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## Molecular engineering of benzothienoisindigo copolymers allowing highly preferential face-on orientations†

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Orientation of conjugated polymers is increasingly important in organic photovoltaics (OPV) to achieve high power conversion efficiency (PCE). The optimized orientation of conjugated backbones for photo-generated charge carriers in OPV devices is in contrast to organic semiconductor devices, demanding new strategies to control and realize face-on orientation of conjugated systems onto substrates. Here we report new conjugated polymers composed of electron-accepting benzothienoisindigo (BTIDG), an asymmetric unit of isoindigo and thienoisindigo. BTIDG was coupled with weakly electron-donating thiazolothiazole or benzobisthiazole, concurrently leading to moderate optical band gaps (1.41–1.52 eV) and the highest occupied molecular orbital (–5.35 to –5.50 eV). The alkylthiophene spacer between BTIDG and the donor unit provided a marked control over the orientation of polymers, among which the degree of face-on orientation as high as 95% was revealed by grazing incidence X-ray diffraction. The maximum PCE was improved up to 4.2% using the system with [6,6]-phenyl-C<sub>71</sub>-butyric acid methyl ester (PC<sub>71</sub>BM). We present a useful basis on the structure (orientation)–property (OPV output) relationship to lay down new guidelines for the design of efficient solar cell materials.

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## Introduction

A low bandgap polymer (LBP) constituting of electron donor and acceptor units is expected to create an efficient organic photovoltaic device (OPV),<sup>1–4</sup> due to its wide photoabsorption and the deep highest occupied molecular orbital (HOMO). Thienoisindigo (TIDG), analogous to isoindigo (IDG),<sup>5–7</sup> is a comparatively new electron acceptor which we<sup>8</sup> and other groups<sup>9,10</sup> have reported in 2012–2013. Replacing the phenyl of IDG by thiophene led to a more planar  $\pi$ -plane and stronger electron accepting ability than IDG. However, TIDG resulted in an excessively narrow optical bandgap ( $E_g^{\text{opt}}$ , 1.0–1.1 eV) in conjugation with the benchmark donors (benzodithiophene and cyclopentadithiophene).<sup>8</sup> An unsatisfactory PCE of 1.4% was obtained as the best value for the blend of [6,6]-phenyl-C<sub>61</sub>-butyric acid methyl ester (PC<sub>61</sub>BM) and fluorene-TIDG polymer ( $E_g^{\text{opt}}$  = 1.6 eV).<sup>11</sup> In addition to their mismatches in energy alignment, we have identified the limiting factors of

photovoltaic performance, *i.e.* the short exciton lifetime, the loss *via* bulk charge recombination, and unfavorable orientation/morphology partly associated with insufficient solubility.<sup>11</sup> Therefore, simultaneous control of these factors through molecular engineering or processing is a crucial prerequisite towards boosting the PCE of TIDG-based polymers; however this is a challenging theme shared by all disciplines.

One of the synthetic routes to the TIDG unit we have reported for the first time is the dimerization of mono-keto and di-keto precursors in acetic acid, the former of which was prepared by reacting the di-keto precursor with hydrazine reagent followed by Wolff–Kishner type reduction.<sup>8</sup> Inspired by this quantitative and heterotic dimerization scheme, we attempted the synthesis of half-TIDG and half-IDG units from their relevant two precursors. This would bring moderate electron accepting nature with a good solubility as a result of its asymmetric half-distorted  $\pi$ -plane. In the course of research, Chen and Fréchet *et al.*<sup>12</sup> reported this benzothienoisindigo (BTIDG)-based polymer synthesized by an almost similar route and demonstrated a large improvement of PCE (4.7%) for the terthiophene-BTIDG copolymer. Notably, they realized selective manipulation of edge-on or face-on orientation of polymers by controlling the backbone planarity (IDG or BTIDG) suitable for field-effect transistors (FETs) or OPV devices, respectively. However, the high-performing BTIDG copolymer still includes some portions of edge-on contribution, suggestive of extra room for enhancing the portion of face-on orientated polymer crystallites where

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considerable improvement has been reported in both charge carrier transport and overall PCE of OPV devices.<sup>13</sup>

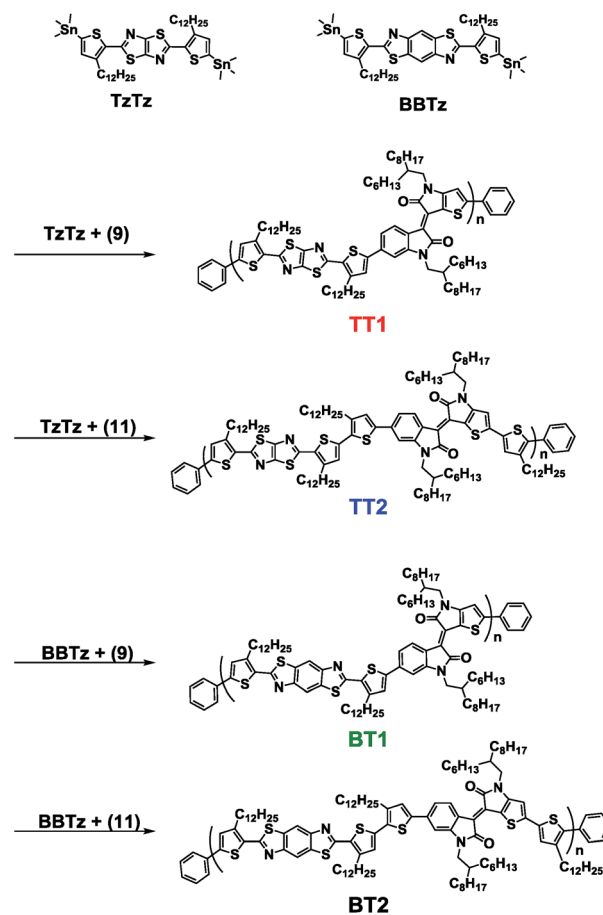
In this work, we synthesized novel BTIDG-based copolymers in combination with benzobisthiazole (BBTz)<sup>14,15</sup> or thiazolothiazole (TzTz),<sup>16,17</sup> where BBTz and TzTz are normally electron acceptors but used as weak donors<sup>18–20</sup> against BTIDG. This molecular design is intended to provide both optimal  $E_g^{\text{opt}}$  (ideally 1.37–1.45 eV from the calculation)<sup>21</sup> and HOMO levels. Furthermore, we exemplified the synergetic impact of the alkylthiophene spacer on the control of polymer orientation and OPV output.

