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Reduced roughness for improved mobility in  
benzodipyrrolidone-based, n-type OFETs†J. W. Rumer,<sup>a</sup> S. Rossbauer,<sup>b</sup> M. Planells,<sup>a</sup> S. E. Watkins,<sup>c</sup> T. D. Anthopoulos<sup>b</sup>  
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A series of six phenyl-flanked benzodipyrrolidone-based copolymers are designed and synthesised: three exhibit backbone in-plane curvature or torsional twisting, while the other three remain planar and linear. While the former may appear less crystalline by X-ray diffraction, they afford a smoother thin-film surface topology and increased electron mobility in top-gate, bottom-contact organic field-effect transistors.

## Introduction

Benzodipyrrolidones are promising building blocks for organic semiconductor materials, in part due to their innate stability and deep colour which lead to their early use as dye pigments.<sup>1</sup> For example, similar bis-lactam-based conjugated materials such as diketopyrrolopyrrole,<sup>2</sup> isoindigo<sup>3</sup> and dihydropyrroloindoleone<sup>4,5</sup> often exhibit high performance in organic field-effect transistors (OFETs)<sup>6–8</sup> and solar cells.<sup>9,10</sup> Such devices are particularly interesting due to their solution processed nature, facilitating printing of large-scale, flexible and light-weight displays.<sup>11</sup> However, recent progress in achieving ultra-high hole mobilities<sup>12,13</sup> and well balanced ambipolarity<sup>14,15</sup> garners renewed interest in electron transporting (*n*-type) organic semiconductor materials.<sup>16</sup> Moreover, as commercially viable performance targets are met,<sup>17</sup> a new focus on material stability has emerged, particularly for air and water stable transistor materials.<sup>18–20</sup>

Recently, a number of copolymers based on phenyl-flanked benzodipyrrolidone (BPP)<sup>21–25</sup> and its analogues<sup>26,27</sup> have been prepared, with charge carrier mobilities as high as 0.1 cm<sup>2</sup> V<sup>-1</sup> s<sup>-1</sup> under optimal processing conditions.<sup>28</sup> With appropriate copolymers, improved solid-state morphologies and high ionisation potentials may be realised, which can confer innate stability to electronic devices, with increased and preferential electron mobility. Indeed, such molecular design plays an important role in governing self-assembly in the solid-state, with resultant texture directly affecting OFET performance.<sup>29</sup>

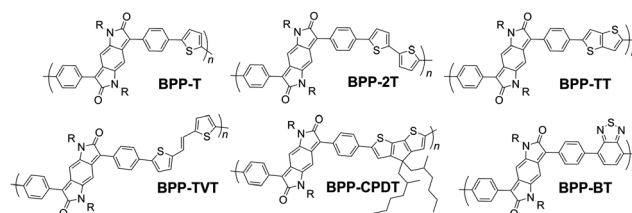


Fig. 1 The series of six BPP copolymers; R = *N*-2-decyl-tetradecane.

To this end we report a series of six BPP copolymers for systematic comparison (Fig. 1). The synthesis, morphology, thermal, optical and electronic properties and computational models of the polymers are presented. Device measurements for organic field-effect transistors (OFETs) are reported, with analysis of the film morphology by atomic force microscopy (AFM). We demonstrate that charge carrier mobility correlates to film roughness, itself being related to the conjugated polymer backbone conformation and molecular architecture.

## Results and discussion

## Synthesis

The BPP polymers were prepared in five steps from *p*-phenylenediamine (Scheme 1). Condensation with 4-bromo-DL-mandelic acid followed by ring closure in concentrated sulphuric acid furnishes the tricyclic core, reduction by sodium hydroxide and potassium persulfate and subsequent alkylation then affording the dibrominated monomer as a dark brown powder solid in 35% overall yield (four steps). Polymerisation with either a bis-boronic ester or bis-stannylated comonomer was then effected by palladium catalysed cross-coupling (Suzuki or Stille, respectively) (Scheme S1†).

In total, five alternating electron rich-poor, push-pull type BPP-based polymers were prepared (see Fig. 1 for structures),

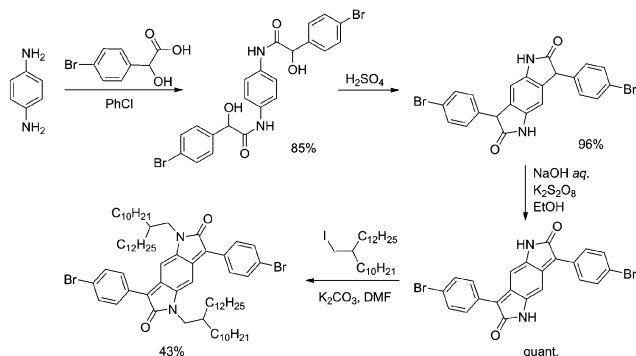
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Scheme 1 Synthesis of the dibrominated BPP monomer unit for use in Pd cat. cross-coupling polymerisations.

with the following electron-rich repeat units: thiophene (**T**), 2,2'-bithiophene (**2T**), thieno[3,2-*b*]thiophene (**TT**), (*E*)-1,2-bisthiophen-2-yl-ethene (**TVT**) and 4,4-bis(2-ethylhexyl)-4*H*-cyclopenta[2,1-*b*:3,4-*b'*]dithiophene (**CPDT**). In addition a copolymer was prepared with the electron-poor benzothiadiazole (**BT**) unit.

