The boron–boron triple bond? A thermodynamic and force field based interpretation of the N-heterocyclic carbene (NHC) stabilization procedure†

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Recently, the NHC→B=B=NHC molecule I has been published in Science where it is described as a stabilized B₂ molecule in its \( ^1\Sigma \) excited state (B₂*) (NHC = N-heterocyclic carbene C₃N₂H₂(C₆H₃Pr'₂-2,6)₂). The bonding of I based on sophisticated calculations and the BB distances of the solid compound was discussed as the first example of a B₂ triple bond in a stable molecule. Now we present an only experimentally based interpretation of I via detailed thermodynamic considerations, including its fragmentation to B₂ molecules. Furthermore, from the vibrational spectrum force constants \( f_{BB} \) for the BB bond and \( f_{BC} \) for the BC bond) were extracted, which are classical examples to indicate single, double and triple CC bonds in organic chemistry. The consequence of both properties of I \((\Delta E \text{ and } \varphi)\) generates a new interpretation which is in contrast to the triple bond donor–acceptor description visualized by arrows and which casts a critical light on the interpretation of any NHC “stabilized” molecule.

Introduction

During the last two decades an unprecedented renaissance of main group chemistry has been developed. High-lighted areas are e.g.: nanoscaled species such as metalloid clusters as intermediates between normal valence compounds and bulk metals on one hand, and, on the other hand, reactive molecules containing multiple bonding stabilized either via bulky ligands or via N-heterocyclic carbenes (NHCs). As an outstanding example in the latter field we will concentrate here on a compound containing an unusual B–B bond.

Though there have already been presented some nice molecules before, which contain electron precise boron–boron bonds as well as BB bonds containing additional \( \pi \)-bonding, two years ago the outstanding impressive molecule NHC→B=B=NHC (I) (NHC = N-heterocyclic carbene C₃N₂H₂(C₆H₃Pr'₂-2,6)₂) (Fig. 1) was published in Science under the title “Ambient-Temperature Isolation of a Compound with a Boron–Boron Triple Bond”.

I has been characterized via its X-ray solid state structure exhibiting a short B–B distance which is, however, longer than expected, via different NMR spectroscopic investigations, via three intensive IR bands, and an elemental analysis. Furthermore, detailed quantum chemical investigations, including the calculated UV/VIS spectrum have been presented. These calculations and the visualization of the orbitals involved in the multiple bonding were the basis for the interpretation summarized in the above-mentioned title and for a highlight article in the same issue of Science. Since this triple bond

![Fig. 1](image-url)  
Fig. 1. The structure of I in the crystalline state and its interpretation as a donor–acceptor molecule symbolized via two arrows along the C₁B bond. The two methyl groups of each of the eight grey C-atoms are omitted for clarity. The following distances (\( \text{Å} \)) are essential for the discussion: \( d(BB) = 1.45; d(BC) = 1.49; d(C₂N) = 1.39. \)
interpretation has also been predicted theoretically\textsuperscript{20} and was also included in recent reviews,\textsuperscript{19,21} we feel that it is time to make a cut and to give a different, only experimentally based interpretation for this nice compound 1 in order to increase our understanding of its unusual bonding and in order to prevent already beginning confusion, especially in the textbooks.\textsuperscript{22} Furthermore we will show that any mainly “orbital based” interpretation in which an NHC stabilization is involved should make us cautious. Compound 1 was chosen for this discussion as an impressive example for the NHC stabilization, since many other molecules containing B–B bonding exhibit suitable reference data because the variation of the B–B vibration indicates the degree of “stabilization”. Thus, we want to discuss 1 with respect to its thermodynamic property (Section 1) and its molecular vibrations (Section 2) for which the bond strengths (force constants) are an essential basis; i.e. we want to discuss two properties which are basic for every bonding discussion in the entire field of chemistry.\textsuperscript{23–25}

**Thermodynamic view**

Today the thermodynamic stability of every species, even of reactive ones, is available via quantum chemical calculations. Sometimes, additional experimental data make this discussion more confident. Unfortunately, this discussion is missing for 1\textsuperscript{26,28} and its interpretation as a donor–acceptor stabilized molecule visualized by two arrows (Fig. 1). In Fig. 2 the calculated and the experimentally obtained thermodynamic data of molecule 1 and of solid boron together with the hot gaseous molecules/atoms B\textsubscript{1} [ground state 2\textPi],\textsuperscript{27,28} B\textsubscript{2} [ground state 3\Sigma\textsubscript{g} \textsuperscript{-}] and B\textsubscript{2} [excited state 1\Sigma\textsubscript{g} \textsuperscript{+}] are summarized.

