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Fluorosolvatochromism of furanyl- and thiophenyl-substituted acetophenones†

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A series of *para*-substituted acetophenones bearing a furanyl or a thiophenyl moiety show a large Stokes-shift, which is a function of various solvent properties. Photophysical properties such as emission lifetime of the compounds have been determined using time-correlated-single photon counting to secure the intrinsic fluorescence behaviour. The solvent dependent position of the UV/Vis emission band $\tilde{\nu}_{\max,em}$ of the compounds has been measured in 26 various solvents. The influence of the solvent on $\tilde{\nu}_{\max,em}$ is of very complex nature and mathematically analysed by multiple square linear solvation energy (LSE)-correlation analysis using Catalán's four-solvent parameter set. Solvent acidity has a strong influence on the bathochromic shift of 2,5-disubstituted furan derivatives compared to the non-5-substituted furan and thiophene derivatives, which show a contrary behaviour. Therefore, the 5-cyanofuranyl-substituted acetophenone derivative is useful as a probe for measuring environmental properties by fluorescence spectroscopy.

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Introduction

In recent years, there has been growing interest in organic fluorescent dyes because of their potential applications as fluorescent probes in chemistry, biochemistry and medicine, and in optoelectronics including organic light emitting diodes (OLEDs), sensors, and lasers.^{1–3} Such organic fluorophores can also act as environmental probes, *e.g.* acid–base, and solvent-sensitive indicators.⁴

In previous work, we have studied and discussed many solvatochromic dyes, which are suitable to measure environmental effects by shift of the position of the UV/Vis absorption band as a function of polarity of the environment. For this purpose, *push–pull* substituted catechol dyes,^{5a} azobenzene chromophores being linked at polyvinylamine backbones,^{5b} and specific aminobenzofuran-dione dyes as basicity probes for ionic liquids^{5c} were used.

Interactions of solvatochromic dyes with pure solvents or solvent mixtures are determined by a combination of many effects.⁶ In order to separate the effects of specific interactions (hydrogen-bonding capacity) from non-specific interactions including

electrostatic effects (dipolarity/polarizability) on the UV/Vis band ($\tilde{\nu}_{\max}$) the concept of linear free solvation energy relationship (LSER) is used.⁶

During the last few decades, multiple polarity effects of various environments have been considered in terms of the Kamlet–Abboud–Taft approach (eqn (1)), which has been established as a suitable tool for this purpose.

$$\tilde{\nu}_{\max} = \tilde{\nu}_{\max,0} + a \cdot \alpha + b \cdot \beta + s \cdot \pi^* \quad (1)$$

Kamlet, Abboud, and Taft used basically three different polarity parameters which describe the overall solvent polarity: acidity α (hydrogen-bond donor capacity), basicity β (hydrogen-bond acceptor capacity) and dipolarity/polarizability π^* .^{6a,b,7} $\tilde{\nu}_{\max}$ is the longest wavelength UV/Vis maxima of the compound measured in a particular solvent; $\tilde{\nu}_{\max,0}$ corresponds to a standard process referenced to cyclohexane as a nonpolar medium; a , b , and s are solvent-independent correlation coefficients that reflect the contribution of each parameter to the overall solute–solvent interactions.

Unfortunately, the Kamlet–Abboud–Taft concept does not distinguish between dipolarity and polarizability effects both of which are involved in a single π^* -parameter. Therefore, the improved solvent polarity parameter scales of Catalán (eqn (2)) should be used which allow dipolarity and polarizability effects of a solvent to be considered independently of each other.⁸

$$\tilde{\nu}_{\max} = \tilde{\nu}_{\max,0} + a \cdot SA + b \cdot SB + d \cdot SP + e \cdot SdP \quad (2)$$

$\tilde{\nu}_{\max}$ describes the solvent-dependent physico-chemical properties in a particular solvent and $\tilde{\nu}_{\max,0}$ the solute

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properties of a reference system, which is defined for the gas phase. The SA (solvent acidity) scale^{8a,b,e} has been established as a measure of the hydrogen-bond donor (HBD) and electron pair acceptor (EPA) properties, whereas the SB (solvent basicity) scale^{8c,e} is responsible for the hydrogen-bond acceptor (HBA) and electron pair donor (EPD) properties. The nonspecific interactions polarizability and dipolarity are reflected by the parameters SP^{8d,e} and SdP,^{8e} respectively. *a*, *b*, *d*, and *e* are the correlation coefficients characterising the sensitivity of the property $\tilde{\nu}_{\max}$ to different solute–solvent interaction mechanisms.

Fluorosolvatochromism, compared to the comprehensively reported solvatochromism of *push–pull* substituted chromophores, is still not widely established for measuring environmental effects because only a few probes are known which show a distinct function of $\tilde{\nu}_{\max,em}$ as a function of Kamlet–Abboud–Taft or Catalán solvent parameter. Allegeable solvent effects on the shift of $\tilde{\nu}_{\max,em}$ indicated as fluorosolvatochromism are reported for a few dyes and interpreted in detail concerning specific effects.^{8e,9} So far, multiple square correlation analyses of $\tilde{\nu}_{\max,em}$ as a function of various solvent properties are still not established taking into account a numerous set of solvents.

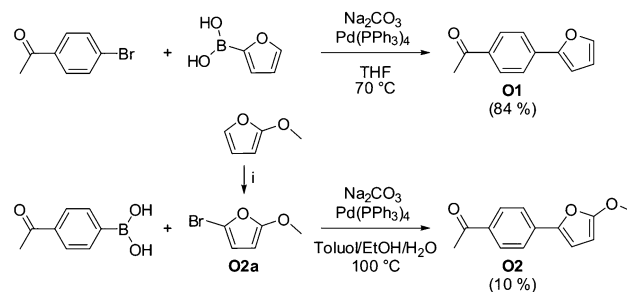
In this work, we present four acetophenone derivatives bearing the heterocycles furan or thiophene in the *para*-position to the acetyl unit as a promising class of fluorosolvatochromic compounds. Cross-coupling reactions such as Heck, Stille, Suzuki, Sonogashira, and Negishi type reactions provide an easy route to link carbon atoms of different aryl compounds. Especially by means of the application of the nucleophilicity–electrophilicity principle, those reactions are well manageable.¹⁰ A nucleophilic (metal) aryl compound reacts with an aryl halide under transition metal catalysis. Palladium has particularly proven to be the transition metal of choice. It bears an extraordinary potential as an activation agent for oxidative addition, insertion, reductive elimination and β -hydride elimination.¹¹

