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Exceptional CO_2 capture in a hierarchically porous carbon with simultaneous high surface area and pore volume†

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A new type of hierarchically porous carbon (HPC) structures of simultaneously high surface area and high pore volume has been synthesised from carefully controlled carbonization of in-house optimised metal–organic frameworks (MOFs). Changes in synthesis conditions lead to millimetre-sized MOF-5 crystals in a high yield. Subsequent carbonization of the MOFs yield HPCs with simultaneously high surface area, up to 2734 $m^2\ g^{-1}$, and exceptionally high total pore volume, up to 5.53 $cm^3\ g^{-1}$. In the HPCs, micropores are mostly retained and meso- and macro- pores are generated from defects in the individual crystals, which is made possible by structural inheritance from the MOF precursor. The resulting HPCs show a significant amount of CO_2 adsorption, over 27 mmol g^{-1} (119 wt%) at 30 bar and 27 °C, which is one of the highest values reported in the literature for porous carbons. The findings are comparatively analysed with the literature. The results show great potential for the development of high capacity carbon-based sorbents for effective pre-combustion CO_2 capture and other gas and energy storage applications.

Broader context

Several energy and environmental technologies require the advance of highly stable porous nanostructures to enhance the performance of core devices, such as in chemical energy storage, purification/separation of gaseous molecules, and pollutant removal in water and air. Particularly, there is great urgency in cutting down the emission of CO_2 from concentrated sources. The adsorptive removal of carbon (CO_2) by nanoporous solids is considered one of the most promising approaches to fulfilling the short-term goals of CCS (carbon capture and sequestration). Control and optimization of pore geometry, size, volume and surface area are crucial to achieve efficient CO_2 capture at pre- or post- combustion power stations. In this study, we describe a method of obtaining highly hierarchical porous carbons (HPCs) with simultaneously high surface area (up to 2730 m^2g^{-1}) and ultra-high total pore volume (up to 5.5 cm^3g^{-1}) from selected and relatively large MOF crystals. HPCs show significantly enhanced CO_2 adsorption (*e.g.*, over 27 mmol g^{-1} at 30 bar and 27 °C) at high-pressures and temperatures up to 75 °C, compared with their MOF counterparts and other porous carbons reported in the literature. Given their rich pore property, the HPCs are also potential candidates for further pore optimization/functionalization for energy storage, *e.g.* as electrode materials and methane storage.

As fossil fuels will continue to dominate the ever increasing energy demand in the coming decades, CO_2 emissions are likely to rise considerably.[1,2] Thus there is a great urgency in cutting down the emissions by carbon (CO_2) capture at anthropogenic point sources[1–4] and effective sorbent materials and methods are highly needed.[2–6] A wide spectrum of materials has been considered for carbon capture by adsorption and/or separation.[5–11] Recently, the physical adsorptive removal of CO_2 by porous solids, such as porous carbons and metal–organic frameworks (MOFs) has attracted much attention.[10–23] Quite a large number of MOFs have been tailor-made and studied extensively for clean energy and environmental applications, including H_2 and CH_4 sorption, storage and separation.[12,13,24–26] Given their structural flexibility and thermal stability, light-weight porous carbons have also emerged as promising adsorbents.[11,14–23,27–41] A variety of porous carbons, such as activated,[14,15,18,21,23,31–33] templated,[16,22,29,35,38] carbide derived,[17,41] polymer derived[14,19,36,39] and recent graphene-oxide derived[11,20,28,30] carbons have been investigated with the aim to improve their energy storage (electrode materials, as well as H_2 and CH_4 storage) and carbon capture performances. The key porous parameters that determine if the adsorbents have potential in being used in practical applications are surface area and pore volume together with pore size. Thus the current research is actively focused on enhancing the pore properties of the adsorbents.

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Here, we describe the synthesis and characterisation of highly hierarchical porous carbons (HPCs) for enhanced carbon dioxide capture at high pressure and temperature, up to 30 bar and 75 °C, respectively. This new class of HPCs (HPC5, HPC74, HPC53) are derived from the direct carbonization of three different MOF structures (MOF-5, MOF-74, MIL-53, respectively).[12,24,42–44] The carbonization process of one of the MOF precursors, MOF-5 is schematically shown in Scheme 1. HPCs essentially consist of local defective carbon structures with mainly sp^2 bonding. We also show that these new materials exhibit simultaneously high surface areas, up to 2734 $m^2\ g^{-1}$, and hierarchical pores (micro, meso (>2 nm) and macro (>50 nm)) with very high total pore volumes, up to 5.53 $cm^3\ g^{-1}$, achieved by controlling synthesis conditions of the MOF precursors. Such porosity characteristics of simultaneously high surface area and pore volume are seldom found in the family of highly porous carbons. Due to the exceptional porous properties, the HPCs exhibit high CO_2 adsorption capabilities, particularly at high-pressure, compared with their parent MOFs and other high surface area porous carbons. Thus the HPCs are ideal sorbents for pre-combustion CO_2 capture.

The prototype framework structures of the MOF precursors are schematically shown in Fig. 1. In our study, all the MOFs were synthesized by the solvothermal method according to the literature and the references therein (see the ESI† for detailed experimental procedure).[42–44] As shown in Fig. 1, in the case of MOF-5, a modification in the synthesis conditions resulted in millimetre-sized crystals (batch 2) with enhanced yield (~140 wt%) compared to the micron-sized crystals from the regular method (batch 1) (Fig. S1†). The powder X-ray diffraction (XRD) patterns (Fig. S2a†), and N_2 adsorption isotherms (Fig. S2b†) measured at liquid nitrogen temperature on the MOFs reveal well ordered, and crystalline pore structures, in very good agreement with the earlier reports.[42–44] The surface area and pore volume values are summarized in Table 1.

