Synthesis of magnetic mesoporous titania colloidal crystals through evaporation induced self-assembly in emulsion as effective and recyclable photocatalysts†

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This study illustrates the directed self-assembly of mesoporous TiO2 with magnetic properties due to its colloidal crystal structure with Fe3O4. The Fe3O4 nanoparticles were synthesized using co-precipitation techniques to a size of 28.2 nm and a magnetic saturation of 66.9 emu g−1. Meanwhile, mesoporous titania nanoparticles (MTNs) with a particle diameter of 373 nm, a specific surface area of 236.3 m2 g−1, and a pore size of 2.8 nm were prepared by controlling the rate of hydrolysis. Magnetic colloidal crystals (a diameter of 10.2 μm) were formed by the aggregation of Fe3O4 and MTNs caused by the interface phenomena during solvent evaporation in emulsion. Even the anatase octahedrite produced from the colloidal crystal after a hydrothermal reaction retained a magnetic saturation of 2.8 emu g−1. This study also investigates the photodegradation activity of our synthesized material as a photocatalyst, while utilizing its capability for magnetic separation to prove its usefulness in catalyst recycling.

1. Introduction

The ability of a catalyst to reduce the activation energy of a chemical reaction and reaction time proves its vital application in various industrial processes.1 Moreover, catalysts are not depleted during a reaction, so methods to recycle and reuse catalysts are also popular research topics.2 General solid catalyst recycling methods include three main techniques: centrifugation, filtration, and magnetic separation. Centrifugation relies on the centrifugal force to separate the solid and liquid from a mixture, but is only appropriate for laboratory development rather than for industrial applications due to its time consuming procedures.3 Filtration relies on the varying particle sizes of a mixture so that only specific sized particles pass through the filter to establish separation between the catalyst and the solution.4 Yet filtration generally requires a vacuum source or high pressure to facilitate the process, which results in the additional cost and energy usage associated with filtration. Additionally, a long-term usage of the same filter causes particle deposition onto the filter pores, which reduces the efficiency of filtration. It is necessary to frequently replace the filter to prevent particle accumulation and maintain minimal filtration efficiency.5 Due to these difficulties presented by centrifugation and filtration methods, magnetic separation is the ideal technique for solid catalyst recycling. It is capable of cheaply and efficiently separating magnetic particles from a solution by applying an external magnetic field.6,7 The main prerequisite for using magnetic separation, however, is the catalyst’s necessary magnetism.

Methods for the incorporation of magnetic particles within mesoporous catalysts are well studied,8 and mainly two general methods are utilized: the sol–gel method that creates a core–shell structure of magnetic cores and another material as a shell,9 and the direct growth of magnetic particles in the pores of another material.10 However, both methods have difficulties to overcome. For example, if the growth rate of titanium oxide for the core–shell structure was too slow, then the titanium oxide shell would be too thin to provide the necessary photoactivity to the core. Yet this would result in low particle dispersion and aggregation. Similarly, the method of direct magnetic particle growth in titanium oxide mesopores has a similar problem.13,14 In addition to the preliminary preparation of mesopores, the method...
requires precise magnetic particle deposition control. Too few magnetic particles would result in the inadequate magnetic attraction during separation, while too many would result in magnetic particle deposition over the catalyst surface to cause low catalytic rates.

This study proposes a directed self-assembly process designed to incorporate magnetic particles within the titanium oxide sphere. We here choose to use mesoporous titania nanoparticles (MTNs) with a high surface area as photocatalysts and iron oxide (Fe₃O₄) as the magnetic particle. As shown in Fig. 1, MTNs and Fe₃O₄ are suspended in the emulsion system in the form of droplets. The MTNs self-assemble into a colloidal crystal structured microparticle due to the interface phenomena during solvent evaporation. Fe₃O₄ is introduced into the colloidal crystal to result in the magnetized microparticle during the directed self-assembly of MTNs into the colloidal crystal structure. The resulting material consists of magnetic Fe₃O₄ and MTNs, and is abbreviated as M-MTNs hereafter. In addition to the structural and physical analysis of the M-MTNs, this study will investigate the activity of the M-MTNs as photocatalysts by measuring their organic degradation, while utilizing their capability for magnetic separation to prove their usefulness in catalyst recycling.

