



CrossMark
click for updates

Cite this: *Chem. Sci.*, 2015, 6, 5526

Received 8th April 2015
Accepted 24th June 2015

DOI: 10.1039/c5sc01266j

www.rsc.org/chemicalscience

Reaction of a diaryldigermynes with ethylene†

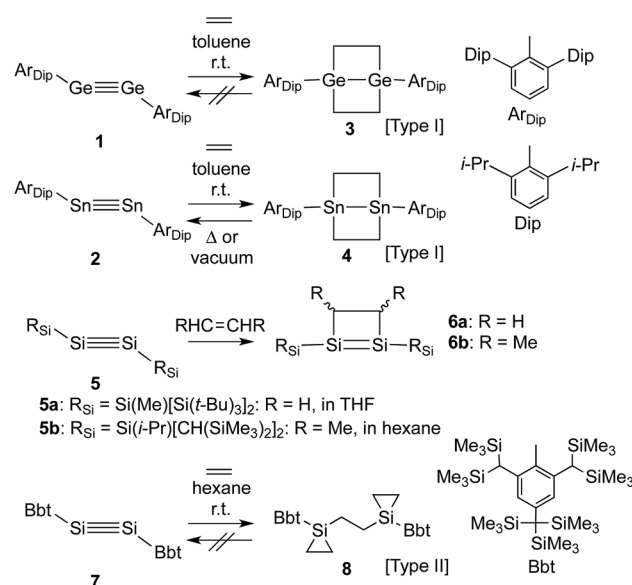
Takahiro Sasamori,^{*a} Tomohiro Sugahara,^a Tomohiro Agou,^a Koh Sugamata,^a Jing-Dong Guo,^{ab} Shigeru Nagase^b and Norihiro Tokitoh^{*a}

Reaction of the stable digermynes BbtGe≡GeBbt (Bbt = 2,6-[CH(SiMe₃)₂]₂-4-[C(SiMe₃)₃]-C₆H₂) with ethylene initially afforded the corresponding 1,2-digermacyclobutene. Depending on the reaction conditions applied, further reaction of this 1,2-digermacyclobutene with ethylene furnished two different reaction products: a 1,4-digermabicyclo[2.2.0]hexane or a bis(germiranyl)ethane. Combined experimental and theoretical results suggested that the 1,4-digermabicyclo[2.2.0]hexane and the bis(germiranyl)ethane are the thermodynamic and kinetic reaction products, respectively. A reaction mechanism in agreement with these results was proposed.

Control over the modification of olefin groups is important in organic synthesis, as a variety of preparative methods for the introduction of functional groups start from C–C multiple bonds. Even though several olefin addition reactions, such as hydrosilylation,¹ hydroboration² and olefin polymerisation,³ are well established, the use of transition metal catalysts is required in many cases. However, divalent or multiple-bonded compounds of heavier group 14 elements have recently received much attention as potential transition metal-free catalysts.⁴ These compounds generally react with olefins or other compounds that have carbon-containing multiple bonds to form the corresponding cycloadducts, tantamount to a strong propensity to activate small inert molecules. Unfortunately, low-coordinate species of heavier main group elements are usually difficult to isolate, mostly due to their inherently high reactivity towards addition reactions involving atmospheric moisture and/or aerobic oxygen and self-oligomerisation. Nevertheless, these compounds can be isolated while retaining their characteristic reactivity when sterically demanding substituents are used to provide kinetic stabilisation.⁵ Power and co-workers have, for example, reported the isolation of the heavier acetylene analogues Ar_{Dip}Ge≡GeAr_{Dip} (**1**)⁶ and Ar_{Dip}Sn≡SnAr_{Dip} (**2**)⁷ as stable compounds. The reactions of **1** and **2** with ethylene proceed smoothly in the absence of any transition metal catalyst at room temperature to afford the corresponding 4-membered cycloadducts **3** and **4** (Type I; Scheme 1).^{8,9} Subsequently, **4** is able to undergo a thermal retro-cycloaddition to generate **2**, concomitant with the release of two molecules of ethylene.

