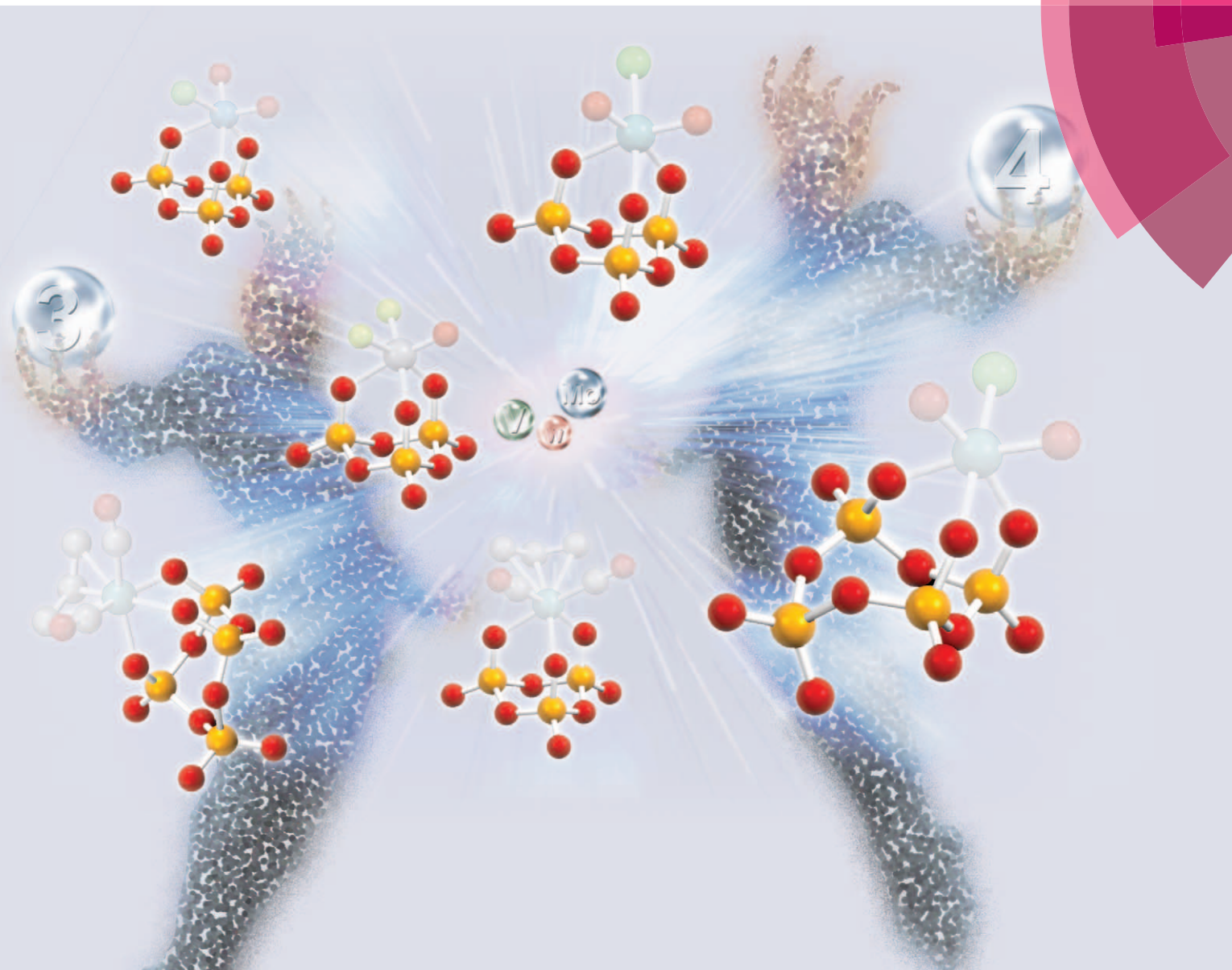


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Facile synthesis of mononuclear early transition-metal complexes of κ^3 cyclo-tetrametaphosphate ($[P_4O_{12}]^{4-}$) and cyclo-trimetaphosphate ($[P_3O_9]^{3-}$)

Facile synthesis of mononuclear early transition-metal complexes of κ^3 *cyclo*-tetrametaphosphate ($[\text{P}_4\text{O}_{12}]^{4-}$) and *cyclo*-trimetaphosphate ($[\text{P}_3\text{O}_9]^{3-}$)[†]

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We herein report the preparation of several mononuclear-metaphosphate complexes using simple techniques and mild conditions with yields ranging from 56% to 78%. Treatment of *cyclo*-tetrametaphosphate ($[\text{TBA}]_4[\text{P}_4\text{O}_{12}] \cdot 5\text{H}_2\text{O}$, TBA = tetra-*n*-butylammonium) with various metal sources including $(\text{CH}_3\text{CN})_3\text{Mo}(\text{CO})_3$, $(\text{CH}_3\text{CN})_2\text{Mo}(\text{CO})_2(\eta^3\text{-C}_3\text{H}_5)\text{Cl}$, $\text{MoO}_2\text{Cl}_2(\text{OSMe}_2)_2$, and VOF_3 , leads to the clean and rapid formation of $[\text{TBA}]_4[(\text{P}_4\text{O}_{12})\text{Mo}(\text{CO})_3] \cdot 2\text{H}_2\text{O}$, $[\text{TBA}]_3[(\text{P}_4\text{O}_{12})\text{Mo}(\text{CO})_2(\eta^3\text{-C}_3\text{H}_5)]$, $[\text{TBA}]_3[(\text{P}_4\text{O}_{12})\text{MoO}_2\text{Cl}]$ and $[\text{TBA}]_3[(\text{P}_4\text{O}_{12})\text{VOF}_2] \cdot \text{Et}_2\text{O}$ salts in isolated yields of 69, 56, 68, and 56% respectively. NMR spectroscopy, NMR simulations and single crystal X-ray studies reveal that the $[\text{P}_4\text{O}_{12}]^{4-}$ anion behaves as a tridentate ligand wherein one of the metaphosphate groups is not directly bound to the metal. *cyclo*-Trimetaphosphate-metal complexes were prepared using a similar procedure *i.e.*, treatment of $[\text{PPN}]_3[\text{P}_3\text{O}_9] \cdot \text{H}_2\text{O}$ (PPN = bis-(triphenylphosphine)iminium) with the metal sources $(\text{CH}_3\text{CN})_2\text{Mo}(\text{CO})_2(\eta^3\text{-C}_3\text{H}_5)\text{Cl}$, $\text{MoO}_2\text{Cl}_2(\text{OSMe}_2)_2$, MoOCl_3 , VOF_3 , WOCl_4 , and $\text{WO}_2\text{Cl}_2(\text{CH}_3\text{CN})_2$ to produce the corresponding salts, $[\text{PPN}]_2[(\text{P}_3\text{O}_9)\text{Mo}(\text{CO})_2(\eta^3\text{-C}_3\text{H}_5)]$, $[\text{PPN}]_2[(\text{P}_3\text{O}_9)\text{MoO}_2\text{Cl}]$, $[\text{PPN}]_2[(\text{P}_3\text{O}_9)\text{MoOCl}_2]$, $[\text{PPN}]_2[(\text{P}_3\text{O}_9)\text{VOF}_2] \cdot 2\text{CH}_2\text{Cl}_2$, and $[\text{PPN}]_2[(\text{P}_3\text{O}_9)\text{WO}_2\text{Cl}]$ in isolated yields of 78, 56, 75, 59, and 77% respectively. NMR spectroscopy, NMR simulations and single-crystal X-ray studies indicate that the trianionic ligand $[\text{P}_3\text{O}_9]^{3-}$ in these complexes also has κ^3 connectivity.

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1. Introduction

Bimetalated *cyclo*-tetrametaphosphate ($[\text{P}_4\text{O}_{12}]^{4-}$) and monometalated *cyclo*-trimetaphosphate ($[\text{P}_3\text{O}_9]^{3-}$) moieties have a wide range of applications including use as pigments, catalysts, food additives, and fluorescent materials.^{1–6} The *cyclo*-tetraphosphate anion is of special interest due to its flexible eight-membered ring structure that can produce complexes with different conformations.^{1,7–9} Currently, two different strategies are employed for the preparation of bimetallic complexes of *cyclo*-tetrametaphosphate. One strategy is based on acidification of MCl_2 ($\text{M} = \text{Co}, \text{Fe}, \text{Mn}, \text{Ni}, \text{Cu}$) by $[\text{H}_2\text{PO}_4]^-$, followed by thermal treatment at high temperatures (≥ 600 °C), leading to binary oligomeric systems of the general formula $\text{M}_2\text{P}_4\text{O}_{12}$.^{2–4,10–19}

A second strategy developed by Kamimura *et al.* comprises stirring a solution containing a salt of the *cyclo*-tetrametaphosphate ligand with the desired noble transition metal (such as rhodium, iridium, ruthenium, and palladium) at room temperature overnight. The desired product can then be obtained by further extraction and crystallization from dichloromethane and diethyl ether.⁷ They have also described the synthesis of di- and trinuclear μ -oxo titanium(IV) *cyclo*-tetraphosphate complexes.¹ Even though the preparation of noble and non-noble metal complexes using this ligand has been reported, a strategy for the straightforward preparation of mononuclear complexes of *cyclo*-tetrametaphosphate using mild conditions and non-noble metals represents a vacancy in the literature. In addition to filling a vacancy, we recognize that metals such as molybdenum and vanadium are useful in oxidation catalysis, such that cyclophosphate complexes of these metals may turn out to be valuable as precatalysts.

