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# Tuning the surface structure of supported PtNi<sub>x</sub> bimetallic electrocatalysts for the methanol electro-oxidation reaction†

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**The structures of PtNi<sub>x</sub> nanoalloy particles were modified through thermal annealing in different atmospheres. The evolution of surface structures was uncovered by advanced transmission electron microscopy, and the structure–function correlation in methanol electro-oxidation was probed. It provided new insights into the design and synthesis of highly efficient electrocatalysts.**

Nowadays, fuel cells have attracted extensive interest in the search for a solution to energy shortage and environmental pollution.<sup>1–3</sup> Direct methanol fuel cells (DMFCs) have become increasingly important because of their low cost, high-efficiency energy conversion, less pollution and safe storage as well as transport.<sup>4</sup> Platinum (Pt) as the most generally accepted active metal for the cathode electrode material<sup>5</sup> plays a very important role in fuel cells, but the high cost and poor stability limit its application. The introduction of other metal elements, especially the low-cost 3d transition metals (*e.g.* Fe, Co and Ni), not only reduces the usage of Pt, but also modifies the electronic and chemical properties, which could result in enhanced electrocatalytic activity.<sup>6–10</sup> It has been traced and illustrated that electrocatalytic reactions take place on the surface of catalysts,<sup>11,12</sup> where adsorption and desorption occur. Therefore, the activity of electrocatalysts highly depends on their surface properties,<sup>13</sup> and higher catalytic activity can be obtained by tuning the surface structure and the composition of Pt-based nanomaterials.<sup>14,15</sup> In order to control the surface structure and composition at the atomic scale, various methods have been developed, including multistep synthesis<sup>16</sup> and electrochemical dealloying.<sup>17–19</sup> For instance, Stamenkovic *et al.* designed and synthesized a novel Pt<sub>3</sub>Ni catalyst for the oxygen reduction

reaction (ORR).<sup>1</sup> The outermost layer and the second atomic layer of Pt<sub>3</sub>Ni nanoparticles (NPs) are Pt-rich and Ni-rich, respectively, and the weak interaction between them makes the Pt<sub>3</sub>Ni(111) more active. However, there are few reports regarding the surface/interface structural evolution induced by thermal annealing in different atmospheres on Pt-based nanoalloy NPs in detail, although it is significant for a more precise and easy control of the fine structures of catalysts.<sup>20–24</sup> Thanks to the developments in transmission electron microscopy (TEM), by which the structural investigation of the solid catalyst at the atomic level becomes possible during the synthesis process and catalysis reaction.<sup>25,26</sup>

Herein, the polyhedral PtNi<sub>2</sub> NPs synthesized using a solvo-thermal method were supported on multilayer graphene (PtNi<sub>2</sub>/Gr), and then PtNi<sub>2</sub>/Gr was thermally annealed in a distinct atmosphere. The carbon monoxide (CO)-induced surface reconstruction of the polyhedral PtNi<sub>2</sub> NPs led to a core-shell structure with a Pt-rich surface. The methanol electro-oxidation was selected as the probing reaction for testing the performance of PtNi<sub>2</sub> nanoalloy NPs with different structures. The results indicate that the interaction between CO and the metal (Pt and Ni) during the thermal treatment provides a facile approach to tune the surface structure of the PtNi<sub>2</sub> catalyst.

Fig. 1 shows representative low-magnification TEM images of PtNi<sub>2</sub> NPs. The particle size distribution (PSD) of PtNi<sub>2</sub> NPs is from 5.4 nm to 12.6 nm, and the average size is *ca.* 9.6 nm based

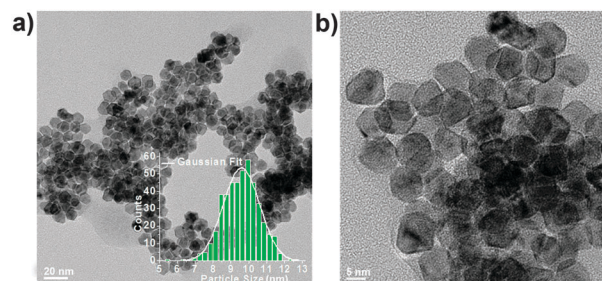


Fig. 1 Representative TEM images of PtNi<sub>2</sub> NPs before supporting. The inset in (a) is the corresponding particle size distribution.

