

Hexagonal closest-packed spheres liquid crystalline phases stabilised by strongly hydrated counterions†

Connie K. Liu and Gregory G. Warr*

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The sequence and structure of lyotropic liquid crystals formed in C₁₂–C₁₆ alkyltrimethylammonium surfactants with hydrolysable and multivalent phosphate (PO₄^{3−}, HPO₄^{2−} and H₂PO₄[−]), oxalate (HC₂O₄[−] and C₂O₄^{2−}), and carbonate (HCO₃[−]/CO₃^{2−}) counterions were determined using a concentration gradient method coupled with polarising optical microscopy and small angle X-ray scattering. In addition to the discrete cubic (I₁, space group *Pm3n*) and hexagonal (H₁, *p6m*) phases, almost all of these surfactants also formed the (previously) rare hexagonally closest-packed spheres (HCPS, *P6₃/mmc*) phase at compositions between the *Pm3n* cubic and L₁ micellar phases. This structure has not been previously observed in cationic surfactants, but is readily achieved by using strongly hydrated counterions to stabilise spherical micelles at high concentrations.

Introduction

The structure of liquid crystalline phases formed in aqueous solutions by surfactant self-assembly above the critical micelle concentration (CMC), depends on the packing of the surfactant molecules into micelles. The hydrocarbon chain length (*l_c*) and volume (*v*), together with the effective head group area (*a₀*) determine the value of the packing parameter, *v/a₀l_c*, and the propensity of a surfactant to form spherical micelles (*v/a₀l_c* ≤ 1/3), cylinders (1/3 < *v/a₀l_c* ≤ 1/2), or bilayers (1/2 < *v/a₀l_c* ≤ 1) under different solution conditions.¹ As the surfactant concentration increases, self-assembly that usually begins with spherical micelles evolves into structures with lower curvature. Micelles adopt liquid-like order in isotropic solution (L₁) transition first into discrete cubic (I₁), then hexagonal (H₁), bicontinuous cubic (V₁) and lamellar (L_α) phases, as typified by the well-known behavior of dodecyltrimethylammonium chloride (DTAC).²

The optically isotropic, “discrete micellar” cubic phase is found at compositions between the micellar solution and hexagonal phase. Here micelles are usually arranged into 3D cubic symmetries including body-centered cubic (BCC, *Im3m*), face-centered cubic (FCC, *Fm3m*, *Fd3m*) and tetrahedrally close-packed (*Pm3n*) structures. *Pm3n* is by far the most common of these phases, and since it was first reported in DTAC² it has been found in numerous other cationic,³ anionic,⁴ ternary,⁵ non-ionic surfactant^{6,7} and phospholipid^{8,9} systems.¹⁰ Also

sometimes counted as an I₁ phase is the hexagonally close-packed spheres (HCPS, *P6₃/mmc*) structures, however this is only rarely observed, and then often classified separately.^{11–13}

The discrete micellar cubic phase is something of a puzzle. Often described in texts as an ordered array of spherical micelles, it curiously does not have either the fcc or hcp symmetry of a closest-packed lattice. Three structures have been proposed for the *Pm3n* cubic phase, as illustrated in Fig. 1. Two of these suggest that at least some of the micelles on the cubic lattice are non-spherical. From NMR studies on the *Pm3n* cubic phase formed in the palmitoylsophosphatidylcholine (PLPC)-water system, Fontell *et al.* and Eriksson *et al.* proposed a unit cell of 8 identical, slightly elongated micelles where those on the cube vertices are able to rotate freely in all directions while those on the cube faces have only one axis of rotation.^{14,15} From X-ray diffraction, freeze fracture electron microscopy and electron density mapping of the *Pm3n* cubic phase of DTAC, Vargas, Luzzati, and Delacroix proposed a unit cell containing two spherical and 6 disc shaped micelles.^{9,16} As the NMR results for PLPC yield a time-average structure close to the DTAC X-ray structure, it is possible that these two non-spherical micelle models are essentially equivalent (see Fig. 1).