The chemistry of complexes incorporating the trianionic *cyclo*-trimetaphosphate ligand $[\text{P}_3\text{O}_9]^{3-}$ has also been investigated, specifically in studies of catalytic applications such as vinylidene rearrangement of general internal alkynes *via* the 1,2-migration of alkyl, aryl, and acyl groups.^{5,6} The methodology used for the synthesis of metal-*cyclo*-trimetaphosphate compounds was first

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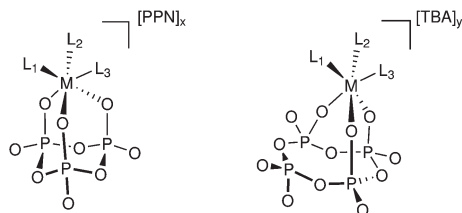


Fig. 1 General structure of *cyclo*-tetrametaphosphate (right) and *cyclo*-trimetaphosphate (left) salts. The letter M represents one of the transition metals used in this work ($M = V, W$ or Mo) and L_n represents the remainder of the metal's coordination sphere.

reported by Klemperer and co-workers in 1981, and since then only a handful of publications reporting the preparation of mononuclear *cyclo*-trimetaphosphate-transition-metal complexes has been published.^{5,6,20–29} Additionally, the coordination chemistry of early transition-metal- $[P_3O_9]$ complexes has not been explored extensively. Our group has recently reported the preparation of the first tridentate-*cyclo*-tetrametaphosphate complex.⁸ Montag *et al.* reported the preparation of $[Na][[(P_4O_{12})Co(TACN)]]$ (TACN = 1,4,7-triazacyclononane) in aqueous media. The structure determination of this salt revealed that the ligand is coordinated to the metal in a κ^3 fashion in which one of the phosphate groups is not directly connected to the metal center.

Herein we report the coordination chemistry of mononuclear complexes of *cyclo*-tetrametaphosphate and *cyclo*-trimetaphosphate ligands that incorporate non-noble metals (Fig. 1) along with the development of a simple synthetic procedure, using readily accessible precursors and common glovebox techniques.

2. Results and discussion

2.1. Synthesis and characterization

$[TBA]_4[P_4O_{12}] \cdot 5H_2O$ (**1**) was prepared from the known salt $[Na]_4[P_4O_{12}] \cdot 5H_2O$ ³⁰ by dissolving the latter in deionized water and passing the solution through a DOWEX® column which was previously conditioned with $[TBA][OH]$. The eluent was evaporated to afford a white solid which was dissolved in acetonitrile and then precipitated with diethyl ether to yield the desired compound. In general, salts **2**, **3**, **4** and **5** were prepared by dissolving $[TBA]_4[P_4O_{12}] \cdot 5H_2O$ in dry acetonitrile and adding an appropriate transition-metal reagent (Fig. 2). Typically, the reaction mixture was stirred for one hour, leading to complete and clean conversion into the desired complex with isolated yields ranging from 56 to 69%. The observed less-than-quantitative yields are attributed mainly to losses incurred in the purification process (crystallization).

The salts **7**, **8**, **9**, **10** and **11** were obtained by adding the corresponding metal precursors to a stirring solution of $[PPN]_3[P_3O_9] \cdot H_2O$ in dry acetonitrile, affording complete conversion to the final products (Fig. 3). The compounds $[PPN]_2[(P_3O_9)Mo(CO)_2(\eta^3-C_3H_5)]$ (**7**) and $[PPN]_2[(P_3O_9)MoO_2Cl]$ (**8**) spontaneously precipitate *ca.* two minutes after the

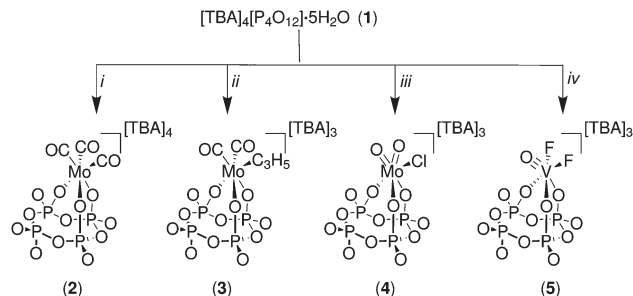


Fig. 2 Synthesis of molybdenum and vanadium *cyclo*-trimetaphosphate salts. All reactions were carried out in a glovebox using dry solvents. $i = (CH_3CN)_3-Mo(CO)_3$, $ii = (CH_3CN)_2Mo(CO)_2(\eta^3-C_3H_5)Cl$, $iii = MoO_2Cl_2(OSMe_2)_2$, $iv = VOF_3$.

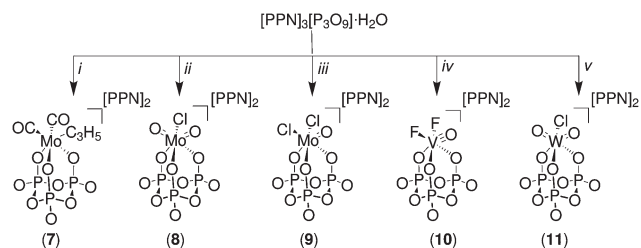


Fig. 3 Synthesis of molybdenum, tungsten, and vanadium *cyclo*-trimetaphosphate salts. All reactions were carried out in a glovebox using dry solvents. $i = (CH_3CN)_2Mo(CO)_2(\eta^3-C_3H_5)Cl$, $ii = MoO_2Cl_2(OSMe_2)_2$, $iii = MoOCl_3$, $iv = VOF_3$, $v = WOCl_4$.

addition of the corresponding precursor to the acetonitrile solution of the metaphosphate ligand. The isolated yields of these complexes are very similar to those obtained in the case of the *cyclo*-tetrametaphosphate complexes, varying from 56 to 78%.

The products were characterized by solid-state FT-IR spectroscopy; stretching frequencies for signature oscillators of *cyclo*-tetrametaphosphate complexes are summarized in Table S1.† The IR spectrum of complex $[TBA]_4[(P_4O_{12})Mo(CO)_3] \cdot 2H_2O$ (**2**) features three carbonyl bands at 1874, 1708, and 1702 cm^{-1} consistent with C_1 symmetry. The compound $[TBA]_3[(P_4O_{12})Mo(CO)_2(\eta^3-C_3H_5)]$ (**3**) features two stretching bands at 1898 and 1709 cm^{-1} for the carbonyl groups which are higher in energy compared to those observed in the case of salt **2**, as would be expected as a function of the higher oxidation state of the molybdenum center.^{31,32} The salt $[TBA]_3[(P_4O_{12})MoO_2Cl]$ (**4**) shows two stretches for the MoO_2 moiety (896 and 880 cm^{-1}), consistent with the presence of a pair of *cis* oxo ligands.³³ $[TBA]_3[(P_4O_{12})VOF_2] \cdot Et_2O$ (**5**) shows only one stretch (912 cm^{-1}) for the $V=O$ bond. The value observed for this stretch is in the expected range.^{34–37} ⁵¹V-NMR shows a broad peak at -594.5 ppm similar to the observations reported by Roesky *et al.* for the complex Ph_3NVOF_2 .³⁴

Complex $[PPN]_2[(P_3O_9)Mo(CO)_2(\eta^3-C_3H_5)]$ (**7**) has two characteristic $C=O$ stretches at 1918 and 1812 cm^{-1} which, as expected, appear at higher energy compared to those of the already reported salt $[PPN]_3[(P_3O_9)Mo(CO)_3]$ (**6**).³⁸ Complex $[PPN]_2[(P_3O_9)MoO_2Cl]$ evinces two IR bands at 923 and 898 cm^{-1} , characteristic of a molybdenum dioxo moiety, while



Table 1 Comparison of selected IR data between compounds reported in this work and those reported for other tripodal ligands (cm^{-1}). Counter ions and complex charges omitted for simplification

M	MP_4O_{12}	MP_3O_9	$\text{M}[\eta^5\text{-CpCo}\{\text{P}(\text{O})(\text{R})_2\}_3]$	$\text{MBC}_9\text{H}_{10}\text{N}_6$	$\text{MC}_3\text{H}_6\text{N}_3\text{Me}_3$
$\text{Mo}(\text{CO})_3$	1874	1883	1880 ³⁹	1890 ⁴⁰	1900 ⁴¹
	1708	1723	1738 ³⁹	1750 ⁴⁰	1785 ⁴¹
	1702		1710 ³⁹		1750 ⁴¹
$\text{Mo}(\text{CO})_2\text{C}_3\text{H}_5$	1898	1928	1921 ³⁹	1934 ⁴²	nr
	1709	1812	1822 ³⁹	1842 ⁴²	nr

complex $[\text{PPN}]_2[(\text{P}_3\text{O}_9)\text{MoOCl}_2]$ (**9**) has a single stretch for the $\text{Mo}=\text{O}$ bond at 932 cm^{-1} . The FT-IR spectrum of compound $[\text{PPN}]_2[(\text{P}_3\text{O}_9)\text{VOF}_2]\cdot 2\text{CH}_2\text{Cl}_2$ (**10**) manifests one signal at 930 cm^{-1} , characteristic of a $\text{V}=\text{O}$ stretch (Table S2†). To the best of our knowledge, there is only one previous report of a fully characterized and stabilized monomeric oxo fluoride vanadium compound, reported by Roesky *et al.* The complex $\text{Ph}_3\text{PNVOF}_2$ was prepared by treating the ligand ($\text{Ph}_3\text{PNSiMe}_3$) with the precursor (VOF_3) at low temperature ($-73\text{ }^\circ\text{C}$) and stirring for twelve hours.³⁴ The salt $[\text{PPN}]_2[(\text{P}_3\text{O}_9)\text{WO}_2\text{Cl}]$ (**11**) exhibits two absorption bands for the WO_2 moiety at 938 and 907 cm^{-1} .

