Gold trifluoromethyl complexes

Juan Gil-Rubio* and José Vicente

This article reviews the synthesis, reactivity and applications of gold trifluoromethyl complexes, which are the only isolated perfluoroalkyl complexes of gold. The most reported examples are neutral Au(I) complexes of the type [Au(CF\textsubscript{3})L], whereas only two Au(II) trifluoromethyl complexes have been reported, both being diamagnetic and containing a strong Au–Au bond. A number of Au(III) trifluoromethyl complexes have been prepared by oxidative addition of halogens or iodotrifluoromethane to Au(I) complexes or, in a few cases, by transmetalation reactions. Owing to the limitations of the available synthetic methods, a lower number of examples is known, particularly for the oxidation states (II) and (III). Gold trifluoromethyl complexes present singular characteristics, such as thermal stability, strong Au–C bonds and, in some cases, reactive α-C–F bonds. Some of the Au(III) complexes reported, show unusually easy reductive elimination reactions of trifluoromethylated products which could be applied in the development of gold-catalyzed processes for the trifluoromethylation of organic compounds.

Introduction

Perfluoroalkyl complexes form a special class among transition metal organometallic compounds. After the first examples were synthesized in the middle of the past century, it has been realized that their properties are remarkably different from those of their non-fluorinated analogues.\textsuperscript{1,2} Thus, it is commonly accepted that they are generally more stable than their alkyl counterparts and that, owing to this stability, the metal–perfluoroalkyl bond presents low reactivity, particularly in reductive elimination or migratory insertion reactions,\textsuperscript{3–5} which are the key steps in metal-promoted coupling reactions.

Efficient synthetic methods for the introduction of fluorinated substituents into organic molecules are currently demanded because fluoroorganic compounds are scarce in nature,\textsuperscript{6–8} and they offer industrial applications as advanced...
materials,9–12 in fluorour chemistry13–17 and, most importantly, in medicine.18,19 Owing to their pharmaceutical applications, trifluoromethylated organic compounds have received special attention. This demand has stimulated the search for reactive perfluoroalkyl metal complexes, giving rise to the recent development of a plethora of new metal mediated or catalyzed trifluoromethylation reactions.20–27

Most metal-promoted synthetic routes for trifluoromethylated organic compounds make use of copper trifluoromethyl complexes,22,24,26,28–31 some of which were used as trifluoromethylating agents in organic synthesis for a long time.32,33 Methods based on other metals, such as silver34–39 or palladium,40–47 have also been developed, although they are less used. Overall, these advances have revealed an unforeseen synthetic potential for perfluoroorganotransition metal compounds, which should be further developed by the investigation of new types of transition metal perfluoroalkyl complexes.

There is still a limited knowledge of the chemistry of gold perfluoroalkyl complexes, which contrasts with the remarkable progress made in the organometallic and catalytic chemistry of gold over the last 25 years.48–55 The first gold perfluoroalkyls were prepared rather late (1973) by Puddephatt and co-workers.56,57 Since then, a relatively low number of trifluoromethyl complexes has been reported. For instance, no higher methyl complexes has been reported. For instance, no higher perfluoroalkyls of gold have been described, with the exception of [Au(CF3)2][PF6] which was spectroscopically observed but not isolated.58 It is also noteworthy that the synthetic routes to key steps in metal mediated- or catalyzed coupling reactions have received considerable interest, because reductive elimination is one of the steps.

The pioneering work on the synthesis of fluorocarbon metal complexes was compiled in three reviews by Stone et al.1,2,69 The field of transition metal trifluoromethyl complexes has been reviewed by Morrison (1993),5 and Menjón et al. (2012).70 In this Perspective, we have reviewed the synthesis, properties and reactivity of gold trifluoromethyl compounds, putting emphasis on recently reported reactions with potential synthetic interest.

