

Si–H addition followed by C–H bond activation induced by a terminal thorium imido metallocene: a combined experimental and computational study†

Cite this: *Chem. Sci.*, 2014, 5, 3165Wenshan Ren,^{ab} Enwei Zhou,^a Bo Fang,^a Guofu Zi,^{*a} De-Cai Fang^{*a}
and Marc D. Walter^{*c}

The Si–H bond addition to a terminal actinide imido complex was comprehensively studied. The base-free thorium imido [η^5 -1,2,4-(Me₃C)₃C₅H₂]₂Th=N(*p*-tolyl) (**1**) activates Si–H bonds in PhSiH₃ or Ph₂SiH₂ to give the thorium amido hydrido metallocenes [η^5 -1,2,4-(Me₃C)₃C₅H₂]₂Th(H)[N(*p*-tolyl)SiH₂Ph] (**2**) and [η^5 -1,2,4-(Me₃C)₃C₅H₂]₂Th(H)[N(*p*-tolyl)SiHPh₂] (**3**), respectively. Complex **2** readily inserts unsaturated molecules into the Th–H bond, whereas complex **3** reversibly activates an intramolecular aromatic C–H bond to yield [η^5 -1,2,4-(Me₃C)₃C₅H₂]₂Th[η^2 -N,C-(*p*-MeC₆H₃)(SiHPh₂)] (**4**) and H₂. The experimental results have been complemented by density functional theory (DFT) studies and provide a detailed understanding of the observed reactivity. In addition, a comparison between Th and early transition metals reveals that the Th⁴⁺ behaves more like an actinide than a transition metal.

Received 23rd February 2014
Accepted 25th May 2014

DOI: 10.1039/c4sc00576g

www.rsc.org/chemicalscience

Introduction

Terminal imido complexes of actinide-metals containing an An=N functionality have received widespread attention over the last 20 years because of their unique structural properties and their potential application in group transfer reactions and catalysis.^{1–3} Whereas many uranium imido complexes have been prepared and structurally characterized, only a few of them show significant reactivity.^{2*h-k,s,t,x*,3*f,g,j,t,o,y*} In contrast, thorium imido complexes have remained rare and not much is known about their reaction chemistry.^{2*d,3h,k,q*} This is surprising since thorium has a 7s² 6d² ground state electron configuration, and one might expect a similar reactivity to that of group 3 and 4 metals, such as Sc, Ti, Zr and Hf, for which several complexes with M=N bond have been prepared.^{4,5} However, the underlying question remains whether 5f-orbitals contribute to the bonding in thorium organometallics and whether Th⁴⁺ should be considered as a transition metal or as an actinide element.⁶ To answer this question we have recently prepared the base-free terminal thorium imido complex [η^5 -1,2,4-(Me₃C)₃C₅H₂]₂Th=

N(*p*-tolyl) (**1**).^{7,8} Complex **1** shows a rich reaction chemistry such as the activation of elemental sulfur (S₈).⁷ Furthermore it is an important intermediate in the catalytic hydroamination of internal acetylenes,⁷ an efficient catalyst for the trimerization of PhCN,⁷ and a useful precursor for the preparation of oxido and sulfido thorium metallocenes [η^5 -1,2,4-(Me₃C)₃C₅H₂]₂Th=E (E = O, S) by cycloaddition–elimination reactions with Ph₂C=E (E = O, S) or CS₂.⁹ Encouraged by this broad reactivity, we are now focusing on small molecule activation. Herein, the first Si–H bond activations by an actinide imido complex and the reactivity of the resulting thorium amido hydrido complexes are reported. Furthermore, the difference between thorium and early transition metal imido complexes will be addressed.

Experimental

General methods

All reactions and product manipulations were carried out under an atmosphere of dry dinitrogen with rigid exclusion of air and moisture using standard Schlenk or cannula techniques, or in a glove box. All organic solvents were freshly distilled from sodium benzophenone ketyl immediately prior to use. PhSiH₃ and Ph₂SiH₂ were freshly distilled from CaH₂ immediately prior to use. [η^5 -1,2,4-(Me₃C)₃C₅H₂]₂Th=N(*p*-tolyl) (**1**) was prepared according to literature methods.⁷ All other chemicals were purchased from Aldrich Chemical Co. and Beijing Chemical Co. and used as received unless otherwise noted. Infrared spectra were obtained from KBr pellets on an Avatar 360 Fourier transform spectrometer. ¹H and ¹³C NMR spectra were recorded on a Bruker AV 400 spectrometer at 400 and 100 MHz, respectively. All chemical shifts are reported in δ units with reference

^aDepartment of Chemistry, Beijing Normal University, Beijing 100875, China. E-mail: gzi@bnu.edu.cn; dcfang@bnu.edu.cn; Fax: +86-10-58802075; Tel: +86-10-58806051

^bCollege of Chemistry and Chemical Engineering, Southwest University, Chongqing 400715, China

^cInstitut für Anorganische und Analytische Chemie, Technische Universität Braunschweig, Hagenring 30, 38106 Braunschweig, Germany. E-mail: mwalter@tu-bs.de; Fax: +49-531-3915387; Tel: +49-531-3915312

† Electronic supplementary information (ESI) available: NBO analysis of complex **1**, cartesian coordinates of all stationary points optimized at B3PW91-PCM + D3 level and kinetic study. CCDC 977309–977311. For ESI and crystallographic data in CIF or other electronic format see DOI: 10.1039/c4sc00576g



to the residual protons of the deuterated solvents, which are internal standards, for proton and carbon chemical shifts. Melting points were measured on an X-6 melting point apparatus and were uncorrected. Elemental analyses were performed on a Vario EL elemental analyzer.

Syntheses

Preparation of $[\eta^5\text{-}1,2,4\text{-(Me}_3\text{C)}_3\text{C}_5\text{H}_2]_2\text{Th(H)[N}(p\text{-tolyl)SiH}_2\text{Ph}]$ (**2a**)

Method A. A toluene (5 mL) solution of PhSiH_3 (68 mg, 0.622 mmol) was added to a toluene (10 mL) solution of $[\eta^5\text{-}1,2,4\text{-(Me}_3\text{C)}_3\text{C}_5\text{H}_2]_2\text{Th}=\text{N}(p\text{-tolyl})$ (**1**; 500 mg, 0.622 mmol). After this mixture was stirred at room temperature for one hour, the solvent was removed. The residue was extracted with *n*-hexane (10 mL \times 2) and filtered. The volume of the filtrate was reduced to ca. 2 mL and colorless crystals of **2a** were isolated when this solution stood at room temperature for 2 days. Yield: 465 mg (82%) (found: C, 61.72; H, 8.18; N, 1.63. $\text{C}_{47}\text{H}_{73}\text{NSiTh}$ requires C, 61.88; H, 8.07; N, 1.54%). Mp: 116–118 °C (decomp.). $^1\text{H NMR}$ (C_6D_6): δ 18.07 (s, 1H, ThH), 7.59 (d, $J = 5.6$ Hz, 2H, aryl), 7.13 (m, 3H, aryl), 7.06 (d, $J = 5.6$ Hz, 2H, aryl), 7.02 (m, 2H, aryl), 6.14 (d, $J = 3.6$ Hz, 2H, ring CH), 6.01 (d, $J = 3.6$ Hz, 2H, ring CH), 5.06 (s, 2H, SiH₂), 2.15 (s, 3H, tolylCH₃), 1.57 (s, 18H, (CH₃)₃C), 1.48 (s, 18H, (CH₃)₃C), 1.43 (s, 18H, (CH₃)₃C). $^{13}\text{C}\{^1\text{H}\}$ NMR (C_6D_6): δ 143.1 (phenyl C), 142.2 (phenyl C), 135.9 (phenyl C), 135.3 (phenyl C), 134.9 (phenyl C), 130.0 (phenyl C), 129.6 (phenyl C), 122.9 (phenyl C), 119.0 (ring C), 115.7 (ring C), 115.5 (ring C), 115.3 (ring C), 113.7 (ring C), 35.1 ((CH₃)₃C), 34.6 ((CH₃)₃C), 34.1 ((CH₃)₃C), 34.0 ((CH₃)₃C), 33.7 ((CH₃)₃C), 33.2 ((CH₃)₃C), 20.4 (CH₃). IR (KBr, cm^{-1}): ν 2959 (s, C–H), 2124 (w, Si–H), 1606 (m, C=C), 1458 (s, C–C), 1385 (s, C–C), 1258 (s, C–N), 1094 (s, C–N), 1019 (s, C–N). The NMR sample was maintained at 70 °C and monitored periodically by $^1\text{H NMR}$ spectroscopy. After 4 days, resonances of the *anti*-isomer $[\eta^5\text{-}1,2,4\text{-(Me}_3\text{C)}_3\text{C}_5\text{H}_2]_2\text{Th(H)[N}(p\text{-tolyl)SiH}_2\text{Ph}]$ (**2b**) ($^1\text{H NMR}$ (C_6D_6): δ 16.93 (s, 1H, ThH), 8.00 (s, 1H, aryl), 7.92 (d, $J = 6.8$ Hz, 2H, aryl), 7.39 (d, $J = 7.4$ Hz, 2H, aryl), 7.23 (d, $J = 7.4$ Hz, 2H, aryl), 7.13 (m, 2H, aryl), 6.41 (d, $J = 3.2$ Hz, 2H, ring CH), 6.20 (d, $J = 3.2$ Hz, 2H, ring CH), 5.23 (s, 2H, SiH₂), 2.38 (s, 3H, tolylCH₃), 1.39 (s, 18H, (CH₃)₃C), 1.38 (s, 18H, (CH₃)₃C), 1.36 (s, 18H, (CH₃)₃C)) were observed by $^1\text{H NMR}$ spectroscopy (*syn/anti* ratio 1 : 9). However, on prolonged heating some degradation was also observed.