The main focus of the work concentrates on the synthesis, photophysical properties and fluorosolvatochromism¹² of the acetophenone derivatives. Furan (*para* electrophilic substituent constant $\sigma_p^+ = -0.39$)¹³ and thiophene ($\sigma_p^+ = -0.43$)¹³ are well-known electron-donating heterocycles that are widely used to lower the energy band of conjugated materials. These heterocycles are indeed weaker donor substituents as the methoxy group ($\sigma_p^+ \text{OCH}_3 = -0.78$),¹³ but they offer the option to enhance the length of the chromophoric π -electron system. This implicates a decrease in the energetic HOMO–LUMO gap and results in a shift towards an even lower-energy absorption range.

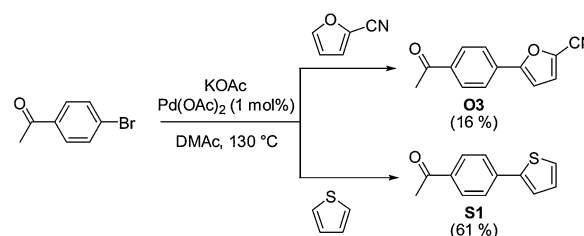
Results and discussion

Synthesis of furanyl- or thiophenyl-substituted acetophenones

In this work, the palladium-catalysed Suzuki and Heck coupling served as a synthesis tool for the C–C bond linkage between



Scheme 1 Synthesis of the functionalized acetophenones **O1** and **O2** using the palladium-catalysed Suzuki coupling. Reaction conditions for (i): *N*-bromosuccinimide, toluene, cat. *p*-toluenesulfonic acid, 70 °C, 4 h.



Scheme 2 Synthesis of the functionalized acetophenones **O3** and **S1** using the palladium-catalysed Heck coupling.

acetophenone and a furan or thiophene moiety. By modifying the furan part with electron donor (–OMe) or acceptor (–CN) moieties it was possible to achieve a structure–property relationships.

In the Suzuki reaction, an aryl halide (in the case of **O1**: 4-bromoacetophenone, **O2**: 2-bromo-5-methoxyfuran) was reacted with organoboron compounds (**O1**: 2-furanboronic acid, **O2**: 4-acetylphenylboronic acid) in the presence of base and tetrakis(triphenylphosphine) palladium(0) (Scheme 1).

However, the functionalized acetophenones **O3** and **S1**¹⁴ were prepared by Heck coupling. Analogous to ref. 14, 2-cyanofuran was converted with the electron-deficient 4-bromoacetophenone to compound **O3** (Scheme 2).

Theoretical calculations

DFT theoretical studies have been carried out to gain an insight into the electronic structure of acetophenones. All geometries have been optimized using density functional theory with the B3-LYP functional¹⁵ in combination with the def2-TZVPP basis set.¹⁶ The calculated HOMO–LUMO energy gaps of compounds **O1**, **O2**, **O3**, and **S1** (see ESI†, Table S1) present values of the HOMO ranging between 4.045 eV for **O1** and 3.737 eV for **O2**. A comparison of the furanyl- or thiophenyl-substituted acetophenone derivatives with acetophenones bearing electron-donating groups (*e.g.* –OCH₃) shows a significant decrease of the energy gap between HOMO and LUMO. The linear dependency between the calculated HOMO–LUMO energy gaps and the measured UV/Vis absorption maxima $\tilde{\nu}_{\max}$ (Fig. 1) indicates that UV/Vis absorption of acetophenones under investigation is mainly justified by a HOMO–LUMO transition.



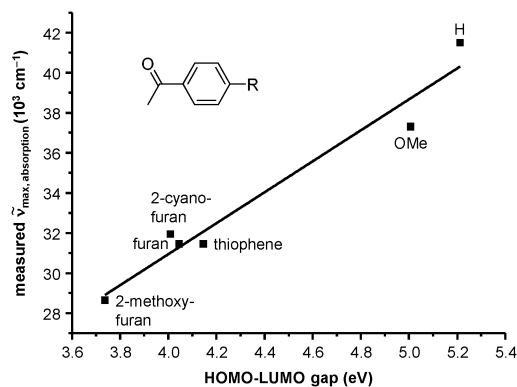


Fig. 1 Experimentally measured UV/Vis absorption maxima of acetophenone derivatives (in DCM) as a function of the theoretically calculated HOMO–LUMO energy differences.

Solvent effects on the UV/Vis absorption spectra

Optical properties such as UV/Vis absorption and fluorescence of functionalized acetophenones **O1**, **O2**, **O3**, and **S1** were investigated both in solution and in the solid state. UV/Vis absorption and emission maxima of the acetophenone derivatives in a variety of solvents are given in the ESI† (Table S3). Fig. 2 exemplarily shows the UV/Vis absorption and emission spectra of the respective compounds recorded in four selected solvents.

All compounds show the shortest wavelength UV/Vis absorption band in *n*-hexane. The compounds **O1** and **S1** exhibit the longest wavelength UV/Vis absorption band in acetone, while **O3** shows the strongest bathochromic shift in hexamethylphosphoramide (HMPA). Methoxy-functionalized acetophenone **O2** presents the longest wavelength UV/Vis absorption band in alcohols such as 1-butanol or 1-propanol. These UV/Vis shifts correspond to a positive solvatochromism with a solvatochromic range of $\Delta\tilde{\nu}(\mathbf{O1}) = 2140 \text{ cm}^{-1}$, $\Delta\tilde{\nu}(\mathbf{O2}) = 1340 \text{ cm}^{-1}$, $\Delta\tilde{\nu}(\mathbf{O3}) = 2350 \text{ cm}^{-1}$, and $\Delta\tilde{\nu}(\mathbf{S1}) = 2050 \text{ cm}^{-1}$. Considering the furan derivatives, it is obvious that the solvatochromic range increases with increasing electron-withdrawing character of the substituent in the 5-position ($\text{OMe} < \text{H} < \text{CN}$).