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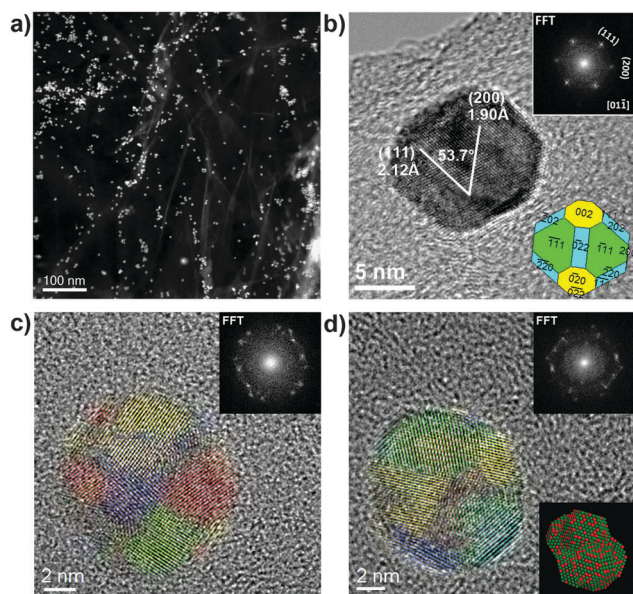
on the statistical analysis of the PSD histogram (the inset in Fig. 1a). Subsequently, the synthesized PtNi<sub>2</sub> NPs were supported on multilayer graphene. The obtained supported PtNi<sub>2</sub> NPs were heated to 300 °C for 4 hours in argon (Ar) and CO gases, named as PtNi<sub>2</sub>/Gr-Ar and PtNi<sub>2</sub>/Gr-CO, respectively.

X-ray diffraction (XRD) patterns (Fig. S1, ESI†) show that all of the PtNi<sub>2</sub> catalysts exhibit face-centred cubic structures, without a single Pt or Ni phase. The characteristic diffraction peaks of PtNi<sub>2</sub> NPs between Pt and Ni diffraction-peak positions can be indexed (111), (200) and (220) peaks based on JCPDS cards of Pt (No. 04-0802) and Ni (No. 04-0850). The full-width half-maximum (FWHM) values of (111), (200) and (220) peaks decrease after thermal treatment, indicating an increase in the crystallinity of PtNi<sub>2</sub> NPs. The main peaks of PtNi<sub>2</sub>/Gr-CO shift toward a higher angle (Ni phase) compared with that of PtNi<sub>2</sub>/Gr, which indicates that some Pt atoms segregate and the other Pt and Ni atoms form the PtNi<sub>x</sub> (*x* > 2) nanoalloy.

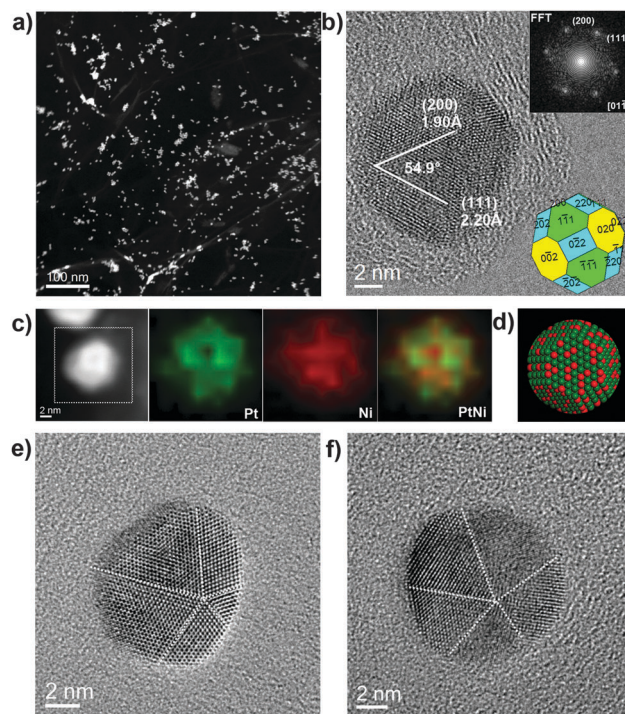
Advanced electron microscopy techniques were used for uncovering the shape, composition and fine structures of the supported PtNi<sub>2</sub> NPs. The PtNi<sub>2</sub> NPs are highly dispersive on the multilayer graphene in the PtNi<sub>2</sub>-Gr sample, as shown in the high-angle annular dark field-scanning transmission electron microscopy (HAADF-STEM) image (Fig. 2a). High-resolution TEM (HRTEM) observations reveal that the shape of PtNi<sub>2</sub> NPs is polyhedral (Fig. 2b). The lattice fringes (200) and (111), *d*-spacing 1.90 Å and 2.12 Å, of PtNi<sub>2</sub> with a characteristic acute angle of 53.7° were identified on the NPs, referring to the JCPDS cards (Pt: No. 04-0802, Ni: No. 04-0850, space group: *Fm*3̄*m*). Summarizing the analyses of HRTEM images, the exposed facets of PtNi<sub>2</sub> NPs can

