



Cite this: *CrystEngComm*, 2015, 17, 4712

Systematic solvate screening of trospium chloride: discovering hydrates of a long-established pharmaceutical†

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The ability of the antispasmodic agent trospium chloride (TCl) to form solvates was investigated by applying conventional solvate screening methods on 20 solvents. According to the solubility of TCl, different approaches were considered (slow evaporation, slurring and anti-solvent addition). Five solvates, with the solvents methanol, acetonitrile, propionitrile, *N,N*-dimethylformamide, nitromethane and dihydrate, were identified and characterized by various analytical techniques. Moreover, a solvate with isopropyl alcohol and TCl sesquihydrate was prepared circumstantially outside the systematic screening. The structures of all the solvates were determined by single crystal X-ray diffraction. The reproducible forms were further characterized by powder X-ray diffraction and desolvation behaviour was observed by thermoanalytical (TGA/DSC) methods. Structural features of novel solvates and of previously described polymorphs and cocrystals of TCl were compared, presented by a tree diagram which classifies the structures according to their molecular packing.

Received 16th April 2015,
Accepted 24th May 2015

DOI: 10.1039/c5ce00744e

www.rsc.org/crystengcomm

Introduction

Polymorph screening, which inherently includes solvate screening, is nowadays a necessary step taken from the early development stages of an API (active pharmaceutical ingredient). It is also often supplemented for long-known APIs already established on the market, as new forms might provide unexpected advantageous properties or explain unanswered problems which arise during production, transfer, generic development or reformulation to a new drug with the combination of more APIs.

The precipitation of a new solid form also presents a risk which could endanger a robust production of an API¹ and might appear even if minor changes in the technology are installed, an appearing polymorph of ritonavir representing a very challenging scenario.²

Creative methods applied for polymorph and solvate screens are being introduced,³ but basically the solvent-based techniques used in the solid form screening narrow down to solvent evaporation,⁴ anti-solvent addition⁵ and slurring.^{6,7}

The possibility of the preparation of various polymorphs can be also investigated/controlled *via* desolvation of the solvate as was reported in the case of furosemide,⁸ sulfamerazine⁹ or spironolactone¹⁰ polymorphs.

One of the long-established pharmaceutical substances which lacks thorough solid state screening is trospium chloride. Trospium chloride (TCl, Fig. 1), a muscarinic antispasmodic drug used for the treatment of an over-active bladder,

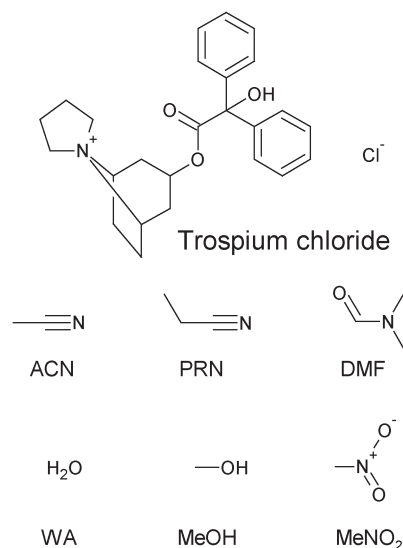


Fig. 1 Molecular structure of trospium chloride and solvents, with which it forms solvates.

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† Electronic supplementary information (ESI) available: X-ray crystallographic information files (CIF) for all solvates, PXRD patterns of the bulk samples compared with the calculated ones, detailed solubility measurement data, further structural comparisons and tables of hydrogen bonds. See DOI: 10.1039/c5ce00744e



exhibits high solubility in water and other polar solvents (ethanol, methanol) and is insoluble in nonpolar solvents. Recently, the structural characterization of new cocrystals, as well as the polymorphs of TCl, has been reported.^{11,12} An ethanol solvate has been reported as well.¹¹

The motivation of this study was to further investigate its propensity to include small molecules within the crystal structure, thus forming solvates, as trospium chloride seems to be a very accommodating molecule providing a whole range of different phases.

We proposed a conventional solvate screening investigating the solvatomorphism of TCl, using the three most common techniques: slurring, slow evaporation and anti-solvent addition on 20 solvents. The chosen solvents belong to classes 2 (“solvents to be limited”) and 3 (“solvents with low toxic potential”) according to the Q3C guideline classification¹³ and have a polar protic or polar aprotic character. We questioned the possible relation between TCl solubility in the solvents and its propensity to form solvates with them; are there any limits restricting the solvate formation?

Experimental

Materials

Trospium chloride was kindly provided by the Interpharma Praha, a.s. company (Prague, Czech Republic) as form I and was used without further purification. Solvents were purchased from various suppliers and were used as received. Due to the hygroscopicity of some of the used solvents, only freshly opened bottles of solvents were used for solubility measurements, to assure the declared water content.

Systematic screening

TCl (20 mg) was placed in a glass vial and solvent was gradually added until TCl dissolved or until 5 mL of solvent was added. The suspensions with TCl as an undissolved powder were treated as slurry screening samples. When TCl dissolved in less than 5 mL of solvent, the solutions were transferred to 50 mL round-bottomed flasks and left at ambient temperature until crystalline material was formed. The approximate solubility of TCl at room temperature was determined. All slurry experiment solvents (where the estimated solubility was around or below 4 mg per mL) were selected as being suitable for anti-solvent addition experiments.

Slow evaporation (SE)

In the slow evaporation experiments, the solutions of TCl (20 mg) in cyclohexanol, 1-butanol (*n*BuOH), ethanol (EtOH), isobutyl alcohol (*i*BuOH), methanol (MeOH), 2-methoxyethanol, *N*-methylpyrrolidone, 1-propanol (*n*PrOH) and water (WA) were placed in 50 mL round-bottomed flasks and left open at ambient temperature until crystalline material was formed.

Slurring (SI)

Slurries of TCl (20 mg) in acetone, acetonitrile (ACN), *N,N*-dimethylformamide (DMF), ethyl acetate (EtAc), isopropyl alcohol (*i*PrOH), methyl acetate (MeAc), methyl ethyl ketone (MEK), nitromethane (MeNO₂), 1-octanol, propionitrile (PRN), tetrahydrofuran (THF) and water (WA) in glass vials with caps and parafilm were placed in an orbital shaker (IKA MS 3 digital, 500 rpm) for 48 hours at ambient temperature, then filtered and analyzed.

