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Simulating preferential sorption of tartrate on prismatic calcite surfaces†

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Understanding the influence of additives on crystal growth is required to engineer the crystal properties according to their functional applications. In this work, the sorption behavior of tartrate on calcite surfaces is investigated employing molecular dynamics simulations to understand additive-mediated crystal growth. The free energy landscapes for the sorption of tartrate are calculated using metadynamics. The adsorption binding energies of favorable conformations, orientations and positions of tartrate near the (104) and (1–10) calcite surfaces are determined. The obtained results provide a molecular-level explanation of the experimentally observed tartrate-stabilized exposure of prismatic {1–10} faces during calcite growth. The simulations show that tartrate preferentially adsorb directly to the (1–10) calcite surface, whereas tartrate is more loosely adsorbed on the (104) surface, mainly by solvent-mediated binding. The (1–10) geometry of calcite surface sites closely matches the structure of tartrate, with a specific role of carboxylate and hydroxyl groups in recognizing the calcium and carbonate ions, respectively. Two stable adsorption configurations are identified for the (1–10) face: (1) adsorbed tartrate with the effect of surface-induced conformational change and (2) incorporated tartrate into the surface by fitting one of the carboxylate groups into lattice position normally occupied by carbonate ions and additionally stabilized by binding of both hydroxyl groups to neighboring carbonate ions. The results indicate that surface energetics, structural matching and adsorbed water layer play a major role in the strength of the interactions and hence in the expression of calcite morphology. Preferential adsorption of tartrate on {1–10} surfaces could stabilize these otherwise fast-growing faces and thus inhibit crystal growth in {1–10} directions.

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1. Introduction

The processing and application of crystalline products strongly depends on their morphology and size distribution. The control and design of a desired crystal morphology and size can be achieved by employing specific additives during crystal growth processes.^{1,2} In many cases additives are not desired in the crystallization process since they reduce the purity of the crystalline product. It is important to understand how the additives are incorporated into the crystal to control their uptake. A detailed understanding of the influence of additives on crystal growth is required to control the crystallization process and to engineer the crystal properties according to their functional applications. However, the mechanisms of interactions between additives and crystal surfaces are not yet fully understood because of the complexity

of the processes occurring at the crystal surface–aqueous solution interface. Once the crystal/solution interface is understood on the atomic resolution, it may be possible to design molecules that have specific interactions with one or more of the crystal faces. Challenges are in the complexity of the process that comprises growth, dissolution and sorption occurring at the solid/liquid interfaces. Given the increasing computational power, molecular dynamics (MD) is emerging as a viable approach toward building a comprehensive picture of molecular phenomena involved in additive-controlled crystallization processes.^{3–5}

In recent years, calcium carbonate has been intensively studied to understand how polymorphism, structural features and morphology can be controlled by organic additives.⁶ Calcium carbonate forms different polymorphs, including calcite, aragonite and vaterite. Among them calcite is the thermodynamically most stable modification under standard conditions and it appears in various morphologies, typically as rhombohedral crystals bounded by the most stable {104} calcite faces. Precipitated calcium carbonate is an important synthetic mineral widely used in various technologies, mainly as multifunctional filler or pigment (paper, plastics, pharmaceuticals, foods).^{7–9} The organic-mediated crystal growth

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control of calcium carbonate is also important in biomineralization, where nature engineers materials with substantially enhanced mechanical properties.⁶ Namely, recent studies showed that calcite has high affinity to incorporate organic molecules into the structure, thereby affecting its composition.^{10–12} It is believed that the adsorption and incorporation of organic molecules is controlled by the additive recognition for a specific surface structure, and consequently organic additives are selectively incorporated onto specific crystal faces. Some experimental studies showed that inside biogenic calcite crystals, acidic proteins are preferentially incorporated on the set of crystal planes parallel to its crystallographic *c*-axis.^{13,14} Ionic interactions are important in the crystallization of mineral compounds. The so-called *multifunctional additives*, *i.e.* organic molecules with various functional groups, are capable of forming several bonds with (ionic) surface sites. Calcite exhibits exceptional chemical affinity for organic molecules having polar functional groups, such as carboxylate and hydroxyl groups.⁶ One of the hypotheses^{15,16} is that the binding of the acidic macromolecules to calcite surfaces could occur by fitting carboxylate groups into lattice positions occupied by carbonate groups. In order to understand how particular functional groups bind to the surface, breaking down the interactions of these large molecules into more basic units, has shown to be a good strategy.^{5,17,18} Experimental studies showed that small organic molecules, such as carboxylates or amino acids, can be used as effective regulators of calcite crystal morphology. In particular, dicarboxylates substituted with the amino or hydroxy group on α -carbon has been identified to show high affinity and selectivity for calcite surfaces.¹⁷ Previous experimental results indicated specific interactions of tartrate with newly stabilized calcite surfaces, thus modulating the morphology.¹⁹ The co-addition of tartrate during steady-state calcite growth system led to a clear selective expression of the prismatic $\{1-10\}$ side faces, parallel to the *c*-axis (Fig. 1). This suggests the importance of surface energetics and structural matching in tartrate interactions with calcite surfaces. The prismatic $\{1-10\}$ surface is of higher energy, than the most stable rhombohedral $\{104\}$ surface, and if these surfaces are to be expected in the final morphology, they must be significantly stabilized and prevented

from growing out. It was suggested that charged carboxylate groups recognize the calcium ions on $\{1-10\}$ with specific spacing that matches the structure of tartrate. Hydroxyl groups might also have a specific role binding to surface carbonate groups, thus enhancing the adsorption.²⁰ Still, detailed thermodynamic and structural information controlling the growth mechanism, such as the binding energies and sorption modes of organic additives at specific faces and sites are not understood.

MD simulations can provide useful information at molecular level and have already substantially contributed to our understanding of interactions between organics and mineral surfaces. Earlier simulation studies mainly applied simplified approaches, not accounting explicitly for water molecules.^{16,21,22} However, the effect of water cannot be ignored in studies of additive-controlled crystallization from solution, and recent simulation studies emphasized the importance of structured interfacial water layers on additive adsorption behavior.^{5,23–25} One of the challenges of classical MD simulations remains in the development of reliable force fields which govern the behavior of solid/liquid interface of a particular system. In the last years there has been considerable effort directed toward the development of force fields for calcium carbonate, capable of describing the bulk crystal as well as interface of this material with water and organics (with the aim to address biomineralization), and to model the nucleation and growth of calcium carbonate.²⁶ MD simulation studies have mainly focused on the most stable rhombohedral $\{104\}$ calcite surface/liquid interface, obtaining valuable information regarding the structure and dynamics of adsorbed water layers,^{26,27} but seldom included the presence of organics.^{5,24} The latest developed calcium carbonate force field²⁸ uses Lennard–Jones potentials which makes the combination with an organic force fields straightforward by applying common combination-rule approach.