be identified as the {002}, {111} and {220} planes (see the insets in Fig. 2b). However, only a few PtNi<sub>2</sub> NPs with good crystallinity exist in PtNi<sub>2</sub>/Gr, and most of PtNi<sub>2</sub> NPs are polycrystalline with the size of the constituent single crystalline domains in the range between 1.0 and 8.0 nm, as shown in the colored HRTEM images (Fig. 2c and d). It is consistent with the XRD pattern that the diffraction peaks with the broad FWHM values in Fig. S1 (ESI†). According to above structural information, a structural model was proposed, showing the random distributions of Pt and Ni in most of polycrystalline PtNi<sub>2</sub> NPs (the lower-right inset in Fig. 2d).

Aiming to modify the surface structure of PtNi<sub>2</sub> NPs, the PtNi<sub>2</sub>/Gr sample was thermally treated (300 °C) in Ar and CO gases. Fig. 3 shows the microscopic features of the PtNi<sub>2</sub>/Gr sample treated in Ar gas. The PtNi<sub>2</sub> NPs remained highly dispersed on the supports except for slight agglomeration (Fig. 3a and Fig. S3a, ESI†). Most of PtNi<sub>2</sub> particles were rearranged to polyhedral NPs with good crystallinity, which has been identified by HRTEM (Fig. 3b and Fig. S3b, ESI†). According to HRTEM analyses, the PtNi<sub>2</sub> NPs are also encapsulated by the {002}, {111} and {220} planes (the inset in Fig. 3b). Moreover, there are some multiply twinned NPs in the PtNi<sub>2</sub>/Gr-Ar, displayed in Fig. 3e and f. This result is in accordance with the XRD results that the FWHM value of diffraction peaks decreased slightly compared with that before annealing. EDS elemental maps show the distributions of Pt and Ni are uniform in PtNi<sub>2</sub> NPs (Fig. 3c and Fig. S3c, S5a, ESI†),



**Fig. 2** Low-magnification HAADF-STEM image of PtNi<sub>2</sub>/Gr (a), HRTEM images of polyhedral PtNi<sub>2</sub> NPs with good crystallinity (b), and polycrystalline PtNi<sub>2</sub> NPs (c and d). The top-right insets in (b–d) are the corresponding local Fast Fourier Transformation (FFT), and the lower-right insets in (b and d) are the shape and relevant atomic model (Pt–green, Ni–red) of PtNi<sub>2</sub> NPs. The constituent single-crystalline domains are indicated in different colours in (c) and (d).



**Fig. 3** Low-magnification HAADF-STEM image of PtNi<sub>2</sub>/Gr-Ar (a), HRTEM images of polyhedral PtNi<sub>2</sub> NPs with good crystallinity (b), multiply twinned PtNi<sub>2</sub> NPs (e and f), EDS elemental maps (c) and relevant atomic models of PtNi<sub>2</sub> NPs (d: Pt–green, Ni–red). The insets in (b) are the corresponding local FFT and shape of the PtNi<sub>2</sub> NPs. The dashed white lines in (e) and (f) are the twin boundaries.