Anti-solvent addition (AS)

The saturated solution of TCl in water (25 μL) was pipetted into a vial with 1 ml of solvent in which the solubility of TCl was below 4 mg per mL.

Solubility measurement

A more precise determination of the solubility of TCl in solvents in which it forms solvates was measured by UV-vis spectrophotometry and by gravimetry.

UV-vis spectrophotometry

Using a solution of a known concentration of TCl, the absorbance of the sample was measured on a UV-visible spectrophotometer and an appropriate wavelength for calibration was determined for each solvent. Linear calibration curves were obtained from the measurements of 5 calibration samples (prepared by dilution of a solution of known concentration) with absorbance in the range of 0.3–0.9. Saturated solutions were diluted to fall in the range of the linear calibration and concentrations of the saturated solutions were calculated.

Gravimetry

The saturated solutions of TCl were prepared, 3 mL were pipetted in weighed round-bottomed flasks and the solvents were removed by a rotary vacuum evaporator. The solid precipitates in the flasks were further dried in the oven (24 h, at 50 °C) and the flasks were reweighed. The solid precipitates were then analyzed by PXRD, confirming the presence of anhydrous TCl (form I). The procedure was repeated 3 times, and an average value of solubility is presented.

Analytical techniques

Thermal analysis TGA/DSC. We performed TGA/DSC, a combined analysis of thermogravimetry analysis and differential scanning calorimetry. Calorimeters Setaram and Setsys Evolution were used. The temperature range was 20–230 °C, in air atmosphere and with a heating rate of 10 °C min⁻¹.

Desolvation. The solvates were prepared in a slurry, dried in air and further dried in the oven at 50 °C on petri dishes until no weight loss was observed (2 h). Then the solvates were subjected to increased temperature, determined for each solvate from TGA/DSC analysis, for 6 hours. The resulting powder was reweighed and measured by PXRD.



Dynamic vapour sorption (DVS). Dynamic vapour sorption (DVS) was measured on the device DVS Advantage 1 from Surface Measurement Systems. The sample weight in a quartz crucible was between 20.3–21.1 mg, and the temperature in the device was kept between 25.3–25.4 °C.

A measuring program was used: the sample was measured by one cycle of 0% relative humidity to 90% relative humidity (adsorption) and then of 90% relative humidity to 0% relative humidity (desorption). A measuring gas of 4.0 nitrogen flow of 200 sccm was used.

Powder X-ray diffraction (PXRD). X-ray powder diffraction data for screening experiments were collected at room temperature with a laboratory X'PERT PRO MPD PANalytical diffractometer with parafocusing Bragg–Brentano geometry, using CuK α radiation ($\lambda = 1.54184 \text{ \AA}$), a measurement range 2–40° 2 θ , a step size of 0.01° 2 θ and a counting time of 0.5 s step⁻¹. Samples were ground in an agate mortar and stacked on a Si holder (zero background). Data evaluation was performed in the software package HighScore Plus.¹⁴

Single crystal X-ray diffraction (SXRD). All the structures except TCl dihydrate (diHyd, 8) were measured using an Xcalibur Onyx CCD diffractometer with graphite monochromated Cu-K α radiation. The structure of diHyd (8) was measured using a Gemini Atlas CCD diffractometer with mirror collimated Cu-K α radiation. The data reduction and absorption correction were done with CrysAlis PRO software.¹⁵ The structures were solved using charge flipping¹⁶ methods and refined by full matrix least squares on F

squared value using Jana2006 software.¹⁷ The MCE¹⁸ software was used for visualization of electron density maps. The positions of disordered solvent molecules were found in difference Fourier maps, in the case of MeOH (4) the bond length was restrained, and in the case of iPrOH (3) and PRN (6) the positions were refined using rigid body refinement. The overall occupancy of disordered solvents was constrained to 1. The hydrogen atoms were discernible in difference Fourier maps and could be refined to a reasonable geometry. According to common practice the hydrogen atoms attached to carbon atoms were placed geometrically while positions of hydrogen atoms attached to oxygen were refined with restrained bond lengths. The value of $U_{\text{iso}}(\text{H})$ was set to 1.2 U_{eq} of the parent atom (C, O). The structures were deposited in the Cambridge Structural Database. In the case of the DMF solvate (2), the isotropic extinction correction was refined (extinction coefficient 3800 (600)). For further information on the data collection and structure refinement see Table 1. The structures were compared using software: Mercury¹⁹ and Discovery Studio Viewer Pro.²⁰

Results and discussion

Screening results

We carried out a conventional solvate screening investigating the solvatomorphism of TCl. All the screening samples with a crystalline material were analyzed by PXRD and their