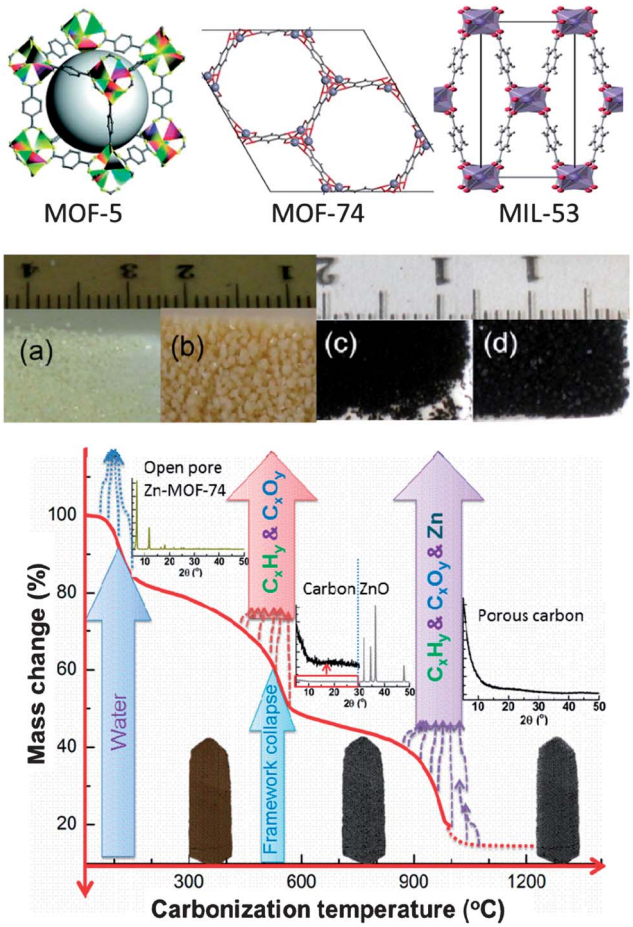
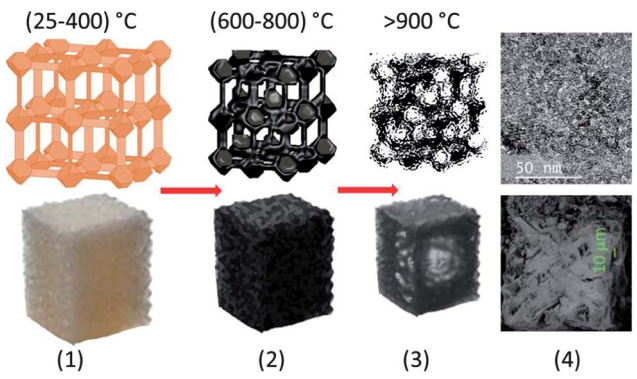


Fig. 1 Top: schematic representation of MOF-5,[42] MOF-74 and MIL-53 framework structures. Middle: digital photographs of MOF-5 and its carbon (carbonized at 1000 °C) crystals: (a) batch 1, (b) batch 2, (c) HPC5b1-1000 and (d) HPC5b2-1000. Bottom: mechanism involved in the carbonization process of Zn-MOF-74. The plot represents the MOF mass change *vs.* carbonization temperature.



Scheme 1 Schematic illustration for steps in fabricating HPCs from a precursor MOF-5; (1) MOF framework/crystal structure (up to 400 °C) with well-defined pore structure in the micropore region, (2) decomposition, and metal-oxide formation, ZnO and carbon between 600 and 800 °C yielding poor pore development, (3) at >900 °C, ZnO reduction and evaporation of Zn and CO yields highly porous carbon with a hierarchical pore structure and (4) surface morphology of HPCs by TEM & SEM.

The HPCs were obtained by the direct carbonization of MOFs at temperatures between 600 °C and 1100 °C under N_2 atmosphere. The carbonization process and mechanism for Zn-MOF-74 is schematically presented in Fig. 1 and is shown through state-of-the-art simultaneous thermogravimetric analysis (TGA) and mass spectroscopy (MS) (Fig. 2a and S3a–S3c†) for all of the MOFs. The first mass-loss, below 150 °C, is due to the evaporation of adsorbed/terminal water molecules within the pores and the second mass-loss between 400 °C and 600 °C, is attributed to the framework decomposition leading to a major release of carbon containing gaseous products (mostly CO_2, CO, C_6H_6 and a small amount of H_2 and C_xH_y hydrocarbon mixtures) and formation of metal-oxides (ZnO and Al_2O_3) occurs. The third mass-loss seen in MOF-5 and MOF-74 starting at ~900 °C is due to further release of CO_2 and CO with Zn through the reduction of ZnO by carbon *via* ZnO + C → Zn(g) + CO.

The carbonization yields materials with either a featureless XRD pattern (Fig. 2b) or very weak and broad XRD peaks between 20° and 25°, and at ~45° of two-theta (Fig. S4†),

Table 1 The surface area, total pore volume, and CO_2 adsorption performance of HPCs and the MOF precursors at PSA (pressure swing adsorption) and VSA (vacuum swing adsorption) conditions for pure CO_2 (PSA: between 6 bar and 1 bar, VSA: between 1.5 bar and 0.05 bar) and flue gas CO_2 (20%; PSA: between 1.2 bar and 0.2 bar, VSA: between 0.3 bar and 0.001 bar). All are obtained from the 27 °C CO_2 adsorption isotherms