Method B

NMR scale. To a J. Young NMR tube charged with $[\eta^5\text{-}1,2,4\text{-(Me}_3\text{C)}_3\text{C}_5\text{H}_2]_2\text{Th}=\text{N}(p\text{-tolyl})$ (**1**; 16 mg, 0.02 mmol) and C_6D_6 (0.5 mL), an excess of PhSiH_3 was added. Upon addition complete conversion to **2a** was observed by $^1\text{H NMR}$ spectroscopy.

Preparation of $[\eta^5\text{-}1,2,4\text{-(Me}_3\text{C)}_3\text{C}_5\text{H}_2]_2\text{Th(H)[N}(p\text{-tolyl)SiHPh}_2]$ (**3a**)

Method A. This compound was prepared as colorless microcrystals from the reaction of $[\eta^5\text{-}1,2,4\text{-(Me}_3\text{C)}_3\text{C}_5\text{H}_2]_2\text{Th}=\text{N}(p\text{-tolyl})$ (**1**; 500 mg, 0.622 mmol) and Ph_2SiH_2 (114 mg, 0.622 mmol) in toluene (15 mL) and recrystallization from an *n*-hexane solution by a similar procedure as outlined for the

synthesis of **2a**. Yield: 516 mg (84%) (found: C, 64.52; H, 8.02; N, 1.43. $\text{C}_{53}\text{H}_{77}\text{NSiTh}$ requires C, 64.41; H, 7.85; N, 1.42%). Mp: 120–122 °C (decomp.). $^1\text{H NMR}$ (C_6D_6): δ 17.84 (s, 1H, ThH), 7.74 (s, 2H, aryl), 7.50 (d, $J = 7.2$ Hz, 2H, aryl), 7.11 (m, 8H, aryl), 7.03 (d, $J = 7.2$ Hz, 2H, aryl), 6.02 (s, 4H, ring CH), 5.08 (s, 1H, SiH), 2.18 (s, 3H, tolylCH₃), 1.58 (s, 18H, (CH₃)₃C), 1.43 (s, 36H, (CH₃)₃C). $^{13}\text{C}\{^1\text{H}\}$ NMR (C_6D_6): δ 136.2 (phenyl C), 135.8 (phenyl C), 134.5 (phenyl C), 131.5 (phenyl C), 129.9 (phenyl C), 129.7 (phenyl C), 129.6 (phenyl C), 128.2 (phenyl C), 127.8 (ring C), 127.5 (ring C), 119.0 (ring C), 115.5 (ring C), 115.2 (ring C), 35.0 ((CH₃)₃C), 34.8 ((CH₃)₃C), 34.4 ((CH₃)₃C), 34.1 ((CH₃)₃C), 33.3 ((CH₃)₃C), 33.1 ((CH₃)₃C), 20.5 (CH₃). IR (KBr, cm^{-1}): ν 2961 (s, C–H), 2110 (w, Si–H), 1587 (m, C=C), 1454 (m, C–C), 1384 (s, C–C), 1260 (s, C–N), 1090 (s, C–N), 1018 (s, C–N). The NMR sample was maintained at 70 °C and monitored periodically by $^1\text{H NMR}$ spectroscopy. After 1 h, new resonances of $[\eta^5\text{-}1,2,4\text{-(Me}_3\text{C)}_3\text{C}_5\text{H}_2]_2\text{Th}[\eta^2\text{-N,C}\{N(p\text{-MeC}_6\text{H}_4)\text{SiHPh}_2\}]$ (**4**) (see below) were observed by $^1\text{H NMR}$ spectroscopy with 20% conversion. After 3 h, 27% conversion was observed. However, no change in this ratio was detected on prolonged heating.

Method B

NMR scale. To a J. Young NMR tube charged with $[\eta^5\text{-}1,2,4\text{-(Me}_3\text{C)}_3\text{C}_5\text{H}_2]_2\text{Th}=\text{N}(p\text{-tolyl})$ (**1**; 16 mg, 0.02 mmol) and C_6D_6 (0.5 mL), an excess of Ph_2SiH_2 was added. Complete conversion to **3a** was observed by $^1\text{H NMR}$ spectroscopy.

Preparation of $[\eta^5\text{-}1,2,4\text{-(Me}_3\text{C)}_3\text{C}_5\text{H}_2]_2\text{Th}[\eta^2\text{-N,C}\{N(p\text{-MeC}_6\text{H}_4)\text{SiHPh}_2\}]$ (4**).** After a toluene (10 mL) solution of $[\eta^5\text{-}1,2,4\text{-(Me}_3\text{C)}_3\text{C}_5\text{H}_2]_2\text{Th(H)[N}(p\text{-tolyl)SiHPh}_2]$ (**3a**; 247 mg, 0.25 mmol) was stirred at 70 °C for three days, the solvent was removed. The residue was extracted with *n*-hexane (10 mL \times 2) and filtered. The volume of the filtrate was reduced to ca. 1 mL, and yellow crystals of **4** were isolated when this solution stood at room temperature for 2 days. Yield: 197 mg (80%) (found: C, 64.52; H, 7.82; N, 1.43. $\text{C}_{53}\text{H}_{75}\text{NSiTh}$ requires C, 64.54; H, 7.66; N, 1.41%). Mp: 128–130 °C (decomp.). $^1\text{H NMR}$ (C_6D_6): δ 7.97 (m, 5H, aryl), 7.49 (d, $J = 8.0$ Hz, 1H, aryl), 7.25 (m, 4H, aryl), 7.20 (d, $J = 6.8$ Hz, 2H, aryl), 6.81 (d, $J = 8.0$ Hz, 1H, aryl), 6.35 (s, 2H, ring CH), 6.18 (d, $J = 3.2$ Hz, 2H, ring CH), 5.97 (s, 1H, SiH), 2.35 (s, 3H, tolylCH₃), 1.42 (s, 18H, (CH₃)₃C), 1.36 (s, 18H, (CH₃)₃C), 1.31 (s, 18H, (CH₃)₃C). $^{13}\text{C}\{^1\text{H}\}$ NMR (C_6D_6): δ 190.4 (ThC), 144.9 (phenyl C), 144.2 (phenyl C), 142.6 (phenyl C), 142.1 (phenyl C), 141.9 (phenyl C), 137.8 (phenyl C), 136.3 (phenyl C), 130.5 (phenyl C), 129.6 (phenyl C), 129.3 (ring C), 127.3 (ring C), 118.3 (ring C), 116.3 (ring C), 115.9 (ring C), 35.0 ((CH₃)₃C), 34.4 ((CH₃)₃C), 34.3 ((CH₃)₃C), 34.3 ((CH₃)₃C), 34.0 ((CH₃)₃C), 32.5 ((CH₃)₃C), 21.2 (CH₃). IR (KBr, cm^{-1}): ν 2962 (s, C–H), 2113 (w, Si–H), 1457 (m, C–C), 1384 (s, C–C), 1260 (s, C–N), 1091 (s, C–N), 1019 (s, C–N).