The third explanation of the observed symmetry, due to Imai *et al.* and based on studies of the C₁₂EO₈-water system,¹⁷ is that the spherical micelles in this phase pack so as to minimise the total interfacial contact area between neighbours, rather than to maximise packing density (whence fcc or hcp). This effect may be specific to the compression of the hydrophilic chains of polyoxyethylene in the micelle corona of these nonionic surfactants.

In contrast, the hexagonally closest-packed spheres phase (HCPS), which might be expected on packing density grounds, is rare. First identified in the binary non-ionic surfactant,

School of Chemistry F11, The University of Sydney, NSW, 2006 Australia. E-mail: gregory.warr@sydney.edu.au; Fax: +612 9351 3329; Tel: +612 9351 2106

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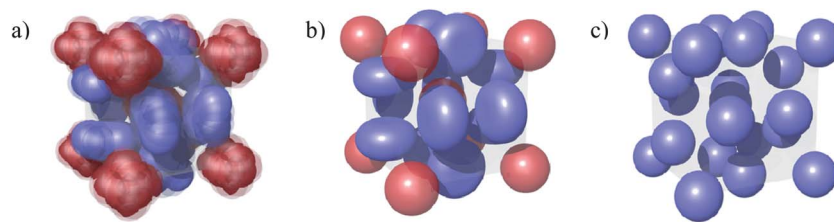


Fig. 1 Comparison of proposed $Pm3n$ unit cell structures comprised of: (a) 8 identical, slightly elongated rods illustrating free rotation of (red) micelles at the corner and in the centre of the unit cell and restricted rotation of (blue) micelles on its faces;^{14,15} (b) 2 spherical (red) and 6 disc-shaped (blue) micelles;^{9,16} (c) 8 identical spherical micelles.¹⁷

$C_{12}EO_8$ –water system at 31.5–39 wt% surfactant, and below 18 °C⁶ electron density mapping has confirmed that micelles are spheres with their polar ethylene oxide moieties in nearest-neighbour contact.¹² Since then, the HCPS phase has only been found in one other binary system, comprising anionic choline carboxylate–water mixtures in a phase sequence that included the $Pm3n$ cubic, hexagonal, bicontinuous cubic and lamellar phases at higher concentrations.¹⁸

HCPS has also been observed in two other, multicomponent systems. Soni *et al.* identified the HCPS phase in the triblock copolymer system poly(ethyleneoxide)₂₀–poly(propylene oxide)₇₀–poly(ethyleneoxide)₂₀, (Pluronic® P123) with ethanol and water⁵ and also recently, inverse micelles adopting the HCPS arrangement was found by Sherman *et al.* in a mixture of dioleoylphosphatidylcholine, dioleoylglycerol, cholesterol and water.¹⁹ HCPS has also been reported as surface structures in silica films formed from the cationic surfactant CTAB,²⁰ and in copolymer gels of methacrylate (MAA) and *N*-isopropylacrylamide (NIPAM), P(MAA/NIPAM), associated with tetradecyl- and dodecyltrimethylammonium surfactant cations.²¹

The primary surfactant packing condition that favours the formation of the HCPS phase over $Pm3n$ (or hexagonal) structures is thus the preservation of spherical micelles at high surfactant concentration. As a secondary condition, steric interactions, or any interactions which might favour minimizing contact area between neighbouring micelles, are best avoided. This suggested to us that ionic surfactants with weakly binding counterions which would resist sphere-to-cylinder transformation by added electrolyte should stabilize the HCPS phase. We have recently shown²² that resiliently-spherical micelles are formed by cationic surfactants with hydrolysable phosphate (PO_4^{3-} , HPO_4^{2-} and $H_2PO_4^-$), oxalate ($HC_2O_4^-$ and $C_2O_4^{2-}$) and carbonate (HCO_3^-/CO_3^{2-}) counterions. Although weakly binding in the monovalent state, these anions remain strongly hydrated even when hydrolysed into their divalent or trivalent forms and are strongly electrostatically bound to the cationic micelles, and exhibit negligible micelle growth even at high common-ion electrolyte concentrations. This makes them excellent candidates for formation of HCPS phases.

