



Cite this: *Phys. Chem. Chem. Phys.*,
2015, 17, 28088

Gas phase selective hydrogenation over oxide supported Ni–Au

Fernando Cárdenas-Lizana and Mark A. Keane*

The chemoselective continuous gas phase ($T = 573\text{ K}$; $P = 1\text{ atm}$) hydrogenation of nitroarenes (*p*-chloro-nitrobenzene (*p*-CNB) and *m*-dinitrobenzene (*m*-DNB)) has been investigated over a series of oxide (Al_2O_3 and TiO_2) supported Au and Ni–Au (1:10 mol ratio; 0.1–1 mol% Au) catalysts. Monometallic supported Au with mean particle size 3–9 nm promoted exclusive formation of *p*-chloroaniline (*p*-CAN) and *m*-nitroaniline (*m*-NAN). Selective hydrogenation rate was higher over smaller Au particles and can be attributed to increased surface hydrogen (from TPD measurements) at higher metal dispersion. (STEM analysis has confirmed an equivalent metal particle size for the supported bimetallics at the same Au loading where TPR indicates Ni–Au interaction and EDX surface mapping established Ni in close proximity to Au on isolated nanoparticles with a composition (Au/Ni) close to the bulk value ($= 10$). Increased spillover hydrogen due to the incorporation of Ni in the bimetallics resulted in elevated $-\text{NO}_2$ group reduction rate. Full selectivity to *p*-CAN was maintained over all the bimetallic catalysts. Conversion of *m*-DNB over the lower loaded Ni–Au/ Al_2O_3 generated *m*-NAN as sole product. An increase in Ni content (0.01 \rightarrow 0.1 mol%) or a switch from Al_2O_3 to TiO_2 as support resulted in full $-\text{NO}_2$ reduction (to *m*-phenylenediamine). Our results demonstrate the viability of Ni-promotion of Au in the continuous production of functionalised anilines.

Received 16th January 2015,
Accepted 2nd March 2015

DOI: 10.1039/c5cp00282f

www.rsc.org/pccp

1. Introduction

The combination of two metallic elements has been shown to be an effective means of improving catalytic performance in a range of hydrogen mediated reactions, notably chemoselective hydrogenation (of nitro-compounds¹), partial hydrogenation (of alkynes²) and hydrogenolysis (of chlorophenols³). Renewed interest in Au as a catalytic agent has encompassed Au in bimetallic formulations, as noted in the recent review by Villa *et al.*⁴ Gold has shown unique selectivity in the hydrogenation of polyfunctional reactants but delivers low reaction rates.⁵ This drawback has been ascribed to a restricted capacity for H_2 activation by dissociative chemisorption.⁶ The combination of Au with transition metals that exhibit greater H_2 chemisorption capability offers a possible means of enhancing hydrogenation rate.

The hydrogenation response over bimetallic catalysts is determined by the structure and physico-chemical properties of the supported nanoparticles, which are dependent on the distribution of both metals within the crystal nanostructure. In order to achieve any degree of catalytic synergy, the two metals must be in close proximity (if not in direct contact) on the support. The size, metal ratio and nature of the support are

critical variables. Bimetallic particles at the nano-scale ($\sim 3\text{ nm}$) can form solid solutions (homogeneous alloy particles) regardless of the miscibility gap between the two metallic elements.⁷ The incorporation of a second metal with Au (X–Au, where X = Pd,⁸ Pt⁹) has been shown to increase hydrogenation activity but at low Au/X ratios the selectivity response is affected or even governed by the second metal.^{8,9} The redox character and acid–base properties of the (oxide) carrier can act to stabilize small Au nanoparticles as a result of electron transfer across the metal–support interface and induce geometric and electronic modifications that impact on catalysis.⁵

Prior research has focused on binary metals that form (ordered or random) bulk alloys with no miscibility gap in the corresponding bulk phase diagram. This is the case with Pd–Au as the most widely studied Au-containing bimetallic.¹⁰ Interest has shifted to systems such as Ni–Au that do not mix in the bulk but can form stable alloys in the outermost surface layers.^{11,12} The incorporation of Au onto partially oxidised Ni resulted in partial oxidation of Au atoms and the formation of $\text{Au}_n\text{–O–Ni}_m$ ensembles (from XPS),¹³ which were suggested as active sites in the isomerisation of methylstyrenes over Ni–Au/ SiO_2 .¹⁴ Nishikawa *et al.*,¹⁵ studying the promotional effect of Au on Ni, demonstrated (by XRD, TEM, XAFS and ¹⁹⁷Au Mössbauer) the formation of an Ni–Au alloy, following reduction (to 673 K in H_2) of Au/NiO co-precipitates. These were active in the hydrogenolysis of benzylic alcohols with superior catalytic activity compared with RANEY[®] Ni.

Chemical Engineering, School of Engineering and Physical Sciences, Heriot-Watt University, Edinburgh, EH14 4AS, Scotland. E-mail: M.A.Keane@hw.ac.uk; Tel: +44 (0)1314514719



Chin *et al.*¹⁶ demonstrated by EXAFS/XANES analysis of Ni–Au/MgAl₂O₄ (prepared by reductive deposition of Au on Ni/MgAl₂O₄) Au → Ni electron transfer with surface alloy formation that served to inhibit carbon deposition during *n*-butane steam reforming. The addition of Au to Ni (1:4) resulted in the formation of nano-crystals with a Au-core and Ni-enriched shell structure (based on EXAFS, XPS and UV-vis) with increased hydrogenolysis activity in the conversion of 2-phenoxy-1-phenylethanol.¹⁷

We have previously examined the catalytic action of supported Ni–Au with Ni/Au mol/mol = 10 in promoting gas phase nitroarene hydrogenation.^{18,19} That work was directed at the continuous selective production of industrially important functionalised amines under mild reactions conditions. Supported Ni–Au delivered higher hydrogenation activities but was less selective than supported Au. We have refocused our attention and have considered the possible role of Ni (introduced as the minor component; Ni/Au mol/mol = 0.1) to enhance the chemoselective hydrogenation action of Au to deliver enhanced rates to the target amine product. We assess the feasibility of controlling hydrogenation rate by varying Au particle size and the nature of the oxide support.

