Nano and micro-TiO₂ for the photodegradation of ethanol: experimental data and kinetic modelling

Claudia L. Bianchi, Carlo Pirola, Federico Galli, Marta Stucchi, Sara Morandi, Giuseppina Cerrato and Valentino Capucci

With TiO₂-photocatalysis being an effective alternative to other more expensive Advanced Oxidation Processes (AOPs), the possibility of using micro-sized TiO₂ materials rather than the well-known nano-sized powders is an important goal in terms of both handling safety and cost saving. In this work the photodegradation of ethanol, used as a model VOC (Volatile Organic Compound) molecule, was investigated, comparing the efficiency of both commercial nano- and micro- sized TiO₂ samples. In all cases the same degradation pathway was observed, namely, a consecutive first-order reaction with acetaldehyde as an intermediate product and CO₂ and water as the final products. All photocatalysts were characterized by means of XRD, TEM, IR, BET and XPS analysis. A kinetic model was also developed considering the collected experimental data and a regression of both adsorption and kinetic constants was made using MATLAB software. The optimized parameters were used for simulating the experimental data using an ode15s algorithm.

1 Introduction

Photocatalysis is a well-known process in which a chemical reaction occurs when a semiconductor is hit by radiation that complies with its band gap. In particular, when the semiconductor is irradiated by a light source with an appropriate wavelength, some of its electrons are transferred from the valence band to the conduction band, creating electron–hole pairs that are capable of causing reduction and oxidation reactions.

TiO₂ is an excellent photocatalyst used for degrading organic compounds, which are oxidized according to the aforementioned mechanism. In particular, UV irradiation induces the formation of electron–hole pairs, whose charge carriers react with chemical species such as water molecules (humidity) and molecular oxygen present in the air to produce hydroxyl radicals (·OH) and superoxide radical anions (O₂⁻), respectively, which contribute to the decomposition of organic molecules at the TiO₂ surface.

Chemical stability and also the high oxidizing power of the photogenerated holes are the main characteristics that make TiO₂ the most used and investigated photocatalyst and also the key factors for the pollutant oxidation.

Pollution abatement is something that is very important nowadays, but worldwide research needs to develop new “green” technologies in order to limit energy consumption and save money, and also to take advantage of processes with no environmental drawbacks. In a sense, the TiO₂-photodegradation of pollutants, in particular organic compounds, nitrogen and sulphur oxides, is an effective alternative to the much more expensive advanced oxidation processes (AOPs). Moreover, as most of human life spent in an indoor environment, and most of the soilage of the interior of buildings comes from VOCs, their degradation is a very crucial point in order to improve both air and human health quality.

In this study, two different commercial TiO₂ samples were tested in the photodegradation reaction of ethanol, which was chosen as a model molecule. Furthermore, considering it as an important atmospheric pollutant, commonly used as an industrial solvent and a fuel additive, its oxidation is of interest because it is emitted from many industrial processes such as breweries and bakeries. Ethanol emissions accounted for about 4% of the total VOC anthropogenic emissions in the UK in 1993.

Moreover, the atmospheric concentrations of this pollutant are rising as a consequence of the use of ethanol as a biofuel in the automotive sector and it is well established that this fact implies an increase in the atmospheric levels of acetaldehyde, which is far more toxic and reactive in the atmosphere than ethanol. Besides this, acetaldehyde, also present in the ethanol photodegradation pathway, is an important indoor pollutant itself, released by some building materials such as polyurethane foams, and some consumer products such as...
cigarettes, adhesives, coatings and inks.\textsuperscript{11} In fact, ethanol is first oxidized to acetaldehyde and then completely degraded to carbon dioxide and water.

