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Ion-specific adsorption and electroosmosis in charged amorphous porous silica†

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Monovalent and divalent aqueous electrolytes confined in negatively charged porous silica are studied by means of molecular simulations including free energy calculations. Owing to the strong cation adsorption at the surface, surface charge overcompensation (overscreening) occurs which leads to an effective positive surface next to the Stern layer, followed by a negatively charged diffuse layer. A simple Poisson–Boltzmann model in which the single-ion potential of mean force is introduced is shown to capture the most prominent features of ion density profiles near an amorphous silica surface. Nevertheless, due to its mean-field nature, which fails to account for correlations, this simple model does not predict overscreening corresponding to charge inversion at the surface. Such an overscreening drastically affects the transport of confined electrolytes as it leads to flow reversal when subjected to an electric field. A simple continuum theory is shown to capture how the electro-osmotic flow is affected by overscreening and by the apparent enhanced viscosity of the confined electrolytes. Comparison with available experimental data is discussed, as well as the implications of these phenomena for ζ -potential measurements.

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1 Introduction

While electrokinetic methods such as electroosmosis and electrophoresis^{1–3} allow controlling the flow of confined electrolyte solutions, the large surface to volume ratio in nanofluidic devices leads to a variety of new adsorption and transport phenomena.^{4,5} Understanding the complex structure and flow properties of electrolytes near charged interfaces is therefore of paramount importance to develop and improve nanotechnologies in the field of drug delivery, filtration, energy conversion, *etc.* Among the possible nanofluidic devices, silica-based materials are important as they can be easily synthesized, tuned and shaped, as well as used to prepare lab-on-a-chip devices.

As far as molecular simulation techniques are concerned, in parallel to the abundant literature on pure water confined in crystalline^{6–9} and amorphous^{10–14} porous silica, several authors have investigated the adsorption of alkali halide and alkaline

earth halide ions at idealized interfaces^{15,16} or crystalline materials.^{17,18} More recently, the interfacial properties of aqueous electrolyte solutions with large ion concentrations adsorbed at the surface of amorphous silica have received increasing attention.^{19–23} A few studies related to electrolyte solutions in amorphous silica have focused on the migration of ions between reservoirs connected by a short neutral^{20,22,23} or charged^{24,25} silica pore. The effect of the pore size was investigated in these studies, in addition to structural, dynamical, or dielectric properties. Videla *et al.*²⁰ compared the densities of NaCl aqueous electrolytes confined in cylindrical pores with radii of 10.0, 15.0, and 17.5 Å in neutral amorphous silica. Axial density profiles of water and ions were found to be correlated with the overall distribution of silanol groups, which indicates that certain silanol groups form more favorable adsorption sites than others. Renou *et al.*²³ found, for amorphous silica with a cylindrical pore of radius 6 Å, that ion polarizability has little effect on the axial ion density profiles of NaCl, while a larger effect was found for NaI in the form of rejection of polarized anions at the pore entrance. Such ion-specific effects can be key to selective removal and separation of ions. Cruz-Chu *et al.*²² observed for 1 mol L^{−1} KCl solutions in cylindrical pores of radii of 10 to 30 Å a preferred cation adsorption to dangling oxygen atoms at the silica surface, especially when multiple dangling oxygen atoms appeared next to each other. These sites are easily accessible and provide sufficiently strong interaction energy to disturb the hydration shell of the ion. Consequently, the distribution and flow of electrolyte solutions can be controlled *via*

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the distribution of surface groups, as we will further investigate in this study. The flow of ions through pores in contact with bulk-like reservoirs form an important class of fluid–solid interface problems. However, the pore lengths that can be studied using molecular dynamics (MD) simulations are too short to obtain a region at which the influence of the entrance can be neglected and lateral homogeneity can be assumed. Instead, simulations of fully periodic pores allow for the study of confined fluids without any pore entrance effects. Lorenz *et al.*¹⁹ have performed MD simulations of streaming current and electro-osmotic flow of NaCl and CaCl₂ solutions in a charged amorphous silica slit pore of 75 Å width. They showed that the electrokinetic flow depends on the ionic strength and that the obtained ion distribution depends on the silica force field and the preparation method. A multiscale approach was proposed, which couples charge distributions from MD simulations with a continuum model. While this approach proved promising, the comparison between simulations and theory was hindered by limited understanding of electroviscous effects. Argyris *et al.*¹⁸ compared the ion distributions of 1 mol L^{−1} NaCl and CsCl solutions confined in a neutral crystalline silica slit pore of a width of 10.68 Å. Ion-specificity was found to be manifested as a dense layer of sodium ions formed close to the surface while cesium ions do not show such strong layering and are located further from the surface. Chen *et al.*²⁶ recently studied carbohydrate–calcite interactions for calcite slit pores of widths between 36.5 and 54.8 Å. The carbohydrate was immersed in an aqueous solution with brine concentrations of Na⁺, Ca²⁺, Mg²⁺, and Cl[−]. The authors found that different terminations of the calcite crystal structure resulted in different ion adsorption behavior, including the occurrence of overscreening. Interestingly, it was found that an increase of the brine concentration resulted in a reduction of the carbohydrate–calcite hydrogen-bond interaction due to more screening from the salt adsorbed onto the calcite surface.

Despite many insights gained from earlier studies, many questions surrounding ion-specificity and liquid–solid interfaces remain unanswered. This includes gaining a full understanding of the underlying mechanisms and consequences of overscreening and local viscosity enhancement. Another open problem is the development of a theoretical description of the unique properties of nano-confined electrolytic fluids. This paper aims at determining the behavior of various aqueous electrolyte solutions confined in charged amorphous silica. In addition to conventional electrolytes (NaCl and KCl), we consider radionuclides, such as CsCl and SrCl₂, which are relevant to decontamination of nuclear effluents (such as in the framework of nuclear accidents). The selective removal of ¹³⁷Cs and ⁹⁰Sr ions from contaminated effluents is indeed crucial as they tend to be confused by bio-organisms with K⁺ and Ca²⁺, respectively, which possess similar solvation properties. Here, we consider the possibility to selectively filter these harmful radionuclides using ionic exchange in nanoporous silica^{27–29} as a promising alternative to classical liquid–liquid separation techniques. We first determine, by means of molecular simulations, ion adsorption at the charged surface of porous silica. We show that ion adsorption is relatively well accounted for by a modified Poisson–Boltzmann equation in which the potential

of mean force is introduced for a single solvated ion near a silica surface. In contrast, owing to its mean-field nature, this simple model does not capture overscreening (*i.e.* overcompensation of the surface charge by the counterions) which is observed in our simulations as a result of ion–ion correlations and static localized surface charges considered in our model of porous silica. By simulating the electro-osmotic flow induced by an electric field, we study the effect of such overscreening on the transport properties of the confined electrolytes. The remaining part of this paper is organized into five sections. In Section 2 we discuss the preparation and details of our simulations. Equilibrium molecular dynamics results are shown in Section 3. A modified Poisson–Boltzmann model is presented and discussed in Section 4. We perform electro-osmotic flow simulations and apply a continuum model in Section 5. Finally, Section 6 summarizes our findings and conclusions.

2 Computational methods

2.1 Models

Fig. 1 shows a typical molecular configuration of porous silica considered in this work. *y* is the axis perpendicular to the surface, defined as *y* = 0 located at the pore center. The procedure for the preparation of silica walls is described in detail elsewhere.³⁰ In brief, a block of cristobalite is first melted at *T* = 4000 K. Using a simulated annealing technique, the material is then quenched until 300 K to reach an amorphous state at which the structure factor is validated against experimental measurements. A surface is created by cutting the silica block and removing the un-coordinated Si atoms, as well as the O atoms that are no longer connected to any Si atom. The dangling bonds of the O atoms at the surface are then saturated by adding an H atom to form a silanol group. As explained by Villemot *et al.*³¹ the number of OH bonds is adjusted to reach a final silanol density of 5OH nm^{−2} (based on the frontal

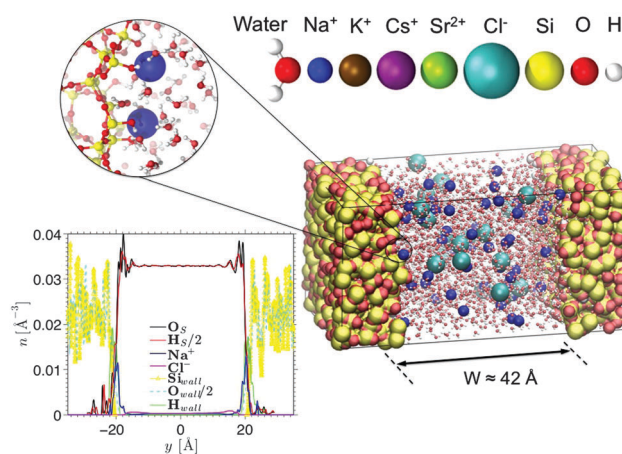


Fig. 1 An aqueous electrolyte solution confined in an amorphous silica slit pore with a surface charge density of $\sigma_s = -0.1 \text{ C m}^{-2}$. The surface charge is created by deprotonating some of the silanol groups, shown in the magnification on the left. The legend on top shows the atoms and ions considered. Four aqueous solutions are considered, where the electrolytes are NaCl, KCl, CsCl, and SrCl₂ at a concentration of 0.6 M.