2. Experimental

2.1 Catalyst preparation and activation

The Al₂O₃ (Puralox, Condea Vista Co.) and TiO₂ (Degussa) supports were used as received. High (HL, 0.1 mol%) and low loaded (LL, 0.01 mol%) Ni/Al₂O₃ and/or Ni/TiO₂ catalyst precursors were synthesised by standard impregnation, where the support (6 g) was contacted with an aqueous Ni(NO₃)₂ solution (Aldrich, 7 × 10^{−5}–7 × 10^{−4} M, 85 cm³). The slurry was heated (2 K min^{−1}) to 353 K and maintained under constant agitation (600 rpm) in a He purge. The solid was dried in a flow of He at 383 K for 5 h and sieved (ATM fine test sieves) to mean particle diameter = 75 μm. The Au–Ni catalyst precursors were prepared by first reducing a batch of (LL and HL) Ni/Al₂O₃ and (LL) Ni/TiO₂ in a (60 cm³ min^{−1}) H₂ stream at 2 K min^{−1} to 723 K, which was maintained for 1 h to ensure formation of zero valent Ni.²⁰ The gas flow was switched to He, cooled to ambient temperature and the samples passivated in 1% v/v O₂/He. This treatment served to provide a protective oxide layer over the surface Ni that prevented bulk oxidation upon exposure to the atmosphere. The passivated samples were treated with an aqueous HAuCl₄ solution (Aldrich, 3 × 10^{−5}–5 × 10^{−4} M, 85 cm³) to deliver a 10/1 Au/Ni mol ratio, post-treatment as above. Monometallic (HL, 1 mol% and LL, 0.1 mol%) Au/Al₂O₃ and (LL, 0.1 mol%) Au/TiO₂ were prepared by standard impregnation with aqueous HAuCl₄ (Aldrich, 7 × 10^{−4}–7 × 10^{−3} M, 85 cm³). Bulk metal loading was determined by ICP-OES (Vista-PRO, Varian Inc.). The Au-containing samples were stored under He in the dark at 277 K to avoid the deleterious effects of light (photodecomposition of cationic gold to Au⁰)²¹ and temperature (aggregation of Au precursor species).²² Before reaction, the catalyst precursors were activated in 60 cm³ min^{−1} H₂ at 2 K min^{−1}

to 723 K (Ni/Al₂O₃ and Ni/TiO₂) or 603 K (Au/Al₂O₃, Au/TiO₂, Ni–Au/Al₂O₃, Ni–Au/TiO₂), which was maintained for 1 h.

2.2 Catalyst characterisation

Specific surface area (SSA) and total pore volume measurements were made on a Micromeritics Flowsorb II 2300 unit. Prior to analysis, the samples were outgassed at 423 K for 1 h in 20 cm³ min^{−1} N₂. SSA was obtained in a 30% v/v N₂/He flow (20 cm³ min^{−1}) with at least three cycles of N₂ adsorption–desorption using the standard single-point BET method. Total pore volume was obtained at a relative N₂ pressure (P/P_0) = 0.95. The data were reproducible to ±3% and values quoted in this paper are the mean.

Temperature programmed reduction (TPR), H₂ chemisorption and temperature programmed desorption (TPD) were determined using the CHEM-BET 3000 (Quantachrome) unit. The samples were loaded into a U-shaped Quartz cell (3.76 mm i.d.) and heated in 17 cm³ min^{−1} 5% v/v H₂/N₂ (Brooks mass flow controlled) to 603 ± 1 K (Au/Al₂O₃, Au/TiO₂, Ni–Au/Al₂O₃ and Ni–Au/TiO₂) or 723 ± 1 K (Ni/Al₂O₃ and Ni/TiO₂) at 2 K min^{−1}. The effluent gas passed through a liquid N₂ trap and H₂ consumption was monitored by a thermal conductivity detector (TCD) with data acquisition/manipulation using the TPR Win™ software. The reduced samples were maintained at the final temperature in H₂ until return to baseline. After the reduction step, the samples were cooled in a He flow and subjected to H₂ chemisorption (at ambient temperature) using a pulse (10 μl) titration procedure, as described elsewhere.²³ TPD was conducted in a N₂ flow (65 cm³ min^{−1}) at 50 K min^{−1} to 873 K with an isothermal hold until the signal returned to the baseline. The support alone was subjected to an equivalent TPR and subsequent TPD, which was used to correct for H₂ desorption from Al₂O₃.²⁴

Metal particle morphology, size distribution and surface composition were determined by (scanning) transmission electron microscopy using a JEOL JEM-2100F unit operating at an accelerating voltage of 200 kV with resolution to 0.14 nm and an EDAX Genesis XM 4 system 60. Samples were dispersed in 1-butanol by ultrasonic vibration, deposited on a lacey-carbon/Cu grid (200 Mesh) and dried at 383 K. Up to 300 individual metal particles were counted for each catalyst and the surface area-weighted metal diameter (d) was calculated from:

$$d = \frac{\sum_i n_i d_i^3}{\sum_i n_i d_i^2} \quad (1)$$

where n_i is the number of particles of diameter d_i .

2.3 Gas phase nitroarene hydrogenation

2.3.1 Materials. The reactants (*p*-chloronitrobenzene (*p*-CNB) and *m*-dinitrobenzene (*m*-DNB), Aldrich, ≥98%) and solvent (1-butanol, Riedel-de Haën, ≥99.5%) were used as supplied, without further purification.

2.3.2 Catalytic system. Catalytic reactions were carried out under atmospheric pressure, *in situ* immediately after activation, in a fixed bed vertical continuous flow glass reactor (i.d. = 15 mm) at 573 K. The reactions were conducted under conditions of



negligible heat/mass transport limitations. A preheating zone (layer of borosilicate glass beads) ensured that the nitroarene reactant was vaporised and reached reaction temperature before contacting the catalyst. Isothermal conditions (± 1 K) were maintained by thoroughly mixing the catalyst with ground glass (75 μm) before insertion into the reactor. Temperature was continuously monitored by a thermocouple inserted in a thermowell within the catalyst bed. The *p*-CNB (or *m*-DNB) reactant was delivered, in a co-current flow of H_2 , via a glass/teflon air-tight syringe and a teflon line, using a microprocessor controlled infusion pump (Model 100 kd Scientific) at a fixed calibrated flow rate, with an inlet $-\text{NO}_2$ molar flow ($F_{-\text{NO}_2}$) in the range 1.1×10^{-4} – 3.1×10^{-4} mol h^{-1} , where the molar metal to inlet $-\text{NO}_2$ feed rate ratio spanned the range 1.1×10^{-3} – 16.0×10^{-3} h. The H_2 content was far in excess of the stoichiometric requirement for $-\text{NO}_2$ reduction (160–430), the flow rate of which was monitored using a Humonics (Model 520) digital flowmeter; $\text{GHSV} = 2 \times 10^4$ h^{-1} . In a series of blank tests, passage of *p*-CNB or *m*-DNB in a stream of H_2 through the empty reactor or over the supports alone did not result in any detectable conversion.