Table 1 summarizes the carbonyl stretches of the compounds reported in this work (*i.e.* **2**, **3**, **6**, and **7**) and compounds that contain a tripodal ligand (Kläui, (pyrazolyl)borate and 1,3,5-triazacyclohexane ligands) bearing the same metal-carbonyl moiety. From the values reported, we may conclude that the *cyclo*-tetrametaphosphate ligand is a better donor than the *cyclo*-trimetaphosphate ligand and the other tridentate ligands listed in Table 1. This is reflected in decreased $\nu_{\text{C}=\text{O}}$ stretching frequencies corresponding to weaker CO bonds.^{43,44}

In addition to IR spectroscopy, ^{31}P -NMR spectroscopy is a helpful tool for determining the connectivity of the metaphosphate ligand with respect to the metal center. All the *cyclo*-tetrametaphosphate-based complexes reported herein exhibit three signals in their ^{31}P -NMR spectra in an intensity ratio of 1 : 2 : 1. This supports that the polyanionic ligand is binding in a tripodal fashion, leading to κ^3 connectivity (similar to a previous report from our group).⁸ On the other hand, $[\text{TBA}]_3[(\text{P}_4\text{O}_{12})\text{VOF}_2]\cdot\text{Et}_2\text{O}$ exhibits only two peaks in a ratio of 1 : 3, possibly due to a fluxional process or to a change to κ^2 coordination in solution (this could possibly be attributed to the lability of the oxygen trans to the oxo ligand), leading to broad peaks instead of the expected triplets. Variable temperature NMR experiments confirmed the behavior described above. At lower temperatures (Fig. 4) the spectrum for complex **4** shows a new pattern with a spin system AM_2X represented by three triplets in a ratio of 1 : 2 : 1 while at room temperature the spectrum consists of a singlet.

Fig. 5 shows the spectrum observed for complex $[\text{TBA}]_3[(\text{P}_4\text{O}_{12})\text{MoO}_2\text{Cl}]$ at $-60\text{ }^\circ\text{C}$ (a) along with the corresponding simulation (b) which was computed using the program gNMR.⁴⁵ The simulation reinforces the proposed four spin system AM_2X in line with the existence of two phosphate moieties which are similar to each other and strongly coupled. These coupled phosphate units have two different phosphate neighbors which do not show any coupling

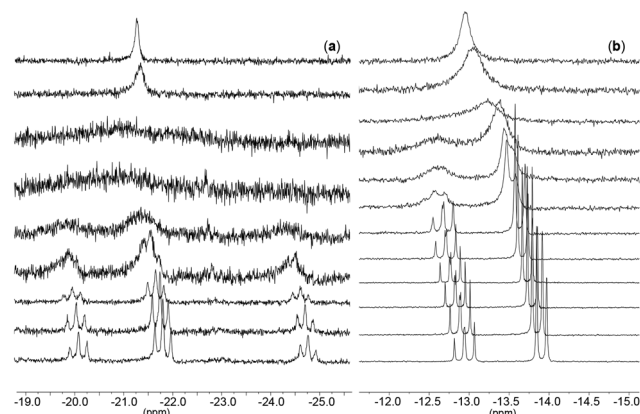


Fig. 4 VT- ^{31}P -NMR of (a): $[\text{TBA}]_3[(\text{P}_4\text{O}_{12})\text{MoO}_2\text{Cl}]$ and (b): $[\text{PPN}]_2[(\text{P}_3\text{O}_9)\text{Mo}(\text{CO})_2(\eta^3\text{-C}_3\text{H}_5)]$. The spectra were taken with an increase of $10\text{ }^\circ\text{C}$ from $+20\text{ }^\circ\text{C}$ (top) to $-60\text{ }^\circ\text{C}$ (bottom) in case of spectra a and from $+70\text{ }^\circ\text{C}$ (top) to $-40\text{ }^\circ\text{C}$ (bottom) in case of spectrum b.

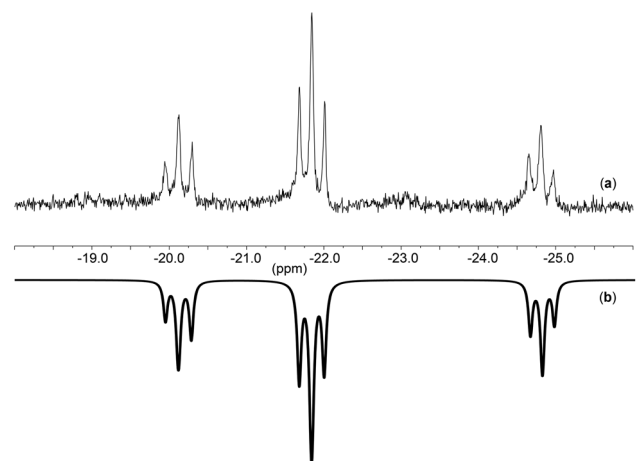


Fig. 5 ^{31}P -NMR of (a): $[\text{TBA}]_3[(\text{P}_4\text{O}_{12})\text{MoO}_2\text{Cl}]$ at $-60\text{ }^\circ\text{C}$ and (b): simulated spectrum using gNMR.⁴⁵

between them. This suggests, for example, a κ^3 connectivity of the phosphate to the metal in a fashion similar to that observed by Montag *et al.*⁸

$^{31}\text{P}\{^1\text{H}\}$ -NMR spectra of compounds **7**, **8** and **10** show three signals in a ratio of 4 : 1 : 2 corresponding to $[\text{PPN}]^+$ and the respective anion. The ^{31}P -NMR spectrum of complex **7** was recorded at different temperatures (Fig. 4) showing that at high temperatures it consists of a singlet and at low temperatures it



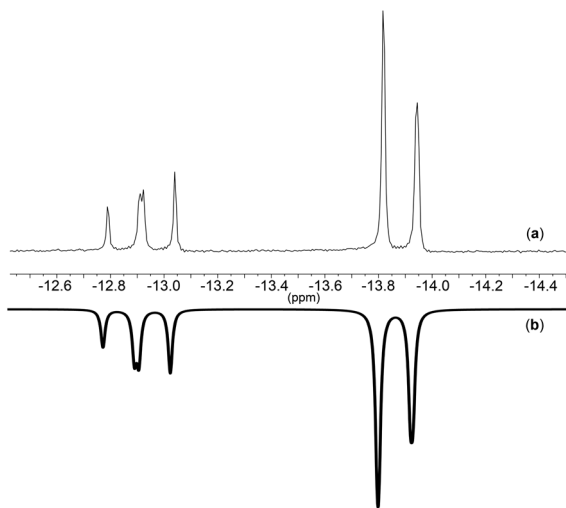


Fig. 6 ^{31}P -NMR of (a): $[\text{PPN}]_2[(\text{P}_3\text{O}_9)\text{Mo}(\text{CO})_2(\eta^3\text{-C}_3\text{H}_5)]$ at $-30\text{ }^\circ\text{C}$ and (b): spectrum simulated using gNMR.⁴⁵

evolves into a triplet and a doublet. This behavior could be related to fluctuation of the ligand around the metal. The ^1H -NMR spectrum of salt 7 shows the presence of four signals corresponding to the hydrogens of the allyl group and $[\text{PPN}]^+$, wherein the apical and *syn* protons appear at very similar chemical shifts. The $^{31}\text{P}\{^1\text{H}\}$ -NMR spectrum of $[(\text{P}_3\text{O}_9)\text{-MoO}_2\text{Cl}]^{2-}$ is characteristic of an A_2B spin system indicating a κ^3 connectivity of the metaphosphate ligand. The $^{31}\text{P}\{^1\text{H}\}$ -NMR spectrum of salts 9 and 11 features only two signals in a 4 : 3 ratio, corresponding to $[\text{PPN}]^+$ and the *cyclo*-trimetaphosphate-metal complex respectively.

The simulated spectrum for the complex $[\text{PPN}]_2[(\text{P}_3\text{O}_9)\text{-Mo}(\text{CO})_2(\eta^3\text{-C}_3\text{H}_5)]$ at $-30\text{ }^\circ\text{C}$ is shown in Fig. 6. The simulation demonstrates that the spin system of this molecule is as proposed A_2B giving rise to a doublet coupled to an apparent triplet. This spin system suggests that one out of the three metaphosphate moieties has a unique chemical environment.

2.2. Crystallographic studies

X-ray diffraction quality crystals of $[\text{TBA}]_3[(\text{P}_4\text{O}_{12})\text{Mo}(\text{CO})_2(\eta^3\text{-C}_3\text{H}_5)]$ (Fig. 7) were grown from a mixture of diethyl ether and dichloromethane. The symmetry of the complex is C_1 and the connectivity of the metaphosphate is κ^3 . This result is consistent with the conclusions drawn from ^{31}P -NMR and FT-IR spectroscopies studies. As the NMR simulations suggested, there are three different phosphorous environments and two of these are not coupled (the dangling metaphosphate and the farthest counterpart). The configuration of the allyl group is *endo*, and the bond length between the metal and the central carbon is, as expected, shorter than those between the metal and the terminal carbons.⁴⁶ The distances between molybdenum and the oxygen atoms are all similar and only differ by 0.035 \AA . The crystal structure of this salt also contains a dichloromethane molecule which is hydrogen bonded to the dangling metaphosphate, partially stabilizing the negative

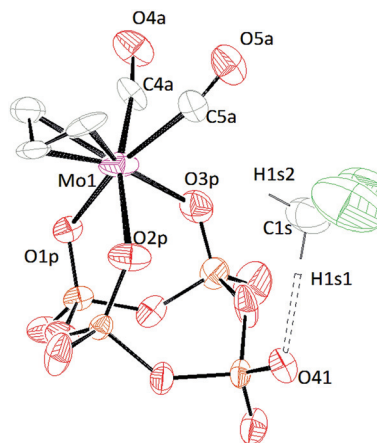


Fig. 7 Solid-state molecular structure of anion $[(\text{P}_4\text{O}_{12})\text{Mo}(\text{CO})_2(\eta^3\text{-C}_3\text{H}_5)]^{3-}$ rendered using ORTEP⁴⁸ with ellipsoids at the 50% probability level; the $[\text{TBA}]^+$ cations and minor disorder are omitted for clarity.

charge on the metaphosphate. The $\text{D}\cdots\text{A}$ distance is 3.065 \AA (Table S3[†]) suggesting that this is a weak hydrogen bond.⁴⁷ This structure suggests that it may be possible to pre-organize a substrate by interaction with the dangling phosphate residue.