All **BPP**-based polymers bore the same *N*-alkyl side chain 2-decyltetradecane. The polymers were typically purified by precipitation from methanol followed by Soxhlet extraction using methanol, acetone, hexane, chloroform and, if necessary, chlorobenzene. The latter fraction(s) were then separately heated and vigorously stirred with aqueous sodium diethyldithiocarbamate to remove residual catalytic metal impurities. Lower molecular weight oligomers were readily removed in the acetone and hexane fractions, affording satisfactory molecular weights – as determined by gel permeation chromatography (GPC) – for all six polymers (see Table 1 for data and Fig. S1–S6† for GPC elution traces).

### Thermal properties & morphology

Thermal properties and thin-film morphology were probed by thermogravimetric analysis (TGA), differential scanning calorimetry (DSC) and X-ray diffraction (XRD). The decomposition temperatures (5% weight loss) were all in the range 335–395 °C (Fig. S7†). However, no obvious transitions were observed by DSC (in the range –30 to 350 °C) (Fig. S8†), with the exception of

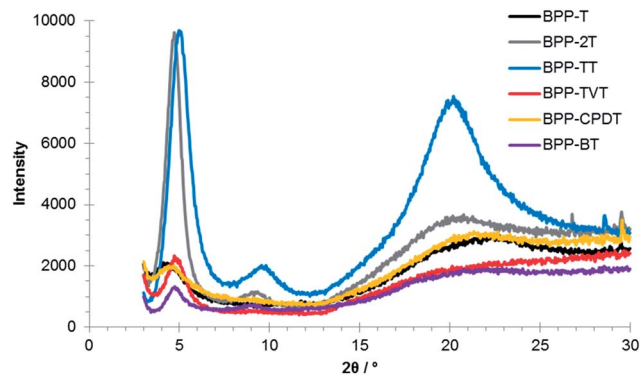


Fig. 2 X-ray diffraction (XRD) diagrams of the **BPP**-based copolymers as thin-films drop-cast on Si substrates.

**BPP-2T**, which exhibited a melt at ~230 °C and corresponding crystallisation at ~200 °C (Fig. S9†).

From X-ray diffraction experiments (Fig. 2), the polymers were all determined to be crystalline as-cast, with an out-of-plane reflection peak at 4–5°, corresponding to the (100) peak diffraction. For **BPP-2T** and **BPP-TT**, where the (100) diffraction peak is more intense, the second order reflection peaks (200) were also visible at 9–10°. The low angle diffraction peaks are characteristic of lamellar-type crystallinity in  $\pi$  stacked conjugated polymers. In addition to the (100) and (200) diffraction peaks, the (010) peak associable with co-facial  $\pi$ -stacking between polymeric backbones is distinctly observable for **BPP-TT** at ~20°.

Table 2 showed the interlayer *d*-spacing distance obtained from XRD experiments. Calculated lamellar stacking distances were ~19 Å for all polymers, expect for **BPP-TT** (21 Å) and **BPP-T** (21 Å).

The shorter *d*-spacing distance of **BPP-TT** is consistent with the increased intensity of the  $\pi$ -stacking peak, indicating a high degree of crystallinity. In comparison to **BPP-2T**, which also exhibits a high intensity lamellar packing peak, the increased  $\pi$ -stacking in **BPP-TT** may be attributed to the fused nature of the **TT** comonomer unit improving coplanarity along the polymer backbone. Moreover, the phenyl-**TT**-phenyl linkage is strictly linear in comparison to the other five comonomers which all

Table 1 Properties of the **BPP**-based copolymers

Polymer	Soxhlet fraction	$M_n^a$ (kDa)	$M_w^a$ (kDa)	PDI <sup>a</sup>	$T_{5\% \text{ wt loss}}^b$ (°C)	$T_{\text{melt}}^c$ (°C)	$\lambda_{\text{max}}^{\text{abs } d}$ (nm)		$E_g^{\text{opte}}$ (eV)	Energy levels <sup>f</sup> (eV)	
							Solution	Thin-film		$E_{\text{HOMO}}$	$E_{\text{LUMO}}$
<b>BPP-T</b>	CHCl <sub>3</sub>	25	25	1.64	386	—	602	622	1.66	–5.68	–4.02
<b>BPP-2T</b>	PhCl	17	42	2.47	393	~230	602	624	1.58	–5.63	–4.05
<b>BPP-TT</b>	CHCl <sub>3</sub>	15	51	3.41	337	—	613	631	1.61	–5.68	–4.07
<b>BPP-TVT</b>	CHCl <sub>3</sub>	20	46	2.28	391	—	622	647	1.49	–5.55	–4.06
<b>BPP-CPDT</b>	CHCl <sub>3</sub>	21	41	1.96	394	—	661	687	1.49	–5.44	–3.95
<b>BPP-BT</b>	CHCl <sub>3</sub>	54	66	1.25	376	—	543	554	1.85	–5.84	–3.99