Materials and methods

Alkyltrimethylammonium surfactants with hydrolysable carbonate, phosphate and oxalate counterions were prepared and characterized as described previously.²² These solutions

will undergo some degree of anion speciation by hydrolysis, but are identified throughout according to the parent salt *e.g.* (DTA)₃PO₄, DTAHC₂O₄, *etc.* Although prepared as carbonate salts, these surfactants exist primarily as bicarbonates in solution due to atmospheric CO₂. Polarizing optical microscopy and small angle X-ray scattering (SAXS) were both used to identify the sequence and structure of lyotropic phases formed in binary surfactant–water mixtures.

In polarising optical microscopy, a Leica DM LB microscope fitted with cross polarisers, quarter-wave plate and a DC 300 digital camera was used. All experiments were performed at room temperature (23 °C). Concentration gradients were prepared by placing a small amount of solid surfactant sample onto a microscope slide, a cover-slip pressed on top and a few drops of water are added under the edge of the cover-slip. Water is drawn under the cover-slip through capillary action towards the surfactant sample creating a concentration gradient forming the full sequence of phases present in the surfactant–water mixture. Phases form separate bands which are distinguished by refractive index discontinuities. The phases can then be identified by their viscosity, birefringence and characteristic optical textures.^{11,23,24}

Small angle X-ray scattering experiments were performed on the SAXS/WAXS beamline at The Australian Synchrotron, Victoria, Australia. The high intensity of a synchrotron X-ray beam allows for much higher resolution of scattering peaks and the detection of weak reflections. Concentration gradient samples were prepared by packing solid surfactant into the bottom of a 1.5 mm diameter glass capillary (Charles Supper Company). Water was then added and allowed to penetrate the surfactant and establish a concentration gradient over a 12–24 hour period prior to being scanned with the synchrotron X-ray beam. This experimental set up was inspired by a similar methodology developed by Ricoul *et al.* in the determination of ternary phase diagrams of cationic surfactant–glycolipid–water systems.²⁵ The entire length of the capillary containing the sample was scanned at 0.80 mm (vertical) increments with synchrotron X-rays ($\lambda = 0.6888$ Å or 1.0332 Å). Typically, 40 points along the capillary was scanned in order to ensure that all the phases formed were sampled. The scattered X-rays were collected on a Dectris-Pilatus 1 M detector over a q -range of 0.008–0.7 Å^{−1}. Spectra were collected for 10 or 30 s and measurements were performed at ambient temperature in the instrument hutch, 27 °C. The resulting characteristic diffraction patterns were indexed and the space group symmetry and consequently the structure of liquid crystalline phases are identified.^{6,11}



Individual SAXS patterns are obtained for discrete points along a similarly-established gradient but within a sample capillary, allowing each phase formed to be examined. However, as phase boundaries formed within a capillary are not necessarily horizontal, and the beam itself is 0.250 mm (horizontal) \times 0.150 mm (vertical) (FWHM), two adjacent phases were sometimes sampled simultaneously. Phases that occur at a narrow concentration range may be completely missed by the scanning synchrotron X-ray beam. Thus, all SAXS scans are interpreted in conjunction with the corresponding polarising optical microscopy image. The complete set of polarising optical microscopy flooding experiments may be found in the ESI†