Another challenge of the methods based on MD is adequate sampling of the conformational space. Most recent simulations of adsorption indicated that regular MD simulations are often not efficient at overcoming the problem of multiple free energy minima separated by significant barriers which could trap a molecule in one of the local minima. Especially this was shown to be valid for the case of

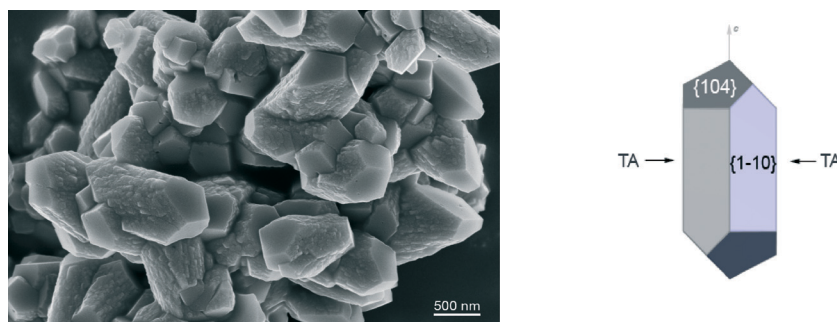


Fig. 1 SEM image of prismatic calcite crystals grown in the presence of tartrate, TA (left) and crystal shape model (right) combined with the two forms, prismatic $\{1-10\}$ side faces and basal $\{104\}$ faces. Tartrate stabilizes new $\{1-10\}$ calcite faces and modulates the crystal morphology.



highly-charged mineral surfaces, such as phosphates²³ and carbonates.⁵

In this work, the adsorption behavior of tartrate on calcite surfaces is studied employing MD simulations to understand additive-mediated crystal growth. Tartrate ($\text{O}_2\text{C}-\text{CH}(\text{OH})-\text{CH}(\text{OH})-\text{CO}_2$)²⁻ served as a simple model compound of hydroxycarboxylates, giving valuable information about basic molecular interactions of these polar functional groups with different calcite surfaces, needed for rational development of novel growth-controlling additives. First, we investigated the energetics and structure of interfacial calcite–water interface for (104) and (1–10) calcite surfaces. Then, free energy landscapes for the adsorption and desorption of tartrate are calculated using metadynamics,²⁹ thus avoiding the pitfalls of MD. Using metadynamics, the adsorption energetics of favorable conformations, orientations and positions of the tartrate molecule near the (104) and (1–10) calcite surfaces are determined. We show the importance of surface energetics, structural matching and preorganization of functional groups as well as structuring of the interfacial water on the adsorption strength, and hence calcite morphology.

2. Simulation methods

2.1. Simulation systems setup

MD simulations were performed with the GROMACS³⁰ software package (version 4.6.3). For the force field parameters of calcite and the calcite/water–organics interactions, CaCO_3 –TIP3P–solution interfacial model was used, as recently developed by Xiao *et al.*²⁸ Calcite–tartrate pairwise interaction parameters were deduced from atomwise CaCO_3 parameters and atomwise tartrate parameters by applying the geometric mixing rule. The general AMBER force field (GAFF)³¹ was used to model tartrate. Consistent charges indicated compatibility of GAFF and the chosen calcium carbonate force field.²⁸

Tartrate was modeled as a dianion (deprotonated state of both carboxylic groups), since this is the predominant species at experimental conditions¹⁹ of calcite crystallization under slightly basic conditions (pH 8). The dihedral parameters were re-parameterized by quantum chemical calculations using the *Schrödinger's Jaguar*³² software package. Density functional theory (DFT) with B3LYP functional and 6-31G(d,p) basis set was chosen to stay consistent with the GAFF parameterization procedure. The partial atomic charges were derived by the R.E.D Server,³³ which implements standard GAFF approach, using Restrained Electrostatic Potential (RESP)³⁴ method taking into account multiple conformations.

Initial structures of the (104) and (1–10) faces were generated using GDIS.³⁵ The resulting hexagonal calcite unit cells were periodically replicated in 3D space, thus generating a crystal slab of approximately $5 \times 3 \times 2.5 \text{ nm}^3$. Extending the simulation box in the direction perpendicular to the surface gave a gap of about 5 nm into which a tartrate molecule and about 2500 water molecules were inserted. To keep the entire system charge-neutral, Na^+ counter ions were added in the presence of negatively charged tartrate. Two Na^+ ions were

initially placed in bulk solution. By doing so, they remained mostly in the solution. Na^+ ions might interfere between tartrate and calcite crystal surface, but this was not observed in our performed simulations. Na^+ ions were not in close contact with tartrate. This was checked by the convergence test (see the *Metadynamics* section) and by tracking and inspecting Na^+ and tartrate trajectories. Periodic boundary conditions were applied in all directions. Ideal calcite surfaces were assumed: calcium and carbonate ions were position restrained by a harmonic potential with a force constant of 1000 kJ mol^{-1} . Previous simulation studies reported no significant relaxation of surface ions on (104) calcite–water interface,^{26,27} and only a slight relaxation of (1–10) surface.³⁶ In pure systems (without additives), the (1–10) surface is in contact with water molecules and the top surface ions relax, in order to minimize excess surface free energy. However, in our simulations unrelaxed (1–10) surfaces are assumed, since they are supposed to be stabilized with adsorbed tartrate molecules (adsorbed tartrate layer). All bond lengths of carbonate, water and tartrate were constrained to their equilibrium positions with the LINCS algorithm. Non-bonded interactions were evaluated with a cutoff of 1.0 nm. The Particle Mesh Ewald method (PME) was used to treat long-range electrostatic interactions, with a grid spacing of 0.16 nm and PME order of 4. Simulations were performed using a 2.0 fs time step in the isothermal–isobaric (NPT) ensemble at 300 K and 1 bar. Nose–Hoover temperature coupling was used with a relaxation time of 2.0 ps. Constant pressure was obtained by coupling to a semi-isotropic Parrinello–Rahman barostat, allowing only scaling of the simulation box dimension normal to the surface, with a relaxation time of 2.0 ps. The system was equilibrated by performing 2.0 ns of NPT simulation, before proceeding with production runs or metadynamics. In order to visualize the simulation and to generate molecular snapshots, *Visual Molecular Dynamics* (VMD)³⁷ was used.

2.2. Metadynamics

Metadynamics applies a time-dependent bias to a predefined set of collective variables (CV) that describe characteristics of a system. A detailed review of metadynamics can be found elsewhere.^{29,38,39} In short, the algorithm biases a MD simulation through the periodic addition of a small repulsive potential, defined as a sum of Gaussian distributions, to the overall potential energy. By introducing the bias potential it enables escaping the system's free energy minima and simultaneously quantitatively evaluates the free energy landscape. To properly explore and describe the sorption behavior of tartrate on different calcite faces, we selected three CVs: (1) distance from the crystal surface defined as the center of mass of tartrate in the direction orthogonal to the crystal surface, (2) the tartrate orientation represented by the cosine of the angle ($\cos(\theta)$) between the tartrate vector, defined by the two carbon atoms to which the polar functional groups are attached, and the axis perpendicular to the surface (3) tartrate



molecular conformation, defined as the dihedral angle of the carbon chain. The distance from the surface is defined relative to the average position of the Ca surface sites in the outermost layer. The selected choice of CVs allowed resolving bonded/unbonded states in the CV-space, differently oriented bound configurations and preferred conformations, *i.e.* the conformational change of tartrate, upon sorption onto the surface. The biasing Gaussian potential functions with a height parameter of 0.05 kJ mol^{-1} were deposited every 0.5 ps, with a width of 0.01 nm, 0.025 and 5° along CV1, CV2 and CV3, respectively. To ensure the convergence of the system, but not to overfill the free energy landscape in CV-space (oversampling the bulk solution), the simulation run time was optimized to 160 ns. The computational time needed for the simulation on 64 cores was approx. 48 hours. The binding free energy was estimated by reconstructing the free energy landscape using *sum_hills* utility provided with the *PLUMED*³⁹ plugin. The free energy difference between the adsorbed state and the desorbed state obtained from the free energy profile is defined as the binding free energy. As a reference level for the binding free energy the most stable *trans* conformation of the desorbed tartrate in the bulk water was taken. The convergence of simulation was assessed by comparing the reconstructed free energy profiles in the CV space as a function of simulation time. The system was considered to be sampled properly (converged) when the difference between binding energies over time approached zero.