Reaction of $[\eta^5\text{-}1,2,4\text{-(Me}_3\text{C)}_3\text{C}_5\text{H}_2]_2\text{Th}[\eta^2\text{-N,C}\{N(p\text{-MeC}_6\text{H}_4)\text{SiHPh}_2\}]$ (**4**) with H_2

NMR scale. To a J. Young NMR tube charged with $[\eta^5\text{-}1,2,4\text{-(Me}_3\text{C)}_3\text{C}_5\text{H}_2]_2\text{Th}[\eta^2\text{-N,C}\{N(p\text{-MeC}_6\text{H}_4)\text{SiHPh}_2\}]$ (**4**; 20 mg, 0.02 mmol) and C_6D_6 (0.5 mL) an excess of H_2 (at 1 atm) was added. $^1\text{H NMR}$ spectroscopy indicated complete conversion to **3a** and **3b** (ratio **3a/3b** ca. 10 : 1; $^1\text{H NMR}$ (C_6D_6): δ 16.84 (s, 1H, ThH), 7.84 (d, $J = 6.8$ Hz, 1H, aryl), 7.60 (d, $J = 7.5$ Hz, 2H, aryl), 7.34 (s, 2H, aryl), 7.11 (m, 6H, aryl), 6.95 (s, 1H, aryl H), 6.77 (d, J



= 8.1 Hz, 2H, aryl), 6.21 (d, $J = 3.0$ Hz, 2H, ring CH), 6.18 (d, $J = 3.0$ Hz, 2H, ring CH), 5.14 (s, 1H, SiH), 2.19 (s, 3H, tolylCH₃), 1.54 (s, 18H, (CH₃)₃C), 1.41 (s, 18H, (CH₃)₃C), 1.39 (s, 18H, (CH₃)₃C).

Preparation of $[\eta^5-1,2,4-(\text{Me}_3\text{C})_3\text{C}_5\text{H}_2]_2\text{Th}[\text{N}(p\text{-tolyl})-\text{C}(=\text{NC}_6\text{H}_{11})-\text{N}(\text{C}_6\text{H}_{11})]$ (6**)**

Method A. A toluene (5 mL) solution of *N,N'*-dicyclohexylcarbodiimide (DCC; 103 mg, 0.50 mmol) was added to a toluene (10 mL) solution of $[\eta^5-1,2,4-(\text{Me}_3\text{C})_3\text{C}_5\text{H}_2]_2\text{Th}(\text{H})[\text{N}(p\text{-tolyl})\text{SiH}_2\text{Ph}]$ (**2a**; 228 mg, 0.25 mmol). After this solution was stirred at room temperature for 2 hours, the solvent was removed. The residue was extracted with *n*-hexane (10 mL \times 2) and filtered. The volume of the filtrate was reduced to *ca.* 2 mL, and colorless crystals of **6** were isolated when this solution stood at room temperature for 2 days. Yield 217 mg (86%) (found: C, 64.32; H, 8.58; N, 4.03. C₅₄H₈₇N₃Th requires C, 64.19; H, 8.68; N, 4.16%). Mp: 210–212 °C (decomp.). ¹H NMR (C₆D₆): δ 7.30 (d, $J = 8.0$ Hz, 2H, aryl), 7.09 (d, $J = 8.0$ Hz, 2H, aryl), 6.68 (s, 2H, ring CH), 6.35 (s, 2H, ring CH), 3.84 (m, 1H, NCH), 3.31 (m, 2H, CH₂), 3.24 (m, 1H, NCH), 2.36 (m, 2H, CH₂), 2.25 (s, 3H, tolylCH₃), 2.12 (m, 4H, CH₂), 1.84 (m, 8H, CH₂), 1.53 (m, 4H, CH₂), 1.52 (s, 18H, (CH₃)₃C), 1.50 (s, 18H, (CH₃)₃C), 1.27 (s, 18H, (CH₃)₃C). ¹³C{¹H} NMR (C₆D₆): δ 152.0 (C=N), 145.9 (phenyl C), 144.2 (phenyl C), 143.6 (phenyl C), 143.6 (phenyl C), 130.8 (ring C), 129.3 (ring C), 127.5 (ring C), 125.9 (ring C), 118.3 (ring C), 62.1 (CHN), 53.8 (CHN), 36.4 ((CH₃)₃C), 35.0 ((CH₃)₃C), 34.9 ((CH₃)₃C), 34.8 ((CH₃)₃C), 34.5 ((CH₃)₃C), 33.2 ((CH₃)₃C), 32.4 (CH₂), 31.2 (CH₂), 28.0 (CH₂), 27.2 (CH₂), 27.1 (CH₂), 25.6 (CH₂), 20.8 (CH₃). IR (KBr, cm⁻¹): ν 2961 (s, C–H), 1618 (m, C=N), 1574 (s, C=C), 1448 (s, C–C), 1385 (s, C–C), 1259 (s, C–N), 1093 (s, C–N), 1019 (s, C–N).

Method B

NMR scale. To a J. Young NMR tube charged with $[\eta^5-1,2,4-(\text{Me}_3\text{C})_3\text{C}_5\text{H}_2]_2\text{Th}(\text{H})[\text{N}(p\text{-tolyl})\text{SiH}_2\text{Ph}]$ (**2a**; 18 mg, 0.02 mmol) and C₆D₆ (0.5 mL), DCC (8.2 mg, 0.04 mmol) was added. Complete conversion to **6** and (C₆H₁₁)N=CHN(C₆H₁₁)(SiH₂Ph) (**5**) (¹H NMR (C₆D₆): δ 7.37 (s, 1H, N=CH), 7.10 (m, 5H, aryl), 4.22 (s, 2H, SiH₂), 3.12 (m, 2H, NCH), 1.84 (m, 16H, CH₂), 1.32 (m, 4H, CH₂)) was observed by ¹H NMR spectroscopy.

Method C

NMR scale. To a J. Young NMR tube charged with $[\eta^5-1,2,4-(\text{Me}_3\text{C})_3\text{C}_5\text{H}_2]_2\text{Th}=\text{N}(p\text{-tolyl})$ (**1**; 16 mg, 0.02 mmol) and C₆D₆ (0.5 mL), DCC (4.1 mg, 0.02 mmol) was added. Complete conversion to **6** was observed by ¹H NMR spectroscopy.

Reaction of $[\eta^5-1,2,4-(\text{Me}_3\text{C})_3\text{C}_5\text{H}_2]_2\text{Th}(\text{H})[\text{N}(p\text{-tolyl})\text{SiH}_2\text{Ph}]$ (2a**) with DCC**

NMR scale. To a J. Young NMR tube charged with $[\eta^5-1,2,4-(\text{Me}_3\text{C})_3\text{C}_5\text{H}_2]_2\text{Th}(\text{H})[\text{N}(p\text{-tolyl})\text{SiH}_2\text{Ph}]$ (**2a**; 18 mg, 0.02 mmol) and C₆D₆ (0.5 mL), DCC (4.1 mg, 0.02 mmol) was added. Resonances due to **6** along with those of **5** and unreacted **2a** were observed by ¹H NMR spectroscopy (50% conversion based on **2a**).

X-ray crystallography

Single-crystal X-ray diffraction measurements were carried out on a Bruker SMART CCD diffractometer using graphite monochromated Mo K α radiation ($\lambda = 0.71073$ Å). An empirical

absorption correction was applied using the SADABS program.¹⁰ All structures were solved by direct methods and refined by full-matrix least squares on F^2 using the SHELXL-97 program package.¹¹ The hydride atom in **2a** was located from a difference-Fourier map and refined isotropically. Hydrogen atoms were geometrically fixed using the riding model. Disordered solvents in the voids of **2a** and **6** were modeled or removed by using the SQUEEZE program.¹² Crystallographic details for **2a**, **4** and **6** are summarized in Table 1.

