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# Hypercrosslinked porous polymer nanosheets: 2D RAFT agent directed emulsion polymerization for multifunctional applications†

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Two-dimensional hyper-cross-linked microporous polymers (2D HCPs) have been readily synthesized by emulsion polymerization using trithiocarbonate functionalized graphene oxide as a 2D reversible addition–fragmentation chain transfer (RAFT) agent. After a gel-type and permanently porous poly(vinyl-benzyl chloride-co-divinylbenzene) precursor resin was grafted to the graphene surface, an FeCl<sub>3</sub>-promoted Friedel–Crafts reaction was carried out for benzyl chloride groups to construct a porous structure confined within 2D nanosheets. The resulting HCPs exhibited superhydrophobic behavior, and had a predominance of micropores with a specific surface area of up to 1224 m<sup>2</sup> g<sup>−1</sup>. Moreover, they showed improved thermal stability in comparison with unadorned HCPs obtained without using the graphene template. Their H<sub>2</sub> and CO<sub>2</sub> capacities at 77 and 273 K reached 1.27 and 9.74 wt%, respectively, at a relative pressure of 0.99. In addition, the as-synthesized 2D HCPs were used as carbon precursors to generate 2D porous carbons with a high specific surface area (871 m<sup>2</sup> g<sup>−1</sup>) and high-performance electrochemical energy storage (144 F g<sup>−1</sup> at 0.5 A g<sup>−1</sup>).

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## Introduction

Porous polymers<sup>1–6</sup> with organic porous structures have attracted tremendous attention owing to their controllable porous features associated with prominent physical properties and potential applications, such as in light harvesting,<sup>7</sup> sensing,<sup>8</sup> gas separation<sup>9</sup> and storage,<sup>10</sup> catalysis,<sup>11–13</sup> and energy storage and conversion.<sup>14</sup> Among them, hyper-cross-linked porous polymers (HCPs) represent one of the fastest developing types of porous polymers—not only because of their high specific surface areas and gas storage properties but also due to the easy availability of many controllable polymerization methods and the wide variety of functional monomers.<sup>15</sup> Most of the reported HCPs have been amorphous, and were synthesized by traditional radical<sup>10,16–18</sup> and oxidation<sup>19</sup> polymerization methods, followed by an additional step for

further cross-linking. Most of the previous studies on HCPs have focused on pore construction, tuning the gas<sup>19–21</sup> or small-molecule<sup>22,23</sup> separation/storage properties through optimizing the reaction conditions such as the monomer ratio, catalyst percentages, the temperature, and the reaction time.<sup>10,24</sup> However, effort dedicated to morphology control has rarely been reported for HCPs. Only quasi-zero-dimensional microsphere<sup>17,25</sup> and three-dimensional monolith<sup>22,23</sup> HCP types have been successfully synthesized. Two-dimensional (2D) HCPs are still under exploration owing to the lack of synthetic methods or 2D templates and the thermodynamic instability of 2D porous polymers in comparison with their zero- and three-dimensional porous counterparts.

Graphene, which consists of a honeycomb-like hexagonal array of sp<sup>2</sup>-bonded carbon atoms, exhibits many remarkable properties such as superior electrical conductivity, a large surface area, excellent mechanical flexibility, and high thermal/chemical stability.<sup>26,27</sup> Graphene oxide (GO), one of the most widely synthesized derivatives of graphene and made from graphite by a modified Hummers' method, is an ideal 2D template for the preparation of 2D porous polymers because GO can be easily functionalized with many functional groups and solution processing.<sup>28–33</sup> It is well known that the properties of graphene-based nanocomposites<sup>34</sup> are highly dependent on the good dispersibility and interfacial interactions of graphene. In order to produce a homogeneous graphene

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dispersion during the preparation procedure, many attempts have been made to attach small molecules and linear polymers to the graphene surface by means of covalent bonds and weak interactions.<sup>35</sup> Several grafting routes have been explored using either GO or functionalized GO with good dispersibility in various organic solvents as templates for the preparation of graphene-based 2D brushes and “sheet-coil” polymers.<sup>36</sup> Ruoff and co-workers have prepared GO-based macroinitiators by the functionalization of GO with  $\alpha$ -bromoisobutryl bromide for controlled styrene grafting *via* atom transfer radical polymerization (ATRP).<sup>37,38</sup> Almost all the reported studies have focused on linear polymers, such as polystyrene,<sup>39,40</sup> poly(*N*-vinylcarbazole),<sup>41</sup> polymethacrylamide,<sup>42,43</sup> and poly(*N*-isopropylacrylamide),<sup>44</sup> grafting from or grafting to graphene surfaces. To the best of the authors' knowledge, the preparation of 2D porous polymers using the 2D chain-transfer agent (CTA) promoted reversible addition–fragmentation chain transfer (RAFT) polymerization method has not yet been reported.