Accordingly, distannyne **2** can, in contrast to digermynes **1**, be considered as an ethylene-storage molecule. In this context, the reaction of a comparable disilyne with ethylene should also be of great interest. Independently, the groups of Wiberg and Sekiguchi have reported the stereoselective [2+2] cycloaddition of stable disilynes **5a,b** (R_{Si}Si≡SiR_{Si}; **5a**: R_{Si} = Si(Me)[Si(*t*-Bu)₃]₂,¹⁰ **5b**: R_{Si} = Si[CH(SiMe₃)₂](*i*-Pr)¹¹) with alkenes (RHC=CHR) to afford disilenes **6a,b** (**6a**: R = H, **6b**: R = Me).^{10,12} However, neither the further reaction of **6a,b** with ethylene, nor any possible retro-reaction were reported.

Previously, we have reported the synthesis of the stable diaryldisilyne BbtSi≡SiBbt (**7**, Bbt = 2,6-[CH(SiMe₃)₂]₂-4-[C(SiMe₃)₃]-C₆H₂).¹³ The reaction of **7** with ethylene resulted in the unexpected formation of **8** (Type II; Scheme 1), containing



Scheme 1 Reactions of dimetallynes **1**, **2**, **5** and **7** with ethylene.

^aInstitute for Chemical Research, Kyoto University, Gokasho Uji, Kyoto 611-0011, Japan. E-mail: sasamori@boc.kuicr.kyoto-u.ac.jp

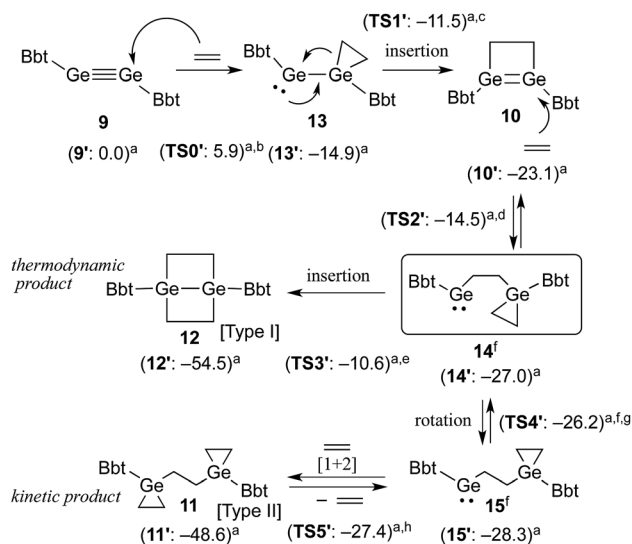
^bFukui Institute for Fundamental Chemistry, Kyoto University, Kyoto 606-8103, Japan

† Electronic supplementary information (ESI) available: experimental and computational details, as well as X-ray crystallographic data for **10–12** are available. CCDC 1054594–1054596. For ESI and crystallographic data in CIF or other electronic format see DOI: 10.1039/c5sc01266j



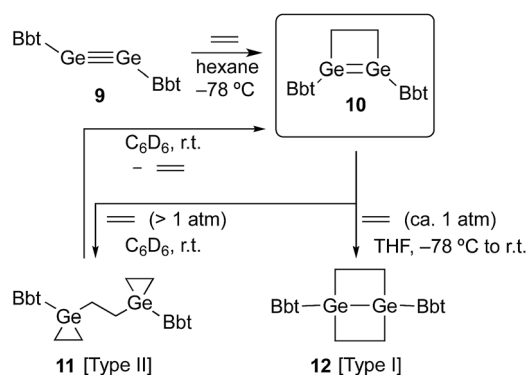
two silacyclopropane moieties. Compound **8** was found to be remarkably stable, as decomposition of these silacyclopropane moieties was not observed, even upon heating.¹⁴ Subsequently, we began to investigate the reactivity difference between diaryldisilynes and diaryldigermynes. Herein, we report the reaction of the stable diaryldigermine $\text{BbtGe}\equiv\text{GeBbt}$ (**9**)¹⁵ with ethylene to afford the corresponding 1,2-digermacyclobutene (**10**), which is the formal [2+2] cycloadduct of **9**. Depending on the reaction conditions, further treatment of **10** with ethylene resulted in the formation of two products, specifically a four-membered cycloadduct (**12**, Type I; Scheme 2) and a three-membered cycloadduct (**11**, Type II; Scheme 2).

