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A straightforward implementation of *in situ* solution electrochemical ^{13}C NMR spectroscopy for studying reactions on commercial electrocatalysts: ethanol oxidation

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Identifying and quantifying electrocatalytic-reaction-generated solution species, be they reaction intermediates or products, are highly desirable in terms of understanding the associated reaction mechanisms. We report herein a straightforward implementation of *in situ* solution electrochemical ^{13}C NMR spectroscopy for the first time that enables *in situ* studies of reactions on commercial fuel-cell electrocatalysts (Pt and PtRu blacks). Using ethanol oxidation reaction (EOR) as a working example, we discovered that (1) the complete oxidation of ethanol to CO_2 only took place dominantly at the very beginning of a potentiostatic chronoamperometric (CA) measurement and (2) the PtRu had a much higher activity in catalysing oxygen insertion reaction that leads to acetic acid.

Solution NMR spectroscopy is probably the method of choice in identifying and quantifying reaction-related species in mechanistic studies of homogeneous reactions.¹ However, it has rarely been applied to delineating electrochemically-driven reactions due largely to the intrinsic incompatibility between conventional NMR detection using inductive coils and the electrical conduction necessary for electrochemical (EC) reactions.^{2,3} Still, there has been a continuous effort over the last decade to bring this powerful analytical method into play in studying EC redox reactions *in situ*.^{4–8} In order to overcome the aforementioned detection incompatibility and also maintain the high-resolution feature of solution NMR, sophisticated designs of EC-NMR cells have been proposed in which deposition of cylindrically symmetric ultra-thin (<100 nm) Au film as the working electrode is a prerequisite.^{5,8} While the feasibility of such EC-NMR cells has been demonstrated, their configurations are not easily amenable to studying electrocatalysis in which the focus is on the performance of electrocatalysts as the working electrode. Recently, while still limited to ^1H NMR, Dunsch and co-workers developed a rather simple

EC-NMR cell that used carbon fibres as the working electrode.⁷ We report herein our adaptation of the EC-NMR cell for the first *in situ* solution ^{13}C NMR study of ethanol oxidation reaction (EOR) on commercial (Johnson Matthey or J-M) Pt black and PtRu black electrocatalysts by taking advantage of the fact that the electrocatalysts can be easily loaded onto the carbon fibres by simple drop-casting or a dip-adsorption process. We hope that this simple advancement will encourage the broader use of *in situ* EC-NMR in electrochemistry in general and electrocatalysis in particular.

Fig. 1 shows the schematic of the EC-NMR cell that we adapted based on a 10 mm standard commercial NMR tube used in routine solution NMR measurements. It consisted of a carbon fiber bundle as the working electrode (WE), a Pt gauze as the counter electrode (CE) and a thin chlorinated silver (Ag/AgCl)

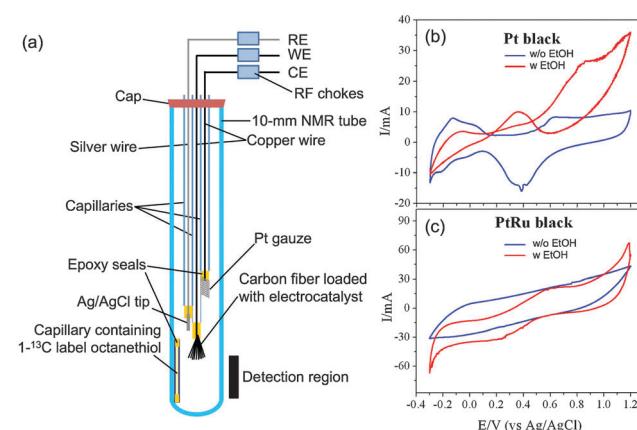


Fig. 1 (a) Schematic of the EC-NMR cell adapted from Dunsch et al.'s work. The cell is based on a standard 10 mm commercial NMR tube used in high-resolution solution NMR so it can be used directly with a routine solution NMR spectrometer. RF chokes are used to isolate the NMR detection from the EC potentiostat. The position of the carbon fibre bundle as the working electrode is critically important to preserve the original Q-factor of the NMR probe. (b) Normal (blue) and EOR (red) CVs on the activated Pt black. (c) Normal (blue) and EOR (red) CVs on the activated PtRu black.