Crystals of $[\text{TBA}]_3[(\text{P}_4\text{O}_{12})\text{MoO}_2\text{Cl}]$ were obtained from a mixture of dichloromethane and diethyl ether, and were subjected to study *via* X-ray crystallography. The conformation of the oxo atoms is *cis* and the connectivity of the metaphosphate is κ^3 , which is again in complete agreement with FT-IR and NMR spectroscopy studies (Fig. 8). The values observed for the interatomic distances of compound $[\text{TBA}]_3[(\text{P}_4\text{O}_{12})\text{MoO}_2\text{Cl}]$ are very similar to the values for $\text{MoO}_2\text{Cl}_2(\text{OSMe}_2)_2$.⁴⁹ As was the case for 3, this derivative (complex 4) possesses C_1 symmetry. The crystal structure of this compound also contains a dichloromethane molecule that is hydrogen bonded with one of the oxygen atoms in the metaphosphate ligand. The $\text{D}\cdots\text{A}$ distance is 3.253 \AA (Table S3[†]), indicative of a weak hydrogen bond.⁴⁷ Selected interatomic distances for salts 3 and 4 are summarized in Table 2.

The decreased $\text{Mo}\text{-OP}$ distance in 4 *versus* that of 3 correlates with the increased oxidation state in the former. Even

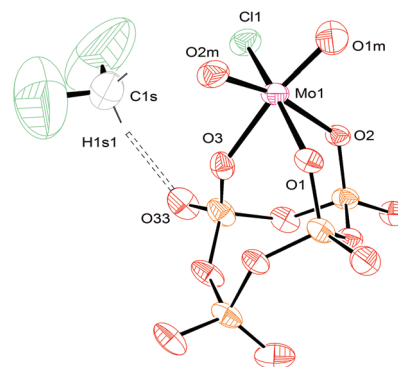


Fig. 8 Solid-state molecular structure of anion $[(\text{P}_4\text{O}_{12})\text{MoO}_2\text{Cl}]^{3-}$ rendered using ORTEP⁴⁸ with ellipsoids at the 50% probability level; the $[\text{TBA}]^+$ cations and minor disorder are omitted for clarity.



Table 2 Selected interatomic distances (Å) of salts 3 and 4

Bond	3	4
Mo–O ₁	2.181(4)	2.0764(4)
Mo–O ₂	2.216(4)	2.163(4)
Mo–O ₃	2.200(5)	2.172(4)
Mo–CO	1.979(11)–2.038(13)	
Mo–Cl		2.396(2)
Mo–O _m		1.691(4)–1.794(8)

3: [TBA]₃[(P₃O₁₂)Mo(CO)₂(η³-C₃H₅)], 4: [TBA]₃[(P₃O₁₂)MoO₂Cl].

though there was no crystal structure reported for the starting material (CH₃CN)₂Mo(CO)₂(η³-C₃H₅)Cl, the crystal structure of (CH₃CN)₂Mo(CO)₂(η³-2-methyl-allyl)Cl was reported by Baker *et al.* The data from the solid-state structure of this complex show that the values of the interatomic distances of Mo=O and Mo–Cl are very similar to those observed in our complexes.⁵⁰

Crystal structures were obtained for all the *cyclo*-trimetaphosphate salts reported herein. In general, all the complexes have C_s symmetry. Crystals of [PPN]₂[(P₃O₉)Mo(CO)₂(η³-C₃H₅)] were grown by vapor diffusion of diethyl ether into a dilute dichloromethane solution (Fig. 9). The solid-state structure of [(P₃O₉)Mo(CO)₂(η³-C₃H₅)]²⁻ exhibits κ³ connectivity (consistent with the ³¹P{¹H}-NMR spectrum) wherein all of the Mo–OP bonds differ by 0.085 Å. As suggested by the NMR simulations, one of the Mo–O distances is shorter than the other two (Mo–O₁) leading to a slightly different phosphorus environment. The Mo–CO interatomic distances are similar to those reported for [(P₃O₉)Mo(CO)₃]³⁻.³⁸ The configuration of the allyl group in the solid state is *endo*, and the bond length between the metal and the central carbon is shorter than the bond between the metal and the external carbons by 0.14 Å.

Crystals of [PPN]₂[(P₃O₉)MoO₂Cl] and [PPN]₂[(P₃O₉)WO₂Cl] were grown by vapor diffusion of diethyl ether into a saturated acetonitrile solution affording colorless crystals (Fig. 10 and 11

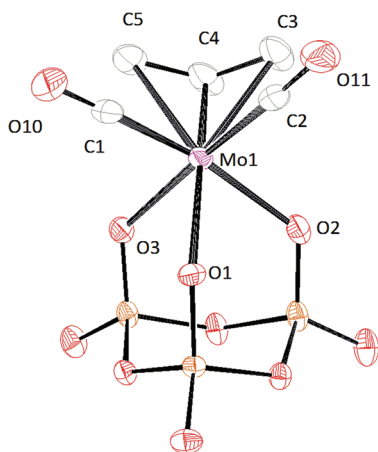


Fig. 9 Solid-state molecular structure of anion [(P₃O₉)Mo(CO)₂(η³-C₃H₅)]²⁻ rendered using ORTEP⁴⁸ with ellipsoids at the 50% probability level; the [PPN]⁺ cations are omitted for clarity.

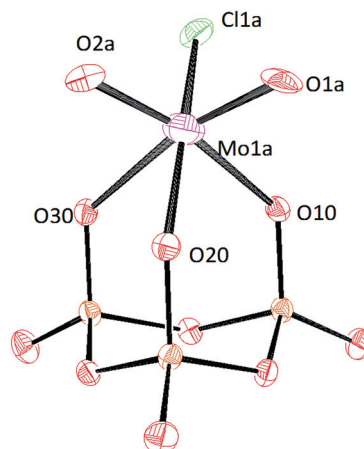


Fig. 10 Solid-state molecular structure of anion [(P₃O₉)MoO₂Cl]²⁻ rendered using ORTEP⁴⁸ with ellipsoids at the 50% probability level; the [PPN]⁺ cations and minor disorder are omitted for clarity.

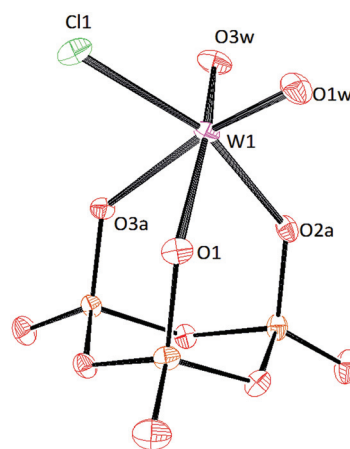


Fig. 11 Solid-state molecular structure of the anion [(P₃O₉)WO₂Cl]²⁻ rendered using ORTEP⁴⁸ with ellipsoids at the 50% probability level; the [PPN]⁺ cations and solvent are omitted for clarity.

respectively). The crystal structure of the compound [PPN]₂[(P₃O₉)WO₂Cl] shows that the configuration of the two oxo atoms is *cis* which is consistent with the result obtained by IR spectroscopy. Crystals of [PPN]₂[(P₃O₉)MoOCl₂] were grown from a mixture of dichloromethane–toluene (2 : 1) affording green crystals (Fig. 12).

Crystals of [PPN]₂[(P₃O₉)VOF₂]₂·2CH₂Cl₂ were grown from a mixture of dichloromethane and diethyl ether (Fig. 13). The V≡O distance is 1.6242(19) Å indicating a bond order of three.⁵¹ In the solid-state structure of salt 10, two solvent molecules are present. One dichloromethane solvent molecule engages in a weak hydrogen bond (D⋯A distance 3.317 Å) to one of the fluoride ions attached to the vanadium center (Table S3†).⁴⁷

Selected interatomic distances for all the crystal structures of *cyclo*-trimetaphosphate complexes are summarized in Table 3. In general, the higher the oxidation state of the metal,



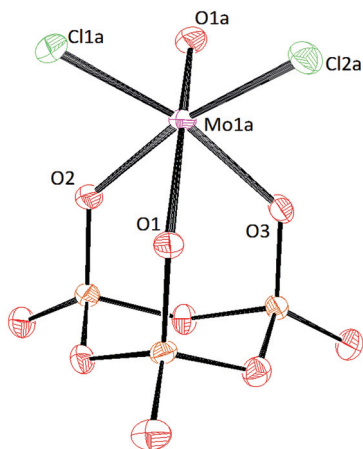


Fig. 12 Solid-state molecular structure of anion $[(P_3O_9)MoOCl_2]^{2-}$ rendered using ORTEP⁴⁸ with ellipsoids at the 50% probability level; the $[PPN]^+$ cations and minor disorder are omitted for clarity.