2.3.3 Analytical method and activity/selectivity evaluation.

The reactor effluent was frozen in a liquid nitrogen trap for subsequent analysis using a Perkin-Elmer Auto System XL gas chromatograph equipped with a programmed split/splitless injector and a flame ionization detector, employing a DB-1 50 m \times 0.20 mm i.d., 0.33 μm film thickness capillary column (J&W Scientific). Repeated catalytic runs with different samples from the same batch of catalyst delivered product compositions that were reproducible to within $\pm 6\%$ with a carbon balance of $\pm 5\%$. Catalyst activity is quantified in terms of nitroarene (*m*-DNB ($X_{m\text{-DNB}}$) or *p*-CNB ($X_{p\text{-CNB}}$)) fractional conversion

$$X_{\text{nitroarene}} = \frac{[\text{nitroarene}]_{\text{in}} - [\text{nitroarene}]_{\text{out}}}{[\text{nitroarene}]_{\text{in}}} \quad (2)$$

where the subscripts “in” and “out” refer to inlet and outlet streams, respectively. Selectivity in terms of (*e.g.*) *m*-phenylenediamine (*m*-PDM) from *m*-DNB ($S_{m\text{-PDM}}$) is given by

$$S_{m\text{-PDM}} = \frac{[m\text{-PDM}]_{\text{out}}}{[m\text{-DNB}]_{\text{in}} - [m\text{-DNB}]_{\text{out}}} \times 100 \quad (3)$$

3. Results and discussion

3.1 Catalyst characterisation

The metal content, specific surface area (SSA), pore volume, H_2 uptake/release during TPR/TPD and mean metal particle size for the (mono- and bi-metallic) catalysts used in this study are given in Table 1. The SSA and pore volume of the Al_2O_3 support (191 $\text{m}^2 \text{g}^{-1}$; $450 \times 10^{-3} \text{cm}^3 \text{g}^{-1}$) are close to values reported for commercial Puralox mesoporous $\gamma\text{-Al}_2\text{O}_3$ (157 $\text{m}^2 \text{g}^{-1}$; $420 \times 10^{-3} \text{cm}^3 \text{g}^{-1}$).²⁵ Incorporation of Au and Ni resulted in a measurable decrease in SSA and pore volume that can be ascribed to partial pore filling by the metal(s) or Al_2O_3 dissolution during impregnation resulting in pore collapse.²⁶ The TiO_2 supported systems exhibited similar SSA and pore volumes to the starting support (52 $\text{m}^2 \text{g}^{-1}$; $120 \times 10^{-3} \text{cm}^3 \text{g}^{-1}$) that are in accord with Degussa P25 (61 $\text{m}^2 \text{g}^{-1}$; $120 \times 10^{-3} \text{cm}^3 \text{g}^{-1}$).²⁷

The TPR profiles generated for Au/ Al_2O_3 -HL (I), Au/ Al_2O_3 -LL (II) and Au/ TiO_2 -LL (III) can be compared in Fig. 1. The three samples present a principal H_2 consumption peak with an

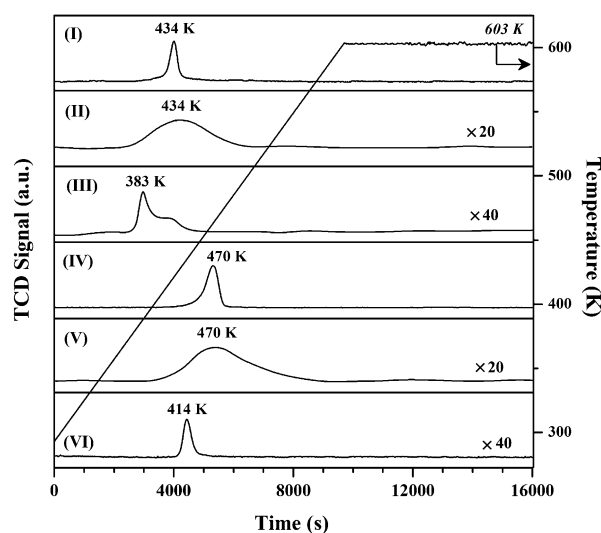


Fig. 1 TPR profiles generated for: (I) Au/ Al_2O_3 -HL, (II) Au/ Al_2O_3 -LL, (III) Au/ TiO_2 -LL, (IV) (1:10) Ni-Au/ Al_2O_3 -HL, (V) (1:10) Ni-Au/ Al_2O_3 -LL and (VI) (1:10) Ni-Au/ TiO_2 -LL.

Table 1 Physico-chemical characteristics of Al_2O_3 and TiO_2 supported mono-(Au) and bi-metallic (Ni-Au) catalysts

Catalyst ^a	SSA ^b ($\text{m}^2 \text{g}^{-1}$)	Pore volume $\times 10^{-3}$ ^c ($\text{cm}^3 \text{g}^{-1}$)	TPR T_{max} (K)	TPR H_2 ^{d,e} ($\mu\text{mol g}_{\text{catalyst}}^{-1}$)	H_2 TPD ($\text{mmol mol}_{\text{Au}}^{-1}$)	d^f (nm)
Au/ Al_2O_3 -HL	161	427	434	102 ^d /104 ^e	473	9
Ni-Au/ Al_2O_3 -HL	189	444	470	151 ^d /145 ^e	3118	9
Au/ Al_2O_3 -LL	169	425	434	9 ^d /16 ^e	1101	4
Ni-Au/ Al_2O_3 -LL	190	447	470	9 ^d /8 ^e	2471	4
Au/ TiO_2 -LL	49	117	383	11 ^d /15 ^e	1233	3
Ni-Au/ TiO_2 -LL	47	115	414	12 ^d /15 ^e	1765	3

^a Au metal content (mol%): LL = 0.1, HL = 1.0, 1:10 Ni:Au molar ratio in bi-metallics. ^b Specific surface area (SSA), $\text{Al}_2\text{O}_3 = 191 \text{m}^2 \text{g}^{-1}$, $\text{TiO}_2 = 52 \text{m}^2 \text{g}^{-1}$. ^c Pore volume, $\text{Al}_2\text{O}_3 = 450 \times 10^{-3} \text{cm}^3 \text{g}^{-1}$, $\text{TiO}_2 = 120 \times 10^{-3} \text{cm}^3 \text{g}^{-1}$. ^d H_2 required for the reduction of the metal precursor. ^e Experimentally determined H_2 consumption. ^f Surface area weighted mean metal particle size from TEM/STEM analysis.



associated temperature maximum (T_{\max}) over 383–434 K, which is within the range recorded elsewhere (328–465 K) for the activation of Au/Al₂O₃ and Au/TiO₂ and attributed to precursor reduction.²⁸ The lower T_{\max} recorded for Au/TiO₂-LL (relative to Au/Al₂O₃) is consistent with the literature.²⁹ Moreover, Delannoy *et al.*³⁰ have demonstrated (by *in situ* XAFS and DRIFTS) greater reducibility of cationic gold on TiO₂ compared with Al₂O₃. Hydrogen consumption during catalyst activation was close to that required for Au³⁺ reduction (Table 1). Supported bimetallic catalyst preparation involved reductive deposition where Ni with a lower electrochemical potential (ECP = −0.27) acts to reduce the Au precursor (HAuCl₄, ECP = 1.00).³¹ The TPR profiles generated for Al₂O₃ and TiO₂ supported Ni (not shown) exhibited a single broad H₂ consumption peak at the final isothermal hold (723 K) that can be associated with a combined decomposition-reduction of the precursor to metallic nickel, as noted elsewhere.¹ A single stage reduction was observed for the supported Ni–Au systems at a higher T_{\max} (414–470 K) than that recorded for the corresponding monometallic Au catalysts. Modification to the TPR response for Au containing bimetallic catalysts with respect to monometallic counterparts has been reported and linked to interaction between both metals.³² Pu *et al.*³³ have recently described a stabilisation of Au species (on activated carbon spheres) with the introduction of Ni by co-impregnation that inhibited reduction, displacing the temperature for Au^{III} → Au⁰ (from 583 K to 588 K) and Au^I → Au⁰ (from 619 K to 659 K) to higher values.

