

Dicyanopyrazine-derived push–pull chromophores
for highly efficient photoredox catalysis†Cite this: *RSC Adv.*, 2014, 4, 30062Yu Zhao,^{‡a} Chenhao Zhang,^{‡a} Kek Foo Chin,^c Oldřich Pytela,^b Guo Wei,^a
Hongjun Liu,^a Filip Bureš^{*b} and Zhiyong Jiang^{*a}

Here, we report dicyanopyrazine (DPZ)-derived push–pull chromophores, easily prepared and tunable organic compounds, as new kinds of photoredox catalysts. In particular, the DPZ derivative H, containing 2-methoxythienyl as electron-donating moiety, exhibits a broad absorption of visible light with an absorption edge up to 500 nm and excellent redox properties, and has been demonstrated as a desirably active and efficient photoredox catalyst in four challenging kinds of photoredox reactions. The amount of catalyst in most reactions is less than 0.1 mol% and even 0.01 mol%, representing the lowest catalyst loading in the current photoredox organocatalysis.

Received 15th May 2014
Accepted 30th June 2014

DOI: 10.1039/c4ra05525j

www.rsc.org/advances

Introduction

Towards designing new reactions, mild and green conditions are always favoured to minimize energy consumption on the standpoint of green chemistry.¹ In this context, photoredox catalysis (photo-induced single electron transfer, SET), which utilizes visible light as a naturally abundant and clean energy source and does not need any special instrumentation nor apparatus, has attracted considerable attention from synthetic organic chemists.^{2–8} In particular, the visible light absorbing photoredox catalysts, which are able to transform solar to chemical energy, play the most important role in photocatalytic reactions.

In recent years, various readily available compounds, such as polypyridyl complexes of ruthenium^{9,10} and iridium^{11,12} as well as organic dyes,^{13,14} have been utilized as the photoredox catalysts. Other designed photocatalysts with tailored structures include platinum(II),¹⁵ gold(III),¹⁶ palladium(II)¹⁷ and copper(II)^{18,19} complexes and acridinium ions^{20,21} (Fig. 1). Obviously, the photoredox catalyst simultaneously equipped with the following features still remains highly desirable: (1) a readily available/prepared organic compound. The organic compound should

more likely reduce the cost and the potential contamination of target products with toxic heavy metals;¹⁴ (2) tunable structure. The easily modified structure of photocatalyst will benefit to optimize the photophysical properties, such as the absorption edge and intensity in the visible range as well as the redox potentials. In comparison with the well-established metal complexes, the organic photoredox catalyst with tunable structure is still deficient;^{20–22} (3) high activity and efficiency. A universal substance able to catalyse a variety of challenging and unprecedented photoredox reactions is always desirable, which is determined by the activity of photocatalyst. Meanwhile, the amount of the catalyst loaded in the reaction is always an important index to evaluate its efficiency. To the best of our knowledge, few examples^{23–25} with less than 0.1 mol% of the catalyst were presented to date unless using high power light sources^{17,26} or in flow.^{27–29} Most important, the example of photoredox reaction with less than 0.1 mol% organic catalyst is not reported. In principle, the activity and efficiency of photocatalyst could be improved by optimizing its photophysical properties.

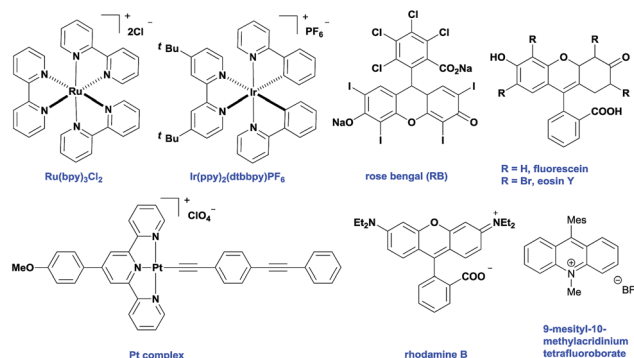


Fig. 1 Selected examples of photoredox catalysts reported to date.