Therefore, fundamental thermodynamic considerations could be the starting point for every bonding discussion before any other investigation will go into detail. The most remarkable conclusion from Fig. 2 is that, with respect to the gaseous species B atoms, B\textsubscript{2} (\Sigma\textsubscript{g}) and B\textsubscript{2} (\Sigma\textsubscript{g} \textsuperscript{-}), solid boron and the gaseous compound 1 are nearly on the same thermodynamic level: e.g. two boron atoms are 11.6 eV (1120 kJ mol\textsuperscript{-1}) above solid boron and 11.0 eV above 1.\textsuperscript{29,30} However, since the boiling point of boron at 3900 °C exhibits its robustness based on strong covalent B–B bonds, it is hard to believe that the gaseous B\textsubscript{1}, B\textsubscript{2}, B\textsubscript{2} species can be obtained from 1 via heating, i.e. energy transfer from outside.\textsuperscript{24} Nevertheless, at least theoretically in a Gedanken experiment the dissociation of 1 to B\textsubscript{2} species and 2NHCs can be allowed. However, the observation of a similar energy transfer starting from boron and from 1 to get gaseous B\textsubscript{1}, B\textsubscript{2} and B\textsubscript{2}\textsuperscript{+} should be alarming!

Now we will discuss the reverse Gedanken experiment, i.e. we look at the gaseous species B\textsubscript{1}, B\textsubscript{2}, and B\textsubscript{2}\textsuperscript{+}, and allow B\textsubscript{2}\textsuperscript{+} (a) to condense to solid boron or (b) to react with 2NHC molecules to 1 in the gas phase.

At about 2000 K the following gaseous boron species are in equilibrium with solid boron: i.e. the relative concentration of B\textsubscript{1}, B\textsubscript{2} and B\textsubscript{2}\textsuperscript{+} is 10\textsuperscript{18} : 10\textsuperscript{12} : 1 (see ESI). Now we concentrate on the excited B\textsubscript{2}\textsuperscript{+} molecule\textsuperscript{22} though its relative concentration in the gas phase is extremely low. Nevertheless this B\textsubscript{2}\textsuperscript{+} molecule was the basis for the theoretical discussion of 1 and its triple bond character.

(a) When B\textsubscript{2}\textsuperscript{+} is condensed to form solid boron, 13.2 eV (1275 kJ mol\textsuperscript{-1}) are gained, because the multiple bond of B\textsubscript{2}\textsuperscript{+} is changed to solid boron containing only single bonds like in a polymerization process. Thus, there is a dramatic exothermic rearrangement of atoms and electrons, and one of the strong covalent bonded allotropes of boron with high thermal robustness is formed. The large energy gain which is connected with this process will cause a strong heating of boron, which is no problem for this material, and which finally heats the environment. The conclusion of this Gedanken experiment is, that solid boron and its structure has nothing to do with the educts (B\textsubscript{2}, B\textsubscript{2}\textsuperscript{+} and 2NHC molecules), but only with this process which will cause a strong heating of boron, which is no problem for this material, and which finally heats the environment.