3. Results and discussion

3.1. Calcite–water interfaces

Before proceeding to tartrate adsorption behavior on calcite surfaces, we first examine the energetics and structure of calcite interfaces, particularly focusing on the comparison of the (1–10) surface with the most stable (104) surface.

Surface free energies, γ , were determined according to the procedure described by de Leeuw *et al.*⁴⁰ using the relation, $\gamma = \Delta E/(A \cdot N_a)$, where E is the total potential energy, A is the surface area, N_a the Avogadro number; the energy difference was calculated as $\Delta E = E_{\text{interface}} - E_{\text{bulk}} - E_{\text{water}}$. The surface energies of dehydrated surfaces are calculated to be $\gamma_{(104)} = 0.5 \text{ J m}^{-2}$ and $\gamma_{(1-10)} = 0.8 \text{ J m}^{-2}$. The (104) surface is more stable and less reactive than the (1–10) surface. Both surfaces are stabilized when water is introduced, with calculated surface energies $\gamma_{(104)} = 0.1 \text{ J m}^{-2}$ and $\gamma_{(1-10)} = 0.4 \text{ J m}^{-2}$. The fact that the (104) surface is more stable than the (1–10) surface means that the more reactive (1–10) surface should grow more quickly and may well grow out of the thermodynamic crystal morphology altogether, leaving large {104} faces expressed in the crystal morphology. Indeed, the rhombohedral {104} crystal form is the most often experimentally observed in additive-free crystallization systems.^{17,41} Our calculated surface energies are in good agreement with other theoretical studies^{36,40} who have demonstrated that this surface is significantly lower in energy ($\Delta\gamma = 0.3 \text{ J m}^{-2}$) than the (1–10) surface. Thus, if the {1–10} faces are to be expected in

the final morphology, it must be significantly stabilized and prevented from growing out.

The higher energy of the (1–10) surface, *i.e.* higher instability and therefore higher reactivity, can be related in terms of surface structure and electrostatic interactions. Each surface ion, calcium and carbonate, on the (104) face has one broken bond that resulted with the unsaturated coordination and partial charge of $z = +1/3$ and $-1/3$ felt in the vacant site above the Ca surface sites and CO_3 groups, respectively. On the other hand, the prismatic (1–10) calcite face exhibit three dangling bonds above each surface ion with partial charge of $z = \pm 1$. Consequently, the adsorption of charged organic molecules, such as tartrate with polar functional groups, may neutralize partial surface charge, wherein the interactions are stronger with the prismatic faces that are of higher energy due to its three dangling bonds. On contrary, the unsaturated coordination of lattice ions of the rhombohedral {104} calcite surface are mainly stabilized by adsorption of water molecules.

On the most stable calcite (104) surface, water forms several ordered high-density layers, maxima at 0.22, 0.34 and 0.49 nm (Fig. 2(a)), which is consistent with previous simulations^{5,26} and experimental findings.²⁷ The first stable water layer, with a density of 3000 kg m^{-3} is formed due to strong interactions between water molecules and calcite surface ions. These water molecules are oriented parallel to the surface, directly bind to the outermost Ca groups, and are also able to form weak hydrogen bonds with the top-surface CO_3 groups. The second layer of water molecules are oriented perpendicular to the surface and strictly form hydrogen bonds with the oxygen atoms of the outermost CO_3 groups. The radial distribution function of water around the outermost Ca sites showed that in average (simulation time 10 ns) 1.7 water molecules form direct interactions with each of the outermost Ca sites with an average Ca–O distance of 0.238 nm (Fig. 2(a, inset)).

The interfacial water at the (1–10) surface on the other hand is much less structured. Three major water density peaks are located at 0.04, 0.18 and 0.44 nm from the surface (Fig. 2(b)). The density of the first and second water layers are 1500 and 2500 kg m^{-3} , respectively. The water molecules interact directly with the outermost Ca, the underlying Ca layer and outermost CO_3 groups. Four water molecules are within the first coordination shell of each outermost Ca site, at an average Ca–O distance of 0.240 nm (Fig. 2(b, inset)). The third water layer interacts mainly with the water molecules in the second layer, additionally neutralizing the high charge of the surface ions, resulting in a density of 1300 kg m^{-3} .

The lower densities of water layers and their less structuring could originate from much lower surface ion density of 2.4 ions nm^{-2} on the (1–10) face, in comparison to 5.0 ions nm^{-2} on the (104) calcite face. In addition, the partial surface charge is also higher on the (1–10) face (three dangling bonds), so more water molecules coordinate to highly charged surface ions ($z = \pm 1$) leading to less structured layers.



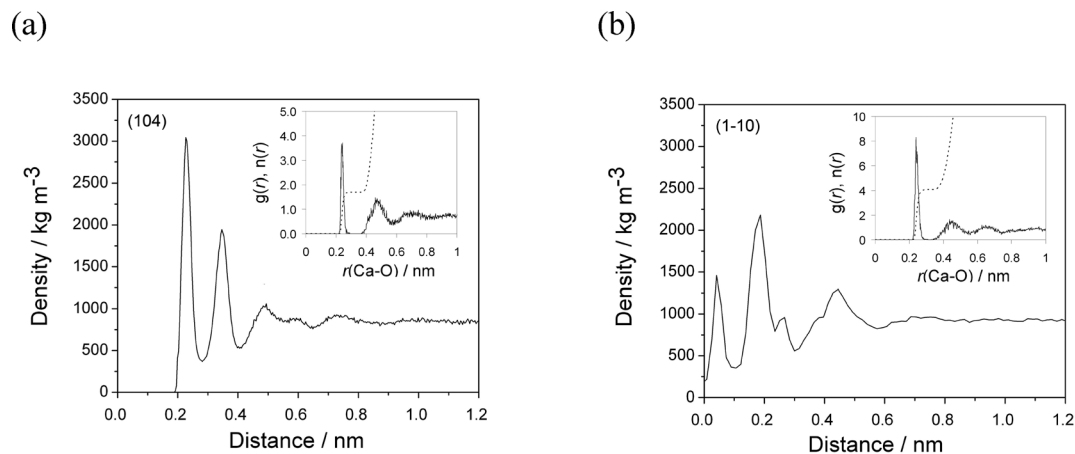


Fig. 2 Water density profiles normal to the calcite (104) (a) and (1-10) surfaces (b). The distance from both surfaces is defined relative to the average position of the Ca sites in the outermost layer. Insets show Ca-O_{water} pair distribution functions, $g(r)$, of water molecule around the outermost surface calcium ions (insets); running integration numbers, $n(r)$, are presented as dashed lines.

