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A reduced radial potential energy function for the halogen bond and the hydrogen bond in complexes $B \cdots XY$ and $B \cdots HX$, where X and Y are halogen atoms

It is shown by considering 76 halogen- and hydrogen-bonded complexes B···XY and B···HX (where B is a Lewis base N₂, CO, C₂H₂, C₂H₄, H₂S, HCN, H₂O, PH₃ or NH₃ and X, Y are F, Cl, Br or I) that the intermolecular stretching force constants k_{σ} (determined from experimental centrifugal distortion constants *via* a simple

model) and the intermolecular dissociation energies D_{σ} (calculated at the CCSD(T)(F12*)/cc-pVDZ-F12 level of theory) are related by $D_{\sigma} = C_{\sigma}k_{\sigma}$, where $C_{\sigma} = 1.50(3) \times 10^3 \text{ m}^2 \text{ mol}^{-1}$. This suggests that one-

dimensional functions implying direct proportionality of D_{σ} and k_{σ} , (e.g. a Morse or Rydberg function)

might serve as reduced radial potential energy functions for such complexes.

Anthony C. Legon

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1. Introduction

During the last decade there has been a rapid growth of interest in the halogen bond across the disciplines of Chemistry, Materials Science and Biology,¹ especially in its parallels with the hydrogen bond.² The halogen bond³ is represented conventionally by the three centred dots in $B \cdots X-R$, where the halogen atom X of the molecule X-R interacts with a nucleophilic acceptor atom/centre Z of a simple Lewis base B or of a much larger molecule. There has naturally followed discussion of both the radial and angular potential energy functions associated with such interactions.^{4,5} This article is concerned with the characteristics of the one-dimensional function that describes the variation of the energy with the intermolecular distance $r(Z \cdots X)$, that is the intermolecular radial potential energy function. Attention will be focussed initially on several series of halogen-bonded complexes $B \cdot \cdot XY$, where B is one of the Lewis bases N₂, CO, C₂H₂, C₂H₄, H₂S, HCN, H₂O, PH₃ and NH₃ and XY is one of the dihalogen molecules F₂, ClF, Cl₂, BrCl, Br₂ and ICl. Thereafter, the corresponding series of hydrogen-bonded complexes B···HX, where X = F, Cl, Br or I, will be discussed.

Two important characteristics of such a one-dimensional radial potential energy function are the intermolecular dissociation energy D_{σ} and the intermolecular quadratic stretching force constant k_{σ} , both of which provide a measure of the strength of the (generally weak) interaction of B and XY. D_{σ} is the energy required to take the complex from the equilibrium distance $r_{\rm e}$ along r to infinite separation, while k_{σ} is the curvature of the

function at r_e and provides a measure of the restoring force per unit infinitesimal displacement from r_e along the same path.

In this article, a direct proportionality of D_{σ} and k_{σ} is established for many complexes of the halogen-bonded type B···XY and the hydrogen-bonded type B···HX. The values of D_{σ} were obtained by means of *ab initio* calculations at the explicitly correlated level of theory CCSD(T)(F12*)/cc-pVDZ-F12, after correction for basis set superposition error, while the k_{σ} values were those established from the rotational spectra of B...XY or B···HX through interpretation of experimental zero-point centrifugal distortion constants in terms of a simple model. The k_{σ} are therefore zero-point, rather than equilibrium, values and are also subject to errors introduced by the assumption of rigid subunits B and XY (or HX). The fact that the equation $D_{\sigma} = C_{\sigma}k_{\sigma}$ describes, with the same constant of proportionality C_{σ} , the behaviour of a large number of complexes B···XY and B····HX suggests that, for example, a reduced Morse function $V(r) = D_{\sigma} [1 - e^{-a(r-r_c)}]^2$, with $a = (2C_{\sigma})^{-\frac{1}{2}}$, or a reduced Rydberg function $V(r) = -D_{\sigma}[1 + b(r - r_e)]e^{-b(r - r_e)}$, with $b = C_{\sigma}^{-\frac{1}{2}}$, could be useful to describe the radial intermolecular potential energy functions in such molecules.

