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A monolithic device for CO₂ photoreduction to generate liquid organic substances in a single-compartment reactor†

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A solar to chemical energy conversion efficiency of 4.6% was demonstrated for CO₂ photoreduction to formate utilizing water as an electron donor under simulated solar light irradiation to a monolithic tablet-shaped device. The simple CO₂ photoreduction system was realized by exploiting the effect of the carbon substrate on selective CO₂ reduction in the presence of oxygen and selective H₂O oxidation over IrO_x catalysts in the presence of formate.

The development of clean energy sources and CO₂ recycling are crucial approaches to the depletion of fossil resources and climate change induced by an increase in CO₂ emissions; solar energy is an ideal and realistic way to achieve these goals. Direct photoconversion from CO₂ to chemical fuels or organic raw materials for chemical products using water is an ultimate way to realize a sustainable carbon-neutral society. Interest in the direct solar reduction of CO₂ to liquid organic substances has grown due to the higher energy density and accessible storage advantages compared with hydrogen energy. However, direct solar reduction of CO₂ to liquid organics by electrons and protons extracted from water molecules, similar to photosynthesis in plants, has been considered much more difficult than the hydrogen generation reaction. The CO₂ molecule is highly stable, so that challenges to the realization of photoinduced CO₂ recycling using water include the higher potential required for CO₂ reduction than that of hydrogen generation and the low product selectivity in the presence of water. For example, the potentials for CO₂ reduction to CO and formic acid were calculated from the Gibbs free energy change to be −0.10 V (vs. SHE) and −0.17 V (vs. SHE). There have been several reports on highly efficient catalysts for selective CO₂ photoreduction;^{1–4} however, they required a sacrificial reagent as an electron

donor to facilitate CO₂ reduction. More importantly, electrical coupling of reactions in a closed system, in which selective water oxidation to extract electrons and selective CO₂ reduction are functionally coupled, is extremely difficult. Few studies on CO₂ photoreduction utilizing water as an electron donor have been reported for semiconductor photocatalysts;^{5–8} however, some of them were conducted with an external electrical or chemical bias to assist the reaction, and others were active only under ultraviolet light irradiation. While in the case of complex catalysts, other molecules as sacrificial electron donors are required in place of H₂O to facilitate the reaction. Recently, we reported a solar-to-chemical conversion efficiency of 0.14% for formate generation without an external electrical bias through the combination of a photocathode consisting of a p-type semiconductor, zinc-doped indium phosphide (InP), coated with a [Ru{4,4'-di(1*H*-pyrrolyl-3-propylcarbonate)2,2'-bipyridine}(CO)₂]_n polymer (RuCP)⁹ catalyst for CO₂ reduction and a SrTiO₃ photoanode for water oxidation in a two-compartment reactor separated by a proton exchange membrane, in which water is utilized as both an electron donor and a proton source.^{10,11} However, this system should be improved further for practical application. Solar CO₂ recycling requires not only high efficiency and selectivity, but also consideration of the design of simplified low-cost systems. The reduction of CO₂ to formate using a CO₂ electrolyzer equipped with an indium electrode and a Si photovoltaic cell has also been reported.¹² However, an applied voltage of 4–5 V is necessary to operate the CO₂ electrolyzer due to the substantial negative potential for CO₂ reduction over the indium electrode.^{12,13} The solar-to-chemical energy conversion efficiency of the system was less than 2%, while the solar power conversion efficiency was 8–9% because the combination of the CO₂ electrolyzer and photovoltaic cells is accompanied by a potential drop due to various resistances (resistances in the solar cell, between the solar cell and the electrode, and permeance at the proton exchange membrane) and also requires a complicated potential transformer for impedance matching. These issues could be overcome using a monolithic tablet-shaped device, which composed of a light absorber, a cathode for CO₂ reduction and an anode

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for H₂O oxidation. The device is an independent standalone system immersed in a single-compartment reactor filled with water and CO₂. Furthermore, this device is applicable on a large scale by simply setting the devices in an array configuration because the impedance loss caused by scale-up is very small and determined by the direction of tablet thickness. It is also advantageous that replacement of the catalytic components is easier and cheaper.