In this work, a highly soluble trithiocarbonate-based CTA was used to functionalize GO to act as a 2D CTA as well as a 2D template. Then, poly(vinylbenzyl chloride-*co*-divinylbenzene) (PVD) grafted GO was prepared by RAFT emulsion polymerization based on the as-prepared 2D CTA. After further hyper-cross-linking by an FeCl<sub>3</sub>-promoted Friedel–Crafts reaction of the benzyl chloride groups, graphene-based sandwich-type HCPs with a 2D morphology (GHCPs) were successfully prepared on a large scale (~100 g per synthesis). The resulting GHCPs exhibited superhydrophobic behavior (contact angle >150°) and had high specific surface areas of up to 1224 m<sup>2</sup> g<sup>−1</sup>. They also exhibited an enhanced H<sub>2</sub> capacity at 77 K (1.27 *versus* 0.85 wt%) and CO<sub>2</sub> capacity at 273 K (9.74 *versus* 8.80 wt %) at a relative pressure of 0.99 in comparison with unadorned HCP obtained without using the graphene template. The high specific surface areas and the enhanced H<sub>2</sub> and CO<sub>2</sub> capacities can be the contributions of the 2D graphene template. Owing to the carbon backbone of graphene, GHCPs were used to prepare porous carbon nanosheets without removing the template. The as-prepared porous carbon nanosheets exhibited the typical 2D morphology and a high surface area (871 m<sup>2</sup> g<sup>−1</sup>), and were successfully used in electrochemical capacitors (144 F g<sup>−1</sup> at 0.5 A g<sup>−1</sup>). These results indicate that this method is a good approach for the preparation of 2D HCPs and 2D porous carbons with multifunctional applications.

## Experimental section

### Reagents and chemicals

Natural flake graphite, 1-dodecanethiol, carbon disulfide, aliquot 336, 1,3-diaminopropane, *N*-hydroxy-succinimide (NHS), 1-(3-dimethylaminopropyl)-3-ethylcarbodiimide hydrochloride (EDC-HCl), divinylbenzene (DVB), 2,2'-azobis(2-methylpropio-nitrile) (AIBN), and poly(vinyl alcohol) (PVA) were purchased from Aladdin (Shanghai). Vinylbenzyl chloride (VBC) was purchased from TCI. 1,2-Dichloroethane (DCE),

tetrahydrofuran (THF), FeCl<sub>3</sub>, and other chemicals were purchased from Sinopharm Chemical Reagent Co. (Shanghai). THF, DCE and triethylamine were dried by standard methods before use.

### Functionalization of graphene oxide with amine groups<sup>45</sup>

GO was prepared from natural graphite flakes by a modified Hummers' method.<sup>46</sup> GO sheets (1.0 g) were added to dry DMF (250 mL). After ultrasonication for 24 h, NHS (3.42 g, 30 mmol) and EDC-HCl (5.75 g, 3.2 mmol) were added to GO dispersion at 0 °C. After stirring for 2 h, 1,3-diaminopropane (5.0 mL) was added, and the solution was stirred overnight at room temperature. After filtration and washing with water and ethanol, the resulting powder was freeze-dried and denoted as GO-NH<sub>2</sub>.

### Preparation of 2D CTA

2-(Dodecylthiocarbonothioylthio)-2-methylpropionic acid (DDAT), which is a widely used and highly soluble chain transfer agent (CTA) for RAFT polymerization, was synthesized according to reported methods.<sup>47,48</sup> GO-NH<sub>2</sub> (1.0 g) was added to dry DMF (250 mL) and sonicated for a few hours. Then, triethylamine (5.0 mL) was added to the dispersion. Acylchloride-DDAT was prepared as follows: DDAT (5.0 g) and excess thionyl chloride (20 mL) were added to a 50 mL three-neck round-bottom flask and refluxed overnight under a nitrogen atmosphere. Then, unreacted thionyl chloride was completely removed under reduced pressure and washed with anhydrous THF (8 mL × 3). Lastly, the acylchloride-DDAT in THF (10 mL) was transferred to the above GO-NH<sub>2</sub> dispersion very slowly under vigorous stirring in an ice bath. The mixtures were stirred at room temperature for 24 h. The resulting DDAT functionalized GO (GO-DDAT) was washed with CH<sub>2</sub>Cl<sub>2</sub> and ethanol, and dried overnight under vacuum at 50 °C.

### RAFT emulsion polymerization

VBC/DVB modified GO precursor resin was prepared by conventional RAFT emulsion polymerization.<sup>49,50</sup> The aqueous phase consisted of distilled water (50 mL), PVA (0.38 g) and NaCl (0.17 g). The organic phase comprised GO-DDAT (100 mg), VBC (6.37 g, 41.7 mmol), DVB (0.16 g, 1.2 mmol), and AIBN (33 mg). The organic phase was suspended in the aqueous phase at 80 °C by a stirring speed of 425 rpm under a N<sub>2</sub> atmosphere. After 8 h, the beads were filtered and washed three times with water, methanol and diethyl ether. After drying under vacuum overnight at 60 °C, the copolymer grafted GO (GO-PVD) was produced. Bare PVD without using a graphene template was also synthesized using the same method.