surface area), which agrees well with experimental measurements.³² Following previous studies on porous silica,^{33,34} the surface roughness of the slit pore is determined by simulating small angle neutron scattering (Fig. S1 in the ESI[†]); a Porod exponent of ~ 4 is found, indicating a smooth surface. This is confirmed by a small variation in the height of the oxygen atoms at the surface which has a standard deviation of 1.2 Å (although, as expected for glass, it is a bit larger than for cristobalite ~ 0.7 Å). The size of our surface is 35.8 Å \times 35.8 Å, which corresponds to 64 surface groups per wall. Some of these (1 out of 8) are deprotonated to create a negative surface charge density of $\sigma_s = -0.1 \text{ C m}^{-2}$. This surface charge in combination with an ionic strength of 0.6 M corresponds to a pH value of ~ 7.0 – 8.0 , depending on the electrolyte species. The walls are approximately 15 Å thick, which are chosen such that there is no notable effect of finite thickness on the interfacial fluid properties. Two silica walls are placed parallel to each other at a distance of approximately 42 Å in order to create a slit pore (the directions parallel to the walls are infinite thanks to the use of periodic boundary conditions). The distance between the walls is chosen to be sufficiently large so that the electric double layers from both walls are separated and the electro-osmotic flow velocity becomes independent of the channel width.

Based on the approximate volume of the silica pore, 20 Cl^- ions and 36 monovalent cations (or 18 divalent cations) are inserted into the channel to compensate the negative surface charge while reaching an average concentration of 0.6 M (mol L^{-1}) (close to that of seawater). Water molecules, described using the TIP4P/2005 rigid model,³⁵ are then added using Grand Canonical Monte Carlo (GCMC) simulations. This water model is selected as it has been shown to be successful in reproducing the transport properties of water.^{9,36} The number of water molecules in our systems ranges from 1736 for the CsCl system to 1822 for the SrCl_2 system. GCMC simulations (constant chemical potential and temperature simulations in which the number of water molecules varies) are needed to prepare the confined electrolytes, as the number of water molecules does not scale exactly with the accessible volume. An interesting result of our GCMC simulations (not shown here) is that by making the surface more negatively charged and adding cations to compensate (thus reducing the volume accessible to water) the number of water molecules actually increases. This is explained by the ordering of water molecules around ions.

2.2 Molecular dynamics

MD simulations are performed using LAMMPS³⁷ with a simulation time step of 1 fs. The fluid temperature is kept at 300 K *via* a Nosé–Hoover thermostat with a coupling time constant of 100 fs. Dispersion interactions are described using a Lennard-Jones potential, with a cutoff length of 10 Å and a tail correction to compensate for the truncated part. The particle–particle–particle–mesh (P^3M)³⁸ method is used to calculate electrostatic interactions, where we truncate the real part at 10 Å. The SHAKE algorithm is used to preserve the rigid structure of water molecules. The force field parameters optimized for the TIP4P/Ew water model by Joung and Cheatham are used for monovalent ions.³⁹ The parameters for

Na^+ and Cl^- have been confirmed to be transferable to TIP4P/2005 water,⁴⁰ while transferability of the parameters for K^+ and Cs^+ is shown in Fig. S2 in the ESI[†]. The parameters recently presented by Mamatkulov *et al.* are used for strontium.⁴¹ These parameters have been optimized for SPC/E and shown to be transferable to TIP4P/2005.⁴² The Lorentz–Berthelot combination rules are used to calculate cross-species interactions, where the parameters for silica are taken from Lee and Rossky.⁴³ The force field parameters are listed in Table SI in the ESI[†]. The polarizability is not included as electrostatic forces near hydrophilic charged surfaces are much larger than the polarization forces.^{23,44} It should be noted that possible charge transfer between the different cations and the oxygen atoms of the silica surface occurs to some limited yet non-negligible extent. However, considering that the force fields used in the present work rely on effective parameters, which are adjusted against available experimental data, they include in an implicit fashion this contribution. Similarly, these force fields include other contributions such as 3-body and 4-body interactions as well as polarizability, which correspond to the electron cloud response to the local electrical field created by other surrounding atoms. Overall, these other interactions, which contribute to a lesser extent to the total internal energy than the Coulombic and dispersive interactions, are usually omitted in molecular simulations of aqueous electrolyte solutions.

2.3 Potential of mean force

A large adsorption energy of ions can make it difficult to sample phase space *via* equilibrium MD simulations. This applies also to a single ion near an uncharged wall (which will be considered in Section 4). We therefore use the Umbrella Sampling technique to efficiently calculate the free energy profile as a function of the position across the channel.⁴⁵ This method deters the ion from being far from the position y_0 (called a ‘sampling window’) by imposing an energy penalty as a function of the distance from this position. A short simulation is run to collect data on where the ion resides when subject to the imposed energy penalty. The position y_0 is then shifted and the process is repeated several times. This eventually results in many histograms, each resulting from a simulation with a known energy bias. Using the weighted histogram analysis method (WHAM) the unbiased density of states (and thus the corresponding free energy profile) can be iteratively reconstructed from the histograms.⁴⁶ We apply Umbrella Sampling using a harmonic energy bias with a spring constant of 0.12 eV Å^{-2} (1.95 N m^{-1}). We use 60 sampling windows with a spacing of 0.75 Å. The sampling windows are located in the domain $y \in [-22.125, 22.125]$, with $y = 0$ being at the center of the channel. The spring stiffness in combination with the width of the sampling windows results in partially overlapping histograms, which is needed for the reconstruction of the unbiased profile. The system is equilibrated for 200 ps after the bias potential is imposed. We subsequently collect data over the next 500 ps. We confirm that the equilibration time and the sampling time are sufficiently long. We also validate that the spring stiffness is large enough so that large energy barriers can be sampled accurately. This is done by performing Umbrella Sampling with an increased stiffness of 0.2 eV Å^{-2} and



comparing the reconstructed energy profile. Furthermore, since the walls and the solvent are charge neutral, the single-ion system has a net charge equal to the charge of the ion. A correction is applied to the treatment of the electrostatic interactions to prevent artifacts in the energy and pressure due to non-neutrality of the simulation system.⁴⁷

The harmonic energy bias is applied in the direction normal to the walls, while ion movement in the other directions is suppressed. This is needed since the ion would otherwise drift to local energy minima at the surface. In order to capture the effect of the surface structure we probe the density of states (DoS) profiles $\rho(y)$ at 9 randomly chosen x - z positions in front of the silica surface. Mapping each profile to the left half of the channel ($y < 0$) gives us $p = 18$ unique profiles since the structure of both walls is different. The potential of mean force (PMF) is then calculated from the arithmetic average of various DoS profiles as follows:

$$U(y) = \beta^{-1} \ln \left(\frac{1}{p} \sum_{i=1}^p \rho_i(y) \right), \quad (1)$$

where $\beta = (k_B T)^{-1}$ is the inverse of the thermal energy. The fact that the PMF is not simply calculated from the arithmetic average of the 18 energy profiles can be understood by the following thought: imagine a hard wall with an opening that covers half the surface area. An ion trying to penetrate the hard wall will result in a very large repulsion, while no repulsion might be found at the same y -position in front of the open part. The arithmetic average of the energy profiles would then result in a PMF with a huge energy value at the y -position of the surface, implying that the penetration of the open area is also not possible (since the one-dimensional PMF does not distinguish between different x - z positions). On the other hand, calculating the PMF from the average DoS (as shown in eqn (1)) correctly accounts for the part of the surface area that is penetrable in our example. A graphical representation of the calculation of the PMF is shown in Fig. S9 in the ESI† in which an average potential of mean force is calculated from 4 profiles by averaging the DoS. We have validated that $p = 18$ is large enough for our silica surface by comparing the PMF to those calculated from 14 and 16 profiles. The addition of more profiles has little effect on the resulting PMF, implying that $p = 18$ is sufficient to represent the amorphous structure of the silica surface. The surface structure and the x - z positions, at which the density of states profiles were computed, are the same for each of the ions to allow for a direct comparison.