Table 1 Summarized crystallographic data

Compound	ACN (1)	DMF (2)	iPrOH (3)	MeOH (4)	MeNO ₂ (5)	PRN (6)	sesHyd (7)	diHyd (8)
Empirical formula	C ₂₅ H ₃₀ NO ₃ ·C ₂ H ₃ N·Cl	C ₂₅ H ₃₀ NO ₃ ·C ₃ H ₇ NO·Cl	C ₂₅ H ₃₀ NO ₃ ·C ₃ H ₈ O·Cl	C ₂₅ H ₃₀ NO ₃ ·CH ₄ O·Cl	C ₂₅ H ₃₀ NO ₃ ·CH ₃ NO ₂ ·Cl	C ₂₅ H ₃₀ NO ₃ ·C ₃ H ₅ N·Cl	2(C ₂₅ H ₃₀ NO ₃)·2(Cl)·3(H ₂ O)	C ₂₅ H ₃₀ NO ₃ ·Cl·2(H ₂ O)
Formula wt.	469.0	501.1	488.1	460.0	489.0	483.1	910.0	464.0
Crystal system	Monoclinic	Monoclinic	Monoclinic	Monoclinic	Monoclinic	Monoclinic	Monoclinic	Monoclinic
Space group	<i>P</i> ₂ ₁ / <i>n</i>	<i>P</i> ₂ ₁ / <i>n</i>	<i>P</i> ₂ ₁ / <i>n</i>	<i>P</i> ₂ ₁ / <i>n</i>	<i>P</i> ₂ ₁ / <i>n</i>	<i>P</i> ₂ ₁ / <i>n</i>	<i>P</i> ₂ ₁ / <i>c</i>	<i>P</i> ₂ ₁ / <i>c</i>
<i>a</i> (Å)	14.9067(4)	15.1865(6)	9.9394(6)	13.654(2)	9.9289(3)	15.1562(7)	39.7859(5)	19.1988(7)
<i>b</i> (Å)	9.4670(3)	9.6504(3)	9.3964(8)	9.7390(14)	9.4290(3)	9.5360(3)	6.95750(10)	13.1107(3)
<i>c</i> (Å)	18.6466(5)	19.3822(6)	28.5819(17)	18.621(2)	26.8357(10)	18.8677(8)	17.1898(4)	9.5387(3)
β (°)	108.471(2)	112.605(3)	98.318(5)	106.609(12)	97.556(3)	109.782(4)	100.0985(14)	98.175(3)
<i>V</i> (Å ³)	2495.88(13)	2622.35(17)	2641.3(3)	2372.8(6)	2490.53(14)	2566.02(19)	4684.60(14)	2376.59(13)
<i>Z</i>	4	4	4	4	4	4	4	4
<i>T</i> (K)	180	180	180	150	180	150	180	120
ρ_{calc} (g cm ⁻³)	1.248	1.269	1.227	1.287	1.304	1.250	1.290	1.297
μ (mm ⁻¹)	1.59	1.58	1.54	1.68	1.68	1.57	1.72	1.71
<i>F</i> (000)	1000	1072	1048	984	1040	1032	1944	992
Reflns. collected	35 012	11 578	24 465	13 045	16 871	24 850	98 885	9602
Indep. reflns. (<i>R</i> _{int})	5017 (0.042)	5307 (0.029)	5573 (0.043)	4800 (0.034)	5114 (0.046)	5239 (0.032)	9845 (0.061)	4145 (0.024)
GOF	1.86	1.59	1.53	1.80	1.82	1.83	1.91	1.56
<i>R</i> ₁ , <i>wR</i> ₂ [<i>I</i> > 3 σ (<i>I</i>)]	0.045, 0.119	0.038, 0.101	0.045, 0.116	0.047, 0.120	0.051, 0.131	0.042, 0.111	0.046, 0.120	0.036, 0.096
<i>R</i> ₁ , <i>wR</i> ₂ (all data)	0.056, 0.125	0.043, 0.105	0.058, 0.125	0.055, 0.126	0.059, 0.137	0.049, 0.116	0.051, 0.125	0.041, 0.099
Res. el. dens. (e Å ⁻³)	0.40, -0.29	0.19, -0.18	0.24, -0.17	0.22, -0.25	0.38, -0.27	0.24, -0.27	0.42, -0.37	0.26, -0.17
CCDC number	1056149	1056150	1056151	1056152	1056153	1056154	1056155	1056156



diffraction patterns were compared with the diffraction patterns of pure trospium chloride form I.

A total of 8 new crystal structures are reported, 6 of which were obtained repeatedly. All the obtained crystalline phases were analyzed by PXRD or SXR. The solid state purity was confirmed by Rietveld refinement (see ESI†); the single crystal structure was successfully fitted to an experimental powder diffraction pattern.

In the systematic screening, the material consumption was taken into account during making the choice of the applied screening method. Where the solubility of TCl in the solvent was low, slurring (Sl) and anti-solvent addition (AS) were applied. Due to the higher solubility of TCl in most protic solvents and, therefore, due to the increased material consumption needed for slurring, only the slow evaporation (SE) method was applied in some protic solvents.

The slurries gave solvates with ACN (1), DMF (2), MeNO₂ (5) and PRN (6) at single crystal quality for structure determination.

For anti-solvent addition (AS) experiments, the saturated solution of TCl in water was prepared, which was added into solvents where TCl had a low solubility. Solvates with ACN (1) and PRN (6) were thus prepared. When the solution was added into MEK, 1-octanol and EtAc, the dihydrate (8) was formed. In MeNO₂ three phases were prepared: solvate (5), diHyd (8) and anhydrous TCl (form I). In the other solvents anhydrous TCl form I was prepared. Thus, the method shown was more efficient for the dihydrate (8) preparation rather than for the preparation of the solvates, also the DMF solvate was not prepared by AS. Single crystals of diHyd (8) were also collected at the bottom of the vial with a saturated solution in water (slurry).

By slow evaporation, the solvate with MeOH (4) was prepared. Its single crystals for SXR structure determination were obtained from repeated experiments.

Two more structures are presented in the paper, which were prepared serendipitously. The single crystal of sesquihydrate (7) was found during cocrystal screening experiments (but its structure was not included in the paper).²¹ The single crystal of the iPrOH solvate (3) was discovered in the mixture with pure TCl recrystallized from iPrOH and the solvate has not been recollected since.⁹ All the repeated experiments to obtain sesquihydrate (7), isopropyl alcohol solvate (3) and the previously published ethanol solvate, led to the formation of anhydrous TCl form I.

Hydrate appearance

Interestingly, no hydrate form was prepared during any preceding crystallization experiments. The hydrate appearance during our systematic screening was therefore unexpected. This pseudopolymorph, appearing more than 50 years after the discovery of the parent compound, could be reproduced easily in the slurry and presented opportunities for a new drug formulation.

To characterize it, we performed the dynamic vapour sorption experiment of both TCl form I and of TCl dihydrate (8). While the anhydrous TCl was not hygroscopic, diHyd (8) absorbed more than 13 wt% at 90% RH and its crystals visibly dissolved. The absorption curve of TCl form I showed no indication of the sesqui- or di-hydrate formation.

Solubility of TCl

We questioned the possible relation between TCl solubility in the solvents and its propensity to form solvates with them.

To observe any relation between the solubility of TCl in the solvent and the solvate formation, the solubilities of TCl in the solvents, with which it forms solvates, were determined by gravimetry or UV-vis spectrophotometry. From the solubilities estimated during the screening, the formation of some solvates seemed improbable.