## Results and discussion

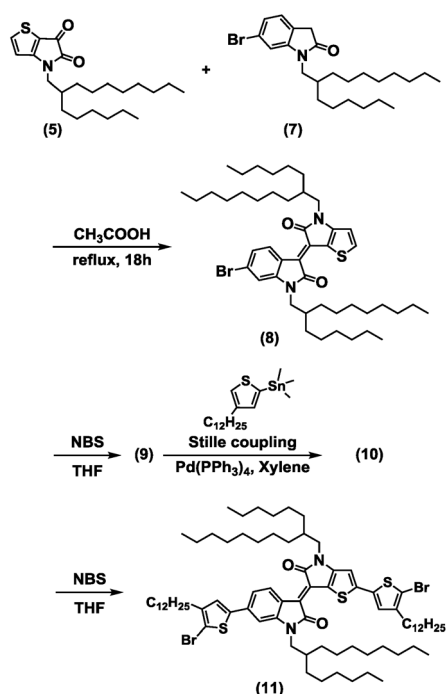
### Optical and electrochemical properties

The brominated BTIDG monomer (9) and alkylthiophene-BTIDG monomer (11) were synthesized according to the route of Scheme 1. Synthesis and characterization of the compounds (1)–(8) including <sup>1</sup>H NMR and mass spectroscopy data are provided in the ESI and Scheme S1.† BTIDG-based copolymers were polymerized by the Stille coupling reaction (Scheme 2).

Fig. 1a shows the chemical structures of BTIDG-based copolymers: TT1, TT2, BT1, and BT2, in which the number (1 or 2) represents the length of the alkylthiophene spacer (thiophene or bithiophene). Electronic absorption spectra of BTIDG polymers were measured in chloroform solution (Fig. 1b) and thin film (Fig. 1c). The BTIDG polymers indicate the absorption edges at *ca.* 820–880 nm corresponding to the  $E_g^{\text{opt}}$  of 1.41–1.52 eV (Table 1). These  $E_g^{\text{opt}}$  values are finely tuned to the aforementioned ideal value for single cell OPV polymers,<sup>21</sup> which is a



Scheme 2 Synthesis of BTIDG polymers.



Scheme 1 Synthetic procedure of the BTIDG unit. Those of (5) and (7) are provided in the ESI.†

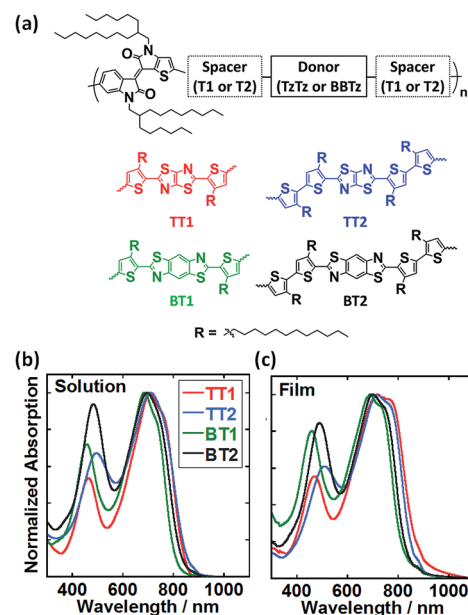


Fig. 1 (a) Chemical structures of BTIDG polymers. (b) Electronic absorption spectra of TT1 (red line), TT2 (blue line), BT1 (green line) and BT2 (black line) in chloroform solution at room temperature, respectively. (c) Electronic absorption spectra of TT1, TT2, BT1 and BT2 as films dropcast from chloroform solution.



Table 1 Optical and electrochemical properties of BTIDG polymers

Polymer	$M_w/\text{kg mol}^{-1}$ (PDI)	HOMO <sup>a</sup> /eV	LUMO <sup>b</sup> /eV	$\lambda_{\text{max}}^c/\text{nm}$	$E_g^{\text{optd}}/\text{eV}$
TT1	32 (1.7)	−5.35	−3.94	718	1.41
TT2	127 (3.5)	−5.43	−3.98	719	1.45
BT1	142 (2.1)	−5.50	−3.98	690	1.52
BT2	231 (3.9)	−5.41	−3.92	699	1.49

<sup>a</sup> Determined by PYS of drop-cast films. <sup>b</sup> HOMO +  $E_g^{\text{opt}}$ . <sup>c</sup> In film states.