3.2. Tartrate in water

Tartrate forms a number of stable conformations, namely *trans* and *gauche* conformers, with the corresponding dihedral angle of about 60° and 180° , respectively (Fig. 3(up)). It is now well established by experimental⁴² and quantum chemical simulations⁴³ that *trans* conformation of the carbon chain is the most stable both in vacuum and water, independent of pH conditions. The original GAFF parameters failed to reproduce the right conformation of tartrate in water (missing appropriate dihedral parameters for diol/dicarboxylate), thus the carbon chain dihedral parameters and partial

charges were re-parameterized by quantum chemical calculations, as mentioned in section 2.1. To obtain dihedral parameters, energies of the optimized tartrate structures in vacuum were calculated for the carbon chain dihedral angles within $0-360$ degree range (Fig. 3(down)). The Ryckaert-Bellemans (RB) dihedral potential function ($E_{\text{dihedral}} = \sum C_n (\cos \varphi - 180^\circ)^n$) was least-squares fitted to minimize the deviations from DFT energy profile, with the following dihedral parameters obtained: $C_1 = 47.48 \text{ kJ mol}^{-1}$; $C_2 = 9.59 \text{ kJ mol}^{-1}$; $C_3 = 28.46 \text{ kJ mol}^{-1}$; $C_4 = 31.60 \text{ kJ mol}^{-1}$. The calculated charges are given in ESI.†

MD simulation of tartrate in water, using the newly obtained parameters, showed that the most stable tartrate conformation is *trans*, with carbon chain dihedral angle of 178° , which is in agreement experimental results.⁴² The relaxed tartrate structure did not significantly differ from the structure in vacuum. One of the hydroxyl groups is bound to form an intramolecular hydrogen bond with the proximal carboxylate oxygen atom while the other hydroxyl group is more free to rotate and more likely to be involved in intermolecular hydrogen bonds, interacting with water molecules (Fig. 4). This relaxed structure of tartrate dianion in aqueous

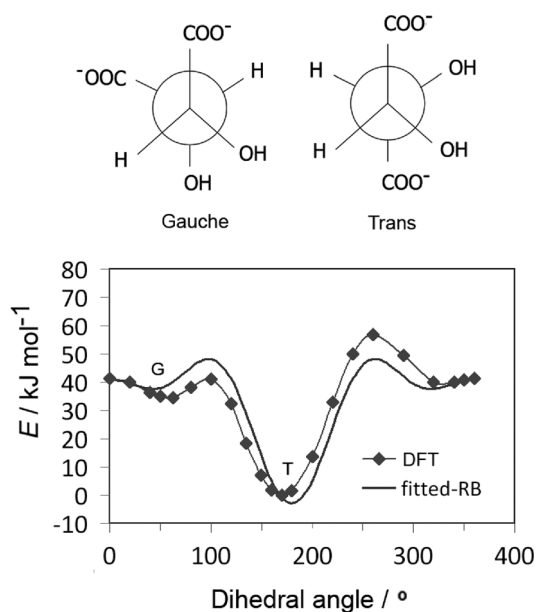


Fig. 3 The two principal tartrate conformations: *gauche* (G) and *trans* (T) conformers (up). Energy profile of tartrate dihedral scan obtained from DFT calculations and the fitted RB function used in MD simulations (down). Dihedral angle of tartrate is defined as carbon chain dihedral angle.

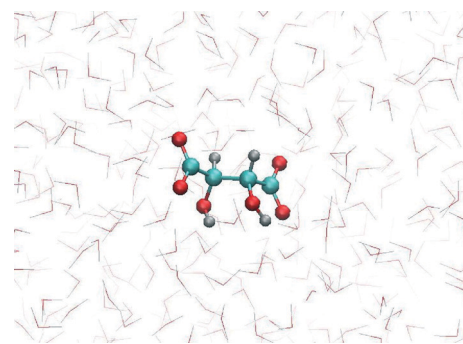


Fig. 4 MD snapshot of relaxed molecular structure of tartrate dianion immersed in water solution.



solution is consistent with other DFT simulation study⁴³ which indicated the same non-symmetrically hydration, while preserving the conformation of the carbon skeleton. This indicated the validity of our parameterization for tartrate dianion, which is critical to investigate accurately the interactions of tartrate with calcite surfaces.

To calculate the energy barrier of the conformational change in solution, the metadynamics was employed; the dihedral angle of the carbon chain was used as CV. The change in energy of different conformational states in water was found to be relatively high, about 13 kJ mol⁻¹ between *gauche* and the most stable *trans* conformation. This energy difference is significantly higher if compared with the thermal energy barrier, about 2.5 kJ mol⁻¹ at 300 K, so the conformational change could not be seen in unbiased MD simulations.

3.3. Tartrate sorption on calcite surfaces

The results of the metadynamics provided the free energy landscape for the sorption of tartrate on different calcite surfaces as a function of the three selected CVs: distance from the surface, molecular orientation ($\cos(\theta)$) and conformation (carbon backbone dihedral angle). In addition to the distance from the surface and the molecular orientation we chose the molecular conformation as the third CV, since the conformational change could not be seen in unbiased MD simulations, as mentioned above. At convergence, the four-dimensional (4D) free energy landscape of the system was obtained. This surface was then projected down to a 2D free energy profile, for which the binding free energy is determined (as described in the *methods* section). The favorable orientation and conformation at discrete positions of tartrate from the surface is explored by the 3D free energy contour plots.

Adsorption on (104) surface. The free energy profile of tartrate adsorption on calcite (104) face, presented as a function of the distance from the surface, is shown on Fig. 5(a). The free energy reaches a zero value plateau, beyond 0.8 nm from the surface, when the tartrate molecule moves into the bulk

water phase. Near the (104) surface, the energy profile displays two binding free energy minima of -6.5 and -2.9 kJ mol⁻¹ located at 0.48 and 0.66 nm from the surface, respectively. Both minima refer to the adsorption configurations where tartrate interacts with the structured water layers between the molecule and the surface. The tartrate molecule do not significantly penetrate the water layer on the (104) surface.

The free energy landscape shown for the first energy minimum as a function of molecular orientation and dihedral angle indicates a global minimum at $\cos(\theta) \approx 0$ and dihedral angle of 172° (Fig. 5(b)). In this most stable adsorption configuration, tartrate is oriented parallel to the surface in its *trans* conformation. The corresponding structure is shown in Fig. 6(a). Both carboxylate groups interact with the interfacial water molecules, located above the two Ca surface sites, whereas hydroxyl groups penetrate the water layer interacting with surface CO₃ groups. At the second energy minimum, tartrate is oriented almost perpendicular to the surface ($\cos(\theta) = 0.8$), with one of the hydroxyl groups pointing towards the bulk solution (Fig. 6(b)). The hydroxyl group penetrates into the first water layer, however the direct interactions between charged carboxylate groups and surface ions were not observed in any of the identified adsorption configurations for the (104) surface. It seems that the presence of the surface carbonate ions specifically arranged and exposed on the (104) surface makes the direct binding of the carboxylate groups to the calcium ion unfavorable. Nevertheless, stable adsorption of the tartrate molecule to the (104) surface in a water-separated configuration is favored by the possibility to form strong hydrogen bonds between the ordered water molecules and the carboxylate groups. Moreover, hydroxyl groups could form hydrogen bonds with outermost CO₃ groups as well as interactions within the first two interfacial water layers.

Adsorption on (1-10) surface. The free energy profile on the (1-10) surface is clearly different from that on the (104) surface. Three free energy minima of -38.2 , -24.7 and -2.5 kJ mol⁻¹ are obtained at 0.28, 0.38 and 0.55 nm from the surface, respectively (Fig. 7(a)). Evidently, tartrate is able to penetrate

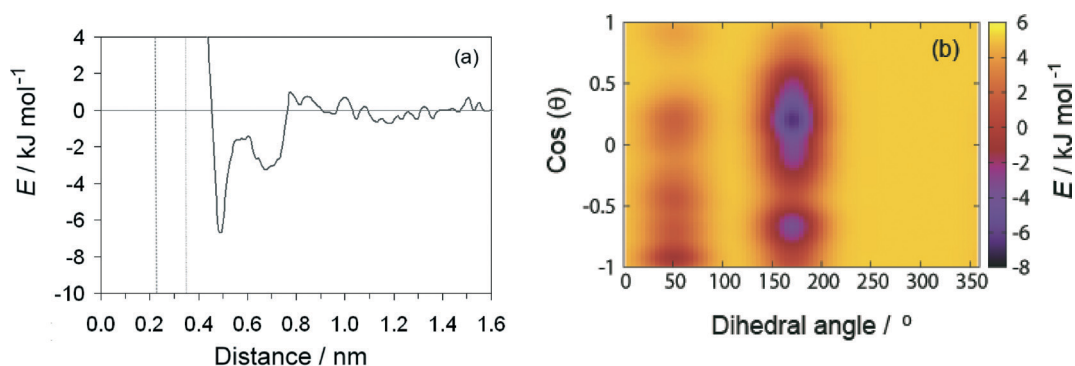


Fig. 5 (a) Free energy as a function of tartrate distance from the calcite (104) surface for the most stable *trans* conformation with dihedral angle of 172°. The distance from the surface is defined relative to the average position of the Ca sites in the outermost layer. The dashed and dotted lines indicate the positions of the water density maxima on (104) surface. (b) Free energy landscape as a function of orientation, $\cos(\theta)$, and dihedral angle at fixed distance of 0.48 nm from the (104) surface.