A hexane solution of digermine **9** was frozen ($-196\text{ }^\circ\text{C}$) and degassed in a J-Young tube, before being charged with ethylene.¹⁶ The colour of the solution changed from dark red to purple. Removal of the solvent from the reaction mixture afforded 1,2-digermacyclobutene **10**. The formation of **10** from the reaction of digermine **9** with ethylene can be explained by the same mechanism used to describe the reaction of disilynes with olefins:¹² initially, interaction between ethylene and one of the Ge atoms in the $\text{Ge}\equiv\text{Ge}$ bond generates germirane-substituted germylene **13** as an intermediate,¹⁷ which subsequently inserts intramolecularly into the Ge–C bond of the germirane moiety (Scheme 3). X-ray crystallographic analysis of **10** revealed a non-planar structure for the four-membered $\text{Ge}=\text{Ge}-\text{C}-\text{C}$ ring (Fig. 1).¹⁸ The two Bbt groups were found to be oriented in opposite directions, resulting in a *trans*-bent geometry for the $\text{Ge}=\text{Ge}$ moiety with *trans*-bent angles of 39.5° (Ge1) and 39.7° (Ge2). A $\text{Ge}=\text{Ge}$ bond length of 2.4132(5) Å was observed, which is slightly shorter than a typical $\text{Ge}-\text{Ge}$ single bond (*ca.* 2.44 Å),¹⁹ but consistent with previously reported $\text{Ge}=\text{Ge}$ double bonds in digermenes (*ca.* 2.2–2.5 Å).¹⁹ These structural features suggested that the $\text{Ge}=\text{Ge}$ double bond in **10** should be weakened by the severe intrinsic strain of the four-membered $\text{Ge}=\text{Ge}-\text{C}-\text{C}$ ring and the highly *trans*-bent geometry. The ^1H NMR spectrum of **10** exhibited signals commensurate with two identical Bbt groups, as well as signals consistent with two equivalent SiMe_3 groups at the *ortho*-positions of the Bbt groups, thus confirming a fast inversion of the *trans*-bent geometry of the $\text{Ge}=\text{Ge}$ bond in **10** in solution.



Scheme 3 Proposed mechanism for the reaction of **9** with ethylene. (a) Calculated relative energies (kcal mol⁻¹) for model compounds bearing 2,6-[CH(SiMe₃)₂]₂-C₆H₃ (Bbp) groups instead of Bbt groups. (b) ΔE^\ddagger (**9**–**13'**) = 5.9 kcal mol⁻¹. (c) ΔE^\ddagger (**13'**–**10'**) = 3.4 kcal mol⁻¹. (d) ΔE^\ddagger (**10'**–**14'**) = 8.6 kcal mol⁻¹. (e) ΔE^\ddagger (**14'**–**12'**) = 16.4 kcal mol⁻¹. (f) **14** and **15** are rotational isomers with respect to a rotation around the central $\text{GeH}_2\text{C}-\text{CH}_2\text{Ge}$ bond. (g) ΔE^\ddagger (**14'**–**15'**) = 0.8 kcal mol⁻¹. (h) ΔE^\ddagger (**15'**–**11'**) = 0.9 kcal mol⁻¹.

In order to induce a further reaction of **10** with a second molecule of the alkene, ethylene was condensed into a sealed vessel, which contained a frozen and degassed C_6D_6 solution of **10** at $-196\text{ }^\circ\text{C}$. Subsequently, the reaction mixture was allowed to warm to r.t. in this sealed tube, and based on the volume of the tube, **10** was treated with an excess of ethylene (*ca.* 5 atm). The purple colour of **10** disappeared immediately,²⁰ and **11** (Type II; Scheme 2) was obtained as a colourless precipitate.²¹ Upon opening the sealed tube in an argon-filled glove box, the colourless powder turned purple again, and on the basis of its ^1H NMR spectrum it could be established that **11** retroconverted quantitatively to afford **10** within a few minutes at r.t. Accordingly, the reaction of **10** with ethylene to furnish **11** is, depending on the ethylene pressure, reversible. On the other hand, exposure of a degassed THF solution of **10**



Scheme 2 Reaction of diaryldigermine **9** with ethylene.