<sup>a</sup> Determined by GPC using polystyrene standards and PhCl as the eluent at 80 °C. <sup>b</sup> Determined by thermogravimetric analysis (TGA). <sup>c</sup> As observed by differential scanning calorimetry (DSC). <sup>d</sup> Solutions (dilute) in PhCl; thin-films drop-cast from 10 mg mL<sup>–1</sup> PhCl solutions on glass substrates. <sup>e</sup> Determined from the absorption onset of the thin-film. <sup>f</sup>  $E_{\text{HOMO}}$  found by AC2 (PESA) measurement;  $E_{\text{LUMO}} = E_{\text{HOMO}} + E_g^{\text{opt}}$ .



Table 2 X-ray diffraction data for the BPP-based copolymers<sup>a</sup>

Substrate	Polymer	Lamellar packing					
		$2\theta/^\circ$	Intensity	$d$ -spacing/ $\text{\AA}$		$d$ -spacing/ $\text{\AA}$	
				$n = 1$	$2\theta/^\circ$	Intensity	$n = 2$
As cast	<b>BPP-T</b>	4.11	$2.1 \times 10^3$	21.5	—	—	—
	<b>BPP-2T</b>	4.72	$9.6 \times 10^3$	18.7	9.33	$1.2 \times 10^3$	18.9
	<b>BPP-TT</b>	4.98	$9.7 \times 10^3$	17.7	9.80	$2.0 \times 10^3$	18.0
	<b>BPP-TV</b>	4.72	$2.3 \times 10^3$	18.7	—	—	—
	<b>BPP-CPDT</b>	4.68	$2.0 \times 10^3$	18.9	—	—	—
	<b>BPP-BT</b>	4.75	$1.3 \times 10^3$	18.6	—	—	—

<sup>a</sup> Out-of-plane reflections observed by X-ray diffraction for drop-cast films (from  $10 \text{ mg mL}^{-1}$  PhCl solution) on Si substrates dried in air.

exhibit some degree of torsional or in-plane deviation. The larger lamellar  $d$ -spacing for **BPP-T** could be due to the decreased donor–acceptor spacing aspect ratio along the conjugated backbone hindering packing. However, the similar lamellar spacing for the other polymers indicates that the alkyl side-chains on the benzodipyrrolidone unit are the dominant steric factor in dictating lamellar packing interactions ( $d$ -spacing distance), as opposed to contributions from the comonomer.

### Optical & electronic properties

The optical and electronic properties of the **BPP**-based copolymers were investigated by UV-Vis absorption spectroscopy (Fig. 3) and photo-electron spectroscopy in air (PESA) and are summarised in Table 1. The thin-film absorption maxima are all in the range 620–650 nm with the exception of **BPP-BT** which is blue-shifted to  $\sim 550$  nm and **BPP-CPDT** which is red shifted to  $\sim 690$  nm. This trend can be explained by the dual-band theory for semiconducting polymers.<sup>30</sup> For the **BPP-BT** polymer both monomeric units are considered electron-deficient. For that reason, there is not a strong donor–acceptor character and the intermolecular charge transfer band (ICT) associated to the

**BPP-BT** polymer is blue shifted compared to our series. Therefore, as we introduce stronger donor moieties (**T**, **2T**, **TT** and **TVT**) the ICT band is red-shifted up to 650 nm. The far-reaching case is for **BPP-CPDT**, where the  $\lambda_{\text{max}}$  is further red-shifted to 690 nm. In this case, **CPDT** is a stronger donor due to the planarization of the bithiophene unit.<sup>31</sup> Consequentially, while the **BPP-T**, **-2T** and **-TT** polymers exhibited an effective band gap of  $\sim 1.6$  eV as determined from the thin-film absorption onset, this was increased to  $\sim 1.9$  eV for **BPP-BT** and decreased to  $\sim 1.5$  eV for **BPP-TV** and **BPP-CPDT** polymers.

In addition, the thin-film absorption maxima are typically red-shifted by some 20–25 nm in comparison to the solutions (dilute in chloroform) (Fig. S10<sup>†</sup>), indicating that solid-state interactions likely improve ordering, suppressing torsional twisting and dihedral angles, which results in increased delocalisation and subsequent red-shifting. Such orbital overlap is necessary for charge transport between conjugated backbones, with stronger interactions likely increasing mobilities in field-effect transistor devices. However, this effect is limited for **BPP-BT** which exhibits only an 11 nm red-shift from solution to solid, possibly due to the torsional twisting due to the phenyl–phenyl link inhibiting  $\pi$ -stacking. Notably the polymers are free from aggregation shoulders in both their solution and thin-film absorption profiles, implying good solubility and solution processability, and that the observed properties are inherent to the molecular architecture and interactions.