We were unable to determine the precise solubility of TCl in iPrOH, as no wavelength maxima were observed in the range of 200–700 nm. The measurement by gravimetry was not suitable, as solubility was lower than 1 mg per 100 mL of iPrOH from a freshly opened bottle. The occurrence of an iPrOH solvate (3) seemed therefore to be an unexpected anomaly.

The solubility of TCl in MeOH and MeNO₂ was also not determined by UV-vis spectrophotometry due to the lack of suitable wavelength maxima, but gravimetry was successfully applied. The results are shown in Table 2.

Structural comparison

The crystal structures containing TCl were compared in several ways. For the newly solved TCl, the conformations of the trospium molecules as well as the positions of the chloride anions and solvent molecules were compared. The comparison of basic crystallographic parameters, molecular packing and crystal habitus were done for the novel structures together with the crystal structures of TCl described previously.

In Fig. 2, the 8 newly solved TCl solvates are compared with regard to the trospium conformation as well as the positions of the chloride anion and solvent molecule. In all those aspects, seven structures are very similar to each other (Fig. 2a). The only conformationally different structure is TCl diHyd (8) (Fig. 2b). The torsion of the ester chain is different and results in a significant change in the position of the phenyl rings.

Also, TCl dihydrate (8) is the only structure where the (trospium OH)⋯(Cl⁻) H-bond is not present. For protic solvates, where the solvent has an H-bond donor (*e.g.*, alcohols), the H-bonding extends to the solvent molecule: (trospium OH)⋯(Cl⁻)⋯(solvent OH). In diHyd (8), the order is reversed and the trospium hydroxyl is H-bonded to one of the water molecules, which then forms an H-bond to the chloride anion.

All solvates have the ratio of 1:1, except for the hydrates (2:3 and 1:2).



Table 2 List of solvents used in the systematic screening. Solvents with which TCl forms solvates are presented at the top and are separated from the other solvents

Solvent	Abbreviation	Class (ICH)	Polar P/A	Approx. solubility (mg mL ⁻¹)	Precise solubility (mg mL ⁻¹)	Applied screening method
<i>N,N</i> -Dimethylformamide	DMF	2	Aprotic	<4	1.85 ^a	AS, Sl
Isopropyl alcohol	iPrOH	3	Protic	<4	<0.01	AS, Sl
Methanol	MeOH	2	Protic	>100	159.7 ^b	SE, Sl ^c
Nitromethane	MeNO ₂	2	Protic	<4	0.82 ^b	AS, Sl
Propionitrile	PRN	—	Aprotic	<4	0.17 ^a	AS, Sl
Water	WA (Hyd)	—	Protic	>500	606 ^a	SE, Sl ^c
Acetone		3	Aprotic	<4	—	AS, Sl
1-Butanol	<i>n</i> BuOH	3	Protic	19	—	SE
Cyclohexanol		—	Protic	6	—	SE
Ethanol	EtOH	3	Protic	>30	—	SE
Ethyl acetate	EtAc	3	Aprotic	<4	—	AS, Sl
Isobutyl alcohol	iBuOH	3	Protic	5	—	SE
Methyl acetate	MeAc	3	Aprotic	<4	—	AS, Sl
Methylethyl ketone	MEK	3	Aprotic	<4	—	AS, Sl
2-Methoxyethanol		2	Protic	30	—	SE
1-Octanol		—	Protic	<4	—	AS, Sl
1-Pentanol		3	Protic	10	—	SE
1-Propanol	<i>n</i> PrOH	3	Protic	20	—	SE
Tetrahydrofuran	THF	2	Aprotic	<4	—	AS, Sl

^a Index indicates solubility determination *via* UV-vis spectrophotometry. ^b Index indicates solubility determination *via* gravimetry. AS: anti-solvent addition, Sl: slurry, SE: slow evaporation. ^c The slurries of TCl in MeOH and WA were prepared even though slow evaporation as the main screening technique was applied. The slurry in MeOH was prepared for gravimetry, and the slurry in WA was prepared to be used during anti-solvent experiments. For the precise UV-vis solubility measurements, the standard deviation was calculated from 3 repeated measurements in WA, ACN and DMF; the %RSD ($q_{0.95}$) was below 5%. It was assumed that the RSD for PRN was comparable, since the identical measurement methodology was followed.

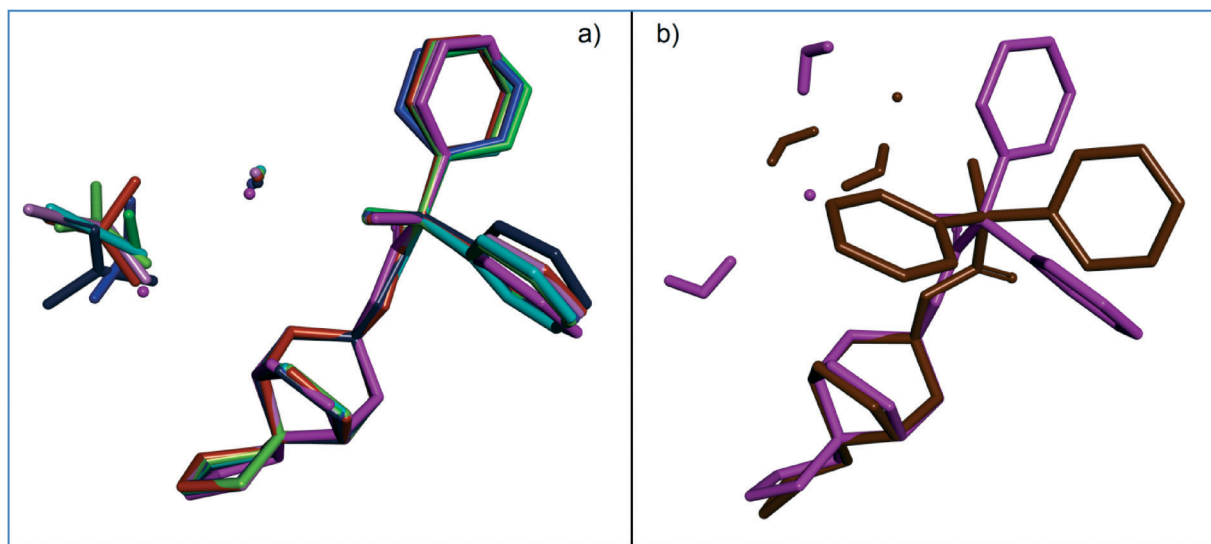


Fig. 2 The conformation of trospium and the position of the chloride anion and the solvent molecule in the structures of TCl solvates. a) Similar structures of TCl ACN (1), DMF (2), iPrOH (3), MeOH (4), MeNO₂ (5) and PRN (6) solvates and sesquihydrate (7); b) a very different conformation of trospium in TCl dihydrate (8) (brown) in contrast with TCl sesquihydrate (7) (magenta, two molecules of water are displayed instead of one and a half according to the stoichiometry).