2. Results and discussion

Fig. 2 shows the XRD patterns of the different samples. The Fe₃O₄ patterns indicate the spinel crystalline, which is typical for iron oxide prepared by co-precipitation. After hydrothermal treatment, Fe₃O₄ is converted into α-Fe₂O₃, which shows less apparent peaks than that of MTNs due to its lower starting amount during synthesis. The UV-Vis absorption of the samples was analyzed (data not shown). The adsorption wavelength for MTNs could be illustrated according to the following equation:

\[ E = \frac{h\lambda}{c} \]  

where \( E \) is the band gap energy [eV], \( h \) is Planck’s constant \((6.6256 \times 10^{-34} \text{ J s})\), \( c \) is the speed of light \((2.9989 \times 10^8 \text{ m s}^{-1})\), and \( \lambda \) is the light’s wavelength [nm]. The MTNs after hydrothermal treatment without the addition of Fe₃O₄ (MTN_HT) exhibit characteristic peaks corresponding to anatase, which is an appropriate candidate as a photocatalyst. Since the TiO₂ anatase has a band gap energy of 3.2 eV, it is possible to use eqn (1) to deduce that MTN_HT best absorbs and begins its photodegradation at a wavelength of 387.5 nm. For the duration of this study, MTN_HT will serve as our standard photocatalyst to be compared with M-MTNs.

Fig. 3 shows the SEM image of all samples. Co-precipitated Fe₃O₄ mainly has a size distribution of 28.2 nm, as shown in Fig. 3(a). The MTNs used here were prepared by a sol–gel method, using titanium tetraisopropoxide (TTIP) as a precursor for the hydrolysis-condensation reactions to yield TiO₂. The overall reaction could be illustrated as:

\[ \text{Ti(OCH(CH₃)₂)₄} + 2\text{H₂O} \rightarrow \text{TiO}_2 + 4\text{(CH₃)₂CHOH} \]

During the early stages of the reaction, the newly synthesized titanium oxide crystals would slowly and steadily form spheres of MTNs with a size of 373 nm, as evaluated by SEM images in Fig. 3(b). The MTNs used here were prepared by a sol–gel method, using titanium tetraisopropoxide (TTIP) as a precursor for the hydrolysis-condensation reactions to yield TiO₂. The overall reaction could be illustrated as:

\[ \text{Ti(OCH(CH₃)₂)₄} + 2\text{H₂O} \rightarrow \text{TiO}_2 + 4\text{(CH₃)₂CHOH} \]
size and the added amount of Fe₃O₄ are small compared to MTNs, the introduction of Fe₃O₄ does not hinder the formation of MTNs. The mesopores of MTNs are approximately 10–20 nm in size, which match the pore size distribution. As seen from Fig. 3, the cracks of the colloidal crystal are filled by Fe₃O₄ to give the microsphere magnetic properties (Fig. S1, ESI†).

Fig. 4 shows that both the magnetic properties of Fe₃O₄ and M-MTNs have no corresponding hysteresis loops, which indicate that both are superparamagnetic. Fe₃O₄ has a magnetic saturation of 66.8 emu g⁻¹, while M-MTNs have a magnetic saturation of 2.8 emu g⁻¹. The magnetic saturation of M-MTNs is an order less than that of Fe₃O₄, because the mass ratio of Fe₃O₄ in MTNs is less than a tenth. Additionally, the hydrothermal reaction causes Fe₃O₄ to be converted into α-Fe₂O₃, which further results in MTN’s lower magnetic saturation.

Fig. 5 illustrates the nitrogen gas absorption and the pore size distribution of different samples, which are summarized in Table 1 along with the other characteristics. According to the IUPAC classification, MTNs prior to hydrothermal treatment are classified as type I as deduced by their hysteresis loop. This indicates that gases are absorbed from the particle’s micropores when under low air pressure. The pore size of MTNs was calculated to be 2.8 nm using Barrett–Joyner–Halenda (BJH) calculations. On the other hand, MTN_HT that has been obtained through hydrothermal treatment showed a type IV classification. This indicates that the majority of MTN_HT’s pores are mesopores, which is supported by the BJH pore size calculations of 10.9 nm. The difference in pore sizes may be due to the difference between the mass transfer coefficient of TiO₂ and the vacancy during hydrothermal treatment. This difference results in the Kirkendall effect, which causes the aggregation of cracks to form mesopores and the conversion of TiO₂ from amorphous to anatase.¹⁶ As the pore size increased, the material’s specific surface area increased inversely from 236.3 m² g⁻¹ prior hydrothermal treatment to 142.7 m² g⁻¹ after hydrothermal treatment (Table 1).