3 Molecular dynamics results

Equilibrium MD simulations are performed for 80 ns, the last 30 ns of which are used to accumulate statistics by storing data every 25 ps. As discussed in detail below, for each simulation, two different initial configurations are considered in which the ions are distributed near the center of the channel. Fig. 2 shows the number density profile of the different electrolyte solutions, where $y = 0$ corresponds to the center of the channel. Fig. 2a shows the ion number density profiles of the cations (full lines)

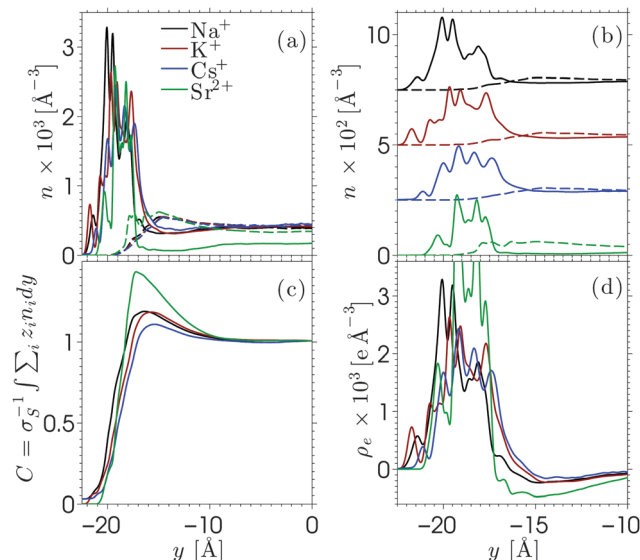


Fig. 2 Equilibrium ion distribution profiles of the various electrolyte solutions. The density profiles are shown in (a), while the same profiles are shifted vertically in (b) for the sake of clarity. The full lines denote the cation number density for Na^+ (black), K^+ (red), Cs^+ (blue), and Sr^{2+} (green). The dashed lines denote the chloride density profiles corresponding to the cation profile in the same colors. The relative screening of the surface charge is shown in (c), while the charge density profiles of the electrolyte solutions are shown in (d). Note on the horizontal axis of (b) and (d) that these figures show the part of the data that correspond to the interfacial region, while the horizontal axes in (a) and (c) shows also the center of the channel.

and the chloride ions (dashed lines). These profiles are shifted vertically in Fig. 2b for the sake of clarity. Fig. 2c shows the charge screening factor, calculated as:

$$C(y) = \sigma_S^{-1} \int_0^y \sum_{i=+,-} z_i n_i(y') dy', \quad (2)$$

where z_i is the valence of the cation or anion species and $n_i(y)$ is its number density profile. The screening factor indicates to what extent the surface charge is screened by the electrolyte adjacent to the wall. A screening factor of 1 means that the surface charge is exactly screened, which is typically the case sufficiently far from the interface. Fig. 2d shows charge density profiles $\rho_e = z_+ n_+ + z_- n_-$.

The ion density profiles in Fig. 2 show a dense layer of cations adsorbed at the negatively charged silica walls, while the anions are repelled by the walls. The adsorbed cations form the Stern layer in the classical model of the electric double layer (EDL).⁴⁸ The thickness of the adsorbed layer is larger than the Debye length of the solution ($\lambda_D \approx 3 \text{ Å}$) and larger than the Lennard-Jones diameters of the cations (see Table SI in the ESI†). This is partially due to disorder at the surface, which induces broadening of the Stern layer. While the cation layer is located at a similar position for each electrolyte solution, it is slightly shifted towards the pore center as the cation size increases. The area under the strontium number density profile is smaller than that of the monovalent cations, since only half



the number of divalent ions are needed to compensate the negative charges of silica and chloride ions.

For each electrolyte, next to the adsorbed layer, the confined electrolyte is partially depleted of cations while anions are in excess. Such overscreening occurs when the cations overcompensate the negative surface charge (*i.e.* $C(y) > 1$ in Fig. 2c), thus causing an effective positive surface near the silica surface. The transition between a positively and a negatively charged layer, which is often referred to as charge inversion (CI), is typically observed for multivalent electrolytes or concentrated monovalent electrolytes. Overscreening, which is mostly pronounced for SrCl_2 followed by NaCl , KCl , and CsCl , increases with the charge density of the cations and the size asymmetry of the electrolytes.⁴² While the microscopic origins of overscreening remain poorly understood,⁴⁹ Faraudo and Travestet⁵⁰ explain that overscreening results from the correlation free energy between ions in the Stern layer, as well as from ion-wall correlations. The correlation free energy required to cause overscreening decreases strongly with increasing ionic strength. In other words, overscreening occurs typically above a critical ion concentration which is strongly dependent on the exact structure and hydrophilicity of the surface.²⁰ Overscreening has even been shown to occur at millimolar concentrations of organic monovalent ions near a hydrophobic colloidal particle.⁵¹ The solvation free energy was said to be the dominant mechanism in this case, while this was also found to be the mechanism that prevented overscreening near a hydrophilic colloid in the same electrolyte solution. The experimental results suggest that overscreening for monovalent electrolytes (especially KCl and CsCl) near a silica surface occurs at larger concentrations than those considered here, while various computational studies find overscreening under conditions similar to ours.^{15,16,21,52–56} Several factors and uncertainties can be responsible for quantitative differences between experimental and theoretical studies. In the first place, from an experimental point of view, the critical concentrations at which overscreening occurs for monovalent ions are often very high and measuring electric properties at such large ionic strengths can lead to a significant dependence on the measurement technique.^{57,58} van der Heyden *et al.*⁵² introduced streaming current measurements as a new experimental approach to study overscreening. They measured the effective surface charge and the streaming current near a silica surface for monovalent, divalent, and trivalent electrolytes at concentrations up to 1 M. Their data confirmed that overscreening occurs above a critical concentration. The onset for CaCl_2 and MgCl_2 solutions in their system was located around 0.4 M. No charge inversion was observed for KCl (the only monovalent electrolyte considered in the study), but the data suggest that the onset of overscreening for KCl is not far above 1 M. These experiments give a good insight into the qualitative dependence on ion concentration and an indication of the concentrations at which overscreening could occur under the conditions of these experiments. However, a quantitative comparison with other studies remains challenging since values can depend for example on how the samples are prepared.

As for molecular simulations, the local properties of the electrolyte solution depend strongly on the ion concentration as

well as the influence of confinement and the simulation model. An example of the effect of different force fields is given by Lorenz *et al.*,¹⁹ who compared the density distribution of CaCl_2 near charged amorphous silica for different force fields and material preparation methods. Both simulations showed overscreening and almost the same charge inversion point but the distributions of cations in the Stern layer were significantly different. Considering that the force fields used in our work were developed to reproduce both the solvation free energy and activity of the ions,⁴¹ it can be assumed that they describe correctly the experimental solubilities in water. Ion precipitation at the surface, which leads to overscreening, can be seen as a surface solubility phenomenon in the spirit of recent studies on enhanced gas solubility in solvents confined in nanoporous media.^{59,60} To further confirm that overscreening observed in the present work is physical, we checked that the problem of accessible timescales in MD simulations does not affect our results. As briefly mentioned at the beginning of this section, each profile in Fig. 2 has been obtained from the average between two simulations with different initial configurations. The comparisons between the independent simulations are shown in Fig. S3–S5 in the ESI.† The two simulations show very good agreement for the monovalent electrolytes. We also run additional simulations in which Na^+ and Cl^- ions are initially distributed in a way that corresponds to no overscreening (these configurations are obtained from simulations in which the surface was homogeneously charged). We then relax the systems and determine the ion density profiles shown in Fig. S6 in the ESI.† From these additional simulations we find that each of them converges rapidly to a configuration that shows overscreening, which is consistent with the data shown in Fig. 2. The deprotonated silanol groups form highly preferable adsorption sites for the cations. The number of charge sites equals the number of excess cations in the system and we observe that almost each charge site forms a direct contact pair with at least one cation. Some additional cations might also adsorb (and thus overscreen the surface charge) at open spaces where the repulsion from other cations is small. As for SrCl_2 (Fig. S5 in the ESI.†), the two independent simulations converge to different yet similar and equivalent ion density profiles. The in-plane structure of the adsorbed cation layer shows that cations adsorb in different sites, due to the fact that there are fewer divalent cations than accessible sites. Once a divalent ion adsorbs to a deprotonated silanol group, it stays there for a long time due to the large binding energy. Averaging between multiple simulations (as was done in the present work) is a way to improve sampling for a system with ergodic hindrances.

4 Electrostatic model

The Poisson–Boltzmann (PB) equation is the combination of the Poisson equation, which relates the Laplacian of the electric potential ϕ to the distribution of charges n , and the Boltzmann distribution, which describes how the distribution of ions is related to the electric potential. The one-dimensional PB equation is given by:

$$\frac{d^2\phi(y)}{dy^2} = -\frac{e}{\epsilon_0\epsilon_w} \sum_{i=+, -} z_i n_i^0 \exp(-z_i e \beta \phi(y)), \quad (3)$$



where $z_i = \pm\{1,2\}$ is the ion valence and n^0 the bulk number density. This mean-field equation contains no information about the discrete nature of the ions or solvent. The latter is treated as a continuum with a constant relative permittivity $\epsilon_w = 58$, the dielectric constant of TIP4P/2005 water.⁶¹ This value is lower than the experimental permittivity of water, but is used for consistency since we are interested in comparing the theoretical data with our simulation results. The surface charge density of the wall σ_s enters the differential equation *via* the boundary conditions: $\epsilon_0 \epsilon_w \hat{n}_y \cdot \nabla_y \phi(\text{wall}) = -\sigma_s$ and $\nabla_y \phi(0) = 0$, where $\hat{n}_y = \pm 1$ indicates the sign of the y -direction normal to the wall.