The possibility of using the common pigmented micro-sized TiO\textsubscript{2} as a photocatalyst has already been proved,\textsuperscript{12} and, although a reduced crystal size ensures a higher surface area (that is positive in terms of photocatalytic activity), it is a disadvantage in terms of safety and product handling; in fact, nano-powders can be very dangerous for health, as reported in several recent papers.\textsuperscript{13,14} A nanomaterial consists of one or more components, present in various forms, that possesses at least a one-dimensional structure with an average diameter in the 1–100 nm range. A quick search through the literature shows that the bulk of the research articles on nanomaterials genotoxicity were published within the past 3 years, even though their cytotoxicity and carcinogenic responses, both \textit{in vitro} and \textit{in vivo}, were assessed previously.\textsuperscript{15} As the development of nanotechnological applications continue to grow, it is anticipated that there will be an even greater demand for safety, health and risk assessment studies in the forthcoming years.\textsuperscript{16}

In the present paper, two micrometric commercial samples and the well-known nano-sized P25 by Evonik have been chosen with a two-fold aim: \textit{i.e.}, not only to evaluate the semiconductor efficiency, considering the different crystallite sizes, but also to verify the ethanol photodegradation pathway with all the tested materials.

From a kinetic point of view, the whole process can be simplified considering a first reaction on the catalyst surface, in which the adsorbed ethanol is converted to the adsorbed acetaldehyde, followed by a second reaction, in which the adsorbed acetaldehyde is mineralized to carbon dioxide and water, following a Langmuir model. The experimental data gathered were used for the regression of the characteristic kinetic parameters, assuming pseudo first order kinetic constants.

2 Experimental

Ethanol is a Fluka product at a high purity grade (99.9\%). Three powdered commercial TiO\textsubscript{2} samples were selected without further treatment before their characterization and use: (i) a nanometric sample by Evonik (P25), always used as a photocatalytic reference material and (ii) two micrometric samples (1077 by Kronos and A-HR by Hunsdman), sold as pigments and chosen with the following features: pure anatase, untreated surfaces, an undoped material, and a powdery form.

2.1 Sample characterization

The surface area of all the catalysts was determined by conventional N\textsubscript{2} adsorption (BET) at 77 K using a Sorptometer (Costech Mod. 1042).

X-ray photoelectron spectra (XPS) were taken in an M-probe apparatus (Surface Science Instruments). The source was monochromatic Al K\textsubscript{α} radiation (1486.6 eV).

The morphology of the catalysts was inspected by means of high-resolution transmission microscopy (HR-TEM): images were recorded using a JEOL 3010-UHR instrument (acceleration potential: 300 kV; LaB\textsubscript{6} filament).

Absorption/transmission IR spectra were obtained on a Perkin-Elmer FT-IR System 2000 spectrophotometer equipped with a Hg–Cd–Te cryo-detector, working in the 7200–580 cm\textsuperscript{-1} range at a resolution of 2 cm\textsuperscript{-1} (number of scans ~ 20). Before the IR analysis, powdery samples were compressed in self-supporting discs (of about 10 mg cm\textsuperscript{-2}) and placed in a homemade quartz cell, equipped with KBr windows, connected to a conventional high-vacuum line (UHV). The spectra were recorded at room temperature (RT) on the samples in air or after prolonged outgassing at RT.

For the characterization of the light absorption features and band-gap determinations, diffuse reflectance spectra (DRS) of the powders were collected using a UV-vis scanning spectrophotometer (PerkinElmer, Lambda 35), which was equipped with a diffuse reflectance accessory.

The XRD spectra were collected using a PW 3830/3020 X Pert Diffractometer from PANalytical working Bragg–Brentano, using the Cu Ka\textsubscript{1} radiation (λ = 1.5406 Å).

2.2 Photocatalytic tests

Photocatalytic degradations were conducted in a cylindrical glass reactor (diameter = 200 mm, volume = 5 l). The amount of catalyst used in each test was 0.05 g and it was deposited on a glass support as a thin film. The latter is obtained by preparing a suspension of isopropanol and TiO\textsubscript{2}, and depositing it on the glass slide through a Pasteur pipette.\textsuperscript{17} The reactor was initially charged with chromatographic air humidified at 40% RH. The starting ethanol concentration was set at 400 ppm. Photon sources were provided by a 500 W UV lamp (Jelosil model HG 500) emitting in the 315–400 nm wavelength range (UV-A) at 30 Wm\textsuperscript{-2} on the sample surface and each test lasted 4 h.\textsuperscript{18} The ethanol photodegradation was monitored, using on-stream gas chromatography (Agilent 3000 A microGC), directly connected to the reactor, that automatically measures the internal concentration of ethanol, acetaldehyde and carbon dioxide.