<sup>d</sup> Determined from the onset of the electronic absorption spectra in film states.

notable advantage over the previous near-infrared absorbing TIDG polymers. In addition, the weight-averaged molecular weights ( $M_w$ ) of BTIDG polymers are as high as 127–231 kg mol<sup>−1</sup> (Table 1) except for TT1 (32 kg mol<sup>−1</sup>), which is much higher than those of the previous TIDG polymers (15–87 kg mol<sup>−1</sup>).<sup>8</sup> This is because the solubility was significantly improved by virtue of the long linear alkyl chains of the thiophene spacer and branched ones of asymmetric BTIDG.

Photoelectron yield spectroscopy (PYS, Fig. S1†) was performed to estimate the HOMO levels of polymer films, giving slightly varying HOMOs (−5.35 to −5.50 eV) and almost identical lowest unoccupied molecular orbitals (LUMO, −3.92 to −3.98 eV). These HOMOs are indeed much deeper than those of our previous TIDG polymers (HOMO: −4.7 to −5.4 eV, LUMO: −3.8 eV),<sup>8</sup> while the LUMO–LUMO offsets between polymer and PCBM (−4.2 eV)<sup>22</sup> became close to or even less than the empirically required minimum value (0.3 eV).<sup>23,24</sup>

Electron withdrawing nature of the BTIDG unit is clearly evident from the density functional theory (DFT) calculation shown in Fig. S2.† The LUMO is localized on the BTIDG unit, while the HOMO is spread over the polymer backbone of the donor (BT or TT), spacer (thiophene), and BTIDG, similar to the TIDG<sup>8</sup> and IDG<sup>25</sup> polymers. Thanks to the inherently high coplanarity and low LUMO, TIDG-thiophene and -naphthalene copolymers have uncovered ambipolar charge transport ( $\mu_h = 0.18 \text{ cm}^2 \text{ V}^{-1} \text{ s}^{-1}$  and  $\mu_e = 0.03 \text{ cm}^2 \text{ V}^{-1} \text{ s}^{-1}$ )<sup>26</sup> and ultrahigh hole mobility up to  $14 \text{ cm}^2 \text{ V}^{-1} \text{ s}^{-1}$ ,<sup>27</sup> respectively. In addition, nonbonding Coulomb interactions arising from the difference in the electronegativity of the sulfur atom of spacer thiophene and the nitrogen atom of TTz and BBTz could facilitate coplanarity of the  $\pi$ -conjugated polymer backbone. This is supported by the DFT calculation exhibiting that the total energy of N $\cdots$ S conformation is 7 kJ mol<sup>−1</sup> lower than that of S $\cdots$ S conformation (Fig. S3†), consistent with that of chalcogen (S $\cdots$ O)<sup>26</sup> and N $\cdots$ S interactions<sup>28</sup> reported previously.

### Polymer orientation and hole mobilities

Polymer orientation in pristine and PC<sub>61</sub>BM blend films was characterized by two-dimensional grazing-incidence X-ray diffraction (2D-GIXRD) shown in Fig. 2. Thin spin-coated pristine films (ca. 60 nm) contained an intense tail of direct X-ray overlapping the weak diffraction of polymer in the out-of-plane low angle diffraction region. This hampers the quantitative

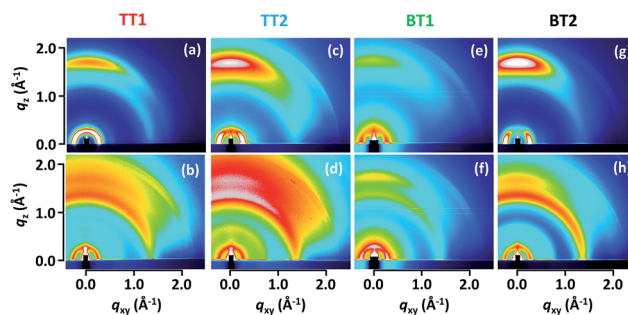


Fig. 2 2D-GIXRD patterns of pristine TT1 (a), TT2 (c), BT1 (e) and BT2 (g) drop-cast on the ITO substrate. (b), (d), (f) and (h) are the patterns of optimal OPV devices (polymer and PC<sub>61</sub>BM blends).

analysis of face-on and edge-on orientation (Fig. S4†), and thus the thick pristine films (ca. 1  $\mu\text{m}$ ) prepared by drop-casting were evaluated. It needs to be mentioned, though, that the  $\pi$ – $\pi$  stacking distance and face-on/edge-on orientations were mostly unchanged between spun-coat and drop-cast films. The blend films were prepared in the same way with the optimal OPV device fabrication (*vide infra*). Surprisingly, all of the pristine polymers indicate profound face-on orientation with negligibly small edge-on peaks. The  $\pi$ – $\pi$  stacking distances are moderate, ranging from 3.57 Å for BT1 to 3.67 Å for TT2 (Table 2). Although the  $\pi$ – $\pi$  stacking distances did not display obvious variations by the molecular structure, the lamellar packing distances were rationally increased from the mono-thiophene spacer (24.8 Å for both TT1 and BT1) to the bis-thiophene one (26.1 Å for TT2 and 25.4 Å BT2).