Computational methods

All calculations were carried out with the Gaussian 09 program (G09),¹³ employing the B3PW91 method, plus polarizable continuum model (PCM) and D3 (ref. 14) (denoted as B3PW91-PCM + D3), with standard 6-31G(d) basis set for C, H, N and Si elements and Stuttgart RLC ECP from EMSL basis set exchange (<https://bse.pnl.gov/bse/portal>) for Th element,¹⁵ to fully optimize the geometries of reactants, complexes, transition state, intermediates, and product structures, and to mimic experimental toluene-solvent conditions (dielectricity constant $\epsilon = 2.379$). All stationary points were subsequently characterized by vibrational analyses, from which their respective zero-point (vibrational) energy (ZPE) were extracted and used in the relative energy determinations; in addition to ensure that the reactant, complex, intermediate, product and transition state structures resided at minima and 1st order saddle points, respectively, on their potential energy hyper surfaces.

Results and discussion

Complex **1** reacts rapidly with silanes such as PhSiH₃ or Ph₂SiH₂ to give the amido hydrido metallocenes $[\eta^5-1,2,4-(\text{Me}_3\text{C})_3\text{C}_5\text{H}_2]_2\text{Th}(\text{H})[\text{N}(p\text{-tolyl})\text{SiH}_2\text{Ph}]$ (**2a**) and $[\eta^5-1,2,4-(\text{Me}_3\text{C})_3\text{C}_5\text{H}_2]_2\text{Th}(\text{H})[\text{N}(p\text{-tolyl})\text{SiHPh}_2]$ (**3a**), respectively (Scheme 1). Consistent with a Si–H bond addition across the Th=N bond the *syn*-isomer is formed exclusively. However, in contrast to scandium^{4f} and titanium^{5k,l} imido complexes, the reaction of **1** with silanes is irreversible, which is presumably a consequence of the more polarized actinide imido bond (see ESIT[†]). Previous studies have clearly established that the 5f orbitals play a key role in the bonding of actinide complexes, which results in a very polarized An=E bond, whereas this is not the case for group 4 metal complexes.^{6c} Therefore, actinide An=E bonds exhibit different reactivity patterns compared to those of early transition metals, as illustrated by reactivity of imido complexes (M=NR) with silanes, in which the less polarized M=N bond (M is early transition metal) shows a reversible Si–H addition, whereas the more polarized Th=N bond undergoes an irreversible Si–H addition because of a more pronounced thermodynamic stabilization of the Si–H addition products. Heating of **2a** at 70 °C for four days forms the *anti*-isomer **2b** (*ca.* 10%). Under similar reaction conditions, complex **3a** is in equilibrium with the aromatic C–H bond activation product $[\eta^5-1,2,4-(\text{Me}_3\text{C})_3\text{C}_5\text{H}_2]_2\text{Th}[\eta^2\text{-}N,C\text{-}\{N(p\text{-MeC}_6\text{H}_5)(\text{SiHPh}_2)\}]$ (**4**) and H₂ (Scheme 1). Moreover, the addition of H₂ (1 atm) to **4** at room temperature results in a mixture



Table 1 Crystal data and experimental parameters for compounds 2a, 4 and 6

Compound	2a	4	6
Formula	C ₄₇ H ₇₃ NSiTh	C ₅₃ H ₇₅ NSiTh	C ₅₄ H ₈₇ N ₃ Th
Fw	912.19	986.27	1010.31
Crystal system	Orthorhombic	Monoclinic	Triclinic
Space group	<i>Pbca</i>	<i>P2₁/c</i>	<i>P</i> ($\bar{1}$)
<i>a</i> (Å)	10.425 (1)	16.578 (3)	10.436 (2)
<i>b</i> (Å)	20.765 (1)	16.572 (3)	15.509 (3)
<i>c</i> (Å)	44.289 (3)	17.515 (3)	20.314 (4)
α (deg.)	90	90	92.79 (1)
β (deg.)	90	97.936 (3)	92.25 (1)
γ (deg.)	90	90	109.07 (1)
<i>V</i> (Å ³)	9587.6 (10)	4765.9 (13)	3098.4 (10)
<i>Z</i>	8	4	2
<i>D</i> _{calc} (g cm ⁻³)	1.264	1.375	1.083
μ (Mo/ <i>K</i> α) _{calc} (cm ⁻¹)	3.164	3.188	2.436
Size (mm)	0.39 × 0.21 × 0.11	0.25 × 0.18 × 0.10	0.45 × 0.20 × 0.11
<i>F</i> (000)	3728	2016	1044
2 θ range (deg.)	3.92 to 55.34	4.40 to 55.22	3.54 to 50.50
No. of reflns, collected	11 151	26 706	10 908
No. of obsd reflns	8989	7619	9578
No. of variables	528	527	542
Abscorr (<i>T</i> _{max} , <i>T</i> _{min})	0.72, 0.37	0.74, 0.50	0.78, 0.41
<i>R</i>	0.032	0.051	0.078
<i>R</i> _w	0.064	0.121	0.224
<i>R</i> _{all}	0.066	0.146	0.230
Gof	1.10	1.07	1.12
CCDC	977309	977310	977311

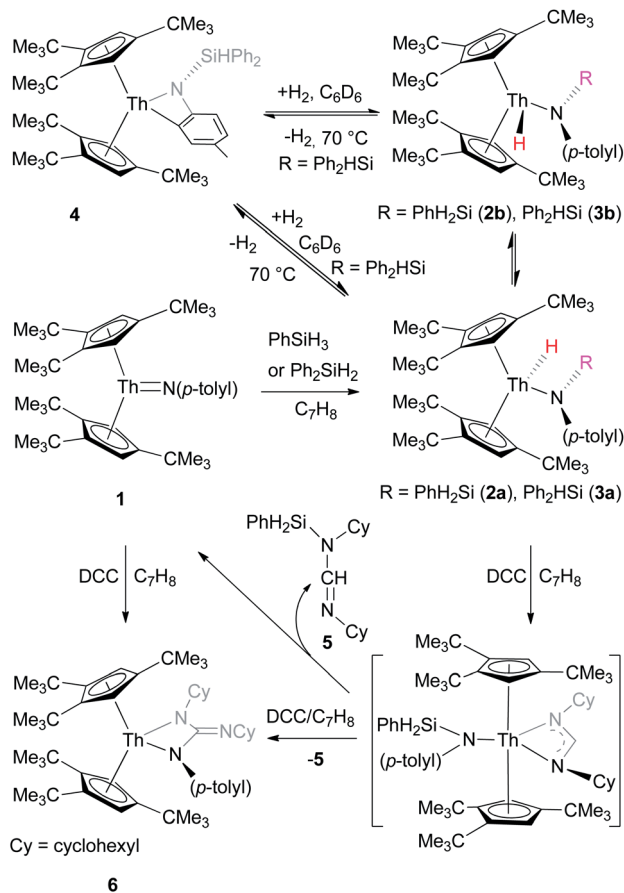
of the *syn*-/*anti*-isomers **3a** and **3b** in a 10 : 1 ratio (Scheme 1), as established by ¹H NMR spectroscopy. Although complex **1** reacts rapidly with PhSiH₃ or Ph₂SiH₂, no reaction was observed with tertiary silanes such as Et₃SiH and Ph₃SiH most likely because of the steric hindrance.

The molecular structure of [η^5 -1,2,4-(Me₃C)₃C₅H₂]₂Th(H)[N(*p*-tolyl)SiH₂Ph] (**2a**) is shown in Fig. 1. The cyclopentadienyl rings adopt a nearly eclipsed conformation with an average Th–C (ring) distance of 2.858(3) Å, and a Cp (cent)–Th–Cp (cent) angle of 138.1(3)°. The Th–N distance of 2.387(2) Å is comparable to the values found in [η^5 -1,2,4-(Me₃C)₃C₅H₂]₂Th[(bipy)(SCPh₂)] (2.435(1) Å)¹⁶ and [η^5 -1,2,4-(Me₃C)₃C₅H₂]₂Th(N-*p*-tolyl)₂ (2.228(3) Å).⁷ The Th–H distance is 2.01(3) Å, which is in line with reported terminal and bridging Th–H distances, (2,6-*t*-Bu₂C₆H₃O)₆Th₃(μ_3 -H)₂(μ -H)₄ (2.0(1)–2.6(1) Å),¹⁷ [(η^5 -C₅Me₅)₂ThH(μ -H)]₂ (2.03(1) (terminal)–2.29(3) (bridging) Å),¹⁸ (η^5 -C₅Me₅)₃ThH (2.33(13) Å),¹⁹ [η^5 -1,3-(Me₃C)₂C₅H₃]₃ThH (1.99(5) Å),²⁰ and (η^5 -C₅Me₄H)₄[μ - η^5 -C₅Me₃H(CH₂)₂- κ C]₂Th₄(μ -H)₄(μ_3 -H)₄ (2.503(4) and 2.508(4) Å).²¹

