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## Can we beat the biotin–avidin pair?: cucurbit[7]uril-based ultrahigh affinity host–guest complexes and their applications

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The design of synthetic, monovalent host–guest molecular recognition pairs is still challenging and of particular interest to inquire into the limits of the affinity that can be achieved with designed systems. In this regard, cucurbit[7]uril (CB[7]), an important member of the host family cucurbit[n]uril (CB[n],  $n = 5–8, 10, 14$ ), has attracted much attention because of its ability to form ultra-stable complexes with multiple guests. The strong hydrophobic effect between the host cavity and guests, ion–dipole and dipole–dipole interactions of guests with CB portals helps in cooperative and multiple noncovalent interactions that are essential for realizing such strong complexations. These highly selective, strong yet dynamic interactions can be exploited in many applications including affinity chromatography, biomolecule immobilization, protein isolation, biological catalysis, and sensor technologies. In this review, we summarize the progress in the development of high affinity guests for CB[7], factors affecting the stability of complexes, theoretical insights, and the utility of these high affinity pairs in different challenging applications.

### 1. Introduction

The high affinity and fidelity in molecular recognition and transformation is a key component of the most common biomolecules (e.g. between protein and ligand, antigen and antibody, sugar and lectin, RNA and ribosome) in the generation of life-like complexity and function. Apart from understanding the complexity of life, high affinity host–guest pairs have tremendous values in both biotechnology and chemical applications. The strongest and best known natural noncovalent interaction pair is the

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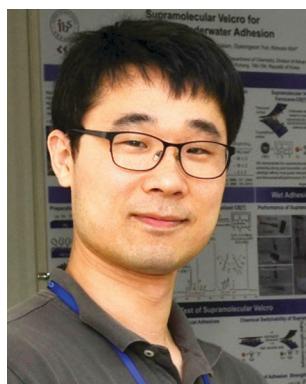
biotin-avidin (or streptavidin) system with a binding constant in the range of  $10^{13}$ – $10^{15}$  M<sup>-1</sup>.<sup>1,2</sup> This affinity pair has been exploited in many applications including enzyme-linked immunosorbent assays, protein and nucleic acid detection and purification, and immobilization of biomolecules on surfaces.<sup>3–9</sup> However, this natural robust binding pair suffers some downsides such that the avidin-biotin complex cannot be easily dissociated (although there has been a considerable amount of work on the site-directed mutagenesis of avidin to temper its affinity for biotin)<sup>10–14</sup> unless operated in elevated temperatures, which causes the denaturation of the proteins. Also, the interaction between avidin and biotin can be undesirably affected by bio-systems, thus limiting their biological applications.<sup>15–17</sup> In order to mimic some of the complexity of natural systems, researchers have been trying to create high affinity synthetic receptor-ligand complexes. In this regard, the central goal has been the development of host-guest systems with binding affinities as high as those reached in biological systems, which often permit the selective binding at concentrations in the nanomolar or picomolar regimes. However, the number of synthetic host families that display biomolecular range affinity ( $K_d < 1 \mu\text{M}$ ,  $K_a > 10^6 \text{ M}^{-1}$ ) and selectivity in aqueous solution is limited to certain cyclophanes,<sup>18</sup> self-assembled receptors,<sup>19–24</sup> and cucurbiturils.<sup>25–29</sup>

Among them, cucurbit[n]uril (CB[n],  $n = 5$ – $8$ ,  $10$ ,  $14$ ), a family of synthetic macrocyclic host molecules composed of  $n$  glycoluril units with a hydrophobic cavity and two identical carbonyl-fringed portals, attracted tremendous interest, which is depicted in regular reviews from different perspectives.<sup>30–39</sup> Isaacs, Kim, and others reported that this host family forms exceptionally stable host-guest complexes with specific guest molecules.<sup>40–44</sup> This complexation mechanism contrasts with that found in other cavitands, such as cyclodextrins (CDs), in which the hydrophobic interactions alone are believed to dominate the complexation process.<sup>45</sup> Consequently, the multiple binding forces involved in the complex formation of CB[n] with different guests give rise to highly stable inclusion complexes in solution as well as in the solid state.

CB[7], an important member of the CB[n] family with an internal cavity volume of 279 Å<sup>3</sup>, which is close to that of β-CD (262 Å<sup>3</sup>), has drawn much attention in host-guest systems because of its ability to complex with a variety of guest molecules with high binding affinity and selectivity.<sup>40–44,46</sup> There are some unique features that separate CB[7] from its other family members. Unlike other extensively studied members – CB[6] and CB[8] – which show poor solubility in common solvents ( $< 10^{-5}$  M), CB[7] has a moderate solubility in water (20–30 mM). More importantly, in contrast to the other synthetic receptors, CB[7] commonly displays binding affinities to its guests between  $10^7 \text{ M}^{-1}$  and  $10^{17} \text{ M}^{-1}$ . These binding constant values are equal to or greater than that of the natural biotin-avidin complex. In 2005, the ultrahigh affinity of CB[7] toward ferrocene<sup>40,41</sup> and adamantane<sup>41</sup> derivatives was first measured by Isaacs using <sup>1</sup>H NMR competition experiments and independently measured by our group in collaboration with Inoue and Kaifer by competition ITC methods. In the subsequent years, researchers explored new high affinity guest scaffolds such as bicyclo[2.2.2]octane<sup>43</sup> and diamantane derivatives,<sup>44</sup> while computational scientists demonstrated the importance of new parameters such as the release of high-energy water molecules during the formation of CB[7]-guest complexes.<sup>47–50</sup> Considering the importance of high affinity host-guest pairs, this review first summarizes and presents the development of high affinity guest molecules for CB[7]. In the later parts, we describe the origins of high-affinity binding and the details about enthalpy/entropy compensation, hydrophobic effects, and their contributions to the unusual binding phenomenon. At the end, we discuss the applications of these high-affinity binding pairs in different fields of research.

## 2. High-affinity host-guest pairs

Even though CB homologues share characteristic features such as a hydrophobic cavity and polar carbonyl groups surrounding



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research focuses on the development of novel supramolecular materials and exploring their applications.

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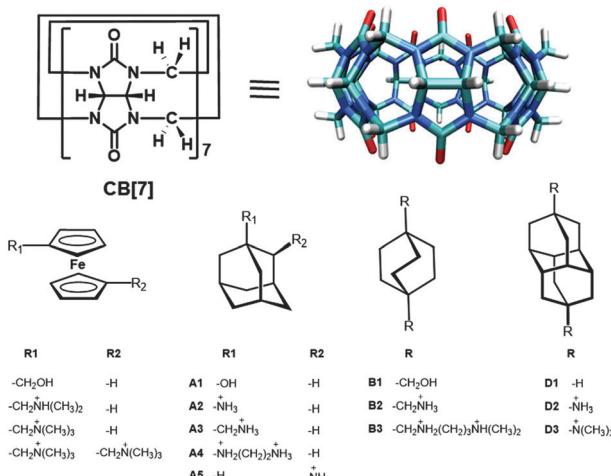
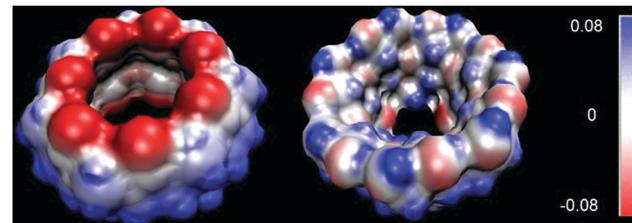


Fig. 1 Molecular structures of CB[7] and its high affinity guests.

the portals, their varying cavity and portal sizes lead to molecular recognition properties different from each other. In general, CB[6] forms very stable complexes with protonated diaminoalkanes whereas CB[7] forms 1:1 complexes with larger guest molecules including ferrocene (and cobalotocene),<sup>26,40–42,46,51</sup> adamantane,<sup>41,43</sup> diamantane,<sup>44,52</sup> and bicyclooctane derivatives,<sup>43,53</sup> which are too large to be included in CB[6] (Fig. 1). The reported affinities of such molecules can approach or even exceed those of the strongest noncovalent interaction pairs found in nature.<sup>3,54</sup> On the other hand, CB[8] has a cavity large enough to accommodate two guest molecules to form 1:2 or 1:1:1 complexes. In general, the affinity of CB[7] complexes in water is higher than buffer solutions because of the ability of metal ions (such as  $\text{Na}^+$ ) to bind at the carbonyl portals of CB[7] which reduces the observed values of association equilibrium constants ( $K_a$ ) by competition.<sup>55,56</sup>

Host–guest properties of CBs have often been compared with those of CDs because of the similarity in their cavity sizes. However, molecular composition and guest binding properties of these two families of molecular containers are quite different from each other. In general, CBs bind guest molecules with higher affinity and higher selectivity than CDs do. For example, the  $K_a$  values for  $\beta$ -CD rarely get over  $10^6 \text{ M}^{-1}$  whereas the corresponding member of the CB family, CB[7] displays much higher binding affinity up to  $10^{17} \text{ M}^{-1}$ . Furthermore, CBs show strong affinity toward cationic guests, moderate affinity toward neutral guests, and little affinity toward anionic guests,<sup>40</sup> while CDs show moderate affinity to hydrophobic guests regardless of their charges, which can be explained by the electrostatic potential of the host families (Fig. 2). In CBs, the regions around the carbonyl oxygens are found to be quite negative (red colored) and the inner surface of the cavity is also significantly negative while the outer surface is somewhat positive (blue colored). On the other hand, the portal and cavity of CDs are almost neutral. We should also note that the high structural rigidity of CBs in comparison with CDs allows highly selective recognition processes for the former. Furthermore, the inner cavity of CB[7] has extremely low polarizability, which is closer to the gas-phase

Fig. 2 The surface electrostatic potential for CB[7] (left) and  $\beta$ -CD (right). The color ranges from negative (red) to positive (blue).

than to that of any other known solvents,<sup>57,58</sup> and a very high quadrupole moment,<sup>59</sup> which allows higher-order electrostatic interactions.