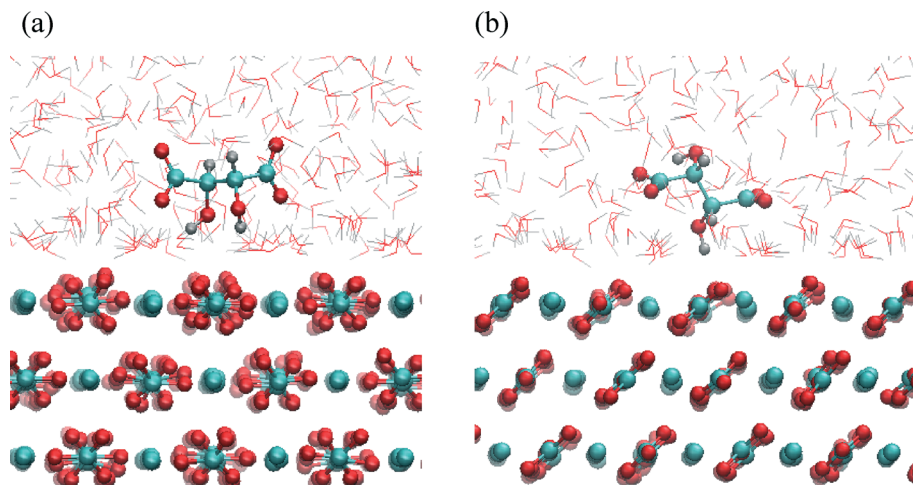


Fig. 6 Characteristic snapshots of calcite/water–tartrate interface for (104) surface. Water-mediated adsorption states of tartrate in the most stable *trans* conformation oriented parallel (a) and perpendicular (b) to the surface.

the interfacial water layers, on the (1–10) face, thus binding more closely and with much stronger interactions than at the (104) face. Significant energy barriers with heights of 8 and 25 kJ mol^{-1} separate the free energy minima of sorption. Moving away from the surface along the sorption pathway, the free energy approaches zero beyond 0.8 nm from the surface, where the tartrate molecule effectively is in bulk water.

At the global minimum of $-38.25 \text{ kJ mol}^{-1}$, the tartrate molecule lies 0.28 nm from the surface adsorbed only in *trans* conformation, whereas at the local minima both *gauche* and *trans* conformations were observed. Interestingly, for the local minima positioned at 0.38 nm from the surface the *gauche* conformation is significantly stabilized when adsorbed on the (1–10) calcite surface. The tartrate molecule changes its conformation in order to match the surface geometry. This is clearly evident from the free energy landscape plotted as a function of dihedral angle and molecular orientation at fixed tartrate positions corresponding to the two minima 0.28 nm (Fig. 7(b)) and 0.38 nm (Fig. 7(c)). For the first minimum tartrate lies 0.28 nm with the dihedral angle of 161° indicating *trans* conformational state when adsorbed. Beside the global minimum at $-38.25 \text{ kJ mol}^{-1}$, where the tartrate molecule is oriented parallel to the surface, $\cos(\theta) \approx 0$, two additional local minima exist with tartrate in perpendicular orientation, $\cos(\theta) = 0.8$ and -0.9 , respectively. However, the surface-induced conformational change is evident for the adsorbed configuration positioned at 0.38 nm, where *gauche* conformation, with the dihedral angle of 54° , is favored. To closely match the outermost surface ions, the tartrate molecule is oriented parallel to the surface, $\cos(\theta) \approx 0$.

The structures corresponding to the two energy minima are shown in Fig. 8. At the binding configuration closest to the surface, the polar functional groups of tartrate are not only bonded to the outermost surface ions but also to the underlying Ca sites (in the second layer) (Fig. 8(a, b)). In this case we can define tartrate as incorporated into the (1–10) surface since one of the polar functional groups matches the

position and orientation of the carbonate lattice ions.⁴⁴ In the most stable configuration one of the carboxylate groups fits into the (1–10) surface, coordinated by three different Ca surface sites (Fig. 8(b)). Specifically, the incorporated carboxylate group forms a stable interaction with two outermost Ca sites at a Ca–O distance of 0.22 and 0.24 nm, and with one underlying Ca site at a Ca–O distance of 0.23 nm. Both hydroxyl groups interact strongly with oxygen atoms of a two outermost CO_3 surface groups at H– OCO_2 distances of 0.16 and 0.22 nm. Apart from that, one of the hydroxyl groups additionally binds *via* oxygen atom with an underlying Ca site with a Ca–O distance of 2.33 nm. Another, similar stable configuration is shown on Fig. 8(b), where one of the carboxylate groups is incorporated in the (1–10) surface. In this state tartrate is oriented parallel to the surface as well, forming strong interactions with each of the hydroxyl groups at H– OCO_2 distances of 0.18 and 0.22 nm. The incorporated carboxylate group is coordinated with four Ca sites, two outermost and two underlying, with Ca–O distances 0.24, 0.23 and 0.24, 0.26 nm, respectively. Tartrate incorporation in the surfaces is additionally stabilized by the possible hydrogen bonding between hydroxyl groups and outermost CO_3 groups. In both configuration cases described above, the carboxylate group oriented perpendicular to the surface perfectly fits into the lattice positions normally occupied by carbonate ions of (1–10) surface. This finding is in good agreement with the experimental results showing that organic molecules, with carboxylate side groups, could preferentially adsorb and incorporate into (1–10) calcite interfaces.^{13–16} The incorporation of organic molecules leads to a distortion of the crystal lattice in *c*-direction, and high resolution powder diffraction studies¹⁴ supported the model that the incorporation of acidic macromolecules in biogenic calcite occurs on this specific crystal planes that are parallel to the crystallographic *c*-axis. It was proposed that this should occur by the specific interaction of the acidic residues, *i.e.* carboxylate groups that can replace the lattice