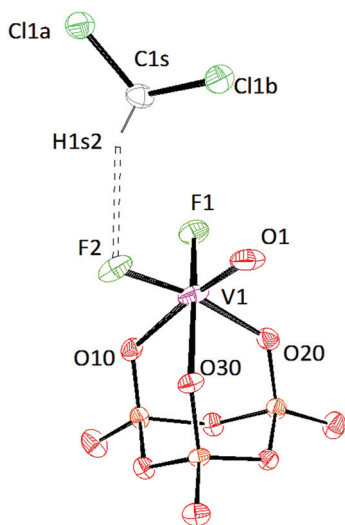


Fig. 13 Solid-state molecular structure of the anion $[(P_3O_9)VOF_2]^{2-}$ rendered using ORTEP⁴⁸ with ellipsoids at the 50% probability level; the $[PPN]^+$ cation and solvent are omitted for clarity.

the shorter the Mo–OP bond. The M–O_m bond (O_m = terminal oxo atom) becomes shorter as the oxidation state of the metal increases. The bond length of V=O is shorter than the other

M–O_m bonds, lending credence to its interpretation as a triple bond.⁵¹

3. Conclusions

In summary, the facile synthesis of a series of early transition-metal-cyclophosphate complexes (bearing labile ligands) is reported. The complexes (2–5 and 7–11) were isolated in good yield and were fully characterized by means of IR and NMR spectroscopy. The obtained crystal structures and NMR spectroscopy data for these complexes showed that both *cyclo*-triphosphate and *cyclo*-tetraphosphate behave as tridentate ligands connected in κ^3 mode: in the case of $[P_4O_{12}]^{4-}$, one of the metaphosphate units is uncoordinated.

4. Experimental

General

Unless stated otherwise, all the operations were performed in a Vacuum Atmospheres drybox under an atmosphere of purified nitrogen. Mo(CO)₆ was purchased from Strem and used without further purification. Tetrabutylammonium hydroxide 1 M and DOWEX® 50 WX4-200 ion-exchange resin were purchased from Sigma-Aldrich. $[Na]_4[P_4O_{12}] \cdot 5H_2O$,³⁰ $[PPN]_3[P_3O_9] \cdot H_2O$,⁵² $(CH_3CN)_3Mo(CO)_3$,⁵³ $(CH_3CN)_2Mo(CO)_2(\eta^3-C_3H_5)Cl$,⁵⁴ $MoO_2Cl_2(OSMe_2)_2$,⁴⁹ $MoOCl_3$,⁵⁵ $WO_2Cl_2(CH_3CN)_2$,⁵⁵ $WOCl_4$ ⁵⁵ and $[PPN]_3[(P_3O_9)Mo(CO)_3]$ (6)³⁸ were prepared according to published procedures. Aqueous solutions were prepared using reagent grade deionized water ($r \geq 18$ MΩ cm; Rica Chemical Company, USA). Diethyl ether, acetonitrile, dichloromethane and tetrahydrofuran were dried and deoxygenated by the method of Grubbs using a system built by SG Water USA, LLC and stored over 4 Å molecular sieves. 4 Å molecular sieves were dried under reduced pressure at a temperature above 200 °C over the course of one week. ¹H, ¹³C{¹H} and ³¹P{¹H}-NMR spectra were recorded on Varian Mercury-300 or Bruker AVANCE-400 spectrometers. NMR solvents were obtained from Cambridge Isotope Laboratories. ¹H and ¹³C chemical shifts are reported with respect to internal solvent resonances (DMSO-*d*₆, δ : 2.50 and 39.52 ppm respectively). ³¹P-NMR chemical shifts are reported with respect to an external reference (85% H₃PO₄, δ : 0.0 ppm). ⁵¹V-NMR chemical shifts are reported with respect to an

Table 3 Selected interatomic distances (Å) of salts 6, 7, 8, 9, 10, and 11

Bond	6	7	8	9	10	11
M–O ₁	2.278(3)	2.1467(11)	2.1966(14)	2.204(7)	2.1725(16)	2.1904(11)
M–O ₂	2.272(3)	2.2204(13)	2.1938(14)	2.068(7)	2.0395(17)	2.0453(17)
M–O ₃	2.300(3)	2.2325(14)	2.0587(15)	2.083(8)	1.9692(16)	2.1790(18)
M–X	1.940(4)–1.961(2)	1.942(2)–1.945(2)	2.3628(11)	2.352(10)–2.360(18)	2.347(18)–2.35(4)	2.3562(6)
M–O _m			1.687(4)–1.694(2)	1.73(5)	1.6242(19)	1.7152(13)–1.711(3)

6: $[PPN]_3[(P_3O_9)Mo(CO)_3]$; 7: $[PPN]_2[(P_3O_9)Mo(CO)_2(\eta^3-C_3H_5)]$; 8: $[PPN]_2[(P_3O_9)MoO_2Cl]$; 9: $[PPN]_2[(P_3O_9)MoOCl_2]$; 10: $[PPN]_2[(P_3O_9)VOF_2] \cdot 2CH_2Cl_2$; 11: $[PPN]_2[(P_3O_9)WO_2Cl]$; X = CO, Cl, F; O_m = terminal oxo atom.



external reference (VOCl_3 , δ : 0.0 ppm). J values are given in Hz. Infrared spectra were recorded on a Bruker TENSOR37 FT-IR spectrometer. X-ray data collections were carried out on a Siemens Platform three-circle goniometer with a CCD detector using $\text{Mo-K}\alpha$ radiation, $\lambda = 0.71073 \text{ \AA}$. An Agilent Technologies 5975C Mass Selective Detector operating in electron impact ionization mode was used to collect mass spectrometric data. Mass Spectrometry was performed on a Micromass Q-TOF ESI spectrometer in anionic mode. Samples were prepared in a glovebox and placed into a 1.5 mL GC vial with a septum to prevent exposure to moisture and oxygen and immediately analyzed by ESI mass spectrometry. All samples were prepared at a concentration of 0.1 mg mL^{-1} in acetonitrile. Prior to injection of the samples, the capillary tubing was flushed with 3 mL of dry acetonitrile (dried over 4 \AA molecular sieves for 3 days). The anions of interest were observed to be sensitive to instrument voltages. Optimum instrument settings were obtained with a capillary voltage of 3200 V, a MCP detector voltage of 2500 V, source temperature of $100 \text{ }^\circ\text{C}$ and a desolvation gas temperature of $150 \text{ }^\circ\text{C}$. EPR measurements were performed using a Bruker EMX EPR spectrometer, an ER 4199HS cavity and a Gunn diode microwave source producing X-band ($8\text{--}10 \text{ GHz}$) radiation.

[TBA]₄[P₄O₁₂]₅H₂O (1). Dowex® (18.0 g) was stirred with a solution of 0.5 M tetrabutylammonium hydroxide (120 mL) for 12 hours. The Dowex was loaded into a column and washed with distilled water (40 mL). An aqueous solution of $[\text{Na}]_4[\text{P}_4\text{O}_{12}]\cdot 5\text{H}_2\text{O}$ (4.0 g in 200 mL) was loaded into the column. The first 40 mL of eluent was discarded and the remainder was combined with effluent generated by washing the column with 40 mL of water after the elution was finished. The eluent was evaporated to dryness under vacuum at $40\text{--}50 \text{ }^\circ\text{C}$ yielding a white powder. The powder was washed with diethyl ether ($3 \times 20 \text{ mL}$) and further dried in a Schlenk line at $50 \text{ }^\circ\text{C}$ overnight. The solid was extracted with acetonitrile (20 mL) and the insoluble solids were filtered out and washed with acetonitrile (5 mL). The filtrate was concentrated down to 5 mL. A large excess of diethyl ether (250 mL) was added leading to the formation of a white solid that was collected by filtration, washed with diethyl ether ($3 \times 10 \text{ mL}$) and dried *in vacuo* in the Schlenk line overnight at $50 \text{ }^\circ\text{C}$ (6.0 g, 57% yield). Elem. Anal. Found: C, 55.68; H, 11.24; N, 4.13%. Calc. for $\text{C}_{64}\text{H}_{154}\text{N}_4\text{O}_{17}\text{P}_4$: C, 55.87; H, 11.28; N, 4.07%. δ_{H} (400 MHz; $\text{DMSO-}d_6$) 0.95 (t, $^3J_{\text{HH}} = 7.2$, $\text{CH}_3\text{--CH}_2$, 48H), 1.34 (m, $\text{CH}_2\text{--CH}_2$, 32H), 1.59 (m, $\text{CH}_2\text{--CH}_2$, 32H), 3.21 (t, $^3J_{\text{HH}} = 7.2$, $\text{CH}_2\text{--N}$, 32H) ppm. δ_{C} (100 MHz; $\text{DMSO-}d_6$) 57.2 (CH_2 , 16C), 23.2 (CH_2 , 16C), 19.1 (CH_2 , 16C), 13.4 (CH_3 , 16C) ppm. δ_{P} (121.5 MHz; CH_3CN) -23.2 ppm (s, 4P).