The ionisation potentials (HOMO levels) were obtained by PESA and shown in Table 1. The **BPP-BT** polymer exhibits the deepest HOMO energy level ( $-5.84$  eV) due to the inductive effect from both electron-poor **BT** and **BPP** units. The addition of electron-donating units (**T**, **2T** and **TT**) raised the HOMO energy level by 0.2 eV to approximately  $-5.65$  eV. Finally, stronger electron-donating **TVT** and **CPDT** units further raised the HOMO to  $-5.55$  and  $-5.44$  eV, respectively. The corresponding LUMO energy levels obtained by the addition of the optical band gaps to the HOMO values were in the (narrow) range of  $-3.95$  to  $-4.07$  eV, being suitably deep to permit electron injection in transistor devices. Furthermore, the deep HOMO energy levels (below  $-5.30$  eV) could indicate good stability towards unintentional doping by atmospheric species such as oxygen and water.

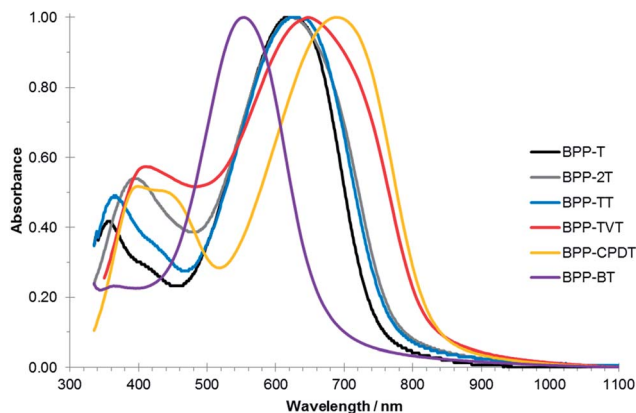


Fig. 3 Normalised UV-Vis absorption spectra of the **BPP**-based copolymers in solid-state thin-films (spin-cast from  $10 \text{ mg mL}^{-1}$  chlorobenzene solutions on glass substrates).



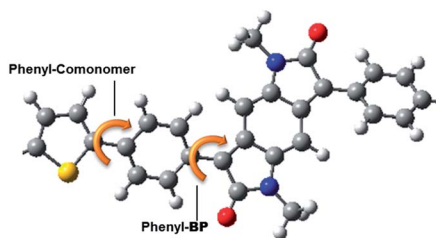


Fig. 4 The phenyl-comonomer and phenyl-BP C–C–C–C dihedral angles are indicated for a repeat unit of BPP-T.

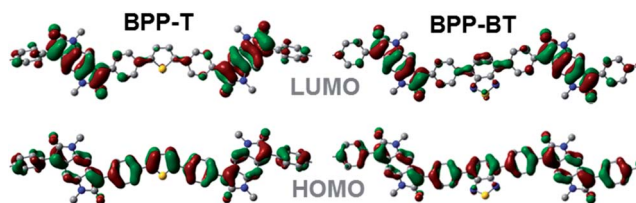


Fig. 5 Segments of the computational models of the BPP-based copolymers BPP-T and BPP-BT, showing the energy minimised structures with *N*-methyl substitution and visualisation of the HOMO and LUMO energy distributions (hydrogen atoms are omitted from view for clarity).

### Computational models

The optimized structures of the BPP-based copolymers and visualisations of the HOMO and LUMO molecular orbital distributions are shown in Fig. S11,<sup>†</sup> as calculated by density functional theory (DFT) using Gaussian 09 software at the B3LYP/6-31G\* level of theory with *N*-methyl substitution, due to the weak influence of the alkyl side-chains on the electronic structure of the conjugated core. Each polymer is modelled as a tetramer, except BPP-2T and BPP-TV<sub>2</sub>T, which are modelled as trimers due to the larger size of the systems.

The dihedral angle between the carbon atoms of the lactam rings on the central benzodipyrrolidone (BP) unit and flanking phenyl rings is found to be  $\sim 30^\circ$  in all cases (Fig. 4 and Table 3; phenyl-BP C–C–C–C). Meanwhile the dihedral angle between the carbon atoms of the phenyl rings and the comonomer (phenyl-comonomer) is  $\sim 20^\circ$  in all cases except BPP-BT, where

Table 3 Calculated dihedral angles of the BPP-based copolymers

Polymer	Dihedral angle <sup>a</sup>	
	Phenyl-comonomer	Phenyl-BP
BPP-T	21°	27°
BPP-2T	21°	28°
BPP-TT	22°	28°
BPP-TV <sub>2</sub> T	20°	27°
BPP-CPDT	19°	27°
BPP-BT	34°	30°

<sup>a</sup> Measured on the optimized structures shown in Fig. S11;<sup>†</sup> the phenyl-comonomer and phenyl-BP dihedral C–C–C–C angles are illustrated for BPP-T in Fig. 4.