Photolysis tests were also performed for both pure ethanol and acetaldehyde to monitor the possible molecule degradation under UV light in the absence of a photocatalyst. The same procedure was repeated using acetaldehyde. Adsorption tests were also carried out by monitoring the concentration for 4 hours for both ethanol and acetaldehyde in the dark to obtain the adsorption constants (k) separately, in order to reduce the model parameters to be estimated. This approach has been already used in the literature for the independent determination of adsorption and kinetic constants.\textsuperscript{19,20}

All the photocatalytic tests carried out on each powder have been repeated at least 5 times, exhibiting the same behavior for each run.

The regression of both the adsorption and the kinetic constants were made using MATLAB software, minimizing the sum of the squared errors between the calculated and the experimental data. The optimized parameters were used for simulating the experimental data using a differential equation solver able to handle stiff problems, \textit{i.e.} the ode15s algorithm.\textsuperscript{21}
3 Results and discussion

3.1 Morphological characterization

The morphological characteristics of all the used catalysts are reported in Table 1.

XRD analysis confirms the well-known P25 phase composition, 80:20 in anatase/rutile ratio, and shows that anatase is the unique polymorph for the other two samples, 1077 and A-HR. One of the most critical drawbacks of both the micrometric samples is their low specific surface area. However, the presence of the bare anatase polymorph is confirmed, and this is a fundamental feature in terms of the photocatalytic performance.

As for the general morphology of the various samples, low magnification TEM images are displayed in Fig. 1. It is evident that P25 totally reveals its nano-sized nature (average particle size ~ 20–30 nm) if compared to both the 1077 and A-HR powders (for which the average particles size is much larger, lying in the 100–180 nm range). The high-resolution images relative to all the samples (see the inset in each image reported in Fig. 1) confirm the high crystallinity exhibited by the powders and the phase composition, as well.

BET surface area data are consistent with the above considerations. In fact, a particle size located in the nano-size range, as in case of P25, implies a much higher surface area and vice versa the opposite trend is expected for the micro-sized powders (see Table 1, column 2 and 3).

The surface state of the TiO$_2$ particles was analyzed by XPS. No significant differences can be observed in the Ti 2p region concerning the binding energies (BEs). The peak Ti 2p$_{3/2}$ is always regular and the BE (see Table 1, column 4) compares well with the data for Ti(IV) reported in the literature for TiO$_2$ materials.

The analysis of the oxygen peaks reveals the presence of more than one component, which can be attributed to lattice oxygen in TiO$_2$ (529.9 eV) and to surface OH species (>531.5 eV).

The hydrophilicity/hydrophobicity characteristics of the photocatalyst surface plays a crucial role in determining the ethanol adsorption and thus the final photocatalytic activity, at least in the degradation of pollutants.

3.2 Kinetic parameters regression

The degradation reaction is followed considering this pathway:

\[
\begin{align*}
&\text{(1)} \quad CH_3CH_2OH_{ads} \rightarrow K_1 CH_2COH_{ads} \\
&\text{(2)} \quad CH_2COH_{ads} \rightarrow K_2 2CO_2 + H_2O
\end{align*}
\]

Photocatalyzed reactions in which a gas phase reactant is involved are usually interpreted using the Langmuir–Hinshelwood mechanism, where only the reactant molecules previously adsorbed on the catalytic surface take part in the reaction. This kinetic interpretation is well-known for photocatalyzed reactions, as reported in some previous papers.