(b) Now we look at the reaction of B\textsubscript{2}\textsuperscript{+} with 2NHC molecules in the gas phase: this formation of 1 is strongly exothermic with 12.6 eV (1221 kJ mol\textsuperscript{-1}) because the electrons are rearranged, new bonds are formed, and the original bonds are changed. The large energy gain of this gas-phase reaction should result in the heating of 1 and its fragmentation; i.e. this fragmentation process would start by breaking the weakest bonds: however, in principle, this fragmentation can be avoided\textsuperscript{15–28} – at least theoretically – if the heat can be transferred fast enough to the environment, e.g. via radiation. Anyway, if 1 really would survive in this exothermic reaction starting from B\textsubscript{2}\textsuperscript{+} and 2NHC, one has to conclude: the final state of this reaction (i.e. 1) is energetically far below the starting point (B\textsubscript{2}\textsuperscript{+} + 2NHC) (12.6 eV), i.e. electron distribution and bonding of the educts and products must be extremely different. In one word, the final state 1 has nearly nothing to do with the educts (B\textsubscript{2}\textsuperscript{+} and 2NHC molecules), like in the case of solid boron! Consequently, the bonding situation in 1 can hardly be symbolized as a slight modification of the educts via arrows (Fig. 1), which suggest only a weak donor (NHC)–acceptor (B\textsubscript{2}\textsuperscript{+}) interaction in which the bonding of the educts is still visible.

![Energy diagram for solid boron, B-atoms, B\textsubscript{2} molecules (B\textsubscript{2}, B\textsubscript{2}\textsuperscript{+}) and the decomposition of 1 to B\textsubscript{2} and 2NHC.](https://example.com/energy-diagram.png)
However, to look more carefully to the energetic situation of the final electron distribution within the “C−B−B−C” core, especially with respect to the BC bonding, one has to start with the ground state of B₂ (8 eV (772 kJ mol⁻¹) above 1); with its BB single bond and without BC contacts to 2NHCs. If the BB bond is not changed during the reaction with 2NHCs, the formation of two BC single bonds will consume 772 kJ mol⁻¹, i.e. each BC bond consumes 386 kJ mol⁻¹, nearly the expected value of the BC single bond energy 372 kJ mol⁻¹. However, this intermediate electronic situation of 1 with one BB and two BC single bonds has to drop into a more stable energetic valley, if the energy of the B₂ molecule is increased after its formation from B₂⁺ as starting point. Subsequently, this intermediate possessing only single bonds has to consume 445 kJ mol⁻¹ (4.61 eV), which is nearly quantitatively possible by the formation of two BC double bonds, which are each estimated to be at 230 kJ mol⁻¹, i.e. slightly more stable than the BC single bond. Thus the following situation for the central C−B−B−C core results, which is in line with the most prominent neutral resonance structure presented in Scheme 1. Finally, after distribution of the π electrons from the BC bonds to the whole CBBC core, this thermodynamic Gedanken experiment results in a 4-electron−4-center π bond, NHC−B−B−NHC (Scheme 2).

In order to confirm the conclusion of a rearrangement of electrons in 1 which is completely different from that of the educts, we look at the forces between the atoms of 1 which are visible by its vibrational spectrum i.e. by its IR and Raman spectrum and compare this situation with that of the B₂⁺ molecule.

### Determination and discussion of the force constants of 1

For the bonding discussion of 1, the most convincing structured property concerning the BB bond was its short distance.³¹,³⁹,⁴¹−⁴³ More much more reliable and sensitive for the experimental characterization of a bond and especially of multiple bonds are the force constants which reflect directly the situation of the bonding electrons between two nuclei. More accurately, the force constant \( f \) (spring constant: force = \( f \Delta r \)) represents the restoring force which resists to a small elongation (\( \Delta r \)) of the atoms from the equilibrium distance.⁴⁴ The relevance of force constants for the discussion of bond properties has recently been shown for the Zn–Zn bond in Zn₂CP₂⁴,⁴⁵ has been highlighted for the discussion of S–S multiple bonds⁴⁶ and, as a classical example, is convincingly demonstrated in fundamental organic chemistry, where the relation of the force constants of a CC single bond to a double and a triple bond is about 1 : 2 : 3.⁴⁷,⁴⁸

Therefore, a discussion of force constants of bonds between carbon and its direct neighbor element boron should also be a convincing measure to discuss BB and BC bonds, especially whether or not multiple bonding is involved.