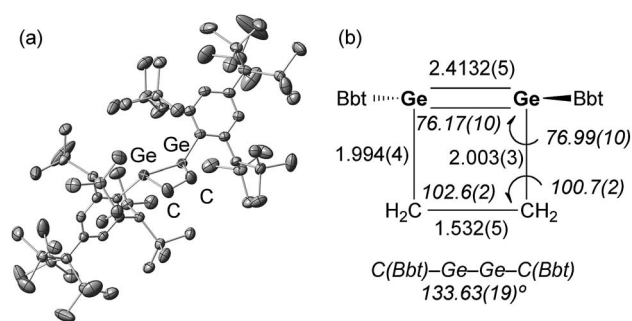


Fig. 1 (a) Molecular structure of **10** (thermal ellipsoids at 50% probability; hydrogen atoms omitted for clarity), and (b) selected metric parameters for the digermacyclobutene core in **10**.



to ethylene at ambient pressure (*ca.* 1 atm; $-78\text{ }^{\circ}\text{C}$ to r.t.; 1 d) afforded stable colourless crystals of **12** (Type I; Scheme 2) in quantitative yield. Depending on the reaction conditions, the reaction of **10** with ethylene thus delivers different reaction products. The molecular structures of **11** and **12** were determined unambiguously by spectroscopic and X-ray crystallographic analyses.²²

In order to elucidate the underlying reaction mechanism (Scheme 3), the reaction between digermynes **9** and ethylene was monitored by ^1H NMR spectroscopy in THF-*d*₈. After exposing a degassed THF-*d*₈ solution of **9** to ethylene (*ca.* 1 atm) at $-78\text{ }^{\circ}\text{C}$ and then allowing it to warm to r.t., the colour of **9** disappeared and only signals associated with **10** were observed. After 10 min, the intensity of these signals decreased, and additional signals consistent with the formation of **11** were observed (**10** : **11** = *ca.* 1 : 1). After 20 min, signals in agreement with the formation of **12** appeared, and after 5 hours, the quantitative formation of **12** was observed. These experimental results suggested that the reaction of **10** with ethylene furnishes **11** and **12** as the kinetic and thermodynamic products, respectively.

Taking all the previously discussed results into consideration, the reaction mechanism for the reaction between digermynes **9** and ethylene can most likely be interpreted as follows: the reaction is initiated by a nucleophilic attack of ethylene towards the LUMO of **9** to afford **13**, which readily undergoes an intramolecular ring-expansion, affording **10** *via* a gerymylene-digermene rearrangement.²³ Subsequently, nucleophilic attack of another molecule of ethylene towards the LUMO of **10** affords gerymylene **14**, which is expected to easily undergo a [1+2] cycloaddition reaction between a further molecule of ethylene and the second gerymylene moiety. While this [1+2] cycloaddition reaction should be reversible,²⁴ considering the results of the NMR monitoring reactions, the intramolecular C–Ge insertion of **14** is expected to proceed irreversibly to provide the thermodynamically stable product 1,4-digerma-bicyclo[2.2.0]hexane (**12**). The solubility of **11** in benzene was found to be limited, and the precipitation of **11** in the form of a colourless solid was observed when the reaction was conducted in this solvent. When the same reaction was carried out in THF, the kinetic product **11** was generated at an early stage in the reaction, and subsequently both **10** and **11** were converted to the thermodynamic product **12**. It can thus be concluded that the reactions of such digermynes are mostly initiated by nucleophilic attack of π -electrons towards the in-plane π^* orbital (LUMO) of the Ge–Ge triple bond, which is consistent with the previously reported reactivity of π -bond compounds containing heavier group 14 elements.