Sample	BET surface area ($m^2\ g^{-1}$)	Total pore volume ($cm^3\ g^{-1}$)	Pure CO_2		Flue CO_2	
			PSA (mmol g^{-1})	VSA (mmol g^{-1})	PSA (mmol g^{-1})	VSA (mmol g^{-1})
HPC5b2-1100	2734	5.23	7.33	3.78	2.44	1.4
HPC5b2-1000	2517	5.53	7.01	3.93	2.52	1.5
HPC5b2-900	2495	5.41	6.97	4.00	2.60	1.5
HPC5b1-1000	2462	3.79	6.62	3.79	2.49	1.4
HPC5b1-900	2109	3.26	5.36	3.23	2.12	1.4
HPC74-1000	1900	1.00	5.59	3.88	2.52	1.6
HPC74-900	1828	0.95	5.37	3.76	2.37	1.5
HPC74-800	1574	1.48	4.50	3.67	2.29	1.7
HPC74-600	1269	1.31	4.82	3.12	2.00	1.5
HPC53-1000	1260	2.47	3.49	1.96	1.29	0.7
MOF-5-b1	3224	1.31	4.52	1.16	0.77	0.2
MOF-74	1000	0.40	3.83	5.60	3.67	2.8
MIL-53	1193	0.44	2.57	1.98	1.42	1.2

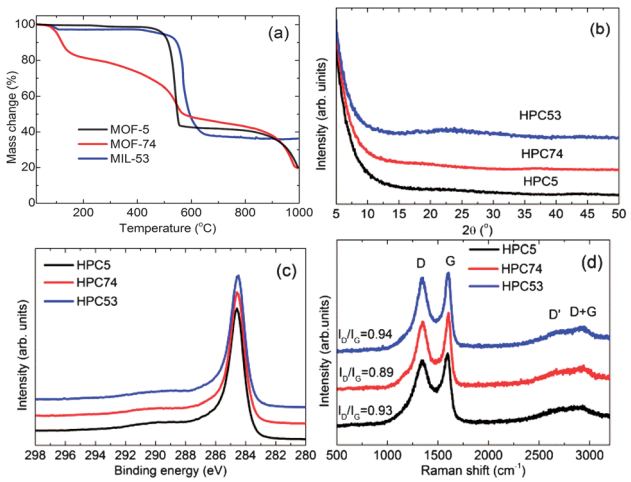


Fig. 2 (a) TGA plots for the MOF precursors; and (b) powder XRD patterns, (c) C 1s core level XPS spectra, (d) Raman spectra for the HPCs (HPC5b1-1000, HPC74-1000, HPC53-1000).

indicating the disorderly oriented tiny graphenic type fragments in the structures. It is worth noting that the reduction of ZnO by carbon during carbonization yields a pure porous carbon, whereas MIL-53 (Al) derived carbon always contains metal particles (Fig. S3 and S4†). The X-ray photoemission spectroscopy (XPS) C 1s core level spectra shown in Fig. 2c indicate that the HPCs are graphenic in nature with a majority of sp^2 carbons.[20] Clearly no Zn is detected in Zn 2p core level XPS spectra in samples obtained at ≥900 °C carbonization or after acid treatment (Fig. S4c†). The broad and highly intense disorder-induced D-band and featureless second-order bands between 2600 cm^{-1} and 2900 cm^{-1} in the Raman spectra (Fig. 2d and S4d†) represent the highly disordered/defective or porous nature of the HPCs.[28,45] The SEM images and digital photographs presented in Fig. 1, 3, S1 and S5† reveal that the HPCs retain mostly similar crystallite shapes and surface structures to their MOF precursors. It is also noted that the parent MOF crystallites shrink after carbonization as evident from the microcracks and the sponge-like surface morphology.[34] Furthermore, the transmission electron microscopy (TEM) images shown in Fig. 4 clearly reveals the highly defective, randomly oriented graphenic type porous carbon.

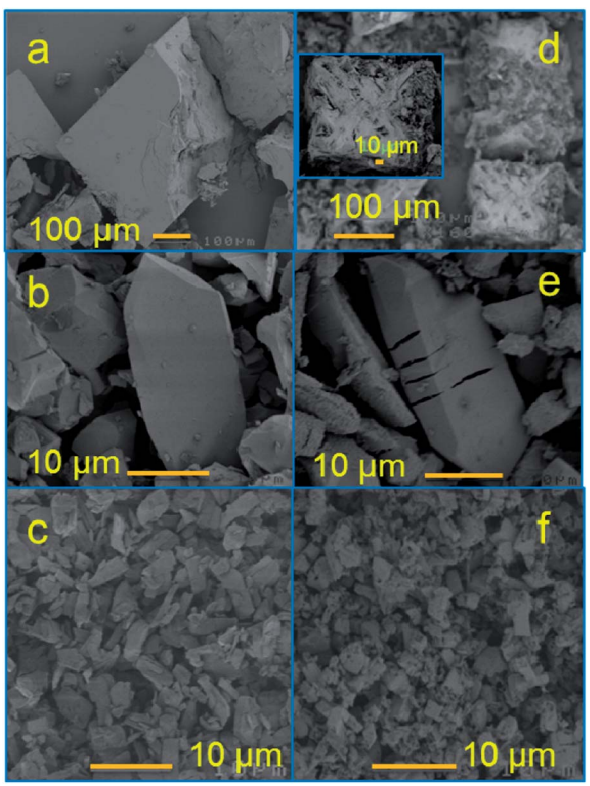


Fig. 3 SEM images of (a) MOF-5b1, (b) MOF-74, (c) MIL-53 and their HPCs; (d) HPC5b1-1000, (e) HPC74-1000, (f) HPC53-1000.