Unfortunately, force constants cannot directly be obtained from the vibrational spectra, because the observed spectra are the result of interactions of the hypothetical isolated motions within every bond or special entities of a molecule. However, the so-called normal-coordinate analysis⁴² allows extracting the force constants for every bond, if the molecule and the number of vibrations are not too large. Therefore, instead of the determination of the force constant of 1 (there are 390 vibrations) we have chosen the model compound 1a in which the NHC of 1 is substituted by the most simple NHC containing only H ligands: C(NH)₂(CH)₂, and only 54 vibrations are obtained. Since the vibrations and the vibrational coupling within the CBBC moieties of 1 and 1a are similar⁴⁹ and since the relevant structural data of the \( N₂CBBNC₂ \) unit of 1a and 1 are nearly identical (cf. below), this simplification is valid. The normal coordinate analysis of the vibrations of the \( N₂CBBNC₂ \) core of 1a is collected in Table 1.

The potential energy distribution (PED)⁵⁰ exhibits the strong coupling of the BB, BC and CN vibrations, which is a first hint for a strong BC bonding! Furthermore, the resulting force constants (mdyn Å⁻¹) \( f_{BB}, f_{BC} \) and \( f_{BB/BC} \) give a convincing picture of the bonding in the core of 1a and also of 1.²⁴

\[
\begin{align*}
 f_{BB} &= 6.0 \ (1.49 \ \text{Å}) \\
 f_{BC} &= 5.2 \ (1.49 \ \text{Å}) \\
 f_{BB/BC} &= 0.16
\end{align*}
\]

In order to get a feeling for these values and to discuss them, we will compare them with force constants of some species summarized in Table 2.

In order to decide about the possible multiple bond strength of the BB bond in 1 we must obtain reliable reference data for a BB single bond. As far as we know, only a single experimental value with 3.4 mdyn Å⁻¹ has been published (for B₂Cl₄).²⁸ Perhaps this value represents a weak BB single bond, because the BB distance in B₂Cl₄ with 1.73 Å is relatively large. The value for \( f_{BB} \) of the B₂ triplet molecule (\( ^{3}Σ \)) in its ground state with 3.6 mdyn Å⁻¹ is a little bit larger; however, the distance of the BB bond is significantly shorter (1.59 Å). Both parameters of this B₂ molecule are difficult to access because of the triplet character of B₂ and two “binding” electrons in orthogonal
π-bonds.\textsuperscript{56,55} In order to have a BB bond situation similar to that of 1 with a linear X–B–B–X moiety,\textsuperscript{56} we have finally calculated the force constants [mdyn Å\textsuperscript{-1}] within the O=B–B=O\textsuperscript{57} species (O=\textsuperscript{12}C=CH\textsubscript{2}) for which already the vibrational spectrum of the matrix isolated species was obtained two decades ago:\textsuperscript{58}

\[
\begin{align*}
\text{BB:} & \quad 3.5 \text{ (3.3)} \quad 1.59 \text{ (1.64)} \\
\text{BO:} & \quad 2.8 \text{ (3.7)} \quad 1.73 \text{ (1.70)} \\
\text{BC:} & \quad 7.7 \quad 1.40 \\
\end{align*}
\]

The force constant \(f_{\text{BB}}\) of 3.5 mdyn Å\textsuperscript{-1} corresponds to a BB single bond though the BB distance is shorter than in \(\text{B}_2\text{Cl}_4\) and longer than in \(\text{B}_3\text{(^1S)}\) (Table 1). However, the Lewis formula O=B–B=O is in accordance with the values of the BB and BC force constants. The value of the interaction force constant \(f_{\text{BB/BO}}\) is, like that of 1a, unexpectedly low (mostly about 10% of the stretching force constants), which demonstrates that the changes (stretching/compressing) of the BB bond have only a small influence on BC bonds (1a) or BO bonds in \(\text{B}_2\text{O}_2\). However, these interactions are, as expected, significantly larger for 1a than for \(\text{B}_2\text{O}_2\), i.e. for a more ionic compound.