The following structural comparisons were done with 8 new solvates together with all of the previously published crystal structures of TCl. Those include two polymorphs of pure TCl (form I, TCl-I and form II, TCl-II), the EtOH solvate and five cocrystals with salicylic acid (Sal), adipic acid (Adip),

glutaric acid (Glut), oxalic acid (Oxal) and urea. That makes 16 structures containing TCl in total.

The next comparison is a tree diagram (Fig. 3) based on the results from the Crystal Packing Similarity tool in the Materials module of Mercury. The diagram is used for a clear



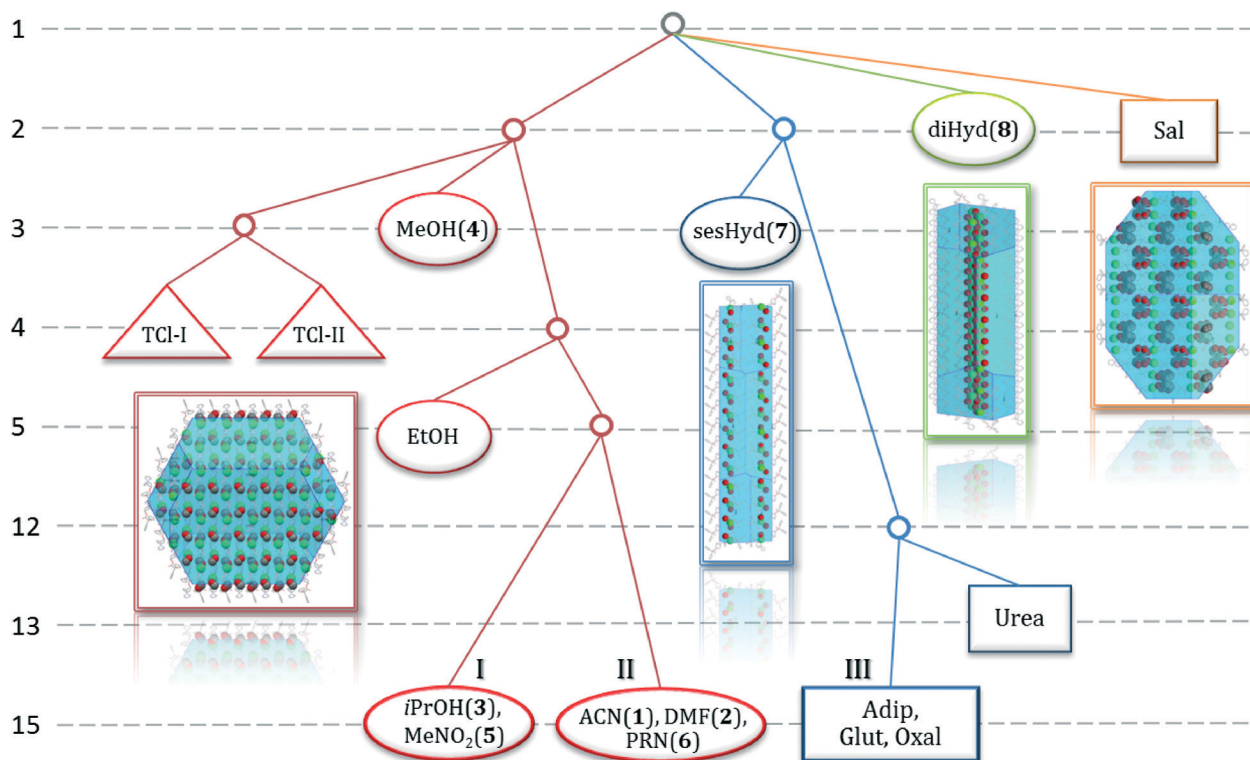


Fig. 3 A tree diagram based on the results from the Crystal Packing Similarity tool in the Materials module of Mercury.¹⁹ All 16 known crystal structures containing TCl are compared. Pure polymorphs are marked as triangles, cocrystals as rectangles and solvates as ovals. Structures grouped on the bottom row (I, II and III) indicate that 15 out of 15 trospium molecules in a cluster are similar and these are considered to be isostructural. The calculated crystal shape and packing for each of the four branches is indicated. For the very left and very right branches, the crystal shape is a prism (TCl MeOH solvate (4), TCl salicylic acid cocrystal). For the two middle branches, it is a plate (side view; TCl sesquihydrate (7), TCl dihydrate (8))

analysis of similarity in the crystal packing of structures of one compound. It can compare the crystal packing of only the largest molecule in the structure (in our case the trospium cation). Therefore, it can be easily used to analyse a family of polymorphs, hydrates, solvates, salts and cocrystals.

A tree diagram is constructed in several steps. For detailed instructions for the diagram construction, see ref. 22. Every row represents a separate calculation. For every level, a cluster of molecules is calculated for all structures. The size of the cluster (number of molecules) corresponds to the row number, *i.e.*, 1 to 15. Then, all clusters are overlaid over each other in pairs and it is considered whether all molecules fit within a pre-set error. We used a 20% margin for both position and the torsion angles. The larger the calculated cluster, the more the structures must be similar for all the molecules to fit. This results in branching and the formations of a tree diagram. Structures grouped (*e.g.*, *i*PrOH (3), MeNO₂ (5)) on the bottom row indicate that 15 out of 15 trospium molecules in a cluster are similar and these are considered to be isostructural (Fig. 7 in ESI[†]). We can see three such groups (I–III) in Fig. 3. In the tree diagram, the solid forms of TCl are divided into four main structural groups (red, blue, green and orange). It is very interesting that all solvates (except hydrates) belong to the same structural group (red). It also includes both pure polymorphs of TCl (TCl-I and TCl-II).