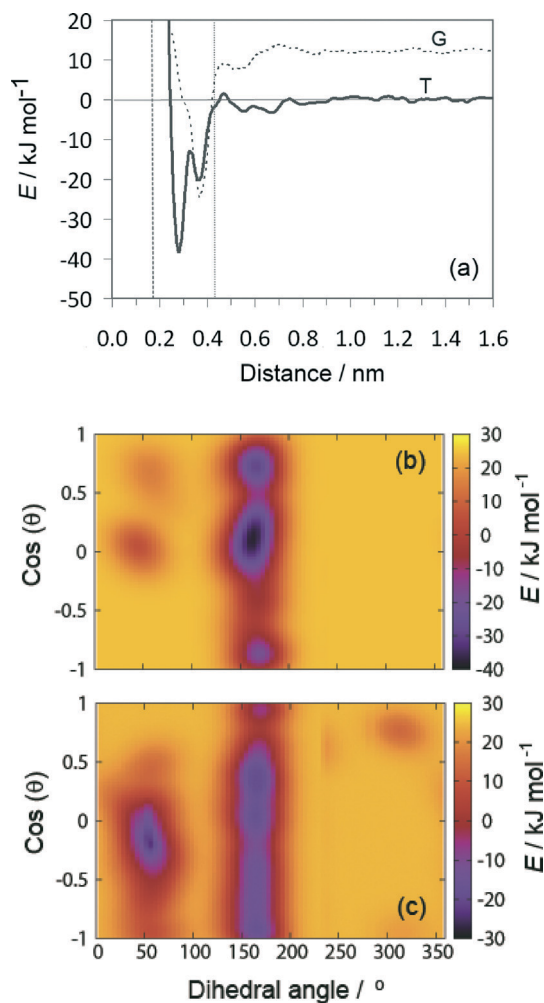


Fig. 7 Free energy as a function of tartrate distance from the calcite (1-10) surface (a). The distance from the surface is defined relative to the average position of the Ca sites in the outermost layer. Both *trans* (T) and *gauche* (G) conformations of tartrate are shown for the dihedral angle of 54° and 161° , respectively. The dashed and dotted lines indicate the positions of the water density maxima on the (1-10) surface. Free energy landscape as a function of orientation, $\cos(\theta)$, and dihedral angle shown for fixed distances of 0.28 nm (b) and 0.38 nm (c) from the (1-10) surface.

carbonate ions in these planes. Our simulation results directly support this proposal.

At the local minimum, where tartrate lies 0.38 nm from the surface, the tartrate molecule changes its conformation in order to match the surface geometry to maximize its interactions with the outermost surface ions (Fig. 8(c, d)). In this particular configuration tartrate lies parallel to the surface and forms four stable interactions with the surface. Each of the carboxylate groups interacts strongly with the outermost Ca surface sites; the one in a monodentate mode with Ca–O distance of 0.22 nm and the other in bidentate mode with Ca–O distances of 0.22 and 0.34 nm. Each of the hydroxyl groups directly binds to the oxygen atoms of the outermost CO_3 groups at H– OCO_2 distances of 0.22 and 0.23 nm. Thus, the (1-10) geometry of the outermost surface sites closely

matches the particular *gauche* conformation of tartrate with the corresponding dihedral angle of 54° . Despite the higher energy of the *gauche* conformation, adsorption at the (1-10) surface is energetically more favorable than for the otherwise more stable *trans* conformation in solution, indicating that the close interaction at the surface overcompensates the higher conformational energy. This emphasizes the importance of preorganization of functional groups to closely match the surface structure to stabilize the adsorption. Such mechanism of interaction could also be important for other organic molecules, for examples in the case of highly efficient multifunctional growth-modifiers, such as organic polymers or biomolecules, capable of forming multiple conformations. Similar study simulating tri-peptide adsorption onto hydroxyapatite,⁴⁵ also pointed to such a possible mechanism of adsorption with the surface-induced conformational change of tri-peptide.

Tartrate adsorption behavior. The obtained relatively low adsorption binding free energies (-6.5 kJ mol^{-1}) and the fact that bonded tartrate is separated by interfacial water molecules indicate the physical adsorption of tartrate onto the (104) surface. The tartrate molecule is loosely bonded to the (104) calcite surface and can easily desorb from the surface and diffuse into the bulk solution. The barrier height of about $3\text{--}4 \text{ kJ mol}^{-1}$ is comparable to thermal fluctuations at room temperature ($\sim 2.5 \text{ kJ mol}^{-1}$), indicating that tartrate is capable of kinetically overcoming the free energy barrier of sorption. On the other hand, the sorption barrier height identified in the case for the (1-10) surface is significantly higher, about 8 and 25 kJ mol^{-1} , and indicate strong irreversible adsorption. To test this hypothesis we performed unbiased MD simulations (20 ns) and found that tartrate, once adsorbed on the (104) surface, easily diffuses away from the surface. On contrary, tartrate was immediately adsorbed, even after the equilibration time, and stayed irreversibly in one position on the (1-10) surface during the performed simulation time (20 ns). This is driven by the strong electrostatic interactions between negatively charged carboxylate groups and positively charged Ca sites on the (1-10) calcite surface. The obtained binding configuration corresponded to the local minimum adsorption state with a distance from the surface of 0.38 nm, lying perpendicular to the surface. One carboxylate group was interacting with the outermost Ca site and one hydroxyl group with outermost CO_3 group, both staying firmly adsorbed in one position during the simulation. The other two hydroxyl and carboxylate groups were oriented towards the solution, freely and randomly rotating and bending due to thermal motion. Thus tartrate is trapped in a local minimum, due to high free energy barriers, preventing it to transfer from the surface to bulk solution or to another adsorption configuration corresponding to the global minimum.

The results of interfacial structure analysis and metadynamics showed that the interactions between tartrate and different calcite faces strongly relate to the structure of (104) and (1-10) calcite–water interfaces. The differences in tartrate adsorption affinities for these two surfaces can be interpreted



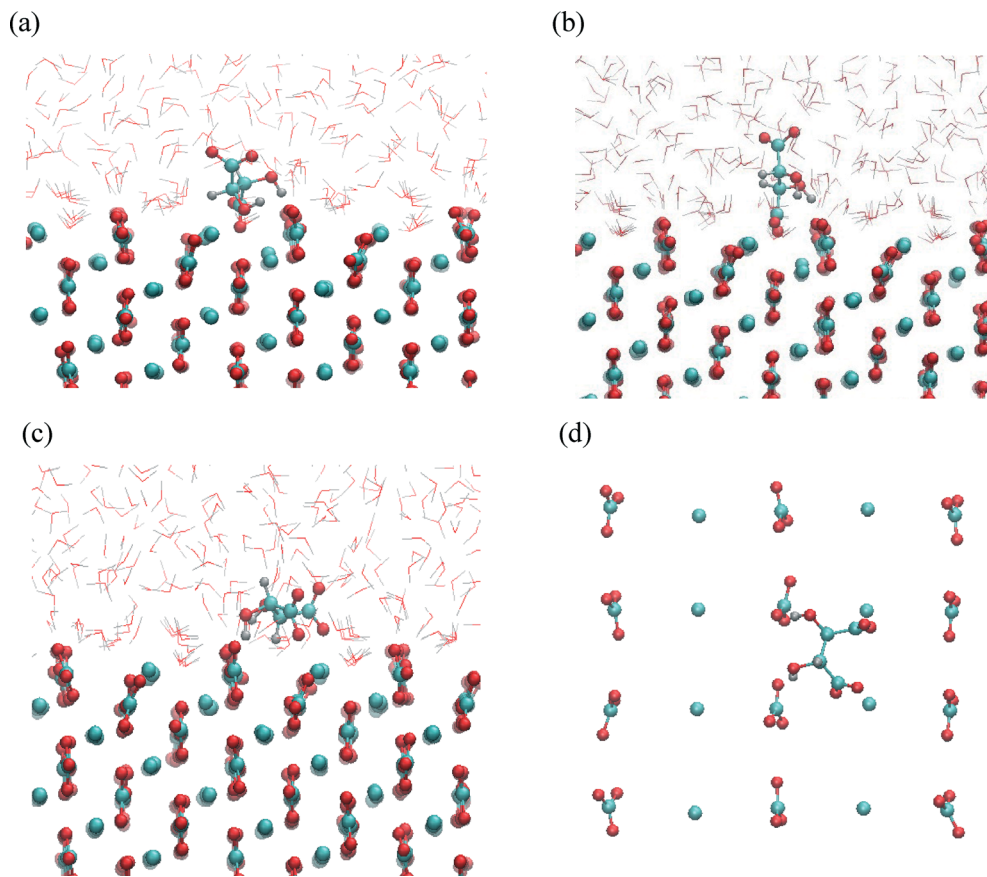


Fig. 8 Characteristic snapshots of calcite/water–tartrate interface for (1–10) surfaces. Tartrate incorporated into the surface in *trans* conformation (a, b) where carboxylate group substitutes lattice carbonate ion coordinated by (a) three Ca sites and (b) four Ca sites, and adsorbed on the surface in *gauche* conformation perfectly matching the geometry of outermost surface ions (c, d): (c) side view and (d) top view.