The BET surface area and pore volume values for the HPCs and MOFs were obtained from the nitrogen adsorption–desorption isotherms (Fig. 5a, S2 and S6–S8†) and are listed in Table 1 and S1.† The qualitative behaviour of the isotherms (Langmuir or type- I for micropores and S-shaped or type-II for micro–meso–macropores) represents hierarchical pores with dominant microporosity in HPC74 and a large fraction of mesopores in HPC5s. Figs. S6–S8† represent the pore size distribution and cumulative pore volume plots, obtained by the QSDFT and BJH methods. It is worth noting that the HPCs show a considerable enhancement in their total pore volumes compared with their MOF precursors. The total pore volumes between 5.23 $cm^3\ g^{-1}$ and 5.53 $cm^3\ g^{-1}$ in HPC5b2s are some of the highest values reported so far for porous carbons.[23,34,38,46–53] It is important to note that the HPC5b2s also show simultaneously high surface areas between 2495 $m^2\ g^{-1}$ and 2734 $m^2\ g^{-1}$. Such combination of properties (simultaneously large total pore volume and high surface area) is unique for MOF derived porous carbons and has not been observed yet in any other type of porous carbons. A few of the carbons show pore volumes over 5 $cm^3\ g^{-1}$ but with surface areas limited to well below 2000 $m^2\ g^{-1}$ or *vice versa*.[23,34,38,46–53] We also note that the measured total pore volumes of over 5 $cm^3\ g^{-1}$ in HPC5b2s are the highest among MOF derived porous carbons reported so far (Table S2† and references therein).[34,47,48,53] The highest reported values were 4.3 $cm^3\ g^{-1}$ and ~4 $cm^3\ g^{-1}$ in Al-MOF and MOF-5 derived carbons, respectively.[47,48,53] However, the surface area and total pore volume of HPC5b1s are well within the range of literature values (Table S2†).

The hierarchical pore formation in the HPCs is understandable from their pre-defined well-ordered pore structures of MOF precursors and combined TGA-MS measurements. During the carbonization process (≥900 °C), considerable C and O mass-loss occurs from the framework ligands through the reduction of ZnO that leads to the evolution of Zn, CO and CO_2, leaving a more defective or hollow carbon network. This process is a simple self-activation for the carbon, similar to the well-known physical or chemical activation of the carbonaceous precursors with either CO_2 or KOH to achieve porous activated carbons.[11,19,33] The simultaneously enhanced surface area and pore volume in HPC5b2 are attributed to the large, millimetre-sized crystallites of the MOF precursor, that leaves more void space due to the extended pores within the crystals after carbonization, compared with the smaller crystals in HPC5b1. The influence of the crystallite size of the MOF-5 precursor on the porosity parameters is clearly observed in Fig. S6† and Table 1. Similarly, the framework structure, metal-centres and initial porosity properties of the MOF precursor plays a crucial role in developing desirable porosity in HPCs. The low-density, spherical pore structure and cubic crystallites of the MOF-5 precursor always yield high porosity in the HPCs. However, the carbons derived from Zn-MOF-74 of densely packed, one

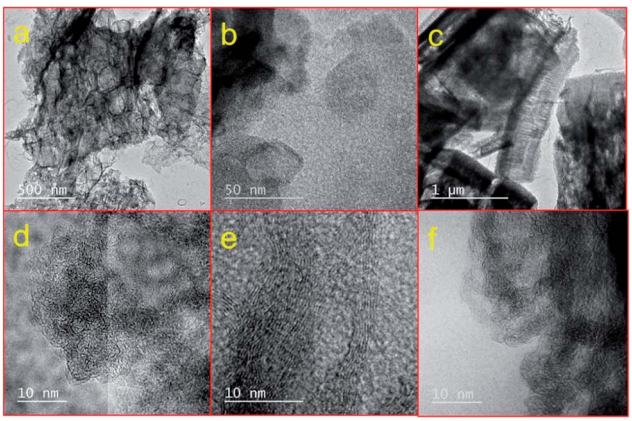


Fig. 4 TEM images of HPC5b2-1000 (a and d), HPC74-1000 (b and e) and HPC53-1000 (c and f). (a–c) and (d–f) represent low and high magnification respectively.

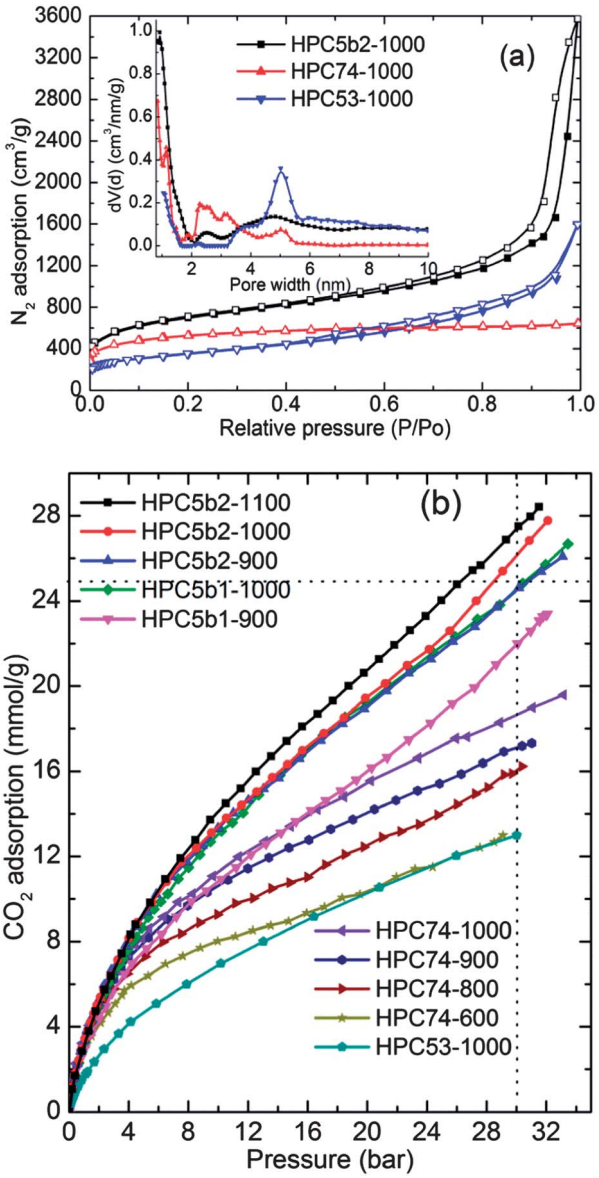


Fig. 5 (a) N_2 adsorption–desorption isotherms of HPCs measured at −196 °C. Inset shows the QSDFT pore size distribution plots. (b) CO_2 adsorption isotherms of HPCs measured at 27 °C. The numbers between 600 and 1100 represent the carbonization temperature in °C. HPC74-600, HPC74-800, and HPC53-1000 are after acid treatment.