The PB equation successfully predicts the charge distribution in systems with small ion concentrations, surface charge densities, and ion valencies, such as in the case of some colloidal systems or dilute electrolytes. On the other hand, systems with large charge concentration or surface charge densities give rise to effects that are generally not accounted for in the PB model. For instance, in situations with a large surface charge density, or large counter-ions: due to steric repulsions it might not be energetically preferable to screen the surface charge *via* a dense Stern layer. This results in ion crowding in a wider region near the surface.⁶² The PB equation does not account for this possibility, since information about the shapes and sizes of ions is not included. Similarly, ion-specific information is needed to explain specific (non-Coulombic) adsorption effects. Furthermore, electrostatic ion-ion correlations are typically very important in the case of ionic liquids or concentrated multivalent electrolytes. Finally, the boundary conditions described above require the surface charge to be located at a fixed location at the edge of the domain using which the differential equation is solved. However, in reality the charge is often discretely distributed among the sites on a rough surface and ions might penetrate into the amorphous matrix.

A rich variety of modified Poisson-Boltzmann (MPB) equations have been proposed in the literature, to predict ion distributions more accurately under certain conditions. Examples of modifications that have been studied in the literature are: the addition of ion specificity,^{63–65} the inclusion of electrostatic and non-electrostatic ion-ion correlations,^{66–68} and accounting for variations in the dielectric constant of the solvent.^{69,70} Most of these studies involved a model electrolyte near a flat wall. Here, our objective is to use a simple model that is capable of predicting the ion distribution near an amorphous discretely charged wall. The structure and charge distribution of the solid surface affect the effective ion-solid interaction. This now becomes dependent not only on the y -position of the ion (*i.e.* the position with respect to the center of the channel), but also on its lateral position relative to the wall. A number of other effects are important to accurately describe the fluid-solid interface. For example, when a solvated ion approaches the interface within a few angstroms, the hydration shell around the ion forms a resistance for the ion to approach the surface closer. Whether an ion and the surface remain separated by a solvent layer or a direct ion-surface contact is formed depends on the balance between the interactions between ions, water, and the surface.⁷¹ The balance between these competing interactions

governs to what extent ion adsorption and the corresponding EDL are ion-specific or dependent on the surface properties. For example, hydrophilicity is known to have an influence on ion adsorption.^{20,51} Another interfacial effect that could be accounted for in a continuum model is the image force that an ion feels near the interface between two media with different dielectric constants, such as the silica-water interface. Attempts have been made to predict the effective adsorption energy by including each of these contributions individually in a model.^{69,72} However, this requires a large number of approximations and empirical quantities and has shown to be only moderately successful. Alternatively, we suggest to extend the PB model by including the Helmholtz free energy associated with a single solvated ion near a charge-neutral surface. The simulated ion distributions at finite ion concentration (see Fig. 2) are the results of the combined effect of ion-specific adsorption, ion-ion correlations, dielectric variations, and the surface charge density. By calculating the free energy profile for a single ion we investigate how important specific adsorption is for the ion distribution normal to the surface. Furthermore, the surface needs to be charge-neutral since the effect of the negative surface charge on the ion distribution is already accounted for *via* the boundary conditions. To express the free-energy in terms of a potential of mean force (PMF) we modify the Boltzmann distribution as follows:

$$n_i = n_i^0 \exp(-z_i e \beta \phi - \beta U_i^{\text{PMF}}), \quad (4)$$

where U_i^{PMF} is the PMF corresponding to ion species i . This modified distribution contains now the sum of the electrostatics due to the surface charge density and the contribution of the single-ion free energy profile near a charge-neutral wall. This approach is strictly equivalent to the one-body potential contribution in density functional theory (DFT). In a similar fashion, an ion-ion correlation contribution can be added to the modified Boltzmann distribution.⁷³

The surface charge density can be expressed as an average accessible surface area per unit charge: $159 \text{ \AA}^2 \text{ e}^{-1}$, which is much larger than the specific charge area of cesium ($8.9 \text{ \AA}^2 \text{ e}^{-1}$), the largest ion in our study. Therefore, steric effects between ions are assumed to be minor, while the influence of the ion size on the ion-surface energy is accounted for in the PMF. The low packing density of the ions in combination with low valencies suggests that electrostatic ion correlations will also not be dominant in our case. The presence of a (polar) solvent can affect the properties of the EDL since the dielectric tensor of the solvent varies locally. Dielectric variations are not explicitly taken into account in the MPB equation, but the solvent structure near the interface is indirectly included in the PMFs. However, since the PMFs are calculated near a neutral surface, the dielectric variations that are accounted for in our model do not include the effect that the surface charge has on the orientation of interfacial water molecules. Studies that include dielectric variations in the electrostatic model can be found in the literature.^{74–77}

The single-ion PMFs, which are calculated using the Umbrella Sampling technique (see Section 2.3), are shown in Fig. 3 for each of the ion species considered in this work. Each of the PMFs shows a minimum around $y = -18 \text{ \AA}$. These minima are



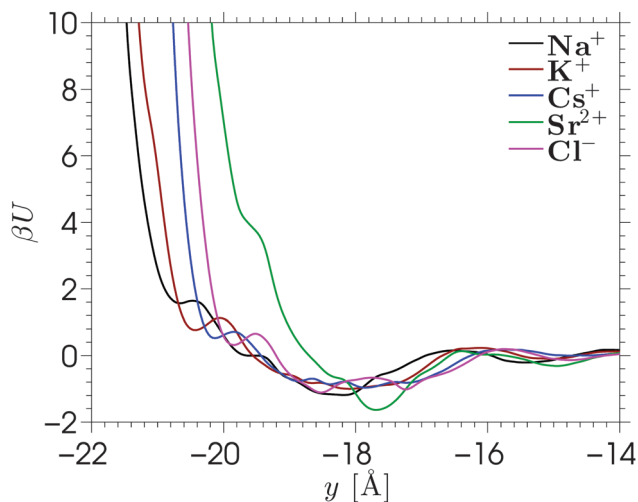


Fig. 3 The potential of mean force for single ions near a neutral hydrophilic wall. The profiles correspond to Na^+ (black), K^+ (red), Cs^+ (blue), Sr^{2+} (green) and Cl^- (magenta). These profiles are obtained by Umbrella Sampling at different x - z positions. The resulting PMF is an average over 18 profiles to represent the amorphous structure of the silica.

smooth and shallow compared to the free energy profiles corresponding to single x - z positions (shown in Fig. S7 in the ESI,[†] with further details shown in Fig. S8 and S9). The free energy converges to zero further from the wall, where the ion-wall interaction and the influence of the wall on the solvent structure are small. The profiles are calculated over the whole pore width, but the figure shows the interesting region near the silica surface. Very close to the wall the free energy asymptotes due to direct contact with wall atoms, leading to steric repulsions. The location of the free energy asymptotes for the monovalent cations shows an ordering based on the sizes of the ions ($\text{Na}^+ < \text{K}^+ < \text{Cs}^+$), with the smallest ion approaching nearest to the silica surface. These PMFs overlap almost perfectly if they are shifted horizontally to the left by the Lennard-Jones radii of the respective ions, implying that we observe no size-selective penetration of pockets at the surface. The PMF corresponding to Sr^{2+} asymptotes further from the silica surface, despite the small ionic radius. This counterintuitive result is caused by the strong hydration shell of the divalent ion. Hence, the asymptote of this PMF corresponds to a solvent-separated contact, while direct contact with the silica surface would require a greater energy to sufficiently distort the hydration shell.

The MPB equation is solved for each of the electrolytes by substituting the PMFs corresponding to the cations and Cl^- into eqn (4). We show the result corresponding to NaCl in Fig. 4. The results for other electrolyte solutions, which are qualitatively similar, can be found in Fig. S10 in the ESI.[†] Solutions of the conventional nonlinear Poisson-Boltzmann (PB) equation and the MPB model are compared against the MD simulation data for a surface charge density of $\sigma_s = -0.1 \text{ C m}^{-2}$. Fig. 4 shows the number density profiles of Na^+ (full lines) and Cl^- (dash-dotted lines). The inset of the figure shows a magnification of the diffuse part of the EDL. As expected, the PB density profiles show a diffuse regime in which the cation density decays monotonically

towards the bulk density, while the anion density increases monotonically towards this value. The edge of the domain for the PB equation is chosen at $y_{\text{wall}} = -21.2 \text{ \AA}$, which corresponds to the average y -position of the hydrogen atoms at the surface. While the actual location of the amorphous wall is ambiguous, the choice of the location of the edge of the domain (where the boundary condition is applied) has a strong effect on the solution of the standard PB equation.

