



Cite this: *Org. Biomol. Chem.*, 2015, **13**, 3341

A new multicomponent reaction: unexpected formation of derivatizable cyclic α -alkoxy isothioureas†

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An unexpected formation of cyclic α -alkoxy isothioureas has been achieved. As is known, the heterocyclic imines 2,5-dihydro-1,3-thiazoles are convertible to bisamides with the aid of a carboxylic acid and an isocyanide (Ugi reaction). Herein, it is shown that 2,5-dihydro-1,3-thiazole *S*-monoxides—the respective α -sulfinyl imines—are characterized by an altered reaction behavior. In a hitherto unknown multicomponent reaction the α -sulfinyl imines react with an isocyanide under acidic conditions in an alcoholic solution to the respective α -alkoxy isothioureas in good yields. In addition to the investigations on this unexpected synthesis the regioselectivity of the acylation of the synthesized compounds is described. A rearrangement, which is accelerated by EDC and HOBt, between both possible regioisomers was found.

Received 15th December 2014,
Accepted 16th January 2015

DOI: 10.1039/c4ob02608j

www.rsc.org/obc

Introduction

For more than 60 years the 2,5-dihydro-1,3-thiazole (3-thiazoline) scaffold has been an important industrial key element in the course of the preparation of pharmaceutically active molecules.¹ Furthermore, other remarkable investigations on this family of heterocycles characterize fields of research in general.² The synthesis of 3-thiazolines was first reported by F. Asinger in 1956.³ Following a modified protocol of the Asinger reaction the cyclic imines are preparable by a four-component reaction (A 4-CR) where an α -chloro aldehyde is reacted with a carbonyl compound, ammonia, and sodium hydrosulfide.⁴

3-Thiazolines could be converted into a large number of different products by functionalizing their reactive C=N double bond.⁵

One of the most important reactions involving imines is the Ugi three-component reaction (U 3-CR). In the U 3-CR an imine is treated with an isocyanide and a carboxylic acid to

form the desired bisamide.⁶ As our earlier research had shown, 3-thiazolines are unproblematically employable as the imine component in the U 3-CR.⁷

In one of our recent studies, we had additionally investigated the chemoselective oxidation of 3-thiazolines.⁸ An oxidation of the sulfur atom leading to 3-thiazoline *S*-monoxides represents one possible and realizable reaction pathway of the oxidation of these substrates.

In order to examine the reaction behaviour of these new 3-thiazoline *S*-monoxides, we then tried to convert these compounds to the respective bisamides with the aid of the U 3-CR (Fig. 1). However, as presented in this paper, we observed that under the reaction conditions mentioned above, the 3-thiazoline *S*-monoxides react contrary to our expectations to cyclic isothioureas. The products are characterized by an alkoxy moiety next to the endocyclic nitrogen atom of the isothiurea (Fig. 2). Despite the existence of many syntheses of cyclic isothioureas,⁹ the class of cyclic α -alkoxy isothioureas has not yet been reported before.

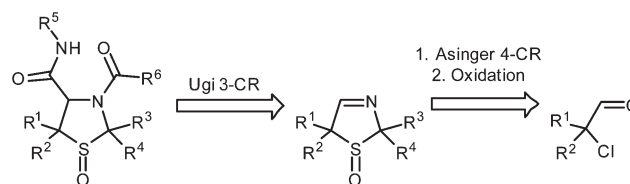


Fig. 1 Retrosynthetic consideration of the designed synthesis route.

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† Electronic supplementary information (ESI) available: Experimental procedures and analysis data for all isolated new compounds. ¹H and ¹³C NMR spectra of all isolated new compounds as well as ¹H NMR spectra (stacked) of the monitored rearrangement of **5a** and **5d**. CCDC 1037548. For ESI and crystallographic data in CIF or other electronic format see DOI: 10.1039/c4ob02608j



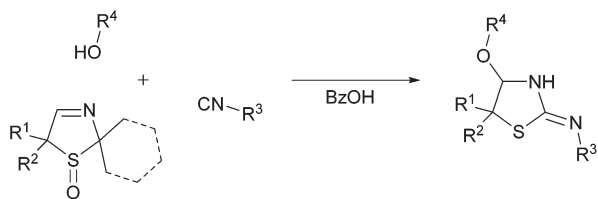


Fig. 2 Unexpected three-component synthesis of cyclic isothiureas.