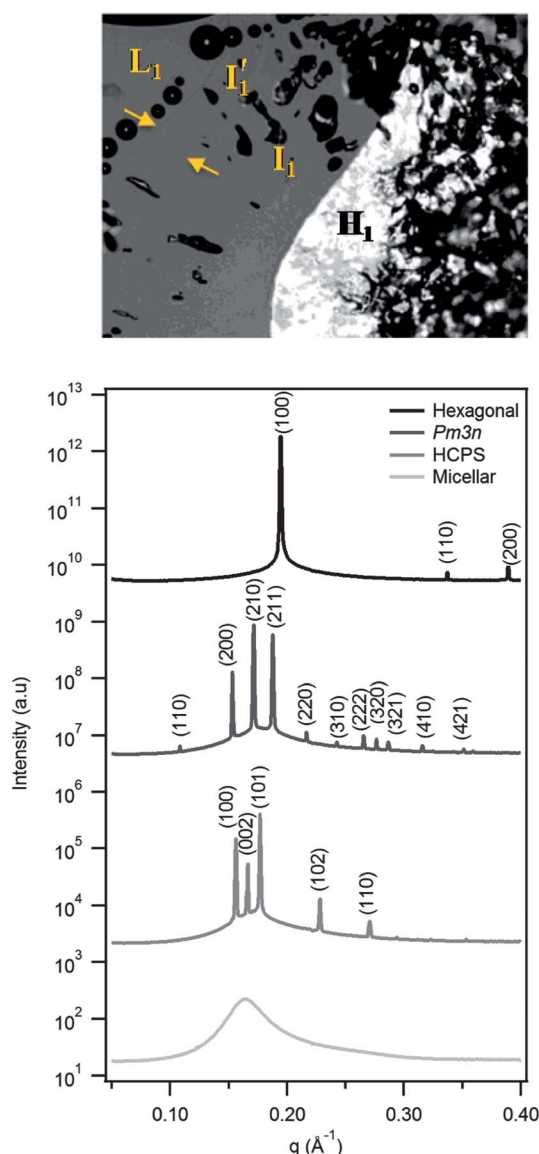


Fig. 2 Polarising optical microscopy and corresponding synchrotron SAXS diffraction patterns showing the phase sequence for DTA_2CO_3 in a concentration gradient scan. SAXS patterns show increasing concentration from bottom to top. The phases and symmetry/space group are identified as micelles (L_1), I_1' HCPS ($P6_3/mmc$), I_1 cubic ($Pm3n$) and H_1 hexagonal ($p6m$). Curves are offset for clarity.

Results and discussion

Fig. 2 shows a representative polarising optical micrograph of a concentration gradient experiment on DTA_2CO_3 , together with synchrotron SAXS patterns of a similar gradient sample. The two adjacent isotropic lyotropic phases denoted I_1' and I_1 are separated by a refractive index discontinuity. The corresponding SAXS patterns, from bottom to top, confirm the structure sequence as micellar solution, then HCPS ($P6_3/mmc$), discrete cubic ($Pm3n$) and hexagonal phase ($p6m$), and confirm the existence of two distinct I_1 structures; the more dilute corresponds to HCPS and the more concentrated to $Pm3n$. This sequence is unusual, and few systems with the adjacent discrete micellar phases have been found.^{6,7,18,26,27}

In fact all dodecyl-, tetradecyl- and hexadecyl-trimethylammonium surfactants with hydrolysable phosphate, oxalate and carbonate counterions (except CTAHC_2O_4) showed two such adjacent, I_1 phases no matter what the hydrolysis state of the counterion. (See ESI† for complete set of polarizing optical micrographs.) Table 1 summarises the phases observed, including comparative literature data for bromide, chloride and sulfate counterions. For the alkyltrimethylammonium halides and cetyltrimethylammonium sulfate (CTA_2SO_4), bicontinuous

Table 1 Summary of the liquid crystalline phases formed in alkyl-trimethylammonium surfactants various counterions. The phases identified in order of increasing concentration are isotropic (micellar), L_1 , isotropic (hexagonal micellar) HCPS, isotropic (cubic), $Pm3n$, and hexagonal, H_1 . The presence of phases detected by both polarising optical microscopy and SAXS is denoted with "X" while phases observed by microscopy only are indicated with "M"^a

