Electroorganic synthesis of nitriles via a halogen-free domino oxidation–reduction sequence†

Marius F. Hartmer and Siegfried R. Waldvogel*

A direct electroorganic sequence yielding nitriles from oximes in undivided cells is reported. Despite the fact that intermediate nitrile oxides might be formed, the method is viable to prepare benzonitriles without substituents ortho to the aldoxime moiety. This constant current method is easy to perform for a broad scope of substrates and employs common electrodes, such as graphite and lead.

In the past, the formation of nitriles from oximes has been realized by the use of harsh dehydration agents such as phosphorous pentoxide, thionyl chloride, benzenesulfonyl chloride, acetic anhydride, and many others.1,2 Unfortunately, these strongly corrosive agents have to be used stoichiometrically or even in excess. This results in unattractive waste generation and considerable safety issues upon work up. Recently, catalytic approaches have emerged which reduce the waste generation significantly.3 However, recycling of the expensive and toxic catalysts still remains an issue. Consequently, new methods for the aldoxime dehydration are highly desired.

Electrochemistry has become an attractive alternative to conventional methods in organic synthesis, since only electricity is applied which might even originate from renewable resources, and therefore no reagent waste is produced.4 Shono et al. reported an electrochemical protocol to generate nitriles from oximes with halogenides as a mediator.5 However, such halogen species cause corrosion at the precious platinum electrodes.6,7 Additionally, low current yields make this protocol in particular unattractive. We report a direct, halogen-free procedure to generate nitriles from oximes by the application of inexpensive and easily available electrode materials.

We chose mesitylaldoxime (1a) as test substrate. Within the course of transformation we observe stable nitrile oxide as intermediate (2a, Scheme 1). Initially, we investigated the anodic reaction using methyltriethyl ammonium methylsulfate (MTES) in acetonitrile as electrolyte. Graphite and glassy carbon were tested as carbon anode materials at room temperature. Glassy carbon served as cathode material in both cases. Graphite turned out to be an excellent anode leading to almost quantitative consumption of the starting material upon application of 2.5 F (Fig. 1).

However, the conversion at glassy carbon electrodes was poor. Only the intermediate nitrile oxide was formed, indicating a possible deactivation of the cathodic surface. It turned out that the deoxygenation is not promoted. As outlined, the intermediate was observed and identified via GC-MS. The anodic conversion leads to nitrile oxide which is the same intermediate as in the...
halogenide-mediated electrolysis by Shono et al. Hence, we consider
the transformation as a domino oxidation–reduction sequence with
respect to the definition of domino reactions, since a stable inter-
mediate is formed which in turn reacts without manipulating the
reaction conditions to the final product (Scheme 2).8

Nevertheless, the key for the desired reaction sequence
seemed to be the deoxygenation of nitrile oxide formed. Having
some recent experience with lead electrodes and variations
thereof,9 we tested lead, stainless steel (A5), nickel, platinum,
and boron-doped diamond as cathode materials. Indeed, lead
turned out within the screening (see ESI† for GC-MS data) as an
excellent cathode material, whereas the deoxygenation was less
favored at the other electrode materials. Surprisingly, platinum
showed a very poor performance although it is known as the
cathode material of choice for this particular deoxygenation
reaction.5

Subsequently, the influence of temperature and current density
was studied (Table 1). Performing the electrolysis at 22 °C proved
to be an optimum concerning the temperature, whereas raising (50 °C)
or lowering (0 °C) the temperature led to a drop in yield (entries 1
and 5). Investigating the current density rendered the best yields
with 10 mA cm$^{-2}$ (entry 3), whereas lower (5 mA cm$^{-2}$) or higher
(20 mA cm$^{-2}$) current densities resulted in decreased yields and
selectivity, respectively (entries 2 and 4).

With the optimized electrolysis parameters in hand, we
applied the method to a range of substrates in order to
elucidate the scope of the reaction (Scheme 3). These substrates
vary in stability for the intermediately formed nitrile oxides
as well as their preference for dimerization. Stabilization of
reactive intermediates such as nitrile oxides can be achieved by
bulky groups ortho to the aldoxime moiety.