Fig. 2 shows the molecular structure of [η^5 -1,2,4-(Me₃C)₃C₅H₂]₂Th[η^2 -N,*C*-{N(*p*-MeC₆H₃)(SiHPh₂)}] (**4**). The average Th–C (ring) distance is 2.882(6) Å, and the Cp (cent)–Th–Cp (cent) angle is 140.6(2)°. The Th–N distance of 2.310(5) Å is shorter than that (2.387(2) Å) found in **2a**. Furthermore, the Th–C(40) distance of 2.420(3) Å is also significantly shorter than the reported Th–C_{aryl} bond distances (2.548(2)–2.654(14) Å).^{22–24} Interestingly, only four structurally characterized thorium aryl complexes have been reported, and in all of them the thorium–aryl bond is supported by chelating ligands.^{22–24}

As demonstrated above, complex **1** irreversibly activates Si–H bonds to yield the amido hydrido complexes **2** and **3**. Complex **3** undergoes a reversible intramolecular C–H bond activation to give **4** and H₂. DFT calculations were performed at the B3PW91 level of theory to further understand the observed process. To properly describe the observed reactivity it was necessary to include solvent and dispersion effects (see ESI†). The reaction of **1** with Ph₂SiH₂ to **3a** is exergonic with ΔG (343 K) = –17.6 kcal mol⁻¹ and proceeds concerted *via* the 4-membered transition state **TS** formed by the Th=N and Si–H moieties. The Si–N and Th–H bonds in **TS** are 2.343 and 2.299 Å, respectively, and *ca.* 0.61 and 0.23 Å longer than those in the product **3a** (see ESI†). The barrier for this reaction is ΔG^\ddagger = 16.7 kcal mol⁻¹ (343 K) (15.2 kcal mol⁻¹ at 298 K) (Fig. 3); and these values are consistent with the rapid and irreversible formation of **3a** at ambient temperature. The transformation of **3a** to **4** + H₂ is more complicated and involves the *anti*-intermediate **3b** and two transition states (**TSa** and **TSb**). The barrier for the conversion of **3a** to **3b** *via* **TSa** is 33.7 kcal mol⁻¹ at 343 K and also represents the rate-limiting step in this reaction. The computed barrier is consistent with the experimental observation that this reaction only occurs at elevated temperatures, and it is also in reasonable agreement with the experimentally estimated barrier of 26.8 kcal mol⁻¹ (see ESI†). The deviation between the computed and experimental values is mainly attributed to the low solubility of H₂ in organic solvent, which complicates the kinetic evaluation because the actual H₂ concentration cannot be determined with certainty (see ESI†). However, for **TSa** a closer inspection reveals an interesting





Scheme 1 Synthesis of compounds 2–4 and 6.

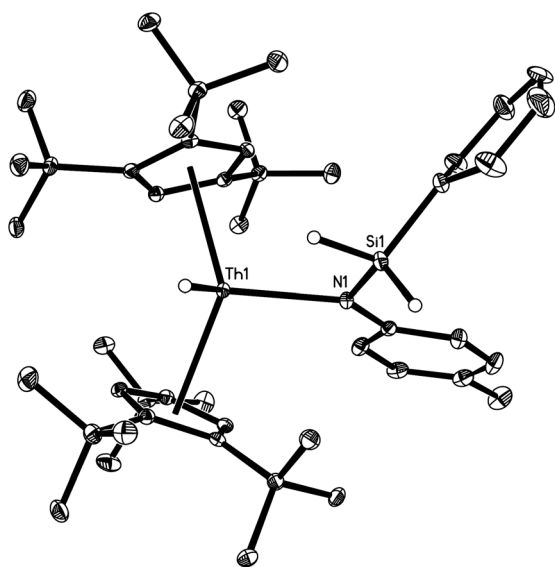


Fig. 1 Molecular structure of 2a (thermal ellipsoids drawn at the 35% probability level). Selected bond lengths [Å] and angles [°]: Th–C(Cp) (av.) 2.858(3), Th–C(Cp) (range) 2.870(3) to 2.946(3), Th–Cp(cent) (av.) 2.592(3), Th–N 2.387(2), Th–H 2.01(3), Th–Si 3.256(1), Cp (cent)–Th–Cp (cent) 138.1(3), N–Th–H 111.9(11).

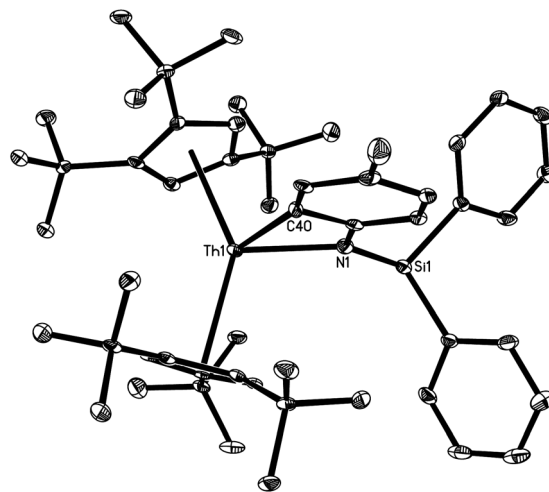


Fig. 2 Molecular structure of 4 (thermal ellipsoids drawn at the 35% probability level). Selected bond lengths [Å] and angles [°]: Th–C(Cp) (av.) 2.882(6), Th–C(Cp) (range) 2.803(6) to 2.956(6), Th–Cp (cent) (av.) 2.617(6), Th–N(1) 2.310(5), Th–C(40) 2.420(7), Cp (cent)–Th–Cp (cent) 140.6(2), N(1)–Th–C(40) 61.8(2).

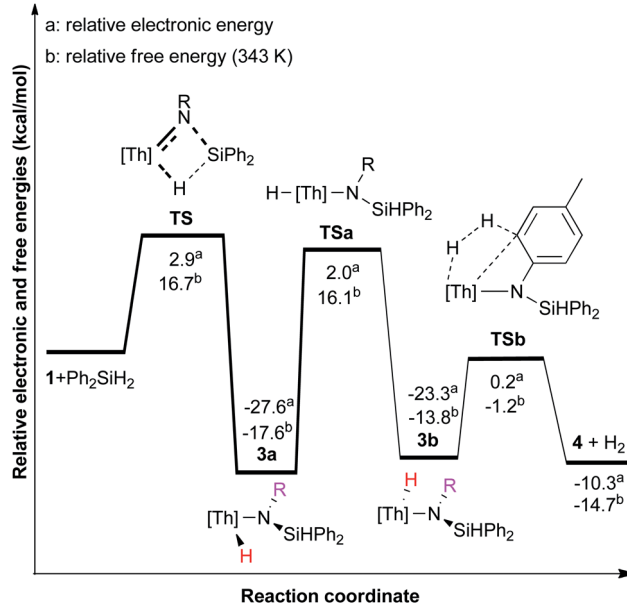


Fig. 3 Energy and free energy profile (kcal mol⁻¹) for the reaction of 1 with Ph₂SiH₂, obtained with B3PW91-PCM + D3 method. [Th] = [η⁵-1,2,4-(Me₃C)₃C₅H₂]₂Th. R = *p*-tolyl.

feature, since the *syn*-/*anti*-intermediates are interconverted not by rotation of the –N(R)(*p*-tolyl) moiety along the Th–N axis, but by a Th–H movement (Fig. 3). This is a direct consequence of the severe steric bulk of the two 1,2,4-(Me₃C)₃C₅H₂ ligands, which hinders the rotation along the Th–N axis. Nevertheless, once complex 3b is formed, H₂ is released *via* TSb and with a low barrier of 12.6 kcal mol⁻¹ (Fig. 3). This makes the *anti*-isomer 3b difficult to isolate, since H₂ release from the reaction mixture shifts the equilibrium to the final product 4. On the contrary, H₂ addition to complex 4 forms a mixture of 3b and 3a.