The ratio of face-on orientation ( $R_{\text{face-on}}$  in %) was evaluated by  $I_{\text{ip}}/(I_{\text{ip}} + I_{\text{op}})$ , where  $I_{\text{ip}}$  and  $I_{\text{op}}$  are the peak intensities of (100) diffraction in the in-plane and out-of-plane directions, respectively (Fig. S5†). It should be noted that  $R_{\text{face-on}}$  is an approximation of face-on fraction in the XRD-active crystalline part,<sup>17,29</sup> and thus  $R_{\text{face-on}}$  is not directly associated with the crystallinity of polymer. Besides, out-of-plane diffraction tends to appear with higher intensity than its actual composition, due to the combination of the thin film geometry and the constraint of one axis to the sample plane in the polar figure.<sup>30</sup> By taking these into account,  $R_{\text{face-on}}$  evaluation in this study is supposed to be

Table 2 2D-GIXRD and SCLC data of BTIDG polymers

Film (p/n ratio)	$\pi$ – $\pi^a/\text{\AA}$	IL <sup>b</sup> /Å	$R_{\text{face-on}}^c$	$\mu_h/10^{-5} \text{ cm}^2 \text{ V}^{-1} \text{ s}^{-1}$
TT1	3.64	24.8	67%	1.0
TT1 : PC <sub>61</sub> BM (1 : 1)	3.64	24.8	40%	6.8
TT2	3.67	26.1	80%	5.9
TT2 : PC <sub>61</sub> BM (1 : 2)	3.71	26.1	73%	10
BT1	3.57	24.8	93%	0.18
BT1 : PC <sub>61</sub> BM (1 : 2)	3.54	23.7	68%	4.9
BT2	3.66	25.4	95%	120
BT2 : PC <sub>61</sub> BM (1 : 2)	3.64	24.8	75%	12

<sup>a</sup> Calculated from the (010) peaks in the out-of-plane direction.

<sup>b</sup> Interlamellar distance calculated from the (100) peaks in the in-plane direction. <sup>c</sup> Calculated by  $I_{\text{ip}}/(I_{\text{ip}} + I_{\text{op}})$ .



slightly underestimated, since the out-of-plane diffraction at the small scattering vector attributed to edge-on orientation could be stronger than is the actual case. As listed in Table 2, pristine **BT2** shows a pronounced  $R_{\text{face-on}}$  as high as 95% (the ratio of face-on to edge-on,  $I_{\text{ip}}/I_{\text{op}}$  is *ca.* 19). To the best of our knowledge, this is the top-class preferential face-on orientation reported so far in OPV polymers. Examples of high  $R_{\text{face-on}}$  polymers are thiazolothiazole (91%)<sup>17</sup> and dicarboximide-dithiophene (PBTI3T, 93%, face-on : edge-on = 14)<sup>31</sup> polymers. The pioneering LBPs of benzodithiophene (BDT)-thienopyrrolidone (PBDTTPD)<sup>32,33</sup> and BDT-thienothiophene (PTB7 and PBDTTT-C-T)<sup>34–36</sup> of which the PCE exceeded 6–7% in 2010 also have high face-on orientation. We extracted an  $R_{\text{face-on}}$  of 94% for pristine PTB7 from its 2D-GIXRD image (Fig. S5†),<sup>37</sup> corroborating the high face-on orientation of the successful polymer. Other early LBPs such as carbazole-benzothiaziazole (PCDTBT)<sup>38</sup> and cyclopentadithiophene-benzothiaziazole (PCPDTBT)<sup>39</sup> also indicate the face-on preferred 2D-GIXRD images along with rather amorphous features. Since the face-on orientation is advantageous for vertical charge transport in OPV cells,<sup>13,17,31</sup> recent development of high performance LBPs is directed to control the polymer orientation as well as bandgap engineering. The second highest  $R_{\text{face-on}}$  of 93% was found for **BT1** followed by **TT2** (80%) and **TT1** (67%).

Due to the complicated factors affecting the polymer orientation, a rational molecular design of face-on polymers remains poorly understood. Solubility, shape of solubilizing alkyl chains (branched one is likely more favorable for face-on than linear one), and the length of alkyl chains have been supposed as the premiere factors that control polymer orientation. In the present study, the sizes of the  $\pi$ -plane (BT > TT) and spacer (bithiophene : 2 > thiophene : 1) were found to have an impact on improving face-on orientation. In sharp contrast, mixing with PC<sub>61</sub>BM significantly decreased the face-on orientation. The similar trend appeared in all blend films, where  $R_{\text{face-on}}$  dropped by 10–20% (Table 2). For instance, the most face-on-preferred **BT2** indicates a  $R_{\text{face-on}}$  of 75% in the PCBM blend, accompanied by a decrease of diffraction intensity.

Hole mobilities ( $\mu_{\text{h}}$ ) of BTIDG polymers and their blends with PC<sub>61</sub>BM were measured by using a space-charge limited current (SCLC) technique<sup>40</sup> (Table 2 and Fig. S6†). Remarkably, the most face-on-orientated pristine **BT2** demonstrated the highest  $\mu_{\text{h}}$  of  $1.2 \times 10^{-3} \text{ cm}^2 \text{ V}^{-1} \text{ s}^{-1}$ , while the other polymers gave 2–3 orders of magnitude smaller values ( $10^{-5}$  to  $10^{-6} \text{ cm}^2 \text{ V}^{-1} \text{ s}^{-1}$ ). The **BT2** : PC<sub>61</sub>BM blend film showed a  $\mu_{\text{h}}$  of  $1.2 \times 10^{-4} \text{ cm}^2 \text{ V}^{-1} \text{ s}^{-1}$  decreased by one order of magnitude; however, this is still higher than the other polymer blends ( $10^{-4}$  to  $10^{-5} \text{ cm}^2 \text{ V}^{-1} \text{ s}^{-1}$ ). Despite the reduced diffraction intensities upon mixing with PC<sub>61</sub>BM, the  $\pi$ - $\pi$  stacking and lamellar packing distances were mostly unchanged, indicative of no intercalation into the polymer crystallites. Interestingly, the  $\mu_{\text{h}}$  values of PC<sub>61</sub>BM-blended **TT1**, **TT2**, and **BT1** are comparable to or even higher than those of pristine polymers, regardless of the deceased  $R_{\text{face-on}}$ . Although the cause of the enhanced  $\mu_{\text{h}}$  in the blend films is unclear, this might merit the OPV performance which prefers a high  $\mu_{\text{h}}$  in the vertical direction.