The best results were obtained with mesitylaldoxime (1a, 81%),
closely followed by 2,4,6-trimethoxybenzaldoxime (1c, 75%) and
2,6-dimethylbenzaldoxime (1b, 73%). The unexpectedly low yield
of the stabilized 2,6-dichlorobenzaldoxime (1f, 41%) can be
attributed to the competing dechlorination reaction (see ESI†
for GC-MS data). The analogue debromination reaction was even
more dominant and disadvantageous for the nitrile formation,
whereas 2,6-difluorobenzonitrile (3e, 47%) was isolated in a higher
yield due to the higher Ar–F bond strength (Ar–F: 526 kJ mol$^{-1}$;
Ar–Cl: 400 kJ mol$^{-1}$; Ar–Br: 336 kJ mol$^{-1}$).10

The stabilized 2-methoxy-1-naphthaldoxime (1h) resulted
in a yield of 55%, whereas the partially stabilized naphthyl-
aldoxime (1g) led to a yield of 40%. Surprisingly the non-
stabilized 4-methoxybenzonitrile (3c) was synthesized in an
acceptable yield of 53%.

Table 1 Optimization of reaction parameters

<table>
<thead>
<tr>
<th>Entry</th>
<th>T [°C]</th>
<th>j [mA cm$^{-2}$]</th>
<th>Q [F]</th>
<th>Isolated yield$^a$ [%]</th>
</tr>
</thead>
<tbody>
<tr>
<td>1</td>
<td>0</td>
<td>10</td>
<td>2.5</td>
<td>70 (4)</td>
</tr>
<tr>
<td>2</td>
<td>22</td>
<td>5</td>
<td>2.5</td>
<td>63 (13)</td>
</tr>
<tr>
<td>3</td>
<td>22</td>
<td>10</td>
<td>2.5</td>
<td>81 (0)</td>
</tr>
<tr>
<td>4</td>
<td>22</td>
<td>20</td>
<td>2</td>
<td>77 (0)</td>
</tr>
<tr>
<td>5</td>
<td>50</td>
<td>10</td>
<td>2</td>
<td>66 (0)</td>
</tr>
</tbody>
</table>

$^a$ Yield of nitrile oxide determined by $^1$H NMR in parentheses.
In general, stabilized substrates result in higher yields than non-stabilized congeners. However, the corresponding aldehyde was detected in all electrolyses by reaction monitoring.

It is noteworthy, that the aldehyde formation was significantly higher for the non-stabilized substrates. It is known that aldazines are oxidized in a divided cell to iminoxyl radicals which dimerize to aldazine bis-N-oxides (5, Scheme 4).\(^1\)

These aldazine bis-N-oxides react further to aldehydes by loss of nitrogen. We anticipate that the dimerization takes place prior to the deoxygenation of the intermediate. In contrast, only small amounts of aldehyde have been observed for stabilized substrates. We anticipate that the stability of the iminoxyl radical formed is based on the sterical hindrance and therefore, no or very little dimerization occurs. Hence the formation of nitrile oxide is favored which then undergoes deoxygenation at the cathode leading to the nitrile in high yield.

In conclusion, we developed a direct, halogen-free protocol for the synthesis of nitriles from oximes with yields up to 81% by applying inexpensive and readily available electrode materials such as lead\(^12\) and graphite. This protocol is applicable for a range of substrates under constant current and ambient conditions. Furthermore, some insight into the mechanistic scenery can be deduced confirming a domino oxidation–reduction sequence.

We wish to acknowledge financial support from the Federal Ministry of Education and Research (CARLOTTA, FKZ 01DM14005).

Notes and references


6 Average pricing at London Metal Exchange (May 2015): 36716 S per kg.


12 Average pricing at London Metal Exchange (May 2015): 2 S per kg.