Gold particle size, from TEM analysis (Fig. 2 and 3 and Table 1), is dependent on metal loading and the nature of the support. The particle size in Au/Al₂O₃-HL was in the range 1–20 nm (Fig. 3(I)) with a surface area weighted mean diameter of 9 nm. The lower Au loaded sample (Au/Al₂O₃-LL) exhibited a narrower size distribution (1–8 nm, Fig. 3(II)) and smaller mean (4 nm). This effect is consistent with literature that has shown wider size range and the formation of larger nanoparticles with increasing Au content.³⁴ This can be linked to mobility of the chloride precursor, resulting in Au agglomeration during thermal treatment.³⁵ Enhanced Au dispersion on TiO₂ (Fig. 3(III)) is a consequence of a difference in metal/support interaction relative to Au/Al₂O₃.³⁶ Loss of lattice oxygen with the formation of surface vacancies is a feature of thermal treatment in H₂ for reducible oxides such as TiO₂ at $T \geq 573$ K.³⁷ These surface defects act as Au nucleation sites with electronic interactions that limit particle growth, leading to enhanced dispersion.^{38,39}

Representative (S)TEM images of Ni–Au/Al₂O₃-HL (I), Ni–Au/Al₂O₃-LL (II) and Ni–Au/TiO₂-LL (IV) can be compared in Fig. 4. The metal size range in Ni–Au/Al₂O₃-HL (1–20 nm) coincided with Au/Al₂O₃-HL and there was no obvious alteration due the inclusion of Ni. This response extended to Ni–Au/Al₂O₃-LL and Ni–Au/TiO₂-LL with an equivalent mean metal size relative to the monometallic system (see Table 1). This can be attributed to the low Ni content in the bimetallic samples. An unchanged metal particle size has been reported previously for TiO₂ supported Au and Pt–Au with low Pt content (0.01–0.03 wt%).⁹ EDX mapping (see representative analysis for Ni–Au/Al₂O₃-HL in Fig. 4(II)) demonstrated that individual particles contained both Au and Ni in proportions close to the bulk value (Au/Ni = 10).

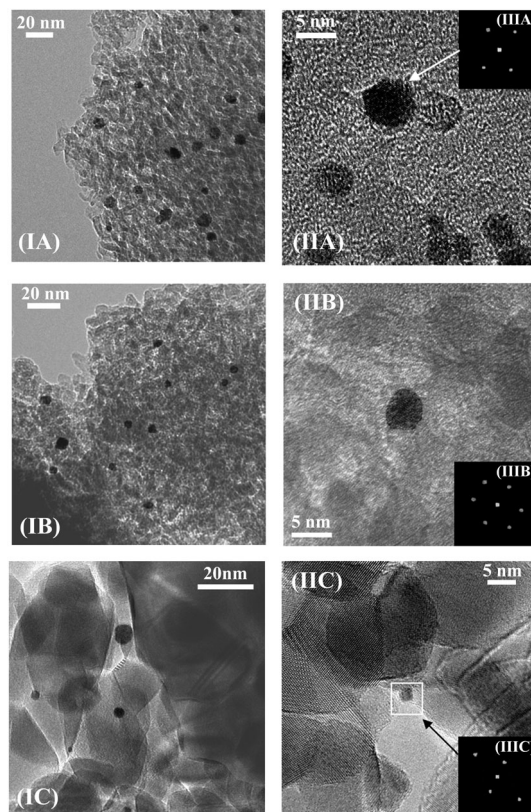


Fig. 2 Representative (I) medium and (II) high resolution TEM images with (III) associated diffractogram pattern of an isolated single Au nanoparticle in (A) Au/Al₂O₃-HL, (B) Au/Al₂O₃-LL and (C) Au/TiO₂-LL.

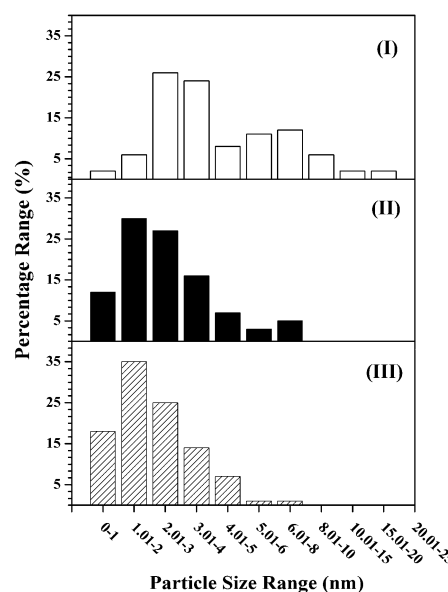


Fig. 3 Gold particle size distributions associated with (I) Au/Al₂O₃-HL (open bars), (II) Au/Al₂O₃-LL (solid bars) and (III) Au/TiO₂-LL (hatched bars).

EDX analysis cannot establish unequivocally the exact nature of Ni–Au interaction or rule out the occurrence of bimetallic clusters⁴⁰ or surface alloy.⁴¹ Both possibilities have been suggested



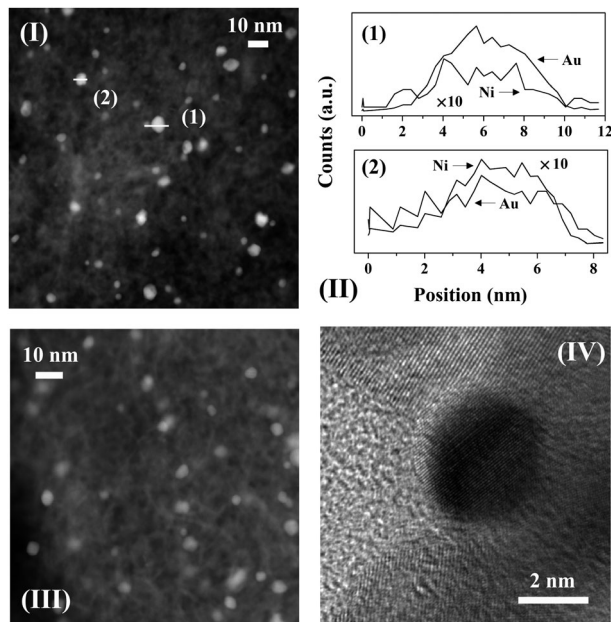


Fig. 4 Representative TEM/STEM images of (I) Ni-Au/Al₂O₃-HL showing (1) 12 nm and (2) 8 nm segments across individual metal particles subjected to (II) EDX mapping with the associated Ni (counts $\times 10$) and Au distribution, (III) Ni-Au/Al₂O₃-LL and (IV) Ni-Au/TiO₂-LL.

for Ni-Au catalysts prepared in an analogous manner and with a similar metal dispersion to that in this study.^{16,42,43} Deghedi and co-workers⁴² proposed the formation of Ni particles covered with Au adatoms (on the basis of EXAFS and TEM-EDX measurements) for Ni-Au/SiO₂ synthesised by redox deposition that contained metal ensembles of 5.1 nm. Maniecki and co-workers⁴⁴ reported an alloy phase (on the basis of XRD) for Ni-Au/Al₂O₃ prepared by co-impregnation (with Ni(NO₃)₂ and HAuCl₄) and reduction in H₂. Surface alloy formation was suggested by EXAFS/XANES for Au-Ni/MgAlO₄ (10.1 nm) prepared by redox deposition.¹⁶ Molenbroek and Nørskov⁴³ also demonstrated the occurrence of a surface alloy in Ni-Au/SiO₂ (2.5–6.0 nm) and Ni-Au/MgAl₂O₄ (3.0–15.0 nm) by Monte Carlo simulations, which was experimentally verified by EXAFS, TEM and *in situ* XRD.