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wire as the quasi-reference electrode (RE). The carbon fiber or Pt gauze was adjoined to a capillary-housed copper wire by conductive adhesive. The carbon-fiber-protruding end of the capillary was sealed by epoxy resin so that only carbon fiber could contact the electrolyte. The other end of the capillary was simply wrapped with poly(tetrafluoroethylene) (PTFE) tape to hold the connecting copper wire. The chlorinated Ag/AgCl quasi-RE was made by immersing a silver wire that was pre-cleaned in an aqua regia solution in a saturated KCl solution. To quantify the reaction species, we used an internal reference consisting of a short capillary sealed with 1^{13}C labeled octanethiol inside as indicated in Fig. 1. Its integrated NMR signal at 23 ppm was used to normalize all the NMR signals reported here. Radio frequency chokes were used to isolate NMR detection from any electrical interference using the potentiostat (VoltaLab PGZ100, Radiometer).

As a testing case of the above setup, we investigated the EOR on two different electrocatalysts, J-M Pt black and PtRu black, comparatively. All NMR measurements were carried out on a hybrid Bruker/Tecmag 300 MHz spectrometer using a standard 10 mm probe. ^{13}C -labeled ethanol $\text{CH}_3^{13}\text{CH}_2\text{OH}$ (99%, Cambridge Isotope) was used in the study. The normal Q-factor of the probe is ~ 75 , which gave a full width at half maximum (FWHM) of ~ 8 Hz for the ^{13}C -labeled ethanol peak at 58 ppm. We found that the Q-factor was very sensitive to the position of the carbon fibre bundle with respect to the NMR detection coil when the EC-NMR cell was used. If the carbon fibre bundle was just above the detection coil, as shown in Fig. 1, the Q factor was not affected, so a FWHM of 8 Hz could still be obtained for the ethanol peak at 58 ppm. Lowering further the carbon fibre bundle into the detection coil, the Q-factor dropped dramatically to ~ 25 , leading to lower sensitivity, making ^{13}C NMR unattainable. We used the former position for all the NMR measurements reported here; 128 scans were acquired for signal averaging for a given ^{13}C NMR spectrum using a standard proton-decoupling pulse sequence with a repetition delay of 5 s.

0.1 mL of the Pt or PtRu black suspension (20 mg mL^{-1}) was drop-casted onto the carbon fibre bundle and was then air-dried. Before any NMR measurements, the electrocatalyst adsorbed onto the carbon fibre bundle was activated by holding the electrode potential at -0.5 V and 0.8 V (vs. Ag/AgCl wire) for 5 s alternatively for 10 cycles in a 0.1 M HClO_4 + 0.1 M $\text{CH}_3\text{CH}_2\text{OH}$ solution. Fig. 1b and c present the normal (blue curve) and EOR (red curve) cyclic voltammograms (CVs) of the activated samples. After the activation, the original ethanol-containing electrolyte was exchanged to a 0.1 M HClO_4 + 0.1 M $\text{CH}_3^{13}\text{CH}_2\text{OH}$ solution for NMR measurements. Fig. 2 compares the NMR spectra obtained before (blue) and after (red) the EOR on Pt black (Fig. 2a) and PtRu black (Fig. 2b). In agreement with what has been known previously,⁶ two EOR-generated species were observed: acetic acid ($\text{CH}_3^{13}\text{COOH}$) at 177 ppm and acetaldehyde ($\text{CH}_3^{13}\text{CHO}$) at 208 ppm. As can be seen from the spectra, the EOR on the PtRu black produced a much higher percentage of acetic acid than on the Pt black while the latter produced a higher percentage of acetaldehyde, which showcased different reaction mechanisms on these two different electrocatalysts.