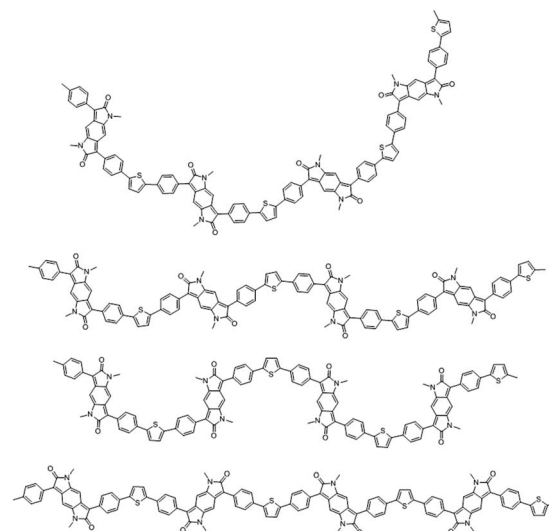


Fig. 6 Some possible chain conformations for the BPP-T polymer shown over four repeat units to illustrate the various polymer chain shapes that can arise.

the torsional twist is much greater ( $34^\circ$ ), due to the increased steric repulsion in phenyl-phenyl couples compared to phenyl-thiophene.

The HOMO molecular orbital distribution is largely delocalised over the conjugated  $\pi$ -system for all the BPP-based polymers while the LUMO is mainly located on the BPP electron deficient unit (Fig. 5).

Notably the BPP and different comonomer units have the possibility for a range of conformations, orienting in various different directions, as exemplified for BPP-T in Fig. 6. BPP-T is able to theoretically adopt such a range of conformations as the thiophene comonomer does not afford a co-linear connection to its neighbouring BPP units. Instead, the thiophene introduces an in-plane kink angle to the conjugation pathway of the polymer backbone. The same is true of the CPDT comonomer. Allowing for perpendicular bonds to the neighbouring units,

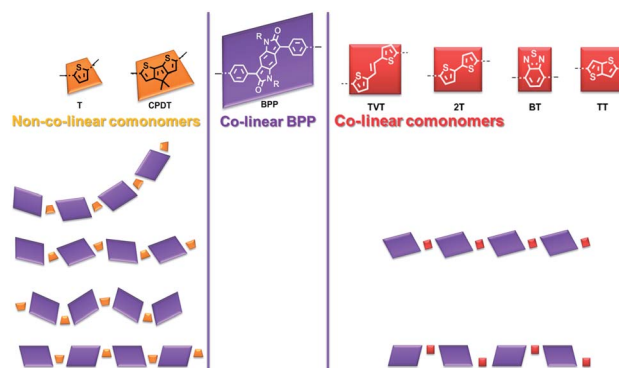


Fig. 7 The BPP comonomers can be modelled as either trapeziums which introduce an in-plane kink angle to the conjugation pathway of the polymer backbone (non-co-linear), or rectangles which afford a co-linear join. The possible in-plane conformations of the BPP-based polymers over four repeat units are shown for the non-co-linear and co-linear comonomers.



the six comonomers can be crudely modelled as trapeziums (**T** and **CPDT**) or rectangles (**TVT**, **2T**, **BP**, **TT**), the latter of four units being co-linear (Fig. 7). The number of different theoretically possible conformations increases exponentially with the number of repeat units considered. However, even with just four repeat units, a simple but meaningful model can be considered. In all cases an overall linear conformation of the conjugated polymer backbone is possible. However, when the comonomer is non-co-linear and introduces a kink (**T** and **CPDT**) a curved shape can also evolve along the conjugation pathway in the plane of the backbone, which may hamper solid-state ordering.

In all computational models the polymers have been modelled in the likely lowest energy state, which would allow the alkyl side-chains to point out into free space, reducing steric clashes. For **BPP-T** and **BPP-CPDT** this is an overall linear conformation with the benzodipyrrolidone units alternately 'pointing' in different directions, while for the other four polymers this is an overall linear conformation where the benzodipyrrolidone unit consistently orients in the same direction.

While in solution the polymer may be able, with sufficient energy, to transition between different shapes, the conformation will be kinetically frozen on casting into the solid-state. With appropriate casting conditions, such as high boiling point solvents or additives, the polymer may be able to settle into a thermodynamically preferred low energy orientation that affords ordered solid-state packing, as indicated by X-ray diffraction.