2. Methods

2.1 k_{σ} values from centrifugal distortion constants

In the quadratic approximation and with the assumption of rigid subunits B and XY unperturbed by the weak interaction, Millen⁶ showed that the quartic centrifugal distortion constant D_I for a linear or symmetric top complex B···XY (or B···HX) is

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School of Chemistry, University of Bristol, Cantock's Close, Bristol BS8 1TS, UK. E-mail: a.c.legon@bristol.ac.uk

simply related to the intermolecular stretching force constant by the expression

$$k_{\sigma} = (16\pi^2 \mu B^3 / D_J) [1 - B/B_{\rm B} - B/B_{\rm XY}], \tag{1}$$

in which *B*, *B*_B and *B*_{XY} are strictly the equilibrium rotational constants of the complex, B and XY, respectively, and $\mu = m_{\rm B}m_{\rm XY}/(m_{\rm B} + m_{\rm XY})$ is the reduced mass for the intermolecular motion in question. When B···XY is an asymmetric-top molecule of *C*_{2v} symmetry, in which XY lies along the *C*₂ axis, the corresponding centrifugal distortion constant Δ_J , obtained by fitting the rotational transitions by means of a Hamiltonian that employs the Watson A reduction in the I^r representation, is given by

$$k_{\sigma} = (8\pi^2 \mu / \Delta_J) [B^3(1-b) + C^3(1-c)], \qquad (2)$$

in which $b = (B/B_B) + (B/B_{XY})$ and $c = (C/B_B) + (C/B_{XY})$, and *B* and *C* are equilibrium rotational constants of the complex. When Δ_J is used, eqn (2) holds whether the C_{2v} molecule $B \cdots XY$ is planar, as in $C_2H_2 \cdots XY$, or non-planar, as in $C_2H_4 \cdots XY$ where XY is perpendicular to the plane containing the ethene nuclei.

Identical equations, with XY replaced by HX, apply to the corresponding members of the series of hydrogen-bonded complexes $B \cdots HX$.

Equilibrium values of the spectroscopic constants required for use in eqn (1) and (2) to obtain k_{σ} have not been determined experimentally for complexes of the type considered here and in general only zero-point quantities are available. To allow progress, we invoke a type of 'effective' rigid-rotor approximation, namely the use of zero-point centrifugal distortion constants and rotational constants in these equations in place of their equilibrium counterparts. The utility of this approximation can be judged, in general, by the conclusions presented in this article and, in particular, by reference in Section 4 to the examples of the simple linear complexes OC···HX and OC···XY, for which tests of the approximation are available.

Values of k_{σ} calculated by means of the appropriate eqn (1) or (2) for a wide range of complexes B···XY,⁷⁻⁵² where the Lewis base B is one of N₂, CO, C₂H₂, C₂H₄, H₂S, HCN, H₂O, PH₃ and NH₃ and XY is one of the dihalogen molecules F₂, ClF, Cl₂, BrCl, Br₂ and ICl are given in Tables 1 and 2, while Table 3 collects together the values of k_{σ} for the corresponding set of

Table 1 Some observed and calculated properties of halogen-bonded complexes $B \cdots X_2$ involving non-polar dihalogen molecules X_2^a