The ion density profiles predicted by the MPB model show a much better agreement with the MD data for each of the electrolytes, especially for KCl and CsCl . Due to the addition of the PMFs, the cation density at the edge of the domain is now no longer the maximum value, instead a density peak is formed. This also allows us to locate the edge of the domain slightly further into the wall ($y_{\text{wall}} = -23 \text{ \AA}$) to account for the fact that ions can penetrate the surface beyond the location of the surface groups. The figure shows that MPB reproduces the magnitude and location of the cation distribution near an amorphous wall. However, a few differences between the model and the MD simulations can be seen. The difference between the shape of the cation layer obtained from simulations and theoretical prediction is explained by the localized charges at the surface. The cations adsorb to the wall at energetically preferable locations, being at the charge sites. The one-dimensional mean-field theory does not account for the fact that the electrostatic energy near the wall varies across the x - z plane, and that ions are located where it is energetically favorable. Instead, MPB assumes that ions will be homogeneously distributed across the x - z plane. Similarly, in the calculation of the PMF it was assumed that each x - z position is equally probable (hence the arithmetic average of the densities of state). A different approach to solving the MPB problem would be to solve the model separately for each of the energy profiles (corresponding to different x - z

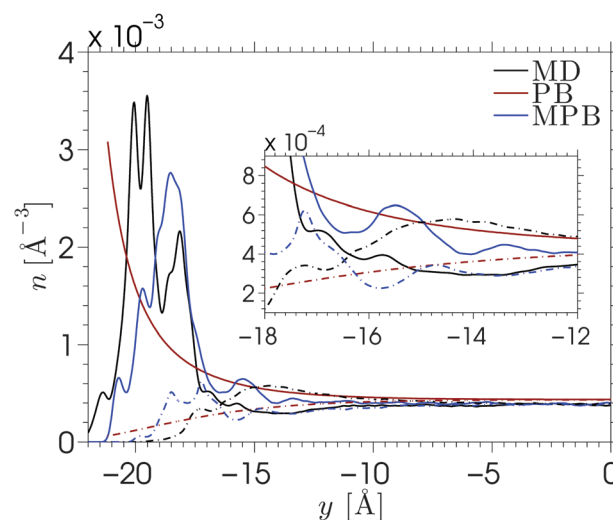


Fig. 4 Ion distribution profiles for a NaCl solution near a surface with charge density $\sigma_s = -0.1 \text{ C m}^{-2}$. The full lines correspond to cation distributions, while the dash-dotted lines correspond to anion distributions. The inset shows a magnification of the diffuse region of the EDL, where it can be seen that the anion density exceeds the cation density for the MD results, while both theoretical predictions do not show this behavior.



positions) and the average between the solutions of the model. This would then allow for applying different boundary conditions to the different calculations, effectively providing a method to include the inhomogeneous distribution of the surface charge in the model. This approach is not attempted in the current work.

The predicted location of the cation density peak is a consequence of the balance between the free energy of a single ion near an uncharged wall and the electrostatic cation–wall attraction that enters the equation *via* the boundary conditions. Since the boundary condition is not directly applied at the fluid–solid interface, the electrostatic field that is felt in the channel will be slightly smaller when the boundary condition is applied closer to the fluid–solid interface. We find that the solution of the MPB is dominated by the PMF, such that the location of the edge of the domain has little influence on the density profiles. However, since the PMFs were calculated for single-ion systems, the prediction of the anion peak is based on the chloride–wall interaction, while in reality the adsorbed layer of cations would strongly affect the location of the anions in the electric double layer.[‡] The fact that it is energetically preferable for the chloride ions to be located near the cation layer rather than near the silica surface shows important consequences for overscreening. Our MPB model is not able to account for this ion–ion effect. More generally, our model fails to predict overscreening since dielectric variations and electrostatic correlations are not taken into account. Nonlocal electrostatic correlations would enable the model to predict the partial depletion of cations and enhancement of anions directly adjacent to the Stern layer.^{16,78} Such corrections were not considered here since our main objective is to extend the simple mean-field theory to account for the specific adsorption at the rugged silica surface.

5 Electro-osmotic flow

When an electric field \mathbf{E} is applied to the confined aqueous electrolyte, a force $\mathbf{f}_i = ez_i\mathbf{E}$ acts on the ions, where z_i is the valence of ion species i . Consequently, cations are accelerated in the direction of the field, and anions in the opposite direction. Ions create a fluid flow by dragging the surrounding water along, where the rate at which energy is transferred onto the water depends on the velocities of the ions as well as the interaction forces with the surrounding water molecules. The resulting electro-osmotic flow (EOF) profile is a product of the competing positive and negative contributions of ions to the flow. The electrolyte solutions in this study have an overall positive charge, which would suggest a net flow in the direction of the electric field. However, the adsorption of cations onto the silica surface causes many cations near the interface to be partially or fully immobilized (which means that these ions transfer less or no energy to the water), such that the effective flow velocity is not linearly proportional to the excess of positive charges present in the fluid. In fact, the ion mobility and density

distribution in the EDL determine the EOF velocity (the velocity in the center of the channel), since the fluid beyond the EDL is charge-neutral and thus has no net contribution to the flow. Interestingly, this means that the EOF velocity is independent of the channel width, as long as the double layers of both channel walls are separated from each other (typically for concentrated solutions and/or large channels). The ion distributions in Fig. 2 showed that for our data charge-neutrality is reached approximately at 8 Å from the center of the channel.

We apply an electric field $E_x = -d\phi/dx = 0.2 \text{ V nm}^{-1}$ in the positive x -direction. We have confirmed that the response to this field strength is still in the linear regime and that the flow velocity induced by this electric field remains very small compared to the velocity corresponding to thermal energy. A Nosé–Hoover thermostat, with a relaxation time of 100 fs, is coupled to the directions perpendicular to the electric field to ensure that the thermostat does not interfere with the streaming motion. We start from an equilibrium configuration and equilibrate the system for 10 ns with the field turned on. We then average over a steady state for another 100 ns by storing the data every 25 ps. Such long simulations are needed because the competing ion contributions result in a small EOF velocity and thus also a small signal-to-noise ratio.⁷⁹ The cation, anion, and water velocity profiles are shown in Fig. 5. The statistical uncertainty in the data is of similar magnitude as the fluctuations in the profiles and thus omitted for clarity. The anion

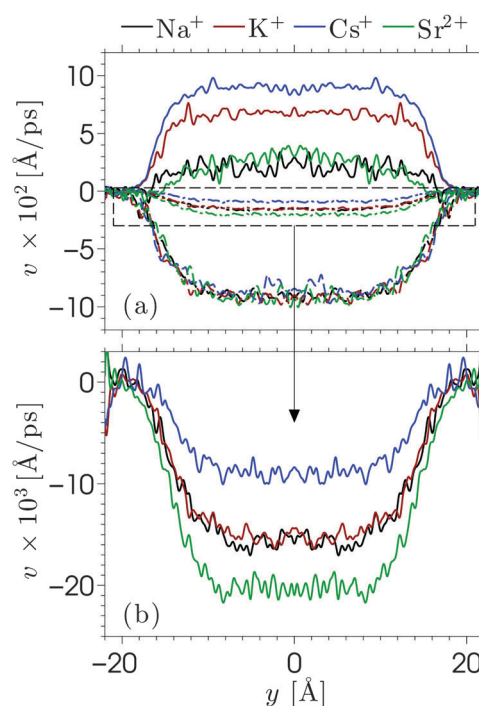


Fig. 5 Electro-osmotic velocity profiles under the influence of an electric field of $E = 0.2 \text{ V nm}^{-1}$. The top (a) shows the velocities of the cations (full lines: Na^+ (black), K^+ (red), Cs^+ (blue), and Sr^{2+} (green)), and the anion (dashed lines) and water (dash-dotted lines)) corresponding to the cation profile in the same colors. The bottom (b) shows a magnification of the water velocity profiles. The fluid temperature is kept at 300 K by thermostating the directions perpendicular to the flow.

[‡] To demonstrate this point we show in Fig. S11 in the ESI† the free energy profiles of the ions calculated from MD simulations. The chloride free energy profiles asymptote further from the wall than in Fig. 3.



velocities are approximately the same for each of the electrolytes, while the cation velocities vary strongly between the different electrolytes. The cation velocities increase with their diffusivity in bulk water taken at similar ion concentration.⁴² This suggests that the ion migration in water is the key-mechanism for ion transport, even under the influence of an electric field. The water velocity profiles are in between the cation and anion profiles, with a total mass flux in the negative directions. The ordering of these profiles indicates that flow reversal is connected to overscreening, but this does not fully explain the EOF profiles. We have seen that the overscreening for NaCl is larger than for KCl, while their EOF velocity profiles are quite comparable. Ion-specific hydration properties and differences in the enhanced apparent viscosity effect for both simulation systems play a role. If we only consider the cations, the velocity profile corresponding to Na⁺ is larger than that of K⁺. However, the number of Na⁺ ions that contribute to the flow is slightly smaller due to stronger adsorption at the surface. Furthermore, fully hydrated Na⁺ has a smaller coordination number than K⁺ (6 and 7, respectively), but the binding strength with the surrounding water molecules is larger for Na⁺. All these factors contribute to some extent to the resulting EOF velocity.