Furthermore, similar to the amido hydrido scandium complex,^{4f} amido hydrido thorium complexes can also insert unsaturated substrates such as carbodiimides into the Th–H bond. For example, treatment of **2a** with 2 equiv. of *N,N'*-dicyclohexylcarbodiimide (DCC) rapidly forms the metallacycle $[\eta^5-1,2,4-(\text{Me}_3\text{C})_3\text{C}_5\text{H}_2]_2\text{Th}[\text{N}(p\text{-tolyl})\text{C}(=\text{NC}_6\text{H}_{11})-\text{N}(\text{C}_6\text{H}_{11})]$ (**6**) and azomethine ($\text{C}_6\text{H}_{11}\text{N}=\text{CHN}(\text{C}_6\text{H}_{11})(\text{SiH}_2\text{Ph})$) (**5**) in quantitative conversion (Scheme 1). Furthermore, analogous to group 4 imido complexes,²⁵ thorium imido **1** can react with DCC to yield metallacycle **6** (Scheme 1), whereas **2a** reacts with 1 equiv. of DCC to give **6** in only 50% conversion. These observations suggest that the insertion product $[\eta^5-1,2,4-(\text{Me}_3\text{C})_3\text{C}_5\text{H}_2]_2\text{Th}[\text{N}(p\text{-tolyl})\text{SiHPh}_2][\eta^3-\text{N}(\text{C}_6\text{H}_{11})\text{CHN}(\text{C}_6\text{H}_{11})]$ is unstable and eliminates azomethine **5** to form the imido complex **1**, which then reacts with DCC to the metallacycle **6** by an [2 + 2] cycloaddition reaction (Scheme 1).

The molecular structure of $[\eta^5-1,2,4-(\text{Me}_3\text{C})_3\text{C}_5\text{H}_2]_2\text{Th}[\text{N}(p\text{-tolyl})\text{C}(=\text{NC}_6\text{H}_{11})-\text{N}(\text{C}_6\text{H}_{11})]$ (**6**) is shown in Fig. 4. The Th^{4+} ion is η^5 -bound to two Cp-rings and σ -coordinate to the two nitrogen atoms of the $[\text{N}(p\text{-tolyl})\text{C}(=\text{NC}_6\text{H}_{11})-\text{N}(\text{C}_6\text{H}_{11})]$ group in a distorted-tetrahedral geometry with an averaged Th–C (ring) distance of 2.877(11) Å. The dihedral angle defined by the planes containing the Cp (cent)–Th–Cp (cent) and N(1)–C(42)–N(3) is 75.2(3)°, and the N(1)–Th(1)–N(3) angle is 58.5(3)°. These metric parameters resemble those found in the alkyne addition complexes $[\eta^5-1,2,4-(\text{Me}_3\text{C})_3\text{C}_5\text{H}_2]_2\text{Th}[\text{N}(p\text{-tolyl})\text{C}(\text{R}')=\text{C}(\text{R}')]$ ($\text{R}' = \text{Me}, \text{Ph}$),⁷ while the Th–N distances (2.346(9) and 2.337(8) Å) are similar to those found in **2a** and **4**.

Conclusions

In conclusion, the first example of a Si–H bond activation by a terminal actinide imido complex has been comprehensively studied. In contrast to scandium^{4f} and titanium^{5k,l} imido complexes, silanes such as PhSiH_3 and Ph_2SiH_2 add irreversibly to the thorium imido $[\eta^5-1,2,4-(\text{Me}_3\text{C})_3\text{C}_5\text{H}_2]_2\text{Th}=\text{N}(p\text{-tolyl})$ (**1**),

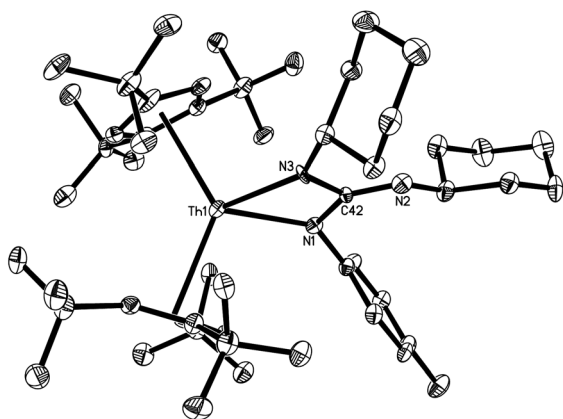


Fig. 4 Molecular structure of **6** (thermal ellipsoids drawn at the 35% probability level). Selected bond lengths [Å] and angles [°]: Th–C(Cp) (av.) 2.877(11), Th–C(Cp) (range) 2.772(10) to 3.024(11), Th–Cp (cent) (av.) 2.613(9), Th–N(1) 2.346(9), Th–N(3) 2.337(8), Cp (cent)–Th–Cp (cent) 128.6(3), N(1)–Th–N(3) 58.5(3).

supporting the notion that Th^{4+} behaves more like an actinide than a transition metal.^{6c} DFT studies reveal that the 1,2-addition proceeds in a concerted, 4-membered transition state to give amido hydrido complexes **2** and **3**. These compounds are reactive species, as illustrated by the formation of H_2 and $[\eta^5-1,2,4-(\text{Me}_3\text{C})_3\text{C}_5\text{H}_2]_2\text{Th}[\eta^2-N,C-\{\text{N}(p\text{-MeC}_6\text{H}_3)(\text{SiHPh}_2)\}]$ (**4**) via an intramolecular aromatic C–H bond activation in **3**, and by the insertion of unsaturated substrates such as DCC into the Th–H bond of **2**. The development of new actinide imido complexes and the exploration of thorium amido hydrido complexes in catalysis are ongoing projects in these laboratories.

Acknowledgements

This work was supported by the National Natural Science Foundation of China (Grant no. 21172022, 21302155, 21373030), and the Deutsche Forschungsgemeinschaft (DFG) through the Emmy-Noether program (WA 2513/2). We thank Dr Xuebin Deng for his help with the crystallography, and Prof. Richard A. Andersen for helpful discussions.

Notes and references

- For selected recent reviews, see: (a) M. Ephritikhine, *Dalton Trans.*, 2006, 2501–2516; (b) E. Barnea and M. S. Eisen, *Coord. Chem. Rev.*, 2006, **250**, 855–899; (c) A. R. Fox, S. C. Bart, K. Meyer and C. C. Cummins, *Nature*, 2008, **455**, 341–349; (d) T. Andrea and M. S. Eisen, *Chem. Soc. Rev.*, 2008, **37**, 550–567; (e) C. R. Graves and J. L. Kiplinger, *Chem. Commun.*, 2009, 3831–3853; (f) T. W. Hayton, *Dalton Trans.*, 2010, **39**, 1145–1158; (g) P. L. Arnold, *Chem. Commun.*, 2011, **47**, 9005–9010; (h) T. W. Hayton, *Chem. Commun.*, 2013, **49**, 2956–2973; (i) T. W. Hayton, *Nat. Chem.*, 2013, **5**, 451–452.
- For selected papers on actinide non-metallocenes containing terminal imido groups, see: (a) C. J. Burns, W. H. Smith, J. C. Huffman and A. P. Sattelberger, *J. Am. Chem. Soc.*, 1990, **112**, 3237–3239; (b) D. R. Brown and R. G. Denning, *Inorg. Chem.*, 1996, **35**, 6158–6163; (c) P. Roussel, R. Boaretto, A. J. Kingsley, N. W. Alcock and P. Scott, *J. Chem. Soc., Dalton Trans.*, 2002, 1423–1428; (d) I. Korobkov, S. Gambarotta and G. P. A. Yap, *Angew. Chem., Int. Ed.*, 2003, **42**, 814–818; (e) I. Castro-Rodriguez, K. Olsen, P. Gantzel and K. Meyer, *J. Am. Chem. Soc.*, 2003, **125**, 4565–4571; (f) T. W. Hayton, J. M. Boncella, B. L. Scott, P. D. Palmer, E. R. Batista and P. J. Hay, *Science*, 2005, **310**, 1941–1943; (g) T. W. Hayton, J. M. Boncella, B. L. Scott, E. R. Batista and P. J. Hay, *J. Am. Chem. Soc.*, 2006, **128**, 10549–10559; (h) T. W. Hayton, J. M. Boncella, B. L. Scott and E. R. Batista, *J. Am. Chem. Soc.*, 2006, **128**, 12622–12623; (i) I. Castro-Rodriguez, H. Nakai and K. Meyer, *Angew. Chem., Int. Ed.*, 2006, **45**, 2389–2392; (j) L. P. Spencer, P. Yang, B. L. Scott, E. R. Batista and J. M. Boncella, *J. Am. Chem. Soc.*, 2008, **130**, 2930–2931; (k) S. C. Bart, C. Anthon, F. W. Heinemann, E. Bill, N. M. Edelstein and K. Meyer, *J. Am. Chem. Soc.*, 2008, **130**,