The proposed reaction pathways were also examined by density functional theory (DFT) calculations (see Fig. S11†), using appropriate model compounds (**9'**–**15'**) bearing Bbp (Bbp = 2,6-[CH(SiMe₃)₂]₂-C₆H₃) instead of Bbt groups (Scheme 3).²⁵ The results suggested that intermediate **13'** is formed with a small reaction barrier of 5.9 kcal mol⁻¹, and is thermodynamically more stable than **9'** + ethylene by 14.9 kcal mol⁻¹. Subsequently, **13'** can afford **10'** (8.2 kcal mol⁻¹ more stable) with a small reaction barrier of 3.4 kcal mol⁻¹. Following that, the reaction of **10'** with ethylene can provide key intermediate

14' (3.9 kcal mol⁻¹ more stable) with a reaction barrier of 8.6 kcal mol⁻¹. The second molecule of ethylene can then react smoothly with **14'** to give product **11'** *via* intermediate **15'**, which is a rotational isomer with a very low reaction barrier (<1.0 kcal mol⁻¹), while product **12'** is produced with a large barrier of 16.4 kcal mol⁻¹, which is 5.9 kcal mol⁻¹ more stable than product **11'**. The results of these DFT calculations corroborated the hypothesis that the reaction of **10** with ethylene should furnish **11** and **12** as the kinetic and thermodynamic products, respectively.²⁶

Finally, the reactivity difference between the reaction of ethylene with digermynes (**1** and **9**) and that with disilyne (**7**) can be explained as follows: for the reaction with **1** (Ar_{Dip}-Ge≡GeAr_{Dip}), the calculations draw the conclusion that the corresponding Type II product with three-membered rings should be the kinetic product, while the Type I product **3**, *i.e.* 1,4-digerma-bicyclo[2.2.0]hexane, should be the thermodynamic product, indicating that the observation of the kinetic product under these reaction conditions is unlikely.²⁷ These conclusions are in agreement with our experimental observations. For the reaction of **7** with ethylene, theoretical calculations indicated that Type II product **8** should be both the kinetically and the thermodynamically favoured product.²⁸ These results could be interpreted in terms of the relative stability of the Ge- or Si-containing three-membered rings.

Conclusions

In summary, we found that the reaction of digermynes **9** with ethylene affords two different reaction products (**11**, **12**), depending on the reaction conditions applied. The stable digermacyclobutene **10**, which is an intermediate in this reaction, could be isolated and subsequently treated under controlled reaction conditions with a second molecule of ethylene. A combined theoretical and experimental investigation of these reactions allowed the assignment of **11** and **12** as the kinetic and thermodynamic reaction products, respectively.

Acknowledgements

This work was partially supported by the following grants: Grants-in-Aid for Scientific Research (B) (no. 2235001722350017), Scientific Research on Innovative Areas, “New Polymeric Materials Based on Element-Blocks” [#2401] (no. 2510251925102519), “Stimuli-Responsive Chemical Species for the Creation of Functional Molecules” [#2408] (no. 2410901324109013), and the MEXT Project of Integrated Research on Chemical Synthesis from the Ministry of Education, Culture, Sports, Science and Technology (MEXT), Japan, as well as the “Molecular Systems Research” project of the RIKEN Advanced Science Institute and the Collaborative Research Program of Institute for Chemical Research, Kyoto University. We would like to thank Mr Toshiaki Noda and Ms Hideko Natsume at Nagoya University for the expert manufacturing of custom-tailored glassware. The manuscript for this publication was written at the Rheinische Friedrich-Wilhelms-Universität



Bonn during the tenure of an Alexander von Humboldt Research Award (NT) and a Friedrich Wilhelm Bessel Research Award (TS). NT and TS would like to express their gratitude to Prof. Rainer Streubel and his co-workers for their kind hospitality, as well as to the Alexander von Humboldt Stiftung for their generosity.