Continuum theory can help to elucidate the effect of overscreening and viscosity enhancement on the electro-osmotic flow. The one-dimensional conservation of momentum equation for a system at the steady-state is given by the Stokes equation:

$$\eta \frac{d^2 v(y)}{dy^2} + \sum_{i=+,-} e z_i n_i(y) E_x = 0, \quad (5)$$

where η is the shear viscosity constant for a bulk solution ($\eta = 1.07$ mPa s for 0.6 M NaCl).⁴² This equation uses the assumption that the fluid properties are homogeneous in the directions parallel to the wall. By substituting the density profiles of Fig. 2 into eqn (5) and integrating twice with respect to y we obtain a prediction for the velocity profile. This is shown in Fig. 6 for a NaCl solution. The lines labeled 'EMD' indicate velocity profiles that are predicted from the conservation of momentum by using the ion distribution profiles obtained from equilibrium MD simulations (Fig. 2). The data corresponding to eqn (5) are labeled as ' $\eta = c$ ', which indicates that the shear viscosity is assumed to be constant across the channel. The figure shows that the predicted flow velocity is positive while the simulated EOF velocity profile is predominantly negative. The overestimation of eqn (5) stems from the inaccurate assumption that the cations near the walls fully contribute to the flow. We can account for the enhanced viscosity effect by inserting a location-dependent shear viscosity in the momentum balance equation:

$$\frac{d}{dy} \left(\eta(y) \frac{d}{dy} v(y) \right) + \sum_{i=+,-} e z_i n_i(y) E_x = 0. \quad (6)$$

The shear viscosity profile is calculated from the local parallel diffusion coefficient $D_{\parallel}(y)$ via the Stokes–Einstein relation:

$$\eta(y) = \frac{k_B T}{3\pi\sigma D_{\parallel}(y)}, \quad (7)$$

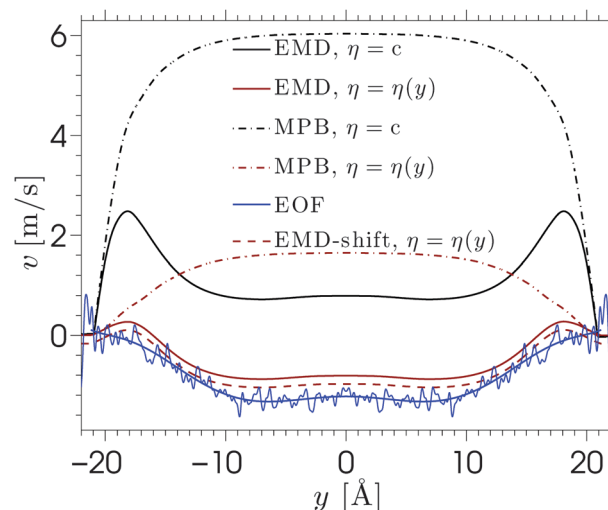


Fig. 6 The electro-osmotic velocity profile for a NaCl solution under the influence of an electric field $E_x = 0.2$ V nm⁻¹. Non-equilibrium MD results are shown in blue. The data are shown with a 6th order polynomial fit through it as a guide to the eye. The predictions from equilibrium MD data are shown by the black and red full lines, where the black line shows the result of the Stokes equation, which assumes a constant viscosity, and the red line is the solution if a calculated viscosity profile is used. The dash-dotted lines are obtained by solving the same equations using the density profiles predicted by the MPB model. The dashed red line obtained from the equilibrium MD data, the viscosity profile, and the knowledge of the location of the shear plane.

where σ signifies a hard-sphere diameter. The diffusion coefficient profile is computed from the mean square displacement in slabs, as described by Předota *et al.*⁸⁰ The effective diameter in eqn (7) is chosen as $\sigma = 2$ Å, which results in a viscosity in the center of the channel that agrees with the bulk viscosity. The viscosity and diffusion profiles for a NaCl solution are shown in Fig. 7.

The viscosity near the wall increases with respect to the bulk value and shows an asymptote around $y = \pm 22$, which is beyond the point where any flow is observed in the EOF simulations (see Fig. 5). The velocity profile predicted from eqn (6) shows good qualitative agreement with the EOF velocity measured from the MD simulations. The large difference between the two predicted velocity profiles illustrates the effect that the local viscosity enhancement has on the electrokinetic flow velocity. Note that the predicted velocity profile is very sensitive to the viscosity profile, while an accurate calculation of a local viscosity profile is hard and controversial.⁸¹ Nevertheless, the effect that the enhanced viscosity has on the velocity profile is confirmed even when the viscosity profile is assumed to be a simple step function.⁷⁰ So far, the only data used for the predicted velocity profile have been obtained from equilibrium MD simulations. Rather than using equilibrium MD data, we can solve eqn (5) and (6) using the density profiles predicted using the MPB equation. These results are also shown in Fig. 6. Since the MPB equation did not predict overscreening, the contribution of cations exceeds that of the anions everywhere across the channel. Therefore, the velocity is positive everywhere and monotonically increases towards the middle of the channel. Accounting for the enhanced viscosity effect decreases



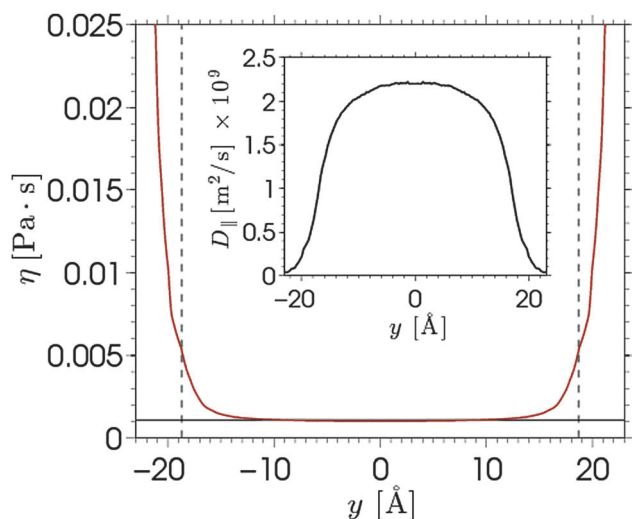


Fig. 7 The equilibrium shear viscosity profile calculated from the parallel diffusion coefficient profile via the Stokes–Einstein relation. The black horizontal line represent the bulk viscosity, while the dashed vertical lines indicate the location of the shear plane in electro-osmotic flow simulations. The inset shows the parallel diffusion coefficient profile of water.

somewhat the positive contribution of the cations, but still it cannot capture the qualitatively different trend seen in the simulation data. These results illustrate the large effect that even a small amount of overscreening or viscosity enhancement has on the EOF velocity profile.

Further predictive improvement can be achieved if the location of the shear plane is known. This is the location at which the fluid velocity is zero, which is an important concept in experimental studies.[§] This could serve as a boundary condition for eqn (6). The dashed line in Fig. 6 shows the predicted velocity profile shifted down such that its shear plane coincides with the one observed in the EOF simulations. This leads to very good agreement. Thus, the first layer of solvent and ions appear to be almost immobile.⁸² The fact that these immobile ions cannot transmit charges or momentum to the fluid strongly reduces, and even inverts, the electro-osmotic flow. This phenomenon can be modelled by a space-dependent viscosity, or by a traditional method which considers a specific location for the plane of shear. It should be noted that no slip occurs at the silica surface and stick boundary conditions can be applied for hydrophilic silica surfaces, as has been demonstrated using Poiseuille flow.¹⁴

Another important quantity in electrokinetics is the ζ -potential, which represents the electric potential at the location of the shear plane. This quantity is typically derived from electrokinetic experiments by measuring the flow velocity u_{EOF} in the bulk region of the channel and applying the Helmholtz–Smoluchowski (HS) equation:

$$\zeta = -\frac{u_{\text{EOF}}\eta}{\varepsilon_0\varepsilon_w E_x} \quad (8)$$

[§] Note that there could be additional positions with zero streaming velocity, located between the cation and anion layers moving in opposing directions. This might be expected in the case of concentrated multivalent electrolytes.