The hysteresis loop of M-MTNs is similar to that of MTN_HT, where both are classified as type IV.¹⁷ As seen from the SEM images in Fig. 3(d), the cracks of the colloidal crystal particles are around 10–20 nm. This indicates that the peak for pore size distribution of M-MTNs is more towards the right as compared to that of MTNs. M-MTNs exhibit an average pore diameter of 12.3 nm and a specific surface area of 129.4 m² g⁻¹. If the Fe₃O₄ contents were to be removed, then the specific surface area of M-MTNs would be 142.34 m² g⁻¹, which is similar to MTN_HT. Since the particle size of Fe₃O₄ (28.2 nm) is an order larger than the diameter of MTN’s mesopores (2.8 nm), particles of Fe₃O₄ would not occupy MTN’s mesopores during self-assembly, which thus means that the photocatalyst retains its high specific surface area.

The general information on the photocatalyst ability to photodegrade Orange II is summarized in Fig. 6, where C is the concentration of Orange II, C∗ is the starting concentration of Orange II, C₀ is the starting concentration of Orange II prior to UV radiation, and t is the duration of UV radiation. Orange II degradation was not observed for the blank sample (without any addition of photocatalyst) under UV radiation. It is therefore deduced that the photocatalyst is the main factor for photodegradation upon its addition to the sample. As illustrated in Fig. 6(a), MTN_HT absorbed 44% of Orange II while M-MTNs absorbed 29%, which is 59% of what MTN_HT absorbed. It has been well known that particles are usually absorbed and aggregated around the pore opening during the absorption process for porous materials. For M-MTNs, the colloidal crystal structure of MTNs causes more closely packed particles, which result in fewer pores in contact with the solvent and thus lower absorption. Additionally, the possibility that

<table>
<thead>
<tr>
<th>Sample</th>
<th>Surface area (m² g⁻¹)</th>
<th>Pore size (nm)</th>
<th>Crystallinity</th>
<th>Adsorption ability (µg mg⁻¹)</th>
<th>Reaction constant (k, min⁻¹)</th>
</tr>
</thead>
<tbody>
<tr>
<td>MTN</td>
<td>236.3</td>
<td>2.8</td>
<td>Amorphous</td>
<td>N/A</td>
<td>N/A</td>
</tr>
<tr>
<td>MTN_HT</td>
<td>142.7</td>
<td>10.9</td>
<td>Anatase</td>
<td>33.9</td>
<td>0.0448</td>
</tr>
<tr>
<td>M-MTN</td>
<td>129.4</td>
<td>12.3</td>
<td>Anatase + α-Fe₂O₃</td>
<td>18.5</td>
<td>0.0183</td>
</tr>
</tbody>
</table>

* indicates the starting concentration of Orange II prior to UV radiation.
Fe₃O₄ particles covered the openings of the pores could also be another reason for the lowered absorption.¹⁸

Fig. 6 illustrates the degradation kinetics after UV radiation. MTN_HT and M-MTNs both possess the ability for photocatalytic degradation, which could be described by the following first-order equation:

\[ r = \frac{dC}{dt} = kC \]

where \( r \) is the rate of Orange II degradation, \( C \) is the concentration of Orange II, \( k \) is a rate constant, and \( t \) is the reaction time. Eqn (2) could be integrate to be:

\[ \ln \frac{C_0}{C} = kt \]

where \( C_0 \) is the concentration of Orange II prior to UV radiation. Fig. 6(c) could be derived from Fig. 6(b) using the above equations to calculate the rate constant (\( k \)). The \( k \) value for M-MTNs was determined to be 0.0183 min⁻¹, which is 40% of that for MTN_HT. Once again, this difference could be the result of \( \alpha \)-Fe₂O₃ covering the surface of MTNs, and thus resulting in blocked UV absorption and decreased photocatalytic degradation of M-MTNs.