Results and discussion

Synthesis of the precursors

First, we synthesized four known 3-thiazolines *via* the A 4-CR. Treatment of the respective α -chloro aldehyde with sodium hydrosulfide, an aqueous ammonia solution and acetone or cyclohexanone led to the cyclic imines 2,2,5,5-tetramethyl-2,5-dihydro-1,3-thiazole (**1a**),^{5c} 5,5-diethyl-2,2-dimethyl-2,5-dihydro-1,3-thiazole (**1b**),^{7b} 2,2-dimethyl-1-thia-3-azaspiro[4.5]dec-3-ene (**1c**),¹⁰ and 2,2-dimethyl-1-thia-4-azaspiro[4.5]dec-3-ene (**1d**).^{5c}

The oxidation of 3-thiazolines **1** to sulfoxides **2** was described by us recently.⁸ In line with this, all four 3-thiazolines **1a–d** were converted to the respective sulfoxides **2a–d** with the aid of *m*CPBA (Fig. 3).

Unexpected preparation of the cyclic isothiureas

On the basis of our earlier work,⁷ we wanted to examine the influence of the sulfoxide moiety on the reactivity of the cyclic imine by converting the 3-thiazoline *S*-monoxides **2** to bisamides. To validate the applicability of unoxidized 3-thiazolines **1** in the U 3-CR, we reacted the 3-thiazoline **1a** with 1-chloro-4-isocyanobenzene and benzoic acid in MeOH. The bisamide **3** was formed according to our expectations. Afterwards, the oxidized analog of **1a** (3-thiazoline *S*-monoxide **2a**) was converted analogously. It was found that after 72 h the substrate **2a** was

fully reacted. However, it came as a surprise that no bisamide was formed and the corresponding cyclic α -methoxy isothiurea **4a** was isolated as the sole product (Scheme 1). This unexpected result implies that the reaction of the sulfoxide **2a** is accompanied by a formal elimination of acetone and a cyclization realized by an addition of the isocyanide. Moreover, an addition of MeOH, which was initially applied as a solvent, to the imine bond takes place. Apart from using NMR, IR, and mass analysis, the structure of the product **4a** was also verified by X-ray crystal structure analysis.¹¹ The crystal structure reveals that the double bond is exocyclic and (*Z*)-configured. Contrary to our expectations, even at -40 °C in CDCl₃ only the tautomer (and only one diastereomer) bearing an exocyclic double bond was observable, so that, consequently, a tautomeric equilibrium could not be proven.

To explore the influence of the carboxylic acid, which is not a part of the product, on the formation of the cyclic isothiureas **4**, we examined the transformation of 3-thiazoline **2a** with 1-chloro-4-isocyanobenzene to isothiurea **4a** under different conditions by varying the type and the equivalent of the acid (Table 1). The entries 1–4 show that an acidic medium is significant for the course of the reaction and that, therefore, the amount of acid has an enormous effect on the obtained yield. The best result was received by using 1.0 equivalent of the acid (Table 1, entry 1). The acidity of benzoic acid in comparison with acetic acid, trifluoroacetic acid, and sulfuric acid seems to be optimal. Thus, the use of benzoic acid led to the highest yield. A further increase of the amount of acid did not lead to an enhanced formation of the product **4a** (Table 1, entry 4).

The observable formation of **4a** in the absence of a carboxylic acid or an inorganic acid could probably be explained by the acidity of methanol. Furthermore, benzoic acid was found to be the most effective acid to convert the 3-thiazoline **2a** in the presence of an isocyanide and MeOH (Table 1, entries 1, 2, and 4) in comparison with acetic acid (Table 1, entry 5), trifluoroacetic acid (Table 1, entry 6), and sulfuric acid (Table 1, entry 7).