In the present study, we demonstrate a simple CO₂ photo-reduction reaction that utilizes a monolithic tablet-shaped device. The device is composed of a porous ruthenium complex polymer (p-RuCP) as a CO₂ reduction catalyst, iridium oxide (IrO_x) as a water oxidation catalyst, and a triple-junction of amorphous silicon-germanium (SiGe-jn, Fig. S1, ESI†) as a light absorber. p-RuCP was developed by chemical polymerization of RuCP on a porous carbon substrate (see the ESI†) and connected to the stainless steel side (narrower bandgap side of the junction) of the SiGe-jn. An IrO_x nanocolloid containing no organic substances was synthesized according to a previously reported method¹⁴ and coated on the indium tin oxide (ITO) surface of the SiGe-jn (see the ESI†). These components were functionally coupled to realize CO₂ photoreduction with a high solar-to-chemical conversion efficiency in a single-compartment reactor. Two essential technologies were developed to realize the monolithic tablet-shaped device for CO₂ photoreduction; selective CO₂ reduction, even in the presence of H₂O and O₂, and selective H₂O oxidation, even in the presence of organic substances. Without separation functions, such as in the photosynthesis of plants, the organic substances produced from CO₂ in a liquid phase can be re-oxidized and competes with the H₂O oxidation to O₂ reaction. This is also the case for the other side reaction, in which H⁺ and O₂ produced from H₂O can be reduced and compete with the CO₂ reduction reaction in a single compartment reactor. Both these reactions cancel the products generated at both sides; therefore, selective CO₂ reduction and H₂O oxidation are necessary for the artificial photosynthesis system to produce liquid chemicals from CO₂ and H₂O in a single-compartment reactor.

The first key technology is the selective CO₂ photoreduction in aqueous media, which was established using the semiconductor-metal-complex hybrid system. We have previously reported the combination of a band-controlled semiconductor for visible-light excitation with a metal-complex to catalyze selective CO₂ reduction, in which photoexcited electrons are transferred from the conduction band of the semiconductor to the lowest unoccupied molecular orbital (LUMO) of the metal-complex catalyst within tens of picoseconds, which resulted in highly selective CO₂ photoreduction.^{15,16} Because the concept of the semiconductor-metal-complex hybrid catalyst is highly versatile, we have constructed the InP/RuCP photocathode catalyst for CO₂ reduction.¹⁷

The CO₂ reduction potential (E_{CR}) of p-RuCP was estimated to be -0.18 V (vs. RHE) (Fig. S2 and Table S1, ESI†), which is substantially lower than that of metal electrodes. For example, the CO₂ reduction potential for formate generation over an indium electrode was reported to be -0.90 V (vs. RHE).¹³ This is the advantage of using a metal-complex catalyst, but there still remain issues such as the competing reduction reaction of O₂

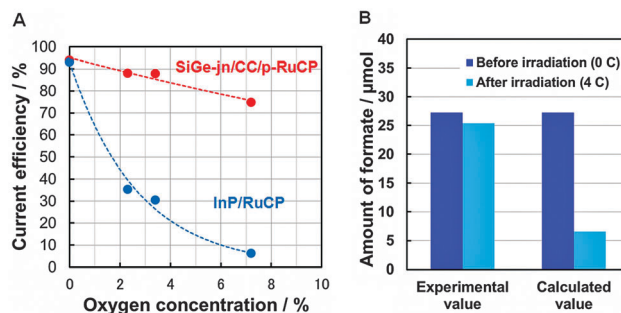


Fig. 1 (A) Current efficiency for formate formation as a function of oxygen concentration in CO₂ flow over SiGe-jn/CC/p-RuCP and InP/RuCP photocathodes under light irradiation for 1 h in 0.1 M phosphate buffer electrolyte. (B) Amount of formate observed before and after photodegradation over the IrO_x/SiGe-jn photoanode.

evolved simultaneously over the water-oxidation site with the CO₂ reduction reaction over InP/RuCP. O₂ is more easily reduced than CO₂ and the elimination of O₂ is thus necessary for CO₂ reduction using a metal complex catalyst.