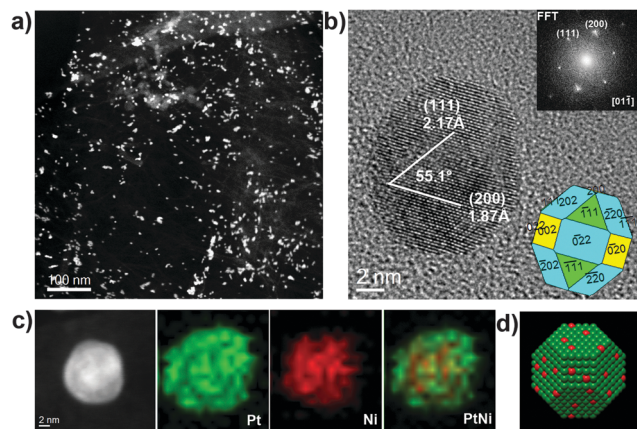


Fig. 4 Low-magnification HAADF-STEM image of PtNi<sub>2</sub>/Gr-CO (a), HRTEM image of PtNi<sub>2</sub> NP (b), EDS elemental maps (c) and relevant atomic model (d: Pt-green, Ni-red) of PtNi<sub>2</sub> NPs. The insets in (b) are the corresponding local FFT and shape of PtNi<sub>2</sub> NP.

indicating the Pt and Ni rearranged and formed a randomly mixed bimetallic type after thermal annealing in Ar gas. The atomic model is shown in Fig. 3d, displaying a three-dimensional structure with a random distribution of Pt and Ni in PtNi<sub>2</sub> NPs.

The structures of the PtNi<sub>2</sub>/Gr treated in CO gas exhibited distinct changes. Low-magnification (S)TEM images show that the PtNi<sub>2</sub> NPs are well distributed on the supports (Fig. 4a and Fig. S4a, ESI†). HRTEM reveals that all of the PtNi<sub>2</sub> NPs retain the polyhedral shape with exposed {002}, {111} and {220} facets, as shown in Fig. 4b and Fig. S4b (ESI†). On these exposed planes, there exists some surface roughness (*e.g.* corner and step defects), ascribed to the atom migration induced by the CO thermal annealing. These NPs are of good crystallinity. Furthermore, EDS elemental mapping reveals that the Pt element is enriched on the surface of PtNi<sub>2</sub> NPs, which formed the core (PtNi<sub>x</sub>,  $x > 2$ )-shell (Pt) type (Fig. 4c and Fig. S4d, S5b, ESI†). The HR-STEM image in Fig. S4c (ESI†) clearly displays that brighter atoms (Pt) surround the Ni atoms, indicating a core-shell structure. Comparing the fine structures of PtNi<sub>2</sub> NPs treated with thermal annealing in different gases, the shapes and surface structures of PtNi<sub>2</sub> NPs can be tuned at the atomic level. Except for the microscopic features mentioned above, the composition of the series of PtNi<sub>2</sub> samples has been characterized by inductively coupled plasma mass spectrometry (ICP-MS) (Table S1, ESI†). The composition of the PtNi<sub>x</sub> nanoalloy was kept nearly the same (the Pt/Ni atomic ratio was about 1/2) at the macroscopic scale.

For studying the effects of the tuned microstructures and for illustrating the structure-function relationship, the methanol electro-oxidation reaction (MOR) was selected as the probing experiment. The MOR performance on PtNi<sub>2</sub>-based samples, PtRu/C and Pt/C was evaluated by cyclic voltammetry (CV) curves (Fig. 5a), and the normalized activity (current density) at the respective peak potential is shown in Fig. 5b. The better activity with 33% and 47% enhancement is presented in the PtNi<sub>2</sub>/Gr-CO, compared to those of other two PtNi<sub>2</sub>-based catalysts. The enhanced activity of the PtNi<sub>2</sub>/Gr-CO sample can be attributed to the rearrangement of Pt and Ni induced