12536–12546; (l) L. P. Spencer, E. J. Schelter, P. Yang, R. L. Gdula, B. L. Scott, J. D. Thompson, J. L. Kiplinger, E. R. Batista and J. M. Boncella, *Angew. Chem., Int. Ed.*, 2009, **48**, 3795–3798; (m) L. P. Spencer, P. Yang, B. L. Scott, E. R. Batista and J. M. Boncella, *Inorg. Chem.*, 2009, **48**, 2693–2700; (n) L. P. Spencer, P. Yang, B. L. Scott, E. R. Batista and J. M. Boncella, *Inorg. Chem.*, 2009, **48**, 11615–11623; (o) T. Cantat, C. R. Graves, B. L. Scott and J. L. Kiplinger, *Angew. Chem., Int. Ed.*, 2009, **48**, 3681–3684; (p) D. L. Swartz II, L. P. Spencer, B. L. Scott, A. L. Odom and J. M. Boncella, *Dalton Trans.*, 2010, **39**, 6841–6846; (q) L. A. Seaman, S. Fortier, G. Wu and T. W. Hayton, *Inorg. Chem.*, 2011, **50**, 636–646; (r) R. E. Jilek, L. P. Spencer, D. L. Kuiper, B. L. Scott, U. J. Williams, J. M. Kikkawa, E. J. Schelter and J. M. Boncella, *Inorg. Chem.*, 2011, **50**, 4235–4237; (s) E. M. Matson, P. E. Fanwick and S. C. Bart, *Eur. J. Inorg. Chem.*, 2012, 5471–5478; (t) E. M. Matson, M. G. Crestani, P. E. Fanwick and S. C. Bart, *Dalton Trans.*, 2012, **41**, 7952–7958; (u) O. P. Lam, S. M. Franke, H. Nakai, F. W. Heinemann, W. Hieringer and K. Meyer, *Inorg. Chem.*, 2012, **51**, 6190–6199; (v) C. Camp, J. Pécaut and M. Mazzanti, *J. Am. Chem. Soc.*, 2013, **135**, 12101–12111; (w) D. M. King, F. Tuna, E. J. L. McInnes, J. McMaster, W. Lewis, A. J. Blake and S. T. Liddle, *Science*, 2012, **337**, 717–720; (x) D. M. King, F. Tuna, E. J. L. McInnes, J. McMaster, W. Lewis, A. J. Blake and S. T. Liddle, *Nat. Chem.*, 2013, **5**, 482–488.

3 For selected papers about actinide metallocenes containing terminal imido groups, see: (a) R. E. Cramer, K. Panchanatheswaran and J. W. Gilje, *J. Am. Chem. Soc.*, 1984, **106**, 1853–1854; (b) J. G. Brennan and R. A. Andersen, *J. Am. Chem. Soc.*, 1985, **107**, 514–516; (c) R. K. Rosen, R. A. Andersen and N. M. Edelstein, *J. Am. Chem. Soc.*, 1990, **112**, 4588–4590; (d) D. S. J. Arney, C. J. Burns and D. C. Smith, *J. Am. Chem. Soc.*, 1992, **114**, 10068–10069; (e) D. S. J. Arney and C. J. Burns, *J. Am. Chem. Soc.*, 1993, **115**, 9840–9841; (f) D. S. J. Arney and C. J. Burns, *J. Am. Chem. Soc.*, 1995, **117**, 9448–9460; (g) T. Straub, W. Frank, G. J. Reiss and M. S. Eisen, *J. Chem. Soc., Dalton Trans.*, 1996, 2541–2546; (h) A. Haskel, T. Straub and M. S. Eisen, *Organometallics*, 1996, **15**, 3773–3775; (i) B. P. Warner, B. L. Scott and C. J. Burns, *Angew. Chem., Int. Ed.*, 1998, **37**, 959–960; (j) R. G. Peters, B. P. Warner, B. L. Scott and C. J. Burns, *Organometallics*, 1999, **18**, 2587–2589; (k) T. Straub, A. Haskel, T. G. Neyroud, M. Kapon, M. Botoshansky and M. S. Eisen, *Organometallics*, 2001, **20**, 5017–5035; (l) J. L. Kiplinger, D. E. Morris, B. L. Scott and C. J. Burns, *Chem. Commun.*, 2002, 30–31; (m) W. J. Evans, S. A. Kozimor and J. W. Ziller, *Chem. Commun.*, 2005, 4681–4683; (n) G. Zi, L. Jia, E. L. Werkema, M. D. Walter, J. P. Gottfriedsen and R. A. Andersen, *Organometallics*, 2005, **24**, 4251–4262; (o) G. Zi, L. L. Blossch, L. Jia and R. A. Andersen, *Organometallics*, 2005, **24**, 4602–4612; (p) C. R. Graves, B. L. Scott, D. E. Morris and J. L. Kiplinger, *J. Am. Chem. Soc.*, 2007, **129**, 11914–11915; (q) E. J. Schelter, D. E. Morris, B. L. Scott and J. L. Kiplinger, *Chem.*

Commun., 2007, 1029–1031; (r) C. R. Graves, P. Yang, S. A. Kozimor, A. E. Vaughn, D. L. Clark, S. D. Conradson, E. J. Schelter, B. L. Scott, J. D. Thompson, P. J. Hay, D. E. Morris and J. L. Kiplinger, *J. Am. Chem. Soc.*, 2008, **130**, 5272–5285; (s) C. R. Graves, B. L. Scott, D. E. Morris and J. L. Kiplinger, *Organometallics*, 2008, **27**, 3335–3337; (t) C. R. Graves, A. E. Vaughn, E. J. Schelter, B. L. Scott, J. D. Thompson, D. E. Morris and J. L. Kiplinger, *Inorg. Chem.*, 2008, **47**, 11879–11891; (u) L. P. Spencer, R. L. Gdula, T. W. Hayton, B. L. Scott and J. M. Boncella, *Chem. Commun.*, 2008, 4986–4988; (v) W. J. Evans, C. A. Traina and J. W. Ziller, *J. Am. Chem. Soc.*, 2009, **131**, 17473–17481; (w) C. R. Graves, B. L. Scott, D. E. Morris and J. L. Kiplinger, *Chem. Commun.*, 2009, 776–778; (x) W. J. Evans, E. Montalvo, J. W. Ziller, A. G. DiPasquale and A. L. Rheingold, *Inorg. Chem.*, 2010, **49**, 222–228; (y) R. K. Thomson, T. Cantat, B. L. Scott, D. E. Morris, E. R. Batista and J. L. Kiplinger, *Nat. Chem.*, 2010, **2**, 723–729.

4 Selected papers on scandium imido complexes: (a) E. Lu, Y. Li and Y. Chen, *Chem. Commun.*, 2010, **46**, 4469–4471; (b) E. Lu, J. Chu, M. Borzov, Y. Chen and G. Li, *Chem. Commun.*, 2011, **47**, 743–745; (c) J. Chu, E. Lu, Z. Liu, Y. Chen, X. Leng and H. Song, *Angew. Chem., Int. Ed.*, 2011, **50**, 7677–7680; (d) E. Lu, Q. Zhou, Y. Li, J. Chu, Y. Chen, X. Leng and J. Sun, *Chem. Commun.*, 2012, **48**, 3403–3405; (e) Z. Jian, W. Rong, Z. Mou, Y. Pan, H. Xie and D. Cui, *Chem. Commun.*, 2012, **48**, 7516–7518; (f) T. Chu, W. E. Piers, J. L. Dutton and M. Parvez, *Organometallics*, 2013, **32**, 1159–1165; (g) J. Chu, E. Lu, Y. Chen and X. Leng, *Organometallics*, 2013, **32**, 1137–1140; (h) W. Rong, J. Cheng, Z. Mou, H. Xie and D. Cui, *Organometallics*, 2013, **32**, 5523–5529; (i) J. Chu, C. E. Kefalidis, L. Maron, X. Leng and Y. Chen, *J. Am. Chem. Soc.*, 2013, **135**, 8165–8168.