The hydrogenation of nitroarenes is enhanced by contributions due to spillover hydrogen.^{19,36} The generation of spillover species involves dissociative adsorption of H₂ on a (donor) supported metal site with migration of the atomic hydrogen generated to the (acceptor) support.⁴⁵ Hydrogen TPD is a practical approach that can serve to quantify the degree of spillover. Ambient temperature H₂ chemisorption following TPR on all the catalysts was low (≤ 15 mmol mol_{Au}⁻¹) and close to the detection limits. Low uptake on supported Au has been demonstrated elsewhere and attributed to the filled d-band and high activation energy barrier for dissociative adsorption.^{6,46} There was no measurable difference in ambient temperature H₂ uptake on the bimetallics, which may be due to the low Ni content. The H₂ TPD profiles are presented in Fig. 5. In all cases, the profiles show H₂ release with maxima at $T \geq 770$ K where the total amount desorbed (per mol_{Au}, Table 1) was appreciably greater (by two orders of magnitude) than that taken up in the chemisorption step. This response

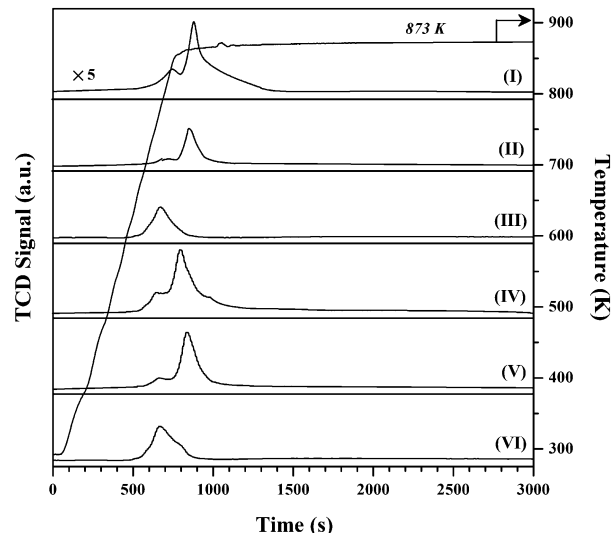


Fig. 5 Hydrogen TPD profiles generated for: (I) Au/Al₂O₃-HL, (II) Au/Al₂O₃-LL, (III) Au/TiO₂-LL, (IV) (1:10) Ni-Au/Al₂O₃-HL, (V) (1:10) Ni-Au/Al₂O₃-LL and (VI) (1:10) Ni-Au/TiO₂-LL.

suggests the release of spillover hydrogen generated during catalyst activation by TPR.⁴⁵ The shift in the TPD peak for TiO₂ systems to a lower temperature (by up to a 100 K) is indicative of a less energetically demanding desorption of H₂. This is consistent with literature^{47,48} that has reported a down-shift (by *ca.* 60–70 K) in the H₂-TPD peak for Ni/TiO₂ and Pd-Au/Al₂O₃-TiO₂ compared with Ni/Al₂O₃ and Pd-Au/Al₂O₃, respectively. Taking the three monometallic Au catalysts presented in Fig. 5, H₂ desorbed (Table 1) increased in the order: Au/Al₂O₃-HL (I) < Au/Al₂O₃-LL (II) < Au/TiO₂-LL (III) where the values obtained are close (54–932 mmol mol_{Au}⁻¹) to those reported elsewhere for Au with similar metal size (3–5 nm) supported on Fe₂O₃.³⁶ This sequence reflects that of decreasing mean Au nanoparticle size (Table 1). There is evidence in the literature²³ that metal size is a crucial variable controlling spillover where a greater (donor) surface area for smaller metal nanoparticles extends the donor/acceptor interface and facilitates spillover transfer. The profiles generated for the bimetallic Ni-Au catalysts (Fig. 5, profiles IV–VI) reveals a significant increase in H₂ desorbed relative to Au samples with the same loading (see Table 1). This suggests a surface Ni/Au synergism that impacts on H₂ adsorption/desorption dynamics. The existing literature on Au containing bimetallics for hydrogenation applications has been from the perspective of Au as a diluent with low H₂ uptake capacity and a resultant decrease in surface hydrogen associated with the bimetallic.⁴⁹ As a result, there is no directly comparable H₂ TPD analysis for supported bimetallic Au-containing systems where the second metal was present in small amounts. We should flag the work of Wojcieszak *et al.*⁵⁰ who observed an increase in H₂ released (0.22 \rightarrow 0.38 mol g⁻¹) during TPD of Ni-Ag/SiO₂ with increasing Ni content (0.50 \rightarrow 0.75 wt%). It is important to note that Ni-Au/Al₂O₃-HL with lower metal dispersion showed significantly greater H₂ desorption (Table 1) relative to Ni-Au/Al₂O₃-LL. STEM-EDX analyses have demonstrated the presence



of both metals on individual nanoparticles. The formation of bimetallic nano-crystals (*i.e.* Au decorating Ni nanoparticles) has been demonstrated (EXAFS and TEM-EDS) for Ni–Au bimetals prepared and activated using a similar methodology.⁴² It is known that H₂ interaction is distinct for monometallic, bimetallic and alloy nanoparticles.^{51,52} The greater Ni loading in Ni–Au/Al₂O₃-HL (0.1 mol% *vs.* 0.01 mol%) must facilitate H₂ dissociative adsorption and spillover during TPR, which is reflected in greater H₂ release during TPD. In the case of the low loading bimetals, H₂ desorption from Ni–Au/Al₂O₃-LL exceeded Ni–Au/TiO₂-LL (Table 1). This can be tentatively linked to modifications induced by the support that impact on H₂ uptake on the bimetallic phase. Likewise, Gu *et al.*⁴⁷ observed a greater H₂ TPD from Pd–Au/Al₂O₃ relative to Pd–Au/Al₂O₃-TiO₂. The characterisation analysis has established the formation of nano-scale (mean = 3–9 nm) Au particles in the monometallics with greater surface hydrogen for catalysts with increasing metal dispersion. TPR measurements suggest interaction of Ni with Au in the bimetals. STEM-EDX analysis has demonstrated the presence of Au and Ni in individual nanoparticles and the TPD results are consistent with increased spillover resulting from Au/Ni synergy that increases the available surface hydrogen.