Both hydrates are structurally different from the other solvates. TCl sesquihydrate (*sesHyd*) is loosely similar in packing to cocrystals of TCl with adipic acid (*Adip*), glutaric acid (*Glut*), oxalic acid (*Oxal*) and urea, which are in the group marked in blue. TCl dihydrate (8), as well as the cocrystal of TCl with salicylic acid (*Sal*), are not similar to any other structure; they separate from the rest of the structures as early as on level two. Surprisingly, both forms (*diHyd* (8) and *Sal* cocrystal) are likely to be very convenient arrangements for TCl. The cocrystal with salicylic acid forms most readily from all the described cocrystals of TCl, even with simple mixing and heating of both components.¹²

For all of the four main branches in the tree diagram, theoretical crystal shape and packing were calculated in Mercury (Bravais, Friedel, Donnay and Harker (BFDH) theoretical crystal morphologies, Fig. 3 and 4).

This visual representation correlates nicely with the results from the tree diagram; the crystal packing is clearly very different in each of the forms. It varies from the regular alternation of trospium and solvent molecules and the prism crystal shape in TCl MeOH (4) (red) to two different layered structures forming plates (TCl sesquihydrate (7) (blue) and TCl dihydrate (8) (green)) and to the column structure in the TCl salicylic acid cocrystal (orange). The calculated crystal shapes were found to be in good correlation with the experimentally observed ones (Fig. 4).



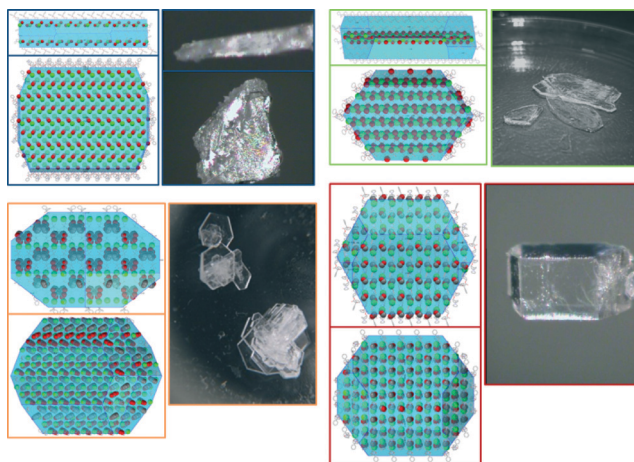


Fig. 4 The observed crystal shape and the calculated crystal shape and packing viewed from two directions for TCl sesquihydrate (7) (blue frame), TCl dihydrate (8) (green frame), TCl salicylic acid cocrystal (orange frame) and TCl MeOH solvate (4) (red frame). Trospium molecules are displayed in wire mode, with chloride anions and solvent/coformer molecules as balls (Cl – green, C – grey, O – red).

The structures of solvates can be divided into three branches (red, blue and green branches in a tree diagram, Fig. 3) depending on structural similarity as was revealed by the crystal packing similarity calculations (the fourth branch (orange) contains only a cocrystal structure).

Both the blue and green branches contain only one solvate structure (blue: 7 and green: 8). The remaining red branch is populated by six new solvate structures (1–6) and is divided into three parts: (4), (3 and 5; *i.e.* isostructural group I) and (1, 2 and 6; *i.e.* isostructural group II).

The red branch is the most complex. Its groups are fairly similar, yet the differences can be observed in the crystal packing as well as the patterns of intermolecular non-covalent interactions. The common feature of this group is the formation of a dominant chain of hydrogen bonds, formed by hydrogen bond O3–H1o3...Cl1 and the weak hydrogen bonds resulting in the formation of columns. These columns are further connected *via* weak hydrogen bonds and $\pi\cdots\pi$ stacking.

TCl MeOH (4) and group II have a surprising amount of similarities. The dominant chain is in the same direction [1 0 1] and it is formed by alternating mirror images of trospium cations and chloride anions. The main difference among these groups is the presence of hydrogen bond O1s–H1o1s...Cl1 in TCl MeOH (4), resulting in a reversed orientation of the solvent and denying the participation of the solvent in C(ar)–H...X (X=N or O) hydrogen bonds observed in group II. The orientation of the solvent did not affect the packing similarity calculation used for the creation of the tree diagram and the arrangement of the trospium cations is really similar in both of these groups, yet the fragment of the three molecules of TCl MeOH (4) fails the similarity check and is excluded from the group of aprotic solvents extremely early. This can be ascribed to a significantly smaller unit cell of TCl MeOH (4).

Group I is apparently different from the other members of the red branch. The dominant chain of hydrogen bonds is in the direction of [1 0 0] and is formed by translation of the same molecule of TCl, resulting in an almost identical layer 1 with group II and TCl MeOH (4) and an inverted layer 2 (for figures on layer arrangement see ESI,† Fig. 10). This group of solvates is excluded from the rest when the calculated fragment reaches 5 molecules. This follows the layer similarities, since four molecules were selected from the same layer, but the fifth one is selected from the second layer populated by mirror images. The solvent is bound by a hydrogen bond to a chloride anion in this case (O1s–H1o1s...Cl1 or C1s–H3c1s...Cl1) and plays no other important role in the crystal packing.

The differences in the structures can be easily demonstrated in the figures of crystal packing, where the mirror images of stacking results in a zigzag line formed by phenyl rings, absent in group I (see Fig. 5).