### Organic field-effect transistor (OFET) device performance

Organic field-effect transistors (OFETs) were prepared with the **BPP**-based polymers. A top-gate, bottom-contact device architecture (see Fig. S13†) was used, fabricated by spin-coating polymer solution (5 mg mL<sup>-1</sup> in hot [75 °C] *ortho*-dichlorobenzene) and drying under a nitrogen atmosphere at 100 °C. In preliminary testing annealing at 200 °C lead to worse performance (as seen for similar **BPP**-based polymers),<sup>22</sup> as did bottom-gate, bottom-contact devices, exhibiting large hysteresis and poor performance. The top-gate devices consisted of Cytop dielectric and Al/Au electrodes on glass substrates. Notably these devices only exhibited electron transport for all six **BPP**-based polymers, with moderate threshold voltages (30–65 V) and satisfactory  $I_{\text{on/off}}$  ratios all  $>10^3$  (Table 4). The output and transfer curves are shown in Fig. 8, S14 and S15.†

The **BPP**-based polymer mobilities are of the magnitude  $\times 10^{-2}$  cm<sup>2</sup> V<sup>-1</sup> s<sup>-1</sup> (**BPP-2T**, **-TT** and **-TVT**) or  $\times 10^{-3}$  cm<sup>2</sup> V<sup>-1</sup> s<sup>-1</sup> (**BPP-T**, **-CPDT** and **-BT**), independent of threshold voltage or  $I_{\text{on/off}}$  ratio. To further investigate this trend the OFET device morphology was probed by atomic force microscopy (AFM) imaging of polymer films prepared in an identical fashion (Fig. S16†).<sup>32</sup> There is a clear correlation to root mean square (RMS) roughness of the films, with the lower mobility polymers exhibiting a roughness at least double that of the higher mobility polymers; for example **BPP-CPDT** has the lowest RMS roughness at 0.40 nm with a corresponding highest mobility of 0.01 cm<sup>2</sup> V<sup>-1</sup> s<sup>-1</sup> whilst **BPP-TVT** exhibits the highest RMS

Table 4 Electron mobilities ( $\mu_e$  [cm<sup>2</sup> V<sup>-1</sup> s<sup>-1</sup>]), threshold voltages ( $V_{\text{th}}$  [V]),  $I_{\text{on/off}}$  ratios and RMS roughness (nm) of the **BPP**-based copolymer OFET devices<sup>a</sup>

Polymer	$V_{\text{th}}$ (V)	$\mu_e$ (cm <sup>2</sup> V <sup>-1</sup> s <sup>-1</sup> )	$I_{\text{on/off}}$	RMS roughness (nm)
<b>BPP-T</b>	60 ± 5	0.007	4 ± 1 × 10 <sup>3</sup>	0.54
<b>BPP-2T</b>	49 ± 3	0.001	2 ± 1 × 10 <sup>3</sup>	1.37
<b>BPP-TT</b>	34 ± 1	0.001	1 ± 0.5 × 10 <sup>3</sup>	1.63
<b>BPP-TVT</b>	59 ± 7	0.002	1 ± 0.6 × 10 <sup>3</sup>	1.82
<b>BPP-CPDT</b>	47 ± 1	0.01	4 ± 2 × 10 <sup>3</sup>	0.40
<b>BPP-BT</b>	43 ± 1	0.01	8 ± 3 × 10 <sup>3</sup>	0.68

<sup>a</sup> Top-gate, bottom-contact architecture devices annealed at 100 °C. Mobility values are accurate to one significant figure.

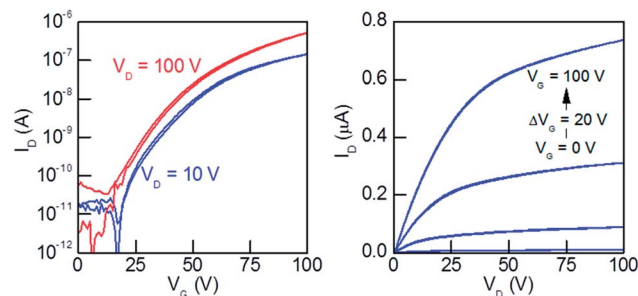


Fig. 8 The **BPP-BT** copolymer OFET device characteristics: (left) the output curves and (right) the transfer curves; (top-gate, bottom-contact architecture; annealed at 100 °C). Of the three higher electron mobility polymers **BPP-BT** also exhibited the highest  $I_{\text{on/off}}$  ratio and lowest threshold voltage.

roughness at 1.82 nm, accompanied by a lower mobility of 0.002 cm<sup>2</sup> V<sup>-1</sup> s<sup>-1</sup> (Fig. 9). However in all cases the roughness is well below what is considered acceptable for OFET devices (up to 25 nm),<sup>8,18</sup> indicating that the trend is attributable to the polymer structure, solid-state interactions and morphology, as opposed to poor processing or fabrication technique.

Interestingly the mobility and roughness is apparently related to the crystallinity and potential ability of the polymer to adopt different conformations, energetically accessible in hot polymer solutions, which may then be kinetically frozen on casting the solid-state film. Both **BPP-T** and **BPP-CPDT** allow for curved or kinked polymer chain shapes (Fig. 7) whilst **BPP-BT** has a greater degree of torsional twisting along the backbone (Fig. 4 and Table 3). In addition, **BPP-2T** exhibits the highest degree of crystallinity with an observable melt by DSC and the most intense low angle XRD peak.