As evidence, the current efficiency η_{C} for formate formation over InP/RuCP in the presence of O₂ is shown in Fig. 1A (see Table S2, ESI† for details). The InP/RuCP photocathode was used as the working electrode in a three-electrode configuration. The Pt wire and a Hg/Hg₂SO₄ electrode were used as counter and reference electrodes, respectively. 0.1 M phosphate buffer was used as the electrolyte. Gaseous CO₂ containing various concentrations of oxygen was continuously bubbled into the reactor during the reaction, and the CO₂ photoreduction reaction was conducted at $+0.21$ V (vs. RHE) for 1 h. η_{C} was significantly decreased from 93% (at 0% O₂) to 6% (at 7% O₂) with an increase in the oxygen concentration due to selective O₂ reduction (O₂ → O₂^{•−}) competing with CO₂ reduction. Therefore, a system was developed to enhance the CO₂ reduction selectivity over RuCP, even in the presence of O₂. A porous carbon cloth (CC) sheet made of carbon fiber was applied, which possesses a low activity for hydrogen generation (Fig. S3, ESI†), and a surface area that is two orders of magnitude larger than that of the flat and smooth surface of a conventional semiconductor film. η_{C} for formate formation over RuCP coated onto CC (CC/p-RuCP) is also shown in Fig. 1A. CO₂ photoreduction reaction was conducted at $+1.41$ V (vs. RHE). An η_{C} of 76% was observed, even in the presence of 7% O₂. Comparison of the current-potential characteristics of CC under Ar and CO₂ atmospheres suggested preferential adsorption of CO₂ on the CC (Fig. S4, ESI†). It was thus assumed that gaseous CO₂ in aqueous solution was concentrated adjacent to the CC on which RuCP was polymerized. When the RuCP was applied onto the surface of the stainless-steel side (lower bandgap side) of the SiGe-jn, where H₂ generation is preferential (Fig. S3, ESI†), η_{C} for formate generation was only 0.3%, while it significantly improved to $94 \pm 5\%$ when using CC/p-RuCP (Table S3, ESI†).

The second key technology is highly selective H₂O oxidation, even in the presence of formate. Photoreduction of CO₂ to liquid organic products in a single-compartment cell is difficult because the products accumulated in the liquid phase can be re-oxidized to CO₂ on the surface of water oxidation catalysts



such as TiO_2 , as previously reported.¹¹ Iridium oxide, an excellent catalyst for H_2O oxidation,¹⁸ is reported to oxidize formic acid on Ti/IrO_2 electrodes in perchloric acid.¹⁹ In the present study, formate was not decomposed in the oxidation reaction over the IrO_x catalyst in a phosphate buffer solution. Fig. 1B shows the photodegradation of formate over the $\text{IrO}_x/\text{SiGe-jn}$ photoanode performed in 0.1 M phosphate buffer electrolyte containing *ca.* 27 μmol (*ca.* 1.4 mM) of formate with a three-electrode configuration at -0.25 V (*vs.* RHE). The bias voltage was set at the operation point of the present device estimated with a two-electrode configuration (details are provided later). Even though the total charge of 4 C observed during the photoanodic reaction over the $\text{IrO}_x/\text{SiGe-jn}$ photoanode was sufficient to decompose *ca.* 21 μmol formate (calculated value in Fig. 1B), the amount of formate decreased was negligible, which indicates that the IrO_x catalyst has very low activity for the photodegradation of formate. It was also reported that the current efficiency for formate degradation over Ti/IrO_2 in a perchloric acid electrolyte decreased from over 90% to less than 10%, according to the decrease in the concentration of formate from *ca.* 550 mM to *ca.* 30 mM, which indicates that the applied current exceeds the mass transport limit of formate.¹⁷ Therefore, it is supposed that the anodic photocurrent over the $\text{IrO}_x/\text{SiGe-jn}$ photoanode also exceeded the mass transport of formate on the surface of the IrO_x catalyst and generated oxygen from water. The negligible photodegradation of formate over the $\text{IrO}_x/\text{SiGe-jn}$ photoanode was also observed in sulfate, borate and carbonate solutions (Fig. S5, ESI†).