### Synthesis of 2D hypercrosslinked microporous polymer

The VBC/DVB modified GO precursor resin beads (3.25 g) were swollen in DCE (40 mL) for 3 h. The mixture was cooled in an ice bath before adding FeCl<sub>3</sub> (2.0 g). After achieving a uniform dispersion of FeCl<sub>3</sub> throughout the precursor beads in 2 h, the mixture was heated to 80 °C for 12 h. After the resulting hyper-



Polym. Chem., 2015, 6, 7171–7178 | 7173



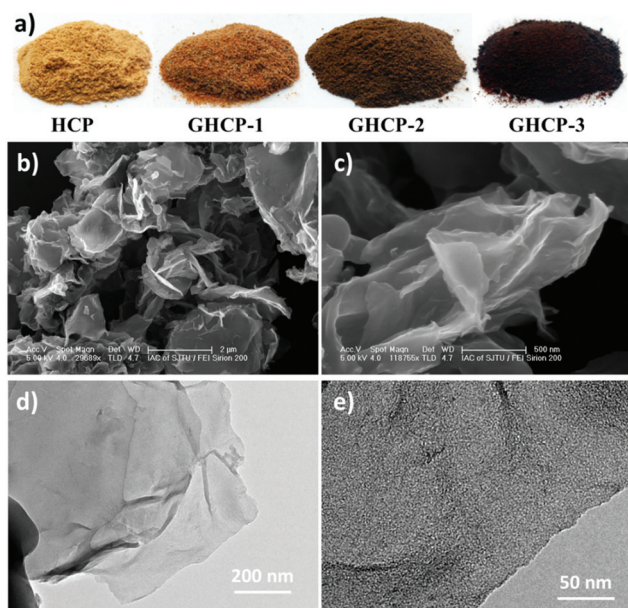


Fig. 1 (a) HCP and the GHCPs. (b, c) SEM and (d, e) TEM images of GHCP-3.

To analyze the molecular structure of the HCPs, Fourier transform infrared (FT-IR) spectroscopy was employed. Typical FT-IR spectra for GO, GO-DDAT, unadorned HCP, and GHCP-3 are shown in Fig. 2. The broad band at  $3430\text{ cm}^{-1}$  is attributed to O–H stretching vibrations,<sup>54</sup> which might originate from carboxylic acid and hydroxyl groups on the GO sheets or from adsorbed water. GO shows two adjacent peaks at  $1720$  and  $1645\text{ cm}^{-1}$ , which correspond to C=O and C=C stretching vibrations.<sup>55</sup> The peak at  $1065\text{ cm}^{-1}$  represents C–O stretching vibrations,<sup>56</sup> which indicate the presence of the epoxide group in GO.<sup>57</sup> In Fig. 2, unadorned HCP and GHCP-3 show a significant peak at  $1610\text{ cm}^{-1}$  (skeletal vibrations from graphitic domains), attributable to aromatic C=C.<sup>54</sup> Additionally, the

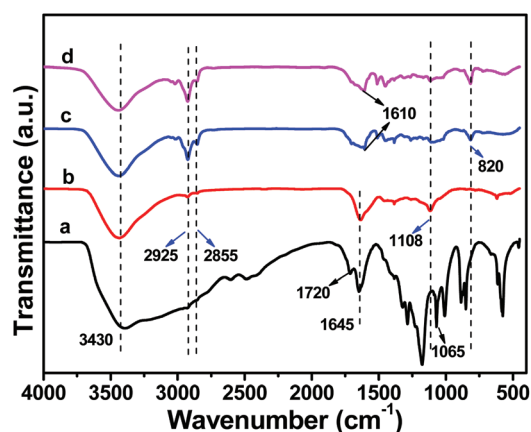


Fig. 2 FT-IR spectra of (a) GO, (b) GO-DDAT, (c) unadorned HCP and (d) GHCP-3.

peaks at  $2925$  and  $2855\text{ cm}^{-1}$  assigned to the methylene stretch indicate the existence of  $\text{CH}_2$  or  $\text{CH}$  groups.<sup>54</sup> In particular, the peak at  $1065\text{ cm}^{-1}$  for GO shifted to  $1108\text{ cm}^{-1}$  for GO-DDAT and GHCP-3, which might be attributable to hydrogen bond interactions.<sup>58</sup> Compared with GO and GO-DDAT, unadorned HCP and GHCP-3 have a distinctive peak at  $820\text{ cm}^{-1}$ , corresponding to C–Cl stretching,<sup>59</sup> which suggests the existence of a few unreacted  $\text{Ar-CH}_2\text{-Cl}$  groups (see the model reaction in Scheme 1).

In order to reveal the porous structures of the 2D HCPs, the  $\text{N}_2$  adsorption/desorption properties of GHCPs and unadorned HCP were studied (Fig. 3 and Table 1). All the samples showed a distinct step in the  $\text{N}_2$  adsorption isotherm at a low relative pressure ( $P/P_0 < 0.1$ ), which corresponds to gas sorption in micropores.<sup>60</sup> According to the IUPAC classification,<sup>61</sup> these samples give rise to type I nitrogen gas sorption isotherms with H3 hysteresis loops, indicating that the materials contain micro- and mesopores. The specific surface areas of  $1224$  and  $1693\text{ m}^2\text{ g}^{-1}$  (the former calculated from the Brunauer–Emmett–Teller (BET) theory, the latter from Langmuir’s theory) for GHCP-1 are larger than those of HCP ( $600$  and  $832\text{ m}^2\text{ g}^{-1}$ ), respectively. This result could be attributable to the graphene template and the resulting 2D morphology. This analysis was further confirmed by a pore size study. Pore size distributions calculated by the nonlocal density functional theory (NL-DFT) method for HCP and the GHCPs are illustrated in Fig. 3b. All of the samples exhibited pore diameters centering around  $1.6\text{ nm}$ , implying that they contained mainly micropores ( $<2.0\text{ nm}$ ). The GHCPs also showed several broad peaks between  $2.3$  and  $6.0\text{ nm}$ , implying the existence of a few mesopores ( $2.0\text{--}50.0\text{ nm}$ ). These results—when considered together—indicate that the GHCPs prepared in this study are new hierarchical porous organic materials with a 2D morphology.