This indicates that while crystallinity in conjugated polymer systems can be beneficial for charge transport,<sup>33</sup>  $\pi$ -stacking is much more important than lamellar packing, as expected from the nature of charge hopping between conjugated backbones *via* delocalised orbitals and their overlap.<sup>34</sup> Polymers that could theoretically adopt a greater range of conjugated backbone shapes, achieved through either out-of-plane twisting or non-co-linearity with the conjugation pathway in this case, are able to



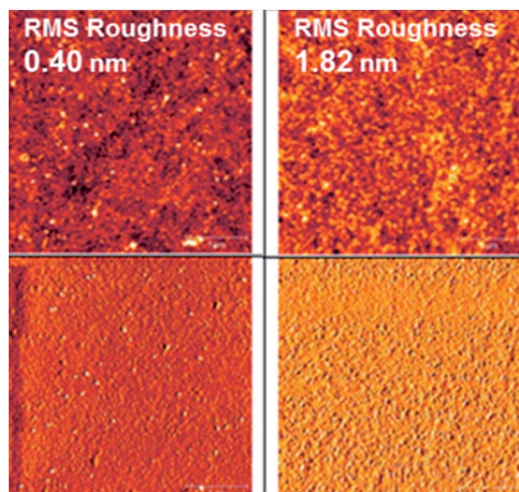


Fig. 9 Atomic force microscopy (AFM) images of (left) the BPP-CPDT and (right) BPP-TVt thin-films to show the highest contrast in film roughness; (top) topography and (bottom) phase images, both  $4 \times 4 \mu\text{m}$ .

achieve smoother topologies than planar and co-linear highly crystalline polymers. While high crystallinity may result in grain boundary defects, the ability of the polymer to coexist in amorphous regions, results in a reduced surface roughness, likely due to denser packing, giving improved charge carrier mobilities.<sup>35</sup>

## Conclusions

In summary a series of six phenyl-flanked benzodipyrrolidone (BPP)-based copolymers were designed and synthesised, all bearing *N*-2-decyl-tetradecyl alkyl side-chains for high solubility in common organic solvents and solution processing. The polymers show good thermal stability, as-cast crystallinity, medium optical band gaps and deep HOMO and LUMO energy levels. Conjugated backbone planarity depends on the comonomer (BPP-BT exhibiting greater out-of-plane twisting), with all polymers exhibiting only electron conduction in top-gate, bottom-contact organic field-effect transistor (OFET) devices.

Charge carrier mobility was observed to be dependent on polymer film surface roughness. Polymers that could theoretically exhibit a greater range of backbone shapes (e.g. curved or twisted: BPP-T, -CPDT and -BT), afford smoother films (RMS roughness' <1 nm) and exhibit mobilities an order of magnitude greater than their strictly linear counterparts (BPP-2T, -TT and -TVT; RMS roughness >1 nm).