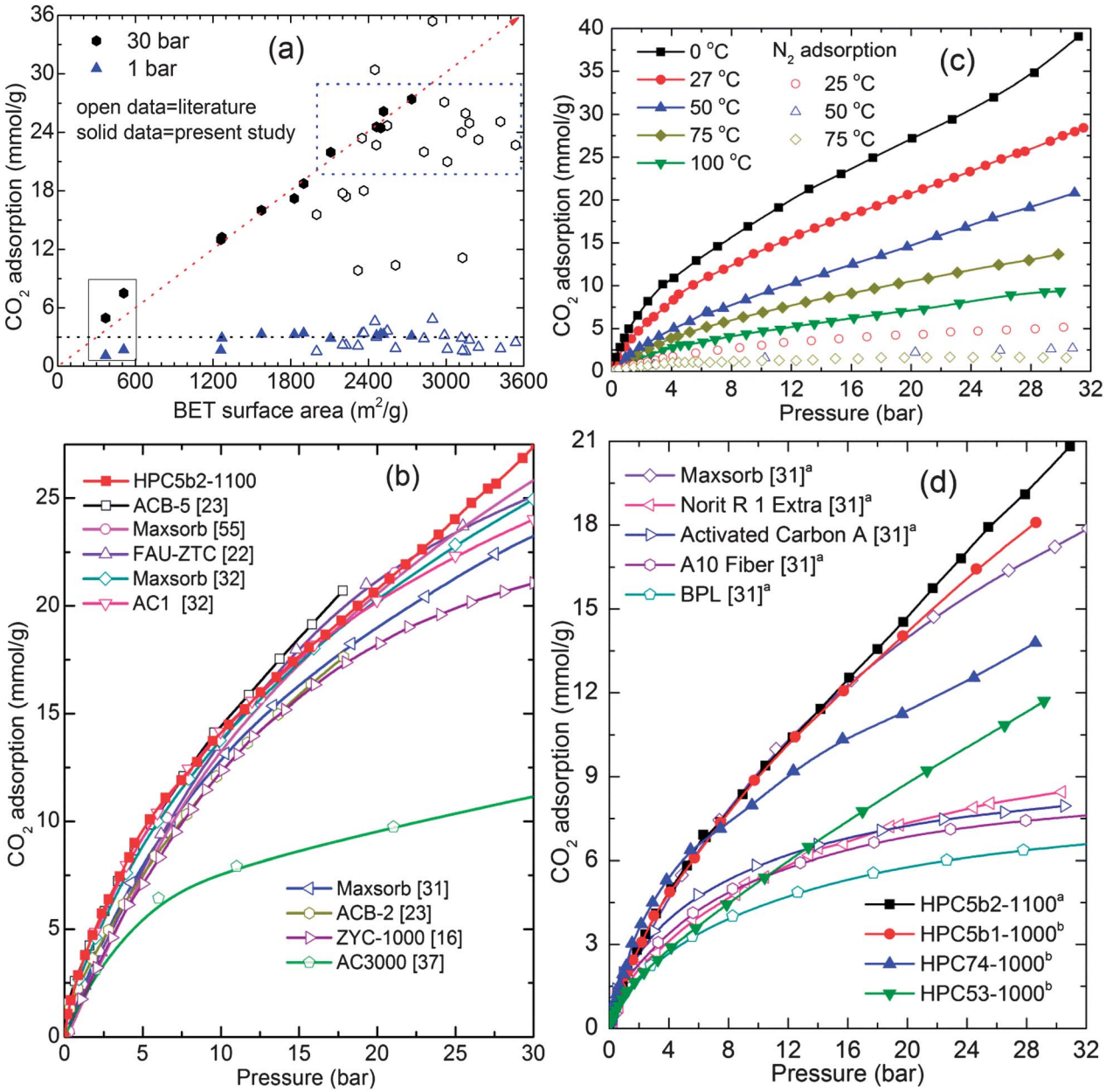


Fig. 6 (a and b) represent the comparative CO_2 adsorption at (25–27) °C: (a) CO_2 adsorption against BET surface area of HPCs and various porous carbons with surface area ≥2000 $m^2\ g^{-1}$. (b) Isotherms of various porous carbons with surface area ≥3000 $m^2\ g^{-1}$ and HPC5b2-1100 with 2734 $m^2\ g^{-1}$. The diagonal dotted line represents a linear relationship between the surface area and CO_2 uptake for HPCs. The horizontal dotted line at 3 mmol g^{-1} is the CO_2 uptake at 1 bar. The dotted box exemplifies the CO_2 performance of HPCs to that of carbons with ≥3000 $m^2\ g^{-1}$ in the literature. The data under the solid line box is from as-synthesized HPC74-600 and HPC74-800 samples. (c) CO_2 and N_2 adsorption in MOF5b2-1100 sample measured in the temperature range of (0–100) °C, and (d) compares the CO_2 adsorption performance of our HPC5b2-1100 sample, represented by superscript *a*, with commercial carbons at 50 °C. The CO_2 adsorption performance of HPC5b1-1000, HPC74-1000 and HPC53-1000 are at 47 °C, represented by superscript *b*. *Note*: in ref. 37, the isotherm on the AC3000 sample was obtained without outgassing, therefore the adsorption capacity is underestimated.

dimensional pore channels and lower surface area contain a high proportion of micro and mesoporosity. Thus designing and developing of desirable HPCs is not straight forward. The best example in this case is carbons obtained from the MIL-53(Al and Fe) (Fig. S8†). The MIL-53(Fe) derived carbon exhibits very poor porosities with a surface area of only 138 $m^2\ g^{-1}$ and a pore volume of 0.22 $cm^3\ g^{-1}$. Clearly the MOF framework structure acts as a self-template, thus the microporosity from the MOFs is mostly retained in the derived HPCs and the meso- and macro-pores are generated by reduction of ZnO with carbon and by evolution of Zn, CO and CO_2. In the case of MIL-53(Fe), the high solubility of carbon in the reduced Fe particles eventually results in growth of graphitic fibre-like structures with poor porosity (see Fig. S8d† for XRD and SEM images).