## Solar cell performance

Fig. 3a shows current–voltage curves of normal cell OPV (PEDOT : PSS/active layer/Ca/Al). The active layers were prepared from chlorobenzene (CB) solution with a 3 vol% solvent additive of either 1,8-diiodooctane (DIO) or 1-chloronaphthalene (CN). We have examined the p/n blend ratio (1 : 1 ~ 1 : 5), film thickness, and thermal annealing effect (80, 120, and 150 °C). The optimal device parameters are summarized in Table 3, where the best blend ratio depends on the polymer (from 1 : 1 to 1 : 3) and no thermal annealing is preferred for all the polymers. The other low-performing device outputs are listed in Tables S1–S4.† Except for the low PCE of **TT1** (0.60%), **TT2**, **BT1**, and **BT2** indicated comparable PCEs from 3.3 to 3.7%. The best PCE of 3.74% (average 3.47%) was achieved for **BT2** : PC<sub>61</sub>BM = 1 : 2 with an open-circuit voltage ( $V_{\text{oc}}$ ) of 0.62 V, a  $J_{\text{sc}}$  of  $10.07 \text{ mA cm}^{-2}$ , and a fill factor (FF) of 0.66. The external quantum efficiency (EQE) spectra of **TT2**, **BT1**, and **BT2** shown in Fig. 3b present almost identical features corresponding to the superposition of polymer and PC<sub>61</sub>BM absorption spectra. The calculated  $J_{\text{sc}}$  ( $= J_{\text{sc}}^{\text{calc}}$ ) from the EQE spectra is in excellent agreement with the  $J_{\text{sc}}$  under 1 sun (Table 3), indicative of the appropriate evaluation. The lowest-performing **TT1** shows a suppressed EQE spectrum and low  $J_{\text{sc}}$ , arising from the lowest face-on orientation ( $R_{\text{face-on}} = 40\%$ ) and molecular weight.

The film morphologies observed by atomic force microscopy (AFM) are almost similar (Fig. 3c and d and S7†).  $V_{\text{oc}}$  is presumably proportional to the energy difference between the HOMO of the p-type polymer and the LUMO of n-type fullerene with an energy loss ranging from 0.8 to 1.3 eV.<sup>41</sup> The highest  $V_{\text{oc}}$  of 0.73 V was obtained for **BT1** (HOMO = −5.50 eV), while those of other polymers having the almost same HOMO of −5.35 to −5.43 eV were *ca.* 0.56–0.62 V, consistent with 0.1 V lower HOMOs.

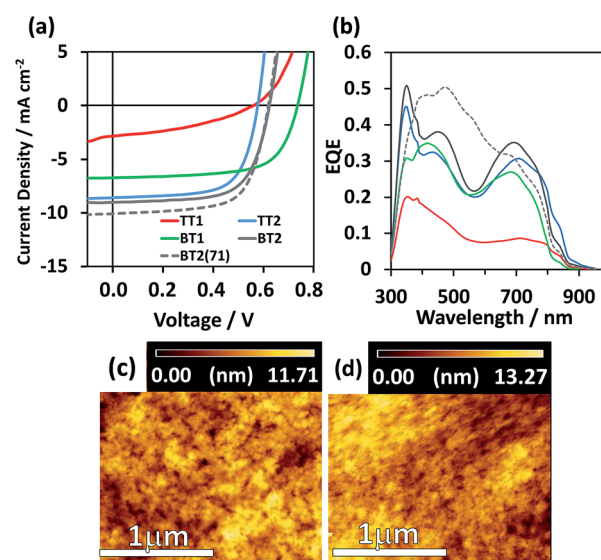


Fig. 3 (a)  $J$ – $V$  curves and (b) EQE spectra of the best-performing OPV devices. AFM topology images ( $2 \times 2 \mu\text{m}$ ) of (c) **BT2** : PC<sub>61</sub>BM = 1 : 2 and (d) **BT2** : PC<sub>71</sub>BM = 1 : 2. The scale bar at the bottom is 1  $\mu\text{m}$ .





Table 3 Summary of optimized OPV device performances of BTIDG polymers

Blend <sup>a</sup>	p/n ratio <sup>b</sup>	L/nm	PCE <sup>d</sup> (average)/%	V <sub>oc</sub> /V	J <sub>sc</sub> /mA cm <sup>-2</sup>	FF	J <sub>sc</sub> <sup>calc</sup> /mA cm <sup>-2</sup> (J <sub>sc</sub> /J <sub>sc</sub> <sup>calc</sup> ) <sup>e</sup>
TT1 : PC <sub>61</sub> BM	1 : 1	79	0.60 (0.48)	0.56	2.83	0.38	2.85 (0.99)
TT2 : PC <sub>61</sub> BM	1 : 2	150	3.40 (3.22)	0.58	8.54	0.69	7.90 (1.08)
BT1 : PC <sub>61</sub> BM	1 : 3	84	3.29 (3.21)	0.73	6.72	0.67	6.74 (1.00)
BT2 : PC <sub>61</sub> BM	1 : 2	89	3.74 (3.47)	0.62	9.01	0.66	8.68 (1.04)
BT2 : PC <sub>71</sub> BM	1 : 2	90	4.18 (3.92)	0.62	10.07	0.67	10.04 (1.00)

<sup>a</sup> Normal cell structure (PEDOT : PSS/active layer/Ca/Al). The active layers were fabricated by chlorobenzene solution with 3 v/v % DIO or CN. No thermal annealing was applied for the active layer. <sup>b</sup> In weight% fraction. <sup>c</sup> Thickness of the active layer. <sup>d</sup> The maximum PCE. The value in the brackets is an average PCE over at least 5 devices. <sup>e</sup> The calculated J<sub>sc</sub> by integrating the EQE spectrum. The value in the brackets is J<sub>sc</sub>/J<sub>sc</sub><sup>calc</sup>.