Notes and references

- 1 D. Troegel and J. Stohrer, *Coord. Chem. Rev.*, 2011, **255**, 1440.
- 2 H. C. Brown and B. Singaram, *Acc. Chem. Res.*, 1988, **21**, 287.
- 3 V. Busico and R. Cipullo, *Prog. Polym. Sci.*, 2001, **26**, 443.
- 4 (a) P. P. Power, *Nature*, 2010, **463**, 171; (b) P. P. Power, *Acc. Chem. Res.*, 2011, **44**, 627; (c) C. M. Cui, M. M. Olmstead, J. C. Fettinger, G. H. Spikes and P. P. Power, *J. Am. Chem. Soc.*, 2005, **127**, 17530.
- 5 (a) R. C. Fischer and P. P. Power, *Chem. Rev.*, 2010, **110**, 3877; (b) T. Sasamori and N. Tokitoh, *Bull. Chem. Soc. Jpn.*, 2013, **86**, 1005; (c) Y. Mizuhata, T. Sasamori and N. Tokitoh, *Chem. Rev.*, 2009, **109**, 3479; (d) S. Nagase, *Bull. Chem. Soc. Jpn.*, 2014, **87**, 167; (e) M. Asay and A. Sekiguchi, *Bull. Chem. Soc. Jpn.*, 2012, **85**, 1245.
- 6 M. Stender, A. D. Phillips, R. J. Wright and P. P. Power, *Angew. Chem., Int. Ed.*, 2002, **41**, 1785.
- 7 A. D. Phillips, R. J. Wright, M. M. Olmstead and P. P. Power, *J. Am. Chem. Soc.*, 2002, **124**, 5930.
- 8 Y. Peng, B. D. Ellis, X. P. Wang, J. C. Fettinger and P. P. Power, *Science*, 2009, **325**, 1668.
- 9 The reduction of Tsi(Cl)Ge: using Mg/MgBr₂ in the presence of ethylene in THF was reported to afford the corresponding 1,4-digerma-bicyclo[2.2.0]hexane, which is the Type I ethylene-adduct of a transient TsiGe≡GeTsi species; see: T. Ohtaki and W. Ando, *Organometallics*, 1996, **15**, 3103.
- 10 N. Wiberg, S. K. Vasisht, G. Fischer and P. Mayer, *Z. Anorg. Allg. Chem.*, 2004, **630**, 1823.
- 11 A. Sekiguchi, R. Kinjo and M. Ichinohe, *Science*, 2004, **305**, 1755.
- 12 R. Kinjo, M. Ichinohe, A. Sekiguchi, N. Takagi, M. Sumimoto and S. Nagase, *J. Am. Chem. Soc.*, 2007, **129**, 7766.
- 13 (a) T. Sasamori, K. Hironaka, Y. Sugiyama, N. Takagi, S. Nagase, Y. Hosoi, Y. Furukawa and N. Tokitoh, *J. Am. Chem. Soc.*, 2008, **130**, 13856; (b) T. Sasamori, J. S. Han, K. Hironaka, N. Takagi, S. Nagase and N. Tokitoh, *Pure Appl. Chem.*, 2010, **82**, 603.
- 14 J. S. Han, T. Sasamori, Y. Mizuhata and N. Tokitoh, *J. Am. Chem. Soc.*, 2010, **132**, 2546.
- 15 Y. Sugiyama, T. Sasamori, Y. Hosoi, Y. Furukawa, N. Takagi, S. Nagase and N. Tokitoh, *J. Am. Chem. Soc.*, 2006, **128**, 1023.
- 16 Experimental details and chemical data for previously unreported compounds are given in the ESI.†
- 17 Given the structure and NBO charge (C₂H₄ part, +0.045) of transition state **TS0'**, the initial interaction between ethylene and digermine **9** can be feasibly explained by an interaction between the HOMO of ethylene and the in-plane π* orbital (LUMO) of **9**. Conversely, the geometry of intermediate **13'** should, considering its geometry and NBO charge (C₂H₄ part, -0.586), be interpreted in terms of an orbital interaction between the π* orbital of ethylene and the out-of-plane π orbital of **9**. For further details, see ref. 15 and the ESI.†
- 18 Torsion angle: 11.06(18)°.
- 19 For example, the Ge–Ge and Ge=Ge bond lengths in Ar_{Dip}H₂Ge–GeH₂Ar_{Dip} and Ar_{Dip}HGe=GeHAr_{Dip} are 2.4019(10) Å and 2.3026(3) Å, respectively; see: G. H. Spikes, J. C. Fettinger and P. P. Power, *J. Am. Chem. Soc.*, 2005, **127**, 12232.