in terms of surface energetics and differently arranged surface ions. The energy difference between the lowest-energy (104) surface and more reactive (1–10) surface is $\Delta\gamma = 0.3 \text{ J m}^{-2}$. Thus, the (1–10) surface must be significantly stabilized and prevented from growing out, to be exposed in the final crystal morphology (Fig. 1). Adsorption of the tartrate molecules, could stabilize this otherwise high-energy and fast-growing face. Interactions between tartrate and the (1–10) are stronger with this particular higher energy surface due to special arrangements of the surface ions and their higher reactivity, *i.e.* higher partial charges. Surface ions specifically arranged and exposed on the (1–10) surface enables the direct binding of carboxylate groups to calcium ions with the simultaneous binding of both hydroxyl groups to neighboring surface carbonate ions. Especially, the position of the surface carbonate group, oriented perpendicular to the surface has a positive effect, where carboxylate functional group fits into the lattice position normally occupied by carbonate ions, and thus coordinate to several calcium ions. On contrary, more than the twice higher surface ion density of the (104) surface ($5.0 \text{ ions nm}^{-2} > 2.4 \text{ ions nm}^{-2}$) makes the direct binding of the charged carboxylate groups to the calcium ion unfavorable due to the close repulsive interactions with the surface carbonate ions exposed on the (104) surface as well as lower partial charges of surface ions on this stable face.

Moreover, local water density variations at the interface affect the adsorption binding modes of tartrate, which significantly differs for the two surfaces. The structuring of the water is higher for the (104) face, most probably due to the higher density of the surface ions which favors the adsorption of water molecules forming multiple layers above the surface. Within these ordered layers water molecules not only interact with the surface ions but also form strong lateral interactions. Keresit *et al.*⁴⁶ demonstrated that water binding energy (-2 kJ mol^{-1}) is small relative to the adsorption energy of water (-45 kJ mol^{-1}) indicating a free-energy loss by an entropic term, *i.e.* significant structuring at the interface. Thus, tartrate not only competes with water molecules for the binding sites on the (104) surface but can also interact with highly-structured interfacial water. However, the extent of such interactions are significantly lower in comparison to the direct interactions with the highly charged surface ions exposed on the more reactive (1–10) surface. Thus, we confirm that interfacial water is crucial for the tartrate adsorption behavior on calcite surfaces, as it has previously been recognized for selective adsorption of polystyrene sulfonate on calcite⁵ and citric acid²⁵ and amino acids^{23,47,48} on hydroxyapatite surfaces.

Growth of calcite is considered to proceed by the movement of steps on the most stable (104) surface. Therefore, the



interaction of additive molecules with step sites on the (104) surface should also be considered, as the previous simulations demonstrated that small carboxylated molecules preferentially adsorb along the step sites of the growing (104) surface, rather than on the (104) terraces.^{6,22,49} Such a selective interaction of the additive with step sites on the (104) surface could be an important mechanism in the early evolution of the morphological change, influencing the topography of the crystal face. Experimental studies indicated that tartrate¹⁹ and dicarboxylated amino-acids^{41,49} exhibit new edges on (104) calcite surface suggesting the formation of specifically oriented steps of the original (104) face. Interactions first occur at the step edges of the {104} surfaces which finally results in the expression of near- $\{hk0\}$ faces, stabilized by an additive.

The strong preferential binding of tartrate to the (1-10) calcite surface in comparison to weak, water-mediated binding on the (104) surface, explains why tartrate stabilizes the higher energy (1-10) surface and inhibits the growth in this direction. This result agrees well with the experimental observation that tartrate modulates the calcite morphology by exposing the {1-10} faces.¹⁹ Thus, the crystal can grow in the direction perpendicular to the (104) face but growth is limited perpendicular to the (1-10) face, resulting in a prismatic morphology combined with two forms, predominantly expressed prismatic {1-10} form and basal {104} faces (Fig. 1).

4. Conclusions

The molecular interactions of tartrate with calcite surfaces in aqueous solution were investigated by molecular dynamics simulations to understand additive-mediated crystal growth. The metadynamics approach was employed to determine the adsorption binding energies as a function of tartrate conformation, orientation and position near (104) and (1-10) calcite surfaces.

The results emphasize the importance of the surface energetics, structural matching as well as interfacial water layers in calcite-tartrate interactions. It was found that the adsorption behavior of tartrate is quite different for the two surfaces. Tartrate was more loosely adsorbed on the (104) surface mainly interacting *via* a water-mediated binding. On the other hand, tartrate was directly and irreversibly bonded to the (1-10) surface with two stable configurations identified: (1) adsorbed tartrate with the effect of surface-induced conformational change, favoring *gauche* conformation to closely match the geometry of the outermost surface ions; (2) incorporated tartrate into the surface where carboxylate group, oriented perpendicular to the surface, coordinate with outermost and underlying Ca sites, substituting the lattice carbonate ion. Incorporation of tartrate, by fitting the carboxylate group into lattice positions, is additionally stabilized by binding of both hydroxyl groups to neighboring surface carbonate ions. Thus, preorganization of both charged carboxylate functional groups as well as polar hydroxyl groups plays important role in the strength of the interactions.

The large difference in the adsorption modes for these two surfaces are responsible for preferential adsorption on {1-10} faces, which thus become more pronounced in the calcite morphology at the expense of the {104} faces, which are less affected by tartrate. These results indicate that the strong adsorption of tartrate on {1-10} surfaces could stabilize these otherwise fast-growing and higher energy faces and thus inhibit crystal growth in {1-10} directions. The results also supported the hypothesis that incorporation of acidic molecules in calcite could occur on (1-10) faces by fitting side chains carboxylate functional groups into carbonate lattice positions. Our study demonstrated that MD simulations coupled with metadynamics can be used as a powerful tool for free energy sampling of all possible adsorbed states giving valuable insight into mechanisms and modes of molecular interactions on crystal/water-additive interfaces. The obtained results contribute to fundamental understanding of the molecular interactions between different crystal surfaces and organic additives needed for rational development of novel growth-controlling additives. Our future work will focus on simulating the growth from supersaturated solutions of specific fast-growing calcite faces, such as (1-10), in the presence of organic additives.