^aKey Laboratory of Natural Medicine and Immuno-Engineering of Henan Province, Henan University, Kaifeng, Henan, 475004, P. R. China. E-mail: chmjzy@henu.edu.cn; filip.bures@upce.cz

^bInstitute of Organic Chemistry and Technology, University of Pardubice, Faculty of Chemical Technology, Studentská 573, Pardubice, 53210, Czech Republic

^cDivision of Chemistry and Biological Chemistry, Nanyang Technological University, 21 Nanyang Link, 637371, Singapore

† Electronic supplementary information (ESI) available: Procedures on the preparation of DPZ derivatives, quantum chemical calculations, procedures for photoredox catalysed reactions, characterization data of data and NMR spectra. See DOI: 10.1039/c4ra05525j

‡ These authors contributed equally to this work.

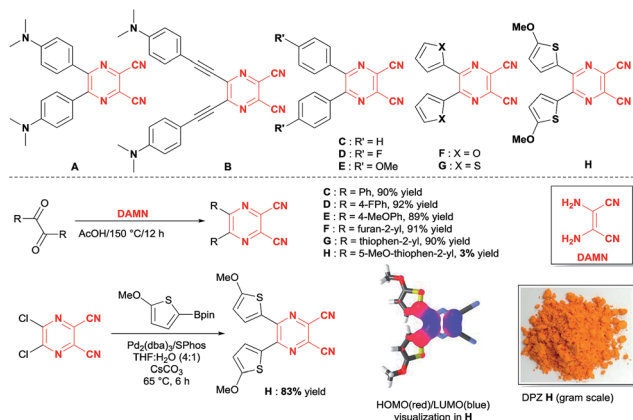


Fig. 2 The structures and convenient/improved preparations of DPZ derivatives A–H.

In 2012, push–pull molecules based on 5,6-disubstituted pyrazine-2,3-dicarbonitrile (dicyanopyrazine, DPZ) were presented and their nonlinear optical properties were investigated by our group (e.g. DPZ derivatives A–B, Fig. 2).³⁰ These organic compounds can be readily prepared from available starting materials such as diaminomaleodinitrile (DAMN) in excellent yields and represent stable D– π –A chromophores with intramolecular charge-transfer (ICT), which resembles the MLCT (metal to ligand charge transfer) in aforementioned metal complexes. Furthermore, their redox potentials are similar to known photoredox catalysts (e.g. $E_{1/2(\text{red})}$ of eosin Y: -1.0 V, rhodamine B: -0.8 V, A: -1.24 V, B: -0.87 V vs. saturated calomel electrode [SCE]) while the absorption maxima locate at the visible light area (λ_{max} of A: 471 nm, B: 499 nm). Most importantly, their optical and electrochemical properties can readily be tuned by the variation of the appended electron-donors. Based on these features, we envisioned that DPZ derivatives could probably serve as aforementioned photoredox catalysts.

Results and discussion

Preparation and properties of DPZ

In this respect, DPZ derivatives C–H were designed and synthesized, possessing either electron withdrawing or

donating (hetero)aromates at the pyrazine C5/C6 positions such as in C–D or E–H (Fig. 2). The method used for the construction of dicyanopyrazine moiety involves condensation of DAMN with 1,2-dicarbonyl compounds (benzils), affording DPZ derivatives (C–G) in 89–92% yields. However, 5-methoxythienyl, as the most electron-rich moiety, made this reaction sluggish and only 3% yield of H was obtained. Our previous findings³⁰ prompted us to attempt twofold cross-coupling reaction of 5,6-dichloropyrazine-2,3-dicarbonitrile with 5-methoxythiophen-2-ylboronic acid pinacol ester (Fig. 2). Under the optimized conditions (see the ESI†), this reaction afforded the desired product H in 83% yield. Notably, the preparation of H in a gram scale could conveniently be carried out *via* this efficient protocol (Fig. 2).