The blue branch is populated by TCl sesquihydrate (7) as well as the known cocrystals of TCl. But the similarities are not very pronounced between the sesquihydrate (7) and the rest of the group. Throughout the group, the main interactions are hydrogen bonds. The largest similar fragment is

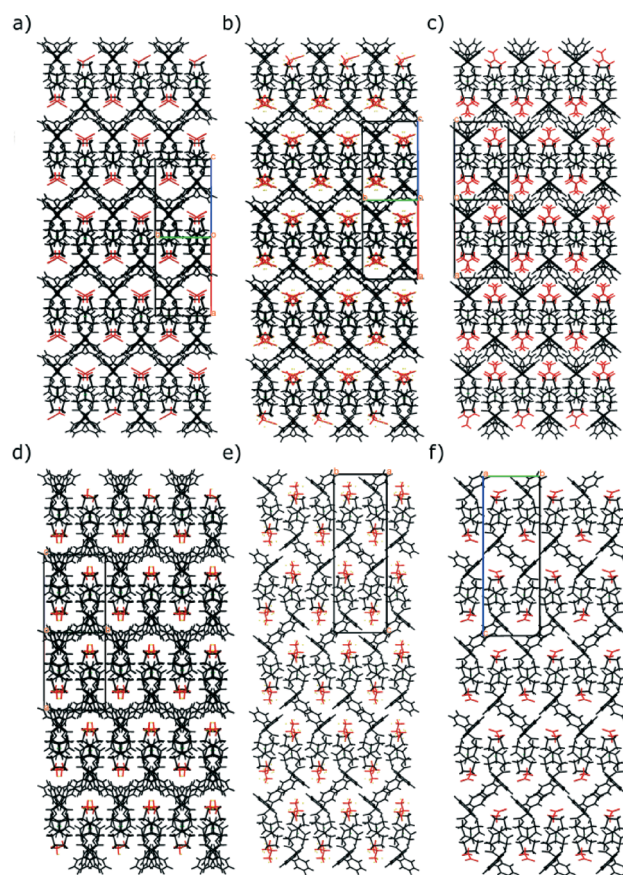


Fig. 5 Crystal packing of TCl solvates viewed in the direction of the main chain of hydrogen bonds. The trospium cation is depicted in black, and the solvent molecule in red. a) TCl ACN (1), b) TCl PRN (6), c) TCl DMF (2), d) TCl MeOH (4), e) TCl iPrOH (3), f) TCl MeNO₂ (5).



formed by two molecules and in this case there is no other distinctive similarity of the cocrystals and the sesquihydrate (7). The hydrogen bond patterns differ significantly and the only similarity of the structure packing is a tendency to form a bilayer with the polar part of the trospium cation chloride anion and solvent (or cofomer) on the inside and aromatic part of the trospium cation on the outside (see Fig. 6). The TCl cocrystals keep the similarity to the threshold of 12 molecules, and then the structure of the urea cocrystal fails the similarity check. This is caused by the positioning of the formed bilayers (see Fig. 7).

In the case of TCl dihydrate (8), the usual hydrogen bond $O3-H1o3 \cdots Cl1$ is no longer observed, as mentioned previously. The chain of the alternating molecules of water and chloride anions is observed instead and the trospium cation is attached to it *via* hydrogen bond $O3-H1o3 \cdots O2w$, creating a firm pillar in the structure. The pillars are further connected with the weaker hydrogen bonds into a bilayer (see Fig. 8), which is further connected into a three-dimensional structure *via* $C(ar)-H \cdots \pi$ and $\pi \cdots \pi$ interactions.

Thermal behaviour of reproducible solvates

The solvate loses firstly its solvent molecules during desolvation, which can drive the parent compound molecules to arrange in a different order, providing a new polymorph. We decided to expel the molecules of the solvent from the solvate structure by heating, as anhydrous TCl is thermally stable. For thermal experiments, we were able to prepare larger sample quantities of sufficient purity only for six of the solvates (see Table 3; the line with water relates to TCl dihydrate).

The temperature of solvent evaporation from the structure was determined from TGA/DSC measurements and is shown in Table 3. The calculation of weight loss during heating confirmed the 1:1 stoichiometry of each solvate and the 1:2 stoichiometry of the dihydrate. Upon desolvation at an increased temperature, the resulting powder diffraction

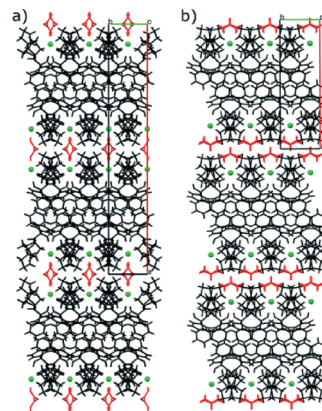


Fig. 7 Bilayer positioning in a) TCl Oxal and b) TCl Urea.

patterns corresponded to pure TCl in its most common polymorph (form I). This was not surprising, as the solvates were shown to be structurally similar to the polymorph form I. However, for the dihydrate (8), whose packing is fundamentally different, there was hope that its desolvation could lead to a new polymorph. Unfortunately, this was not the case.

The DSC experiment of TCl MeOH (4) was erratic at a heating rate of $10 \text{ }^\circ\text{C min}^{-1}$. We found that gentle grinding improved the curve; however, reproducible data were collected only at $1 \text{ }^\circ\text{C min}^{-1}$ (see Fig. 9 and 10).

While it is a common phenomenon for the solvent to leave the structure of its solvate above the boiling point of the present solvent, for two of the presented structures we observed an opposite behaviour. In the DMF solvate (2) and TCl dihydrate (8), the solvents left the structures before their respective boiling points. While we see a reason for this for diHyd (8), we can only speculate about the quality and frequency of the weak interactions for the DMF solvate (2).

In diHyd (8), two molecules of water are bound with weak hydrogen bonds ($d(D \cdots A) > 3.1 \text{ \AA}$) to Cl^- . $O1w$ is connected exclusively to Cl^- , while $O2w$ forms a bridge to trospium $O3$ ($d(D \cdots A) = 2.668(2) \text{ \AA}$, see Table 4. A list of hydrogen bonds is in the ESI†). Upon heating, we expect the $O1w$ molecule to leave rapidly, thus destabilizing the structure packing and

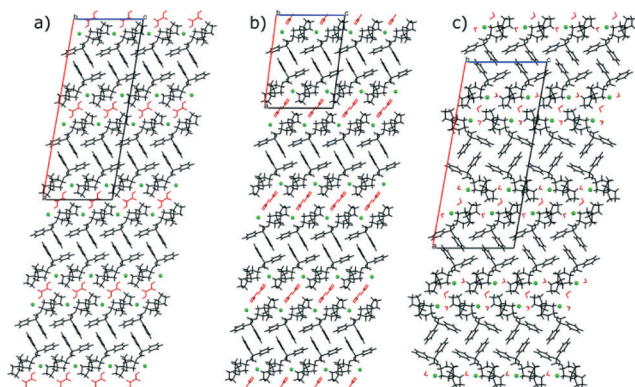


Fig. 6 Crystal packing of selected members of the blue branch (see Fig. 3) viewed in direction $[0\ 1\ 0]$ with the trospium cation depicted in black, the cofomer/solvent in red and the chloride anion in green. a) TCl Oxal, b) TCl Urea, c) TCl sesquihydrate.