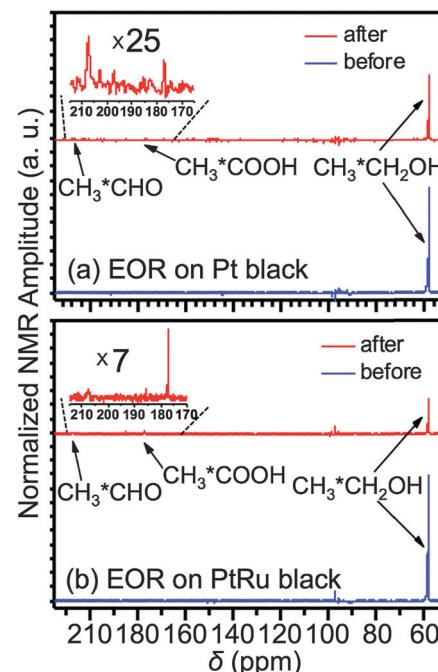


Fig. 2 *In situ* solution ^{13}C NMR spectra obtained before (blue) and after (red) EOR on the Pt black (a) and PtRu black (b). The insets are the zoom-in spectra of acetaldehyde (208 ppm) and acetic acid (177 ppm) produced during the EOR.

Further mechanistic details were gleaned by following the EOR on both samples for 24 h by the *in situ* solution EC ^{13}C NMR spectroscopy. Fig. 3a shows the chronoamperometric (CA) currents recorded over 24 h during the NMR measurements of the EOR at 0.2 V (vs. Ag/AgCl wire) on the Pt and PtRu blacks respectively (The spikes were caused by returning to 0.2 V after potential jumps to -0.5 V for 5 s to replenish the surface.) The CA currents indicate that the EOR activity was higher on the PtRu than on the Pt black. Fig. 3b and c present the NMR measured concentration changes of the three species, ethanol (black squares), acetic acid (red circles) and acetaldehyde (blue triangles), over 24 h by normalizing the peak integrals with that of the internal standard. No other species were observed during that period, which is in contrast with recent solution NMR measurements of ethanol fuel cell exhausts in which ethane-1,1-diol was observed during electrode-potential dependent experiments.⁶

More specifically, Fig. 3b shows that the EOR on the Pt black produced more acetaldehyde than acetic acid over the entire 24 h period. However, this is not the case for the EOR on the PtRu black: the production of acetaldehyde was only slightly higher than acetic acid over the first 7 h, then the amounts of the two species became equal over the next 7 to 8 h, and finally acetic acid took over in the remaining time.

The inverted pink triangles show the remaining total carbon content (RTCC) in the solution as measured with the ^{13}C NMR by adding ethanol, acetic acid and acetaldehyde together. It is intriguing to observe that the RTCC in solution for the EOR on the Pt black was more or less constant after the first hour, as