[TBA]₄[P₄O₁₂Mo(CO)₃]₂H₂O (2). $\text{Mo}(\text{CO})_3(\text{MeCN})_3$ (0.047 g, 0.15 mmol) was added to a solution of $[\text{TBA}]_4[\text{P}_4\text{O}_{12}]\cdot 5\text{H}_2\text{O}$ (0.20 g, 0.16 mmol) in acetonitrile (*ca.* 2 mL). The solution was stirred for one hour, during which it became dark yellow. The solution was concentrated to 0.5 mL and added to excess diethyl ether (*ca.* 20 mL), yielding a yellow oil. The solvent was then decanted and the remaining oil was treated with pentane and then dried under vacuum, affording an orange solid

(0.15 g, 69% yield). Elem. Anal. Found: C, 53.84; H, 10.37; N, 4.02%. Calc. for $\text{C}_{67}\text{H}_{148}\text{MoN}_4\text{O}_{17}\text{P}_4$: C, 53.58; H, 9.93; N, 3.73%. FT-IR (ATR): $\nu_{\text{max}}/\text{cm}^{-1}$ 1702vs (CO), 1708vs (CO) and 1874vs (CO) cm^{-1} . FT-IR (in acetonitrile) $\nu_{\text{max}}/\text{cm}^{-1}$ 1702vs (CO), 1708vs (CO) and 1874vs (CO). δ_{H} (400 MHz; $\text{DMSO-}d_6$) 0.95 (t, $^3J_{\text{HH}} = 7.2$, $\text{CH}_3\text{--CH}_2$, 48H), 1.34 (m, $\text{CH}_2\text{--CH}_2$, 32H), 1.57 (m, $\text{CH}_2\text{--CH}_2 = 7.3$, 32H), 3.17 (t, $^3J_{\text{HH}} = 7.2$, $\text{CH}_2\text{--N}$, 32H) ppm. δ_{C} (100 MHz; $\text{DMSO-}d_6$) 231.5 (s, CO, 3C), 57.4 (CH_2 , 16C), 23.0 (CH_2 , 16C), 19.1 (CH_2 , 16C), 13.4 (CH_3 , 16C) ppm. δ_{P} (121.5 MHz; CH_3CN) -16.6 (b, 1P), -19.6 (b, 2P), -25.6 (b, 1P) ppm.

[TBA]₃[P₄O₁₂Mo(CO)₂($\eta^3\text{-C}_3\text{H}_5$)] (3). $[\text{TBA}]_4[\text{P}_4\text{O}_{12}]\cdot 5\text{H}_2\text{O}$ (0.50 g, 0.36 mmol) was added to a suspension of $(\text{CH}_3\text{CN})_2\text{-Mo}(\text{CO})_2(\eta^3\text{-C}_3\text{H}_5)\text{Cl}$ (0.12 g, 0.39 mmol) in acetonitrile (*ca.* 2 mL). The suspension became homogeneous after addition of the tetrametaphosphate salt. The solution was stirred for 30 min whereupon the solvent was evaporated yielding a yellow solid. The solid was dissolved in dichloromethane (*ca.* 0.5 mL) and diethyl ether was added dropwise up to saturation. The solution was allowed to stand overnight, yielding a yellow crystalline solid which was collected by filtration and washed with tetrahydrofuran ($3 \times 1 \text{ mL}$) (0.25 g, 56%); Elem. Anal. Found: C, 51.49; H, 8.90; N, 3.36%. Calc. for $\text{C}_{53}\text{H}_{113}\text{MoN}_3\text{O}_{14}\text{P}_4$: C, 51.49; H, 9.21; N, 3.40%. FT-IR (ATR): $\nu_{\text{max}}/\text{cm}^{-1}$ 1789vs (CO) and 1898vs (CO). FT-IR (in acetonitrile) $\nu_{\text{max}}/\text{cm}^{-1}$ 1807s (CO) and 1914vs (CO). δ_{H} (400 MHz; $\text{DMSO-}d_6$) 0.66 (d, H^{anti} , $^3J_{\text{HH}} = 9.2$, 2H), 0.95 (t, $^3J_{\text{HH}} = 7.2$, $\text{CH}_3\text{--CH}_2$, 36H), 1.34 (m, $\text{CH}_2\text{--CH}_2$, 24H), 1.57 (m, $\text{CH}_2\text{--CH}_2$, 24H), 3.17 (t, $^3J_{\text{HH}} = 7.2$, $\text{CH}_2\text{--N}$, 24H), 3.17 (b, H^{syn} , 2H), 3.92 (b, H^{apical} , 1H) ppm. δ_{C} (100 MHz; $\text{DMSO-}d_6$) 231.5 (s, CO, 2C), 63.7 (s, C^{allyl} , 1C), 61.0 (s, C^{allyl} , 2C), 57.4 (CH_2 , 12C), 23.0 (CH_2 , 12C), 19.1 (CH_2 , 12C), 13.4 (CH_3 , 12C) ppm. δ_{P} (162 MHz; CH_3CN) -17.3 (t, $^2J_{\text{PP}} = 27.0$, 1P), -21.6 (t, $^2J_{\text{PP}} = 27.0$, 2P), -25.7 (t, $^2J_{\text{PP}} = 27.0$, 1P) ppm. ESI-MS (m/z , CH_2Cl_2): 995.3415 ($[\text{TBA}]_2\text{M}^-$).

[TBA]₃[P₄O₁₂MoO₂Cl] (4). $\text{MoO}_2\text{Cl}_2(\text{OSMe}_2)_2$ (0.052 g, 0.15 mmol) was added to a solution of $[\text{TBA}]_4[\text{P}_4\text{O}_{12}]\cdot 5\text{H}_2\text{O}$ (0.20 g, 0.15 mmol) in acetonitrile (*ca.* 2 mL). The solution was stirred for one hour and the solvent was removed *in vacuo*, affording an oil. The oil was dissolved in dichloromethane (*ca.* 0.5 mL) and diethyl ether was added dropwise up to saturation. The solution was allowed to stand overnight providing light yellow crystals which were collected by filtration and washed with tetrahydrofuran ($3 \times 1 \text{ mL}$) (0.12 g, 68%). Elem. Anal. Found: C, 47.54; H, 8.83; N, 3.44%. Calc. for $\text{C}_{48}\text{H}_{108}\text{ClMoN}_3\text{O}_{14}\text{P}_4$: C, 47.78; H, 9.02; N, 3.48%. FT-IR (ATR): $\nu_{\text{max}}/\text{cm}^{-1}$ 880m ($\text{M}=\text{O}$) and 896m ($\text{M}=\text{O}$). δ_{H} (400 MHz; $\text{DMSO-}d_6$) 0.87 (t, $^3J_{\text{HH}} = 7.2$, $\text{CH}_3\text{--CH}_2$, 36H), 1.27 (m, $\text{CH}_2\text{--CH}_2$, 24H), 1.52 (m, $\text{CH}_2\text{--CH}_2$, 24H), 3.14 (t, $^3J_{\text{HH}} = 7.2$, $\text{CH}_2\text{--N}$, 24H) ppm. δ_{C} (100 MHz; $\text{DMSO-}d_6$) 57.2 (CH_2 , 12C), 23.2 (CH_2 , 12C), 19.1 (CH_3 , 12C), 13.4 (CH_3 , 12C) ppm. δ_{P} (162 MHz; CH_3CN) -21.2 (b, 4P) ppm.

[TBA]₃[P₄O₁₂VOF₂]₂Et₂O (5). VOF_3 (0.021 g, 0.17 mmol) was added to a solution of $[\text{TBA}]_4[\text{P}_4\text{O}_{12}]\cdot 5\text{H}_2\text{O}$ (0.20 g, 0.15 mmol) in acetonitrile (*ca.* 2 mL). The solution was stirred for one hour and the solvent was then removed *in vacuo*. The remaining solid was dissolved in dichloromethane, and



diethyl ether was added dropwise up to saturation. The solution was allowed to stand overnight, affording light yellow crystals which were collected by filtration and washed with tetrahydrofuran (3×1 mL) (0.10 g, 56%). Elem. Anal. Found: C, 51.21; H, 9.77; N, 4.09%. Calc. for $C_{52}H_{118}F_2N_3O_{14}P_4V$: C, 51.10; H, 9.73; N, 3.44%. FT-IR (ATR) $\nu_{\max}/\text{cm}^{-1}$ 912m ($\text{V}=\text{O}$). δ_{H} (400 MHz; DMSO- d_6) 0.88 (t, $^3J_{\text{HH}} = 7.2$, CH_3-CH_2 , 36H), 1.28 (m, CH_2-CH_2 , 24H), 1.53 (m, CH_2-CH_2 , 24H), 3.15 (t, $^3J_{\text{HH}} = 7.1$, CH_2-N , 24H) ppm. δ_{C} (100 MHz; DMSO- d_6) 57.4 (CH_2 , 12C), 23.0 (CH_2 , 12C), 19.1 (CH_2 , 12C), 13.4 (CH_3 , 12C) ppm. δ_{V} (104 MHz; DMSO- d_6) -594.5 ppm.