### High-affinity CB[7]–ferrocene complexes

In 2003, the formation of a stable complex between CB[7] and ferrocene<sup>26</sup> or cobaltocene<sup>46</sup> derivatives was observed by us and Kaifer, respectively. In 2005, a parallel effort by Isaacs<sup>41</sup> and our group<sup>40</sup> in collaboration with Inoue and Kaifer revealed that all neutral (F1) and cationic ferrocene derivatives (F2 and F3) (Fig. 1) form highly stable inclusion complexes with CB[7] with  $K_a$  values of  $\sim 10^9 \text{ M}^{-1}$  and  $\sim 10^{12} \text{ M}^{-1}$ , respectively (Table 1). The binding affinities of the ferrocene derivatives toward CB[7] drop by two or three orders of magnitude upon oxidation of the ferrocene core.<sup>60</sup> On the other hand, quite surprisingly, anionic ferrocene derivatives with a similar size and structure do not bind to CB[7].<sup>40</sup> These results are in sharp contrast to the fact that the  $K_a$  values of  $\beta$ -CD for the same guests are only in the range of  $10^3$ – $10^4 \text{ M}^{-1}$  regardless of the charge on the guests. It was surprising to observe that the simple guest molecules bind to CB[7] with such a high binding affinity. In fact, it was

Table 1 Binding constants ( $K_a$ ) and relevant thermodynamic parameters for the complexation of ferrocene (F), adamantane (A), bicyclo[2.2.2]octane (B), and diamantane (D) derivatives with CB[7]. The values for different host–guest complexes given in the table are measured in  $\text{H}_2\text{O}$  except for diamantane complexes with CB[7], which are measured in  $\text{NaO}_2\text{CCD}_3$  buffer at  $\text{pH} = 4.74$  (see Fig. 1 for molecular structures)

Guest	$\Delta H_{\text{expt}}^\circ [\text{kcal mol}^{-1}]$	$-T\Delta S_{\text{expt}}^\circ [\text{kcal mol}^{-1}]$	$K [\text{M}^{-1}]$
F1 <sup>a</sup>	$-21.5 \pm 0.5$	$8.6 \pm 0.5$	$(3.2 \pm 0.5) \times 10^9$
F2 <sup>a</sup>	$-21.0 \pm 0.5$	$4.1 \pm 0.5$	$(2.4 \pm 0.8) \times 10^{12}$
F3 <sup>a</sup>	$-21.5 \pm 0.5$	$4.3 \pm 0.4$	$(4.1 \pm 1.0) \times 10^{15}$
F4 <sup>a</sup>	$-21.5 \pm 0.2$	$0.5 \pm 0.5$	$(3.0 \pm 1.0) \times 10^{15}$
A1 <sup>a</sup>	$-19.0 \pm 0.4$	$4.9 \pm 0.4$	$(2.3 \pm 0.8) \times 10^{10}$
A2 <sup>a</sup>	$-19.3 \pm 0.4$	$-0.1 \pm 0.5$	$(1.7 \pm 0.8) \times 10^{14}$
A3 <sup>a,c</sup>	$-21.9 \pm 0.4$	$1.7$	$7.7 \times 10^{14}$
A4 <sup>a,c</sup>	$-20.1 \pm 0.4$	$-1.4$	$5 \times 10^{15}$
A5 <sup>a</sup>	$-19.5 \pm 0.4$	$0.4 \pm 0.5$	$(1.0 \pm 0.3) \times 10^{14}$
B1 <sup>a</sup>	$-15.8 \pm 0.2$	$2.4 \pm 0.2$	$(6.1 \pm 0.5) \times 10^9$
B2 <sup>a</sup>	$-15.6 \pm 0.4$	$-3.9 \pm 0.5$	$(2.0 \pm 0.5) \times 10^{14}$
B3 <sup>a</sup>	$-16.3 \pm 0.4$	$-4.3 \pm 0.5$	$(1.2 \pm 0.5) \times 10^{15}$
D1 <sup>b</sup>	—	—	$4.0 \times 10^9$
D2 <sup>b</sup>	—	—	$(1.4 \pm 0.3) \times 10^{11}$
D3 <sup>b</sup>	—	—	$7.2 \times 10^{17}$

<sup>a</sup> From ref. 43. <sup>b</sup> From ref. 44. <sup>c</sup> Due to the extremely slow equilibrium, it was difficult to determine the error by the NMR competition technique.

the tightest binding observed for any monovalent synthetic host-guest systems then. Even though this phenomenon was difficult to understand completely at that time, the much higher guest binding affinity of CB[7] compared to  $\beta$ -CD was attributed at least in part, to the contribution of ion-dipole interactions, which are not favored in the case of  $\beta$ -CD as suggested by the electrostatic potential profiles (Fig. 2). Also, the highly negative regions around the two cavity entrances (Fig. 2) of CB[7] lead to the rejection of negatively charged ferrocene guests. The importance of ion-dipole interactions between the host and guest molecules for the high affinity binding process was further supported by the enhancement of binding affinity ( $\sim 10^3$ ) upon the replacement of the hydroxymethylene group of **F1** ( $K_a = \sim 10^9 \text{ M}^{-1}$ ) with a positively charged ammonium methylene unit (**F2** and **F3**;  $K_a = \sim 10^{12} \text{ M}^{-1}$ ). In our observations it was clear that the ferrocene moiety is nicely located inside the CB[7] cavity through hydrophobic interactions, while the positive charge on the substituents interacts with the negatively charged portals of CB[7] by electrostatic interactions.

Intrigued by the exceptionally high affinity of CB[7] toward **F2** and **F3**, we decided to extend this work to develop even tighter binding guests for CB[7] in collaboration with Inoue, Kaifer, Isaacs, and Gilson. We thought that the addition of a second ammonium unit to the ferrocene unit in **F3**, which can interact with the other portal of CB[7], would further increase the affinity of the guest to CB[7]. Indeed, the binding affinity of CB[7] toward **F4** having two ammonium units further increased to  $3 \times 10^{15} \text{ M}^{-1}$ , as measured by a 4-step competition ITC experiment.<sup>42</sup> The X-ray crystal structure shows that each of the trimethylammoniomethyl arms protrudes through the CB[7] cavity and consequently the dicationic ferrocene guest resides inside CB[7] symmetrically. The tilting of the ferrocenyl residue allows the ideal positioning of each of the trimethylammonium groups to maximize ion-dipole interactions with the carbonyl rims on each of the host portals. The tight fit of **F4** inside the CB[7] cavity was also evident in NMR spectroscopic experiments; the much shorter spin-lattice relaxation time ( $T_1$ ) observed for the ferrocene core protons of **F4@CB[7]** compared to **F4** indicate that the rotational motion of the ferrocene is restricted due to the tight fit of the guest upon complexation. Apart from extensive use of NMR, ITC, electrochemical analysis, and crystallography techniques for the interrogation of these complexes, Raman spectroscopy was also used recently to investigate the solid state structural aspects.<sup>51</sup> The Raman spectra of ferrocene@CB[7] showed significant frequency shifts of specific vibrational modes localized on different parts of the complex due to the structural and electronic interactions between the host and guest components, which confirmed the presence of the ferrocene moiety inside the CB[7] cavity.

Another interesting observation was the enthalpy and entropy changes associated with the guest binding. Unexpectedly, the addition of methylammonium groups onto the ferrocene core does not affect the enthalpy change ( $\Delta H^\circ = \sim -21 \text{ kcal mol}^{-1}$ ) associated with the complexation; instead, it causes an increase in the entropy change by  $3.8\text{--}4.3 \text{ kcal mol}^{-1}$ , which results in a thousand fold enhancement in the corresponding binding constant. The change

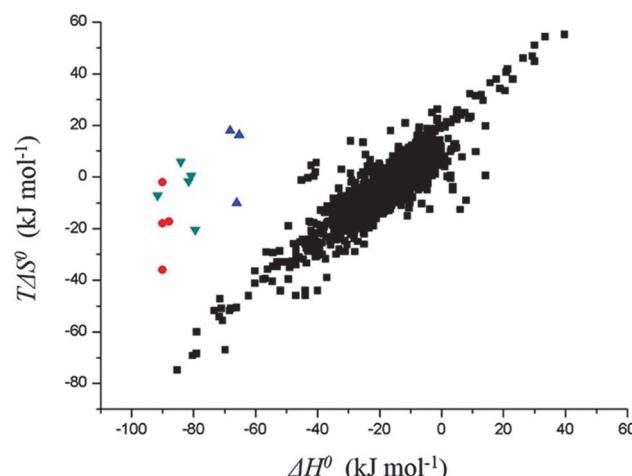


Fig. 3 The thermodynamic data obtained for complexation of ferrocene (●), adamantane (▼), and bicyclo[2.2.2]octane (▲) derivatives with CB[7] which are significantly deviated from the enthalpy–entropy compensation plot for cyclodextrin–guest complexation (■). Data reproduced from ref. 43 and 45.

in solvation entropy ( $\Delta S_{\text{solv}}^\circ = \Delta S^\circ - \Delta S_{\text{config}}^\circ$ ) upon complexation is more favorable for the strong binding dication **F4** ( $18 \text{ kcal mol}^{-1}$ ) compared to monocationic or neutral guests ( $11 \text{ kcal mol}^{-1}$  for **F3** and  $9 \text{ kcal mol}^{-1}$  for **F1**). This observation suggests that the enthalpy–entropy compensation effect, which is commonly observed in other molecular recognition systems,<sup>45,61,62</sup> is not being obeyed in this case. Fig. 3 presents the data points for different high affinity guests along with those for the binding of CDs with a variety of guest molecules, which greatly deviate from the enthalpy–entropy compensation plot for CD–guest complexation.

#### High-affinity complexes of CB[7] with bicyclo[2.2.2]octane and adamantane guests

Right after the successful isolation of CB[7], we observed the formation of a stable host–guest complex between protonated adamantylamine (**A2**) and CB[7].<sup>26</sup> Isaacs and coworkers measured the binding constant for the first time and found that the binding between CB[7] and **A2** is indeed very strong and the binding constant to be  $4.2 \times 10^{12} \text{ M}^{-1}$ .<sup>41</sup> The authors studied the binding properties of CB[7] toward multiple mono-, di-, and tri-substituted adamantanes to probe the size limits and functional group preferences of CB[7]. Among the monosubstituted compounds, cationic ammonium compounds were preferred by CB[7] ( $10^4$ -fold preference) relative to an anionic compound, irrespective of the nature of the ammonium group ( $\text{NH}_3^+$ , quaternary or pyridinium). They also found that the bulkiness of the substituents on the adamantane core in the case of di- or tri-substituted derivatives has a significant effect on the affinity of the guest molecules toward CB[7]. In 2009, Inoue, Gilson and co-workers designed a new series of high affinity guest molecules for CB[7] through theoretical studies.<sup>63</sup> The designed molecules essentially replace the ferrocene moiety with a bicyclo[2.2.2]octane moiety, which is similar to the former in terms of size and hydrophobic character. Two years later,



the same group tested their blinded calculations on designed molecules by experimental methods and found that bicyclo[2.2.2]octane derivatives indeed have high affinity toward CB[7] (Table 1).<sup>43</sup> They synthesized and tested a series of bicyclo[2.2.2]octane and adamantane derivatives to further probe the high-affinity binding regime of CB[7]. These new guests having a rigid and aliphatic core are chosen to fit well into the cavity of CB[7] and outfitted with cationic groups positioned to interact favorably with one or both CB[7] portals. These metal-free guest molecules also show ultrahigh binding affinity toward CB[7], with a binding constant in the range of  $10^{14}$ – $10^{15}$  M<sup>−1</sup>. Similar to our previous observations for the ferrocene derivatives, the neutral bicyclo[2.2.2]octane guest (**B1**) showed less affinity toward CB[7] compared to those having ammonium groups (**B2** and **B3**). Remarkably, the binding affinity of mono-cationic adamantane derivatives (**A2** and **A3**) match with those of **F4** or **B3** having two cationic substituents. However, in these series of guest molecules, the increase in overall charge of the molecules resulted in only a modest increment in binding affinities (Table 1). These observations suggest that a random charge increment to the guest molecule does not warrant a drastic increase in affinity and hence a proper design of guest molecules in accordance with the size of the CB[7] cavity and substitutions at proper positions is more important to achieve the ultra-affinity toward CB[7]. Similar to ferrocene, the remarkably high affinities of these guests are attributed in part to the fact that they lose an unusually small amount of entropy on binding in proportion to their enthalpic changes. A very recent work by Gilson suggests that dispersion forces contribute strongly to the computed binding enthalpies of these complexes, partly due to the fact that their packing density is greater than that of water.<sup>64</sup>

### Ultrahigh affinity CB[7]–diamantane complexes

Very recently, Isaacs, Glaser, Mlinarić-Majerski, and co-workers reported an attomolar-affinity host–guest complex between CB[7] and diamantane quaternary diammonium ion derivative **D3** ( $K_a = 7.2 \times 10^{17}$  M<sup>−1</sup> in pure D<sub>2</sub>O), which is the tightest binding small molecule host–guest complex ever measured for a monovalent molecular recognition to date.<sup>44,52</sup> They studied two diamantane derivatives (**D2** and **D3**) and found that **D3** has much higher binding affinity toward CB[7] than **D2**, though the backbone of the molecules remained the same. The X-ray crystal structure of **D3@CB[7]** shows that the core **D3** is complementary to the cavity of CB[7] and the quaternary ammonium groups of **D3** allow the guest to form optimal ion–dipole interactions with CB[7]. In addition, the authors speculated that the highest binding constant is a combinational outcome of the desolvation of the carbonyl portals within the **D3@CB[7]** complex and the hydrophobic nature of the diamantane core relative to the adamantane and ferrocene cores. It is interesting to note that **D2@CB[7]** has a packing coefficient (79%)<sup>39</sup> much higher than the ideal packing coefficient (55%) for host–guest systems,<sup>21</sup> which may be reflected in its relatively low binding constant ( $K_a = 1.3 \times 10^{11}$  M<sup>−1</sup>) compared to that of **A2@CB[7]** ( $K_a = 4.23 \times 10^{12}$  M<sup>−1</sup>; packing

coefficient = 61%). Nevertheless, the ultrahigh binding affinity of **D3** for CB[7] suggests that the unfavorable effect due to the high packing density may be overcompensated by the positive effect caused by the two quaternary ammonium groups. This work illustrates a possibility of developing synthetic systems whose functions match or exceed their biological counterparts.