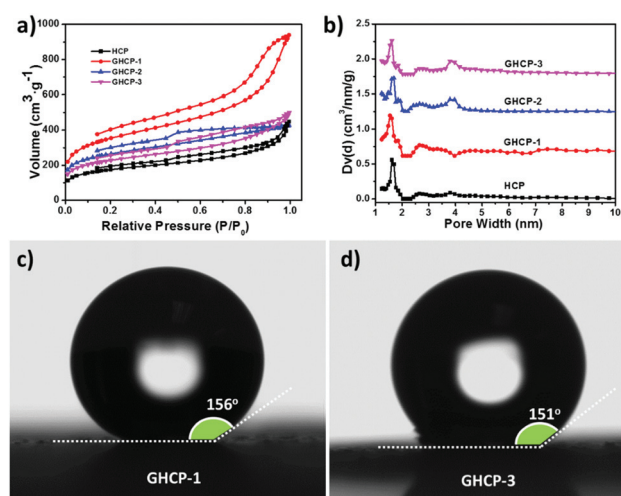


Fig. 3 (a) Nitrogen adsorption and desorption isotherms of unadorned HCP and the GHCPs at  $77\text{ K}$ . (b) The pore size distribution calculated by the NL-DFT method. (c, d) Illustrative images of a water droplet on the superhydrophobic surfaces.



**Table 1** Nitrogen physisorption properties and H<sub>2</sub> and CO<sub>2</sub> uptake for HCP and the GHCPs

	$S_{\text{BET}}^a$ (m <sup>2</sup> g <sup>-1</sup> )	$S_{\text{Lang}}^a$ (m <sup>2</sup> g <sup>-1</sup> )	$V_{\text{tot}}^b$ (cm <sup>3</sup> g <sup>-1</sup> )	H <sub>2</sub> uptake <sup>c</sup>		CO <sub>2</sub> uptake <sup>c</sup>	
				77 K	87 K	273 K	298 K
HCP	600	832	0.69	0.85	0.63	8.80	5.86
GHCP-1	1224	1693	1.45	1.27	0.82	9.74	6.19
GHCP-2	971	1344	0.71	1.16	0.77	9.84	6.22
GHCP-3	780	1083	0.78	1.09	0.67	9.06	5.82

<sup>a</sup> Surface area calculated from the N<sub>2</sub> adsorption isotherm using the BET method and the Langmuir method, respectively. <sup>b</sup> Total pore volume at  $P/P_0 = 0.99$ . <sup>c</sup> Gravimetric gas uptake (wt%) for H<sub>2</sub> and CO<sub>2</sub> at  $P/P_0 = 0.99$ .

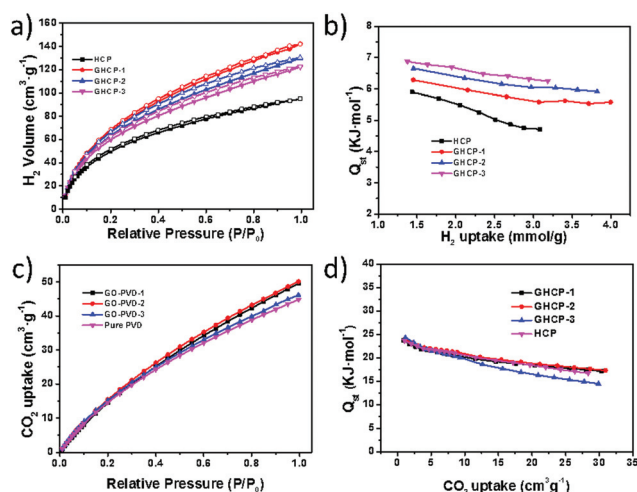
Additionally, the hydrophobic property of the GHCPs was studied to better understand their unique hierarchical porous structures. First, ethanol dispersions of the GHCPs and HCP were sprayed onto glass surfaces. After drying in air at room temperature overnight, water droplets were placed on these surfaces (Fig. 3c and d). All the GHCPs showed a contact angle of  $155 \pm 4^\circ$ , which is typically superhydrophobic, and marginally larger than that of HCP ( $151 \pm 1^\circ$ ). The slightly improved contact angle may result from the 2D morphology, and could prove useful for the construction of better superhydrophobic materials.

Hydrogen is an ideal and attractive candidate for industrial applications, such as ammonia synthesis, automobile fuels, and metallurgy, because of its abundance, renewability, and environmentally "clean" aspect.<sup>62</sup> Hydrogen isotherms for HCP and the GHCPs up to a maximum relative pressure ( $P/P_0$ ) of 0.99 at 77 K were examined, as shown in Fig. 4a. GHCP-1, GHCP-2, and GHCP-3 exhibited hydrogen capacities of 142, 130, and 122 cm<sup>3</sup> g<sup>-1</sup> (1.27, 1.16, and 1.09 wt%), respectively,