Structure, bonding and spectroscopic properties of gold trifluoromethyl complexes

The available experimental evidence suggests that, as observed for other metal perfluoroalkyls, gold trifluoromethyl complexes are more stable than their methyl analogues. For instance, the compounds [Au(CF3)2][CNR] (R = Me,71 Bu61) melt without decomposition at 110 and 119 °C, respectively, whereas their methyl analogues decompose at 95 and 62 °C, respectively.72 The salts [PPh4][Au(CF3)2] (x = 2, 4) start to decompose at 275 °C (x = 2) and 370 °C (x = 4),64 whereas di- and tetramethylaurates Q[AuMe2] (n = 2, Q = Li(MeN(CH2CH3)NMe2)2; n = 4, Q = N4Bu4) readily decompose at 140 °C.73,74 This stability can be attributed to a stronger Au–CF3 bond, as evidenced by the generally shorter Au–CF3 distances compared to those for the Au–CH3 bonds in analogous compounds (Table 1).

In addition, gold trifluoromethyl complexes show two characteristic structural features usually found in other trifluoromethyl complexes: (i) the elongation of the M–C–F distances observed in non-metallic trifluoromethyl derivatives,75 and (ii) the decrease of the F–C–F bond angle and the increase of the M–C–F bond angle76 with respect to the ideal tetrahedral angle. Interestingly, whereas the feature (ii) is always found in all reported structures, the feature (i) is not observed in some Au[n] and Au[nn] trifluoromethyl complexes.

The stability and special structural features of Mn, Rh, Ni and Pt perfluoroalkyl complexes have been addressed by theoretical studies. Thus, in 1972 Hall and Fenske reported the molecular orbital calculations of complexes [Mn(CF3)(CO)3] (X = H, F),83 and in 2012 Macgregor and Grushin reported Natural Bonding Orbital (NBO) analyses of Mn, Rh, Pt and Ni trifluoromethyl and methyl complexes.71 These studies show that: (i) the C 2s character of the M–C bond and the strength of the C–H or C–F bonds; (ii) the high positive charge on the C)/Å and the strength of the C–F bonds; (iii) the high positive charge on the Au and the elongation of the Au–CF3 distance compared to those for the analogous M complexes.

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<table>
<thead>
<tr>
<th>Table 1</th>
<th>Au–C bond distances in analogous Au(i) methyl and trifluoromethyl complexes</th>
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<tbody>
<tr>
<td>Compounds</td>
<td>d(Au–C)/Å</td>
</tr>
<tr>
<td></td>
<td>X = H</td>
</tr>
<tr>
<td>Ph3P–Au–CF3</td>
<td>2.134(28)77</td>
</tr>
<tr>
<td>IPr–Au–CF3</td>
<td>2.065(10)78</td>
</tr>
<tr>
<td>Q[Au(CF3)2]</td>
<td>2.075 (average, Q = N4Bu4)61</td>
</tr>
<tr>
<td>Q[Au(CF3)2]</td>
<td>2.000 (average, Q = N4Bu4)61</td>
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a IPr = 1,3-bis(2,4,6-trimethylphenyl)imidazol-2-ylidene.
increased stability as well as for the Au–C bond shortening and C–F bond weakening observed in these complexes. Accordingly, a comparison of the mean Au–CF3 and C–F distances for the reported structures of Au(i) and Au(III) trifluoromethyl complexes reveals that in Au(m) derivatives, the Au–C distances are longer, and the C–F ones are shorter, than those in Au (i) complexes by about 0.05 Å. This difference could be attributed to rehybridization of the orbitals involved in the Au–C and C–F bonds upon increasing both the coordination number and the oxidation state of the metal. Thus, the Au–C bond distance should be longer for tetracoordinate Au(III) complexes than for dicoordinate Au(i) complexes, whereas the Au–C orbital should decrease its C 2s character in the Au(III) complexes. Consequently, the C–F orbitals should increase their C 2s character.

19F NMR spectroscopy is the most helpful technique to structurally characterize trifluoromethyl complexes in solution. The 19F chemical shift of the Au(i) trifluoromethyl complexes is rather insensitive to the nature of the other ligand (Fig. 1) and, therefore, all the reported values for Au(i) complexes fall in the narrow range of ~24.0 to ~31.5 ppm. In contrast, the chemical shift range of the Au(m) complexes is wider (from ~7.9 to ~44.7 ppm). As noted by Sanner and coworkers, in square-planar Au(m) halo trifluoromethyl complexes the 19F chemical shift is strongly influenced by the nature of the cis halide ligand. Thus, in complexes of the type trans-[Au(CF3)L]X2L (X = halogen, L = phosphine, nitrogen heterocyclic carbene (NHC), CF3−), δ19F increases in the sequence Cl < Br < I (Fig. 1). This effect is also observed in complexes of the type cis-[Au(CF3)L]X(PR3) (Fig. 1), wherein a variation of the halogen X has a stronger impact on the cis-CF3 than on the trans-CF3 ligand.