The overall binding constants for CB[7] in Table 1 can be generally understood by two common features: (1) the hydrophobic cores of the guest molecules are rigid and complementary to the cavity of CB[7] in size and shape, which contributes to the hydrophobic effect and (2) the ammonium groups of the guest molecules are properly positioned, ideally centrosymmetrically, to enable close polar interactions with the CB[7] portals, which results in ion–dipole interactions. The recent experimental and computational thermodynamic analysis give a clearer picture of the factors accounting for the exceptional binding affinities of selected guests toward CB[7]. The removal of all high-energy water molecules and the desolvation of the guests depending on their hydrophobic surface area can greatly contribute to the binding affinity. In general, the hydrogen bonding interactions between the host and guest molecules are not a dominant force in the guest binding in water, which is evident from the fact that a quaternary trimethylammonium group gives a better binding motif than a tertiary or primary ammonium group. In contrast to other supramolecular recognition systems, the CB[*n*] family shows a unique thermodynamic trend by not obeying enthalpy–entropy compensation.

## 3. Theoretical findings on high stability

The extensive theoretical analysis based on density functional theory (DFT) calculations and molecular dynamics (MD) simulations describes the factors responsible for the high-affinity binding phenomenon of CB[7]–guest pairs.<sup>39,50,65</sup> Among such factors, the one related to solvent and thermodynamic relationships are highly considered due to their potential use in elucidating the complexation mechanism. The ITC data (Table 1) suggest that the unusually high affinities of CB[7] with ferrocene, bicyclooctane, and adamantane derivatives can be attributed to their ability to overcome the usual pattern of the enthalpy–entropy compensation.<sup>40,42,43</sup> The molecular modeling studies using the second-generation mining minima algorithm (M2) method by Gilson well supported these experimental observations. According to these studies, the positive solvation entropy ( $\Delta S_{\text{solv}}^\circ = \Delta S^\circ - \Delta S_{\text{config}}^\circ$ ) gained by dehydration from the CB portals and cavity as well as guests is canceled by the negative entropy arising from the increased conformational restriction of the host and guest upon complexation. This unique thermodynamic behavior avoids the large entropic penalty expected on the basis of enthalpy–entropy compensation and produces the extraordinarily high affinity as in **F4@CB[7]**. Thus, in addition to the size–shape complementarity between the host and guest, the release of water molecules to the bulk solvent upon guest binding is also suggested as an



equally important factor for high binding affinity.<sup>42</sup> According to Gilson,<sup>63</sup> in the case of binding of cationic guests toward CB[7], despite the electrostatic complementarity among the guests and the electronegative carbonyl portals of CB[7], these interactions do not provide a significant net driving force for the high binding. Since the strong Coulombic attractions between the host and guests are precisely balanced by the energetic cost of stripping solvating water molecules upon tight binding through ion-dipole interactions and hence, the free energy contributions are almost entirely entropic in nature.<sup>63</sup> Such near-cancelation of electrostatic terms leaves the van der Waals energy component as the main net contributor to the high-affinities among CB[7] complexes. With the central role of solvent entropy in determining the variations in binding free energy across the high-affinity binding systems, recently another theory demonstrating the importance of the removal of high-energy water molecules has also been put-forth. Using MD simulations, De Simone and co-workers,<sup>47</sup> reviving the earlier hypothesis for CDs,<sup>66,67</sup> demonstrated that the major driving force for complexation of hydrophobic residues with CB[n] originates from the release of high-energy water molecules (enthalpically and entropically unfavorable) in the cavities (Fig. 4). The unfavorable water–water interactions in CB[n] ( $n = 5–7$ ) cavities than the bulk phase are evident from the calculated potential energy for water molecules and hydrogen bond count; with increase in the cavity size, the energy of the individual encapsulated water molecules decreases, since the larger cavities allow for the formation of more stable H-bond networks. The potential energy of release of all internal water molecules was noted to be maximum for CB[7], since the size (diameter) of the CB[7] cavity does not allow for the cavity water molecules (7 to 8) to arrange in an energetically stable H-bond network and gains maximum energy by a complete removal of all water molecules from the cavity. This potential energy is higher than that in CB[6] (a smaller cavity) on account of more than twice as large the number of water molecules released upon binding. Also, the larger cavity of CB[8] with 13–14 water molecules optimizes the stable H-bond network to a degree that is structurally similar to the bulk phase and consequently the resulting energetic gain per released water molecule, compared to CB[7], is less for CB[8]. The complementary sizes between the host and the guest are equally necessary to remove the high energy water molecules in the cavity after complexation that would result in a large enthalpic gain as observed in high-affinity CB[7]–ferrocene derivatives.<sup>42</sup> Meanwhile, Gilson and

coworkers also reported a related study on the unfavorable energetic and entropic properties of the confined water molecules in the interior toroidal region of the non-polar cavity of CB[7] that contribute to the unusually high affinity binding.<sup>48</sup>

A recent report on attomolar-affinity CB[7] and diamantane complexes by Isaacs *et al.*<sup>44</sup> is the best example that explains the importance of both favorable solvent entropy and the release of high-energy water molecules for the ultrahigh affinity complexation along with the direct interactions between the host and the guest. Although the inclusion of either ferrocene, adamantane, or diamantane derivatives in the CB[7] cavity results in the release of high energy water molecules, the highest binding constant of D3@CB[7] complexes is a result of additional solvating water molecule release due to the larger and more hydrophobic core of D3 (14 C atoms), which may subsequently give an additional entropic driving force. Overall, multiple studies have demonstrated the importance of various factors driving the high-affinity complexation in CB[7]. The theoretical calculations point that high-affinity binding is a combined effect of removal of all high energy water molecules from cavity, desolvation of guest molecules, and the direct interactions between the host and the guest. With the mentioned factors molecular recognition studies are aiming to control the selectivity and affinity of these host–guest systems.

## 4. Applications of high-affinity CB[7] complexes

### Brief discussion on functionalized CB[7]

Even though the CB[n] family has proved great potential in molecular recognition, self-assembly, and nanotechnology, their practical applications were limited mainly because of their lack of solubility in common organic solvents, and relatively inert internal or external molecular surfaces. Ideally, if the CB[n] family could be modified to improve their solubility and to provide different functional groups that allow further derivatization, the range of potential applications would be significantly expanded. Over the years, thus, significant efforts have been devoted to the functionalization of CB[n] not only to improve their solubility but also to synthesize tailor-made derivatives for various applications.<sup>28</sup> The functionalization was achieved either by use of substituted glycoluril derivatives in CB[n]-forming reactions or by direct substitution at the “equator” positions of CB[n].<sup>68–78</sup> In a major breakthrough in the CB chemistry, we discovered a direct functionalization method to produce perhydroxylated products of CB[n] through oxidation with  $K_2S_2O_8$  in water.<sup>77</sup> This finding represented the first covalent derivatization reactions of the CB[n] family and allowed us to synthesize a wide variety of CB derivatives having many potential applications.<sup>79–95</sup> The reaction efficiency is reasonably good for CB[5] and CB[6] (yield  $\sim 40$ –45%) but the higher homologues CB[7] and CB[8] give a low yield (5% or less) presumably due to instability of the perhydroxylated products. These substituted hydroxy groups at the periphery

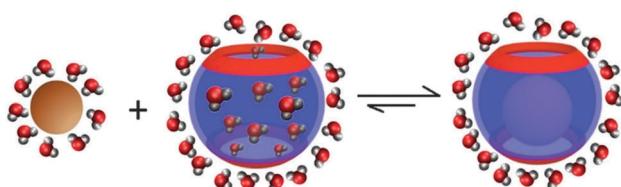


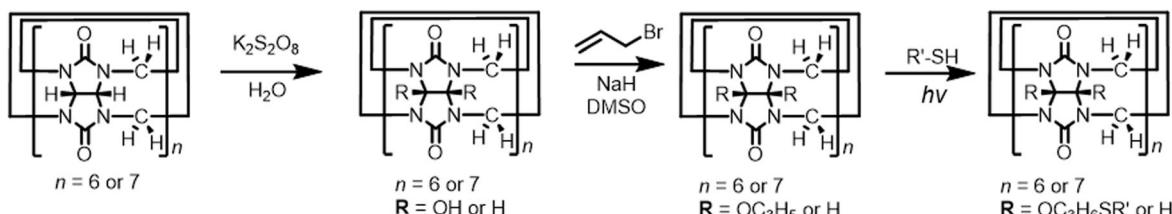
Fig. 4 Schematic representation elucidating the release of high-energy water molecules from the CB[7] cavity upon binding of a hydrophobic guest.



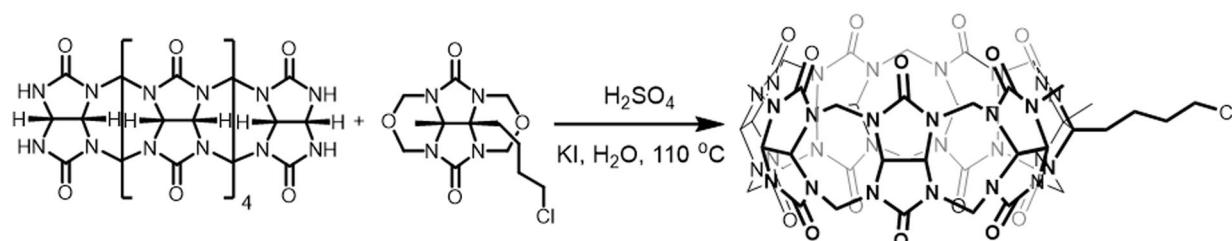
of the CB[n] frameworks do not alter the portal and cavity sizes of the parent molecule and hence the host–guest chemistry pertaining these molecules remains the same. In the subsequent years, by controlling the stoichiometry of the oxidizing agent, we were also able to synthesize multihydroxy<sup>77,80</sup> and monohydroxy-CB[7] derivatives,<sup>96</sup> which are soluble in organic solvents such as DMSO and DMF, which allows further functionalization *via* conventional organic transformation routes including acylation and alkylation (Scheme 1). For example, it can be alkylated with allyl bromide or propargyl bromide to yield mono or multi substituted allyl or propargyl derivatives from corresponding mono or multihydroxyl substituted CB[7], which turned out to be a useful precursor for further transformation and used in many applications that we discussed in this review. We did further modification of allyl substituted CB[7] through photo-reaction with suitably designed thiol compounds.<sup>28</sup> In addition, the multihydroxy CB[7] derivative found to be useful in reducing non-specific binding in biological applications as discussed below.