3.2 Gas phase nitroarene hydrogenation

The hydrogenation of *p*-CNB and *m*-DNB over all the catalysts delivered time-invariant conversions (not shown) and exclusive nitro-group reduction with no evidence of hydrodechlorination, hydrodenitrogenation or aromatic ring reduction. This combined stability/selectivity is important in terms of the viable application of Au catalyst formulations in the production of functionalised anilines. Catalyst deactivation in hydrogenation (of naphthalene,⁵³ 1,3-butadiene⁵⁴ and acetylene^{55,56}) has been reported for Au supported on Al₂O₃^{53,54,57} and TiO₂^{55,56} and ascribed to metal sintering^{54,55} and coking.^{55–57} Moreover, temporal loss of activity in the gas phase hydrogenation of nitroarenes over Pd/Al₂O₃ has been observed and linked to coke formation.^{58,59}

The applicability of pseudo-first order kinetics for the gas phase hydrogenation of *p*-CNB^{1,60} and *m*-DNB³⁶ has been established elsewhere. The extracted specific (per mol_{Au}) pseudo-first order rate constants (*k*) are given in Table 2. The promoting effect of Ni is demonstrated by the significantly increased hydrogenation rate when compared with the monometallic Au catalysts with the same loading. Under the same reaction conditions, Ni/Al₂O₃ and Ni/TiO₂ delivered negligible activity. Duan and co-workers⁶¹ using unsupported nano-crystals reported increased rate (mol_{product} mol_{metal}⁻¹) in the high pressure (40 atm) liquid phase hydrogenation of phenol (283 → 363) and benzene (164 → 261) over Ni–Rh relative to Rh, which was attributed to the bimetallic character of the nanoparticles (based on HRTEM-EDX and XRD) and possible modifications in active site electron density. Lu *et al.*^{62,63} recorded higher activity in the liquid phase hydrogenation of a series of *p*-substituted nitro-derivates over Ni–Pd colloids, ascribed to the formation of a Ni enriched alloy (from XRD, EXAFS, XPS). The elevated rate observed in this study can be linked to surface available reactive hydrogen as shown in Fig. 6 where H₂-TPD is related to the rate constant.

Table 2 Catalysis results: pseudo-first order rate constant (*k*) and product distribution (at *X* ~ 0.07) in the hydrogenation of *p*-CNB and *m*-DNB over Al₂O₃ and TiO₂ supported monometallic (Au) and bimetallic (Ni–Au; 1 : 10 molar ratio) catalysts; reaction conditions: *P* = 1 atm, *T* = 573 K

Catalyst	<i>p</i> -CNB		<i>m</i> -DNB	
	<i>k</i> ^a	Product(s) (<i>S</i> _{product} (%))	<i>k</i> ^a	Product(s) (<i>S</i> _{product} (%))
Au/Al ₂ O ₃ -HL	4	<i>p</i> -CAN (100)	6	<i>m</i> -NAN (100)
Ni–Au/Al ₂ O ₃ -HL	24	<i>p</i> -CAN (100)	40	<i>m</i> -NAN (60) <i>m</i> -PDM (40)
Au/Al ₂ O ₃ -LL	12	<i>p</i> -CAN (100)	16	<i>m</i> -NAN (100)
Ni–Au/Al ₂ O ₃ -LL	21	<i>p</i> -CAN (100)	35	<i>m</i> -NAN (100)
Au/TiO ₂ -LL	18	<i>p</i> -CAN (100)	22	<i>m</i> -NAN (100)
Ni–Au/TiO ₂ -LL	22	<i>p</i> -CAN (100)	25	<i>m</i> -NAN (80) <i>m</i> -PDM (20)

^a mol_{-NO₂} mol_{Au}⁻¹ h⁻¹.

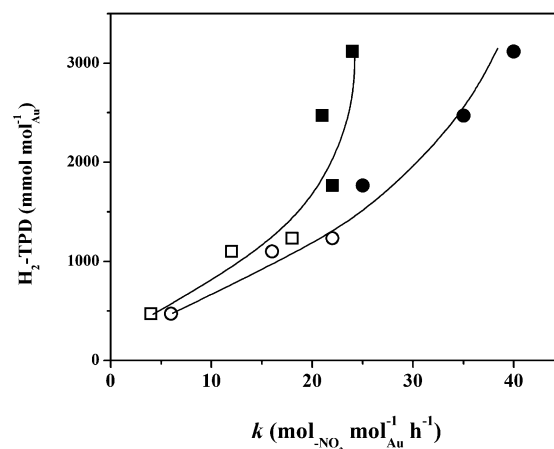


Fig. 6 Variation in hydrogen TPD (H₂-TPD) with pseudo-first order rate constant (*k*) in the hydrogenation of *p*-CNB (□, ■) and *m*-DNB (○, ●) over supported Au (open symbols) and Ni–Au (solid symbols) catalysts.

The increase in surface hydrogen that results from incorporation of Ni with Au was accompanied by a higher hydrogenation rate. Moreover, reaction of *m*-DNB uniformly generated greater rates, a result that can be attributed to the activating effect of the second –NO₂ group, which is consistent with a nucleophilic mechanism, as demonstrated previously.⁶⁴

Reaction selectivity is compared at the same degree of conversion in Table 2. *p*-CNB was converted solely to the target (*p*-chloroaniline) *p*-CAN over all the catalysts, which is consistent with reaction exclusivity in continuous gas phase *p*-CNB → *p*-CAN over oxide supported Au.⁴⁹ In contrast, nitrobenzene,¹ aniline¹ and benzene⁶⁵ have been observed in the gas phase hydrogenation of CNB over Ru,⁶⁵ and Pd¹ catalysts. Variations in product distribution are, however, evident in the hydrogenation of *m*-DNB and sensitive to the nature of the catalyst. Full selectivity to partially reduced (*m*-nitroaniline) *m*-NAN was obtained over monometallic Au systems. In batch liquid phase operation, high pressures (26–34 atm) have been deemed essential to achieve high selectivity (84–98%) to *m*-NAN.^{48,66} Preferential –NO₂ activation on oxide supported Au in the presence of other



reactive (C=C, carbonyl, amide and ester) functional groups has been demonstrated by FTIR analysis.⁶⁷ The incorporation of Ni at low loading (Ni-Au/Al₂O₃-LL) served to increase rate while retaining 100% selective to *m*-NAN. Deghedi *et al.*⁴² reported that the combination of Au and Ni on SiO₂ (Ni:Au molar ratio = 4 : 1) lowered the rate of C=C bond and aromatic ring hydrogenation relative to Ni in the conversion of styrene, which was accounted for in terms of geometric and electronic effects. At higher loadings (Ni-Au/Al₂O₃-HL) complete hydrogenation to *m*-PDM was observed, which is a characteristic of catalysis by Ni.^{18,19} This may be a consequence of increased surface reactive hydrogen which facilitates reduction of both -NO₂ substituents. Alternatively, *m*-DNB activation on Ni-Au/Al₂O₃-HL may be distinct from Ni-Au/Al₂O₃-LL with a planar adsorption on the former through the aromatic ring, resulting in the formation of a resonance structure with two positive localised charges where both -NO₂ groups are activated for nucleophilic attack. It is known that substituted aromatics interact with supported Ni catalysts *via* the π electron-delocalised aromatic ring.⁶⁸ Lonergan *et al.*⁶⁹ have observed *n*-butane selectivity effects in 1,3-butadiene hydrogenation over Ni-Pt/Al₂O₃ with increasing Ni content (1:3 \rightarrow 1:10). The formation of *m*-PDM was also promoted over Ni-Au/TiO₂-LL, which suggests a support effect in terms of *m*-DNB adsorption/activation that acts in tandem with differences in available reactive hydrogen. Wang *et al.*⁷⁰ reported a modified catalytic response in the hydrogenation of 1,3-butadiene for Pt-Ni on Al₂O₃ and TiO₂. DFT calculations revealed stronger metal interaction with Al₂O₃ (binding energy *ca.* -50 kcal mol⁻¹) than TiO₂ (-17 kcal mol⁻¹) with a preferential Ni-Al₂O₃ interaction at the metal-support interface leading to a Pt-terminated bimetallic configuration. A weaker Ni interaction in the case of the TiO₂ carrier can then result in some segregation of the Ni component to the nanoparticle surface.