Fig. 1 19F chemical shifts (ppm) in gold trifluoromethyl complexes.60,61,81,88 IPr = 1,3-bis(2,6-diisopropylphenyl)imidazol-2-ylidene.

Overview of the synthetic methods used for the formation of Au–CF3 bonds

Transmetallation

Owing to the instability of trifluoromethyl lithium or magnesium derivatives, compounds containing less electropositive metals such as Cu, Zn, Cd and Hg have been used as CF3-transfer agents in earlier times. Among these derivatives, the reagent Cd(CF3)2·DME (DME = 1,2-dimethoxyethane) was most popular because of its well-defined composition and high reactivity, which allowed the synthesis of a series of trifluoromethyl complexes of transition metals in low oxidation states, including Au(i) complexes (see below). However, owing to the toxicity of Cd compounds and the need to synthesize the reagent, it was replaced by the commercially available Ruppert–Prakash reagent, Me3SiCF3.89,90 This reagent is typically used in the presence of hard Lewis bases, usually F− or RO−, to generate pentacoordinate Si(iv) species, which act as the nucleophilic CF3-transfer agents.91,92 The volatile Me3SiX (X = F, OR) by-products can be easily separated. A convenient variation of this method is the reaction of Me3SiCF3 with a metal fluorido, hydroxo, or alkoxo complex.

The reactions of the Au(i) chloro complexes with trifluoromethyl silver have been recently used for the synthesis of Au(i) trifluoromethyl complexes (see below). The trifluoromethylating reagent can be in situ generated by using Cd(CF3)2 and a silver salt99 or, more conveniently, by the reaction of AgF and Me3SiCF3 in a N-donor solvent, as reported by Naumann, Tyrra and coworkers.100,101 The NMR studies carried out by the same authors showed that these solutions contain equilibrium mixtures of [Ag(CF3)(solvent)] and [Ag(solvent)2][Ag(CF3)2]99,100 Trifluoromethyl silver has also been used in the synthesis of trifluoromethyl derivatives of various elements of the groups 12–16,100,102 and in the trifluoromethylation of organic substrates.34–36,38,39,103

A few Au(m) trifluoromethyl complexes have been prepared by transmetallation reactions between Cd(CF3)2·DME or Me3SiCF3 and Au(m) precursors, although mixtures of Au(i) and Au(m) complexes frequently result from these reactions (see below).61,88

Oxidative addition of ICF3 to Au(i) complexes

This reaction has been reported for complexes of the type [Au–(R)L] (R = Me,56,57 aryl,67 CF3;58,88 L = phosphine). When R = Me or CF3, it gives mixtures of the products resulting from (i) R by CF3 ligand exchange between the reaction product [Au(CF3)(R)(I)L] and unreacted [Au(R)L], and (ii) reductive elimination to give RI and [Au(CF3)L]. In all these reactions, evidence for the involvement of free radical intermediates has been obtained (see below).

Methods based on the condensation of Au atoms with ‘CF3 radicals or XCF3 (X = Cl, Br)

These methods enabled the synthesis of highly reactive AuCF3 species (see below), although they were obtained in small
amounts and the syntheses require special equipment which is not available in common synthetic laboratories.\textsuperscript{104}

**Au(i) trifluoromethyl complexes**

Puddephatt and coworkers isolated the first gold trifluoromethyl complexes by reacting [Au(Me)L] (L = phosphine) with ICF$_3$ (Scheme 1).\textsuperscript{56,57} Thus, when L = PPh$_3$ the complex [Au(CF$_3$)(PPh$_3$)] was formed, likely by the reductive elimination of MeI from the expected oxidative addition product, [Au(Me)(CF$_3$)(I)L], which was not detected. When L is PMe$_3$ or PMe$_2$Ph, the reactions gave Au(III) trifluoromethyl complexes, and for L = PMe$_2$Ph both Au(i) and Au(III) derivatives were obtained (see below). Evidence to show that these reactions occur at least in part through radical intermediates was obtained. Glockling and coworkers reported the formation of [Au(CF$_3$)(PPh$_3$)] in the reaction of [Au{CH(SiMe$_3$)$_2$}(PPh$_3$)] with ICF$_3$ (Scheme 1).\textsuperscript{105}