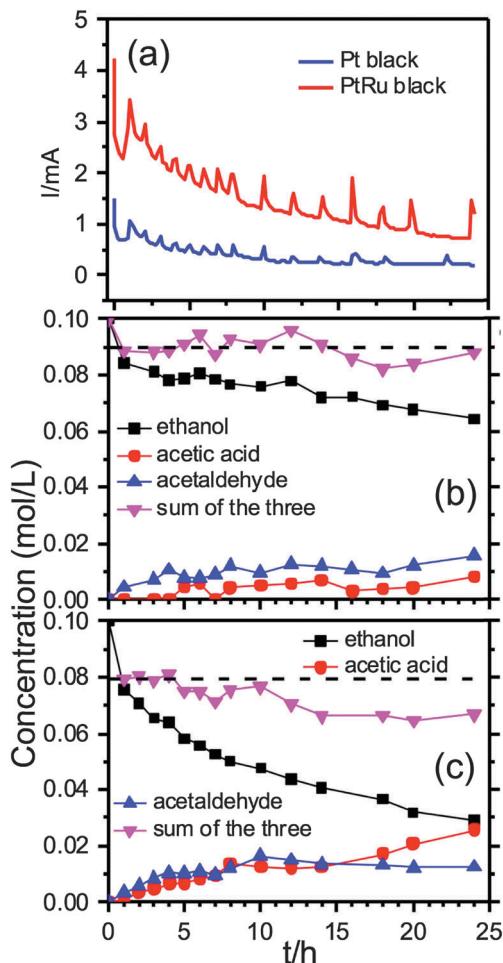


Fig. 3 (a) The CA currents at 0.2 V (vs. Ag/AgCl wire) recorded during the *in situ* ^{13}C NMR measurements of EOR on the Pt black (blue) and the PtRu black (red). The spikes were caused by returning to 0.2 V after potential jumps to -0.5 V for 5 s to replenish the surface. (b)/(c) ^{13}C NMR measured concentrations of ethanol (black squares), acetic acid (red circles) and acetaldehyde (blue triangles) during the EOR on the Pt black (b) and the PtRu black (c) over a period of 24 h. The inversed pink triangles are the remaining total carbon content (RTCC).

indicated by the horizontal dashed line in Fig. 3b. As the missing carbon content indicates presumably the production of volatile CO_2 , the data suggest that the complete oxidation of ethanol to CO_2 on the Pt black ($\sim 10\%$) only took place during the first hour of the reaction. The ensuing current came almost exclusively from reactions to acetaldehyde and acetic acid.

For the EOR on the PtRu black, the activity for complete oxidation of ethanol to CO_2 was also highest at the beginning hour of the reaction ($\sim 20\%$ conversion to CO_2 as indicated by the horizontal dashed line). But it continued, notwithstanding at a much slower rate, ($\sim 10\%$ conversion to CO_2) until the ~ 14 h mark after which the RTCC in solution became constant, again suggesting no further CO_2 generation. Overall, the ethanol conversion rate to other species was much higher, *i.e.*, more active, on the PtRu black ($\sim 70\%$) than on the Pt black ($\sim 30\%$), which is in good agreement with the CA data in

Fig. 3a. It appears that the PtRu is much better than the Pt black in catalysing oxygen insertion reaction that leads to acetic acid.

In summary, we have demonstrated the feasibility of a rather easy and versatile approach to realize *in situ* solution EC-NMR spectroscopy of electrocatalysis by adapting Dunsch *et al.*'s work⁷ and using EOR on commercial Pt and PtRu black as a working example to showcase the first *in situ* solution ^{13}C NMR study of the relevant electrocatalysis. The latter not only shows clearly mechanistic differences in the EOR on the Pt black *vs.* the PtRu black, but also reveals for the first time that the complete ethanol oxidation to CO_2 on these two electrocatalysts only took place dominantly at the very beginning of a potentiostatic CA reaction, a reaction mode that is most relevant to a fuel cell setting and warrants further detailed investigations. However, almost no CO_2 generation beyond the first hour of the EOR is likely caused by the observation that the initially active sites for C-C bond breaking were blocked by the EOR-generated poisonous species, such as adsorbed CO, as the EOR continued.⁹ As to the relatively higher CO_2 production and oxygen insertion reaction rate observed on the PtRu, they could be rationalized by the appearance of the Pt–Ru boundary sites responsible for more C-C bond breaking and the more oxophilic Ru for facilitating oxygen insertion reaction respectively.^{10,11}

Technically, all the components needed to build the EC-NMR cell are readily available so its construction should be amenable to a generic chemistry research lab. Moreover, the NMR measurements can be carried out on any walk-in solution NMR spectrometer that should be available in most if not all research universities. We therefore expect that the demonstrated *in situ* solution EC-NMR approach should be of interest to many practitioners in the field and should provide molecular level information that complements nicely to that obtainable by *in situ* differential EC mass spectrometry, IR and Raman spectroscopy in unraveling important reaction mechanisms in electrocatalysis.

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