To improve the PCE of the best-performing BT2-based OPV cell, PC<sub>61</sub>BM was replaced by PC<sub>71</sub>BM which has higher absorption in the visible region than the former. As a result, BT2 : PC<sub>71</sub>BM boosted the PCE to 4.18% (average 3.92%), due to the increased J<sub>sc</sub> of 10.1 mA cm<sup>-2</sup> without sacrifices of V<sub>oc</sub> (0.62 V) and FF (0.67). The film morphologies of BT2 : PC<sub>61</sub>BM and BT2 : PC<sub>71</sub>BM were almost identical as visualized in the AFM images of Fig. 3c and d. We also examined a thicker film of BT2 : PC<sub>61</sub>BM, because high PCEs of 9–10% were achieved for thick active layers (*ca.* 300 nm) composed of crystalline polymers.<sup>13,42,43</sup> The PCE of 300 nm-thick BT2 : PC<sub>61</sub>BM was 3.51%, being equal to the 120 nm-thick film (3.50%). Despite the identical V<sub>oc</sub> (0.61 V), the J<sub>sc</sub> and FF of the 300 nm-thick film were increased to 13.38 mA cm<sup>-2</sup> and decreased to 0.43, respectively, indicative of the typical trade-off relationship.<sup>17</sup> In the manufacturing stage of OPV cells *via* high-throughput inkjet printing, a polymer without strong thickness dependence is highly favorable<sup>44</sup> in which BT2 could be relevant to this criterion.

## Conclusions

In conclusion, we designed and synthesized new LBPs composed of asymmetric BTIDG acceptor and weak donors, affording suitable E<sub>g</sub><sup>opt</sup> and HOMO levels. The 2D-GIXRD measurements revealed that the polymer orientation is molecularly controlled by the size of the donor  $\pi$ -plane and the number of alkylthiophene spacers, among which the top-class face-on orientation (*R*<sub>face-on</sub> = 95%) was demonstrated in pristine BT2. The best PCE of 4.2% was achieved in BT2 : PC<sub>71</sub>BM without significant thickness dependence. Since the polymer : PC<sub>71</sub>BM films still suffer from a large decrease in face-on orientation by 10–20%, BTIDG-based polymers underpin further evolution of efficient OPV devices in which our work provides a useful design guide on polymer face-on orientation.

## Experimental

### General

Steady-state photoabsorption spectroscopy was performed using a Jasco V-570 UV-vis spectrophotometer. Molecular weights of polymers were determined using the gel permeation chromatography (GPC) method with polystyrene standards.

GPC analysis was performed with chloroform as an eluent at a flow rate of 1 cm<sup>3</sup> min<sup>-1</sup> at 40 °C, on a SHIMADZU LC-20AT, CBM-20A, CTO-20A chromatography instrument connected to a SHIMADZU SPD-M20A UV-vis detector. PYS experiments were carried out by using a Sumitomo Heavy Industry Co. PCR-202. 2D-GIXRD experiments were conducted at the SPring-8 on the beam line BL19B2 or BL46XU using 12.39 keV ( $\lambda$  = 1 Å) X-ray. The 2D-GIXRD patterns were recorded with a 2-D image detector (Pilatus 300K). AFM observations were performed by using a Seiko Instruments Inc. model Nanocute OP and Nano-navi II.

### Materials

The polymers were synthesized according to our previous reports<sup>8</sup> and following procedures. PCBM and solvents were purchased from Frontier Carbon Inc. and Kishida Chemical Inc., respectively, and used as received. All chemicals were purchased from Aldrich, Kanto Chemicals, Tokyo Chemical Inc. (TCI), or WAKO Chemical Co., and were used as received. Column chromatography was carried out on Silica Gel 60N (spherical, neutral) or NH silica gel from Fuji Silysia Chemical Ltd. Air- and water-sensitive synthetic steps were performed in a dried nitrogen or argon atmosphere using a standard Schlenk technique. All monomers were carefully purified prior to use in polymerization.

### Synthesis

The synthetic procedure of BTIDG molecules and polymers is shown in Schemes 1 and 2, respectively. The synthesis of precursors (5) and (7) is provided in the ESI.†

**6-Bromo-1-(2-hexyldecyl)-3-[6-(2-hexyldecyl)-5-oxothieno[2,3-b]pyrrol-4-ylidene]indol-2-one(8).** Compound (5) (127 mg, 0.34 mmol), compound (7) (140 mg, 0.32 mmol), and acetic acid (2.6 mL) were added to a 30 mL two-neck flask at room temperature under the atmosphere of nitrogen. The reaction solution was stirred at 100 °C for 15 h. After the reaction, the reaction mixture was quenched with water and extracted with hexane, and the solvent was removed under vacuum. The residue was purified by flash chromatography on neutral silica gel, eluted with hexane/ethylacetate (10/1). The product was obtained as purple oil (113.6 mg,  $1.4 \times 10^{-4}$  mol, yield: 44%). <sup>1</sup>H NMR (400 MHz, CDCl<sub>3</sub>),  $\delta$  0.85 (m, 12H), 1.24–1.32 (m, 48H), 1.84 (m, 2H), 3.64 (d, *J* = 7.6 Hz, 2H), 3.69 (d, *J* = 7.2 Hz, 2H), 6.74 (d, *J* = 5.6 Hz,