4. Conclusions

Oxide (Al₂O₃ and TiO₂) supported Au (1 and 0.1 mol%) prepared by impregnation with subsequent activation generated nano-scale metal particles of mean size 3–9 nm. A decrease in Au size was accompanied by increased surface hydrogen (from TPD analysis) with higher specific chemoselective hydrogenation rate in the conversion of *p*-CNB \rightarrow *p*-CAN and *m*-DNB \rightarrow *m*-NAN. Supported bimetallic synthesis by reductive deposition of Au on Ni (where Au/Ni = 10) generated an equivalent particle size at the same Au loading but elevates the available surface hydrogen with a significant increase in -NO₂ group reduction rate. TPR analysis suggests Ni-Au interaction where (S)TEM-EDX mapping across isolated single particles has established Au and Ni in close proximity with a surface Au/Ni that coincided with the bulk value. Reaction of *p*-CNB over the bimetallics delivered full selectivity to *p*-CAN. Conversion of *m*-DNB was fully selective to *m*-NAN over monometallic Au and Ni-Au/Al₂O₃-LL. The hydrogenation of both -NO₂ substituents (to *m*-PDM) was promoted over Ni-Au/Al₂O₃-HL and Ni-Au/TiO₂-LL. This is attributed to the higher Ni loading that elevated the surface supply of hydrogen and possible electronic/geometric modifications induced by the

support that impact on the *m*-DNB adsorption/activation. We have demonstrated Au/Ni synergy that serves to enhance -NO₂ reduction rate.

Acknowledgements

The authors are grateful to Dr Santiago Gómez-Quero, Dr Noémie Perret and Pedro Benavente Donayre for their contribution to the work.

References

- 1 F. Cárdenas-Lizana, S. Gómez-Quero, C. Amorim and M. A. Keane, *Appl. Catal., A*, 2014, **473**, 41.
- 2 A. Yarulin, I. Yuranov, F. Cárdenas-Lizana, D. T. L. Alexander and L. Kiwi-Minsker, *Appl. Catal., A*, 2014, **478**, 186.
- 3 G. Yuan, C. Louis, L. Delannoy and M. A. Keane, *J. Catal.*, 2007, **247**, 256.
- 4 A. Villa, D. Wang, D. S. Su and L. Prati, *Catal. Sci. Technol.*, 2015, **5**, 55.
- 5 F. Cárdenas-Lizana and M. A. Keane, in *Gold Catalysis: Preparation, Characterization and Applications in the Gas and Liquid Phase*, ed. L. Prati and A. Villa, Pan Stanford Publishing, Singapore, 2014.
- 6 C. Kartusch and J. A. van Bokhoven, *Gold Bull.*, 2009, **42**, 343.
- 7 G. C. Bond, *Platinum Met. Rev.*, 2007, **51**, 63.
- 8 F. Cárdenas-Lizana, S. Gómez-Quero, A. Hugon, L. Delannoy, C. Louis and M. A. Keane, *J. Catal.*, 2009, **262**, 235.
- 9 D. P. He, X. D. Jiao, P. Jiang, J. Wang and B. Q. Xu, *Green Chem.*, 2012, **14**, 111.
- 10 C. Louis, in *Environmental Catalysis over Gold-based Materials*, ed. G. Avgouropoulos and T. Tabakova, The Royal Society of Chemistry, Cambridge, 2013.
- 11 F. Besenbacher, I. Chorkendorff, B. S. Clausen, B. Hammer, A. M. Molenbroek, J. K. Nørskov and I. Stensgaard, *Science*, 1998, **279**, 1913.
- 12 L. P. Nielsen, F. Besenbacher, I. Stensgaard, E. Laegsgaard, C. Engdahl, P. Stoltze, K. W. Jacobsen and J. K. Nørskov, *Phys. Rev. Lett.*, 1993, **71**, 754.
- 13 A. G. Rybkin, D. Y. Usachev, A. Y. Varykhalov, A. M. Shikin and V. K. Adamchuk, *Phys. Solid State*, 2007, **49**, 984.
- 14 A. V. Naumkin and A. Y. Vasil'kov, *Russ. Chem. Bull.*, 2013, **62**, 2559.
- 15 H. Nishikawa, D. Kawamoto, Y. Yamamoto, T. Ishida, H. Ohashi, T. Akita, T. Honma, H. Oji, Y. Kobayashi, A. Hamasaki, T. Yokoyama and M. Tokunaga, *J. Catal.*, 2013, **307**, 254.
- 16 Y.-H. Chin, D. L. King, H.-S. Roh, Y. Wang and S. M. Heald, *J. Catal.*, 2006, **244**, 153.
- 17 J. Zhang, H. Asakura, J. van Rijn, J. Yang, P. Duchesne, B. Zhang, X. Chen, P. Zhang, M. Saeys and N. Yan, *Green Chem.*, 2014, **16**, 2432.
- 18 F. Cárdenas-Lizana, S. Gómez-Quero, G. Jacobs, Y. Ji, B. H. Davis, L. Kiwi-Minsker and M. A. Keane, *J. Phys. Chem. C*, 2012, **116**, 11166.