The properties of DPZ derivatives C–H were further studied by electrochemical measurements, absorption spectra and DFT calculations (Table 1, Fig. 3 and ESI†). The HOMO–LUMO energies, their differences and the position of the longest-wavelength absorption maxima can be considered as the most important electronic parameters. The first oxidation potential respective the HOMO energy decrease with increasing electron donating ability of the substituents and thus enhanced ICT. As a result, the electrochemical gap (E_g) decreases from 3.70 to 2.82 eV. Likewise, the longest-wavelength absorption maxima λ_{max} shift bathochromically (Fig. 3) and the respective optical as well as the calculated gaps E_g^{DFT} decrease significantly. It is noteworthy, that electrochemical, optical and calculated gaps show the same trend and correlate very tightly (see ESI†). Hence, we can consider both experiments and DFT calculation as reliable tools describing the electronic parameters of C–H. As can be seen, the electronic parameters of DPZ derivative H are exceptional. In contrast to G ($E_g = 3.27$ eV; $\lambda_{\text{max}} = 391$ nm (3.17 eV), $E_g^{\text{DFT}} = 3.37$ eV, $\mu = 12.88$ D), two additional methoxy substituents in H lowered the electrochemical/calculated gap by 0.45/0.39 eV, shifted the λ_{max} bathochromically by 57 nm and increased the molar absorption coefficient as well as the ground state dipole moment. Obviously, the electronic properties of DPZ derivative H ($E_g = 2.82$ eV, $\lambda_{\text{max}} = 448$ nm (2.78 eV), $\epsilon = 21\,500$ mol^{−1} dm³ cm^{−1}, $E_g^{\text{DFT}} = 2.97$ eV, $\mu = 18.26$ D) are well tuned and suited for the photoredox catalysis (e.g. Ru(bpy)₃Cl₂, $\lambda_{\text{max}} = 452$ nm, $E_g = 2.62$ eV). As expected, the HOMO and the LUMO in H are clearly separated and localized on the donor and acceptor parts of the molecule (Fig. 2, for C–G see ESI†).

Table 1 Electrochemical, optical and calculated data for DPZs C–H

DPZ	$E_{1/2(\text{ox})}^a$ [V]	$E_{1/2(\text{red})}^a$ [V]	E_{HOMO}^b [eV]	E_{LUMO}^b [eV]	E_g^c [eV]	λ_{max} [nm (eV)]/ $\epsilon \times 10^{-3}^d$ [mol ^{−1} dm ³ cm ^{−1}]	E_g^{DFTe} [eV]	μ^e [D]
C	+2.30	−1.40	−7.10	−3.41	3.70	340(3.65)/15.3	3.93	11.38
D	+2.27	−1.41	−7.07	−3.39	3.68	345(3.59)/13.2	3.92	7.43
E	+1.90	−1.44	−6.70	−3.36	3.34	383(3.24)/19.7	3.58	10.80
F	+1.95	−1.36	−6.75	−3.44	3.31	379(3.27)/17.4	3.54	12.69
G	+1.92	−1.35	−6.72	−3.45	3.27	391(3.17)/14.6	3.37	12.88
H	+1.37	−1.45	−6.17	−3.35	2.82	448(2.78)/21.5	2.97	18.26

^a $E_{1/2(\text{ox})}/E_{1/2(\text{red})}$ are half-wave potentials of the first oxidation/reduction measured in DCM versus SCE. ^b $E_{\text{HOMO/LUMO}} = E_{1/2(\text{ox/red})} + 4.80$.³¹
^c $E_{\text{HOMO}}-E_{\text{LUMO}}$ electrochemical gap. ^d Longest-wavelength absorption maxima (optical gap)/molar absorption coefficient (DCM). ^e DFT calculated HOMO–LUMO gap and ground state dipole moment (B3LYP/6-311++G(2df,p)//B3LYP/6-311++G(2df,p), scrf = (solvent = dichloromethane)).





Next, we screened a series of CDC reactions between different tertiary/secondary amines and various nucleophiles under the established conditions (28 °C, DMF as solvent, ambient air and irradiation with 9 W fluorescent lamp, Table 3,

^a Detailed conditions for these reactions are described in the ESI.[†] ^b (i) 0.05 mol% of **H** was used, 9 W fluorescent lamp, 28 °C, *t* = 12 h, yield = 86%; (ii) 0.01 mol% of **H** was used, 9 W fluorescent lamp, 28 °C, *t* = 24 h, yield = 85%; (iii) 0.05 mol% of **H** was used, sunlight (30 °C), *t* = 3.5 h, yield = 83%; (iv) 0.01 mol% of **H** was used, sunlight (30 °C), *t* = 3.5 h, yield = 60%. ^c 0.15 mmol scale, DMF as solvent.