| | Dihalogen molecules X ₂ | | | | | | | | | | | |
|-----------------|------------------------------------|----------|----------------------------|---|-------------------------|------|--|--------------------------------|-------------------------|------|----------------------------|--------------------------------|
| Lewis base B | Difluorine | | Dichlorine Cl ₂ | | | | Dibromine Br ₂ | | | | | |
| | $k_{\sigma}/(N m^{-1})$ | Ref. | $D_{\sigma}/(kJ mol^{-1})$ | $r(Z \cdot \cdot \cdot X_i)/\mathring{A}$ | $k_{\sigma}/(N m^{-1})$ | Ref. | $D_{\sigma}/(\mathrm{kJ} \mathrm{mol}^{-1})$ | $r(Z \cdot \cdot \cdot X_i)/Å$ | $k_{\sigma}/(N m^{-1})$ | Ref. | $D_{\sigma}/(kJ mol^{-1})$ | $r(Z \cdot \cdot \cdot X_i)/Å$ |
| OC | _ | _ | _ | _ | 3.68(1) | 12 | 5.19 | 3.145 | 5.03(2) | 20 | 7.26 | 3.111 |
| C_2H_2 | — | _ | — | — | 5.61(1) | 13 | 7.45 | 3.146 | 7.80(3) | 21 | 10.69 | 3.106 |
| C_2H_4 | | | | | 5.88(2) | 14 | 8.61 | 3.092 | 8.8(2) | 22 | 12.93 | 3.004 |
| H_2S | 2.34(1) | 7 | 3.43 | 3.143 | 6.23(2) | 15 | 8.53 | 3.246 | 9.8(2) | 23 | 13.68 | 3.131 |
| H_3P | _ | | — | — | 5.58(2) | 16 | 8.34 | 3.222 | 9.79(3) | 24 | 15.07 | 3.013 |
| HCN | 2.62(1) | 8 | 4.15 | 2.811 | 6.55(2) | 17 | 9.71 | 2.921 | _ `` | _ | | |
| H_2O | 3.66(1) | 9 and 10 | 4.63 | 2.696 | 7.98(3) | 18 | 10.66 | 2.808 | 9.9(2) | 25 | 14.64 | 2.804 |
| H_3N | 4.67(1) | 11 | 6.59 | 2.679 | 12.73(2) | 19 | 17.85 | 2.681 | 18.5(4) | 26 | 27.36 | 2.601 |

^{*a*} Values of k_{σ} are either taken directly from the reference having the number indicated in columns 3, 7 or 11, as appropriate, or are recalculated from the centrifugal distortion constant D_J or Δ_J given therein by using eqn (1) or (2). The quoted error is that transmitted by the error in the distortion constant. D_{σ} and $r(Z \cdots X_i)$ are equilibrium values calculated *ab initio* at the CCSD(T)(F12*)/cc-pVDZ-F12 level of theory (see text). $r(Z \cdots X_i)$ is the distance from the acceptor atom/centre Z in the Lewis base B to the inner halogen atom X_i.

Table 2 Some observed and calculated properties of halogen-bonded complexes B···XY involving polar dihalogen molecules XY^a

| Lewis base B | Dihalogen molecules XY | | | | | | | | | | | |
|------------------|---------------------------|------|--|------------------------------|---------------------------|------|--|--|-------------------------|------|--|--|
| | Chlorine monofluoride ClF | | | | Bromine monochloride BrCl | | | | Iodine monochloride ICl | | | |
| | $k_{\sigma}/(N m^{-1})$ | Ref. | $D_{\sigma}/(\mathrm{kJ} \mathrm{mol}^{-1})$ | $r(Z \cdot \cdot \cdot X)/Å$ | $k_{\sigma}/(N m^{-1})$ | Ref. | $D_{\sigma}/(\mathrm{kJ} \mathrm{mol}^{-1})$ | $r(\mathbf{Z}\cdot\cdot\cdot\mathbf{X})/\mathbf{\mathring{A}}$ | $k_{\sigma}/(N m^{-1})$ | Ref. | $D_{\sigma}/(\mathrm{kJ} \mathrm{\ mol}^{-1})$ | $r(\mathbf{Z}\cdot\cdot\cdot\mathbf{X})/\mathbf{\mathring{A}}$ |
| | | | | | | | | | | | | |
| N_2 | 5.00(3) | 27 | 6.28 | 2.918 | 4.40(2) | 35 | 5.63 | 3.106 | 5.37(2) | 44 | 7.08 | 3.187 |
| OC | 7.04(2) | 28 | 10.56 | 2.772 | 6.27(5) | 36 | 9.20 | 3.006 | 8.00(3) | 45 | 12.73 | 3.003 |
| C_2H_2 | 10.01(2) | 29 | 13.68 | 2.859 | 9.48(6) | 37 | 12.92 | 3.038 | 12.12(8) | 46 | 17.22 | 3.090 |
| C_2H_4 | 11.01(3) | 30 | 17.01 | 2.730 | 10.54(1) | 38 | 15.74 | 2.927 | 14.0(1) | 47 | 21.49 | 2.958 |
| H_2S | 13.40(3) | 31 | 18.13 | 2.835 | 12.07(10) | 39 | 16.65 | 3.057 | 16.55(5) | 48 | 22.65 | 3.120 |
| H ₃ P | _ `` | _ | _ | _ | 11.56(7) | 40 | 19.30 | 2.878 | 20.7(1) | 49 | 28.88 | 2.898 |
| HCN | 12.33(5) | 32 | 18.42 | 2.639 | 11.09(10) | 41 | 16.83 | 2.826 | 14.5(1) | 50 | 23.66 | 2.840 |
| H_2O | 14.24(3) | 33 | 20.14 | 2.544 | 12.08(2) | 42 | 18.05 | 2.735 | 15.9(2) | 51 | 24.65 | 2.776 |
| H_3N | 34.3(5) | 34 | 40.43 | 2.304 | 26.7(3) | 43 | 34.01 | 2.532 | 30.4(3) | 52 | 46.75 | 2.599 |