For the regression of the kinetic parameters in the present paper, two consecutive first order reactions were considered, as shown in the scheme reported in the introduction section. The differential equation system that characterizes the system is reported here:

\[
\begin{align*}
\frac{dn_e}{dt} &= -k_1 \frac{K_e n_a}{1 + K_e n_a + K_n n_a} \\
\frac{dn_a}{dt} &= k_1 K_e n_a - k_2 K_n n_a \\
\frac{dn_{CO_2}}{dt} &= +2k_2 1 + K_e n_a + K_n n_a
\end{align*}
\]

\(K_e\) and \(K_n\) (mol$^{-1}$) are the adsorption constant for ethanol (e) and acetaldehyde (a), respectively, and \(k_1\) and \(k_2\) (mol min$^{-1}$) are the first order kinetic constants. As can be observed, the adsorption of CO$_2$ was not considered.

The adsorption constant \(k_1\) was obtained considering reversible adsorption for both ethanol and acetaldehyde. The differential equation used for this regression is:

\[
\frac{dn_i}{dt} = -k_{ads,i} n_i + k_{des,i} (n_{hi} - n_i)
\]

Table 1 Main features of the selected TiO$_2$ samples

<table>
<thead>
<tr>
<th>Sample</th>
<th>Anatase/Rutile composition</th>
<th>BET surface area (m$^2$ g$^{-1}$)</th>
<th>Average crystallite size (nm)</th>
<th>XPS Ti 2p$_{3/2}$ (eV)</th>
<th>XPS OH/O$_{tot}$</th>
<th>Band gap (eV)</th>
</tr>
</thead>
<tbody>
<tr>
<td>P25</td>
<td>80% anatase 20% rutile</td>
<td>52</td>
<td>26</td>
<td>458.4</td>
<td>0.14</td>
<td>3.21</td>
</tr>
<tr>
<td>1077</td>
<td>100% anatase</td>
<td>11</td>
<td>130</td>
<td>458.4</td>
<td>0.32</td>
<td>3.15</td>
</tr>
<tr>
<td>A-HR</td>
<td>100% anatase</td>
<td>12</td>
<td>130</td>
<td>458.0</td>
<td>0.14</td>
<td>3.15</td>
</tr>
</tbody>
</table>
for either ethanol or acetaldehyde, $n_{0,i}$ is the moles of the component charged in the reactor and $n_i$ is the moles of the component at a certain time. The adsorption constant is then calculated by dividing $k_{\text{ads}}$ by $k_{\text{des}}$ after the regression.

### 3.3 Photolysis and adsorption tests

In order to determine the real role of the photocatalytic process, both photolysis and simple ethanol absorption were investigated.

The photolysis contribution is negligible, and it can be concluded that both ethanol and its oxidized intermediate product, i.e. acetaldehyde, are not sensitive to the bare UV light, and their concentrations remain stable for the entire test.

The role of absorption is different, as displayed in Fig. 3a (P25) and b (Kronos 1077) (AH-R sample gave very similar results to the 1077 adsorption and is not reported for the sake of clarity).

The optimized adsorption constants calculated from these experimental data are reported in Table 2:

The results indicate that P25 possesses a higher adsorption constant for ethanol if compared to that exhibited by both micrometric samples. On the contrary, acetaldehyde results show an opposite trend and the adsorption of these species seems to have no relation with the OH/O surface ratio of the catalyst.

In Fig. 4–6 the experimental data obtained in the photocatalytic runs performed using the different catalysts have been reported, together with the calculated trend using the optimized kinetic parameters.

The regressed kinetic parameters are reported in Table 3:

![Fig. 1 TEM images of the various TiO$_2$ samples.](image1)

![Fig. 2 O$_{1s}$ XPS spectra (A) and FTIR spectra (B) recorded at RT in air for (a) P25, (b) A-HR, (c) 1077.](image2)
The analysis of the kinetic elaboration gives us important information about the rate of reaction using either the nanometric or micrometric samples. The rate of VOC degradation is in general increased if catalysed by a nanometric sample with respect to the micrometric ones. This phenomenon is particularly important for some light compounds (i.e., acetaldehyde). Nevertheless, the catalytic properties of micrometric samples are confirmed from both direct experimental conversion and this kinetic interpretation. In particular, considering the degradation of ethanol, the rate of its conversion is not so different for either the nanometric or micrometric samples.