The CO_2 adsorption–desorption isotherms were measured at several different temperatures between (−53–47) °C and pressure, up to ~30 bar on HPCs (Fig. 5b and S9–S10†) and on their MOF precursors (Fig. S11†). The isotherm data obtained on the MOFs is in good agreement with the results published in the literature.[42–44] HPCs with their hierarchical pores (micro, meso and macro) and large pore volumes clearly show enhanced CO_2 uptake at high pressures, up to 30 bar without any indication of saturation. The HPCs also show considerable enhancement of CO_2 uptake in the low-pressure region, compared to the coordinatively saturated MOF-5 and MIL-53 precursors due to their small slit-pores.[17,38–41,54] The uptake values at low-pressure in MOFs and HPCs correlate well with the initial heat of adsorption (Fig. S11† and ref. 43). The isosteric heat of adsorption in HPCs shows moderate to strong interaction energies. The large mesopore fraction in HPC5 and HPC53 shows a low heat of adsorption.

As shown in Fig. 5b and 6, the measured high-pressure CO_2 uptake capacities of over 25 mmol g^{-1} and a maximum of 27.4 mmol g^{-1} at 30 bar and 27 °C in our HPCs is one of the highest uptake values reported for highly porous carbons.[14,16,18,21–23,31,32,37,55] Furthermore, the room temperature CO_2 uptake in the HPCs at 30 bar show a more or less linear relationship with their BET surface area: a CO_2 uptake of 10 mmol g^{-1} for every 1000 m^2 g^{-1} of surface area. A linear relation between CO_2 uptake at moderate pressure (6 bar) and the micropore volume is also observed (Fig. S12†). In addition, we also show a comparison of 27 °C CO_2 uptake at 1 bar and 30 bar in the present HPCs and other highly porous carbons with surface areas of ≥2000 m^2 g^{-1} (Fig. 6a) and CO_2 uptake isotherms in a wide pressure range for porous carbons with surface areas of ≥3000 m^2 g^{-1} in Fig. 6b.[14,16,18,21–23,27,31,32,37,55] Clearly, the HPCs show superior CO_2 uptakes over other porous carbons. Moreover, the CO_2 adsorption in the HPCs with surface areas between 2000 m^2 g^{-1} and 2700 m^2 g^{-1} are better than or comparable to the porous carbons with surface areas higher than 3000 m^2 g^{-1}. It is also worth noting that no correlation is found between BET surface area and CO_2 uptake at 30 bar for various types of other carbons. This can be primarily attributed to the large differences in their pore volumes as listed in Table S3.† The best example for this case is gas uptake in the well-defined pore structure MOFs (see Fig. S11†). The limited pore volume and pore-filling effect leads no further gas uptake at high pressure. In the HPCs, the hierarchical mesoporosity could play an additional role in obtaining higher CO_2 adsorption in the high-pressure region by multilayer adsorption within the pores (see ESI†). Furthermore, it is anticipated that the carbons with a higher micropore volume perform better for CO_2 adsorption at 1 bar, because of a large caging effect, particularly those smaller than 1 nm.[17,39,56] For example, highly microporous carbon spheres show good CO_2 capture performance, over 4 mmol g^{-1} at 1 bar and 25 °C.[56] However, HPCs dominated by hierarchical mesoporosity show CO_2 adsorption of ~3.0 mmol g^{-1} at ≤1 bar and 27 °C, but their much superior performance at relatively high CO_2 pressures makes them a preferred choice for pre-combustion PSA applications.[43,57] We listed the adsorptive capacities for both the MOF precursors and the HPCs, along with other high surface area carbons, in Table 1 and S3† for possible PSA/VSA conditions in the case of pure CO_2 and flue gas (20% CO_2) at 27 °C.[43,56] In Table S3,† we also list the surface areas, total pore volumes and high-pressure PSA adsorption values for similar materials reported in the literature. Here it is noted that the total pore volumes between 3.26 cm^3 g^{-1} and 5.53 cm^3 g^{-1} in our HPCs are the highest among the listed carbons.

As pre- and post-combustion CO_2 capture temperature conditions fall between 40 °C and 75 °C,[6] we show high-pressure pure component CO_2 and N_2 adsorption performance of the HPC5b2-1100 above room temperature, (0–100) °C in Fig. 6c. Fig. 6d and S13† also compare the 50 °C and 75 °C isotherms of the HPCs with commercial carbons; such as maxsorb with a BET surface area of 3250 m^2 g^{-1}, Norit R1 (1450 m^2 g^{-1}) and BPL (1150 m^2 g^{-1}), *etc.*[27,31] Clearly the present HPCs still absorb considerable amounts of CO_2 and outperform the commercial activated carbons at elevated pressures. We also note that the CO_2 adsorption of 5.7 mmol g^{-1} in the HPC5b2-1100 sample at 75 °C and 7.4 bar is much higher than the recently reported highest CO_2 adsorption of 4.3 mmol g^{-1} under identical experimental conditions in polyethyleneimine impregnated sandwich-like graphene–silica sheets.[58]