Until recently, the method introduced by our group has been the only direct route to obtain functionalized CB[7] compounds. However, Isaacs and co-workers recently reported a building-block approach to obtain monofunctionalized CB[7] derivatives (Scheme 1).<sup>97</sup> This approach involves the condensation of an acyclic glycoluril hexamer with glycoluril bis(cyclic ethers) to produce monofunctionalized CB[7] which can be further transformed into tailor-made CB[7] derivatives for various applications. In addition, very recently Bardelang, Ouari, and co-workers have reported another oxidative method to synthesize high purity monohydroxy-CBs in good yield using hydrogen peroxide and UV light.<sup>98</sup> This method relying on a controlled generation of hydroxyl radicals from hydrogen peroxide seems useful in the rapid generation of monohydroxy derivatives of the main members of the CB[n] family in high yield (> 90%). By taking the advantage of progress in functionalization of CB[7], a number of applications of CB[7] and its functional derivatives are reported which will be covered in the following sections.

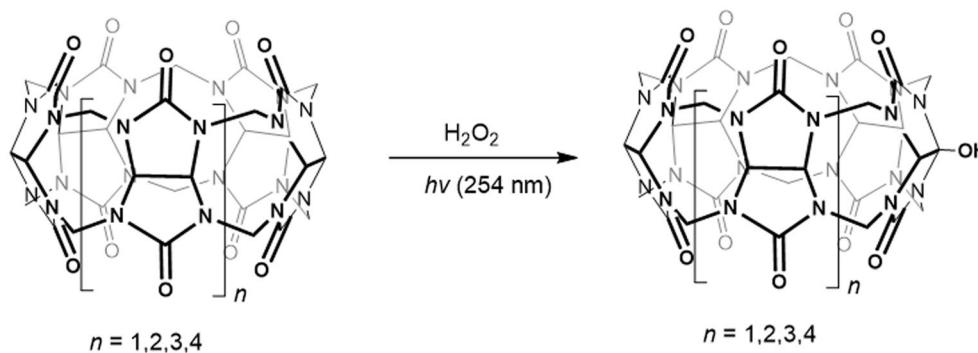
**a) Radical mediated oxidative hydroxylation and subsequent functionalization of CBs (Kim)**



**b) Building block approach for the synthesis of monofunctionalized CB[7] (Isaacs)**



**c) Photo-initiated radical mediated oxidative method for the monohydroxylation of CBs (Bardelang and Ouari)**



**Scheme 1** Synthetic methods developed for the functionalization of CBs.



## Immobilization of biomolecules on a solid surface

Developing a robust yet convenient method to link biomolecules to predefined locations on a reactive surface is of great importance in biotechnology and biochemical applications.<sup>99–101</sup> Site-specific chemical strategies, as opposed to random attachment, provide precise control over immobilizing biomolecules in homogeneously oriented layers, and yield improved performance of biochips.<sup>102–107</sup> For the immobilization of structurally sensitive biomolecules, mild reaction conditions compatible with physiological conditions, and rapid and quantitative conversion are essential. Biotin–streptavidin pair is one of the most widely used tools for such applications because of their strong and specific interaction. However, it suffers some shortcomings including denaturation by organic solvents or elevated temperatures. Right after the measurement of the exceptionally high affinity of ferrocenemethylammonium (F3) toward CB[7] by Isaacs<sup>41</sup> and our group,<sup>40</sup> we immediately realized that the ultrastable host–guest complex can be utilized as a replacement for the biotin–avidin pair in many applications where the latter has been widely used. The synthetic system has advantages including resistance to harsh conditions such as elevated temperature or unwanted enzymatic degradation, easiness to handle and manipulate, long-term durability due to the usage of stable synthetic molecules, and redox-active nature of the ferrocene unit, which allows modulation of binding strength.

With this idea in mind, we first decided to develop a novel method to immobilize proteins on gold using the ultrastable F3@CB[7] (FA@CB[7]) complexes, and explore their applications as a biosensor.<sup>80</sup> The surface immobilization of CB[7] was achieved by olefin metathesis reaction between allyloxy CB[7] and a vinyl-terminated mixed self-assembled monolayer (SAM) on gold. Ferrocene-conjugated proteins can directly interact with CB[7] moieties on the surface (Fig. 5). We chose glucose

oxidase (GOx), one of the most well-known redox proteins extensively studied for glucose sensor applications, as a model protein and the “ferrocenylated” GOx (FA-GOx) was prepared by simple coupling reaction. The protein was immobilized on the CB[7]-anchored gold substrate by direct immersion of the substrate in a solution of FA-GOx. The surface plasmon resonance (SPR) technique confirmed that FA-GOx was anchored with a surface density of  $1.1 \times 10^{12}$  mol cm<sup>-2</sup>, which is approximately 5.5 times higher than that of GOx nonspecifically bound on gold. The protein immobilized on the surface was utilized as a glucose sensor with millimolar sensitivity as the catalytic current increases linearly with increasing analyte concentration.

More recently, we extended this platform utility to immobilize aptamers for sensor applications.<sup>108</sup> By taking advantage of the fact that CB[7] readily forms a self-assembled monolayer (CB[7]SAM) on gold,<sup>109</sup> we developed a new SPR-based apta-sensor chip which can be easily prepared by simple addition of FA-aptamers to CB[7]SAM (Fig. 6). A concentration dependent sensor response was observed when a target protein (thrombin) was injected into a channel of the sensor chip; a nanomolar (sub-ppm) level of sensitivity to thrombin was achieved. Bruns-veld and others utilized the same strategy to immobilize different proteins and tested their applications (Fig. 6).<sup>110–112</sup> Their method allowed us to print stable protein monolayers in well-defined formats with controlled protein orientation and its release from the surface by treating an excess amount of a competitive guest such as (ferrocenylmethyl)trimethylammonium iodide. There is a recent report on supramolecular control of endothelial cell adhesion *via* synthetic integrin binding RGD peptide-ferrocene (Fc-cRGD) conjugates that were immobilized onto CB[7] coated gold surfaces.<sup>111</sup> Supramolecular coating with Fc-cRGD resulted in efficient adhesion and growth of human umbilical vein endothelial cells (hUVECs) on the surfaces as evidenced by the increase in the surface coverage and cell density after 24 h. They also achieved patterned adhesion of hUVECs using glass substrates patterned with gold arrays. In general, these reports suggest that CB[7]-FA pairs can be applied to reversible immobilization of different biomolecules on solid surfaces.

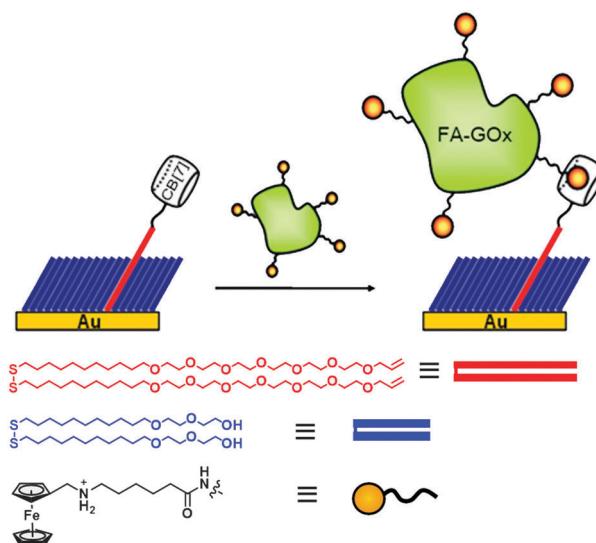


Fig. 5 Immobilization of ferrocenylated proteins on a gold surface using an ultrastable synthetic binding pair system. Reprinted with permission from ref. 80. Copyright 2015 American Chemical Society.

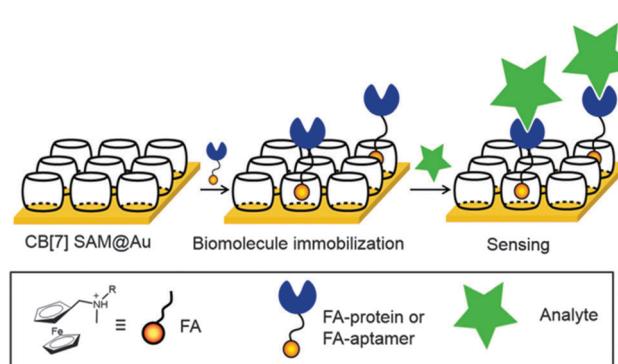


Fig. 6 General strategy for the immobilization of biomolecules on a self-assembled monolayer surface using CB[7]. Reprinted with permission from ref. 108. Copyright 2015 The Royal Society of Chemistry.



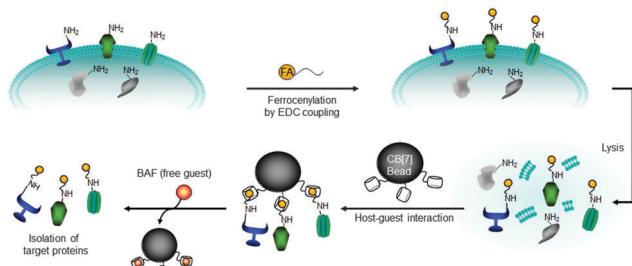


Fig. 7 Strategy for the isolation of plasma membrane proteins using an ultrastable synthetic host–guest binding pair system. Reprinted with permission from ref. 90. Copyright 2011, Nature Publishing Group.

### Plasma membrane protein fishing

The isolation and characterization of plasma membrane proteins (PMPs) and their functions are critical for understanding cell signaling and the effects of stimulation with various signal molecules. However, it is highly challenging because of PMPs inherent hydrophobicity and heterogeneity.<sup>113</sup> A well-known method for isolation of membrane proteins is *via* the use of affinity enrichment, where cell surface membrane proteins are biotinylated on amino acid residues located in exterior domains and subsequently enriched using streptavidin beads.<sup>16,114,115</sup> However, this method has some intrinsic drawbacks, including low yield, contamination with endogenously biotinylated proteins, and interference by streptavidin detached from beads.<sup>15–17,116–118</sup> There have been new approaches for protein purification using synthetic host–guest pairs but they suffered low binding affinities ( $K < 10^5 \text{ m}^{-1}$ ) and poor selectivity.<sup>119,120</sup> To address these issues, we developed a new supramolecular fishing method utilizing ultrastable and highly selective synthetic binding pair, a CB[7]–FA system (Fig. 7).<sup>90</sup>

CB[7]-conjugated sepharose beads (CB[7] beads) were utilized to capture FA-labelled proteins selectively from heterogeneous protein mixtures of various mammalian cells including rat fibroblast (Rat-1) and HEK 293 cells. The capture efficiency for plasma membrane proteins (EGFR, InsR, caveolin and integrin alpha 6) is much higher compared to cytoplasmic proteins (histone H3, cyclin B, Cpn60 and AMPK). In addition, the CB[7] system is not affected by the contaminations from endogenously biotinylated proteins and the captured proteins were easily recovered from the beads at room temperature by treating a strong competitor such as 1',1-bis(trimethylammoniomethyl)ferrocene (BAF), which cannot be achieved by the streptavidin-biotin-based method. These results demonstrated another promising biological application of the ultrastable host–guest pair system – efficient isolation and purification of proteins. This work may be extended to the selective isolation of cytoplasmic proteins, which is currently under investigation.