Furthermore, it is quite evident that an enlargement of the conjugated π -system causes a bathochromic shift of the absorption wavelength as a consequence. The UV/Vis absorption maxima of the furan- and thiophene-functionalized acetophenones **O1** and **S1** are each detected at 318 nm in DCM (non-HBA/HBD-solvent). A bathochromic shift of 5900 cm^{-1} is observed in relation to the UV/Vis absorption maximum of 4-methoxy-acetophenone ($\lambda_{\text{max}} = 268 \text{ nm}$). A stronger *push–pull* system is generated by the additional introduction of a methoxy group on the furan ring (**O2**). In comparison to the unsubstituted heterocyclic compounds ($\lambda_{\text{max}} = 318 \text{ nm}$), this result causes an enormous bathochromic shift of the UV/Vis absorption maximum at 349 nm, measured in DCM. However, the electron-withdrawing substituent $-\text{CN}$ in **O3** effects a slight hypsochromic shift ($\lambda_{\text{max}} = 313 \text{ nm}$, in DCM).

Additionally, the UV/Vis spectroscopic behaviour of **O1**, **O2**, **O3**, and **S1** was investigated in the solid-state. The UV/Vis

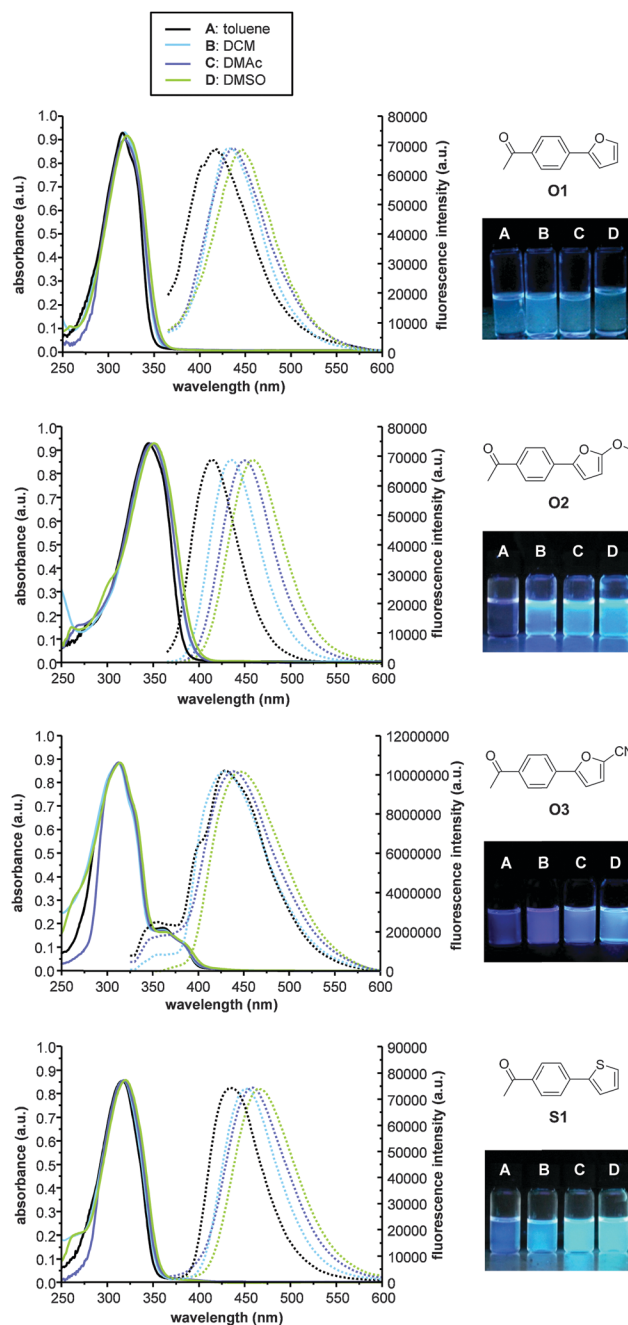


Fig. 2 Normalized UV/Vis absorption and emission spectra of **O1**, **O2**, **O3**, and **S1** measured in toluene (A), DCM (B), DMAc (C), and DMSO (D); $\lambda_{\text{exc}} = 310 \text{ nm}$ (**O1**, **O3**, and **S1**), $\lambda_{\text{exc}} = 350 \text{ nm}$ (**O2**).

absorption and emission spectra of the solids are shown in Fig. 3. Generally, UV/Vis absorption bands are non-symmetric and very broad whereby the band broadening of the compounds with substituents in the 5-position of the furan ring (**O2** and **O3**) is significantly more pronounced.

Solvent effects on the UV/Vis emission spectra

In contrast to the UV/Vis absorption maxima, the UV/Vis emission maxima of the functionalized acetophenones **O1**,



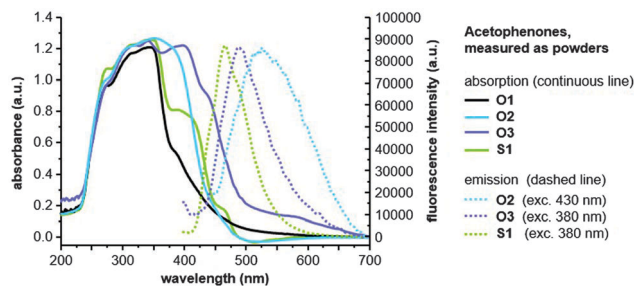


Fig. 3 UV/Vis absorption and emission spectra of **O1**, **O2**, **O3**, and **S1** as powders.

O2, **O3**, and **S1** show a pronounced shift as a function of solvent properties, as illustrated in Fig. 2.