The proposed mechanism for the formation of the cyclic isothiureas **4** is shown in Scheme 2. The sulfoxide **2** is protonated by the carboxylic acid. A ring-opening which is induced by a nucleophilic attack of the alcohol leads to the sulfenic acid **A**. The iminium ion **B** is formed by a substitution of the hydroxyl group by the isocyanide and a subsequent ring-closure. This ion reacts with two molecules of the alcohol under elimination of a proton to the product **4**. The acetal, formed as a co-product, is hydrolyzed to the alcohol and acetone.

Reaction scope

Next, we investigated the generality of this unexpected synthesis of cyclic α -alkoxy isothiureas **4** by varying all three substrates (Table 2). First, the applicability of different isocyanides was studied. It was found that aryl, benzyl and alkyl isocyanides could be used for the conversion of 3-thiazolines **2**. Substituents including aryl chlorides (Table 2, entries 1 and 10) and esters (Table 2, entries 5, 6, 9, and 13–15) were tolerated. Furthermore, it was determined that different alcohols, upon

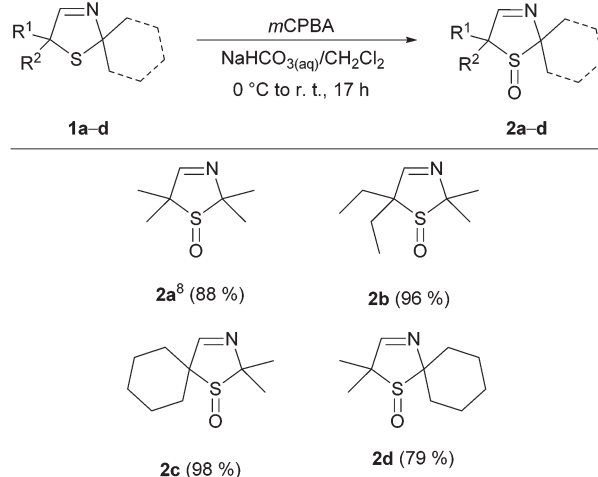
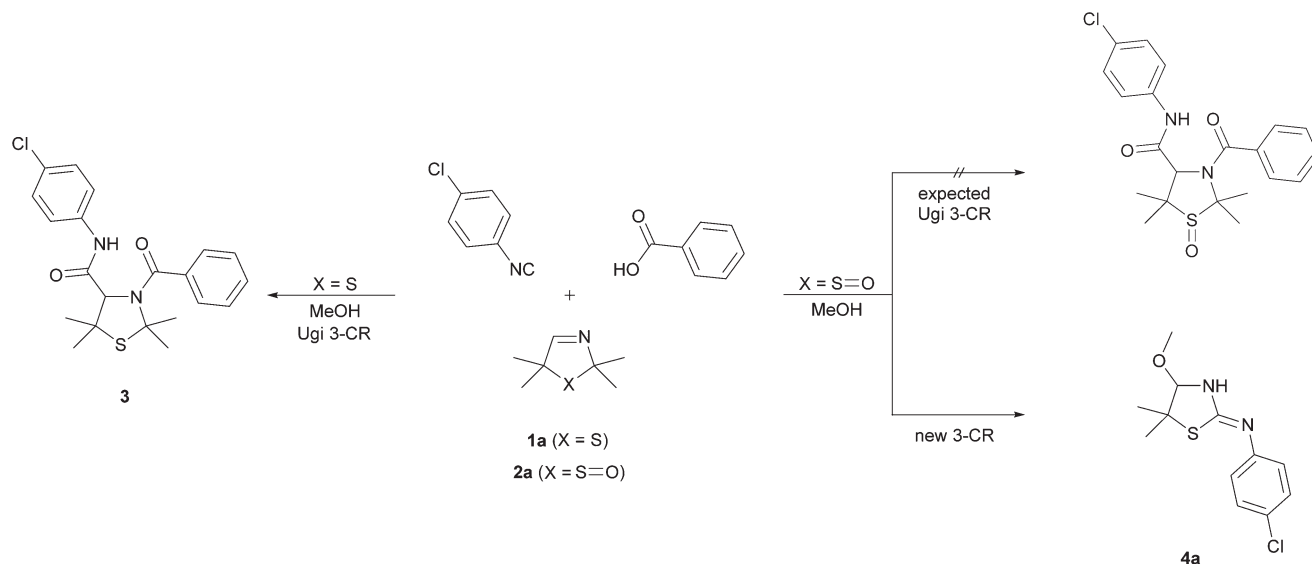


Fig. 3 Synthesis of sulfoxides **2a–d**.