1H), 6.97 (d,  $J = 2.0$  Hz, 1H), 7.23 (dd,  $J = 8.4, 2.0$  Hz, 1H), 7.60 (d,  $J = 5.6$  Hz, 1H), 8.4 (d,  $J = 8.4$  Hz, 1H). Anal. calcd for  $C_{46}H_{71}BrN_2O_2S$ : C, 69.41; H, 8.99; Br, 10.04; N, 3.52; O, 4.02; S, 4.03. Found; C, 68.97; H, 9.12; Br, 10.58; N, 3.50; S, 4.14%. MS (MALDI-TOF):  $m/z$  calcd for  $C_{46}H_{71}BrN_2O_2S$ : 795.04 [ $M^+$ ]; found: 796.16 [ $M^+$ ].

**6-Bromo-3-[2-bromo-6-(2-hexyldecyl)-5-oxothieno[2,3-*b*]pyrrol-4-ylidene]-1-(2-hexyldecyl)indol-2-one(9).** Compound (8) (0.46 g,  $5.8 \times 10^{-4}$  mol), *N*-bromosuccinimide (NBS) (0.10 g,  $5.9 \times 10^{-4}$  mol) and dry tetrahydrofuran (THF) (9.7 mL) were added to a dried 50 mL flask. After stirring 12 h at room temperature under the atmosphere of nitrogen, the reaction mixture was quenched with  $H_2O$  and extracted with hexane. The solvent was removed under vacuum, and the residue was purified by flash chromatography on neutral silica gel, eluted with hexane/ethylacetate (15/1). The product was obtained as purple oil (0.48 g,  $5.5 \times 10^{-4}$  mol, yield: 95%).  $^1H$  NMR (400 MHz,  $CDCl_3$ ),  $\delta$  0.86 (m, 12H), 1.24–1.29 (m, 48H), 1.84 (m, 2H), 3.57 (d,  $J = 7.2$  Hz, 2H), 3.64 (d,  $J = 7.6$  Hz, 2H), 6.74 (s, 1H), 6.91 (d,  $J = 2.0$  Hz, 1H), 7.17 (dd,  $J = 8.4, 2.0$  Hz, 1H), 9.07 (d,  $J = 8.4$  Hz, 1H). Anal. calcd for  $C_{46}H_{70}Br_2N_2O_2S$ : C, 63.15; H, 8.06; Br, 18.27; N, 3.20; O, 3.66; S, 3.66. Found; C, 63.29; H, 8.09; Br, 18.12; N, 3.18; S, 3.76%. MS (MALDI-TOF):  $m/z$  calcd for  $C_{46}H_{70}Br_2N_2O_2S$ : 873.93 [ $M^+$ ]; found: 874.07 [ $M^+$ ].

**1-(2-Hexyldecyl)-3-[6-(2-hexyldecyl)-5-oxo-2-(thiophen-3-yl)thieno[2,3-*b*]pyrrol-4-ylidene]-6-(thiophen-2-yl)indol-2-one(10).** Compound (9) (0.094 g,  $1.1 \times 10^{-4}$  mol), 2-trimethylstannyl-4-bromothiophene (18) (0.089 g,  $2.1 \times 10^{-4}$  mol), and xylene (1.5 mL),  $Pd(PPh_3)_4$  (6.2 mg,  $5.4 \times 10^{-6}$  mol) were added to a dried 20 mL Schlenk flask under a nitrogen atmosphere. After refluxing at 150 °C for 5 h, the reaction mixture was diluted by dichloromethane and filtered with celite. It was purified by flash chromatography on neutral silica gel, eluted with hexane/ethylacetate (20/1). The product was obtained as purple oil (0.113 g,  $9.3 \times 10^{-5}$  mol, yield: 86%).  $^1H$  NMR (400 MHz,  $CDCl_3$ ),  $\delta$  0.86 (m, 18H), 1.24–1.37 (m, 98H), 1.65 (m, 4H), 1.84 (m, 2H), 3.66 (d,  $J = 6.8$  Hz, 2H), 3.77 (d,  $J = 6.4$  Hz, 2H), 6.78 (s, 1H), 6.93 (d,  $J = 2.0$  Hz, 1H), 7.04 (d,  $J = 1.6$  Hz, 1H), 7.30 (d,  $J = 1.2$  Hz, 1H), 7.33 (d,  $J = 1.6$  Hz, 1H), 9.20 (d,  $J = 8.4$  Hz, 1H).

**6-(5-Bromothiophen-2-yl)-3-[2-(5-bromothiophen-3-yl)-6-(2-hexyldecyl)-5-oxothieno[2,3-*b*]pyrrol-4-ylidene]-1-(2-hexyldecyl)indol-2-one(11).** Compound (10) (0.12 g,  $9.3 \times 10^{-5}$  mol), NBS (0.035 g,  $2.0 \times 10^{-4}$  mol), and THF (1.5 mL) were added to a 30 mL two neck flask, and the reaction mixture was stirred for 4 h at room temperature. The blue solution was quenched with  $H_2O$ , and the solvent was removed after the extraction with ethylacetate. The residue was purified by flash chromatography with neutral silica gel, eluted with hexane/ethylacetate (50/1). The product was obtained as blue purple oil (0.12 g,  $8.7 \times 10^{-5}$  mol, yield: 91%).  $^1H$  NMR (400 MHz,  $CDCl_3$ ),  $\delta$  0.88 (m, 18H), 1.24 (m, 98H), 1.62 (m, 4H), 1.83 (m, 2H), 3.53 (d,  $J = 7.2$  Hz, 2H), 3.64 (d,  $J = 6.4$  Hz, 2H), 6.52 (s, 1H), 6.77 (s, 1H), 7.01 (s, 1H), 7.09 (s, 1H), 7.14 (dd,  $J = 8.0, 1.6$  Hz, 1H), 9.07 (d,  $J = 8.4$  Hz, 1H).