SiGe-jn was selected as the semiconductor for photoexcitation, as employed in the previous report.²⁰ SiGe-jn has an open circuit voltage (V_{OC}) of 2.1 V, which is thermodynamically adequate to oxidize water and extract electrons (approximately 1.4 V), and the p-i-n and tunnel junctions in SiGe-jn facilitate the charge separation and transfer of photoexcited electrons and holes.²¹ The conduction band minimum (E_{CBM}) of SiGe-jn was estimated to be -0.52 V (*vs.* RHE) from the current-potential characteristics (Fig. S6, ESI†). The E_{CBM} of SiGe-jn is more negative than the E_{CR} of -0.18 V (*vs.* RHE, see Fig. S2, ESI†) over p-RuCP; therefore, electron transfer from SiGe-jn in a photoexcited state to p-RuCP is thermodynamically possible. Furthermore, the valence band maximum (E_{VBM}) of SiGe-jn was estimated to be 1.58 V (*vs.* RHE) by subtracting V_{OC} from E_{CBM} . The onset potential estimated from the current-potential curve for water oxidation (E_{WO}) over IrO_x was also estimated to be 1.5 V (*vs.* RHE) (Fig. S7, ESI†). Thus, the E_{VBM} of SiGe-jn is also more positive than the E_{WO} over IrO_x , so that $\text{IrO}_x/\text{SiGe-jn}$ can facilitate H_2O oxidation. The semiconductor-metal-complex hybrid system employs a technical advantage of the Ru-complex catalyst, *i.e.*, a low potential required for CO_2 reduction; therefore, the system using the SiGe-jn can demonstrate CO_2 reduction to formate coupled with water oxidation reaction with a very low voltage of less than 2.1 V.

Based on these two key technologies, the $\text{IrO}_x/\text{SiGe-jn}/\text{CC}/\text{p-RuCP}$ monolithic tablet-shaped device was constructed. A schematic illustration of the device is shown in Fig. 2A. The CO_2 photoreduction reaction was conducted by immersing the device in an aqueous phosphate buffer solution saturated with

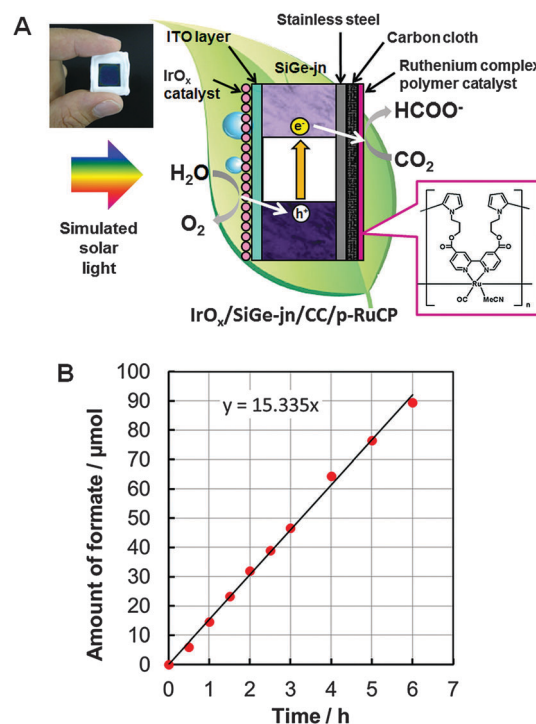


Fig. 2 (A) Schematic illustration of the $\text{IrO}_x/\text{SiGe-jn}/\text{CC}/\text{p-RuCP}$ monolithic tablet-shaped device for CO_2 photoreduction. (B) Time course for the generation of formate during the CO_2 photoreduction reaction using $\text{IrO}_x/\text{SiGe-jn}/\text{CC}/\text{p-RuCP}$ under simulated solar light irradiation (1 sun, AM1.5, 0.25 cm^2). The $\text{IrO}_x/\text{SiGe-jn}/\text{CC}/\text{p-RuCP}$ monolith was immersed in a single-compartment quartz reactor filled with CO_2 -saturated phosphate buffer solution (pH 6.4).

gaseous CO_2 (pH 6.4) in a single-compartment reactor under irradiation with solar simulated light (1 sun, AM1.5, Fig. S8, ESI†). Formate as a liquid organic substance was generated from only CO_2 and H_2O raw materials using sunlight as an energy source. The time course for the generation of formate during the CO_2 photoreduction reaction using the monolithic device with the best performance under simulated solar light irradiation (through a square-shaped slit of 0.25 cm^2) is shown in Fig. 2B. Formate was generated continuously during irradiation for 6 h and the solar-to-chemical conversion efficiency was calculated to be 4.6% from the rate of formic acid generation ($\mu\text{mol HCOOH s}^{-1}$) multiplied by the change in Gibbs free energy per mole of formic acid formation from CO_2 and water (at 298 K, $\Delta G = 270$ kJ mol^{-1}) according to a previous report.²² This result is supported by the photocurrent observed at the operation point shown in the current-potential characteristics of the $\text{IrO}_x/\text{SiGe-jn}$ photoanode and the $\text{CC}/\text{p-RuCP}$ cathode in the three-electrode configuration (Fig. S9, ESI†). A similar photocurrent was also observed at zero bias (*vs.* counter electrode) with the two-electrode configuration using the $\text{IrO}_x/\text{SiGe-jn}$ and $\text{CC}/\text{p-RuCP}$ electrodes (Fig. S10, ESI†). The solar-to-chemical conversion efficiency for the present CO_2 reduction is comparable to that observed for solar hydrogen production utilizing a similar light absorber²⁰ and also reached a level comparable to the theoretical maximal photosynthetic energy conversion