Toste and coworkers have reported that [Au(Me)(PCy$_3$)] does not react with ICF$_3$ in the dark. However, the complex [Au(CF$_3$)-(PCy$_3$)] was quantitatively formed upon irradiating the mixture with ambient light (Scheme 1).\textsuperscript{67} The reaction was fast and no Au(III) intermediates were detected by NMR spectroscopy. Interestingly, HCF$_3$ was formed when the same reaction was run in THF, which may be attributed to H$^+$ abstraction from the solvent, suggesting that the CF$_3$ radicals are formed by photoactivation of ICF$_3$ in the presence of the Au(i) complex.

A series of [Au(CF$_3$)$_3$L] complexes (L = phosphine or fluorophosphine) was obtained by the reactions of [Au(Cl)L] with Cd(CF$_3$)$_2$·DME as reported by the groups of Morrison,\textsuperscript{58} Sanner\textsuperscript{88} and Kruck\textsuperscript{106} (Scheme 2). The volatile complex [Au(CF$_3$)-(CNMe)] was similarly prepared by Puddephatt and coworkers and then used for the preparation of gold films by chemical vapour deposition (Scheme 2).\textsuperscript{71} Decomposition of this complex takes place by CNMe dissociation, followed by Au and C$_2$F$_6$ formation. The only higher gold perfluoroalkyl reported, [Au(C$_2$F$_5$)(PET$_3$)], was detected by NMR spectroscopy in the mixture of compounds resulting from the reaction between [Au(CF$_3$)(PET$_3$)] and ICF$_2$F$_5$, although it was not isolated.\textsuperscript{58}

The reactions of Me$_5$SiCF$_3$ with [Au(OR)L] (R = CH(CF$_3$)$_2$, L = PMe$_3$, PCy$_3$, PPh$_3$, PMe$_2$Ph; R = Ph, L = PCy$_3$, PPh$_3$) were used by Komiya and coworkers to prepare a series of trifluoromethyl complexes of the type [Au(CF$_3$)](phosphine) (Scheme 3).\textsuperscript{98} In contrast, the complex [Au(OH)(IPr)] (IPr = 1,3-bis(2,6-diisopropylphenyl)imidazol-2-ylidene)\textsuperscript{107} did not react with Me$_5$SiCF$_3$ at room temperature or at 80 °C.\textsuperscript{108}

Recently, we reported the synthesis of a family of Au(i) trifluoromethyl complexes by the reaction of Au(i) chloro complexes with AgF and Me$_3$SiCF$_3$ (Scheme 4).\textsuperscript{81} This method allows the synthesis of new examples of [Au(CF$_3$)$_3$L] complexes, where L is phosphine, phosphite, or isonitrile, and the first Au(i) trifluoromethyl complexes containing NHC ligands. The analogous reaction with PPN[AuCl(CF$_3$)$_3$] (PPN$^+$ = Ph$_3$PNPPh$_3$) gave the anionic complexes [Au(CF$_3$)$_3$(CNMe)$_2$]$^-$, [Au(CF$_3$)$_3$]$_2$$^-$ and [Au(C$_6$F$_5$)$_3$]$^-$.$^+$ Single crystals of PPN[Au(CF$_3$)$_3$(C$_6$F$_5$)$_2$]$_2$ were iso-
lated and used for a crystal structure determination (Fig. 2). Notably, this is the only transition metal complex reported that contains a CF$_3$ and a C$_6$F$_5$ ligand attached to the same metal. The isonitrile complexes reacted with NHEt$_2$ to give Au(I) trifluoromethyl complexes containing nitrogen acyclic carbene (NAC) ligands (Scheme 5).