### Supramolecular Velcro system

Another interesting application of high affinity host–guest pairs is the development of an underwater adhesive system. There are growing needs for new adhesion systems that have high strength in water, repeated attachment and switchability on

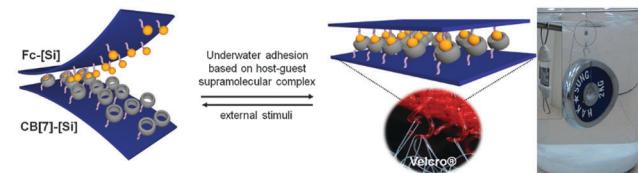


Fig. 8 Supramolecular Velcro® or hook-and-loop strategy for underwater adhesion based on CB[7]- and FA-modified surfaces (above). Reprinted with permission from ref. 96. Copyright 2013 Wiley-VCH Verlag.

demand for biological applications such as surgery sutures and repairing living tissue. Previous efforts to create underwater adhesives have largely focused on biomimetic polymers that are secreted by marine organisms such as mussels, sandcastle worms, barnacles, and sea urchins.<sup>121–125</sup> However, these efforts require curing agents and/or covalent crosslinking that irreversibly crosslink the residues to the surfaces. Another significant problem in developing underwater adhesive surfaces is how to expel the water molecules from the region between the two interacting surfaces. Since the formation of extremely stable complexes between CB[7] and hydrophobic guests is strongly associated with the transfer of water molecules to the bulk solution,<sup>47</sup> we utilized the ultrastable host–guest system to develop functionalized surfaces with strong underwater adhesion properties, which work like a hook and loop fastener, Velcro®. Our system comprises of two complementary silicon wafer surfaces: a ‘loop’ surface that is functionalized with CB[7]; and a ‘hook’ surface that is covered with FA molecules (Fig. 8).<sup>96</sup>

We deposited a layer of branched polyethyleneimine (PEI) on epoxy-functionalized silicon surfaces to generate an amine-terminated surface, which in turn reacts with ferrocene-carboxaldehyde to form ferrocenylmethylamino (FA) groups. A complementary surface was prepared by reaction of thiol groups on surface with monoallyloxy-CB[7] *via* thiol–allyl click reaction. When the two complementary surfaces are pushed together under water, the FA moiety binds tightly inside the CB[7] cavity which is enough to “glue” them together with enough strength to lift a 2 kg weight. The lap shear adhesion strength in air is much higher than that of commercially available adhesives, including double-sided tapes and traditional Velcro®. Interestingly, the adhesive strength can be tuned by decreasing the surface density of FA on one of the interacting surfaces. Unlike existing adhesives, our supramolecular Velcro is not only mechanically reversible but also chemically switchable. The binding affinity of the FA@CB[7] complex can be modulated by the oxidation of the ferrocenyl group.<sup>126</sup> The surfaces can be either mechanically peeled apart or separated chemically with sodium hypochlorite. Exposure to oxidizing agents decreases the adhesive strength between the two surfaces. However, after exposure to a reducing agent, ascorbic acid, the adhesive strength is mostly recovered, although it does not reach the original level, probably because of the partial decomposition of the ferrocenyl groups in the oxidized state. The redox properties of the CB[7]–FA complex afford an interesting method to modulate the strength of adhesion, which is not available in any other adhesive systems.

However, our material lost some of its adhesion strength after repeated cycles of binding and chemical unbinding. We think that an electrochemical method may offer a more efficient alternative to a chemical redox agent for switchable underwater adhesion. Recently, Chinese researchers reported another underwater adhesive based on a similar principle using the FA and  $\beta$ -CD systems.<sup>127</sup> Overall, our work translates a molecular recognition event into macroscopic adhesion between surfaces that function in water. The non-cytotoxic nature of these host and guest molecules permits potential biological applications, but an *in vivo* adhesion test is required to strengthen our claims.

### Development of a high-affinity host–guest FRET pair and its application in single vesicle content mixing assays

Fluorescence resonance energy transfer (FRET) has been widely used as a spectroscopic technique in various areas such as structural elucidation of biological molecules and their interactions, *in vitro* assays, and *in vivo* monitoring in cellular research. Among a series of FRET systems, host–guest FRET pairs using  $\beta$ -CD have recently attracted considerable attention as chemical sensors and biosensors.<sup>128–130</sup> However, because of the moderate affinity between  $\beta$ -CD and guest molecules, it is not desired to use these FRET pairs as a single molecule analysis probe. Therefore, developing a host–guest FRET system with an exceptional affinity may greatly enhance the chances of exploring these pairs in single molecule analysis. Very recently, we developed a high-affinity host–guest FRET pair by utilizing the ultrastable CB[7]–adamantylamine complex.<sup>131</sup> We synthesized the fluorophore-tagged CB[7]-Cy3 and Ad-Cy5, found that these dye-conjugated CB[7]-Ad pairs produced high FRET signals in aqueous solution upon host–guest binding (Fig. 9).

We utilized this host–guest FRET pair to study the challenging SNARE (soluble *N*-ethylmaleimide-sensitive factor attachment protein receptors) mediated membrane fusion. Protein-mediated membrane fusion is a common biological process and plays a principal role in a variety of cellular activities, such as vesicle trafficking, viral infection, exocytosis and neurotransmitter.<sup>132,133</sup> A majority of vesicle fusion is mediated by SNAREs which act as molecular tethers between the two fusing bodies. Single vesicle *in vitro* fusion assays using SNARE-reconstituted vesicles have played a vital role in elucidating the mechanisms of SNARE-mediated vesicle fusion by allowing controlled observations which

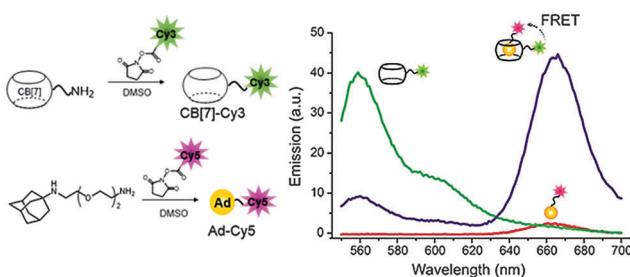


Fig. 9 Synthesis and characterization of CB[7]-Cy3 and Ad-Cy5 high affinity FRET pair. Reprinted with permission from ref. 131. Copyright 2015 American Chemical Society.

are not possible by *in vivo* studies.<sup>134–137</sup> However, the existing assays have limitations such as low sensitivity (signal to noise ratio) and limited ability to detect transient fusion pore openings.

We thus decided to develop a novel FRET-based single-vesicle content assay using our host–guest FRET pair, which combines the benefits of small molecular probes with the high signal-to-noise ratio benefits of FRET based techniques, to study the mechanism of synaptic vesicle fusion mediated by SNARE proteins (Fig. 10).<sup>131</sup> We first encapsulated CB[7]-Cy3 and Ad-Cy5 in SNARE protein reconstituted donor and accepter vesicles (v-vesicle and t-vesicle, respectively) to produce CB[7]-Cy3@v-vesicle and Ad-Cy5@t-vesicle, respectively. We then performed a single-vesicle content mixing assay with these vesicles using a total internal reflection fluorescence microscope (TIRF). The acceptor vesicle, Ad-Cy5@t-vesicle, was immobilized on a quartz surface and a donor vesicle, CB[7]-Cy3@v-vesicle, was injected into the flow chamber of our analysis set-up. Docking of a donor vesicle onto the acceptor vesicle occurred as evidenced by the sudden appearance of Cy3 emission in the time trace. The pore opening led to mixing of the content molecules of the two opposite vesicles to form a strong host–guest complex, which generated the FRET signal of Cy5 emission. The high stability of the CB[7]-Ad complex allowed our technique to be compatible with a wide range of reaction conditions during both vesicle preparation and fusion.

The most important finding of our assay is the direct detection of multistep fusion flickering events, for the first time, in an *in vitro* content mixing assay. It is generally believed that SNARE fusion proceeds through intermediates, which may lead directly to the formation of a fusion pore followed by expansion of the pore or it can open and close in a transient manner (flickering). The merits of our system allowed us to observe the fusion pore flickering events and the intermediate

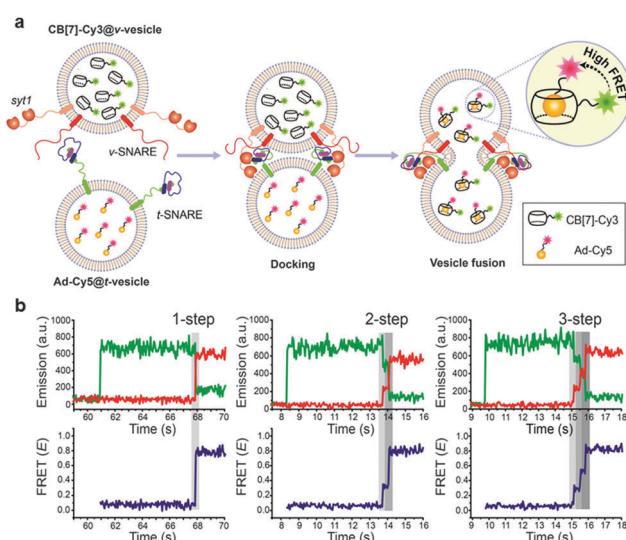


Fig. 10 Schematic illustration of the SNARE-mediated content mixing using a high affinity host–guest binding FRET pair (a) and observation of multi-step fusion pore opening and closure dynamics (b). Reprinted with permission from ref. 131. Copyright 2015 American Chemical Society.



steps involved in the fusion process, which was unsuccessful with existing techniques.<sup>137,138</sup> In major traces, full fusion was observed with the appearance of a plateau corresponding to the FRET efficiency of complete content mixing without any intermediates. In other cases, however, the initial pore formation state is followed by one or more additional plateaus with a FRET efficiency that is intermediate between initial docking and full fusion. This is a characteristic behavior of flickering in which a transient fusion pore opens and closes quickly and hence does not allow full mixing of the contents at once.

This work demonstrates the power of the synthetic system in the field of complex biochemical analysis. Furthermore, this powerful tool may allow us to investigate further to clarify the roles of individual components in the membrane fusion, which is difficult to achieve at the cellular level. This synthetic “host-guest FRET pair” with extremely strong binding affinity and specificity may be applied to various chemical and biological studies as a supramolecular beacon.