Scheme 1 Conversion of **1a** and **2a** under the conditions of the Ugi reaction.

Table 1 Influence of the acid on the formation of the cyclic isothiourea **4a**^a

Entry	Acid	Yield ^b [%]
1	BzOH (1.0 equiv.)	74
2	BzOH (0.5 equiv.)	36
3	—	19
4	BzOH (1.2 equiv.)	73
5	AcOH (1.0 equiv.)	27
6	TFA (1.0 equiv.)	—
7	H ₂ SO ₄ (1.0 equiv.)	—

^a Reaction conditions: **2a** (0.8 mmol), 1-chloro-4-isocyanobenzene (0.8 mmol), acid, in MeOH (5 mL), 72 h at r.t. ^b Isolated yield after chromatography on silica gel.

the condition that they can be used as a solvent, are employable reactants (Table 2, entries 7 and 8).

Besides model substrate **2a**, three further 3-thiazoline *S*-monoxides **2** were examined (Table 2, entries 9–16). It turned out that compound **2c** bearing a spiro-coupled cyclohexane ring at the carbon C⁵ gave the highest yield (up to 87%; Table 2, entry 13). Entry 16 proves that not only 3-thiazoline *S*-monoxides **2a–c** characterized by two methyl groups located at the acetal carbon, but also 3-thiazoline *S*-monoxide **2d** exhibiting a spiro-coupled cyclohexane ring at that carbon atom is an

appropriate substrate. When the substrate **2d** is used, cyclohexanone instead of acetone is the co-product.

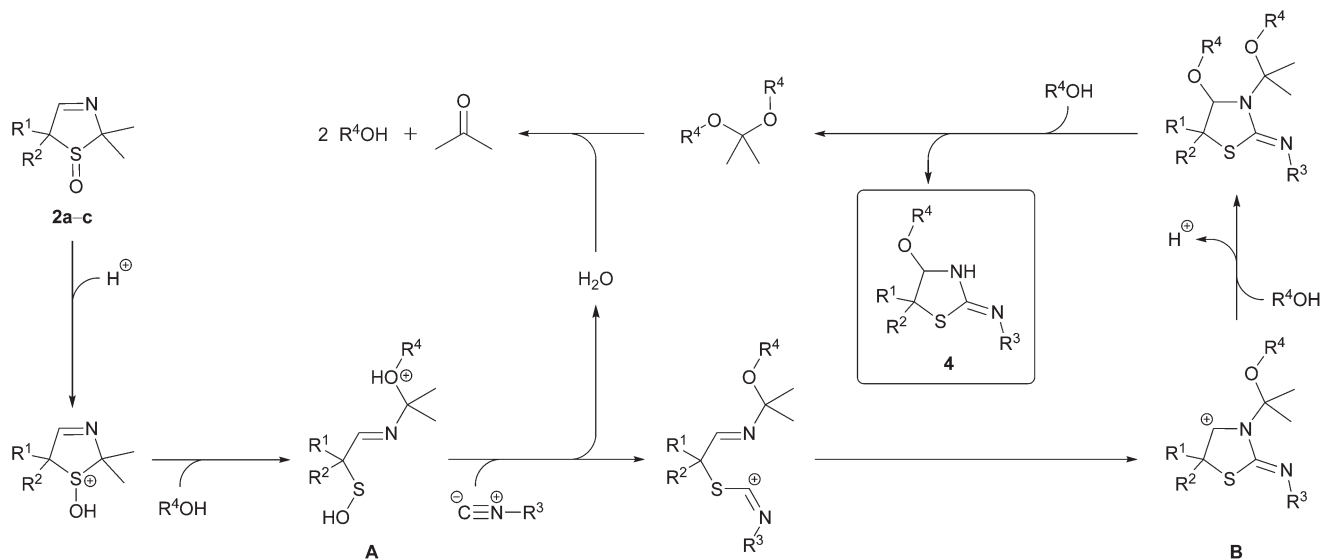
It would therefore appear that the unexpectedly observed formation of the cyclic isothiourea **4a** can be transferred to a broad range of differently substituted substrates. Additionally, the obtained products are suitable for further derivatizations.