entries 10–15). The reaction^{27,35,36,41–43} of *N*-phenyl-tetrahydroisoquinoline **1a** and TMSCN **2d** was firstly processed in the presence of 0.1 mol% of **H**, and the oxidative Strecker adduct **3j** was obtained within 2 hours in 87% yield (Table 3, entry 10). With the same catalyst loading and the help of pyrrolidine/TFA, the powerful catalytic ability of **H** was further demonstrated by the reaction^{16,33,36,38,40,44} of **1a** and acetone **2e**, affording Mannich adduct **3k** in 93% yield after 4 hours (Table 3, entry 11). To introduce a challenging hetero-quaternary stereogenic center adjacent to *N*-aryl-tetrahydroisoquinolines, we attempted the first example of oxidative Mannich reaction of **1a** and 5*H*-oxazol-4-one **2f** via photoredox catalysis (Table 3, entry 12). It was discovered that, in the presence of only 0.01 mol% of **H**, the desired adduct **3l** was obtained in 82% yield with 10 : 1 dr after 4 hours. Use of 0.01 mol% of **H** as catalyst also resulted in the complete reaction^{34,36,45} between **1a** and diethyl phosphonate **2g**, to furnish the P–C bond formation and afford the product **3m** in 81% yield within 7 hours (Table 3, entry 13).

Over the past few years, the CDC reaction^{32,34,40} of *N*-arylpiperidines and nitromethane has been attempted in photoredox catalysis under different conditions; however, the reported reactions were always sluggish with poor yields, and *N*-arylpiperidines were thus recognized as the non-activated amines. Hence, we were intrigued to conduct the formidable reaction of *N*-(4-*tert*-butylphenyl)piperidine **4** and nitromethane **1** (Table 3, entry 14). We were pleased that the desired adduct **5** could be obtained in 76% yield after 45 hours when 2.0 mol% of **H** was used. Due to the longest-wavelength absorption maxima of **H** reaching blue light area, the reaction was also tested under the irradiation of blue Light Emitting Diode (LED, 4 W, household), and completed in 15 hours to deliver adduct **5** in 80% yield. On the other hand, linear secondary amines exhibit very poor activity in photocatalytic CDC reaction, and the singlet oxygen protocol combined with a special photoreactor seems necessary.²⁹ Accordingly, we turned our attention to the reaction of dibenzylamine **6** and TMSCN **2d** (Table 3, entry 15). It was found that the reaction could be finished smoothly in 17 hours by using 0.1 mol% of **H**, to afford the adduct **7** in 84% yield.

Mechanism of the photochemical CDC reactions

We were also interested in fundamental principles of DPZ action in photoredox catalyzed reactions. In the above CDC reactions, H₂O₂ can be envisaged as a reasonable by-product. As Fig. 4 shown, hydrogen peroxide could clearly be detected in the crude reaction mixture of **1a** and **2a** catalysed by **H**. The presence of H₂O₂ was confirmed by the iodide test (KI in glacial acetic acid), in which a colour change to dark red was detected.

Considering three plausible pathways, such as singlet-oxygen oxidation, photoredox reaction or both of these two reactions, we subsequently conducted the study on solvent effects since the lifetime of singlet oxygen highly depends on the solvent used. The singlet oxygen lifetime should be significantly longer in deuterated solvents in comparison to their protonated counterparts.^{46–48} Therefore, a reaction of **1a** and **2a** was performed in the presence of 0.1 mol% of DPZ derivative **H** at 25 °C in both

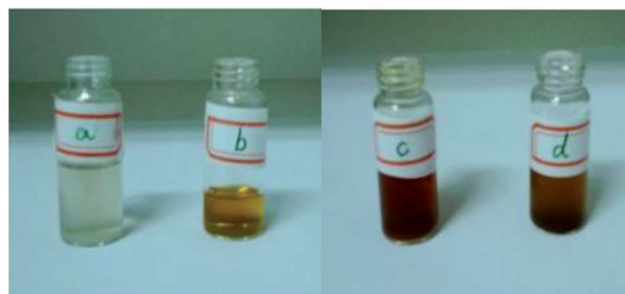


Fig. 4 Test for the production of hydrogen peroxide in the reaction mixture of **1a** and **2a**. (a) KI and starch in glacial acetic acid, colourless. (b) The reaction mixture of **1a** and **2a** after reaction completed, brown. (c) Addition of solution (b) to solution (a), dark red. (d) Addition of 30% of H₂O₂ in solution (a). KI (0.2 M), aqueous acetic acid (0.2 M), starch (4 mg mL⁻¹), DPZ **H** (0.1 μmol, 14.14 μL, 2.5 mg mL⁻¹ in CH₃NO₂), **1a** (0.1 mmol), CH₃NO₂ (1.0 mL), reaction time: 5 h, 28 °C.