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## Notes and references

- C. W. Greenhalgh, J. L. Carey and D. F. Newton, *Dyes Pigm.*, 1980, **1**, 103.
- Y. Li, P. Sonar, L. Murphy and W. Hong, *Energy Environ. Sci.*, 2013, **6**, 1684.
- R. Stalder, J. Mei, K. R. Graham, L. A. Estrada and J. R. Reynolds, *Chem. Mater.*, 2013, **26**, 664.
- J. W. Rumer, S.-Y. Dai, M. Levick, Y. Kim, M.-B. Madec, R. S. Ashraf, Z. Huang, S. Rossbauer, B. Schroeder, L. Biniek, S. E. Watkins, T. D. Anthopoulos, R. A. J. Janssen, J. R. Durrant, D. J. Procter and I. McCulloch, *J. Mater. Chem. C*, 2013, **1**, 2711.
- J. W. Rumer, S.-Y. Dai, M. Levick, L. Biniek, D. J. Procter and I. McCulloch, *J. Polym. Sci., Part A: Polym. Chem.*, 2013, **51**, 1285.
- S. Holliday, J. E. Donaghey and I. McCulloch, *Chem. Mater.*, 2014, **26**, 647.
- M. J. Robb, S.-Y. Ku, F. G. Brunetti and C. J. Hawker, *J. Polym. Sci., Part A: Polym. Chem.*, 2013, **51**, 1263.
- B. Sun, W. Hong, Z. Yan, H. Aziz and Y. Li, *Adv. Mater.*, 2014, **26**, 2636.
- J. W. Rumer, B. C. Schroeder, C. B. Nielsen, R. S. Ashraf, D. Beatrup, H. Bronstein, S. J. Cryer, J. E. Donaghey, S. Holliday, M. Hurhangee, D. I. James, S. Lim, I. Meager, W. Zhang and I. McCulloch, *Thin Solid Films*, 2013, **560**, 82.
- L. Dou, J. Gao, E. Richard, J. You, C.-C. Chen, K. C. Cha, Y. He, G. Li and Y. Yang, *J. Am. Chem. Soc.*, 2012, **134**, 10071.
- H. Dong, X. Fu, J. Liu, Z. Wang and W. Hu, *Adv. Mater.*, 2013, **25**, 6158.
- Y. Yuan, G. Giri, A. L. Ayzner, A. P. Zoombelt, S. C. B. Mannsfeld, J. Chen, D. Nordlund, M. F. Toney, J. Huang and Z. Bao, *Nat. Commun.*, 2014, **5**, 3005.
- C. Luo, A. K. K. Kyaw, L. A. Perez, S. Patel, M. Wang, B. Grimm, G. C. Bazan, E. J. Kramer and A. J. Heeger, *Nano Lett.*, 2014, **14**, 2764.
- A. R. Mohebbi, J. Yuen, J. Fan, C. Munoz, M. F. Wang, R. S. Shirazi, J. Seifert and F. Wudl, *Adv. Mater.*, 2011, **23**, 4644.
- J. W. Rumer, M. Levick, S.-Y. Dai, S. Rossbauer, Z. Huang, L. Biniek, T. D. Anthopoulos, J. R. Durrant, D. J. Procter and I. McCulloch, *Chem. Commun.*, 2013, **49**, 4465.
- Y. Zhao, Y. Guo and Y. Liu, *Adv. Mater.*, 2013, **25**, 5372.
- H. Sirringhaus, *Adv. Mater.*, 2014, **26**, 1319.
- O. Knopfmacher, M. L. Hammock, A. L. Appleton, G. Schwartz, J. Mei, T. Lei, J. Pei and Z. Bao, *Nat. Commun.*, 2014, **5**, 2954.
- M. Yun, A. Sharma, C. Fuentes-Hernandez, D. K. Hwang, A. Dindar, S. Singh, S. Choi and B. Kippelen, *ACS Appl. Mater. Interfaces*, 2014, **6**, 1616.
- Y. Ie, M. Ueta, M. Nitani, N. Tohnai, M. Miyata, H. Tada and Y. Aso, *Chem. Mater.*, 2012, **24**, 3285.
- W. Cui, J. Yuen and F. Wudl, *Macromolecules*, 2011, **44**, 7869.
- K. C. Lee, W.-T. Park, Y.-Y. Noh and C. Yang, *ACS Appl. Mater. Interfaces*, 2014, **6**, 4872.



- 23 W. Yue, X. Huang, J. Yuan, W. Ma, F. C. Krebs and D. Yu, *J. Mater. Chem. A*, 2013, **1**, 10116.
- 24 W. Hong, C. Guo, Y. Li, Y. Zheng, C. Huang, S. Lu and A. Facchetti, *J. Mater. Chem.*, 2012, **22**, 22282.
- 25 P. Deng, L. Liu, S. Ren, H. Li and Q. Zhang, *Chem. Commun.*, 2012, **48**, 6960.
- 26 W. Hong, B. Sun, C. Guo, J. Yuen, Y. Li, S. Lu, C. Huang and A. Facchetti, *Chem. Commun.*, 2013, **49**, 484.
- 27 W. Hong, C. Guo, B. Sun, Z. Yan, C. Huang, Y. Hu, Y. Zheng, A. Facchetti and Y. Li, *J. Mater. Chem. C*, 2013, **1**, 5624.
- 28 G. W. P. van Puijsschen, E. A. Pidko, M. M. Wienk and R. A. J. Janssen, *J. Mater. Chem. C*, 2014, **2**, 731.
- 29 J. W. Ward, R. Li, A. Obaid, M. M. Payne, D.-M. Smilgies, J. E. Anthony, A. Amassian and O. D. Jurchescu, *Adv. Funct. Mater.*, 2014, **24**, 5052.
- 30 P. M. Beaujuge, C. M. Amb and J. R. Reynolds, *Acc. Chem. Res.*, 2010, **43**, 1396.
- 31 S. Barlow, S. A. Odom, K. Lancaster, Y. A. Getmanenko, R. Mason, V. Coropceanu, J.-L. Brédas and S. R. Marder, *J. Phys. Chem. B*, 2010, **114**, 14397.
- 32 Y. Wen, Y. Liu, Y. Guo, G. Yu and W. Hu, *Chem. Rev.*, 2011, **111**, 3358.
- 33 A. Bagui and S. Sundar Kumar Iyer, *Org. Electron.*, 2014, **15**, 1387.
- 34 C. Deibel and V. Dyakonov, *Rep. Prog. Phys.*, 2010, **73**, 096401.
- 35 T. Qin, W. Zajaczkowski, W. Pisula, M. Baumgarten, M. Chen, M. Gao, G. Wilson, C. D. Easton, K. Müllen and S. E. Watkins, *J. Am. Chem. Soc.*, 2014, **136**, 6049.