Conclusions

A new group of highly hierarchical porous carbons with high surface areas up to 2734 m^2 g^{-1} and high total pore volumes up to 5.53 cm^3 g^{-1} have been synthesised by tailored carbonization of various MOFs. The HPCs are highly sp^2-bonded graphenic in nature with a high portion of defective carbon structures. In most cases, the CO_2 uptakes in HPCs are higher than in their MOF counterparts. The high pressure CO_2 uptake of over 27 mmol g^{-1} at 30 bar and 27 °C in HPCs is one of the largest reported in the literature for porous carbons. There seems to be a direct relationship between the CO_2 adsorption capacity and the surface area: a CO_2 uptake of 10 mmol g^{-1} for every 1000 m^2 g^{-1} increase of surface area under 30 bar and 27 °C. Due to their thermal/chemical stability and enhanced CO_2 capture performance the HPCs are preferred over their counterpart MOFs for the PSA/VSA applications. Finally we note that the selection of a MOF structure and the process conditions play a critical role in obtaining the desirable porosities in the carbons. The present results also suggest that MOF derived hierarchical porous carbons may be favourably considered for other energy related applications, such as lithium-ion adsorption–desorption in battery and supercapacitor electrodes, as well as in methane storage.

Acknowledgements

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Notes and references

1 V. Scott, S. Gilfillan, N. Markusson, H. Chalmers and R. S. Haszeldine, *Nat. Clim. Change*, 2013, **3**, 105.

2 K. S. Lackner, S. Brennan, J. M. Matter, A.-H. A. Park, A. Wright and B. van der Zwaan, *Proc. Natl. Acad. Sci. U. S. A.*, 2012, **109**, 13156.
3 T. C. Drage, C. E. Snape, L. A. Stevens, J. Wood, J. Wang, A. I. Cooper, R. Dawson, X. Guo, C. Satterley and R. Irons, *J. Mater. Chem.*, 2012, **22**, 2815.
4 P. Markewitz, W. Kuckshinrichs, W. Leitner, J. Linssen, P. Zapp, R. Bongartz, A. Schreiber and T. E. Muller, *Energy Environ. Sci.*, 2012, **5**, 7281.
5 S. Choi, J. H. Drese and C. W. Jones, *ChemSusChem*, 2009, **2**, 796.
6 D. M. D'Alessandro, B. Smit and J. R. Long, *Angew. Chem., Int. Ed.*, 2010, **49**, 6058.
7 A. Samanta, A. Zhao, G. K. H. Shimizu, P. Sarkar and R. Gupta, *Ind. Eng. Chem. Res.*, 2012, **51**, 1438.
8 S. Wang, S. Yan, X. Ma and J. Gong, *Energy Environ. Sci.*, 2011, **4**, 3805.
9 X. Zhang, X. Zhang, H. Dong, Z. Zhao, S. Zhang and Y. Huang, *Energy Environ. Sci.*, 2012, **5**, 6668.
10 Y. Liu, Z. U. Wang and H.-C. Zhou, *Greenhouse Gases: Sci. Technol.*, 2012, **2**, 239.
11 G. Srinivas, J. Burress and T. Yildirim, *Energy Environ. Sci.*, 2012, **5**, 6453.
12 K. Sumida, D. L. Rogow, J. A. Mason, T. M. McDonald, E. D. Bloch, Z. R. Herm, T.-H. Bae and J. R. Long, *Chem. Rev.*, 2012, **112**, 724.
13 Z. Zhang, Y. Zhao, Q. Gong, Z. Li and J. Li, *Chem. Commun.*, 2013, **49**, 653.
14 J. P. Marco-Lozar, M. Kunowsky, F. Suarez-Garcia, J. D. Carruthers and A. Linares-Solano, *Energy Environ. Sci.*, 2012, **5**, 9833.
15 M. Sevilla and A. B. Fuertes, *Energy Environ. Sci.*, 2011, **4**, 1765.
16 H.-K. Youn, J. Kim, G. Chandrasekar, H. Jin and W.-S. Ahn, *Mater. Lett.*, 2011, **65**, 1772.
17 V. Presser, J. McDonough, S.-H. Yeon and Y. Gogotsi, *Energy Environ. Sci.*, 2011, **4**, 3059.
18 Z. Zhang, M. Xu, H. Wang and Z. Li, *Chem. Eng. J.*, 2010, **160**, 571.
19 M. Nandi, K. Okada, A. Dutta, A. Bhaumik, J. Maruyama, D. Derks and H. Uyama, *Chem. Commun.*, 2012, **48**, 10283.
20 J. W. Burress, S. Gadipelli, J. Ford, J. M. Simmons, W. Zhou and T. Yildirim, *Angew. Chem., Int. Ed.*, 2010, **49**, 8902.
21 J. Silvestre-Albero, A. Wahby, A. Sepulveda-Escribano, M. Martinez-Escandell, K. Kaneko and F. Rodriguez-Reinoso, *Chem. Commun.*, 2011, **47**, 6840.
22 S. Builes, T. Roussel, C. M. Ghimbeu, J. Parmentier, R. Gadiou, C. Vix-Guterl and L. F. Vega, *Phys. Chem. Chem. Phys.*, 2011, **13**, 16063.