CH₃CN and CD₃CN solvent. The results of the kinetic measurements of the product conversions are indicated in Fig. 5. From these dependencies we can conclude that singlet oxygen was likely not a key participant in the reaction as no significant increase in the reaction rate was observed in CD₃CN. Mechanistically, we suggest a plausible mechanism on the above CDC reactions with DPZ derivative **H** as catalyst (Fig. 6), which is a photoredox protocol as Stephenson previously disclosed for the ruthenium or iridium complexes.³² This reaction mechanism involves excitation of the DPZ **H** catalyst by visible light, reductive quench (single-electron transfer, SET) from the amino-substrate to form radical cation/anion R₃N^{•+} and DPZ **H**^{•-}. Subsequent reaction with the nitroalkane or eventually oxygen ([O]) regenerates the **H**-catalyst and formed [O]^{•-} deprotonates the trialkylammonium radical cation R₃N^{•+} to iminium salt, which undergoes nucleophilic attack of the nucleophile used.

On the basis of the above observations and the proposed mechanism, we can deduce the following fundamental structure–catalytic activity relationships:

- Bathochromically shifted absorption maxima (λ_{max}), high molar absorption coefficient (ε) and small HOMO–LUMO gap

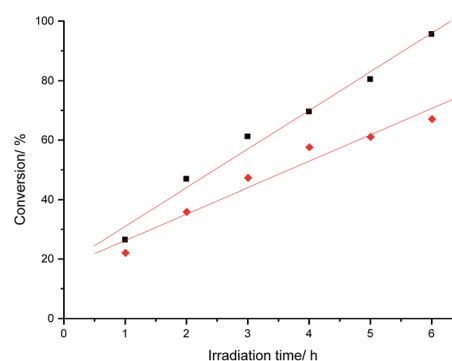


Fig. 5 Photocatalytic CDC reaction of **1a** and **2a** (or **2a**-d₃) to give adduct **3a** in the presence of DPZ derivative **H**; reaction kinetics was monitored by using ¹H NMR spectroscopy; CH₃NO₂ (■), CD₃NO₂ (◆); DPZ **H** (0.05 μmol, 7.0 μL, 2.5 mg mL⁻¹), **1a** (0.5 mmol), CH₃NO₂ or CD₃NO₂ (0.5 mL), reaction time: 6 h, 25 °C.



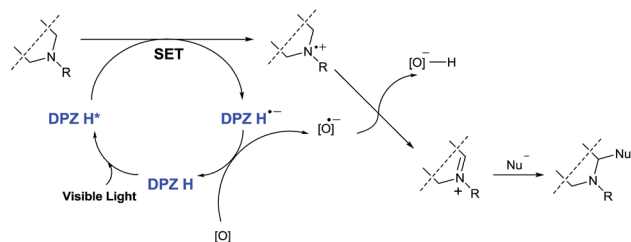


Fig. 6 Plausible mechanism of CDC reaction catalyzed by DPZ H.

(E_g) of the catalyst improve its excitation ability by low-energy visible light (compare for **H** and **C-G**, Table 1).

- Principal changes in the properties of DPZ derivatives **C-H** are caused by the donor-centered HOMO (E_{HOMO} range from -7.10 to -6.17 eV, Table 1).

- The LUMO localized on the dicyanopyrazine acceptor remained almost unaltered in all derivatives (E_{LUMO} range from -3.45 to -3.39 eV, Table 1).

- Reduced HOMO–LUMO gap of the catalyst facilitates the generation of radicals (SET, Fig. 6).

- The conversion of CDC reaction depends linearly and tightly on the electrochemical as well as optical HOMO–LUMO gaps of the used X-shaped push–pull DPZs (see the correlation in Fig. S12 in the ESI†).

- Planar and polarizable π -system of the catalyst help to stabilize the radical anion *e.g.* DPZ $\text{H}^{\bullet-}$ (Fig. 6) through delocalization.

- Well-tuned ICT helps to reduce HOMO–LUMO gap (see charge separation in **H**, Fig. 2) but does not cause high electron saturation of the cyano acceptors.