**General procedure of synthesis of the copolymers.** The synthetic route to copolymers (TT1, TT2, BT1, and BT2) is shown in Scheme 2. The specific condition of each copolymer is given in the ESI.† Polymerization steps for BTIDG polymers

were carried out through the palladium(0)-catalyzed Stille cross-coupling reactions. For the Stille cross-coupling reaction, 1 equiv. of dibromo monomers and 1 equiv. of bistrimethylstannyl monomers were added into anhydrous xylene in a 20 mL Schlenk flask. Tetrakis-(triphenylphosphine)palladium(0) (0.05 equiv.) was transferred into the mixture in a nitrogen atmosphere. The reaction mixture was stirred at 150 °C for 1–2 h, and then an excess amount of bromobenzene and trimethyltin benzene was added to end-cap the trimethylstannyl and bromo groups for 30 min, respectively. After adding a four-fold larger amount of xylene than the reaction solution, the reaction mixture was cooled to 40 °C and added slowly into a vigorously stirred ten-fold larger volume of methanol than the diluted reaction mixture. The polymers were collected by filtration and dried under vacuum. The dried crude polymers were dissolved in  $CHCl_3$  (200 mL) and further purified by column chromatography on neutral silica gel and six layered interleave NH silica with a  $NH_2$  silica gel layer to remove the metallic impurities. The obtained  $CHCl_3$  solution was filtered through celite and the solvent was removed under vacuum. The remaining solid was dissolved in a minimum amount of  $CHCl_3$  and reprecipitated by six-fold larger amount of acetone than dissolved solution to remove oligomers. The precipitate was filtered using a membrane filter and dried under vacuum.

### Organic photovoltaic cell (OPV)

The procedures of fabrication of OPV devices were carried out according to ref. 45. A PEDOT : PSS (Heraeus Clevis P VP AI 4083) layer was cast onto the cleaned ITO layer by spin-coating after passing through a 0.2  $\mu m$  filter. The substrate was annealed on a hot plate at 120 °C for 30 min. An active layer consisting of polymer and PCBM (purchased from Frontier Carbon Inc.) was cast on top of the PEDOT : PSS buffer layer in a nitrogen glove box by spin-coating after passing through a 0.2  $\mu m$  filter. The thickness was around 80–90 nm for TT1, BT1, and BT2, and 150 nm for TT2. No thermal annealing was applied. The solvent and solvent additive were completely removed in a vacuum chamber by keeping the device for at least 30 min. A cathode consisting of 20 nm Ca and 100 nm Al layers was sequentially deposited through a shadow mask on top of the active layers by thermal evaporation in a vacuum chamber. The resulting device configuration was ITO (120–160 nm)/PEDOT : PSS (45–60 nm)/active layer/Ca (20 nm)/Al (100 nm) with an active area of 7.1 mm<sup>2</sup>. Current-voltage curves were measured using a source meter unit (ADCMT Corp., 6241A) under AM 1.5 G solar illumination at 100 mW cm<sup>-2</sup> (1 sun, monitored by using a calibrated standard cell, Bunko Keiki SM-250KD) from a 300 W solar simulator (SAN-EI Corp., XES-301S). The EQE spectra were measured by using a Bunko Keiki model BS-520BK equipped with a Keithley model 2401 source meter. The monochromated light power was calibrated by using a silicon photovoltaic cell, Bunko Keiki model S1337-1010BQ.

### Space-charge limited current (SCLC)

ITO coated glass substrates were subsequently cleaned with detergent, distilled water, acetone and isopropanol for 10 min



each under ultrasonication. Aluminum was vapor deposited on the ITO surface with a patterned mask as the guide electrode. The substrates were further UV-ozone treated and a layer of PEDOT : PSS (Heraeus Clevis P VP AI 4083) was spin-coated after passing through a 0.2 mm filter. They were annealed on a hot plate at 120 °C for 10 min, and then inserted into a glove box filled with nitrogen. A solution of polymer in chlorobenzene was heated at 80 °C under stirring for at least 2 h to dissolve the polymer completely. A thin layer of BTIDG polymer or polymer : PCBM blend was spin-coated from the polymer solution or solution of the OPV best-performing blend ratio, respectively. A gold layer with 70–100 nm thickness was deposited by thermal evaporation in a vacuum chamber. The current–voltage ( $J$ – $V$ ) curves were measured using an ADCMT Corp. 6241A source meter. The hole mobility was determined by fitting the  $J$ – $V$  curve into the Mott–Gurney law,  $J = 9\epsilon_r\epsilon_0\mu(V-V_{bi})^2/(8L^3)$ , where  $\epsilon_0$  is the permittivity of free space,  $\epsilon_r$  is the relative dielectric constant of the material,  $\mu$  is the hole (or electron) mobility,  $V$  is the voltage,  $V_{bi}$  is the built-in voltage, and  $L$  is the layer thickness.<sup>46</sup>

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