The shortest wavelength UV/Vis emission band was detected in 2-propanol ($\lambda_{\text{max}}(\mathbf{O1}) = 405$ nm), tetrachloromethane (CCl_4 , $\lambda_{\text{max}}(\mathbf{O2}) = 398$ nm, $\lambda_{\text{max}}(\mathbf{O3}) = 419$ nm) and triethylamine ($\lambda_{\text{max}}(\mathbf{S1}) = 404$ nm). The compounds **O2** and **O3** bearing substituents on the furan ring exhibit the longest wavelength emission band in alcohols ($\lambda_{\text{max}}(\mathbf{O2}) = 479$ nm in MeOH, $\lambda_{\text{max}}(\mathbf{O3}) = 483$ nm in 2,2,2-trifluoroethanol (TFE)), whereas the furan/thiophene-functionalized acetophenones show the longest wavelength emission band in the solvents γ -butyrolactone and *N,N*-dimethylformamide (DMF, $\lambda_{\text{max}}(\mathbf{O1}) = 467$ nm) as well as in dimethyl sulfoxide (DMSO, $\lambda_{\text{max}}(\mathbf{S1}) = 466$ nm). In this context, the respective UV/Vis shifts correspond to a positive fluorosolvatochromism, but the solvatochromic range of $\Delta\tilde{\nu}(\mathbf{O1}) = 3280$ cm^{-1} , $\Delta\tilde{\nu}(\mathbf{O2}) = 4260$ cm^{-1} , $\Delta\tilde{\nu}(\mathbf{O3}) = 3170$ cm^{-1} and $\Delta\tilde{\nu}(\mathbf{S1}) = 3290$ cm^{-1} is considerably larger.

The UV/Vis spectroscopic results (absorption and fluorescence) of individual compounds demonstrate that **O1**, **O3** and **S1** absorb at shorter wavelengths compared to **O2**, but the fluorescence is observed at similar wavelengths. Thus, an exceptionally large Stokes shift (in DCM: $\Delta\tilde{\nu}(\mathbf{O1}) = 8360$ cm^{-1} , $\Delta\tilde{\nu}(\mathbf{O2}) = 5700$ cm^{-1} , $\Delta\tilde{\nu}(\mathbf{O3}) = 8590$ cm^{-1} , $\Delta\tilde{\nu}(\mathbf{S1}) = 9280$ cm^{-1}) has been obtained (see ESI,† Table S3).

The Stokes shifts of these furanyl- or thiophenyl-substituted acetophenones can be evaluated as disproportionately large compared to other fluorescent dyes, *e.g.* coumarin derivatives ($\Delta\tilde{\nu}$ of 4200–5300 cm^{-1}).¹⁷ However, acetophenone bearing a pyrene function in the *para*-position shows a comparably large Stokes shift in alcohol solutions ($\Delta\tilde{\nu} \sim 10\,000$ cm^{-1}) for **S1**.¹⁸ One of the reasons for the large Stokes shift is certainly that the functionalized acetophenones have a higher dipole moment in the first excited state S_1 than in the ground state S_0 . Associated with this property a change in the dipole moment in strength and orientation is induced. The solvent molecules respond with a reorganization of the solvent shell around the excited acetophenone molecules. Hence, the state S_1 of the acetophenone molecule is better solvated by polar solvents and as a result its energetical stabilization is improved. The electronic state after emission $S_{0,\text{emission}}$ is energetically unfavourably solvated due to the shorter timescale of photon emission in comparison to solvent molecule reorganization. After relaxation of the solvent

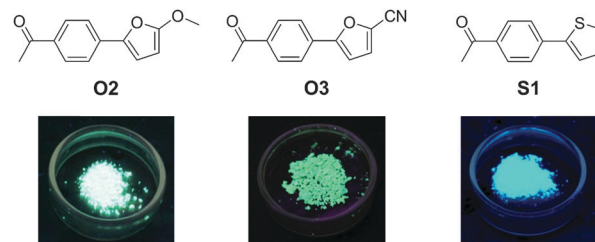


Fig. 4 Photographs of the solid-state fluorescence of acetophenones **O2**, **O3**, and **S1**, excited at 380 nm (**O3**, **S1**) and 430 nm (**O2**).

molecules the equilibrium state S_0 is reached. The difference in the energies of $S_{0,\text{emission}}$ and S_0 adds to a strongly red-shifted fluorescence of the acetophenone systems.

Methoxyfuran-, cyanofuran-, and thiophene-functionalized acetophenones (**O2**, **O3**, and **S1**) also demonstrate a pronounced solid-state fluorescence (Fig. 4). However, compound **O1** shows no fluorescence in the solid-state.

These observations indicate that the substitution in the 5-position of the furan ring or the exchange of oxygen with sulphur plays an important role in tuning the photochemical properties of these types of acetophenones. Fluorescence spectroscopic measurements of the functionalized acetophenones **O2**, **O3**, and **S1** in the solid-state are also shown in Fig. 3. The UV/Vis emission maxima of methoxyfuran-substituted acetophenone (**O2**) is located at 524 nm which is red-shifted by $\Delta\tilde{\nu} = 1500$ cm^{-1} compared to its cyanofuran-substituted analogue **O3**. Compound **S1** is fluorescent with a UV/Vis emission band at 465 nm. It is evident that the solid-state fluorescence of methoxyfuran-functionalized acetophenone (**O2**) is strongly red-shifted by $\Delta\tilde{\nu} = 1800$ cm^{-1} in comparison to the lowest-energy emission band measured in polar solvents (Fig. 2 and 3). The UV/Vis emission maxima of the solids **O3** and **S1** are similar to those in solution (polar solvents, *e.g.* $\lambda_{\text{max}}(\mathbf{O3}) = 483$ nm in TFE, $\lambda_{\text{max}}(\mathbf{S1}) = 466$ nm in DMSO).

Determination of excited-state lifetimes

The fluorescence decay profiles of the furanyl- or thiophenyl-functionalized acetophenones **O1**, **O2**, **O3**, and **S1** were measured by time-correlated single photon counting (TCSPC, Fig. 5). The emission lifetimes are in the order of 0.2 to 2 ns (Table 1) and support the assignment of the emission to be due to fluorescence. For **O2** and **O3** the fluorescence decay-curves were found to be bi-exponential, while for the other substances exponential decays were observed. There is no unique relationship between solvent polarity and emission lifetime. Nevertheless, it is noted that for **O1** and **O2** the lifetime in DCM is longer than in ethanol (EtOH), while the opposite holds true for **O3** and **S1**. This points to significant differences in the electronic structures of the excited states and the impact of specific solute–solvent interactions, *e.g.* hydrogen-bonding ability.