The only example for a BB multiple bond within the molecules of Table 1 is observed in the excited \(\text{B}_2\text{(*)}\) molecule: The large value of \(f_{\text{BB}}\) of 7.7 mdyn Å\textsuperscript{-1} is in line with a small BB distance of 1.4 Å, i.e. a strong double bond is present in this molecule. From all these data we are now able to interpret the \(f_{\text{BB}} = 6.0\) mdyn Å\textsuperscript{-1} and \(f_{\text{BC}} = 5.2\) mdyn Å\textsuperscript{-1} force constant of 1a.

The BB bond should be addressed as a strong 1.5 BB bond and the BC bond as a weak 1.5 BC one. These results are in line with the thermodynamic discussion, from which the BB bond in \(\text{B}_2\text{(*)}\) has lost its strong double bond character and, even more important, the value of the BC force constant forbids to address it as a donor–acceptor bond. This BC bond is even significantly stronger than a 2e2c bond! Therefore, mainly the two strong BC bonds of 1 are responsible for the large energy gain of 12.6 eV discussed in the thermodynamic part.\textsuperscript{54,55,59,60}

The results of the thermodynamic data and of force constant determinations of 1 both accessible by observables show that a reassessment is pending concerning the bond description of 1. For a conclusive interpretation we have to decide whether to rely, besides the measured bond distance, on the predominant occupation of selected calculated MOs\textsuperscript{61} or – and this is the bonding description of multiple bonds we favor – to prefer an interpretation based on observed thermodynamic relationships and the force constants based on the observed vibrational spectra. These force constants reflect the slope of the potential energy curve near the equilibrium distance and are thus a confident measure of the bond strength. This argument has already been impressively demonstrated in the evaluation of CC multiple bonds in the past\textsuperscript{62} so that we had applied it also for assessing the GaGa multiple bonds that were under discussion about 20 years ago.\textsuperscript{62–65} Even at that time, we were able to show that the bond described as a GaGa triple bond was just a slightly stronger single bond.\textsuperscript{66–69}

### Conclusions

Thus, from the thermodynamic and the force constant discussion a new description of bonding results for 1.\textsuperscript{70} To sum up, 1 does not contain a BB triple bond and the description as a stabilization of an excited \(\text{B}_2\text{*}\) molecule via “arrows” of 2NHC molecules is strongly misleading.\textsuperscript{71} Therefore a bonding situation results which can be described by the resonance structures shown in Scheme 1 containing 4 π electrons for the BB and two BC bonds; i.e. the situation for these three bonds is just between single and double ones as represented by a simpler description.

### Tables

#### Table 1
Some vibrational frequencies of the model compound 1a, their assignment via the potential energy distribution (PED), and their isotopic shifts.

<table>
<thead>
<tr>
<th>(\nu/cm^{-1})</th>
<th>(\Delta \nu(\text{^{10}B}/\text{^{11}B}))</th>
<th>(\Delta \nu(\text{^{12}C}/\text{^{13}C}))</th>
<th>(\Delta \nu(\text{^{14}N}/\text{^{15}N}))</th>
<th>PED</th>
</tr>
</thead>
<tbody>
<tr>
<td>(a_1)</td>
<td>1156.94</td>
<td>2.53</td>
<td>9.88</td>
<td>20.27</td>
</tr>
<tr>
<td>(b_1)</td>
<td>1298.74</td>
<td>11.85</td>
<td>13.55</td>
<td>20.29</td>
</tr>
<tr>
<td>(a_2)</td>
<td>1450.95</td>
<td>6.93</td>
<td>8.29</td>
<td>26.56</td>
</tr>
<tr>
<td>(b_2)</td>
<td>1505.32</td>
<td>5.3</td>
<td>22.52</td>
<td>35.34</td>
</tr>
<tr>
<td>(a_3)</td>
<td>1582.80</td>
<td>0.85</td>
<td>0.04</td>
<td>2.55</td>
</tr>
<tr>
<td>(b_3)</td>
<td>1585.54</td>
<td>0.35</td>
<td>1.35</td>
<td>3.85</td>
</tr>
<tr>
<td>(a_4)</td>
<td>1769.50</td>
<td>64.72</td>
<td>10.79</td>
<td>11.09</td>
</tr>
</tbody>
</table>

\(a\) \(\nu^*\) means \(\nu\) (symmetric) but out of phase motion.