- Higher electron saturation of the cyano groups would have detrimental effect on the formation of the radical anion *e.g.* DPZ $\text{H}^{\bullet-}$ (Fig. 6).

- These suppositions can be demonstrated as follows: DPZ derivatives bearing strong electron NMe_2 donors **A** ($E_g = 2.26$ eV) and **B** ($E_g = 1.91$ eV) showed lower E_g values and more bathochromically shifted absorption maxima than **H**, but their catalytic activity is very low (Table 2). This is most likely caused by high D–A coupling between NMe_2 and CN groups and thus higher contribution of the quinoid/zwitterionic form with negatively charged cyano groups (see Fig. S13 in the ESI†).

H catalyses other challenging reactions

To further evaluate the performance of DPZ derivative **H**, other challenging or unprecedented photoredox reactions were successively attempted (Fig. 7). We firstly carried out an oxidation of *N*-methyl-2-(phenylamino)acetamide **8** using 1.0 mol% of **H** in DMF at 28 °C under oxygen atmosphere and irradiation by a 9 W fluorescent lamp (eqn (1)). We found that the reaction was finished within 7 hours to give unsymmetric oxalamide **9**. Indeed, despite the well-established photocatalytic oxidation of tertiary amines,^{49,50} the transformation of activated secondary amine to amide has not been reported yet.

In 2012, Jørgensen and Xiao and co-workers introduced the first oxidative hydroxylation of aryl boronic acids by utilizing

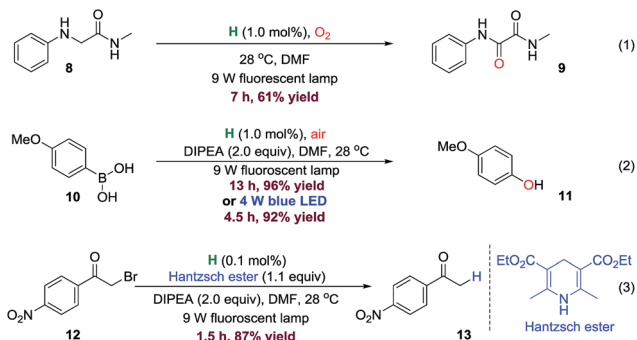


Fig. 7 Photocatalytic oxidation, oxidative hydroxylation and reductive dehalogenation.

2.0 mol% of Ru complex as catalyst, in which 36 W fluorescence lamp was used as light source.⁵¹ Moreover, Acid Red 87, the organic compound, has also been demonstrated as the photo-redox catalyst, but the reaction became slower.⁵¹ Scaiano group⁵² gave a metal-free example of oxidative hydroxylation of aryl boronic acids with methylene blue as catalyst; the catalyst amount could be decreased to 1.0 mol% and the reactions became faster, but two 90 W warm white LEDs as the light source and pure oxygen atmosphere were necessary. Therefore, we investigated this demanding reaction with the DPZ derivative **H** as catalyst. As eqn (2) shown (Fig. 7), the oxidative hydroxylation of 4-methoxyphenylboronic acid **10** was performed with 1.0 mol% of **H** in ambient air and irradiated by a 9 W fluorescent lamp, and the desired product **11** was obtained in 96% yield after 13 hours. The reaction could be accelerated by utilizing 4 W LEDs as the light source, in which the product **11** could be obtained in 92% yield within 4.5 hours.

We subsequently explored the reductive dehalogenation of α -bromoacetophenones. In comparison with the excellent work by Zeitler and co-workers [2.5 mol% eosin Y, 1 W high-power LEDs ($\lambda = 530$ nm), 18 h, 83% yield],¹⁴ the reductive dehalogenation of bromoacetophenone **12** could be finished within 1.5 hours in the presence of 0.1 mol% of **H** under the irradiation by a 9 W house-hold fluorescent lamp, to access the product **13** in 87% yield (Fig. 7, eqn (3)).

Conclusion

Dicyanopyrazine-derived (DPZ) push–pull chromophores have been demonstrated as a new kind of organic catalysts for photoredox catalysis. 2-Methoxythienyl as an electron-donating moiety provides a DPZ derivative (**H**) with a broad absorption edge reaching up to 500 nm as well as the excellent redox properties. Moreover, all the aforementioned features are well balanced and enable the DPZ derivative **H** to become the desirably active and efficient photoredox catalyst, which was demonstrated by more than four kinds of photocatalytic organic reactions. Without the aid of high power light sources or in flow, the amount of catalyst in most reactions is always less than 0.1 mol% and even 0.01 mol%, which represents the lowest catalyst loading in the current photoredox organocatalysis.