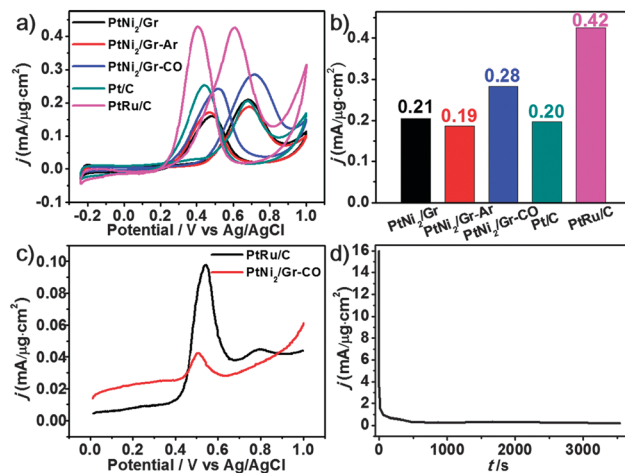


Fig. 5 (a) CV curves of PtNi<sub>2</sub>-based catalysts PtRu/C and Pt/C in N<sub>2</sub>-saturated 0.5 M H<sub>2</sub>SO<sub>4</sub> + 2 M CH<sub>3</sub>OH solutions at a scan rate of 50 mV s<sup>-1</sup>. (b) The normalized current density histogram at the respective peak potential for all catalysts. (c) CO-stripping voltammograms at 50 mV s<sup>-1</sup> of PtNi<sub>2</sub>/Gr-CO and commercial PtRu/C (20 wt% Pt, 10 wt% Ru, Alfa Aesar) catalysts in 0.5 M H<sub>2</sub>SO<sub>4</sub>. (d) Chronoamperometric curve of MOR at 0.7 V on the PtNi<sub>2</sub>/Gr-CO catalyst.

by the CO thermal treatment and the much more amount of Pt atoms segregating on the surface.<sup>11,27,28</sup> The obtained best performance of the PtNi<sub>2</sub>-based catalysts is between those of PtRu/C and Pt/C, which is also comparable to other PtNi-based catalysts reported previously.<sup>29,30</sup> The good CO tolerance of the PtNi<sub>2</sub>-based samples are represented by the ratio (1.1–1.3) of forward ( $I_f$ ) and backward ( $I_b$ ) peak currents.<sup>31</sup> The CO-stripping of the PtNi<sub>2</sub>/Gr-CO sample was also investigated compared to that of the commercial PtRu/C catalyst. From the lower current density and peak potential, it could be observed that the PtNi<sub>2</sub>/Gr-CO sample shows a smaller amount of adsorbed CO and weaker CO adsorption on the surface during the electrocatalytic reaction than that of the commercial PtRu/C catalyst. In this case, the enhancement of CO tolerance can be explained by the electronic effect mechanism and the formation of Pt-rich shells.<sup>29,30</sup> The transition metal Ni changes the electronic properties of the PtNi<sub>x</sub> NPs *via* electron transfer between the Pt and Ni, which weakens the Pt-CO bond and suppresses the CO adsorption on the Pt rich surface. As for the stability, the MOR current density after 3000 s was also about 0.20, which is little lower than that of the 0.28 (Fig. 5a), as shown in Fig. 5d.

Several factors determine the surface segregation of the core-shell structure, which has been unravelled and interpreted based on experimental and theoretical studies,<sup>32–38</sup> such as cohesive energy, surface energy, atomic radius and electronegativity. The surface segregation energy (SE) is the main driving force for the formation of core-shell structures.<sup>39</sup> For instance, according to density functional theory (DFT) calculation, the SE of the 55-atom Ni-core and Pt-shell model is 0.42 eV. The core (Ni)-shell (Pt) structure is preferred when SE is positive, which is dominated by the smaller Wigner Seitz radius of Ni. It means that the compressive strain was relieved while the smaller atom goes into the core. This explanation is also supported by some related investigations.<sup>40–43</sup> In this work, CO molecules play a principal



role in the surface rearrangement. Based on the different CO adsorption energies ( $E_{\text{ad}}$ ) on Pt(111) and Ni(111) calculated by DFT,<sup>42</sup> a reduction surface energy of about 0.67 eV per Pt atom rearranged to the surface was observed. Therefore, the segregation of Pt atoms is favourable during the thermal treatment in the CO surrounding. However, this phenomenon is not obvious in Ar gas due to the inert environment. Otherwise, the CO molecules adsorbed on the PtNi<sub>2</sub> NPs surface can restrain the shape change and retain the polyhedral morphology.<sup>44,45</sup>

In summary, we present a method to obtain a Pt-rich shell structure of PtNi<sub>2</sub> NPs supported on a multilayer graphene that retains the polyhedral shape with exposed {002}, {111} and {220} facets, which exhibits a better activity due to electronic and geometric effects in the MOR. Advanced electron microscopy unambiguously identified the structure evolutions of PtNi<sub>2</sub> NPs from polycrystalline materials to randomly mixed and core-shell bimetallic structures induced by thermal annealing in Ar and CO gases. It offers a good example on the investigation of the structural evolution of bimetallic systems modified by different approaches, and reveals the structure–function correlation in chemical reactions. This work also provides a new insight into the interaction between CO molecules and metals, which could further help in the design of high activity nanoalloy catalysts by precisely tuning the shape, structure and composition.

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