### Recognition-mediated activation of therapeutic gold nanoparticles inside cells

Apart from the utility of high affinity host-guest pairs in the separation of proteins and vesicle fusion, Rotello, Isaacs and co-workers exploited it as a trigger in therapeutic systems. Their system consists of complementary diaminohexane-terminated gold nanoparticles (AuNP-NH<sub>2</sub>) and CB[7] to form a non-toxic nanoparticle assembly that is readily taken up by cells (Fig. 11).<sup>139</sup> Intracellular removal of CB[7] from the nanoparticle surface using an orthogonal competitor guest A2 results in the

endosomal escape of AuNP-NH<sub>2</sub>, which activates the cytotoxicity of the amine-tethered gold nanoparticles and thereby causes cell death. This strategy for triggering therapeutic systems has some advantages in dosage control and potential utility in dual-targeting therapies. The authors used ~2.5 nm sized AuNP-NH<sub>2</sub> as a therapeutic component of the system, which features a self-assembled monolayer of diaminohexane-terminated thiol ligands. It is believed that the terminal diaminohexane moiety renders cytotoxicity to the nanoparticle and also serves as a recognition unit for the formation of a host-guest inclusion complex with CB[7]. Cells treated with AuNP-NH<sub>2</sub> and CB[7]-AuNP-NH<sub>2</sub> alone exhibited 34% and 100% cell viability after 24 h of incubation, respectively. The addition of A2 resulted in the formation of more favorable A2@CB[7] complexes, which triggered the release of CB[7] from the nanoparticles. After 24 h of incubation, AuNP-NH<sub>2</sub> was dispersed in the cytosol whereas CB[7]-AuNP-NH<sub>2</sub> remained in the endosome. Remarkably, the incubation of CB[7]-AuNP-NH<sub>2</sub> treated cells with A2 resulted in the escape of a substantial number of nanoparticles into the cytosol from the endosome. This result demonstrates the use of synthetic host-guest chemistry to provide triggered activation of a therapeutic system by competitive complexation and provides a potential strategy for the construction of synthetic host-guest systems using CB[7], which has an ability to form a strong complex within living cells. This is an interesting engineering of host-guest systems for the regulation of therapeutics that can function inside the cell. Based on the same principle, the authors also reported the regulation of exocytosis of nanoparticles using CB[7]-based host-guest chemistry.<sup>140</sup> This approach provides a potential strategy for prolonged retention of drug carriers within endosomes, enabling a sustained therapeutic effect of the carriers.

### Biological catalysis regulated by CB[7]

It is important to note that the high affinity and selectivity of CB[7] towards selected guest molecules can be comparable to those typically observed in the interaction of proteins and antibodies toward their cognate ligands.<sup>141</sup> By taking this merit into consideration, Isaacs and co-workers demonstrated the influence of CB[7] on the behavior of biological receptors toward their guests and thereby change in their biological activity (Fig. 12).<sup>142</sup> They explored the self-sorting capability of two-faced guests for the regulation of the catalytic activity of bovine carbonic anhydrase (BCA). The active site of the enzyme was initially occupied by a two-faced inhibitor that contains benzenesulfonamide (enzyme inhibitor unit) and trimethylsilylmethylammonium group (CB[7] binding unit). The initial enzymatic activity was recovered by the addition of CB[7], which forms a 1:1 host-guest complex with trimethylsilylmethylammonium part ( $K_a \sim 10^9 \text{ M}^{-1}$ ) of the two-faced guests to release the inhibitor from the enzyme. However, the enzyme activity was inhibited again by the addition of adamantaneammonium derivative, which triggered the dissociation of the CB[7]-enzyme inhibitor resulting in the release of the free two-faced guest, which then re-occupied the catalytic site of the enzyme. Addition of CB[7] and A2 in an alternating sequence turned the

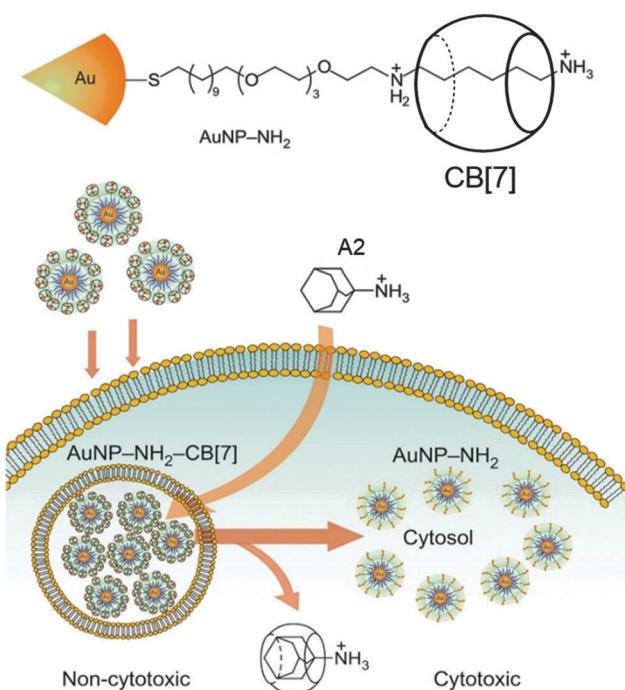


Fig. 11 Schematic illustration of the use of intracellular host-guest complexation to trigger gold nanoparticle (AuNP-NH<sub>2</sub>) cytotoxicity. Reprinted with permission from ref. 139. Copyright 2010, Nature Publishing Group.



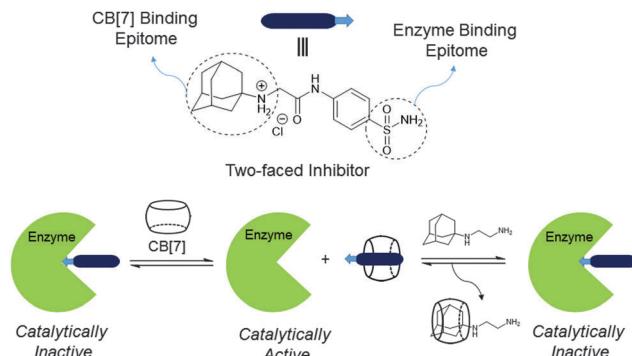


Fig. 12 Schematic illustration of the control of enzymatic activity using CB[7] and an adamantane derivative. Reprinted with permission from ref. 142. Copyright 2010 American Chemical Society.

enzymatic activity off and on over several additional cycles. One of the key controlling factors in this process is the relative binding affinity of the two-faced inhibitors toward enzyme *versus* CB[7]. The interesting aspect of this work is the ability of CB[7] to enhance the rate of dissociation of the enzyme-inhibitor by the transient formation of the enzyme-inhibitor-CB[7] ternary complexes. This strategy differs from the typical allosteric regulation of enzyme activity, in that direct competition between two receptors (enzyme and CB[7]) is involved. Overall, this finding demonstrates the advantage of CB[7] selectivity towards specific guest molecules and hence it may be possible to use this receptor to compete with and thereby control the activity of specific portions of the biological interaction networks.

Very recently, Rotello and co-workers have also explored the benefit of high affinity CB[7]-Ad pairs for the supramolecular regulation of bioorthogonal catalysis in cells.<sup>143</sup> They developed a protein-sized bioorthogonal nanzyme through the encapsulation of hydrophobic transition metal catalysts into the monolayer of water-soluble gold nanoparticles. The activity of the catalyst was reversibly controlled by using CB[7] as the 'gate-keeper' onto the monolayer surface which provides a biomimetic control mechanism that mimics the allosteric regulation of natural enzymes.

## 5. Conclusions

This review summarizes the high affinity host-guest interactions between CB[7] and various guest molecules, as well as their applications in constructing various supramolecular systems. We have described the evolution of different guest molecules that bind CB[7] to water with affinities equivalent to or higher than those of the highest-affinity protein-small molecule systems. In general, the guest binding affinity and selectivity of CB[7] are higher than those of any other synthetic receptors reported to date. These high affinity pairs offer numerous advantages including an attomolar dissociation constant and tunable complexation kinetics over other conventional macrocycles. CB[7] shows remarkable binding characteristics based on the guest size, shape, length, and chemical functionalities. The positive entropy changes

from the extensive dehydration of the host and guest upon complexation and thereby overcoming the enthalpy-entropy compensation critically assists in achieving high binding affinity. The ultrastable host-guest complexes are unique as they rely on monovalency to achieve such high binding affinity. While there have been continuing efforts to develop new high affinity guest molecules, their implementation in challenging applications has rapidly unfolded during the past decade. The inherent high affinity of CB[7] paves its way to deal with challenging applications in materials science, biology, and medicine. With exceptional stability, chemical robustness, biocompatibility, and easy handling, these complexes may replace the biotin-(strept)avidin system in diverse areas of research and therefore the application perspectives of these complexes are looking very bright. Overall, CB[7] brings synthetic receptor systems very close to their natural counterparts with new hope in the development of novel synthetic systems which function equally well or better than their biological counterparts. Furthermore, the recent progress in this field clearly suggests that the strong host-guest complexes are making the transition from a simple research interest to modular components of chemical systems with important practical applications.

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## References

- 1 M. J. Swamy, *Biochem. Mol. Biol. Int.*, 1995, **36**, 219–225.
- 2 B. Kuhn and P. A. Kollman, *J. Am. Chem. Soc.*, 2000, **122**, 3909–3916.
- 3 N. M. Green, *Methods Enzymol.*, 1990, **184**, 51–67.
- 4 M. Wilchek and E. A. Bayer, *Methods Enzymol.*, 1990, **184**, 14–45.
- 5 T. J. Park, K. B. Lee, S. J. Lee, J. P. Park, Z. W. Lee, S. K. Choi, H. C. Jung, J. G. Pan, S. Y. Lee and I. S. Choi, *J. Am. Chem. Soc.*, 2004, **126**, 10512–10513.
- 6 J. Ladd, C. Boozer, Q. Yu, S. Chen, J. Homola and S. Jiang, *Langmuir*, 2004, **20**, 8090–8095.
- 7 S. C. Tao and H. Zhu, *Nat. Biotechnol.*, 2006, **24**, 1253–1254.
- 8 O. H. Laitinen, H. R. Nordlund, V. P. Hytonen and M. S. Kulomaa, *Trends Biotechnol.*, 2007, **25**, 269–277.
- 9 G. N. Ostojic and M. C. Hersam, *Small*, 2012, **8**, 1840–1845.
- 10 A. Chilkoti, P. H. Tan and P. S. Stayton, *Proc. Natl. Acad. Sci. U. S. A.*, 1995, **92**, 1754–1758.
- 11 O. H. Laitinen, K. J. Airenne, A. T. Marttila, T. Kulik, E. Porkka, E. A. Bayer, M. Wilchek and M. S. Kulomaa, *FEBS Lett.*, 1999, **461**, 52–58.
- 12 A. T. Marttila, O. H. Laitinen, K. J. Airenne, T. Kulik, E. A. Bayer, M. Wilchek and M. S. Kulomaa, *FEBS Lett.*, 2000, **467**, 31–36.



13 O. H. Laitinen, H. R. Nordlund, V. P. Hytonen, S. T. Uotila, A. T. Marttila, J. Savolainen, K. J. Airenne, O. Livnah, E. A. Bayer, M. Wilchek and M. S. Kulomaa, *J. Biol. Chem.*, 2003, **278**, 4010–4014.

14 M. Howarth, D. J. Chinnapan, K. Gerrow, P. C. Dorrestein, M. R. Grandy, N. L. Kelleher, A. El-Husseini and A. Y. Ting, *Nat. Methods*, 2006, **3**, 267–273.

15 S. R. Jaffrey, H. Erdjument-Bromage, C. D. Ferris, P. Tempst and S. H. Snyder, *Nat. Cell Biol.*, 2001, **3**, 193–197.