Other excellent performances on the DPZ derivative H-catalyzed reactions include lower power light source, air atmosphere while short reaction time and higher yields. Further investigations, which involve the use of DPZ derivatives as catalyst in novel photocatalytic reactions, are ongoing in our labs and will be reported in due course.

Acknowledgements

This research was supported by the Czech Science Foundation (13-01061S), NSFC (nos. 21072044, U1204207) and the Program for New Century Excellent Talents in University of Ministry of Education (NCET-11-0938).

Notes and references

- P. Tundo, P. Anastas, D. S. Black, J. Breen, T. Collins, S. Memoli, J. Miyamoto, M. Polyakoff and W. Tumas, *Pure Appl. Chem.*, 2009, **72**, 1207.
- J. M. R. Narayanam and C. R. J. Stephenson, *Chem. Soc. Rev.*, 2011, **40**, 102.
- M. A. Ischay and T. P. Yoon, *Eur. J. Org. Chem.*, 2012, 3359.
- L. Shi and W. Xia, *Chem. Soc. Rev.*, 2012, **41**, 7687.
- J. Xuan and W.-J. Xiao, *Angew. Chem., Int. Ed.*, 2012, **51**, 6828.
- C. K. Prier, D. A. Rankic and D. W. C. MacMillan, *Chem. Rev.*, 2013, **113**, 5322.
- J. Zhao, W. Wu, J. Sun and S. Guo, *Chem. Soc. Rev.*, 2013, **42**, 5323.
- X. Lang, X. Chen and J. Zhao, *Chem. Soc. Rev.*, 2014, **43**, 473.
- D. A. Nicewicz and D. W. C. MacMillan, *Science*, 2008, **322**, 77.
- D. A. Nagib and D. W. C. MacMillan, *Nature*, 2011, **480**, 224.
- A. McNally, C. K. Prier and D. W. C. MacMillan, *Science*, 2011, **334**, 1114.
- M. T. Pirnot, D. A. Rankic, D. B. C. Martin and D. W. C. MacMillan, *Science*, 2013, **339**, 1593.
- H. Liu, W. Feng, C. W. Kee, Y. Zhao, D. Leow, Y. Pan and C.-H. Tan, *Green Chem.*, 2010, **12**, 953.
- M. Neumann, S. Földner, B. König and K. Zeitler, *Angew. Chem., Int. Ed.*, 2011, **50**, 951.
- J.-J. Zhong, Q.-Y. Meng, G.-X. Wang, Q. Liu, B. Chen, K. Feng, C.-H. Tung and L.-Z. Wu, *Chem.-Eur. J.*, 2013, **19**, 6443.
- Q. Xue, J. Xie, H. Jin, Y. Cheng and C. Zhu, *Org. Biomol. Chem.*, 2013, **11**, 1606.
- W.-P. To, Y. Liu, T.-C. Lau and C.-M. Che, *Chem.-Eur. J.*, 2013, **19**, 5654.
- J.-M. Kern and J.-P. Sauvage, *J. Chem. Soc., Chem. Commun.*, 1987, 546.
- M. Pirtsch, S. Paria, T. Matsuno, H. Isobe and O. Reiser, *Chem.-Eur. J.*, 2012, **18**, 7336.
- S. Fukuzumi, H. Kotani, K. Ohkubo, S. Ogo, N. V. Tkachenko and H. Lemmetyinen, *J. Am. Chem. Soc.*, 2004, **126**, 1600.
- T. M. Nguyen and D. A. Nicewicz, *J. Am. Chem. Soc.*, 2013, **135**, 9588.
- S. Guo, H. Zhang, L. Huang, Z. Guo, G. Xiong and J. Zhao, *Chem. Commun.*, 2013, **49**, 8689.
- J. M. R. Narayanam, J. W. Tucker and C. R. J. Stephenson, *J. Am. Chem. Soc.*, 2009, **131**, 8756.
- J. D. Nguyen, E. M. D'Amato, J. M. R. Narayanam and C. R. J. Stephenson, *Nat. Chem.*, 2012, **4**, 854.