Linear solvation energy correlation analyses

The influence of the solvent behaviour on the UV/Vis shift of the absorption and emission maxima has been determined by



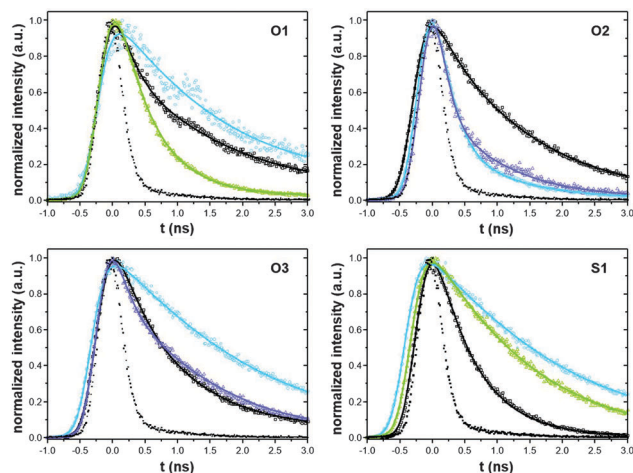


Fig. 5 Emission decay curves (hollow symbols) obtained via TCSPC ($\lambda_{\text{exc}} = 350$ nm) for **O1**, **O2**, **O3**, and **S1** in different solvents: DCM (—), EtOH (—), DMF (—) and CCl_4 (—). The solid lines represent the respective fit curves (see the Experimental section). The solid symbols represent the instrumental response function.

Table 1 Emission lifetimes and fit parameters of **O1**, **O2**, **O3**, and **S1** in selected solvents

Compd	τ (ns)			
	DCM	EtOH	CCl_4	DMF
O1	1.4	0.5	—	0.2 (65%) 1.5 (35%)
O2	0.4 (49%) 1.4 (51%)	0.2 (93%) 1.2 (7%)	< 0.2 (88%) 0.9 (12%)	—
O3	0.6 (86%) 2.0 (14%)	1.9	0.2 (65%) 1.3 (35%)	—
S1	0.5	1.8	—	1.3

means of multiple linear regression analyses using the solvent polarity parameter set of Catalán.¹⁹ The solvatochromism of the furanyl- or thiophenyl-functionalized acetophenones was studied in 26 solvents with dipolarities/polarizabilities and hydrogen-bonding capacities.

However, the HBD/HBA abilities of the environment have only a marginal influence upon the UV/Vis absorption maxima of the solvatochromic probe.

The results of multiple linear regression analysis of the solvent dependent emission band of substituted acetophenones **O1**, **O2**, **O3**, and **S1**, which are qualitatively the best according to the solvent scale of Catalán, are compiled in Table 2. The correlation coefficients r are greater than 0.9 (with the exception of **O1**) for the

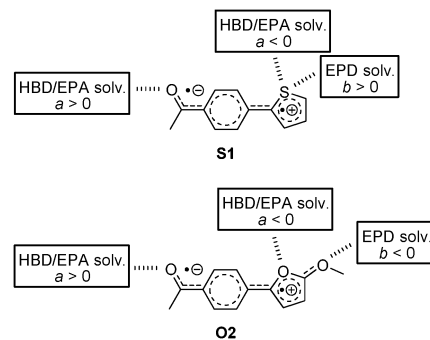


Fig. 6 Possible hydrogen-bond (HB) and electron pair (EP) interactions between solute and solvents, and their influence on the emission as demonstrated for the compounds **S1** and **O2**, both are shown in an assumed excited state.

LSE relationships. These results indicate a very high validity of the obtained multiparameter equations and allow significant conclusions on the fluorosolvatochromic behaviour.

In this context, the solvent independent correlation coefficient a provides information about solvent interactions with the hydrogen-bond acceptor/electron pair donor fragment (*i.e.* lone electron pair of the partially negatively charged oxygen from the keto group) of the respective dye. However, the coefficient b describes specific solvation of the dye by the hydrogen-bond donor/electron pair acceptor group using appropriate HBA/EPD-solvents. Fig. 6 shows, for instance, the assumed electron distribution of compounds **S1** and **O2** in the excited state with possible specific solute-solvent interactions which can have an influence on the fluorosolvatochromic behaviour.

A positive sign indicates that the *push-pull* system is decreased by the solute-solvent interactions and this results in a hypsochromic shift. Consequently, a negative sign is associated with a bathochromic shift. The appearance of the latter is associated with an enhancement of the *push-pull* character of the dye.

Surprisingly, the calculated LSER of the solvent dependent emissions of furan/thiophene-functionalized acetophenones (**O1**, **S1**) show that the solvent independent correlation coefficients a and b are positive. In this case protic solvents, which can act as hydrogen-bond donors, mainly interact with the negatively charged oxygen atom of the keto group ($a > 0$).

As a result, the (+M)-effect of the O^- group is decreased and also the *push-pull* character of the aromatic system is weakened.

A weakening of the *push-pull* system can also be observed by interactions of EPD-solvents with acetophenones **O1** and **S1**.