Surfactant series	Counterion	L_1	HCPS	$Pm3n$	H_1
DTA ⁺ (C_{12})	Br^-	X ¹	—	—	X ¹
	Cl^-	M ²	—	M ² , X ³	M ² , X ⁴
Phosphates	PO_4^{3-}	X	X	X	X
	HPO_4^{2-}	X	X	X	X
	H_2PO_4^-	X	X	X	X
Oxalates	$\text{C}_2\text{O}_4^{2-}$	X	X	X	X
	HC_2O_4^-	X	M	X	X
Carbonate	HCO_3^-	X	X	X	X
	TTA ⁺ (C_{14})	Br^-	—	—	M ⁵
Phosphates	Cl^-	M ²	—	M ^{2b}	M ²
	PO_4^{3-}	X	X	X	X
	HPO_4^{2-}	X	M	X	X
Oxalates	H_2PO_4^-	M	M	M	M
	$\text{C}_2\text{O}_4^{2-}$	X	M	X	M
	HC_2O_4^-	X	M	X	X
Carbonate	HCO_3^-	X	X	X	X
	CTA ⁺ (C_{16})	Br^-	—	—	X ⁶
	Cl^-	X ⁴	—	—	X ⁴
Phosphates	SO_4^{2-}	M ⁷	—	M ^{7b}	M ⁷
	PO_4^{3-}	X	M	X	M
	HPO_4^{2-}	X	M	X	M
Oxalates	H_2PO_4^-	X	M	X	X
	$\text{C}_2\text{O}_4^{2-}$	X	X	X	M
	HC_2O_4^-	X	—	X	X
Carbonate	HCO_3^-	X	M	M	X

^a [1] ref. 32, [2] ref. 24, [3] ref. 2, [4] ref. 28, [5] ref. 29, [6] ref. 31, [7] ref. 30. ^b Cubic phase identified but symmetry not confirmed (*i.e.* not necessarily $Pm3n$).



cubic, intermediate and lamellar phases have also been reported at higher temperatures and surfactant concentrations above the hexagonal phase.^{24,28–32}

Fig. 3 shows the characteristic SAXS patterns of HCPS phases found in the three DTA⁺ phosphate surfactants, irrespective of the hydrolysis state of the counterion. For all the phosphate surfactants examined, both the HCPS and *Pm3n* discrete cubic phases are preserved as the chain length is increased from C₁₂ to C₁₆. The HCPS phase appears to form over a narrower composition range with longer alkyl chains however, as the X-ray scans either missed or showed overlapping SAXS patterns in this region in the majority of samples (see ESI† for an example).

In the HCPS phase, the constituent micelles are closest-packed, identical spheres.^{12,17} Its widespread occurrence in these surfactants is consistent with our recent small-angle neutron scattering investigation of the L₁ phase of these systems, which showed the spherical micelles to be extremely stable and resistant to a sphere-to-rod transformation.²²

Specific counterion effects are one of the best known ways of controlling surfactant self-assembly structure. Changing counterions changes the effective head group area a_0 and therefore the surfactant packing parameter, v/a_0l_c . Conventionally, weaker counterion binding is associated with larger a_0 , and favours

spherical micelles. This is exemplified by comparing chloride counterions with the more strongly binding bromide in DTAC² versus DTAB.³² By binding more strongly (and being less hydrated at the micelle surface), bromide more effectively screens repulsions between cationic headgroups.^{33–35} Thus it reduces a_0 and increases the surfactant packing parameter, favouring the formation of rod-like micelles. Consequently, DTAB forms a hexagonal phase adjacent to its micellar phase at room temperature while DTAC first forms the *Pm3n* cubic phase.

An unusual feature of the counterions in the present study is that they can be strongly binding in their multivalent hydrolysis state(s), as measured by the residual micellar charge from *e.g.* conductivity.²² Both phosphate and oxalate ions are, however, known to remain strongly hydrated and to have large radii of hydration.^{36–39} Thus, despite binding strongly to micelles, these ions are unable to effectively screen the repulsions between neighbouring headgroups within a micelle. This results in the preservation of spherical micelles with a large head group area, a_0 . The formation of HCPS phases in all DTA⁺ phosphates, oxalates and bicarbonate, but not in DTAC, underscores that these counterions at all hydrolysis states behave like a more weakly binding ion than chloride. This is further highlighted by the persistence of the HCPS phase as the alkyl chain length is increased from C₁₂ to C₁₆, where even the *Pm3n* cubic phase is no longer formed by the chloride salt.