5 Selected papers on group 4 imido complexes: (a) P. J. Walsh, F. J. Hollander and R. G. Bergman, *J. Am. Chem. Soc.*, 1988, **110**, 8729–8731; (b) C. C. Cummins, S. M. Baxter and P. T. Wolczanski, *J. Am. Chem. Soc.*, 1988, **110**, 8731–8733; (c) J. E. Hill, R. D. Profflet, P. E. Fanwick and I. P. Rothwell, *Angew. Chem., Int. Ed.*, 1990, **29**, 664–665; (d) H. W. Roesky, H. Voelker, M. Witt and M. Noltemeyer, *Angew. Chem., Int. Ed.*, 1990, **29**, 669–670; (e) P. J. Walsh, F. J. Hollander and R. G. Bergman, *Organometallics*, 1993, **12**, 3705–3723; (f) S. C. Dunn, A. S. Batsanov and P. Mountford, *J. Chem. Soc., Chem. Commun.*, 1994, 2007–2008; (g) J. L. Bennett and P. T. Wolczanski, *J. Am. Chem. Soc.*, 1994, **116**, 2179–2180; (h) J. L. Polse, R. A. Andersen and R. G. Bergman, *J. Am. Chem. Soc.*, 1998, **120**, 13405–13414; (i) H. M. Hoyt, F. E. Michael and R. G. Bergman, *J. Am. Chem. Soc.*, 2004, **126**, 1018–1019; (j) B. Lian, T. P. Spaniol, P. Horrillo-Martínez, K. C. Hultsch and J. Okuda, *Eur. J. Inorg. Chem.*, 2009, 429–434; (k) P. J. Tiong, A. Nova, E. Clot and P. Mountford, *Chem. Commun.*, 2011, **47**, 3147–3149; (l) P. J. Tiong, A. Nova, A. D. Schwarz, J. D. Selby, E. Clot and P. Mountford, *Dalton Trans.*, 2012, **41**, 2277–2288; (m) A. D. Schwarz, A. J. Nielson, N. Kaltsoyannis and P. Mountford, *Chem. Sci.*, 2012, **3**, 819–824.



- 6 Selected recent papers about the bonding of organoactinide complexes, see: (a) T. Cantat, C. R. Graves, K. C. Jantunen, C. J. Burns, B. L. Scott, E. J. Schelter, D. E. Morris, P. J. Hay and J. L. Kiplinger, *J. Am. Chem. Soc.*, 2008, **130**, 17537–17551; (b) N. Barros, D. Maynau, L. Maron, O. Eisenstein, G. Zi and R. A. Andersen, *Organometallics*, 2007, **26**, 5059–5065; (c) A. Yahia and L. Maron, *Organometallics*, 2009, **28**, 672–679; (d) W. Ren, X. Deng, G. Zi and D.-C. Fang, *Dalton Trans.*, 2011, **40**, 9662–9664.
- 7 W. Ren, G. Zi, D.-C. Fang and M. D. Walter, *Chem.–Eur. J.*, 2011, **17**, 12669–12682.
- 8 W. Ren, G. Zi and M. D. Walter, *Organometallics*, 2012, **31**, 672–679.
- 9 W. Ren, G. Zi, D.-C. Fang and M. D. Walter, *J. Am. Chem. Soc.*, 2011, **133**, 13183–13196.
- 10 G. M. Sheldrick, *SADABS, Program for Empirical Absorption Correction of Area Detector Data*, University of Göttingen, Göttingen, Germany, 1996.
- 11 (a) G. M. Sheldrick, *SHELXL-97, Program for the Refinement of Crystal Structure from Diffraction Data*, University of Göttingen, Göttingen, Germany, 1997; (b) G. M. Sheldrick, *Acta Crystallogr.*, 2008, **A64**, 112–122.
- 12 SQUEEZE: P. V. D. Sluis and A. L. Spek, *Acta Crystallogr., Sect. A: Found. Crystallogr.*, 1990, **46**, 194–201.
- 13 M. J. Frisch, G. W. Trucks, H. B. Schlegel, G. E. Scuseria, M. A. Robb, J. R. Cheeseman, G. Scalmani, V. Barone, B. Mennucci, G. A. Petersson, H. Nakatsuji, M. Caricato, X. Li, H. P. Hratchian, A. F. Izmaylov, J. Bloino, G. Zheng, J. L. Sonnenberg, M. Hada, M. Ehara, K. Toyota, R. Fukuda, J. Hasegawa, M. Ishida, T. Nakajima, Y. Honda, O. Kitao, H. Nakai, T. Vreven, J. A. Montgomery Jr, J. E. Peralta, F. Ogliaro, M. Bearpark, J. J. Heyd, E. Brothers, K. N. Kudin, V. N. Staroverov, R. Kobayashi, J. Normand, K. Raghavachari, A. Rendell, J. C. Burant, S. S. Iyengar, J. Tomasi, M. Cossi, N. Rega, J. M. Millam, M. Klene, J. E. Knox, J. B. Cross, V. Bakken, C. Adamo, J. Jaramillo, R. Gomperts, R. E. Stratmann, O. Yazyev, A. J. Austin, R. Cammi, C. Pomelli, J. W. Ochterski, R. L. Martin, K. Morokuma, V. G. Zakrzewski, G. A. Voth, P. Salvador, J. J. Dannenberg, S. Dapprich, A. D. Daniels, O. Farkas, J. B. Foresman, J. V. Ortiz, J. Cioslowski and D. J. Fox, *Gaussian 09, Revision A.02*, Gaussian, Inc., Wallingford CT, 2009.
- 14 S. Grimme, J. Antony, S. Ehrlich and H. Krieg, *J. Chem. Phys.*, 2010, **132**, 154104.
- 15 W. Kuchle, M. Dolg, H. Stoll and H. Preuss, *Mol. Phys.*, 1991, **74**, 1245–1263.
- 16 W. Ren, W. W. Lukens, G. Zi, L. Maron and M. D. Walter, *Chem. Sci.*, 2013, **4**, 1168–1174.
- 17 D. L. Clark, S. K. Grumbine, B. L. Scott and J. G. Watkin, *Organometallics*, 1996, **15**, 949–957.
- 18 R. W. Broach, A. J. Schultz, J. M. Williams, G. M. Brown, J. M. Manriquez, P. J. Fagan and T. J. Marks, *Science*, 1979, **203**, 172–174.
- 19 W. J. Evans, G. W. Nyce and J. W. Ziller, *Organometallics*, 2001, **20**, 5489–5491.
- 20 W. Ren, N. Zhao, L. Chen, H. Song and G. Zi, *Inorg. Chem. Commun.*, 2011, **14**, 1838–1841.
- 21 N. A. Siladke, C. L. Webster, J. R. Walensky, M. K. Takase, J. W. Ziller, D. J. Grant, L. Gagliardi and W. J. Evans, *Organometallics*, 2013, **32**, 6522–6531.
- 22 L. A. Seaman, E. A. Pedrick, T. Tsuchiya, G. Wu, E. Jakubikova and T. W. Hayton, *Angew. Chem., Int. Ed.*, 2013, **52**, 10589–10592.
- 23 I. Korobkov, B. Vidjayacoumar, S. I. Gorelsky, P. Billone and S. Gambarotta, *Organometallics*, 2010, **29**, 692–702.
- 24 I. Korobkov, A. Arunachalampillai and S. Gambarotta, *Organometallics*, 2004, **23**, 6248–6252.
- 25 Selected papers on the reaction of group 4 imido complexes with organic imines, see: (a) K. E. Meyer, P. J. Walsh and R. G. Bergman, *J. Am. Chem. Soc.*, 1994, **116**, 2669–2670; (b) K. E. Meyer, P. J. Walsh and R. G. Bergman, *J. Am. Chem. Soc.*, 1995, **117**, 974–985; (c) R. L. Zuckerman and R. G. Bergman, *Organometallics*, 2000, **19**, 4795–4809; (d) A. E. Guiducci, C. L. Boyd and P. Mountford, *Organometallics*, 2006, **25**, 1167–1187; (e) P. D. Schweizer, H. Wadepohl and L. H. Gade, *Organometallics*, 2013, **32**, 3697–3709; (f) J. Unruangsri, H. Morgan, A. D. Schwarz, A. D. Schofield and P. Mountford, *Organometallics*, 2013, **32**, 3091–3107.