Table 2 Solvent-independent correlation coefficients a , b , d and e of the Catalán parameters SA, SB, SP and SdP, solute property of the reference system $\tilde{\nu}_{\text{max},0}$ gas phase, number of solvents (n), correlation coefficient (r), standard deviation (sd), and significance (f) of the fluorosolvatochromism of the acetophenone-functionalized compounds **O1**, **O2**, **O3**, and **S1**

Entry	Compound	$\tilde{\nu}_{\text{max},0}$ (10^3 cm^{-1})	a	b	d	e	n	r	sd	f
2	O1	26.926	1.198	1.936	−4.204	−2.847	23	0.841	0.591	$\leq 1.175 \times 10^{-4}$
3	O2	24.787	−2.768	−0.993	—	−1.887	25	0.960	0.351	$\leq 9.269 \times 10^{-12}$
4a	O3	23.111	−2.808	—	—	—	31	0.920	0.360	$\leq 2.353 \times 10^{-13}$
4b	O3	22.984	−2.619	—	—	—	12	0.870	0.258	$\leq 9.735 \times 10^{-6}$
5	S1	25.466	1.958	1.917	−3.322	−2.202	24	0.929	0.349	$\leq 6.051 \times 10^{-8}$



Thereby, EPD-solvents interact predominantly with the oxygen or sulphur of the heterocyclic ring due to its positive charge located in the excited state ($b > 0$). The strongest effect of the fluorosolvatochromic behaviour of **O1** and **S1** is exerted by the polarizability (SP) and the dipolarity (SdP) of the solvents. The respective coefficients d and e have a negative value and thus indicate positive fluorosolvatochromism. Here, a bathochromic shift of the emission maxima is observed with increasing polarizability and dipolarity of the solvent. Considering the emission spectra of the functionalized acetophenones **O1**, **O2**, **O3**, and **S1** (Fig. 2) reveal that the compounds **O1** and **S1** show a smaller difference between the emission maxima of non-polar solvents (*i.e.* toluene) and those of polar solvents (*i.e.* DMSO). One possible explanation for these observation could be that the terms of the nonspecific interactions (d and e) exhibit negative values, but the coefficients a and b are positive. The sum of all solvent interactions results in a bathochromic shift, but it is not as large as the shift of acetophenone **O2**, as can be seen in the following interpretation.

In contrast to **O1** and **S1**, the functionalized acetophenone **O2** shows a negative sign of parameter a , which indicates that the interaction of HBD/EPA-solvents with the lone electron pairs of oxygen atoms of the furan ring is significantly stronger than those with oxygen of the keto group ($a < 0$). Therefore, electron density is donated by the furan ring to the respective solvent molecules. This effect induces a reinforcement of the *push-pull* system and finally a bathochromic shift is caused. The correlation coefficient b of **O2** obtained from the LSE correlation analyses also illustrates a bathochromic shift, but it is less pronounced. Furthermore, the correlation of **O2** shows that the absolute e value has a lower influence than the coefficient a . This result demonstrates that the UV/Vis emission spectrum is weakly dependent on the change in the dipolarity of the chromophore environments rather than on the capacity of the solvent to act as a hydrogen-bond donor. Surprisingly, **O2** shows no influence with regard to the polarizability (coefficient d) of the solvent. Fig. 7 exemplarily shows the linear relationship

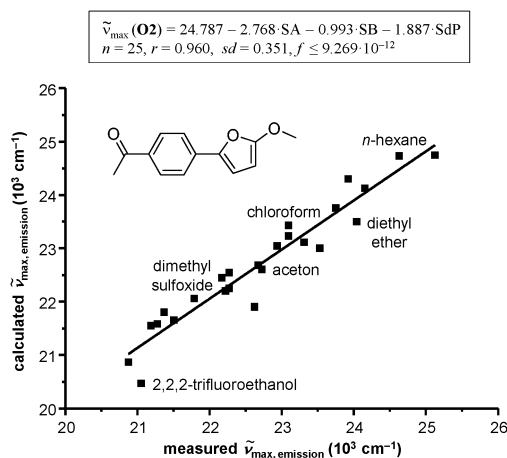


Fig. 7 Linear solvation energy relationship for the fluorescence of **O2** in 25 various solvents. Plots of measured vs. calculated emission maxima, according to the Catalán equation are shown.

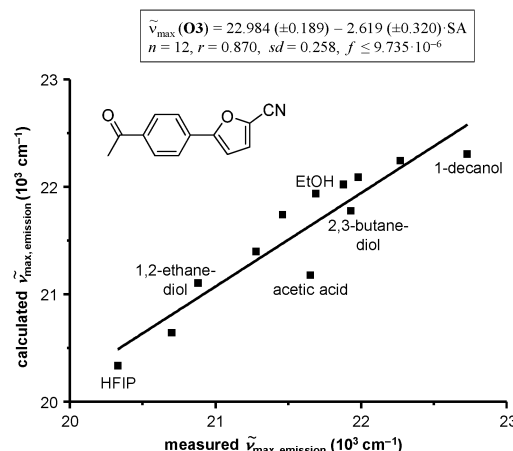


Fig. 8 Linear solvation energy relationship between calculated and measured fluorescence $\nu_{\max,em}$ for **O3** in 12 solvents with hydrogen-bond donating ability.

between the measured and calculated wavenumbers for the solvent dependent UV/Vis emission band of acetophenone **O2**.

In comparison to the methoxyfuran-functionalized acetophenone, the result of multiple linear regression analysis of a CN analogue **O3** shows that its SA sensitivity is of highest significance as shown by the equation in Fig. 8. Therefore, the UV/Vis emission maxima of **O3** were described using the single-parameter equation (Table 2, entry 4a) and show an excellent correlation.

Fig. 8 displays the plot of the calculated UV/Vis maxima for the emission as a function of the corresponding experimental values for only classic HBD-solvents.

Conclusions

In this work, four furanyl- or thiophenyl-functionalized acetophenones have been produced by means of palladium-catalysed cross-coupling reactions such as Heck or Suzuki coupling. Optical properties (UV/Vis absorption and emission) were been studied both in solution and in the solid-state. The represented results show that the UV/Vis absorption spectra of the substituted acetophenones **O1**, **O2**, **O3**, and **S1** do not noticeably shift as the solvent is varied. In contrast, the emission spectra exhibit a strong solvent-dependence.

Due to the dependence of the emission maxima on the nature of the solvent, the fluorosolvatochromic behaviour of the furanyl- or thiophenyl-substituted acetophenones was studied in solvents of different acidity, basicity and polarity. The results of LSER correlations using the Catalán parameter set suggest that the influence of solvent dipolarity/polarizability on the long wavelength emission maximum for **O1** and **S1** is predominant. Furthermore, the influence on the fluorosolvatochromic behaviour of **O2** is reflected by negative signs of a , b and e coefficients, whereby the strongest is exerted by the acidity of the solvent. A single-parameter equation according to Catalán was established for compound **O3**. Accordingly, the fluorosolvatochromic



dye **O3** can be used as a HBD/EPA-sensitive probe for several environments (solvent and solid-state).