- 20 **10**: λ_{max} = 494 nm (ε 10 000).
- 21 Diaryldigermine Ar_{Tip}Ge≡GeAr_{Tip} [Ar_{Tip} = 2,6-bis(2,4,6-triisopropylphenyl)phenyl] was reported to undergo a similar reaction with three molecules of 2,3-dimethyl-1,3-butadiene; see: M. Stender, A. D. Phillips and P. P. Power, *Chem. Commun.*, 2002, 1312.
- 22 Crystallographic data and selected structural parameters of **11** and **12** are given in the ESI.†
- 23 Interconversion between germylgermylenes and digermenes via 1,2-substituent-migration has been reported previously; see: (a) K. M. Baines, J. A. Cooke and J. J. Vittal, *J. Chem. Soc., Chem. Commun.*, 1992, 1484; (b) K. M. Baines, J. A. Cooke, C. E. Dixon, H. W. Liu and M. R. Netherton, *Organometallics*, 1994, **13**, 631.
- 24 The mechanism of formal [1+2] and [1+4] cycloaddition reactions of germylenes has been well studied, see: L. A. Huch and W. J. Leigh, *Organometallics*, 2009, **28**, 6777.
- 25 Level of theory: B3PW91/6-311+G(2d)[Ge,Si] and 6-31G(d,p)[C,H]/B3PW91/Lanl2dz+d[Ge], 3-21G*[Si], and 3-21G[C,H]. M. J. Frisch, *et al.*, *Gaussian 09, revision D.01*, Gaussian, Inc., Wallingford, CT, 2013.
- 26 In general, [2+2] cycloaddition reactions between a C=C π-bond and a Ge=Ge π-bond (with out-of-plane π/π* orbital) are symmetry forbidden according to the Woodward–Hoffmann rules. Thus, the concerted [2+2] cycloaddition pathway from **9** to **10** as well as from **10** to **12** can be excluded. In fact, a reaction barrier of ΔE[‡] = 25.4 kcal mol⁻¹ was calculated for the concerted [2+2] cycloaddition from **10'** to **12'**, which is substantially larger than the barrier in the pathway shown in Scheme 3 (ΔE[‡] = 16.4 kcal mol⁻¹). Another possibility of a [2+2] cycloaddition via a biradical pathway was suggested in a previous paper (K. L. Hurni and K. M. Baines, *Chem. Commun.*, 2011, **47**, 8382), but could be excluded on the basis of the results of attempted calculations on these reaction pathways, which showed substantially larger barriers.
- 27 The Type I product **3**, *i.e.* 1,4-digerma-bicyclo[2.2.0]hexane, was calculated to be 1.3 kcal mol⁻¹ more stable than the Type II product, *i.e.* the bis(germiranyl)ethane. A reaction barrier of *ca.* 19 kcal mol⁻¹ was calculated for the retro-reaction from the Type II product to the Ar_{Dip} analogue of intermediate **14'**, indicating a reversible reaction at r.t. Conversely, a barrier of *ca.* 40 kcal mol⁻¹ was obtained for the conversion of Type I product **3** to the intermediate, suggesting an irreversible reaction. Accordingly, **3** should be the final product under thermodynamically controlled conditions. In our case, the low solubility in non-polar



solvents such as benzene or hexane most probably enabled us to isolate the kinetic product **11**.

28 In the case of Si, the Type II product, *i.e.* **8**, was calculated to be 23.1 kcal mol⁻¹ more stable than the Type I product, *i.e.*

the 1,4-disila-bicyclo[2.2.0]hexane analogue. Levels of theory: 6-311+G(2df)[Si] and 6-31G(d,p)[C,H]//B3PW91/3-21G*.