#### Table 2
Experimentally determined and calculated BB and BC force constants [mdyn Å\textsuperscript{-1}] and distances [Å] for species in this work as well as a weak BC bond of \(\text{H}_2\text{B–CO}\) via a 2e2c single bond of \(\text{H}_2\text{B–CH}_2\) to a double bond in \(\text{H}^–\text{CH}_2^+\).

<table>
<thead>
<tr>
<th>Molecule</th>
<th>(f_{\text{BB}}/f_{\text{BC}})</th>
<th>(r_{\text{BB}}/r_{\text{BC}})</th>
</tr>
</thead>
<tbody>
<tr>
<td>(\text{B}_3\text{(^1S)})</td>
<td>3.6\textsuperscript{a} (3.3)</td>
<td>1.59\textsuperscript{b} (1.64)</td>
</tr>
<tr>
<td>(\text{B}_3\text{(^1S)})</td>
<td>(7.7)</td>
<td>(1.40)</td>
</tr>
<tr>
<td>(\text{B}_2\text{Cl}_4)</td>
<td>3.4\textsuperscript{a} (2.83)</td>
<td>1.73\textsuperscript{b} (1.70)</td>
</tr>
<tr>
<td>(\text{H}_2\text{B–CO})</td>
<td>2.8\textsuperscript{c} (3.7)</td>
<td>1.52\textsuperscript{c}</td>
</tr>
<tr>
<td>(\text{H}_2\text{B–CH}_3)</td>
<td>(7.8)</td>
<td>(1.39)</td>
</tr>
<tr>
<td>(\text{H}_2\text{B–CH}_4)</td>
<td>(4.0)</td>
<td>(1.56)</td>
</tr>
<tr>
<td>1a</td>
<td>(6.0/5.2)</td>
<td>(1.49/1.49)</td>
</tr>
</tbody>
</table>

\(a\) Calculated values are in parentheses (cf. ESI). \(b\) Ref. 49. \(c\) Ref. 50. \(d\) Ref. 51.
for this 4-electron 4-center π-bond as NHC=B=B=NHC (Scheme 2).

To sum up, the bonding within 1 is determined by an overall electron transfer from the “triple” bond of B₂* to the BC bonds. This partial electron transfer is completed in the isolated normal valent B₂O₂ molecule, and consequently a BB single bond results: ⁷²

\[ \overline{O} = B - B = \overline{O} \]

Furthermore, from the discussion of 1 presented here a fundamental conclusion for the reaction of any NHC (cyclic (diamino)carbenes) as well as for cyclic (alkyl)(amino) carbenes such as in 2 during the “stabilization” of a reactive species X has to be drawn, in order to avoid serious problems for the description of bonding in many fields of inorganic chemistry: are the bonds between the X species and the NHC molecule – concluded from thermodynamics and from force constants as significant indicators of bond strength – weak donor–acceptor bonds symbolized by arrows, or are there strong covalent bonds which are possibly increased to have partial multiple bond character. ⁷³ The analysis of the variation of the BB vibration of 1 provides an easy indicator for the degree of its stabilization. Therefore, the bonding description of 1 presented here may also show exemplarily that the bonding discussion of any other NHC stabilized reactive species has to be seen critically. Anyway, the description of such bonding by arrows is at least strongly misleading, ⁷₄ as a more general controversy has already shown. ⁷₅⁻⁷⁷ However, our critical description does not lower the excellent work of H. Braunschweig and G. Robinson, but it increases the understanding of the bonding of their unprecedented compounds.

