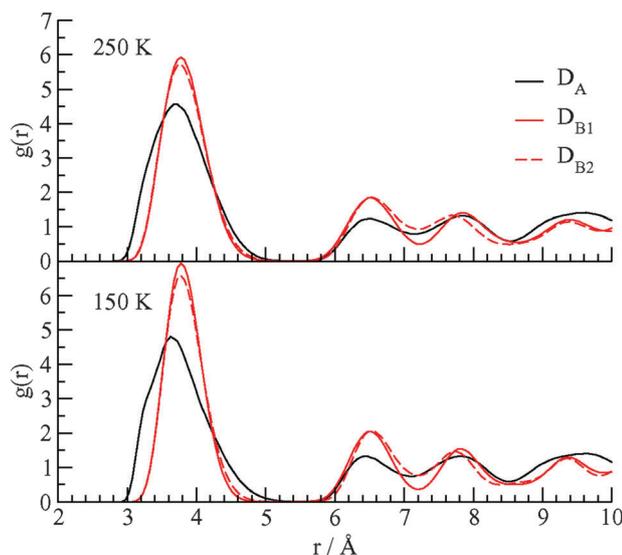


Fig. 4 The RDF for the methane carbon atom–water oxygen atom in the three D cages of the TBAB + CH₄ ionic hydrate at 150 and 250 K.

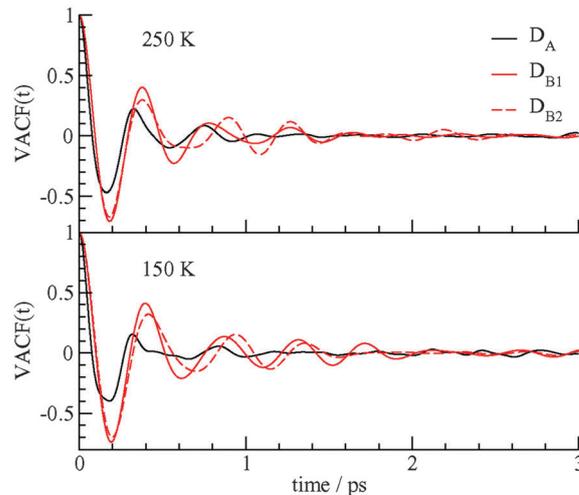


Fig. 5 The velocity autocorrelation function of the methane carbon atoms in the D_A, D_{B1}, and D_{B2} cages at 150 and 250 K.

the D_A cages decays to 0 after about 1 ps, whereas periodic behavior for the motion of the methane guests in the D_B cages continues over times of 2 ps or longer. The dynamics of the methane molecules in the D_{B1} and D_{B2} cages are somewhat different and $\psi(t)$ is sensitive enough to differentiate between these two cage types. The different dynamics of the methane guests in the D_A and D_B cages may lead to different spectroscopic signatures for the guests in the two types of cages. The VACF at different temperatures also reflect the higher amplitude, slightly higher frequency motion, and faster decay of velocity correlations of methane guests at 250 K compared to 150 K. These results of Fig. 4 and 5 confirm that the low symmetry D cages hold the methane guest molecules more loosely than the more structurally regular D_B cages.

The results of methane uptake by TBAB show that physical adsorption properties of clathrate hydrates for developing gas capture technology can be further improved by crystal engineering with ionic guests. For example, to maximize the separation of CO₂ from methane it would be best to have a clathrate host that would have all asymmetric cages to select CO₂ over methane. So far, the drastic differences in D cage capacities for CH₄ and CO₂ have remained unknown and all D cages have been treated as identical. Information as obtained in these structural studies is also critically important for developing quantitative models for the sorption behavior.

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Notes and references

[‡] Crystallographic data for TBAB·38H₂O·0.688CH₄: MoK α radiation ($\lambda = 0.71070$ Å), crystal size $0.6 \times 0.2 \times 0.2$ mm, orthorhombic, space group *Pmma*, $a = 21.0329(15)$ Å, $b = 12.5972(9)$ Å, $c = 12.0333(8)$ Å, $Z = 2$, volume $3188.3(4)$ Å³, density 1.067 g cm⁻³; $\mu(\text{MoK}\alpha) = 0.723$ mm⁻¹, temperature $100.0(1)$ K, scan mode ($1.69 < 2\theta < 29.00$), reflections collected/unique $4515/4169$, 528 parameters. The structure was solved



and refined using the SHELX program⁹ with full matrix least squares on F^2 , GOF (F^2) 1.175; final R indices ($I > 2\sigma(I)$): $R_1 = 0.0549$, $wR_2 = 0.1649$; R indices (all data) $R_1 = 0.0592$, $wR_2 = 0.1670$. CCDC 1431843. The intermolecular van der Waals potentials between atoms in different molecules are considered to be the sum of Lennard-Jones (LJ) and electrostatic point charges. The TIP4P four-charge water molecular model was used.¹⁰ TBA ions were modelled with the general AMBER force field,¹¹ and the force field for bromide by Canongia Lopes and Pádua.¹² The Tse-Klein-McDonald potential^{1h} was chosen for methane helper guest molecules. Partial electrostatic charges on the atoms of the guest molecules were determined from charges using electrostatic potential grid (CHELPG) calculations¹³ with the Gaussian 09 suite of programs¹⁴ at the B3LYP/6-311++G(d,p) level of theory. Temperatures of 100 (X-ray structure determination temperature), 200, and 250 K with ambient pressure were used in the simulations. The used parameters and further detailed methods are given in the ESI.†

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