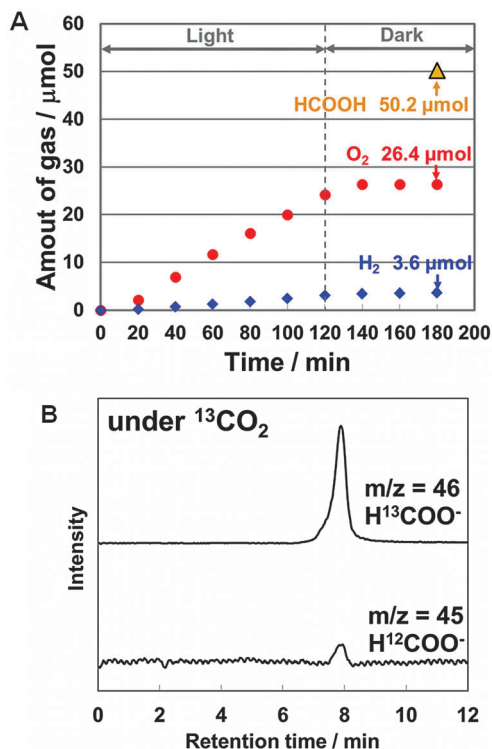


Fig. 3 (A) Experimental verification for formate production from CO_2 and water molecules; time course for oxygen and hydrogen generation during CO_2 photoreduction using a tablet-shaped wireless configuration. The amount of formate was determined at the end of the photoreaction. (B) IC-TOFMS spectra from a tracer experiment utilizing $^{13}\text{CO}_2$.

efficiency for C3 crops (e.g., rice and wheat), estimated to be 4.6%.²³ This reaction was also confirmed to be reproducible. The mean value of the solar-to-chemical conversion efficiency utilizing three monolithic devices was 4.3%.

To verify a stoichiometric reaction, the quantity of oxygen molecules generated over $\text{IrO}_x/\text{SiGe-jn}/\text{CC}/\text{p-RuCP}$ during CO_2 photoreduction was determined from *in situ* measurements. The experiment was conducted in a flow reactor equipped with a single-compartment Pyrex glass cell and a gas chromatograph. The $\text{IrO}_x/\text{SiGe-jn}/\text{CC}/\text{p-RuCP}$ monolithic device was immersed in 0.1 M phosphate buffer saturated with gaseous CO_2 (flow rate 20 mL min^{-1}). A solar simulator equipped with an AM1.5 filter was also used as the light source. The irradiation conditions were different from that used in Fig. 2B due to the experimental setup (see the ESI† for details). During light irradiation, oxygen bubbles were clearly observed at the IrO_x surface. The total amount of oxygen generated after 2 h irradiation was 26.4 μmol (Fig. 3A), which corresponds to 105.6 μmol of photoexcited holes, while 50.2 μmol of formate in the liquid phase and 3.6 μmol of hydrogen in the gas phase were generated simultaneously after 2 h irradiation, which accounts for 107.6 μmol of photoexcited electrons. The amount of electrons was approximately equal to that of photoexcited holes, which strongly suggest that stoichiometric CO_2 reduction is achieved using electrons extracted from water molecules. Furthermore, the ratio of the number of electrons consumed to generate formate to that for the total

reduction products was 93%, which is in good agreement with the current efficiency for formate production observed in a half reaction over the $\text{SiGe-jn}/\text{CC}/\text{p-RuCP}$ photocathode (Table S3, ESI†).