23 X. Shao, Z. Feng, R. Xue, C. Ma, W. Wang, X. Peng and D. Cao, *AIChE J.*, 2011, **57**, 3042.
24 J.-R. Li, Y. Ma, M. C. McCarthy, J. Sculley, J. Yu, H.-K. Jeong, P. B. Balbuena and H.-C. Zhou, *Coord. Chem. Rev.*, 2011, **255**, 1791.
25 H. Furukawa, N. Ko, Y. B. Go, N. Aratani, S. B. Choi, E. Choi, A. O. Yazaydin, R. Q. Snurr, M. O'Keeffe, J. Kim and O. M. Yaghi, *Science*, 2010, **329**, 424.
26 Y. Peng, G. Srinivas, C. E. Wilmer, I. Eryazici, R. Q. Snurr, J. T. Hupp, T. Yildirim and O. K. Farha, *Chem. Commun.*, 2013, **49**, 2992.
27 C. A. Grande, R. Blom, A. Moller and J. Mollmer, *Chem. Eng. Sci.*, 2013, **89**, 10.
28 G. Srinivas, Y. Zhu, R. Piner, N. Skipper, M. Ellerby and R. Ruoff, *Carbon*, 2010, **48**, 630.
29 H.-L. Jiang, B. Liu, Y.-Q. Lan, K. Kuratani, T. Akita, H. Shioyama, F. Zong and Q. Xu, *J. Am. Chem. Soc.*, 2011, **133**, 11854.
30 V. Chandra, S. U. Yu, S. H. Kim, Y. S. Yoon, D. Y. Kim, A. H. Kwon, M. Meyyappan and K. S. Kim, *Chem. Commun.*, 2012, **48**, 735.
31 S. Himeno, T. Komatsu and S. Fujita, *J. Chem. Eng. Data*, 2005, **50**, 369.
32 J. P. Marco-Lozar, J. Juan-Juan, F. Suárez-García, D. Cazorla-Amorós and A. Linares-Solano, *Int. J. Hydrogen Energy*, 2012, **37**, 2370.
33 H. Marsh and F. Rodríguez-Reinoso, *Activated carbon*, Elsevier, London, 2006.
34 W. Chaikittisilp, K. Ariga and Y. Yamauchi, *J. Mater. Chem. A*, 2013, **1**, 14.
35 M. Sevilla and A. B. Fuertes, *J. Colloid Interface Sci.*, 2012, **366**, 147.
36 W. Shen, S. Zhang, Y. He, J. Li and W. Fan, *J. Mater. Chem.*, 2011, **21**, 14036.
37 T. C. Drage, J. M. Blackman, C. Pevida and C. E. Snape, *Energy Fuels*, 2009, **23**, 2790.
38 E. Masika and R. Mokaya, *J. Phys. Chem. C*, 2012, **116**, 25734.
39 Z. Zhang, J. Zhou, W. Xing, Q. Xue, Z. Yan, S. Zhuo and S. Z. Qiao, *Phys. Chem. Chem. Phys.*, 2013, **15**, 2523.
40 G. Srinivas, J. W. Burress, J. Ford and T. Yildirim, *J. Mater. Chem.*, 2011, **21**, 11323.
41 G. Yushin, R. Dash, J. Jagiello, J. E. Fischer and Y. Gogotsi, *Adv. Funct. Mater.*, 2006, **16**, 2288.
42 W. Zhou, H. Wu, M. R. Hartman and T. Yildirim, *J. Phys. Chem. C*, 2007, **111**, 16131.
43 J. M. Simmons, H. Wu, W. Zhou and T. Yildirim, *Energy Environ. Sci.*, 2011, **4**, 2177.
44 S. Bourrelly, P. L. Llewellyn, C. Serre, F. Millange, T. Loiseau and G. Ferey, *J. Am. Chem. Soc.*, 2005, **127**, 13519.
45 G. Srinivas, C. A. Howard, S. M. Bennington, N. T. Skipper and M. Ellerby, *J. Mater. Chem.*, 2009, **19**, 5239.
46 Z. Wen, X. Wang, S. Mao, Z. Bo, H. Kim, S. Cui, G. Lu, X. Feng and J. Chen, *Adv. Mater.*, 2012, **24**, 5610.
47 S. J. Yang, T. Kim, J. H. Im, Y. S. Kim, K. Lee, H. Jung and C. R. Park, *Chem. Mater.*, 2012, **24**, 464.
48 M. Hu, J. Reboul, S. Furukawa, N. L. Torad, Q. Ji, P. Srinivasu, K. Ariga, S. Kitagawa and Y. Yamauchi, *J. Am. Chem. Soc.*, 2012, **134**, 2864.
49 D. Li, F. Han, S. Wang, F. Cheng, Q. Sun and W.-C. Li, *ACS Appl. Mater. Interfaces*, 2013, **5**, 2208.
50 M. A. Worsley, T. Y. Olson, J. R. I. Lee, T. M. Willey, M. H. Nielsen, S. K. Roberts, P. J. Pauzauskie, J. Biener, J. H. Satcher Jr and T. F. Baumann, *J. Phys. Chem. Lett.*, 2011, **2**, 921.

51 R. Wang, P. Wang, X. Yan, J. Lang, C. Peng and Q. Xue, *ACS Appl. Mater. Interfaces*, 2012, **4**, 5800.
52 C. X. Guo, Y. Wang and C. M. Li, *ACS Sustainable Chem. Eng.*, 2013, **1**, 14.
53 K. Xi, S. Cao, X. Peng, C. Ducati, R. V. Kumar and A. K. Cheetham, *Chem. Commun.*, 2013, **49**, 2192.
54 X. Hu, M. Radosz, K. A. Cychosz and M. Thommes, *Environ. Sci. Technol.*, 2011, **45**, 7068.
55 B. B. Saha, S. Jribi, S. Koyama and I. I. El-Sharkawy, *J. Chem. Eng. Data*, 2011, **56**, 1974.
56 N. P. Wickramaratne and M. Jaroniec, *ACS Appl. Mater. Interfaces*, 2013, **5**, 1849.
57 M. T. Ho, G. W. Allinson and D. E. Wiley, *Ind. Eng. Chem. Res.*, 2008, **47**, 4883.
58 S. Yang, L. Zhan, X. Xu, Y. Wang, L. Ling and X. Feng, *Adv. Mater.*, 2013, **25**, 2130.