Acknowledgements

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Notes and references

19 The following convincing argument has been recommended by a referee: given that the covalent radii of C and B are ca. 0.76 and 0.84 Å and that a CC triple bond is ca. 1.18 Å in acetylene, an only slightly longer BB triple bond of ca. 1.30 Å (1.18 × 0.84 = 0.76), which is roughly proportionate to the different covalent radii of B and C, might have been expected (however: r(BB) = 1.45 Å for 1) for B≡B (Fig. 1). On the other hand, the B–C bond length (1.49 Å) is shorter than might be expected for a B–C single bond, and it lies (just) within the range known for B–C double bonds in methylene boranes (1.351–1.488 Å) suggesting that the B–C bond has multiple character.
23 Even two years after the publication of 1 ¹¹⁻¹² a further nice BB containing compound 2 has been published, ²¹ just when we had finished this manuscript. However, even for 2, which does not contain the cyclic di(amino)carbene of 1 but a cyclic (alkyl)(amino)carbene (CAAC), ²² a thermodynamic as well as a force constant discussion based on the vibrational spectra is missing.
27 The excited quartet state of B atoms is 314 kJ mol⁻¹ above the ground state.
28 However, some thermodynamic data are presented for the molecules COB₂CO and N₂B₂N₂ (cf. ref. 20).
29 With respect to this fragmentation, compound 2 is even 1 eV more stable than boron. Obviously, there are stronger BC bonds, as can be expected from the boron carbidines, which are exothermic compounds, cf. ESIF† and ref. 30.


31 More realistic are fragmentation and elimination reactions within the NHC moiety.

32 B2 decomposes exothermically to 2 B atoms in their ground state.

33 However, in all gas phase reactions we have studied in the past, the excited intermediates lose their energy via fragmentation (cf. ref. 34–38). Furthermore, we do not know any molecular example for which a fragmentation is avoided after an excitation of more than 10 eV.


41 However, there are well-known examples that the distance between two atoms can be a misleading property for the characterization of a bond (e.g. the BB distances in B4H6 with 1.77 Å without a direct bond is similar to the 2e2c BB bond in B2Cl4 (1.72 Å); furthermore, the following NaNa distances (Å) show that there is no correlation between bonding and distance: in solid NaCl (3.44), NaNa: 3.72, Na2-molecule: 3.0, (NaCl2)-molecule: 2.81 (see ref. 42 and 43).


44 Larger elongation to infinity distance between the vibrating entities means the dissociation energy ΔE diss mostly is not a representative measure of a bond, as the educt and the (fragmentation) products have different electronic and structural properties. Therefore, the force constant represents a more reliable characterization of a bond!


carbenes, e.g. in 2, are better donors and acceptors (cf. ref. 25).

61 The occupation of single MOs in principle is not directly correlated with a bond, because these separated occupation numbers are no observables. Only the sum of all MOs, their occupation numbers and energies result in observables such as the distance and the force constant.


66 For formal GaGa single/double/triple bonds the following GaGa force constants (mdyn Å⁻¹) and distances (Å; in parentheses) are obtained: Ga₂H₆²⁻: 0.75 (2.58); Ga₂H₄²⁻: 0.98 (2.46); Ga₂H₂²⁻: 1.01 (2.41). In order to show that no heavy-atom effects are responsible for this drastic difference to the situation in the first period (e.g. for organic chemistry), we have studied the situation within the following AsAs-bonded molecules, though arsenic is the element next to the neighbor of Ga in the periodic table; i.e. As and Ga differ by two electrons: As₂H₄: 1.38 (2.49); As₂H₃: 2.61 (2.27); As₂: 4.22 (2.12) (ref. 67).

71 Even the CN₂ vibrations of the NHC moiety of 1 show that they are strongly involved in the bonding and vibration to the “reactive” molecule B₂* (cf. Table 1), i.e. a new covalent bonded molecule is formed with hardly any relation to the “educts” B₂/B₂* and 2NHCs.

72 The same is valid for 2 (cf. ref. 24) which has nothing to do with B₂* and the discussed threefold bonding of 1 (cf. ref. 11); i.e. all argumentation presented here for 1 and its misleading bonding discussion shall be applied to 2. However, for 2 there are slightly stronger BC bonds, which correspond to the higher stability of 2 in comparison to 1. This property is in line with the exothermic behavior of solid boron carbides (ref. 30) and their covalent BC bonds, cf. ESI.

73 Furthermore, a description of this kind generates serious confusion for teaching inorganic chemists. 77


76 D. Himmel, I. Krossing and A. Schnepf, Angew. Chem., 2014, 126, 6159.