Acylation of the synthesized isothioureas

They can be converted to *N*-acyl isothioureas by treating them with the respective acid and EDC¹² [1-ethyl-3-(3-dimethylaminopropyl)carbodiimide] as an activating agent in CH₂Cl₂. Following this route, both possible regioisomers—products **5** acylated at the exocyclic nitrogen atom and products **6** acylated at the endocyclic nitrogen atom—can be observed. Conversion of isothiourea **4j** bearing an aromatic substituent at the exocyclic nitrogen atom with phenylacetic acid resulted in an excess of product **5a**, which is derivatized at the exocyclic nitrogen atom (Table 3, entry 1), (**5a**:**6a** 75:25) in the crude product. Changing the solvent from CH₂Cl₂ to DMF or the reaction temperature from room temperature to –20 °C did not have any influence on the ratio between both regioisomers in the crude product.

After isolation by column chromatography, the ratio was shifted to **5a**:**6a** 27:73. To explain this shift of the ratio between **5a** and **6a** during the isolation, we postulated that a rearrangement of product **5**, once formed by an exocyclic acylation, to **6** is possible. The existence of this rearrangement was proven by a control experiment. Isolated **5a** was kept in CDCl₃ solution at room temperature without convection. As determined by ¹H NMR analysis, the rearrangement to **6a** was fully completed after 24 d (Fig. S1, ESI†). In a repetition of this experiment using additional EDC the rearrangement was already completed after 9 d (Fig. S2, ESI†). Thus, the rearrangement between both regioisomers is accelerated by EDC.





Scheme 2 Postulated mechanism.

Table 2 Substrate scope for the synthesis of cyclic isothioureas **4**^a

Entry	Substrate	R ¹	R ²	R ³	R ⁴	Yield ^b [%]
1	2a	Me	Me	4-Cl-Ph	Me	4a : 74
2	2a	Me	Me	Ph	Me	4b : 65
3	2a	Me	Me	Bn	Me	4c : 69
4	2a	Me	Me	Allyl	Me	4d : 55
5	2a	Me	Me	CH ₂ COOMe	Me	4e : 73
6	2a	Me	Me	(CH ₂) ₃ COOEt	Me	4f : 55
7	2a	Me	Me	4-Cl-Ph	Et	4g : 61
8	2a	Me	Me	Allyl	Allyl	4h : 77
9	2b	Et	Et	CH ₂ COOMe	Me	4i : 63
10	2c	-(CH ₂) ₅ -		4-Cl-Ph	Me	4j : 81
11	2c	-(CH ₂) ₅ -		Allyl	Me	4k : 73
12	2c	-(CH ₂) ₅ -		<i>n</i> -Bu	Me	4l : 71
13	2c	-(CH ₂) ₅ -		CH ₂ COOMe	Me	4m : 87
14	2c	-(CH ₂) ₅ -		(CH ₂) ₂ COOMe	Me	4n : 64
15	2c	-(CH ₂) ₅ -		(CH ₂) ₃ COOMe	Me	4o : 68
16	2d	Me	Me	4-Cl-Ph	Et	4g : 34

^a Reaction conditions: **2** (1.0 equiv.), carboxylic acid (1.0 equiv.), BzOH (1.0 equiv.), in R⁴OH (6 mL per mmol **2**), 72 h at r.t. ^b Isolated yield after chromatography on silica gel.

By using hydroxybenzotriazole (HOBT), which is mainly used as the additional activating agent to suppress the racemization of chiral substrates during peptide syntheses,¹³ in addition to EDC, we improved the outcome of the acylation reaction. Moreover, the use of HOBT in the conversion of **4j** with phenylacetic acid led to the formation of only one regioisomer (**6a**) after 2 h (Table 3, entry 2). In line with this, the

rearrangement of isolated **5a** in CDCl₃ under the addition of EDC and HOBT took less than 20 min (Fig. S3, ESI[†]). This clarifies that the rearrangement between both regioisomers is accelerated more strongly by a combination of EDC and HOBT than by the mere use of EDC. Entries 3 and 4 of Table 3 verify the regioselective acylation of cyclic isothioureas **4** bearing an aromatic substrate at the exocyclic nitrogen atom within the reaction time, when HOBT is added to the reaction mixture.