The FTIR data obtained for the TiO₂ powders, either before and after the photodegradation processes, indicate that ethanol leaves some residues (traces) onto the surface of the supports, independently of their micro- or nanosized nature. See for

### Table 2 Adsorption constants regressed from experimental data

<table>
<thead>
<tr>
<th>Component</th>
<th>P25</th>
<th>1077</th>
<th>A-HR</th>
</tr>
</thead>
<tbody>
<tr>
<td>Ethanol</td>
<td>0.2373</td>
<td>0.1891</td>
<td>0.1884</td>
</tr>
<tr>
<td>Acetaldehyde</td>
<td>0.0786</td>
<td>0.1680</td>
<td>0.1809</td>
</tr>
</tbody>
</table>

### Table 3 Regressed kinetic constant for all the photocatalysts

<table>
<thead>
<tr>
<th>Sample</th>
<th>$k_1$ [mol min$^{-1}$]</th>
<th>$k_2$ [mol min$^{-1}$]</th>
<th>SSE</th>
</tr>
</thead>
<tbody>
<tr>
<td>P25</td>
<td>0.68</td>
<td>10.41</td>
<td>$3.91 \times 10^{-9}$</td>
</tr>
<tr>
<td>1077</td>
<td>0.34</td>
<td>0.42</td>
<td>$1.69 \times 10^{-9}$</td>
</tr>
<tr>
<td>A-HR</td>
<td>0.39</td>
<td>0.42</td>
<td>$6.88 \times 10^{-10}$</td>
</tr>
</tbody>
</table>
instance Fig. 7, in which we reported the results relative to the nanometric P25 powder (top section of the figure) and to one of the two micrometric materials, namely Kronos 1077 (bottom section of the figure), as the A-HR powder exhibited a trend (almost) coincident to that of Kronos 1077.

In fact, if we refer to the spectral range (1800–1500 cm⁻¹) in which the νC=O mode(s) typical of the possible degradation products of ethanol (ketone/alddehyde) are normally located,²⁸ two net components are evident at ~1715 and ~1695.83 cm⁻¹. This confirms that an effective degradation has occurred, but a few (really only traces) of the degraded products are still present at the surface of the TiO₂ powders. Moreover, for the P25 material, other spectral components located at 1602 and 1575 cm⁻¹, reveal the presence of short chain carboxylated species (namely, acetate/formate species) chemisorbed to the TiO₂ surface. This provides evidence of the formation of acetic acid as an intermediate in acetaldehyde degradation, at least in the case of the nano-sized photocatalyst considered here.

This is also in agreement with the calculated kinetic parameters reported in Table 2. In fact, the P25 material, due to its intrinsic nanometric nature, is confirmed to have the best performance as a photocatalyst, particularly for the degradation of acetaldehyde, as the (k₃) kinetic constant is an order of magnitude larger, if compared to the micrometric samples.

However, considering the bare ethanol degradation, the micrometric samples are able to oxidize it almost completely and thus can be employed for its degradation, being them far less dangerous for human health.

4 Conclusions

The photodegradation of ethanol, as a model for VOCs, was performed using a commercial nanometric sample and two different micrometric powders. As well as presenting this model, the possibility to exploit the photocatalytic potential of the pigmentary TiO₂ is proved, contributing to the already useful photocatalytic processes. The micrometric TiO₂, which is less dangerous and much less expensive with respect to the nanometric P25, is an active photocatalyst, being able to degrade VOCs into CO₂. Also, the rate of reaction and the kinetic constants are lower with respect to P25. The good fitting between the experimental and simulated results confirms the assumption of a consecutive first order reaction mechanism degradation pathway that is not influenced by the TiO₂ crystallite dimension.

Acknowledgements

This research was supported by LIFE+ Environment Policy and Governance project LIFE13 ENV/IT/000140.

References