Acetophenones **O2**, **O3**, and **S1** fluoresce in solution as well as in the solid-state. However, the solid-state fluorescence of **O1** is not observed.

Experimental section

Materials

Organic solvents were dried according to standard methods in an argon atmosphere and freshly distilled prior to use. All commercial reagents were used without further purification. They were purchased from the following suppliers: Acros: 4-bromoacetophenone (98%), 2-cyanofuran (99%), thiophene (>99%), potassium acetate (97%), and *p*-toluenesulfonic acid (99%); Alfa Aesar: 2-methoxyfuran (97%); ChemPur: tetrakis-(triphenylphosphine) palladium(0) (99%); Fluka: palladium(II) acetate (47% Pd); Fluorochem: 4-acetylphenylboronic acid (98%) and furan-2-boronic acid (98%); Merck: *N*-bromosuccinimide (99%).

Instrumentation

The melting points were measured on a hot-stage microscope using a digital temperature measuring system from Wagner & Munz and are uncorrected. Differential scanning calorimetry (DSC) measurements were accomplished using a Mettler Toledo, type DSC 30. Elemental analyses were determined using a Vario-EL analyser. Attenuated total reflection infrared (ATR IR) spectra were recorded using a Golden Gate ATR accessory on a Perkin-Elmer Fourier transform 1000 spectrometer at room temperature in the wavenumber range 400 to 4000 cm⁻¹. ¹H NMR (250 MHz) and ¹³C NMR (69.9 MHz) spectroscopy were carried out on a Bruker Avance 250 NMR spectrometer. The residue signals of the solvents were used as internal standards. High-resolution mass spectra were recorded on a Bruker Daltonik microTOF-QII spectrometer. UV/Vis absorption spectra were obtained on a MCS 400 diode array UV/Vis spectrometer from Carl Zeiss, Jena, connected *via* glass-fibre optics. Fluorescence spectra were recorded on a Fluoromax[®]-4 Horiba Jobin Yvon spectrometer. Excitation and emission slits were set to 5 nm. The excitation wavelength was varied between 310 nm (**O1**, **O3**, and **S1**) and 350 nm (**O2**). Correlation analysis: multiple regression analysis was performed using the Origin Pro 9.0G statistics program. The emission decay curves were obtained *via* time-correlated-single-photon-counting. After excitation with a frequency-doubled Ti-sapphire laser adjusted to 700 nm (Tsunami, Newport Spectra-Physics GmbH, pulse-to-pulse repetition rate 800 kHz after passing a pulse selector, model 3980, Newport Spectra-Physics GmbH), *i.e.* at $\lambda_{\text{exc}} = 350$ nm, the luminescence of the sample was collected in a 90°-geometry and detected using a Becker & Hickl PMC-100-4 photon-counting module. Two long-pass filters (355 nm and 395 nm) are inserted into the detection beam path. The instrumental response function (IRF) is measured using a scattering sample (P25 nanoparticles in water). The measured curves are

fitted mono- or bi-exponentially including the convolution with the IRF trace using a home-written algorithm in Scilab.²⁰

Methods

All geometries have been optimized using density functional theory with the B3-LYP functional¹⁵ in combination with the def2-TZVPP basis set.¹⁶ The geometries have been converged to a residual gradient norm of 10⁻³ a.u. or better. The energy was converged to at least 10⁻⁶ a.u., and fine quadrature grids have been used to calculate the exchange and correlation potential (size 5).²¹ All minima have been characterized by normal-mode analysis. All calculations have been performed using the TURBOMOLE program package (for a current version, see <http://www.turbomole.de>).²²

Syntheses

4-(Furan-2-yl)acetophenone (O1). 4-Bromoacetophenone (3.23 g, 16.25 mmol), furan-2-boronic acid (2.00 g, 17.87 mmol), Pd(PPh₃)₄ (1.88 g, 16.25 mmol) and 65 mL of aqueous Na₂CO₃ solution (1.0 mol L⁻¹) were mixed in 120 mL of THF (degassed with argon). The reaction mixture was refluxed under argon for 24 h. To this solution 200 mL of HCl (0.1 mol L⁻¹) was added and the aqueous phase was extracted with DCM (2 × 100 mL). The organic phase was washed with distilled water (3 × 100 mL), dried over MgSO₄, filtered and concentrated. The crude product was purified by column chromatography on silica gel (eluent: DCM) to obtain **O1** (2.54 g, 84%) as a yellow to orange-coloured solid.

Mp: 96–98 °C. DSC (25–350 °C, 10 K min⁻¹, N₂): 100 °C (*endo*), 230 °C (*endo*). Found: C, 77.2; H, 5.4. Calc. for C₁₂H₁₀O₂: C, 77.4; H, 5.4%. $\nu_{\text{max}}/\text{cm}^{-1}$ 3105 (ArCH), 1666 (C=O), 1602 (C=C), 1560 (C=C; furan), 1263 (def vib; 1,4-disubstituted benzene), 839 (def vib; 1,4-disubstituted benzene). δ_{H} (250 MHz; CDCl₃; Me₄Si) 2.61 (3 H, s, Me), 6.52 (1 H, dd, *J* 3.4 Hz, *J* 1.7 Hz, fuH), 6.81 (1 H, d, *J* 3.4 Hz, fuH), 7.54 (1 H, d, *J* 1.7 Hz, fuH), 7.75 (2 H, d, *J* 8.4 Hz, ArH), 7.98 (2 H, d, *J* 8.4 Hz, ArH). δ_{C} (62.9 MHz, CDCl₃, Me₄Si) 26.7, 107.6, 112.3, 123.7, 129.1, 135.0, 135.7, 143.4, 153.0, 197.6.