A conversion of the isothiourea **4l** characterized by an aliphatic substituent at the exocyclic nitrogen atom led mainly to product **6d** (**5d**: **6d** 40:60 in crude product; Table 3, entry 5; coupling agent: exclusive EDC).

The rearrangement of **5d** is very slow. Monitoring the rearrangement of isolated **5d** in CDCl₃ at room temperature without the influence of EDC and HOBT we ascertained that only half the amount of substrate **5d** was rearranged to **6d** after 115 d (Fig. S4, ESI[†]). In the conversion of **4l**, the addition of HOBT only had an influence on the overall yield (raised from 63% to 85%; Table 3, entry 6) and not on the product ratio. These observations were confirmed by the conversion of substrate **4n**, which is also characterized by an aliphatic substituent at the exocyclic nitrogen atom, with Fmoc-*N*-phenylglycine (Table 3, entries 7 and 8).

When a substrate is converted which bears an aromatic substituent at the exocyclic nitrogen atom the acylation at this exocyclic nitrogen atom is kinetically more preferred than the acylation at the endocyclic nitrogen atom. This is probably due to the mesomeric stabilization of the cationic intermediate stage. On the other hand, the thermodynamic stability of the product which is characterized by an exocyclic double bond and an acyl moiety at the endocyclic nitrogen atom in comparison to the regioisomer is enhanced by the aromatic substituent, because an exocyclic double bond results in a conjugated π -system. This results in the observed selectivities and speeds of rearrangements.



Table 3 Acylation of cyclic isothioureas 4

Entry	Substrate	Method ^a	R ¹	R ²	R ³	R ⁴	R ⁵	Ratio ^b 5 : 6	Yield ^c [%]
1	4j	A	-(CH ₂) ₅ -		4-Cl-Ph	Me	Bn	75 : 25	5a: 20; 6a: 55
2	4j	B	-(CH ₂) ₅ -		4-Cl-Ph	Me	Bn	0 : 100	5a: —; 6a: 83
3	4a	B	Me	Me	4-Cl-Ph	Me	Bn	0 : 100	5b: —; 6b: 88
4	4a	B	Me	Me	4-Cl-Ph	Me	Et	0 : 100	5c: —; 6c: 78
5	4l	A	-(CH ₂) ₅ -		<i>n</i> -Bu	Me	Bn	40 : 60	5d: 22; 6d: 41
6	4l	B	-(CH ₂) ₅ -		<i>n</i> -Bu	Me	Bn	40 : 60	5d: 28; 6d: 56
7	4n	A	-(CH ₂) ₅ -		(CH ₂) ₂ COOMe	Me	CH ₂ N(Ph)Fmoc	41 : 59	5e: 18; 6e: 21
8	4n	B	-(CH ₂) ₅ -		(CH ₂) ₂ COOMe	Me	CH ₂ N(Ph)Fmoc	43 : 57	5e: 23; 6e: 28

^a Reaction conditions: (A) 4 (1.0 equiv.), carboxylic acid (1.2 equiv.), EDC (1.3 equiv.), in CH₂Cl₂ (8 mL per mmol 4), 10 h at r.t.; (B) carboxylic acid (1.0 equiv.), carboxylic acid (1.2 equiv.), EDC (1.3 equiv.), HOBT (1.3 equiv.), in CH₂Cl₂ (8 mL per mmol 4), 2 h at r.t.; all reaction times were optimized with a view to the highest overall yield. ^b Ratio was determined by ¹H NMR spectroscopy of the crude product. ^c Isolated yield after chromatography on silica gel.

Conclusions

In summary, 3-thiazoline *S*-monoxides, which are easily preparable by an Asinger 4-CR and a subsequent oxidation, were converted to cyclic α -alkoxy isothioureas. The investigated conversion of the substrates with isocyanides and alcohols under acidic conditions opens unexpected access to the novel class of cyclic α -alkoxy isothioureas in good yields. This procedure may be useful for a general conversion of *S,N*-acetals to the corresponding isothioureas.