[PPN] $_2$ [(P $_3$ O $_9$)Mo(CO) $_2$ (η^3 -C $_3$ H $_5$)] (7). [PPN] $_3$ [P $_3$ O $_9$] \cdot H $_2$ O (0.48 g, 0.26 mmol) was added to a suspension of (CH $_3$ CN) $_2$ -Mo(CO) $_2$ (η^3 -C $_3$ H $_5$)Cl (0.07 g, 0.26 mmol) in acetonitrile (*ca.* 2 mL). Upon addition of the trimetaphosphate salt, the suspension became homogeneous. After stirring for *ca.* 30 seconds yellow solid precipitated. The slurry was stirred for additional 30 minutes in order to ensure completion of the reaction. The solid was collected by filtration and washed with tetrahydrofuran (3×1 mL) (0.25 g, 78% yield). Elem. Anal. Found: C, 61.95; H, 4.45; N, 2.20%. Calc. for $C_{77}H_{63}MoN_2O_{11}P_7$: C, 61.36; H, 4.35; N, 1.86%. FT-IR (ATR): $\nu_{\max}/\text{cm}^{-1}$ 1812vs (CO) and 1918s (CO). FT-IR (in acetonitrile) $\nu_{\max}/\text{cm}^{-1}$ 1818vs (CO) and 1923vs (CO) cm^{-1} . δ_{H} (400 MHz; DMSO- d_6) 0.81 (d, $^2J_{\text{HH}} = 9.2$, H^{anti} , 2H), 3.15 (m, H^{syn} and H^{apical} , 3H), 7.57 (m, H^{Ar} , 48H), 7.72 (m, H^{Ar} , 12H) ppm. δ_{C} (100 MHz; DMSO- d_6) 230.0 (s, CO, 2C), 133.7 (m, *para*, 12C), 132.0 (m, *ortho*, 24C), 129.6 (m, *meta*, 24C), 126.3 (dd, $^1J_{\text{PC}} = 107.1$, $^3J_{\text{PC}} = 1.5$, *ipso*, 12C), 73.7 (s, C^{allyl} , 1C), 59.9 (s, C^{allyl} , 2C) ppm. δ_{P} (162 MHz; CH $_3$ CN) +22.3 (s, 4P), -12.9 (s, 1P), -13.8 (s, 2P) ppm. ESI-MS (m/z), CH $_2$ Cl $_2$: 969.9943 (M - PPN) $^-$, 215.9049 (M) $^{2-}$.

[PPN] $_2$ [(P $_3$ O $_9$)MoO $_2$ Cl] (8). [PPN] $_3$ [P $_3$ O $_9$] \cdot H $_2$ O (1.6 g, 0.86 mmol) was added to a solution of MoO $_2$ Cl $_2$ (OSMe $_2$) $_2$ (0.31 g, 0.84 mmol) in acetonitrile (*ca.* 4 mL). The solution was stirred for 2 hours, during which the formation of white precipitate was observed. The precipitate was collected by filtration, washed with tetrahydrofuran (3×1 mL) and dried *in vacuo* (0.68 g, 56% yield). Elem. Anal. Found: C, 59.18; H, 4.12; N, 2.07%. Calc. for $C_{72}H_{60}ClMoN_2O_{11}P_7$: C, 58.53; H, 4.09; N, 1.90%. FT-IR (ATR): $\nu_{\max}/\text{cm}^{-1}$ 898m (M=O) and 923m (M=O). δ_{H} (400 MHz; DMSO- d_6) 7.57 (m, H^{Ar} , 48H), 7.72 (m, H^{Ar} , 12H) ppm. δ_{C} (100 MHz; DMSO- d_6) 133.7 (m, *para*, 12C), 132.3 (m, *ortho*, 24C), 129.9 (m, *meta*, 24C), 126.33 (dd, $^1J_{\text{PC}} = 107.1$, $^3J_{\text{PC}} = 1.5$, *ipso*, 12C). δ_{P} (121.5 MHz; CH $_3$ CN) +22.3 (s, 4P), -18.2 (t, $^2J_{\text{PP}} = 18.1$, 1P), -19.0 (d, $^2J_{\text{PP}} = 18.1$, 2P) ppm. ESI-MS (m/z , CH $_2$ Cl $_2$): 939.9281 (M - PPN) $^-$, 200.8703 (M) $^{2-}$.

[PPN] $_2$ [(P $_3$ O $_9$)MoOCl $_2$] (9). To a thawing stirring solution of [PPN] $_3$ [P $_3$ O $_9$] \cdot H $_2$ O (0.50 g, 0.27 mmol) in acetonitrile (*ca.* 4 mL) was added dropwise a thawing acetonitrile solution (*ca.* 4 mL) of MoOCl $_3$ (0.058 g, 0.27 mmol). The color changed after 15 min from brown to yellow and then to a green-yellow color. The reaction mixture was stirred overnight, during which time the color turned into green and a solid precipitate was observed. After this time, the solvent was removed *in vacuo* and the residue was dissolved in dichloromethane (*ca.* 4 mL) and the product was precipitated using diethyl ether (*ca.* 8 mL). The

residue was separated from the clear solution by decantation. Dissolution, precipitation, and decantation were repeated three times to get rid of PPnCl coproduct. The residue was then dried under vacuum and then crystallized overnight using a 2 : 1 CH $_2$ Cl $_2$ -toluene mixture to provide the desired product as green crystals which were collected by filtration and washed with a mixture 1 : 3 acetonitrile-diethyl ether (3×2 mL) (0.30 g, 75%). Elem. Anal. Found: C, 57.81; H, 4.20; N, 1.75%. Calc. for $C_{72}H_{60}Cl_2MoN_2O_{10}P_7$: C, 57.77; H, 4.04; N, 1.87%. FT-IR (ATR) $\nu_{\max}/\text{cm}^{-1}$ 932m (M=O). δ_{H} (300 MHz; DMSO- d_6) 7.56 (m, H^{Ar} , 48H), 7.71 (m, H^{Ar} , 12H) ppm. δ_{C} (100 MHz; DMSO- d_6): 133.6 (m, *para*, 12C), 131.8 (m, *ortho*, 24C), 129.7 (m, *meta*, 24C), 126.8 (dd, $^1J_{\text{PC}} = 107.1$, $^3J_{\text{PC}} = 1.5$, *ipso*, 6C) ppm. δ_{P} (121.5 MHz, CH $_3$ CN): +21.54 (s, 4P) ppm. EPR (CH $_2$ Cl $_2$, g-value): 1.934. ESI-MS (m/z , CH $_2$ Cl $_2$): 421.72 (MH) $^-$, 385.74 (M - Cl) $^-$ and 341.75 (M - PO $_3$) $^-$.

[PPN] $_2$ [(P $_3$ O $_9$)VOF $_2$] \cdot 2CH $_2$ Cl $_2$ (10). VOF $_3$ (0.022 g, 0.16 mmol) was added to a solution of [PPN] $_3$ [P $_3$ O $_9$] \cdot H $_2$ O (0.30 g, 0.16 mmol) in acetonitrile (*ca.* 2 mL). The solution was allowed to stir for one hour and the volatile materials were removed *in vacuo*. The remaining yellow oil was dissolved in dichloromethane, and diethyl ether was added up to saturation. The solution was allowed to stand overnight, affording light yellow crystals which were collected by filtration and washed with tetrahydrofuran (3×1 mL) (0.15 g, 59%). Elem. Anal. Found: C, 55.16; H, 4.65; N, 1.52%. Calc. for $C_{74}H_{64}Cl_4F_2N_2O_{10}P_7V$: C, 55.94; H, 4.06; N, 1.76%. FT-IR (ATR): $\nu_{\max}/\text{cm}^{-1}$ 930m ($\text{V}=\text{O}$). δ_{H} (400 MHz; DMSO- d_6) 7.59 (m, H^{Ar} , 60H) ppm. δ_{C} (100 MHz; DMSO- d_6): 133.6 (m, *para*, 12C), 131.9 (m, *ortho*, 24C), 129.5 (m, *meta*, 24C), 126.8 (dd, $^1J_{\text{PC}} = 107.1$, $^3J_{\text{PC}} = 1.5$, *ipso*, 6C) ppm. δ_{P} (121.5 MHz; CH $_3$ CN) +22.16 (s, 4P), -18.19 (s, 1P), -22.15 (s, 2P). δ_{V} (104 MHz; DMSO- d_6): -593.4 ppm.

[PPN] $_2$ [(P $_3$ O $_9$)WO $_2$ Cl] (11), method 1. A thawing acetonitrile solution (*ca.* 4 mL) of [PPN] $_3$ [P $_3$ O $_9$] \cdot H $_2$ O (5.0 g, 0.27 mmol) was added dropwise to a stirring, thawing acetonitrile suspension (4 mL) of WOCl $_4$ (0.091 g, 0.27 mmol). The suspension was allowed to stir for 2 hours during which time the reaction mixture became colorless and homogeneous. After this time, the solvent was removed under vacuum to give a white residue. The obtained white residue was dissolved in acetonitrile (*ca.* 4 mL) and precipitated using diethyl ether (*ca.* 8 mL). The mother liquor was decanted. Dissolution and precipitation were repeated three times, in order to get rid of PPnCl coproduct. The white residue was recrystallized by slow diffusion of diethyl ether into an acetonitrile, or acetonitrile-dichloromethane (1 : 1), solution of the residue to give colorless crystals which were collected by filtration and washed with a mixture 1 : 3 acetonitrile-diethyl ether (3×2 mL) (0.32 g, 77%). Elem. Anal. Found: C, 55.38; H, 3.90; N, 1.75%. Calc. for $C_{72}H_{60}ClN_2O_{11}P_7W$: C, 55.24; H, 3.86; N, 1.79%. FT-IR (ATR) $\nu_{\max}/\text{cm}^{-1}$ 907m (W=O), 938m (W=O). δ_{H} (300 MHz; DMSO- d_6) 7.56 (m, H^{Ar} , 48H), 7.71 (m, H^{Ar} , 12H) ppm. δ_{C} (100 MHz; DMSO- d_6) 133.6 (m, *para*, 12C), 131.8 (m, *ortho*, 24C), 129.7 (m, *meta*, 24C), 126.8 (dd, $^1J_{\text{PC}} = 107.1$, $^3J_{\text{PC}} = 1.5$, *ipso*, 6C) ppm. δ_{P} (121.5 MHz; CH $_3$ CN) +22.20 (s, 4P), -17.79 (m, 3P)



ppm. ESI-MS (m/z , CH_2Cl_2) 488.80 (MH^-), 452.81 ($\text{M} - \text{Cl}^-$), 408.83 ($\text{M} - \text{PO}_3^-$), 243.89 (M^{2-}).