In addition, isotope tracer analysis was conducted using the $\text{IrO}_x/\text{SiGe-jn}/\text{CC}/\text{p-RuCP}$ device in 0.1 M phosphate buffer saturated with gaseous $^{13}\text{CO}_2$ (flow rate 20 mL min^{-1}) to avoid the possibility of experimental error warned by Mul and colleagues.²² Ion chromatography interfaced with time-of-flight mass spectrometry (IC-TOFMS) was used to clarify the formation of $\text{H}^{13}\text{COO}^-$ ($m/z = 46$) (Fig. 3B, mass spectra are shown in Fig. S11, ESI†), which confirmed that the carbon source for formate was the CO_2 molecules. Thus, formate was generated from only CO_2 , H_2O and solar energy over the monolithic device.

Conclusions

A monolithic tablet-shaped device was developed for CO_2 photoreduction to liquid organics in a single-compartment cell. CO_2 photoreduction reaction in almost neutral pH aqueous media containing gaseous CO_2 under irradiation of simulated solar light (1 sun, AM1.5) was successfully demonstrated by immersing the monolithic tablet-shaped device composed of a semiconductor-metal-complex hybrid system in a single-compartment reactor. The solar-to-chemical energy conversion efficiency for formate generation reached 4.6% without external electrical and chemical bias voltages, or a membrane for the separation of products. CO_2 photoreduction to formate utilizing H_2O as an electron donor and a proton source was confirmed by the stoichiometry of reduction/oxidation products and isotope tracer analysis. Highly selective CO_2 reduction and water oxidation in the presence of competitive substrates were identified as the primary technological keys that enable the total reaction to occur in a single-compartment reactor. The secondary key is to facilitate electron transfer from the H_2O oxidation catalyst to the CO_2 reduction catalyst by overcoming the slow CO_2 reduction rate over the metal-complex catalyst. These technologies provide a general and paradigm-changing concept for standalone artificial photosynthesis. The efficiency of the present system can be improved by further optimizing the composition of materials and band engineering of semiconductors and metal complexes because the concept of the semiconductor-metal-complex hybrid system is highly versatile.^{24–28} Therefore, the development of new semiconductor materials and metal-complex catalysts is crucial for further improvement. Furthermore, since a costly iridium oxide catalyst is impractical, development of a water oxidation catalyst is still important. Formic acid, which was generated in the present system, is an industrially useful organic substance²⁹ that has the potential to function as a liquid energy carrier with a high density for CO or H_2 , and as a chemical form for CO_2 fixation with high density. Direct formic acid fuel cells³⁰ and the utilization of formic acid as a hydrogen carrier for fuel cells³¹ have been investigated. A recent report indicated that methanol can be synthesized *via* disproportionation of formic acid catalyzed by a molecular iridium species.³² As an extension,



appropriate replacement of the catalysts in the present system could lead to the direct generation of alcohol.³³