- 19 F. Cárdenas-Lizana, S. Gómez-Quero, C. J. Baddeley and M. A. Keane, *Appl. Catal., A*, 2010, **387**, 155.
- 20 F. Cárdenas-Lizana, S. Gómez-Quero and M. A. Keane, *Appl. Catal., A*, 2008, **334**, 199.
- 21 W.-S. Lee, B.-Z. Wan, C.-N. Kuo, W.-C. Lee and S. Cheng, *Catal. Commun.*, 2007, **8**, 1604.
- 22 R. Zanella and C. Louis, *Catal. Today*, 2005, **107–108**, 768.
- 23 C. Amorim and M. A. Keane, *J. Colloid Interface Sci.*, 2008, **322**, 196.
- 24 J. Chen, X. Liu and F. Zhang, *Chem. Eng. J.*, 2015, **259**, 43.
- 25 E. Kok, N. Cant, D. Trimm and J. Scott, *Catal. Today*, 2011, **178**, 79.
- 26 J. W. Bae, Y.-J. Lee, J.-Y. Park and K.-W. Jun, *Energy Fuels*, 2008, **22**, 2885.
- 27 T. R. Jagannath, PhD dissertation, Saurashtra University, 2005.
- 28 Y.-J. Chen and C.-T. Yeh, *J. Catal.*, 2001, **200**, 59.
- 29 S. Y. Liu and S. M. Yang, *Appl. Catal., A*, 2008, **334**, 92.
- 30 L. Delannoy, N. Weiher, N. Tsapatsaris, A. M. Beesley, L. Nchari, S. L. M. Schroeder and C. Louis, *Top. Catal.*, 2007, **44**, 263.
- 31 D. R. Lide, *Handbook of Chemistry and Physical Properties*, Taylor Francis, Boca Raton, 88th edn, 2007–2008.
- 32 F.-W. Chang, L. S. Roselin and T.-C. Ou, *Appl. Catal., A*, 2008, **334**, 147.
- 33 Y. Pu, J. Zhang, X. Wang, H. Zhang, L. Yu, Y. Dong and W. Li, *Catal. Sci. Technol.*, 2014, **4**, 4426.
- 34 F. Cárdenas-Lizana, S. Gómez-Quero, H. Idriss and M. A. Keane, *J. Catal.*, 2009, **268**, 223.
- 35 A. Schulz and M. Hargittai, *Chem. – Eur. J.*, 2001, **7**, 3657.
- 36 F. Cárdenas-Lizana, S. Gómez-Quero, N. Perret and M. A. Keane, *Catal. Sci. Technol.*, 2011, **1**, 652.
- 37 P. Claus, A. Brückner, C. Mohr and H. Hofmeister, *J. Am. Chem. Soc.*, 2000, **122**, 11430.
- 38 N. Lopez, J. K. Nørskov, T. V. W. Janssens, A. Carlsson, A. Puig-Molina, B. S. Clausen and J.-D. Grunwaldt, *J. Catal.*, 2004, **225**, 86.
- 39 P. V. Kamat, *J. Phys. Chem. C*, 2008, **112**, 18737.
- 40 D. Chen, J. Li, C. Shi, X. Du, N. Zhao, J. Sheng and S. Liu, *Chem. Mater.*, 2007, **19**, 3399.
- 41 S.-H. Zhou, Z. Ma, H.-F. Yin, Z.-L. Wu, B. Eichhorn, S. H. Overbury and S. Dai, *J. Phys. Chem. C*, 2009, **113**, 5758.
- 42 L. Deghedi, J.-M. Basset, G. Bergeret, J.-P. Candy, M. C. Valero, J.-A. Dalmon, A. de Mallmann, A.-C. Dubreuil and L. Fischer, in *10th International Symposium “Scientific Bases for the Preparation of Heterogeneous Catalysts”*, ed. E. M. Gaigneaux, M. Devillers, S. Hermans, P. Jacobs, J. Martens and P. Ruiz, Elsevier, Louvain-la-Neuve, 2010, vol. 175, p. 617.
- 43 A. M. Molenbroek and J. K. Nørskov, *J. Phys. Chem. B*, 2001, **105**, 5450.
- 44 T. P. Maniecki, K. Bawolak, D. Gebauer, P. Mierczynski and W. K. Jozwiak, *Kinet. Catal.*, 2009, **50**, 138.
- 45 W. C. Conner and J. L. Falconer, *Chem. Rev.*, 1995, **95**, 759.
- 46 J. Harris, *Surf. Sci.*, 1989, **221**, 335.
- 47 Z. Gu, L. Luo and M. Teng, *Indian J. Chem.*, 2007, **46A**, 742.
- 48 Y. X. Liu, J. X. Chen and J. Y. Zhang, *Chin. J. Chem. Eng.*, 2007, **15**, 63.
- 49 F. Cárdenas-Lizana and M. A. Keane, *J. Mater. Sci.*, 2013, **48**, 543.
- 50 R. Wojcieszak, S. Monteverdi, J. Ghanbaja and M. M. Bettahar, *J. Colloid Interface Sci.*, 2008, **317**, 166.
- 51 J. Greeley and M. Mavrikakis, *Nat. Mater.*, 2004, **3**, 810.
- 52 J. Greeley and M. Mavrikakis, *J. Phys. Chem. B*, 2005, **109**, 3460.
- 53 B. Pawelec, A. M. Venezia, V. La Parola, S. Thomas and J. L. G. Fierro, *Appl. Catal., A*, 2005, **283**, 165.
- 54 X. Zhang, H. Shi and B.-Q. Xu, *Catal. Today*, 2007, **122**, 330.
- 55 T. V. Choudhary, C. Sivadinarayana, A. K. Datye, D. Kumar and D. W. Goodman, *Catal. Lett.*, 2003, **86**, 1.
- 56 Y. Azizi, C. Petit and V. Pitchon, *J. Catal.*, 2008, **256**, 338.
- 57 F. Cárdenas-Lizana, X. Wang, D. Lamey, M. Li, M. A. Keane and L. Kiwi-Minsker, *Chem. Eng. J.*, 2014, **255**, 695.
- 58 V. Vishwanathan, V. Jayasri, P. M. Basha, N. Mahata, L. M. Sikhwihilu and N. J. Coville, *Catal. Commun.*, 2008, **9**, 453.
- 59 K. K. Yeong, A. Gavriilidis, R. Zapf and V. Hessel, *Catal. Today*, 2003, **81**, 641.
- 60 F. Cárdenas-Lizana, B. Bridier, C. C. K. Shin, J. Pérez-Ramírez and L. Kiwi-Minsker, *ChemCatChem*, 2012, **4**, 668.
- 61 H. Duan, D. Wang, Y. Kou and Y. Li, *Chem. Commun.*, 2013, **49**, 303.
- 62 P. Lu and N. Toshima, *Bull. Chem. Soc. Jpn.*, 2000, **73**, 751.
- 63 P. Lu, T. Teranishi, K. Asakura, M. Miyake and N. Toshima, *J. Phys. Chem. B*, 1999, **103**, 9673.
- 64 F. Cárdenas-Lizana, Z. Martínez de Pedro, S. Gómez-Quero and M. A. Keane, *J. Mol. Catal.*, 2010, **326**, 48.
- 65 C. Xiao, X. D. Wang, C. Lian, H. Q. Liu, M. H. Liang and Y. Wang, *Curr. Org. Chem.*, 2012, **16**, 280.
- 66 V. L. Khilnani and S. B. Chandalia, *Org. Process Res. Dev.*, 2001, **5**, 263.
- 67 K. Shimizu, Y. Miyamoto, T. Kawasaki, T. Tanji, Y. Tai and A. Satsuma, *J. Phys. Chem. C*, 2009, **113**, 17803.
- 68 E.-J. Shin and M. A. Keane, *J. Catal.*, 1998, **173**, 450.
- 69 W. W. Lonergan, D. G. Vlachos and J. G. Chen, *J. Catal.*, 2010, **271**, 239.
- 70 T. Wang, G. Mpourmpakis, W. W. Lonergan, D. G. Vlachos and J. G. Chen, *Phys. Chem. Chem. Phys.*, 2013, **15**, 12156.