Method 2. A thawing solution of $[\text{PPN}]_3[\text{P}_3\text{O}_9]\cdot\text{H}_2\text{O}$ (0.50 g, 0.27 mmol) in acetonitrile (*ca.* 4 mL) was added dropwise to a stirring, thawing solution of $\text{WO}_2\text{Cl}_2(\text{MeCN})_2$ (0.099 g, 0.27 mmol) in acetonitrile (*ca.* 4 mL). The solution was allowed to stir for two hours whereupon the solvent was removed *in vacuo*. The obtained white residue was dissolved in acetonitrile (*ca.* 4 mL) and precipitated using diethyl ether (*ca.* 8 mL). The mother liquor was decanted. Dissolution and precipitation were repeated three times, in order to get rid of PPNCl coproduct. White crystals (0.30 g, 72%) suitable for X-ray studies were obtained by crystallization of the obtained white residue by slow diffusion of diethyl ether into acetonitrile solution. Spectroscopic data for the complex synthesized using this method were identical to those obtained according method 1.

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References

- S. Kamimura, T. Matsunaga, S. Kuwata, M. Iwasaki and Y. Ishii, *Inorg. Chem.*, 2004, **43**, 6127–6129.
- B. Boonchom, *J. Optoelectron. Adv. Mater.*, 2009, **1**, 103–114.
- M. Trojan and P. Sulcová, *Dyes Pigm.*, 2000, **47**, 291–294.
- H. Onoda, K. Okumoto, A. Nakahira and I. Tanaka, *Materials*, 2009, **2**, 1–9.
- K. Kanao, Y. Ikeda, K. Kimura, S. Kamimura, Y. Tanabe, Y. Mutoh, M. Iwasaki and Y. Ishii, *Organometallics*, 2013, **32**, 527–537.
- Y. Ikeda, T. Yamaguchi, K. Kanao, K. Kimura, S. Kamimura, Y. Mutoh, Y. Tanabe and Y. Ishii, *J. Am. Chem. Soc.*, 2008, **130**, 16856–16857.
- S. Kamimura, S. Kuwata, M. Iwasaki and Y. Ishii, *Inorg. Chem.*, 2004, **43**, 399–401.
- M. Montag, C. R. Clough, P. Müller and C. C. Cummins, *Chem. Commun.*, 2011, **47**, 662–664.
- S. Mohamady and S. D. Taylor, *Org. Lett.*, 2013, **15**(11), 2612–2615.
- P. M. Laügt, J. Guitel and I. T. E. G. Bassi, *Acta Crystallogr., Sect. B: Struct. Crystallogr. Cryst. Chem.*, 1972, **28**, 201–208.
- A. G. Nord, *Acta Chem., Scand. Ser. A*, 1983, **37**, 539–543.
- M. Trojan, D. Brandová and Z. Solc, *Thermochim. Acta*, 1987, **110**, 343–358.
- M. Trojan, *Mater. Lett.*, 1989, **8**, 247–252.
- W. Gunsser, D. Fruehauf, K. Rohwer and A. Zimmermann, *J. Solid State Chem.*, 1989, **82**, 43–51.
- M. Trojan and J. Palme, *Thermochim. Acta*, 1993, **224**, 165–175.
- M. Trojan, J. Palme and P. Mazan, *Thermochim. Acta*, 1993, **224**, 177–182.
- B. Boonchoma, M. Thongkamc, S. Kongtaweelert and N. Vittayakorn, *J. Alloys Compd.*, 2009, **486**, 689–692.
- B. Boonchom and N. Vittayakorn, *J. Mater. Sci.*, 2010, **45**, 1459–1463.
- H. Hemissi, M. Rzaigui and Z. A. Al Othman, *Acta Crystallogr., Sect. E: Struct. Rep. Online*, 2010, **66**, m186–m187.
- C. J. Besecker and W. G. Klemperer, *J. Organomet. Chem.*, 1981, **205**, C31–C32.
- C. J. Besecker, V. W. Day and W. G. Klemperer, *Organometallics*, 1985, **4**, 564–570.
- V. W. Day, W. G. Klemperer and D. J. Main, *Inorg. Chem.*, 1990, **29**, 2345–2355.
- V. W. Day, W. G. Klemperer, S. P. Lockledge and D. J. Main, *J. Am. Chem. Soc.*, 1990, **112**, 2031–2033.
- V. W. Day, T. A. Eberspacher, W. G. Klemperer, R. P. Planalp, P. W. Schiller, A. Yagasaki and B. Zhong, *Inorg. Chem.*, 1993, **32**, 1629–1637.
- W. G. Klemperer and B. Zhong, *Inorg. Chem.*, 1993, **32**, 5821–5826.
- K. N. Han, D. Whang, H. J. Lee, Y. Do and K. Kim, *Inorg. Chem.*, 1993, **32**, 2597–2599.
- V. W. Day, T. A. Eberspacher, W. G. Klemperer and B. Zhong, *J. Am. Chem. Soc.*, 1994, **116**, 3119–3120.
- S. Ryu, D. Whang, J. Kim, W. Yeo and K. Kim, *J. Chem. Soc., Dalton Trans.*, 1993, 205–209.
- D. Attanasio, F. Bachechi and L. Suber, *J. Chem. Soc., Dalton Trans.*, 1993, 2373–2378.
- S. Greenfield and M. Clift, *Analytical chemistry of the condensed phosphates*, Pergamon Press, 1975, vol. 57, p. 190.
- J. Nakamoto, *Infrared and Raman Spectra of Inorganic and Coordination Compounds – Part A*, John Wiley & Sons, Inc., 6th edn, 2009, vol. B, pp. 222–223.
- H. H. Kung, *Transition Metal Oxides: Surface Chemistry and Catalysis*, Elsevier Science Publishers B.V., 1st edn, 1989, p. 67.
- J. Nakamoto, *Infrared and Raman Spectra of Inorganic and Coordination Compounds*, John Wiley & Sons, Inc., 5th edn, 1997, vol. B, pp. 168–169.
- H. W. Roesky, I. Leichtweis and M. Noltemeyer, *Inorg. Chem.*, 1993, **32**, 5102–5104.
- J. Harrela, H. Wonga, B. Daveb, B. Dunna and L. Nazar, *J. Non-Cryst. Solids*, 1995, **225**, 319–324.
- K. Kanamoria, K. Nishidaa, N. Miyataa, K. Okamoto, Y. Miyoshic, A. Tamura and H. Sakurai, *J. Inorg. Biochem.*, 2001, **86**, 649–656.
- M. D. Hoops and B. S. Ault, *J. Mol. Struct.*, 2002, **616**, 91–101.
- C. R. Clough, J. S. Silvia, P. Müller and C. C. Cummins, *Inorg. Chim. Acta*, 2012, **382**, 195–198.
- W. Kläui, A. Müller, W. Eberspach, R. Boese and I. Goldberg, *J. Am. Chem. Soc.*, 1987, **109**, 164–169.
- M. D. Curtis and K.-B. Shiu, *Inorg. Chem.*, 1985, **24**, 1213–1218.
- N. L. Armanasco, M. V. Baker, M. R. North, B. W. Skelton and A. H. White, *J. Chem. Soc., Dalton Trans.*, 1998, 1145–1149.



- 42 Y. D. Ward, L. A. Villanueva, G. D. Allred, S. C. Payne, M. A. Semones and L. S. Liebeskind, *Organometallics*, 1995, **14**, 4132–4156.
- 43 J. F. Hartwig, *Organotransition Metal Chemistry: From Bonding to Catalysis*, University Science Books, 2010, pp. 27–32.
- 44 F. A. Cotton, G. Wilkinson, C. A. Murillo and M. Bochmann, *Advanced Inorganic Chemistry*, Wiley-Interscience, New York, 6th edn, 1999, pp. 637–638.
- 45 *gNMR V5*, Adept Scientific plc, Letchworth Herts, UK, 2003.
- 46 R. H. Crabtree, *The Organometallic Chemistry of the Transition Metals*, John Wiley & Sons, Inc., 4th edn, 2005, pp. 131–133.
- 47 G. Desiraju and T. Steiner, *The Weak Hydrogen Bond: In Structural Chemistry and Biology*, Oxford University Press, 2001, pp. 12–16.
- 48 M. N. Burnett and C. K. Johnson, ORTEP-III: Oak Ridge Thermal Ellipsoid Plot Program for Crystal Structure Illustrations, *Oak Ridge National Laboratory Report ORNL-6895*, 1996.
- 49 F. J. Arnaiz, R. Aguado, M. R. Pedrosa and A. D. Cian, *Inorg. Chim. Acta*, 2003, **347**, 33–40.
- 50 P. Baker, M. Drew, A. Johans and M. Meehan, *J. Chem. Crystallogr.*, 1998, **28**, 839–841.
- 51 D. Shriver and P. Atkins, *Inorganic Chemistry*, W. H. Freeman and Company, 5th edn, 2010, pp. 461–462.
- 52 W. G. Klemperer and D. J. Main, *Inorg. Chem.*, 1990, **29**, 2355–2360.
- 53 D. P. Tate, W. R. Knipple and J. M. Augl, *Inorg. Chem.*, 1962, **1**, 433–434.
- 54 D. A. Clark, D. L. Jones and R. J. Mawby, *J. Chem. Soc., Dalton Trans.*, 1980, **4**, 565–569.
- 55 V. C. Gibson, T. P. Kee and A. Shaw, *Polyhedron*, 1990, **9**, 2293–2298.