The derivatization of the synthesized cyclic isothioureas by an acylation leads to two different regioisomeric products. The kinetic product, which is acylated at the exocyclic nitrogen atom, rearranges to the thermodynamic more stable second regioisomer, which is acylated at the endocyclic nitrogen atom. The selectivity and the speed of the rearrangement depend tremendously on the type of the substituent. EDC and HOBT were found to accelerate the rearrangement of *N*-acyl isothioureas.

Experimental

Synthesis of (*RS*)-(Z)-2-(4-chlorophenylimino)-4-methoxy-5,5-dimethyl-1,3-thiazolidine (4a) as a representative example for the synthesis of all cyclic isothioureas 4

The sulfoxide 2a (717 mg, 4.50 mmol) was dissolved in 9 mL methanol. 1-Chloro-4-isocyanobenzene (619 mg, 4.50 mmol), dissolved in 9 mL methanol, and benzoic acid (550 mg, 4.50 mmol), dissolved in 9 mL methanol, were added dropwise. The reaction mixture was vigorously stirred for 72 h at r.t. The solvent was removed on a rotary evaporator. Purification of the crude product by column chromatography on silica gel (EtOAc, *R_f* = 0.63) afforded 4a (902 mg, 74%) as a colorless solid, mp 115–117 °C (from CH₂Cl₂-*n*-hexane); IR (ATR):

$\tilde{\nu}$ 3111, 3051, 2960, 2897, 2827, 1636, 1584, 1484, 1437, 1368, 1341, 1158, 1144, 1063, 843, 824 cm⁻¹; ¹H NMR (500.1 MHz, CDCl₃): δ 1.46, 1.52 [6 H, 2s, C(CH₃)₂], 3.29 (3 H, s, OCH₃), 4.34 (1 H, s, NCH), 6.94–6.95 [2 H, m, 2 *m*-CH_{Ar}(Cl)], 7.22–7.24 [2 H, m, 2 *o*-CH_{Ar}(Cl)] ppm; ¹³C NMR (125.8 MHz, CDCl₃): δ 22.61, 30.66 [C(CH₃)₂], 55.30 (OCH₃), 57.04 [C(CH₃)₂], 95.10 (NCH), 123.48 [2 *m*-CH_{Ar}(Cl)], 128.88 (ClC_{Ar}), 129.12 [2 *o*-CH_{Ar}(Cl)], 148.41 (NC_{Ar}), 162.82 (C=N) ppm; MS (CI-sector): *m/z* 271.2 (M + H⁺, 100%); HRMS (CI-sector): found 271.0673; calc. for C₁₂H₁₆ClN₂OS [M + H]⁺ 271.0672.

Single crystals of 4a were recrystallized from CH₂Cl₂ and *n*-hexane, mounted in inert oil and transferred to the cold gas stream of the diffractometer.

Crystal structure determination of 4a

Crystal data. C₁₂H₁₅ClN₂OS, *M* = 270.77, triclinic, *a* = 6.1948(3), *b* = 10.9643(5), *c* = 11.2530(5) Å, *U* = 675.19(5) Å³, *T* = 120 K, space group *P* $\bar{1}$, *Z* = 2, 6165 reflections measured, 6165 unique (*R*_{int} = 0.0000) which were used in all calculations. The final *wR*(*F*₂) was 0.1262 (all data).

Acknowledgements

The authors gratefully acknowledge Marc Schmidtman for X-ray crystallography.

Notes and references

- W. Keim and H. Offermanns, *Angew. Chem.*, 2007, **119**, 6116–6120, (*Angew. Chem. Int. Ed.*, 2007, **46**, 6010–6013).
- (a) A. Dömling and I. Ugi, *Angew. Chem.*, 1993, **105**, 634, (*Angew. Chem. Int. Ed. Engl.*, 1993, **32**, 563); (b) J. S. Elmore,