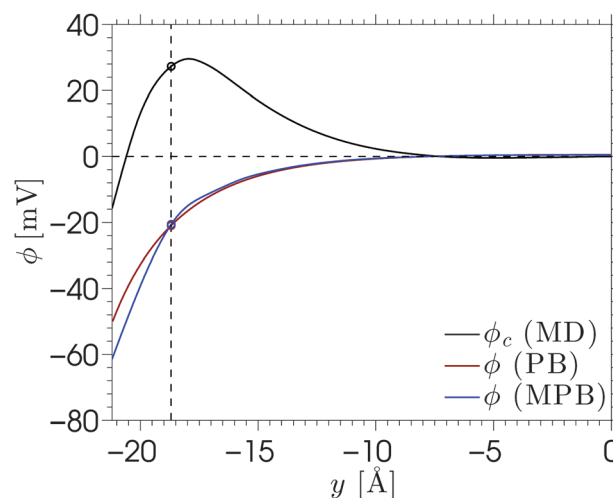


Fig. 8 The electric potential profile ϕ for a NaCl solution near a silica surface with a charge density $\sigma_s = -0.1 \text{ C m}^{-2}$. The black line indicates the MD data, the red line indicates the PB result and the blue line indicates the MPB result. The dashed lines are added as a guide for the eye, with the horizontal line located at $\phi = 0$ and the vertical line indicating the position of the shear plane $y = -18.7 \text{ Å}$.

This equation can be derived by substituting eqn (3) into eqn (5) and integrating the expression with a constant shear viscosity. The assumption of a constant shear viscosity can lead to large deviations in the calculated ζ -potential. While the viscosity enhancement is the largest in the adsorbed layer at the wall, we also found a significant viscosity increase in the diffuse part of the EDL.⁸⁰ The surface potential can be calculated from the ζ -potential if the positions of the shear plane and the wall are known. However, these positions are not always well-defined or easy to determine experimentally. To test the applicability of the HS equation for our system, we compare the ζ -potential calculated using this equation to the one calculated from equilibrium simulations. Using the bulk viscosity for a 0.6 M NaCl solution: $\eta = 1.07 \text{ mPa s}$ and the measured electro-osmotic flow velocity $u_{\text{EOF}} = -1.55 \text{ m s}^{-1}$, the predicted potential is $\zeta = 16.3 \text{ mV}$. We can measure the ζ -potential from simulations by substituting the ion density profiles into eqn (3) and integrating twice. The obtained electric potential profile is shown in Fig. 8 for the MD data as well as the Poisson–Boltzmann models (PB and MPB). At the shear plane location $y = -18.7 \text{ Å}$ we read a potential of $\zeta = 27.3 \text{ mV}$, which shows a 67% difference with respect to the HS equation. This illustrates the importance of local viscosity variations. Finally, the large influence of overscreening on the ζ -potential is clear from the Poisson–Boltzmann models which show negative ζ -potentials, of -20.9 mV (PB) and -20.6 mV (MPB).

6 Conclusions

We have presented a study of ion adsorption and transport of aqueous solutions confined in an amorphous charged silica slit pore. The structure of the electric double layer was investigated, as well as its relationship with electro-osmotic flow. Theoretical



models were presented to predict ion distributions and electro-osmotic velocity profiles. The results presented in this work have important implications for applications ranging from water purification to ion binding to biological membranes. Interesting insights into the dynamics of confined aqueous solutions can be gained from residence times at the pore surface as well as ion pair times. Such typical times, which usually fall in the range between [0.1–1 ns], are attainable using conventional MD simulations and can therefore be used to describe the physical origins at the heart of electrokinetics phenomena such as electroosmosis.

Small differences between the adsorption and dynamical properties of confined electrolytes can potentially be used for applications in which ions need to be separated. We have presented a study of the structure and electrokinetic flow of four aqueous electrolyte solutions (NaCl, KCl, CsCl, and SrCl₂) confined in charged amorphous porous silica. Our simulation results show a typical picture of overscreening in the electric double layer. The amount of overscreening is the greatest for SrCl₂ and much smaller for the monovalent electrolytes NaCl > KCl > CsCl. This ordering is explained on the basis of the sizes and valencies of the cations. A larger ion charge density and localization of charges give rise to stronger correlations, which can lead to overscreening. Particular attention has been devoted to discussing the implications of large interaction energies between ions and surface charge sites. Strong ion binding can cause an ergodic hindrance when the energy barrier associated with unbinding is much larger than the thermal energy of an ion. It has been shown that the monovalent ions in our system do not suffer from such ergodic hindrances, while additional measures are needed to study the adsorption of Sr²⁺ at the charged silica surface. One of the possible measures is to perform multiple simulations which probe different sections of phase space. Averaging between these simulations gives a more reliable sampling of the true density of states. Another possibility is to use Monte Carlo simulations to equilibrate the system and validate our molecular dynamics results. However, equilibrating with MC suffers from some limitations; for instance, specific MC steps have to be developed in order to allow for collective redistributions which are crucial to equilibrate strongly correlated systems such as electrolyte solutions. Moreover, we note that the use of molecular dynamics in the present work would remain required for the study of transport properties.

A simple modified Poisson–Boltzmann model has been used to predict the adsorption of cations onto the negatively charged disordered silica surface. We have used the potential of mean force to add ion-specific information to the Poisson–Boltzmann model and at the same time account for the disordered structure of the silica surface. This approach has shown to be successful in predicting the location of the ions in the mean-field framework, while the in-plane structure caused by the localized surface charges is not captured by the one-dimensional equation.

Electro-osmotic flow simulations have shown that flow reversal occurs for each of the electrolytes. This is caused by

a combination of overscreening and enhancement of the apparent viscosity in the electric double layer. The contributions of both phenomena to the electro-osmotic velocity become clear by comparing the solutions of continuum equations with different profiles for the viscosity and ion density. Small changes in the viscosity or density profiles have large consequences for the predicted electro-osmotic flow profile. Viscosity enhancement reduces the positive contribution to the velocity around the Stern layer, while overscreening contributes negatively to the velocity profile. Viscosity enhancement and overscreening also affect the ζ -potential. We have shown that the classical Helmholtz–Smoluchowski equation leads to a large under-prediction of the ζ -potential for our system, while the true value can be calculated from the ion density profiles.

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References

- 1 B. Rotenberg and I. Pagonabarraga, *Mol. Phys.*, 2013, **111**, 827–842.
- 2 R. B. Schoch and P. Renaud, *Appl. Phys. Lett.*, 2005, **86**, 253111.
- 3 J. Eijkel and A. van den Berg, *Microfluid. Nanofluid.*, 2005, **1**, 249–267.
- 4 L. Bocquet and E. Charlaix, *Chem. Soc. Rev.*, 2010, **39**, 1073–1095.
- 5 B. Coasne, A. Galarneau, R. J.-M. Pellenq and F. Di Renzo, *Chem. Soc. Rev.*, 2013, **42**, 4141–4171.
- 6 D. Argyris, D. R. Cole and A. Striolo, *J. Phys. Chem. C*, 2009, **113**, 19591–19600.
- 7 D. Argyris, N. R. Tummala, A. Striolo and D. R. Cole, *J. Phys. Chem. C*, 2008, **112**, 13587–13599.
- 8 D. Argyris, D. R. Cole and A. Striolo, *Langmuir*, 2009, **25**, 8025–8035.
- 9 A. P. Markesteijn, R. Hartkamp, S. Luding and J. Westerweel, *J. Chem. Phys.*, 2012, **136**, 134104.
- 10 Q. Zhang, K.-Y. Chan and N. Quirke, *Mol. Simul.*, 2009, **35**, 1215–1223.
- 11 A. A. Milischuk and B. M. Ladanyi, *J. Chem. Phys.*, 2011, **135**, 174709.
- 12 A. A. Hassanali and S. J. Singer, *J. Comput. – Aided Mater. Des.*, 2007, **14**, 53–63.
- 13 A. A. Hassanali and S. J. Singer, *J. Phys. Chem. B*, 2007, **111**, 11181–11193.
- 14 B. Siboulet, B. Coasne, J.-F. Dufreche and P. Turq, *J. Phys. Chem. B*, 2011, **115**, 7881–7886.
- 15 H. Yoshida, H. Mizuno, T. Kinjo, H. Washizu and J.-L. Barrat, *J. Chem. Phys.*, 2014, **140**, 214701.