The groups of Tyrra,$^{59,60}$ Forniés$^{59}$ and Menjón$^{59,61}$ independently reported the synthesis of the complex [Au(CF$_3$)$_2$]$^-$ by using the reactions of AuCl or [AuCl(tht)] with Me$_3$SiCF$_3$ and F$^-$ (Scheme 6). This is the only anionic Au(I) trifluoromethyl complex prepared in practical amounts, and it was obtained as PPh$_4^+$, NMe$_4^+$, PNP$^+$, K(18-crown-6)$^+$ or [Ag(py)$_2$]$^+$ salts. Remarkably, this complex is sensitive to hydrolysis, decomposing to form Au nanoparticles. Tyrra and coworkers reported that the treatment of [Au(CF$_3$)$_2$]$^-$ with (CF$_3$CO)$_2$O led to a mixture of CF$_3$COF, [Au(CF$_3$)(CF$_2$OC(O)CF$_3$)]$^-$ and [Au(CF$_2$OC(O)CF$_3$)$_2$]$^-$, which were detected by NMR spectroscopy (Scheme 6).$^{60}$ On the other hand, Forniés, Menjón and coworkers reported that BF$_3$-assisted fluoride abstraction of [Au(CF$_3$)$_2$]$^-$ led to nearly quantitative formation of [Au(CF$_3$)(CO)]$^-$ (Scheme 6).$^{59,61}$ These reactions probably proceed through the extremely electrophilic Au(I) difluorocarbene$^{109}$ CF$_3$–Au=CF$_3$, which was not detected.

Despite its thermal instability and extreme moisture-sensitivity, the carbonyl [Au(CF$_3$)(CO)]$^-$ was isolated and structurally characterized.$^{59}$ The crystal structure shows a rare trigonal arrangement of molecules connected by aurophilic interactions (Fig. 3). The weak nature of these interactions, the bond distances and angles, and the position of the ν(CO) band in the IR spectrum suggest that the CO ligand acts as a weak σ-donor coordinated to a highly electrophilic AuCF$_3$ fragment. Thus, the CO ligand is very labile, being immediately substi-

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**Fig. 2** Molecular structure of the anion in the crystal structure of the PPN[Au(CF$_3$)(C$_6$F$_5$)] salt.$^{81}$

**Scheme 5** Synthesis of Aul(i) trifluoromethyl complexes containing NAC ligands by the reaction of (trifluoromethyl)isonitrile complexes with Et$_2$NH.$^{81}$
Au(n) trifluoromethyl complexes

Only two Au(n) trifluoromethyl complexes have been reported to date (Scheme 7), both being diamagnetic and displaying strong Au–Au bonds.

The first example is [Au₂(CF₃)₂{μ-Ph₂P(CH₂)₂}₂], prepared by using the reaction of [Au₂Cl₂{μ-Ph₂P(CH₂)₂}₂] with Cd(CF₃)₂·DME, which contains two bridging bis(ylide) ligands. The other compound is [Au₂(CF₃)₄(py)₂], obtained after irradiating a solution of [Ag(py)₂][Au(CF₃)₂] with UV light. In this remarkable reaction, the [Ag(py)₂]⁺ cation is reduced to metallic Ag and oxidizes [Au(CF₃)₂]⁻ to Au(II). This compound is a rare example of a dinuclear Au(II) complex where the Au–Au bond is not supported by the bridging ligands, and it has been the object of a theoretical study. The computational results showed that the bond between both the Au(n) centers is of covalent nature. Moreover, the analysis of the bonding orbitals suggested that they can be described by an uncommon 6s₆p₅dₓᵧ hybridization type. [Au₂(CF₃)₄(py)₂] was characterized in the solid state by single crystal X-ray diffraction (Fig. 4), but no data could be obtained in solution because it spontaneously disproportionates into a 1:1 mixture of [Au(CF₃)(py)] and [Au(CF₃)₃(py)], both being detected by NMR spectroscopy.

Au(u) trifluoromethyl complexes

Au(u) trifluoromethyl complexes have been prepared either (i) by oxidative addition of ICF₃ to Au(i) complexes, (ii) by halogenation of Au(I) trifluoromethyl complexes, or (iii) by trifluoromethylation of Au(u) complexes.