which are all higher than that of HCP (95 cm<sup>3</sup> g<sup>-1</sup>, 0.85 wt%). Obviously, the hydrogen capacities decrease along with the content of the graphene template. These results reveal that graphene contributes to the total hydrogen capacity, and underline the importance of a high graphene surface area for maximal hydrogen uptake capacity ( $S_{\text{BET}}$ : GHCP-1 > GHCP-2 > GHCP-3). Based on a variant of the Clausius–Clapeyron equation, the hydrogen isosteric heat of adsorption can be calculated (Fig. 4b).<sup>63</sup> At low coverage, GHCP-1, GHCP-2, and GHCP-3 exhibited the highest adsorption enthalpies at 6.3, 6.7, and 6.9 kJ mol<sup>-1</sup>, respectively (Fig. 4b), which are slightly larger than that of unadorned HCP (5.9 kJ mol<sup>-1</sup>). A more rapid decrease to 4.7 kJ mol<sup>-1</sup> for unadorned HCP was shown in the high hydrogen uptake region. The high heat of adsorption may stem from the narrower pores, which allow stronger overall interactions of the guest gas molecules because of additional interactions with the opposite walls.<sup>64</sup> The adsorption enthalpies of H<sub>2</sub> for the GHCPs increased with the increase in graphene content, indicating that the interaction between H<sub>2</sub> and the sample was increasingly dominating that between H<sub>2</sub> molecules because of the higher content of the graphene template.

In order to study the different gas capacities of the GHCPs, their CO<sub>2</sub> storage capacities were investigated. The CO<sub>2</sub> adsorption isotherms of HCP and the GHCPs at 273 K are presented in Fig. 4c. GHCP-1, GHCP-2, GHCP-3, and unadorned HCP exhibited CO<sub>2</sub> uptake of 9.74, 9.84, 9.06, and 8.80 wt%, respectively (see Table 1). It was found that adsorbed CO<sub>2</sub> was not significantly affected in comparison with H<sub>2</sub> when employing the graphene template, implying that the interactions between CO<sub>2</sub> molecules and samples were unaffected by the graphene template. Unadorned HCP showed CO<sub>2</sub> adsorptions of 44.8 cm<sup>3</sup> g<sup>-1</sup> (273 K) and 29.8 cm<sup>3</sup> g<sup>-1</sup> (298 K) at a relative pressure ( $P/P_0$ ) of 0.99. Under the same pressure conditions, the CO<sub>2</sub> adsorptions of GHCP-1 were 49.6 and 31.5 cm<sup>3</sup> g<sup>-1</sup> at 273 and 298 K, respectively. The isosteric enthalpies ( $Q_{\text{st}}$ ) for CO<sub>2</sub> were also calculated, and are shown in Fig. 4d. The CO<sub>2</sub>  $Q_{\text{st}}$  value at lower coverage mainly reflects the interaction strength between CO<sub>2</sub> and the sorbent.<sup>65</sup> As shown in Fig. 4d, the CO<sub>2</sub>  $Q_{\text{st}}$  values of HCP and the GHCPs were found to be almost the same in the range 0–8.0 cm<sup>3</sup> g<sup>-1</sup>, indicating that (1) the interactions between CO<sub>2</sub> and the porous samples are weak and (2) morphology and the graphene template have a limited impact on CO<sub>2</sub> uptake.<sup>66</sup>

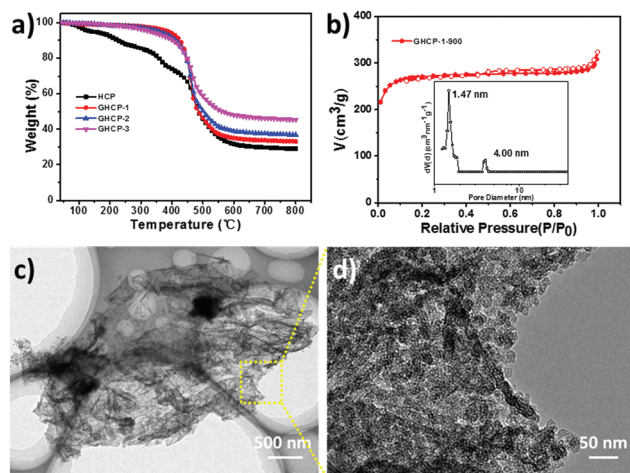
The thermal behaviors of unadorned HCP and the GHCPs were recorded by thermal gravimetric analysis (TGA) at a heating rate of 20 °C min<sup>-1</sup> under a nitrogen atmosphere (Fig. 5a). Unadorned HCP is thermally unstable below 400 °C. Remarkably, the weight loss for the GHCPs started as high as 300 °C, and became rapid only in the range between 400 and 600 °C, which might be attributed to the alkane structure and oxygen-containing groups in the materials. The improved thermal stability of the GHCPs may be attributable to the high thermal stability of the graphene sheets.<sup>40</sup> After 600 °C, the residual weight for all samples was almost constant, which suggests that the samples had been converted to stable carbon



**Fig. 4** Volumetric (a) H<sub>2</sub> and (c) CO<sub>2</sub> isotherms of unadorned HCP and the GHCPs at 77 K and 273 K, respectively. Isosteric heat of adsorption of (b) H<sub>2</sub> and (d) CO<sub>2</sub> for HCP and the GHCPs at different gas loadings.