Only in CTAHC₂O₄ was the HCPS phase not detected. This suggests that the monovalent HC₂O₄[−] more effectively screens neighbouring head groups than its divalent C₂O₄^{2−} form. As with the phosphates, dilute solution measurements have shown that C₂O₄^{2−} neutralizes the micelle charge more effectively than HC₂O₄[−], meaning that hydration of C₂O₄^{2−} must be the determining factor.²² Para and Warszynski have shown in a related system that HSO₄[−] binds more strongly to a surface than SO₄^{2−}, as the divalent ions are more strongly hydrated.⁴⁰

The almost ubiquitous observation of small, spherical micelles in both the L₁ and HCPS phases of these cationic surfactants, together with the recurring phase progression on to I₁ (*Pm3n*) and H₁ suggests to us that curvature changes drive changes in self-assembly structure. It is harder to rationalize a change in spherical micelle packing from maximum density in HCPS to area minimization in I₁ at higher concentration, as proposed by Imai.¹⁷

In these systems the decreased availability of water to hydrate the counterions with increasing surfactant concentration may be the cause of changes in counterion binding, and of the accompanying reduction in a_0 that leads to the formation of less curved aggregates. With increasing concentration we would thus expect a transition from spheres to aspherical micelles (either short rods or disks, as in Fig. 1(a) and (b)) in the *Pm3n* cubic phase, before elongated cylinders form in H₁, as is observed.

Conclusions

The rare hexagonally close-packed spherical micelles (HCPS, *P6₃/mmc* space group) lyotropic liquid crystalline phase is not really so rare. It can be commonly found in alkyl-trimethylammonium surfactants associated with strongly

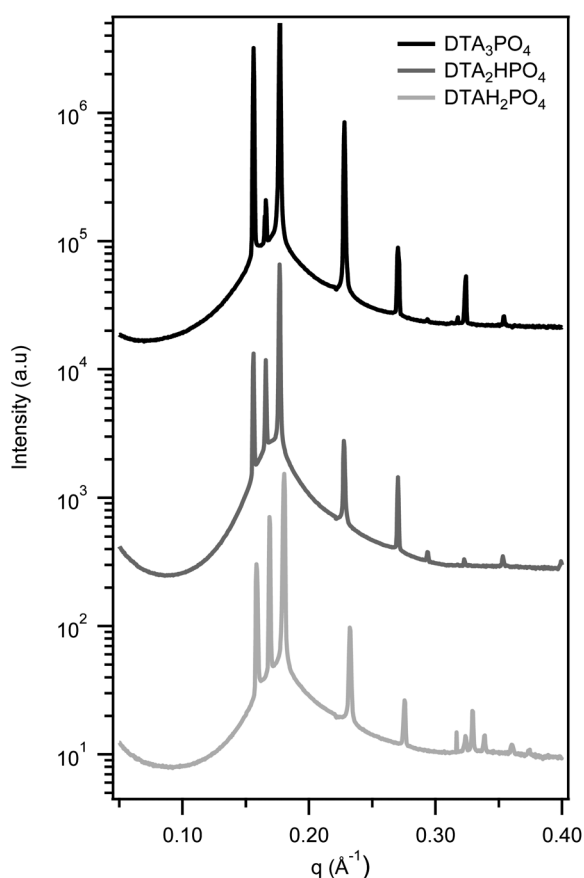


Fig. 3 SAXS patterns showing the HCPS phase with *P6₃/mmc* symmetry is formed in all DTA⁺ phosphate surfactants. HCPS appears before the *Pm3n* cubic phase and after the micellar solution in all cases, irrespective of the hydrolysis state of the phosphate counterion. Curves are offset for clarity.



hydrated counterions such as the hydrolysable phosphate (PO_4^{3-} , HPO_4^{2-} and H_2PO_4^-), oxalate (HC_2O_4^- and $\text{C}_2\text{O}_4^{2-}$), and carbonate ($\text{HCO}_3^-/\text{CO}_3^{2-}$). In this phase, micelles are spherical, consistent with the stability of spherical micelles of these surfactants in more dilute solution.

The HCPS phase persists for alkyl chain lengths up to sixteen carbons for alkyltrimethylammonium phosphate, carbonate and oxalates regardless of the hydrolysis state of the counterion, except for hexadecyltrimethylammonium HC_2O_4^- , which forms the $Pm3n$ cubic phase as its most dilute lyotropic phase. With increasing surfactant concentration all the systems examined form the $Pm3n$ cubic and then the H_1 hexagonal phase.

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