The first reported Au(u) perfluoroalkyls were obtained by Puddephatt and coworkers in the reaction of [Au(Me)L] with ICF₃ (Scheme 8), which gave mixtures of cis- and trans-[Au(CF₃)Me₂L] (L = PMe₃, PMe₂Ph) in variable ratios depending on the solvent and the nature of the ligand L. From these mixtures, only the trans isomers were isolated. These reactions seem to proceed by the oxidative addition of ICF₃ to the Au(i) complexes to give a [Au(CF₃)(Me)L] intermediate, which undergoes Me for I ligand exchange with the unreacted [Au(Me)L]. The intermediate complex was detected in certain experiments and independently prepared by the reaction of trans-[Au(CF₃)Me₂(PMe₃)] with HCl and NaI. Evidence of a radical chain mechanism for these oxidative addition reactions was obtained (Scheme 9).

The photooxidative addition of ICF₃ to [Au(Ar)L] (L = PCy₃, Ar = 4-X-C₆H₄ (X = H, Me, F, MeO), 3,5-F₂C₆H₄) has been reported by Toste’s group (Scheme 10). The complexes [Au(CF₃)(Ar)I(PR₃)] were isolated in high yields as single isomers. Detailed mechanistic studies support a radical chain mechanism as depicted in Scheme 9. The radical chain is initiated by electron transfer from the Au(i) complex or PCy₃ to a photoexcited molecule of ICF₃, giving rise to I⁻ and a CF₃ radical. However, in contrast to the analogous Au(u) complexes containing a methyl ligand (see above), the complexes [Au(CF₃)(Ar)I(PR₃)] are stable against reductive elimination and do not undergo R for I ligand exchange with the starting [Au(R)L] complexes.

Interestingly, the complexes [Au(CF₃)(Ar)I(PR₃)] underwent thermal reductive elimination to afford Ar and [Au(CF₃)(PPH₃)] (Scheme 10, reaction (a)). The complete inhibition of this reaction in the presence of added PPh₃ suggests that it proceeds through PPh₃ dissociation to give a neutral tricoordinate intermediate which eliminates Ar. In contrast, a very fast

![Scheme 7 Formation of Au(i) trifluoromethyl complexes by (i) trans-metallation in an Au(ii) chloro complex, and (ii) photo-disproportionation of [Ag(py)₂][Au(CF₃)₂]. The unstable Au(i) complex [Au₂(CF₃)₄(py)₂] disproportionates in solution to an Au(ii) and an Au(u) complex.]

![Scheme 8 Reactions of Au(i) methyl complexes with ICF₃ which lead to mixtures of Au(i) and Au(u) complexes.]

![Fig. 4 Molecular structure of [Au(CF₃)₄(py)₂]. The Au–Au bond distance is indicated.]

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reductive elimination of ArCF₃ was observed when the complexes [Au(CF₃)(Ar)L] were treated with AgSbF₆ to abstract an I⁻ anion and generate a cationic tricoordinate intermediate (Scheme 10, reaction (b)). The last process is remarkable taking into account (i) the known reluctance of transition-metal trifluoromethyl complexes to undergo reductive elimination of trifluoromethylated products, and (ii) the potential synthetic utility for a gold-catalyzed aromatic trifluoromethylation reaction. In this respect, the authors of this study have demonstrated the feasibility of each individual step in a hypothetical catalytic cycle for the gold-catalyzed photoactivated trifluoromethylation of an arylstannane.

The reactions of the Au(i) trifluoromethyl complexes [Au(CF₃)L] (L = phosphine) or [Au(CF₃)₂] with an excess of Cl₂, Br₂, or I₂ afforded Au(III) complexes resulting from the oxidative addition of the halogens (Scheme 11, reactions (a) and (c)). In these complexes, the halo ligands are disposed mutually in trans (only small amounts of the cis isomers were observed for L = PMe₃ and PEt₃). When equimolar amounts of [Au(CF₃)L] and halogen were used, the additional products resulting from CF₃/X exchange with the remaining starting complex were also observed (Scheme 11, reaction (b)). The Au(III) complexes obtained in these reactions are stable against thermal reductive elimination.

The complexes [Au(CF₃)(IPr)] react similarly with PhICl₂, Br₂, I₂ or ICl, affording exclusively the isomers with the halogens disposed mutually in trans (Scheme 12). These complexes are thermally stable except in the case of trans-[Au(CF₃)I₂(IPr)], which reductively eliminates ICF₃ even at low temperature. However, on irradiating with UV light the compounds [Au(CF₃)XY(IPr)] (X = Y = Cl or Br; X = Cl, Y = I) underwent reductive elimination of halotrifluoromethane (Scheme 12). Remarkably, the chloro(iodo) complex selectively eliminated ICF₃ instead of ClCF₃.