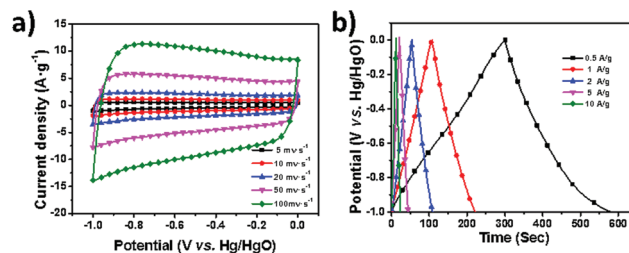




**Fig. 5** (a) TGA curves of HCP and the GHCPs at a heating rate of 20 °C min<sup>-1</sup> under nitrogen flow. (b) Nitrogen adsorption and desorption isotherms at 77 K (inset: the pore size distribution calculated by the NL-DFT method). (c, d) TEM images of GHCP-1-900.

structures. At 800 °C, the residual weights of HCP, GHCP-1, GHCP-2, and GHCP-3 were 28.9, 33.1, 36.8 and 45.4 wt%, respectively, which also indicates that the graphene templates in the GHCPs enhance the thermal stability. The porous nature of GHCP-1-900 was further studied using nitrogen physisorption measurements. It was found that the isotherms for GHCP-1-900 were type I (Fig. 5b). The BET and Langmuir surface areas of GHCP-1-900 were calculated as 871 and 1212 m<sup>2</sup> g<sup>-1</sup>, respectively. The pore size distribution of GHCP-1-900 is presented as an inset in Fig. 5b. The two peaks at 1.47 and 4.00 nm indicate that the hierarchical porous carbon nanosheets can be easily produced *via* the direct pyrolysis of 2D HCPs. These results also indicate that the high surface areas of 2D HCP-derived porous carbon nanosheets can be achieved by taking advantage of the 2D structure of graphene sheets. TEM images of GHCP-1-900 are shown in Fig. 5c and d: free-standing sheets with a morphology similar to that of graphene and several micrometers in size can be observed. It is clear that the 2D morphology with visible wrinkles can be well preserved after high temperature pyrolysis of GHCP-1. Interestingly, alternating light and dark regions can be easily seen in the enlarged TEM image, indicating the possible creation of a porous carbon during pyrolysis.

The high specific surface area and unique 2D nanosheet morphology as well as the porous structure of these newly prepared porous carbon nanosheets hold promise for their application in electrochemical capacitors. Therefore, the electrochemical capacitance of GHCP-1-900 was examined under alkaline conditions (6 M KOH) as a proof of concept. Typically, as shown in Fig. 6a, symmetric and horizontal cyclic voltammetry (CV) curves at a scan rate of 5–100 mV s<sup>-1</sup> were observed for GHCP-1-900, indicating ideal capacitive behavior. The capacitive performance was further investigated in galvanostatic charge/discharge cycling experiments (Fig. 6b). On the



**Fig. 6** (a) CV curves of GHCP-1-900 at 5–100 mV s<sup>-1</sup> in 6 M KOH aqueous solution. (b) Galvanostatic charge/discharge curves for GHCP-1-900 at a current density of 0.5–10 A g<sup>-1</sup>.

basis of the discharge curve, the specific capacitances of GHCP-1-900 were calculated to be 144, 116, 112, 127, and 122 F g<sup>-1</sup> at 0.5, 1, 2, 5, and 10 A g<sup>-1</sup>, respectively. For supercapacitor applications, the graphene layer in sandwich-type GHCP-1-900 can act as both a mini-current collector and a long-distance in-plane charge transporter during the charge and discharge processes, by taking advantage of the high electrical conductivity and the 2D electron transport properties of graphene. Additionally, inter-/intra-layer electrolyte transport also contributes to the superior capacitance, given that the active surface of nanosheets can be efficiently exposed to the electrolyte.<sup>31</sup>

## Conclusions

In summary, GHCPs based on vinylbenzyl chloride and divinylbenzene were successfully prepared by a RAFT emulsion polymerization approach and a FeCl<sub>3</sub>-promoted Friedel–Crafts reaction. A highly soluble trithiocarbonate functionalized GO, which is the key 2D template in this approach, was used as a 2D micro-CTA for RAFT polymerization. The resulting GHCPs exhibited superhydrophobic behavior (contact angle >150°) and a high specific surface area of up to 1224 m<sup>2</sup> g<sup>-1</sup>. They also exhibited an enhanced H<sub>2</sub> capacity (1.27 wt%) at 77 K and CO<sub>2</sub> capacity (9.9 wt%) at 273 K at a relative pressure of 0.99 in comparison with unadorned HCP without using the graphene template. The high specific surface areas and enhanced H<sub>2</sub> and CO<sub>2</sub> capacities might come from the contribution of the 2D graphene template and the 2D morphology. The GHCPs were also used to prepare porous carbon nanosheets without a template-removing operation. The as-prepared porous carbon nanosheets were successfully utilized in electrochemical capacitors, which makes this method a good approach for the preparation of 2D HCPs and 2D porous carbons with wide applications.