- 16 I. Kalcher, J. C. F. Schulz and J. Dzubiella, *J. Chem. Phys.*, 2010, **133**, 164511.
- 17 P. A. Bonnaud, B. Coasne and R. J.-M. Pellenq, *J. Chem. Phys.*, 2012, **137**, 064706.
- 18 D. Argyris, D. R. Cole and A. Striolo, *ACS Nano*, 2010, **4**, 2035–2042.
- 19 C. D. Lorenz, P. S. Crozier, J. A. Anderson and A. Travesset, *J. Phys. Chem. C*, 2008, **112**, 10222–10232.
- 20 P. E. Videla, J. Sala, J. Marti, E. Guardia and D. Laria, *J. Chem. Phys.*, 2011, **135**, 104503.
- 21 H. Zhang, A. A. Hassanali, Y. K. Shin, C. Knight and S. J. Singer, *J. Chem. Phys.*, 2011, **134**, 024705.
- 22 E. R. Cruz-Chu, A. Aksimentiev and K. Schulten, *J. Phys. Chem. C*, 2009, **113**, 1850–1862.
- 23 R. Renou, A. Ghoufi, A. Szymczyk, H. Zhu, J.-C. Neyt and P. Malfreyt, *J. Phys. Chem. C*, 2013, **117**, 11017–11027.
- 24 N. R. Haria and C. D. Lorenz, *Phys. Chem. Chem. Phys.*, 2012, **14**, 5935–5944.
- 25 N. R. Haria and C. D. Lorenz, *J. Phys. Chem. C*, 2015, **119**, 12298–12311.
- 26 H. Chen, A. Z. Panagiotopoulos and E. P. Giannelis, *Langmuir*, 2015, **31**, 2407–2413.
- 27 P. Weisz, *CHEMTECH*, 1973, **3**, 498–505.
- 28 J. L. Mertz, Z. H. Fard, C. D. Malliakas, M. J. Manos and M. G. Kanatzidis, *Chem. Mater.*, 2013, **25**, 2116–2127.
- 29 A. Sachse, A. Merceille, Y. Barré, A. Grandjean, F. Fajula and A. Galarneau, *Microporous Mesoporous Mater.*, 2012, **164**, 251–258.
- 30 B. Coasne, L. Viau and A. Vioux, *J. Phys. Chem. Lett.*, 2011, **2**, 1150–1154.
- 31 F. Villemot, A. Galarneau and B. Coasne, *J. Phys. Chem. C*, 2014, **118**, 7423–7433.
- 32 L. Zhuravlev, *Colloids Surf., A*, 2000, **173**, 1–38.
- 33 S. Bhattacharya, B. Coasne, F. R. Hung and K. E. Gubbins, *Langmuir*, 2009, **25**, 5802–5813.
- 34 B. Coasne and P. Ugliengo, *Langmuir*, 2012, **28**, 11131–11141.
- 35 J. L. F. Abascal and C. Vega, *J. Chem. Phys.*, 2005, **123**, 234505.
- 36 D. Rozmanov and P. G. Kusalik, *J. Chem. Phys.*, 2012, **136**, 044507.
- 37 S. Plimpton, *J. Comput. Phys.*, 1995, **117**, 1–19.
- 38 R. W. Hockney and J. W. Eastwood, *Computer simulation using particles*, Taylor & Francis, Inc., Bristol, PA, USA, 1988.
- 39 I. S. Joung and T. E. Cheatham, *J. Phys. Chem. B*, 2009, **113**, 13279–13290.
- 40 F. Moucka, M. Lisal and W. R. Smith, *J. Phys. Chem. B*, 2012, **116**, 5468–5478.
- 41 S. Mamatkulov, M. Fyta and R. R. Netz, *J. Chem. Phys.*, 2013, **138**, 024505.
- 42 R. Hartkamp and B. Coasne, *J. Chem. Phys.*, 2014, **141**, 124508.
- 43 S. H. Lee and P. J. Rossky, *J. Chem. Phys.*, 1994, **100**, 3334–3345.
- 44 T. A. Ho and A. Striolo, *J. Chem. Phys.*, 2013, **138**, 054117.
- 45 G. Torrie and J. Valleau, *J. Comput. Phys.*, 1977, **23**, 187–199.
- 46 S. Kumar, J. M. Rosenberg, D. Bouzida, R. H. Swendsen and P. A. Kollman, *J. Comput. Chem.*, 1992, **13**, 1011–1021.
- 47 S. Bogusz, T. E. Cheatham and B. R. Brooks, *J. Chem. Phys.*, 1998, **108**, 7070–7084.
- 48 R. B. Schoch, J. Han and P. Renaud, *Rev. Mod. Phys.*, 2008, **80**, 839–883.
- 49 J. Lyklema, *Colloids Surf., A*, 2006, **291**, 3–12.
- 50 J. Faraudo and A. Travesset, *J. Phys. Chem. C*, 2007, **111**, 987–994.
- 51 C. Calero, J. Faraudo and D. Bastos-González, *J. Am. Chem. Soc.*, 2011, **133**, 15025–15035.
- 52 F. H. J. van der Heyden, D. Stein, K. Besteman, S. G. Lemay and C. Dekker, *Phys. Rev. Lett.*, 2006, **96**, 224502.
- 53 M. Chen, D. Jiang, K. Jiang and Y. Qiu, *Proc. – Inst. Mech. Eng.*, 2014, 1–4.
- 54 R. Qiao and N. R. Aluru, *Phys. Rev. Lett.*, 2004, **92**, 198301.
- 55 Y. Chen, Z. Ni, G. Wang and D. Xu, Li, *Nano Lett.*, 2008, **8**, 42–48.
- 56 S. Dewan, V. Carnevale, A. Bankura, A. Eftekhari-Bafrooei, G. Fiorin, M. L. Klein and E. Borguet, *Langmuir*, 2014, **30**, 8056–8065.
- 57 G. Hurwitz, G. R. Guillen and E. M. Hoek, *J. Membr. Sci.*, 2010, **349**, 349–357.
- 58 A. Szymczyk, P. Fievet, M. Mullet, J. Reggiani and J. Pagetti, *J. Membr. Sci.*, 1998, **143**, 189–195.
- 59 L. N. Ho, S. Clauzier, Y. Schuurman, D. Farrusseng and B. Coasne, *J. Phys. Chem. Lett.*, 2013, **4**, 2274–2278.
- 60 S. Clauzier, L. N. Ho, M. Pera-Titus, D. Farrusseng and B. Coasne, *J. Phys. Chem. C*, 2014, **118**, 10720–10727.
- 61 C. Vega and J. L. F. Abascal, *Phys. Chem. Chem. Phys.*, 2011, **13**, 19663–19688.
- 62 M. Z. Bazant, B. D. Storey and A. A. Kornyshev, *Phys. Rev. Lett.*, 2011, **106**, 046102.
- 63 I. Borukhov, D. Andelman and H. Orland, *Electrochim. Acta*, 2000, **46**, 221–229.
- 64 D. M. Huang, C. Cottin-Bizonne, C. Ybert and L. Bocquet, *Phys. Rev. Lett.*, 2007, **98**, 177801.
- 65 L. Joly, C. Ybert, E. Trizac and L. Bocquet, *Phys. Rev. Lett.*, 2004, **93**, 257805.
- 66 M. Jardat, J.-F. Dufreche, V. Marry, B. Rotenberg and P. Turq, *Phys. Chem. Chem. Phys.*, 2009, **11**, 2023–2033.
- 67 I. Kalcher, J. C. F. Schulz and J. Dzubiella, *Phys. Rev. Lett.*, 2010, **104**, 097802.
- 68 M. Boström, F. W. Tavares, D. Bratko and B. W. Ninham, *J. Phys. Chem. B*, 2005, **109**, 24489–24494.
- 69 P.-A. Cazade, R. Hartkamp and B. Coasne, *J. Phys. Chem. C*, 2014, **118**, 5061–5072.
- 70 D. J. Bonthuis and R. R. Netz, *Langmuir*, 2012, **28**, 16049–16059.
- 71 E. Leontidis, M. Christoforou, C. Georgiou and T. Delclos, *Curr. Opin. Colloid Interface Sci.*, 2014, **19**, 2–8.
- 72 D. M. Huang, C. Cottin-Bizonne, C. Ybert and L. Bocquet, *Langmuir*, 2008, **24**, 1442–1450.
- 73 D. Frydel and Y. Levin, *J. Chem. Phys.*, 2012, **137**, 164703.
- 74 A. Ghoufi, A. Szymczyk, R. Renou and M. Ding, *Europhys. Lett.*, 2012, **99**, 37008.



- 75 V. Ballenegger and J.-P. Hansen, *J. Chem. Phys.*, 2005, **122**, 114711.
- 76 D. J. Bonthuis, S. Gekle and R. R. Netz, *Phys. Rev. Lett.*, 2011, **107**, 166102.
- 77 J. Sala, E. Guardia and J. Marti, *J. Chem. Phys.*, 2010, **132**, 214505.
- 78 Y. Rosenfeld, *Phys. Rev. Lett.*, 1989, **63**, 980–983.
- 79 R. Qiao, *Microfluid. Nanofluid.*, 2007, **3**, 33–38.
- 80 M. Předota, P. T. Cummings and D. J. Wesolowski, *J. Phys. Chem. C*, 2007, **111**, 3071–3079.
- 81 R. Hartkamp, A. Ghosh, T. Weinhart and S. Luding, *J. Chem. Phys.*, 2012, **137**, 044711.
- 82 B. Siboulet, J. Molina, B. Coasne, P. Turq and J.-F. Dufreche, *Mol. Phys.*, 2013, **111**, 3410–3417.