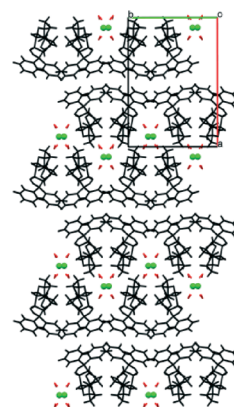


Fig. 8 Formation of the bilayer in TCl diHyd (8).



Table 3 Desolvation behaviour of reproducible solvates

Solvent	Boiling point of the solvent (°C)	Desolvation temperature (°C)	Stoichiometry from TGA/DSC	PXRD of resulting powder
ACN	82	140	1 : 0.89	TCl form I
DMF	152	115	1 : 0.90	TCl form I
MeNO ₂	100	152	1 : 0.69 ^a	TCl form I
MeOH	65	98	1 : 0.94	TCl form I
PRN	96	125	1 : 0.76 ^a	TCl form I
WA	100	93	1 : 1.88	TCl form I

^a Traces of TCl were detected in the samples analysed by TGA/DSC, therefore, the stoichiometry does not match. For desolvation experiments pure solvates were prepared and weight loss corresponded to the calculated ratio of the solvent.

forcing O2w to leave as a consequence. As O2w is not prone to leave the structure so easily, we observe subtle curves in TGA during the dehydration phase in the range of 80–100 °C. The structure rapidly recrystallizes as TCl form I, which is indicated by the weak endotherm at 110 °C in DSC.

The notion of the lower stability of the DMF solvate (2) was also drawn from the observations during the screening experiments, where the DMF solvate (2) was the only solvate that was not obtained from the anti-solvent experiments. The other solvates were successful in the stability competition with TCl form I during the anti-solvent experiment. Hence, the DMF solvate (2) appears to be the least stable form.

Conclusions

We performed systematic solvate screening for a pharmaceutical compound, trospium chloride. Eight solvates were discovered and their crystal structures solved.

We identified three distinct groups of TCl solvates varying in the direction of hydrogen bond chains and solvent orientation.

The packing similarity calculation proved to be a strong tool for the evaluation of a large amount of crystal structures. However, the results can be misleading in the case of large discrepancies in the unit cell volume.

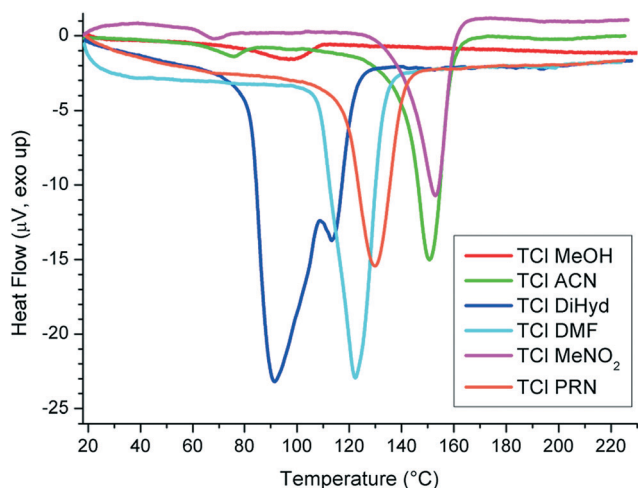
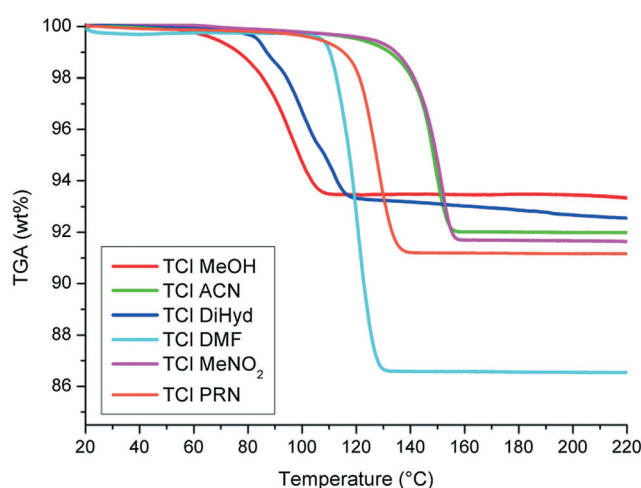
We tried to find a relation between the solubility and the solvate formation, but to no avail. TCl provides solvates both with polar protic and polar aprotic solvents; its solubility in the solvents varies from as low as below 1 mg per mL to as high as above 600 mg per mL.

The serendipitous preparation of 2 out of the 8 novel forms demonstrates that no solid state screening is ever perfect and is able to find all possible crystal forms. The conventional screening revealed forms that were easily reproduced, but did not provide the EtOH solvate previously described in literature nor did it provide the iPrOH solvate and sesquihydrate presented in this work that were prepared outside the systematic screening experiments.

On the other hand, the screening revealed the possibility of TCl hydrate formation, which had been, until now, unknown. Especially for the pharmaceutical hydrates, it is crucial to be aware of their existence, since they can often pose a threat to robust and stable drug formulation.

Acknowledgements

This work was supported by the Grant Agency of Czech Republic, grant no. 106/14/03636S and grant no. 14-03276s, and received financial support from specific university research (MSMT no. 20/2015).

**Fig. 9** DSC curves of the solvates.**Fig. 10** TGA curves of the solvates.

We would like to thank the Interpharma Praha a.s. company for the provided trospium chloride; we have found many peculiarities within this intriguing molecule. We would also like to thank Miša Chrastná, Honza Rohlíček and Ondra Dammer; their contribution to 'the trospium case' is much appreciated.

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