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## Notes and references

- 1 D. Wu, F. Xu, B. Sun, R. Fu, H. He and K. Matyjaszewski, *Chem. Rev.*, 2012, **112**, 3959.
- 2 M. S. Silverstein, N. R. Cameron and M. A. Hillmyer, *Porous Polymers*, John Wiley & Sons, 2011.
- 3 R. Dawson, A. I. Cooper and D. J. Adams, *Prog. Polym. Sci.*, 2012, **37**, 530.
- 4 A. Thomas, *Angew. Chem., Int. Ed.*, 2010, **49**, 8328.
- 5 J. W. Colson and W. R. Dichtel, *Nat. Chem.*, 2013, **5**, 453.
- 6 X. Feng, X. S. Ding and D. L. Jiang, *Chem. Soc. Rev.*, 2012, **41**, 6010.
- 7 L. Chen, Y. Honsho, S. Seki and D. L. Jiang, *J. Am. Chem. Soc.*, 2010, **132**, 6742.
- 8 X. M. Liu, Y. H. Xu and D. L. Jiang, *J. Am. Chem. Soc.*, 2012, **134**, 8738.
- 9 H. A. Patel, S. Hyun Je, J. Park, D. P. Chen, Y. Jung, C. T. Yavuz and A. Coskun, *Nat. Commun.*, 2013, **4**, 1357.
- 10 Y. Zhu, H. Long and W. Zhang, *Chem. Mater.*, 2013, **25**, 1630.
- 11 X. Wang, K. Maeda, A. Thomas, K. Takanabe, G. Xin, J. M. Carlsson, K. Domen and M. Antonietti, *Nat. Mater.*, 2009, **8**, 76.
- 12 R. Palkovits, M. Antonietti, P. Kuhn, A. Thomas and F. Schüth, *Angew. Chem., Int. Ed.*, 2009, **48**, 6909.
- 13 M. G. Schwab, B. Fassbender, H. W. Spiess, A. Thomas, X. Feng and K. Müllen, *J. Am. Chem. Soc.*, 2009, **131**, 7216.
- 14 Y. Kou, Y. Xu, Z. Guo and D. Jiang, *Angew. Chem., Int. Ed.*, 2011, **50**, 8753.
- 15 S. Xu, Y. Luo and B. Tan, *Macromol. Rapid Commun.*, 2013, **34**, 471.
- 16 M. Seo, S. Kim, J. Oh, S.-J. Kim and M. A. Hillmyer, *J. Am. Chem. Soc.*, 2014, **137**, 600.
- 17 F. S. Macintyre, D. C. Sherrington and L. Tetley, *Macromolecules*, 2006, **39**, 5381.
- 18 L. Pan, Q. Chen, J.-H. Zhu, J. Yu, Y. He and B.-H. Han, *Polym. Chem.*, 2015, **6**, 2478.
- 19 J. Germain, J. M. J. Frechet and F. Svec, *Chem. Commun.*, 2009, 1526.
- 20 J. Germain, J. Hradil, J. M. J. Frechet and F. Svec, *Chem. Mater.*, 2006, **18**, 4430.
- 21 A. P. Katsoulidis and M. G. Kanatzidis, *Chem. Mater.*, 2011, **23**, 1818.
- 22 X.-J. Chen, D. Ngoc Phuoc, J. Zhao, Y.-T. Wang, S.-P. Li and F. Svec, *J. Sep. Sci.*, 2012, **35**, 1502.
- 23 Y. Lv, Z. Lin and F. Svec, *Anal. Chem.*, 2012, **84**, 8457.
- 24 C. D. Wood, B. Tan, A. Trewin, F. Su, M. J. Rosseinsky, D. Bradshaw, Y. Sun, L. Zhou and A. I. Cooper, *Adv. Mater.*, 2008, **20**, 1916.
- 25 Y. Ouyang, H. Shi, R. Fu and D. Wu, *Sci. Rep.*, 2013, **3**, 1430.
- 26 K. S. Novoselov, A. K. Geim, S. V. Morozov, D. Jiang, Y. Zhang, S. V. Dubonos, I. V. Grigorieva and A. A. Firsov, *Science*, 2004, **306**, 666.
- 27 A. K. Geim and K. S. Novoselov, *Nat. Mater.*, 2007, **6**, 183.
- 28 X. Zhuang, F. Zhang, D. Wu, N. Forler, H. Liang, M. Wagner, D. Gehrig, M. R. Hansen, F. Laquai and X. Feng, *Angew. Chem., Int. Ed.*, 2013, **52**, 9668.
- 29 X. Zhuang, Y. Chen, L. Wang, K.-G. Neoh, E.-T. Kang and C. Wang, *Polym. Chem.*, 2014, **5**, 2010.
- 30 X. Zhuang, Y. Chen, G. Liu, P. P. Li, C. X. Zhu, E. T. Kang, K. G. Neoh, B. Zhang, J. H. Zhu and Y. X. Li, *Adv. Mater.*, 2010, **22**, 1731.
- 31 X. Zhuang, F. Zhang, D. Wu and X. Feng, *Adv. Mater.*, 2014, **26**, 3081.
- 32 C. A. Cao, X. Zhuang, Y. Su, Y. Zhang, F. Zhang, D. Wu and X. Feng, *Polym. Chem.*, 2014, **5**, 2057.
- 33 Y. Zhang, X. Zhuang, Y. Su, F. Zhang and X. Feng, *J. Mater. Chem. A*, 2014, **2**, 7742.
- 34 X. Huang, X. Qi, F. Boey and H. Zhang, *Chem. Soc. Rev.*, 2012, **41**, 666.
- 35 S. Eigler and A. Hirsch, *Angew. Chem., Int. Ed.*, 2014, **53**, 7720.
- 36 H. J. Salavagione, G. Martinez and G. Ellis, *Macromol. Rapid Commun.*, 2011, **32**, 1771.
- 37 M. Fang, K. Wang, H. Lu, Y. Yang and S. Nutt, *J. Mater. Chem.*, 2009, **19**, 7098.
- 38 S. H. Lee, D. R. Dreyer, J. An, A. Velamakanni, R. D. Piner, S. Park, Y. Zhu, S. O. Kim, C. W. Bielawski and R. S. Ruoff, *Macromol. Rapid Commun.*, 2010, **31**, 281.
- 39 H. M. Etmimi, M. P. Tonge and R. D. Sanderson, *J. Polym. Sci., Part A: Polym. Chem.*, 2011, **49**, 1621.
- 40 F. Beckert, C. Friedrich, R. Thomann and R. Mülhaupt, *Macromolecules*, 2012, **45**, 7083.
- 41 B. Zhang, Y. Chen, L. Xu, L. Zeng, Y. He, E. T. Kang and J. Zhang, *J. Polym. Sci., Part A: Polym. Chem.*, 2011, **49**, 2043.
- 42 Y. Li, X. Li, C. Dong, Y. Li, P. Jin and J. Qi, *Biosens. Bioelectron.*, 2009, **25**, 306.
- 43 Y. Li, X. Li, C. Dong, J. Qi and X. Han, *Carbon*, 2010, **48**, 3427.
- 44 Y. Yang, X. Song, L. Yuan, M. Li, J. Liu, R. Ji and H. Zhao, *J. Polym. Sci., Part A: Polym. Chem.*, 2012, **50**, 329.
- 45 Y. Yang, J. Wang, J. Zhang, J. Liu, X. Yang and H. Zhao, *Langmuir*, 2009, **25**, 11808.
- 46 W. S. Hummers and R. E. Offeman, *J. Am. Chem. Soc.*, 1958, **80**, 1339.
- 47 J. T. Lai, D. Filla and R. Shea, *Macromolecules*, 2002, **35**, 6754.
- 48 B. Zhang, J. Wang, Y. Chen, D. Fruchtl, B. Yu, X. D. Zhuang, N. He and W. J. Blau, *J. Polym. Sci., Part A: Polym. Chem.*, 2010, **48**, 3161.
- 49 J. H. Ahn, J. E. Jang, C. G. Oh, S. K. Ihm, J. Cortez and D. C. Sherrington, *Macromolecules*, 2006, **39**, 627.
- 50 C. D. Wood, B. Tan, A. Trewin, H. Niu, D. Bradshaw, M. J. Rosseinsky, Y. Z. Khimyak, N. L. Campbell, R. Kirk and E. Stöckel, *Chem. Mater.*, 2007, **19**, 2034.