The oxidative addition of ICF₃ to [Au(CF₃)(PR₃)] (R = Me,58,88 Et88) gave mixtures of cis- and trans-[Au(CF₃)₂I(PR₃)], where the cis isomer is the predominant product (Scheme 13). Inhibition of the reactions by a radical scavenger (galvinoxyl) suggests that they also proceed through a radical mechanism. The partial reactions between [Au(CF₃)I₂(PEt₃)] and [Au(CF₃)(IPr)] to give mixtures of [Au(CF₃)₃(PR₃)] and [AuI(PR₃)] have been also reported (Scheme 13).58 No reductive elimination of ICF₃ or CF₃–CF₃ from these Au(III) complexes was described.

Scheme 9 Proposed radical chain mechanism for the oxidative addition of ICF₃ to Au(I) complexes. Subsequent evolution of the formed Au(III) trifluoromethyl complexes with R = Me to give mixtures of Au(I) and Au(III) complexes.67

Scheme 10 Photooxidative addition of ICF₃ to Au(I) aryl complexes. The resulting Au(II) complexes may undergo (a) thermal reductive elimination of ArI or (b) fast reductive elimination of ArCF₃ after removal of an I⁻ ligand.67

Scheme 11 Oxidative addition of halogens to neutral122 and anionic98 Au(i) trifluoromethyl complexes.
The dinuclear complexes \([\text{Au}_2(\text{CF}_3)_4(\mu-\text{X})_2]\) (X = Br, I) were prepared by condensation of Au atoms with XCF\(_3\) (Scheme 14). The isolated compounds are moderately stable and have been characterized by single crystal X-ray diffraction and NMR spectroscopy.\(^{120}\)

An unstable compound formulated as \(\text{Au(CF}_3)_3\) was isolated from the reaction of trifluoromethyl radicals and Au atoms (Scheme 14).\(^{121}\) The NMR data of its PMe\(_3\) adduct are in agreement with those reported for \([\text{Au(CF}_3]_3(\text{PMe}_3)]\)\(^{88}\) (see below).

Most attempts to obtain Au(\(\text{III}\)) trifluoromethyls by transmetallation reactions of Au(\(\text{III}\)) complexes and Cd(CF\(_3\))\(_2\)-DME resulted in the formation of Au(\(\text{I}\)) complexes as the side or main products, although the desired Au(\(\text{III}\)) trifluoromethyls are stable once isolated.\(^{88}\) For instance, the reaction of \([\text{Au(CF}_3]_3(\text{PMe}_3)]\) with Cd(CF\(_3\))\(_2\)-DME gave \([\text{Au(CF}_3])(\text{PMe}_3)]\), ICF\(_3\) and CdI\(_2\) instead of the expected \([\text{Au(CF}_3]_3(\text{PMe}_3)]\). The reaction of \([\text{Au(CF}_3]_3(\text{PMe}_3)]\) (prepared from \([\text{Au(CF}_3])(\text{PMe}_3)]\) and ICF\(_3\)) with Cd(CF\(_3\))\(_2\)-DME gave \([\text{Au(CF}_3]_3(\text{PMe}_3)]\) in high yield (Scheme 15), although the reaction had to be carried out under an ICF\(_3\) atmosphere, because otherwise \([\text{Au(CF}_3]_3(\text{PMe}_3)]\) and ICF\(_3\) could also be formed.\(^{88}\) The tris(trifluoromethyl) complex is a thermally, air and moisture stable crystalline solid, which can be purified by sublimation.\(^{58,88}\) The complexes cis-\([\text{Au(CF}_3](\text{Me})_2\text{L}]\) (L = PMe\(_3\), PET\(_3\), PMePh\(_2\) or PMePh\(_3\)) have been also obtained by Komiya and coworkers in the reaction of cis-\([\text{Au(OPh)](Me)}_2\text{L}\] with Me\(_3\)SiCF\(_3\) (Scheme 15).\(^{98}\)