2-Bromo-5-methoxyfuran (O2a). A mixture of 2-methoxyfuran (1.50 g, 15.29 mmol), *N*-bromosuccinimide (2.72 g, 15.29 mmol) and anhydrous *p*-toluenesulfonic acid (8.0 mg) in toluene (75 mL) was heated for 4 h at 70 °C. After cooling to room temperature the precipitated succinimide was filtered off. The filtrate was used without further purification for the synthesis of 4-(5-methoxyfuran-2-yl)acetophenone (**O2**).

4-(5-Methoxyfuran-2-yl)acetophenone (O2). To the solution of the crude product **O2a** 20 mL of ethanol and 20 mL of distilled water (degassed with argon) were added. Subsequently, 4-acetylphenylboronic acid (2.46 g, 15.00 mmol), Na₂CO₃ (1.75 g, 16.50 mmol) and Pd(PPh₃)₄ (1.39 g, 1.20 mmol) were added. The dark brown reaction mixture was refluxed under argon for about 72 h. After cooling to room temperature 200 mL of aqueous HCl (0.1 mol L⁻¹) was added to this solution and the aqueous phase was extracted with dichloromethane (2 × 100 mL). The organic phase was washed with distilled water (1 × 100 mL), dried over MgSO₄, filtered and the filtrate was evaporated. The resulting



crude product was purified by column chromatography using DCM as an eluent to yield **O2** (0.33 g, 10%) as an orange-coloured solid.

Mp: 88–92 °C. DSC (25–350 °C, 10 K min⁻¹, N₂): 92 °C (*endo*), 181 °C (*exo*). $\nu_{\text{max}}/\text{cm}^{-1}$ 3004 (ArCH), 2838 (–OCH₃), 1667 (C=O), 1569 (C=C), 1551 (C=C; furan), 1250 (def vib; 1,4-disubstituted benzene), 959 (CH def vib; 2,5-disubstituted furan), 816 (def vib; 1,4-disubstituted benzene). δ_{H} (250 MHz; CDCl₃; Me₄Si) 2.59 (3 H, s, Me), 3.93 (3 H, s, Me), 5.31 (1 H, d, *J* 3.5 Hz, fuH), 6.71 (1 H, d, *J* 3.5 Hz, fuH), 7.60 (2 H, d, *J* 8.8 Hz, ArH), 7.94 (2 H, d, *J* 8.8 Hz, ArH). δ_{C} (62.9 MHz, CDCl₃, Me₄Si) 26.5, 57.9, 82.5, 109.5, 122.0, 129.0, 134.5, 134.9, 143.0, 162.4, 197.3. HRMS (ESI-TOF): calc. for C₁₃H₁₃O₃: *m/z* 217.0859, found: *m/z* 217.0864.

4-(Thiophen-2-yl)acetophenone (S1). Compound **S1** was prepared according to the literature.¹⁴ Lemon-coloured solid. 61% yield. Mp: 116–119 °C. DSC (25–350 °C, 10 K min⁻¹, N₂): 120 °C (*endo*), 255 °C (*endo*). $\nu_{\text{max}}/\text{cm}^{-1}$ 3072 (ArCH), 1675 (C=O), 1598 (C=C), 1559 (C=C; thiophene), 1258 (def vib; 1,4-disubstituted benzene), 822 (def vib; 1,4-disubstituted benzene), 714 (def vib; monosubstituted thiophene). δ_{H} (250 MHz; CDCl₃; Me₄Si) 2.61 (3 H, s, Me), 7.12 (1 H, dd, *J* 4.9 Hz, *J* 3.0 Hz, thH), 7.36 (1 H, d, *J* 4.9 Hz, thH), 7.43 (1 H, d, *J* 3.0 Hz, thH), 7.69 (2 H, d, *J* 8.4 Hz, ArH), 7.97 (2 H, d, *J* 8.4 Hz, ArH). δ_{C} (62.9 MHz, CDCl₃, Me₄Si) 26.7, 124.7, 125.8, 126.6, 128.5, 129.3, 135.9, 138.9, 143.1, 197.5. HRMS (ESI-TOF): calc. for C₁₂H₁₁OS: *m/z* 203.0525, found: *m/z* 203.0532 [M + H]⁺.

2-(4-Acetylphenyl)furan-5-carbonitrile (O3). Compound **O3** was synthesized in analogy to the aryl-substituted thiophene derivatives, which were presented in the studies of Roger *et al.*¹⁴ 4-Bromoacetophenone (1.00 g, 5.02 mmol) was dissolved in *N,N*-dimethylacetamide (15 mL), and 2-cyanofuran (0.92 g, 9.90 mmol), potassium acetate (0.99 g, 10.05 mmol), and Pd(OAc)₂ (12.40 mg, 0.06 mmol) were added to this solution. The reaction mixture was heated for 46 h at 150 °C. The resulting solution was added to DCM, the insoluble part was filtered off and the filtrate concentrated. The crude product was purified by column chromatography on silica gel (eluent: DCM/hexane: 4/1) to obtain **O3** (0.172 g, 16%) as an orange-coloured solid.

Mp: 94–108 °C. DSC (25–350 °C, 10 K min⁻¹, N₂): 104 °C (*endo*), 230 °C (*endo*). $\nu_{\text{max}}/\text{cm}^{-1}$ 3135 (ArCH), 2226 (C≡N), 1677 (C=O), 1609 (C=C), 1562 (C=C; furan), 1268 (def vib; 1,4-disubstituted benzene), 961 (CH def vib; 2,5-disubstituted furan), 814 (def vib; 1,4-disubstituted benzene). δ_{H} (250 MHz; CDCl₃; Me₄Si) 2.64 (3 H, s, Me), 6.87 (1 H, d, *J* 3.6 Hz, fuH), 7.21 (1 H, d, *J* 3.6 Hz, fuH), 7.81 (2 H, d, *J* 8.2 Hz, ArH), 8.03 (2 H, d, *J* 8.2 Hz, ArH). δ_{C} (62.9 MHz, CDCl₃, Me₄Si) 26.8, 108.1, 111.7, 124.1, 124.9, 126.2, 129.3, 132.6, 137.4, 157.3, 197.7. HRMS (ESI-TOF): calc. for C₁₃H₁₀NO₂: *m/z* 212.0706, found: *m/z* 212.0714 [M + H]⁺.

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