- M. A. Ischay, M. S. Ament and T. P. Yoon, *Chem. Sci.*, 2012, **3**, 2807.
- S. Samanta, S. Das and P. Biswas, *J. Org. Chem.*, 2013, **78**, 11184.
- J. W. Tucker, Y. Zhang, T. F. Jamison and C. R. J. Stephenson, *Angew. Chem., Int. Ed.*, 2012, **51**, 4144.
- M. Neumann and K. Zeitler, *Org. Lett.*, 2012, **14**, 2658.
- D. B. Ushakov, K. Gilmore, D. Kopetzki, D. T. McQuade and P. H. Seeberger, *Angew. Chem., Int. Ed.*, 2014, **53**, 557.
- F. Bureš, H. Čermáková, J. Kulhánek, M. Ludwig, W. Kuznik, I. V. Kityk, T. Mikysek and A. Růžicka, *Eur. J. Org. Chem.*, 2012, 529.
- B. W. D'Andrade, S. Datta, S. R. Forrest, P. Djurovich, E. Polikarpov and M. E. Thompson, *Org. Electron.*, 2005, **6**, 11.
- A. G. Condie, J. C. González-Gómez and C. R. J. Stephenson, *J. Am. Chem. Soc.*, 2010, **132**, 1464.
- Y. Pan, C. W. Kee, L. Chen and C.-H. Tan, *Green Chem.*, 2011, **13**, 2682.
- D. P. Hari and B. König, *Org. Lett.*, 2011, **13**, 3852.
- Z. Xie, C. Wang, K. E. deKrafft and W. Lin, *J. Am. Chem. Soc.*, 2011, **133**, 2056.
- M. Rueping, J. Zoller, D. C. Fabry, K. Poschary, R. M. Koenigs, T. E. Weirich and J. Mayer, *Chem.-Eur. J.*, 2012, **18**, 3478.
- J.-L. Wang, C. Wang, K. E. deKrafft and W. Lin, *ACS Catal.*, 2012, **2**, 417.
- Q. Liu, Y.-N. Li, H.-H. Zhang, B. Chen, C.-H. Tung and L.-Z. Wu, *Chem.-Eur. J.*, 2012, **18**, 620.
- M. Neumann and K. Zeitler, *Org. Lett.*, 2012, **14**, 2658.
- L. Möhlmann, M. Baar, J. Rieß, M. Antonietti, X. Wang and S. Blechert, *Adv. Synth. Catal.*, 2012, **354**, 1909.
- M. Rueping, C. Vila and T. Bootwicha, *ACS Catal.*, 2013, **3**, 1676.
- Y. Pan, S. Wang, C. W. Kee, E. Dubuisson, Y. Yang, K. P. Loh and C.-H. Tan, *Green Chem.*, 2011, **13**, 3341.
- M. Rueping, S. Zhu and R. M. Koenigs, *Chem. Commun.*, 2011, **47**, 12709.
- M. Rueping, C. Vila, R. M. Koenigs, K. Poschary and D. C. Fabry, *Chem. Commun.*, 2011, **47**, 2360.
- M. Rueping, S. Zhu and R. M. Koenigs, *Chem. Commun.*, 2011, **47**, 8679.
- J. R. Hurst, J. D. McDonald and G. B. Schuster, *J. Am. Chem. Soc.*, 1982, **104**, 2065.
- P. R. Ogilby and C. S. Foote, *J. Am. Chem. Soc.*, 1983, **105**, 3423.
- A. G. Griesbeck, V. Schlundt and J. M. Neudörfl, *RSC Adv.*, 2013, **3**, 7265.
- P. Kohls, D. Jadhav, G. Pandey and O. Reiser, *Org. Lett.*, 2012, **14**, 672.
- J. H. Park, K. C. Ko, E. Kim, N. Park, J. H. Ko, D. H. Ryu, T. K. Ahn, J. Y. Lee and S. U. Son, *Org. Lett.*, 2012, **14**, 5502.
- Y.-Q. Zou, J.-R. Chen, X.-P. Liu, L.-Q. Lu, R. L. Davis, K. A. Jørgensen and W.-J. Xiao, *Angew. Chem., Int. Ed.*, 2012, **51**, 784.
- S. P. Pitre, C. D. McTiernan, H. Ismaili and J. C. Scaiano, *J. Am. Chem. Soc.*, 2013, **135**, 13286.