\([\text{Au(CF}_3]_4^-\) was firstly prepared by Neumann and coworkers as its PPN\(^+\) salt, although the synthetic method was not reported.\(^{122}\) Recently, Menjón and coworkers described the synthesis of this complex by the reaction of AuCl\(_3\), Me\(_3\)SiCF\(_3\) and CsF (Scheme 16).\(^{61}\) During the reaction, partial reduction
of the metal took place to give a mixture of [Au(CF₃)₂]⁻ and [Au(CF₃)₃]⁻. Fortunately, the mixture could be separated by using PPh₄⁺ as the counterion due to the different solubilities of the Au(i) and Au(ii) salts. (PPh₄)[Au(CF₃)₂] is remarkably stable, because it melts at 125 °C and does not show significant weight loss until it is heated at 370 °C. In addition, de-stabilization of the negative charge over 12 fluorine atoms, makes [Au(CF₃)₃]⁻ an uncommon weakly coordinating anion with a flat shape,¹²¹ which has been used as a component of organic superconductors.¹²²,¹²³,¹²⁷

Finally, the reaction of [N⁺Bu₄][Au(CN)]⁺ with ClF afforded a mixture of complexes of the type [AuFₓClₓ(CF₃)₁₄₋ₓ]⁻ (x = 0–4, y = 0–2). Metathesis of this mixture with Me₃SiX (X = Cl or CN) gave Me₃SiF and new mixtures of the composition [AuXₙ(CF₃)₄₋ₙ]⁻ (x = 0–4). The components of these mixtures were identified by NMR spectroscopy but could not be separated.¹²⁸

Summary and outlook

Au(i), Au(ii) and Au(iii) trifluoromethyl complexes have been reported being, in general, more thermally stable than their methyl analogues. All known Au(i) complexes are of the [Au(CF₃)L] type, except [Au(CF₃)₂]⁻ and [Au(CF₃)(CF₃)]⁻. Most of them have been obtained by using Me₃SiF₃ as the nucleophilic trifluoromethylating agent in the presence of fluoride salts as activators of the silane. Among all the methodologies tested, the combination of Me₃SiF₃ and AgF is the most versatile one. Only two Au(i) trifluoromethyl complexes have been described, both displaying an Au–Au bond. Au(iii) complexes containing 1, 2, 3 or 4 trifluoromethyl ligands have been reported. Most of them have been prepared by the oxidative addition of ICF₃ or halogens to Au(i) complexes.

Overall, the available synthetic methods still suffer from important drawbacks, such as the use of expensive reagents, the low yields in selected cases and the narrow scope. Thus, the development of efficient methodologies for the introduction of trifluoromethyl and, specially, higher perfluoroalkyl groups in the coordination sphere of gold is highly desirable, and it should boost the application of these compounds in different fields.

Potentially interesting applications may emerge from the singular structural properties and reactivity of these compounds. Thus, the observation of unusually fast reductive eliminations of trifluoromethylated compounds in the Au(iii) trifluoromethyl complexes could be applied in the catalytic trifluoromethylation of organic compounds. Other opportunities for the application of these compounds in homogeneous catalysis may arise from the strong electron-withdrawing ability of the trifluoromethyl groups, which enhances the Lewis acidity of the metal center, or from the α-fluoride abstraction reactions, which generate highly reactive difluorocarbene intermediates. In addition, the development of efficient methods for the synthesis of the anions [Au(CₘFₙ₊₁)]⁻ would improve their use as flat, weakly coordinating, fluorophilic anions in synthesis and materials science.

Note added in proof

After submission of this review, Toste and coworkers have reported a study of the reductive elimination of ArX and ArCF₃ in complexes [Au(CF₃)(Ar)(X[PPh₄])] (X = F, Cl, Br, I). See M. S. Winston, W. J. Wolf, F. D. Toste J. Am. Chem. Soc. 2015, 137, 7921.

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Notes and references

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75 The mean C–F distances in Au(i), Au(ii) and Au(iii) compounds are 1.368, 1.306 and 1.323 Å, respectively, whereas the mean XF2C–F (X ≠ metal) distance is 1.33 Å (see A. Yokozeki and S. H. Bauer, Top. Curr. Chem., 1973, 53, 71–119 and F. H. Allen, O. Kennard, D. G. Watson,

76 The mean Au–C–F and F–C–F angles are 115.4° and 102.9° (Au(I)), 116.4° and 101.6° (Au(II)), 113.6° and 105.1° (Au(III)).


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