Upon UV radiation for 90 minutes, M-MTNs were capable of removing large quantities of Orange II from the solvent (\( C/C^* = 90\% \)) through the processes of absorption and degradation. As shown in Fig. 7, M-MTNs could be efficiently separated from the solution with an external magnetic field given their magnetic properties from \( \alpha \)-Fe₂O₃. The effectiveness of recycled M-MTNs was tested up to their fifth reuse, as shown in Fig. 6(d). Each sample received UV radiation for 90 minutes, and the recycled M-MTNs were still able to maintain an efficiency of approximately 90% with a standard deviation of 2.8%.

### 4. Experimental section

#### Chemicals

Ammonium hydroxide (\( \text{NH}_4\text{OH}, 28\% \)), hydrochloride acid (HCl, 37%), iron(III) chloride hexahydrate (\( \text{FeCl}_3\cdot6\text{H}_2\text{O} \)), Orange II, potassium chloride (KCl), and titanium(IV) isopropoxide (abbreviated as TTIP, 97%) were purchased from Sigma-Aldrich. Ferrous chloride, 4-hydrate (\( \text{FeCl}_2\cdot4\text{H}_2\text{O} \)) was purchased from J.T.Baker.

#### Synthesis of mesoporous titania nanoparticles (MTNs)

Mesoporous materials have been widely synthesized through various methods such as surfactant-templating processes and layer-by-layer processes.¹⁹–²¹ MTNs were obtained by the controlled hydrolysis of TTIP in an EtOH environment.²² In general, 2.2 mL of TTIP was added dropwise into 100 mL of pure EtOH containing 0.4 mL of KCl solution (0.1 M). After stirring at ambient temperature for 60 min, the white mixture was left stationary for 24 hours for aging. The precipitated product was collected by centrifugation, washed with EtOH, and dried under vacuum.

![Degradation of Orange II and separation of M-MTNs before and after applying by magnetic field.](image)
Synthesis of Fe₃O₄ nanoparticles

Fe₃O₄ nanoparticles were prepared by co-precipitation method.²³ 2 M FeCl₃·6H₂O was prepared in 2 mL of HCl (2M), and then added to 8 mL of Fe³⁺ aqueous solution (1 M). This yellow mixture was mixed with 100 mL of NH₄OH (0.7 M) to precipitate black Fe₂O₃. After stirring at ambient temperature for 30 minutes, Fe₂O₃ was collected by centrifuge and washed with DI water.

Synthesis of Fe₃O₄/MTN colloidal crystals

0.4 mL of the aqueous solution containing 3 mg of Fe₃O₄ and 30 mg of the MTN was added into 20 mL of silica oil, stirred at 2000 rpm for 10 minutes, and then left at 80 °C to evaporate the solvent. After 8 hours, the particles were washed with hexane, collected by centrifuge, and dried at ambient temperature. M-MTNs were obtained after a hydrothermal reaction at 180 °C for 12 hours.

Characterization

X-ray diffraction (XRD) patterns of samples were performed on a Rigaku ultima IV using Cu Kα radiation (λ = 0.154 nm). The surface morphology and internal structure of products were observed using SEM (NovaTM Nano SEM) and TEM (JEOL JEM-1200EX II). The porous properties of samples were analyzed using N₂ adsorption/desorption isotherms on a Micromeritics ASAP 2010 instrument, and all samples were degassed at 100 °C overnight before N₂ adsorption. Surface area and pore size distribution were calculated using Brunauer–Emmett–Teller (BET) and Barrett–Joyner–Halenda methods, respectively. The magnetic properties of samples were studied using a SQUID (MPMS7 from Quantum Design Company of USA) magnetometer at a temperature of 300 K.

Photocatalysis

20 mg of the catalyst and 20 mL of Orange II solution (2 × 10⁻⁵ M) were stirred in the dark overnight to reach adsorption equilibrium. This aqueous slurry was then irradiated using a 4 W UV lamp (wavelength at 254 nm) and kept stirring in a dark environment. The concentration of Orange II was determined using a spectrophotometer (Jasco spectrophotometer V-670) at a wavelength of 484 nm.

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References