- D. S. Mottram, M. Enser and J. D. Wood, *J. Agric. Food Chem.*, 1997, **45**, 3603–3607; (c) A. Dömling, I. Ugi and E. Herdtweck, *Acta Chem. Scand.*, 1998, **52**, 107; (d) X. Fernandez, R. Fellous, L. Lizzani-Cuvelier, M. Loiseau and E. Duñach, *Tetrahedron Lett.*, 2001, **42**, 1519–1521.
- 3 F. Asinger, *Angew. Chem.*, 1956, **68**, 413.
- 4 J. Martens, H. Offermanns and P. Scherberich, *Angew. Chem.*, 1981, **93**, 680–683, (*Angew. Chem. Int. Ed. Engl.*, 1981, **20**, 668).
- 5 For examples, see: (a) K. Harada, in *The chemistry of the carbon-nitrogen double bond*, ed. S. Patai, John Wiley & Sons, London, 1992, ch. 6, pp. 255–298; (b) H. Gröger, Y. Saida, H. Sasai, K. Yamaguchi, J. Martens and M. Shibasaki, *J. Am. Chem. Soc.*, 1998, **120**, 3089–3103; (c) R. Sail, J. Kopf and P. Köll, *Tetrahedron: Asymmetry*, 2000, **11**, 423–433; (d) T. Stalling, F. Brockmeyer, D. Kröger and A. Schwäblein, *Z. Naturforsch., B: Chem. Sci.*, 2012, **67**, 1045–1055; (e) A. Sehlinger, T. Stalling, J. Martens and M. A. R. Meier, *Macromol. Chem. Phys.*, 2014, **215**, 412–420.
- 6 (a) I. Ugi, R. Meyer, U. Fetzer and C. Steinbrückner, *Angew. Chem.*, 1959, **71**, 386; (b) I. Ugi, *Angew. Chem.*, 1962, **74**, 9–22, (*Angew. Chem. Int. Ed. Engl.*, 1962, **1**, 8–21); (c) I. Ugi and E. Wischhöfer, *Chem. Ber.*, 1962, **95**, 136–140.
- 7 For examples, see: (a) J. Kintscher and J. Martens, *Synthesis*, 1992, 837–838; (b) M. Hatam, D. Tehranfar and J. Martens, *Synth. Commun.*, 1995, **25**, 1677–1688; (c) F. Brockmeyer, D. Kröger, T. Stalling, P. Ullrich and J. Martens, *Helv. Chim. Acta*, 2012, **95**, 1857–1870.
- 8 F. Brockmeyer and J. Martens, *ChemSusChem*, 2014, **7**, 2441–2444.
- 9 (a) A. Hirashima, H. Tarui and M. Eto, *Biosci., Biotechnol., Biochem.*, 1994, **58**, 1206–1209; (b) R. F. Ambartsumova, M. G. Levkovich, E. G. Milgrom and N. D. Abdullaev, *Chem. Heterocycl. Compd.*, 1997, **33**, 112–117; (c) B. R. Dixon, C. M. Bagi, C. R. Brennan, D. R. Brittelli, W. H. Bullock, J. Chen, W. L. Collibee, R. Dally, J. S. Johnson, H. C. E. Kluender, W. F. Lathrop, P. Liu, C. A. Mase, A. M. Redman, W. J. Scott, K. Urbahns and D. J. Wolanin, *US Pat*, 6353006 B1, 2002; (d) H. Kai, Y. Morioki, T. Murashi, K. Morita, S. Shinonome, H. Nakazato, K. Kawamoto, K. Hanasaki, F. Takahashi, S. i. Mihara, T. Arai, K. Abe, H. Okabe, T. Baba, T. Yoshikawa and H. Takenaka, *Bioorg. Med. Chem. Lett.*, 2007, **17**, 4030–4034; (e) N. Lui, K. Jelich, M. Littmann, K. Lorenz and B. Lecorre, *WO Pat*, 2009033583 A1, 2009.
- 10 K. Drauz, H. G. Koban, J. Martens and W. Schwarze, *Liebigs Ann. Chem.*, 1985, 448–452.
- 11 The X-ray crystal structure is shown in the ESI.† CCDC 1037548 contains the supplementary crystallographic data.
- 12 EDC is a common coupling agent. For another example, see: M. J. Niphakis, B. J. Turunen and G. I. Georg, *J. Org. Chem.*, 2010, **75**, 6793–6805.
- 13 W. König and R. Geiger, *Chem. Ber.*, 1997, **130**, 788–798.