16 W. Zhang, G. Zhou, Y. Zhao, M. A. White and Y. Zhao, *Electrophoresis*, 2003, **24**, 2855–2863.

17 R. Takechi, A. Taniguchi, S. Ebara, T. Fukui and T. Watanabe, *J. Nutr.*, 2008, **138**, 680–684.

18 W. L. Mock, in *Comprehensive Supramolecular Chemistry, Molecular Recognition: Receptors for Molecular Guests*, ed. F. Vögtle, Pergamon Press, Oxford, UK, 1996, vol. 2, pp. 477–493.

19 C. L. Gibb and B. C. Gibb, *J. Am. Chem. Soc.*, 2004, **126**, 11408–11409.

20 A. V. Davis and K. N. Raymond, *J. Am. Chem. Soc.*, 2005, **127**, 7912–7919.

21 J. Rebek, Jr., *Angew. Chem., Int. Ed.*, 2005, **44**, 2068–2078.

22 K. Takaoka, M. Kawano, M. Tominaga and M. Fujita, *Angew. Chem., Int. Ed.*, 2005, **44**, 2151–2154.

23 R. Custelcean, *Chem. Soc. Rev.*, 2014, **43**, 1813–1824.

24 F. L. Dickert, *Sensors*, 2014, **14**, 22525–22531.

25 W. L. Mock, *Top. Curr. Chem.*, 1995, **175**, 1–24.

26 J. W. Lee, S. Samal, N. Selvapalam, H. J. Kim and K. Kim, *Acc. Chem. Res.*, 2003, **36**, 621–630.

27 J. Lagona, P. Mukhopadhyay, S. Chakrabarti and L. Isaacs, *Angew. Chem., Int. Ed.*, 2005, **44**, 4844–4870.

28 K. Kim, N. Selvapalam, Y. H. Ko, K. M. Park, D. Kim and J. Kim, *Chem. Soc. Rev.*, 2007, **36**, 267–279.

29 X. J. Cheng, L. L. Liang, K. Chen, N. N. Ji, X. Xiao, J. X. Zhang, Y. Q. Zhang, S. F. Xue, Q. J. Zhu, X. L. Ni and Z. Tao, *Angew. Chem., Int. Ed.*, 2013, **52**, 7252–7255.

30 A. E. Kaifer, W. Li and S. Yi, *Isr. J. Chem.*, 2011, **51**, 496–505.

31 E. Masson, X. Ling, R. Joseph, L. Kyeremeh-Mensah and X. Lu, *RSC Adv.*, 2012, **2**, 1213–1247.

32 J. Lü, J.-X. Lin, M.-N. Cao and R. Cao, *Coord. Chem. Rev.*, 2013, **257**, 1334–1356.

33 X. L. Ni, X. Xiao, H. Cong, L. L. Liang, K. Cheng, X. J. Cheng, N. N. Ji, Q. J. Zhu, S. F. Xue and Z. Tao, *Chem. Soc. Rev.*, 2013, **42**, 9480–9508.

34 L. Isaacs, *Acc. Chem. Res.*, 2014, **47**, 2052–2062.

35 X. L. Ni, X. Xiao, H. Cong, Q. J. Zhu, S. F. Xue and Z. Tao, *Acc. Chem. Res.*, 2014, **47**, 1386–1395.

36 L. Peng, A. Feng, M. Huo and J. Yuan, *Chem. Commun.*, 2014, **50**, 13005–13014.

37 N. Vallavoju and J. Sivaguru, *Chem. Soc. Rev.*, 2014, **43**, 4084–4101.

38 H. Yang, B. Yuan, X. Zhang and O. A. Scherman, *Acc. Chem. Res.*, 2014, **47**, 2106–2115.

39 K. I. Assaf and W. M. Nau, *Chem. Soc. Rev.*, 2015, **44**, 394–418.

40 W. S. Jeon, K. Moon, S. H. Park, H. Chun, Y. H. Ko, J. Y. Lee, E. S. Lee, S. Samal, N. Selvapalam, M. V. Rekharsky, V. Sindelar, D. Sobrarsingh, Y. Inoue, A. E. Kaifer and K. Kim, *J. Am. Chem. Soc.*, 2005, **127**, 12984–12989.

41 S. Liu, C. Ruspic, P. Mukhopadhyay, S. Chakrabarti, P. Y. Zavalij and L. Isaacs, *J. Am. Chem. Soc.*, 2005, **127**, 15959–15967.

42 M. V. Rekharsky, T. Mori, C. Yang, Y. H. Ko, N. Selvapalam, H. Kim, D. Sobrarsingh, A. E. Kaifer, S. Liu, L. Isaacs, W. Chen, S. Moghaddam, M. K. Gilson, K. Kim and Y. Inoue, *Proc. Natl. Acad. Sci. U. S. A.*, 2007, **104**, 20737–20742.

43 S. Moghaddam, C. Yang, M. Rekharsky, Y. H. Ko, K. Kim, Y. Inoue and M. K. Gilson, *J. Am. Chem. Soc.*, 2011, **133**, 3570–3581.

44 L. Cao, M. Šekutor, P. Y. Zavalij, K. Mlinarić-Majerski, R. Glaser and L. Isaacs, *Angew. Chem., Int. Ed.*, 2014, **53**, 988–993.

45 M. V. Rekharsky and Y. Inoue, *Chem. Rev.*, 1998, **98**, 1875–1918.

46 W. Ong and A. E. Kaifer, *Organometallics*, 2003, **22**, 4181–4183.

47 F. Biedermann, V. D. Uzunova, O. A. Scherman, W. M. Nau and A. De Simone, *J. Am. Chem. Soc.*, 2012, **134**, 15318–15323.

48 C. N. Nguyen, T. K. Young and M. K. Gilson, *J. Chem. Phys.*, 2012, **137**, 044101.

49 F. Biedermann, M. Vendruscolo, O. A. Scherman, A. De Simone and W. M. Nau, *J. Am. Chem. Soc.*, 2013, **135**, 14879–14888.

50 F. Biedermann, W. M. Nau and H. J. Schneider, *Angew. Chem., Int. Ed.*, 2014, **53**, 11158–11171.

51 Y. Chen, A. Klimeczak, E. Galoppini and J. V. Lockard, *RSC Adv.*, 2013, **3**, 1354–1358.

52 M. Šekutor, K. Molčanov, L. Cao, L. Isaacs, R. Glaser and K. Mlinarić-Majerski, *Eur. J. Org. Chem.*, 2014, 2533–2542.

53 C. Marquez and W. M. Nau, *Angew. Chem., Int. Ed.*, 2001, **40**, 4387–4390.

54 Y. Pazy, T. Kulik, E. A. Bayer, M. Wilchek and O. Livnah, *J. Biol. Chem.*, 2002, **277**, 30892–30900.

55 C. Márquez, R. R. Hudgins and W. M. Nau, *J. Am. Chem. Soc.*, 2004, **126**, 5806–5816.

56 W. Ong and A. E. Kaifer, *J. Org. Chem.*, 2004, **69**, 1383–1385.

57 C. Marquez and W. M. Nau, *Angew. Chem., Int. Ed.*, 2001, **40**, 4387–4390.

58 W. M. Nau, A. Hennig and A. L. Koner, *Fluorescence of Supermolecules, Polymers, and Nanosystems*, Springer, 2008, pp. 185–211.

59 W. M. Nau, M. Florea and K. I. Assaf, *Isr. J. Chem.*, 2011, **51**, 559–577.

60 D. Sobrarsingh and A. E. Kaifer, *Org. Lett.*, 2006, **8**, 3247–3250.

61 M. Rekharsky and Y. Inoue, *J. Am. Chem. Soc.*, 2000, **122**, 4418–4435.

62 C.-E. Chang and M. K. Gilson, *J. Am. Chem. Soc.*, 2004, **126**, 13156–13164.



63 S. Moghaddam, Y. Inoue and M. K. Gilson, *J. Am. Chem. Soc.*, 2009, **131**, 4012–4021.

64 A. T. Fenley, N. M. Henriksen, H. S. Muddana and M. K. Gilson, *J. Chem. Theory Comput.*, 2014, **10**, 4069–4078.

65 S. J. C. Lee, J. W. Lee, H. H. Lee, J. Seo, D. H. Noh, Y. H. Ko, K. Kim and H. I. Kim, *J. Phys. Chem. B*, 2013, **117**, 8855–8864.

66 W. Saenger, *Angew. Chem., Int. Ed.*, 1980, **19**, 344–362.

67 H. J. Schneider, *Angew. Chem., Int. Ed.*, 1991, **30**, 1417–1436.

68 A. Flinn, G. C. Hough, J. F. Stoddart and D. J. Williams, *Angew. Chem., Int. Ed.*, 1992, **31**, 1475–1477.

69 A. Day, A. P. Arnold, R. J. Blanch and B. Snushall, *J. Org. Chem.*, 2001, **66**, 8094–8100.

70 J. Zhao, H. J. Kim, J. Oh, S. Y. Kim, J. W. Lee, S. Sakamoto, K. Yamaguchi and K. Kim, *Angew. Chem., Int. Ed.*, 2001, **40**, 4233–4235.

71 A. I. Day, R. J. Blanch, A. Coe and A. P. Arnold, *J. Inclusion Phenom. Macrocyclic Chem.*, 2002, **43**, 247–250.

72 H. Isobe, S. Sato and E. Nakamura, *Org. Lett.*, 2002, **4**, 1287–1289.

73 C. A. Burnett, J. Lagona, A. Wu, J. A. Shaw, D. Coady, J. C. Fettinger, A. I. Day and L. Isaacs, *Tetrahedron*, 2003, **59**, 1961–1970.

74 C. A. Burnett, D. Witt, J. C. Fettinger and L. Isaacs, *J. Org. Chem.*, 2003, **68**, 6184–6191.

75 A. I. Day, A. P. Arnold and R. J. Blanch, *Molecules*, 2003, **8**, 74–84.

76 K. Jansen, A. Wego, H.-J. Buschmann, E. Schollmeyer and D. Döpp, *Des. Monomers Polym.*, 2003, **6**, 43–55.

77 S. Y. Jon, N. Selvapalam, D. H. Oh, J. K. Kang, S. Y. Kim, Y. J. Jeon, J. W. Lee and K. Kim, *J. Am. Chem. Soc.*, 2003, **125**, 10186–10187.

78 S. Sasmal, M. K. Sinha and E. Keinan, *Org. Lett.*, 2004, **6**, 1225–1228.

79 E. R. Nagarajan, D. H. Oh, N. Selvapalam, Y. H. Ko, K. M. Park and K. Kim, *Tetrahedron Lett.*, 2006, **47**, 2073–2075.

80 I. Hwang, K. Baek, M. Jung, Y. Kim, K. M. Park, D. W. Lee, N. Selvapalam and K. Kim, *J. Am. Chem. Soc.*, 2007, **129**, 4170–4171.

81 D. Kim, E. Kim, J. Kim, K. M. Park, K. Baek, M. Jung, Y. H. Ko, W. Sung, H. S. Kim, J. H. Suh, C. G. Park, O. S. Na, D. K. Lee, K. E. Lee, S. S. Han and K. Kim, *Angew. Chem., Int. Ed.*, 2007, **46**, 3471–3474.