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References

- 1 K. Sangwal, *Additives and Crystallization Processes: From Fundamentals to Applications*, John Wiley & Sons, Chichester, 2007.
- 2 R.-Q. Song and H. Cölfen, *CrystEngComm*, 2011, **13**, 1249–1276.
- 3 M. Salvalaglio, T. Vetter, F. Gilberti, M. Mazzotti and M. Parrinello, *J. Am. Chem. Soc.*, 2012, **134**, 17221–17233.
- 4 C. Schmidt and J. Ulrich, *J. Cryst. Growth*, 2012, **353**, 168–173.
- 5 J. Shen, C. Li, N. F. A. van Vegt and C. Peter, *J. Phys. Chem. C*, 2013, **117**, 6904–6913.
- 6 N. A. J. M. Sommerdijk and G. De With, *Chem. Rev.*, 2008, **108**, 4499–4550.
- 7 Industrial Minerals Association, CCA-Europe, <http://www.cca-europe.eu/mineral-applications.html>, (accessed June 2014).
- 8 M. Ukrainczyk, J. Kontrec and D. Kralj, *J. Colloid Interface Sci.*, 2009, **329**, 89–96.
- 9 J. Kontrec, M. Ukrainczyk, V. Babić-Ivančić and D. Kralj, *Croat. Chem. Acta*, 2011, **84**, 25–32.
- 10 L. Estroff and I. Cohen, *Nat. Mater.*, 2011, **10**, 810–811.
- 11 Y. J. Lee and R. J. Reeder, *Geochim. Cosmochim. Acta*, 2006, **70**, 2253–2263; B. L. Phillips, Y. J. Lee and R. J. Reeder, *Environ. Sci. Technol.*, 2005, **39**, 4533–4539.



- 12 M. Ukrainczyk, M. Gredičak, I. Jerić and D. Kralj, *Cryst. Growth Des.*, 2014, **14**, 4335–4346.
- 13 L. Addadi and S. Weiner, *Angew. Chem., Int. Ed. Engl.*, 1992, **31**, 153–169.
- 14 B. Pokroy, A. N. Fitch and E. Zolotoyabko, *Adv. Mater.*, 2006, **18**, 2363–2368.
- 15 D. B. De Oliveira and R. A. Laursen, *J. Am. Chem. Soc.*, 1997, **119**, 10627–10631.
- 16 A. Wierzbicki, C. S. Sikes, J. D. Madura and B. Drake, *Calcif. Tissue Int.*, 1994, **54**, 133–141.
- 17 S. Mann, J. M. Didymus, N. P. Sanderson, B. R. Heywood and E. J. A. Samper, *J. Chem. Soc., Faraday Trans.*, 1990, **86**, 1873–1880.
- 18 C. L. Freeman, I. Asteriadis, M. Yang and J. H. Harding, *J. Phys. Chem. C*, 2009, **113**, 3666–3673.
- 19 M. Ukrainczyk, J. Stelling, M. Vucak and T. Neumann, *J. Cryst. Growth*, 2013, **369**, 21–31.
- 20 C. Geffroy, A. Foissy, J. Persello and B. Cabane, *J. Colloid Interface Sci.*, 1999, **211**, 45–53.
- 21 W. Shi, C. Ding, J. Yan, X. Han, Z. Lv, W. Lei, M. Xia and F. Wang, *Desalination*, 2012, **291**, 8–14.
- 22 N. H. de Leeuw and T. G. Cooper, *Cryst. Growth Des.*, 2004, **4**, 123–133.
- 23 Z. Xu, Y. Yang, Z. Wang, D. Mhonoto, C. Shang, Z. Liu, Q. Cui and N. Sahai, *J. Comput. Chem.*, 2014, **35**, 70–81.
- 24 B. Zhu, X. Xu and R. Tang, *J. Chem. Phys.*, 2013, **139**, 234705.
- 25 N. H. de Leeuw and J. A. L. Rabone, *CrystEngComm*, 2007, **9**, 1178–1186.
- 26 P. Raiteri, J. D. Gale, D. Quigley and P. M. Rodger, *J. Phys. Chem. C*, 2010, **114**, 5997–6010.
- 27 P. Fenter, S. Keresit, P. Raiteri and J. D. Gale, *J. Phys. Chem. C*, 2013, **117**, 5028–5042.
- 28 S. Xiao, S. A. Edwards and F. Gräter, *J. Phys. Chem. C*, 2011, **115**, 20067–20075.
- 29 A. Barducci, M. Bonomi and M. Parrinello, *Wiley Interdiscip. Rev.: Comput. Mol. Sci.*, 2011, **1**, 826–843.
- 30 D. van der Spoel, E. Lindahl, B. Hess, G. Groenhof, A. E. Mark and H. J. C. Berendsen, *J. Comput. Chem.*, 2005, **26**, 1701–1718.
- 31 J. Wang, R. Wolf, J. Caldwell, P. Kollman and D. Case, *J. Comput. Chem.*, 2004, **25**, 1157–1174.
- 32 A. D. Bochevarov, E. Harder, T. F. Hughes, J. R. Greenwood, D. A. Braden, D. M. Philipp, D. Rinaldo, M. D. Halls, J. Zhang and R. A. Friesner, *Int. J. Quantum Chem.*, 2013, **113**, 2110–2142.
- 33 E. Vanquelef, S. Simon, G. Marquant, E. Garcia, G. Klimerak, J. C. Delepine, P. Cieplak and F.-Y. Dupradeau, *Nucleic Acids Res.*, 2011, **39**, W511–W517.
- 34 C. Bayly, P. Cieplak, W. Cornell and P. Kollman, *J. Phys. Chem.*, 1993, **97**, 10269–10280.
- 35 S. D. Fleming and A. L. Rohl, *Z. Kristallogr.*, 2005, **220**, 580–584.
- 36 B. Kvamme, T. Kuznetsova and D. Uppstand, *J. Math. Chem.*, 2009, **46**, 756–762; M. F. Roberto, B. Marco and A. Dino, *Cryst. Growth Des.*, 2010, **10**, 4096–4100.
- 37 Q. Humphrey, A. Dalke and K. Schulten, *J. Mol. Graphics*, 1996, **14**, 33–38.
- 38 A. Laio and F. L. Gervasio, *Rep. Prog. Phys.*, 2008, **71**, 126601–126623.
- 39 M. Bonomi, D. Branduardi, G. Bussi, C. Camilloni, D. Provasi, P. Raiteri, D. Donadio, F. Marinelli, F. Pietrucci, R. A. Broglia and M. Parrinello, *Comput. Phys. Commun.*, 2009, **180**, 1961–1972.
- 40 N. H. de Leeuw and S. Parker, *J. Phys. Chem.*, 1998, **102**, 2914–2922.
- 41 M. Ukrainczyk, M. Gredičak, I. Jerić and D. Kralj, *J. Colloid Interface Sci.*, 2012, **365**, 296–307.
- 42 M. Hoffmann, J. Grajewski and J. Gawronski, *New J. Chem.*, 2010, **34**, 2020–2026.
- 43 J. Gawronski, K. Gawronska, O. Skowronek, U. Rychlewska, B. Warzajtis, J. Rychlewski, M. Hoffmann and A. Szarecka, *Tetrahedron*, 1997, **53**, 6113–6144.
- 44 E. Elts, M. Greiner and H. Briesen, *J. Chem. Theory Comput.*, 2014, **10**, 1686–1697.
- 45 N. Almora-Barrios and N. H. de Leeuw, *CrystEngComm*, 2009, **12**, 960–967.
- 46 S. Keresit and S. Parker, *J. Am. Chem. Soc.*, 2004, **126**, 10152–10161.
- 47 J. Tao, H. Pan, Y. Zeng, X. Xu and R. Tang, *J. Phys. Chem. B*, 2007, **111**, 13410–13418.
- 48 H. Pan, J. Tao, X. Xu and R. Tang, *Langmuir*, 2007, **23**, 8972–8981.
- 49 C. A. Orme, A. Noy, A. Wierzbicki, M. T. McBride, M. Grantham, H. H. Teng, P. M. Dove and J. J. DeYoreo, *Nature*, 2001, **411**, 775–779.