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Notes and references

- 1 J. Hawecker, J.-M. Lehn and R. Ziessel, *J. Chem. Soc., Chem. Commun.*, 1983, 536–538.
- 2 Y. Tamaki, K. Koike, T. Morimoto and O. Ishitani, *J. Catal.*, 2013, **304**, 22–28.
- 3 T. Morimoto, C. Nishiura, M. Tanaka, J. Rohacova, Y. Nakagawa, Y. Funada, K. Koike, Y. Yamamoto, S. Shishido, T. Kojima, T. Saeki, T. Ozeki and O. Ishitani, *J. Am. Chem. Soc.*, 2013, **135**, 13266–13269.
- 4 S. Sato, T. Morikawa, T. Kajino and O. Ishitani, *Angew. Chem., Int. Ed.*, 2013, **52**, 988–992.
- 5 M. Halmann, *Nature*, 1978, **275**, 115–116.
- 6 T. Inoue, A. Fujishima, S. Konishi and K. Honda, *Nature*, 1979, **277**, 637–638.
- 7 K. Iizuka, T. Wato, Y. Miseki, K. Saito and A. Kudo, *J. Am. Chem. Soc.*, 2011, **133**, 20863–20868.
- 8 S. Yotsuhashi, H. Hashiba, M. Deguchi, Y. Zenitani, R. Hinogami, Y. Yamada, M. Deura and K. Ohkawa, *AIP Adv.*, 2012, **2**, 042160.
- 9 S. Chardon-Noblat, A. Deronzier, R. Ziessel and D. Zsoldos, *J. Electroanal. Chem.*, 1998, **444**, 253–260.
- 10 S. Sato, T. Arai, T. Morikawa, K. Uemura, T. M. Suzuki, H. Tanaka and T. Kajino, *J. Am. Chem. Soc.*, 2011, **133**, 15240–15243.
- 11 T. Arai, S. Sato, T. Kajino and T. Morikawa, *Energy Environ. Sci.*, 2013, **6**, 1274–1282.
- 12 J. L. White, J. T. Herb, J. J. Kaczur, P. W. Majsztrik and A. B. Bocarsly, *J. CO₂ Util.*, 2014, **7**, 1–5.
- 13 Z. M. Detweiler, J. L. White, S. L. Bernasek and A. B. Bocarsly, *Langmuir*, 2014, **30**, 7593–7600.
- 14 Y. Zhao, E. A. Hernandez-Pagan, N. M. Vargas-Barbosa, J. L. Dysart and T. E. Mallouk, *J. Phys. Chem. Lett.*, 2011, **2**, 402–406.
- 15 S. Sato, T. Morikawa, S. Saeki, T. Kajino and T. Motohiro, *Angew. Chem., Int. Ed.*, 2010, **49**, 5101–5105.
- 16 K. Yamanaka, S. Sato, M. Iwaki, T. Kajino and T. Morikawa, *J. Phys. Chem. C*, 2011, **115**, 18348–18353.
- 17 T. Arai, S. Sato, K. Uemura, T. Morikawa, T. Kajino and T. Motohiro, *Chem. Commun.*, 2010, **46**, 6944–6946.
- 18 S. Trasatti, *Electrochim. Acta*, 1984, **29**, 1503–1512.
- 19 S. Fierro, L. Ouattara, E. H. Calderon, E. Passas-Lagos, H. Baltruschat and C. Comninelis, *Electrochim. Acta*, 2009, **54**, 2053–2061.
- 20 S. Y. Reece, J. A. Hamel, K. Sung, T. D. Jarvi, A. J. Esswein, J. J. H. Pijpers and D. G. Nocera, *Science*, 2011, **334**, 645–648.
- 21 X. Deng and E. A. Schiff, in *Handbook of Photovoltaic Science and Engineering*, ed. A. Luque and S. Hegedus, Wiley, Chichester, England, 2003, pp. 505–565.
- 22 Z. Chen, T. F. Jaramillo, T. G. Deutsch, A. Kleiman-Shwarscstein, A. J. Forman, N. Gaillard, R. Garland, K. Takanabe, C. Heske, M. Sunkara, E. W. McFarland, K. Domen, E. L. Miller, J. A. Turner and H. N. Dinh, *J. Mater. Res.*, 2010, **25**, 3–16.
- 23 X.-G. Zhu, S. P. Long and D. R. Ort, *Curr. Opin. Biotechnol.*, 2008, **19**, 153–159.
- 24 C.-C. Yang, Y.-H. Yu, B. van der Linden, J. C. S. Wu and G. Mul, *J. Am. Chem. Soc.*, 2010, **132**, 8398–8406.
- 25 K. Sekizawa, K. Maeda, K. Koike, K. Domen and O. Ishitani, *J. Am. Chem. Soc.*, 2013, **135**, 4596–4599.
- 26 K. Maeda, K. Sekizawa and O. Ishitani, *Chem. Commun.*, 2013, **49**, 10127–10129.
- 27 M. Schreier, P. Gao, M. T. Mayer, J. Luo, T. Moehl, M. K. Nazeeruddin, S. D. Tilley and M. Grätzel, *Energy Environ. Sci.*, 2015, **8**, 855–861.
- 28 B. Kumar, J. M. Smieja, A. F. Sasayama and C. P. Kubiak, *Chem. Commun.*, 2012, **48**, 272–274.
- 29 W. Leitner, *Angew. Chem., Int. Ed. Engl.*, 1995, **34**, 2207–2221.
- 30 J. Jiang and A. Wieckowski, *Electrochem. Commun.*, 2012, **18**, 41–43.
- 31 J. F. Hull, Y. Himeda, W.-H. Wang, B. Hashiguchi, R. Periana, D. J. Szalda, J. T. Muckerman and E. Fujita, *Nat. Chem.*, 2012, **4**, 383–388.
- 32 A. J. M. Miller, D. M. Heinekey, J. M. Mayer and K. I. Goldberg, *Angew. Chem., Int. Ed.*, 2013, **52**, 3981–3984.
- 33 D. J. Boston, C. Xu, D. W. Armstrong and F. M. MacDonnell, *J. Am. Chem. Soc.*, 2013, **135**, 16252–16255.