82 J. Kim, Y. Ahn, K. M. Park, Y. Kim, Y. H. Ko, D. H. Oh and K. Kim, *Angew. Chem., Int. Ed.*, 2007, **46**, 7393–7395.

83 K. M. Park, K. Suh, H. Jung, D. W. Lee, Y. Ahn, J. Kim, K. Baek and K. Kim, *Chem. Commun.*, 2009, 71–73.

84 K. Baek, Y. Kim, H. Kim, M. Yoon, I. Hwang, Y. H. Ko and K. Kim, *Chem. Commun.*, 2010, **46**, 4091–4093.

85 D. Kim, E. Kim, J. Lee, S. Hong, W. Sung, N. Lim, C. G. Park and K. Kim, *J. Am. Chem. Soc.*, 2010, **132**, 9908–9919.

86 E. Kim, D. Kim, H. Jung, J. Lee, S. Paul, N. Selvapalam, Y. Yang, N. Lim, C. G. Park and K. Kim, *Angew. Chem., Int. Ed.*, 2010, **49**, 4405–4408.

87 S. K. Kim, K. M. Park, K. Singha, J. Kim, Y. Ahn, K. Kim and W. J. Kim, *Chem. Commun.*, 2010, **46**, 692–694.

88 K. M. Park, D. W. Lee, B. Sarkar, H. Jung, J. Kim, Y. H. Ko, K. E. Lee, H. Jeon and K. Kim, *Small*, 2010, **6**, 1430–1441.

89 H. Jung, K. M. Park, J. A. Yang, E. J. Oh, D. W. Lee, K. Park, S. H. Ryu, S. K. Hahn and K. Kim, *Biomaterials*, 2011, **32**, 7687–7694.

90 D. W. Lee, K. M. Park, M. Banerjee, S. H. Ha, T. Lee, K. Suh, S. Paul, H. Jung, J. Kim, N. Selvapalam, S. H. Ryu and K. Kim, *Nat. Chem.*, 2011, **3**, 154–159.

91 K. M. Park, J. A. Yang, H. Jung, J. Yeom, J. S. Park, K. H. Park, A. S. Hoffman, S. K. Hahn and K. Kim, *ACS Nano*, 2012, **6**, 2960–2968.

92 K. Baek, G. Yun, Y. Kim, D. Kim, R. Hota, I. Hwang, D. Xu, Y. H. Ko, G. H. Gu, J. H. Suh, C. G. Park, B. J. Sung and K. Kim, *J. Am. Chem. Soc.*, 2013, **135**, 6523–6528.

93 G. Yun, Z. Hassan, J. Lee, J. Kim, N. S. Lee, N. H. Kim, K. Baek, I. Hwang, C. G. Park and K. Kim, *Angew. Chem., Int. Ed.*, 2014, **53**, 6414–6418.

94 J. Kim, K. Baek, D. Shetty, N. Selvapalam, G. Yun, N. H. Kim, Y. H. Ko, K. M. Park, I. Hwang and K. Kim, *Angew. Chem., Int. Ed.*, 2015, **54**, 2693–2697.

95 J. Yeom, S. J. Kim, H. Jung, H. Namkoong, J. Yang, B. W. Hwang, K. Oh, K. Kim, Y. C. Sung and S. K. Hahn, *Adv. Healthcare Mater.*, 2015, **4**, 237–244.

96 Y. Ahn, Y. Jang, N. Selvapalam, G. Yun and K. Kim, *Angew. Chem., Int. Ed.*, 2013, **52**, 3140–3144.

97 B. Vinciguerra, L. Cao, J. R. Cannon, P. Y. Zavalij, C. Fenselau and L. Isaacs, *J. Am. Chem. Soc.*, 2012, **134**, 13133–13140.

98 M. M. Ayhan, H. Karoui, M. Hardy, A. Rockenbauer, L. Charles, R. Rosas, K. Udachin, P. Tordo, D. Bardelang and O. Ouari, *J. Am. Chem. Soc.*, 2015, **137**, 10238–10245.

99 M. Dufva and C. B. Christensen, *Expert Rev. Proteomics*, 2005, **2**, 41–48.

100 S. F. Kingsmore, *Nat. Rev. Drug Discovery*, 2006, **5**, 310–320.

101 D. Weinrich, P. Jonkheijm, C. M. Niemeyer and H. Waldmann, *Angew. Chem., Int. Ed.*, 2009, **48**, 7744–7751.

102 M. B. Soellner, K. A. Dickson, B. L. Nilsson and R. T. Raines, *J. Am. Chem. Soc.*, 2003, **125**, 11790–11791.

103 P. C. Lin, S. H. Ueng, M. C. Tseng, J. L. Ko, K. T. Huang, S. C. Yu, A. K. Adak, Y. J. Chen and C. C. Lin, *Angew. Chem., Int. Ed.*, 2006, **45**, 4286–4290.

104 T. Govindaraju, P. Jonkheijm, L. Gogolin, H. Schroeder, C. F. Becker, C. M. Niemeyer and H. Waldmann, *Chem. Commun.*, 2008, 3723–3725.

105 L. S. Wong, F. Khan and J. Micklefield, *Chem. Rev.*, 2009, **109**, 4025–4053.

106 D. Weinrich, P. C. Lin, P. Jonkheijm, U. T. Nguyen, H. Schroder, C. M. Niemeyer, K. Alexandrov, R. Goody and H. Waldmann, *Angew. Chem., Int. Ed.*, 2010, **49**, 1252–1257.

107 Y. X. Chen, G. Triola and H. Waldmann, *Acc. Chem. Res.*, 2011, **44**, 762–773.

108 D. W. Lee, K. M. Park, B. Gong, D. Shetty, J. K. Khedkar, K. Baek, J. Kim, S. H. Ryu and K. Kim, *Chem. Commun.*, 2015, **51**, 3098–3101.



109 Q. An, G. Li, C. Tao, Y. Li, Y. Wu and W. Zhang, *Chem. Commun.*, 2008, 1989–1991.

110 J. F. Young, H. D. Nguyen, L. Yang, J. Huskens, P. Jonkheijm and L. Brunsved, *ChemBioChem*, 2010, **11**, 180–183.

111 P. Neirynck, J. Brinkmann, Q. An, D. W. van der Schaft, L. G. Milroy, P. Jonkheijm and L. Brunsved, *Chem. Commun.*, 2013, **49**, 3679–3681.

112 D. Wasserberg, D. A. Uhlenheuer, P. Neirynck, J. Cabanas-Danes, J. H. Schenkel, B. J. Ravoo, Q. An, J. Huskens, L. G. Milroy, L. Brunsved and P. Jonkheijm, *Int. J. Mol. Sci.*, 2013, **14**, 4066–4080.

113 D. Josic and J. G. Clifton, *Proteomics*, 2007, **7**, 3010–3029.

114 Y. Zhao, W. Zhang, Y. Kho and Y. Zhao, *Anal. Chem.*, 2004, **76**, 1817–1823.

115 G. Elia, *Proteomics*, 2008, **8**, 4012–4024.

116 L. Zhang, X. e. Wang, X. Peng, Y. Wei, R. Cao, Z. Liu, J. Xiong, X. Ying, P. Chen and S. Liang, *J. Proteome Res.*, 2007, **6**, 34–43.

117 M. Fonovic, S. H. Verhelst, M. T. Sorum and M. Bogyo, *Mol. Cell. Proteomics*, 2007, **6**, 1761–1770.

118 C. E. Chivers, E. Crozat, C. Chu, V. T. Moy, D. J. Sherratt and M. Howarth, *Nat. Methods*, 2010, **7**, 391–393.

119 T. Nguyen, N. S. Joshi and M. B. Francis, *Bioconjugate Chem.*, 2006, **17**, 869–872.

120 J. A. Chung, J. W. Wollack, M. L. Hovlid, A. Okesli, Y. Chen, J. D. Mueller, M. D. Distefano and T. A. Taton, *Anal. Biochem.*, 2009, **386**, 1–8.

121 H. Lee, B. P. Lee and P. B. Messersmith, *Nature*, 2007, **448**, 338–341.

122 R. Santos, G. da Costa, C. Franco, P. Gomes-Alves, P. Flammang and A. V. Coelho, *Mar. Biotechnol.*, 2009, **11**, 686–698.

123 R. J. Stewart and C. S. Wang, *Biomacromolecules*, 2010, **11**, 969–974.

124 Q. Ye, F. Zhou and W. Liu, *Chem. Soc. Rev.*, 2011, **40**, 4244–4258.

125 C. Zhong, T. Gurry, A. A. Cheng, J. Downey, Z. Deng, C. M. Stultz and T. K. Lu, *Nat. Nanotechnol.*, 2014, **9**, 858–866.

126 W. Li and A. E. Kaifer, *Langmuir*, 2012, **28**, 15075–15079.

127 J. Guo, C. Yuan, M. Guo, L. Wang and F. Yan, *Chem. Sci.*, 2014, **5**, 3261–3266.

128 T. Ogoshi and A. Harada, *Sensors*, 2008, **8**, 4961–4982.

129 P. Li, Y. Liu, X. Wang and B. Tang, *Analyst*, 2011, **136**, 4520–4525.

130 L. Feng, C. Tong, Y. He, B. Liu, C. Wang, J. Sha and C. Lü, *J. Lumin.*, 2014, **146**, 502–507.

131 B. Gong, B.-K. Choi, J.-Y. Kim, D. Shetty, Y. H. Ko, N. Selvapalam, N. K. Lee and K. Kim, *J. Am. Chem. Soc.*, 2015, **137**, 8908–8911.

132 R. Jahn and T. C. Sudhof, *Annu. Rev. Biochem.*, 1999, **68**, 863–911.

133 S. Martens and H. T. McMahon, *Nat. Rev. Mol. Cell Biol.*, 2008, **9**, 543–556.

134 J. A. McNew, F. Parlati, R. Fukuda, R. J. Johnston, K. Paz, F. Paumet, T. H. Sollner and J. E. Rothman, *Nature*, 2000, **407**, 153–159.

135 S. Martens, M. M. Kozlov and H. T. McMahon, *Science*, 2007, **316**, 1205–1208.

136 T. Ohya, M. Miaczynska, U. Coskun, B. Lommer, A. Runge, D. Drechsel, Y. Kalaidzidis and M. Zerial, *Nature*, 2009, **459**, 1091–1097.

137 M. Kyoung, A. Srivastava, Y. Zhang, J. Diao, M. Vrljic, P. Grob, E. Nogales, S. Chu and A. T. Brunger, *Proc. Natl. Acad. Sci. U. S. A.*, 2011, **108**, E304–E313.

138 J. Diao, Z. Su, Y. Ishitsuka, B. Lu, K. S. Lee, Y. Lai, Y. K. Shin and T. Ha, *Nat. Commun.*, 2010, **1**, 54.

139 C. Kim, S. S. Agasti, Z. Zhu, L. Isaacs and V. M. Rotello, *Nat. Chem.*, 2010, **2**, 962–966.

140 C. Kim, G. Y. Tonga, B. Yan, C. S. Kim, S. T. Kim, M.-H. Park, Z. Zhu, B. Duncan, B. Creran and V. M. Rotello, *Org. Biomol. Chem.*, 2015, **13**, 2474–2479.

141 K. N. Houk, A. G. Leach, S. P. Kim and X. Zhang, *Angew. Chem., Int. Ed.*, 2003, **42**, 4872–4897.

142 S. Ghosh and L. Isaacs, *J. Am. Chem. Soc.*, 2010, **132**, 4445–4454.

143 G. Y. Tonga, Y. Jeong, B. Duncan, T. Mizuhara, R. Mout, R. Das, S. T. Kim, Y.-C. Yeh, B. Yan, S. Hou and V. M. Rotello, *Nat. Chem.*, 2015, **7**, 597–603.