- 51 S. Yang, X. Feng, L. Wang, K. Tang, J. Maier and K. Müllen, *Angew. Chem., Int. Ed.*, 2010, **49**, 4795.
- 52 B. Li, F. Su, H.-K. Luo, L. Liang and B. Tan, *Microporous Mesoporous Mater.*, 2011, **138**, 207.
- 53 X. Zhuang, Y. Mai, D. Wu, F. Zhang and X. Feng, *Adv. Mater.*, 2015, **27**, 403.
- 54 P.-G. Ren, D.-X. Yan, X. Ji, T. Chen and Z.-M. Li, *Nanotechnology*, 2011, **22**, 055705.
- 55 N. N. Chai, J. Zeng, K. G. Zhou, Y. L. Xie, H. X. Wang, H. L. Zhang, C. Xu, J. X. Zhu and Q. Y. Yan, *Chem. – Eur. J.*, 2013, **19**, 5948.
- 56 M. Xue, F. Li, J. Zhu, H. Song, M. Zhang and T. Cao, *Adv. Funct. Mater.*, 2012, **22**, 1284.
- 57 K. K. Sadasivuni, A. Saiter, N. Gautier, S. Thomas and Y. Grohens, *Colloid Polym. Sci.*, 2013, **291**, 1729.
- 58 K. Haubner, J. Murawski, P. Olk, L. M. Eng, C. Ziegler, B. Adolphi and E. Jaehne, *ChemPhysChem*, 2010, **11**, 2131.
- 59 B.-S. Ko, J.-Y. Sohn and J. Shin, *Polymer*, 2012, **53**, 4652.
- 60 J. Rouquerol, F. Rouquerol and K. S. Sing, *Absorption by Powders and Porous Solids*, Academic Press, 1999.
- 61 R. Pierotti and J. Rouquerol, *Pure Appl. Chem.*, 1985, **57**, 603.
- 62 Y. Luo, S. Zhang, Y. Ma, W. Wang and B. Tan, *Polym. Chem.*, 2013, **4**, 1126.
- 63 Z. Wang, S. W. Yuan, A. Mason, B. Repogle, D. J. Liu and L. P. Yu, *Macromolecules*, 2012, **45**, 7413.
- 64 W. G. Lu, D. Q. Yuan, D. Zhao, C. I. Schilling, O. Plietzsch, T. Muller, S. Brase, J. Guenther, J. Blumel, R. Krishna, Z. Li and H. C. Zhou, *Chem. Mater.*, 2010, **22**, 5964.
- 65 Y. Zhao, X. Liu, K. X. Yao, L. Zhao and Y. Han, *Chem. Mater.*, 2012, **24**, 4725.
- 66 M. G. Rabbani and H. M. El-Kaderi, *Chem. Mater.*, 2012, **24**, 1511.