^{*a*} Values of k_{σ} are either taken directly from the reference having the number indicated in columns 3, 7 or 11, as appropriate, or are recalculated from the centrifugal distortion constant D_J or Δ_J given therein by using eqn (1) or (2). The quoted error is that transmitted by the error in the distortion constant. D_{σ} and $r(Z \cdots X)$ are equilibrium values calculated *ab initio* at the CCSD(T)(F12*)/cc-pVDZ-F12 level of theory (see text). $r(Z \cdots X)$ is the distance from the acceptor atom/centre Z in the Lewis base B to the inner halogen atom X.

Table 3 Some observed and calculated properties of hydrogen-bonded complexes B···HX^a

| | Hydrogen hal | Hydrogen halide molecules HX | | | | | | | | | | | |
|------------------|--|------------------------------|--|--|-------------------------|-----------|--|------------|--|--|--|--|--|
| Lewis | Hydrogen fluo | oride HF | | Hydrogen chloride HCl | | | | | | | | | |
| base B | $k_{\sigma}/(\mathrm{N} \mathrm{m}^{-1})$ Ref. | | $D_{\sigma}/(\mathrm{kJ} \mathrm{\ mol}^{-1})$ | $r(\mathbf{Z}\cdot\cdot\cdot\mathbf{H})/\mathbf{\mathring{A}}$ | $k_{\sigma}/(N m^{-1})$ | Ref. | $D_{\sigma}/(\mathrm{kJ} \mathrm{\ mol}^{-1})$ | r(Z···H)/Å | | | | | |
| N_2 | 5.13(3) | 53 | 9.26 | 2.099 | 2.55(1) | 62 and 63 | 5.12 | 2.400 | | | | | |
| OC | 8.48(9) | 54 | 14.21 | 2.103 | 3.88(1) | 64 | 7.78 | 2.393 | | | | | |
| C_2H_2 | _ `` | — | — | — | 6.4(3) | 65 | 11.03 | 2.378 | | | | | |
| C_2H_4 | — | — | — | — | 5.88(16) | 66 | 11.32 | 2.396 | | | | | |
| H_2S | 12.0(2) | 55 and 56 | 20.36 | 2.284 | 6.81(1) | 67 | 12.73 | 2.480 | | | | | |
| H ₃ P | 10.94(4) | 57 | 19.38 | 2.354 | 6.01(2) | 68 | 11.89 | 2.569 | | | | | |
| HCN | 18.26(5) | 58 | 30.33 | 1.859 | 9.25(4) | 69 | 18.33 21.34 | 2.092 | | | | | |
| H_2O | 24.51(2) | 59 and 60 | 35.34 | 1.721 | 12.72(12) | 70 | | 1.912 | | | | | |
| H_3N | 32.8 | 61 | 50.98 | 1.703 | 18.2(3) | 71 | 32.72 | 1.820 | | | | | |
| | Hydrogen bro | mide HBr | | | Hydrogen iodide HI | | | | | | | | |
| N_2 | 1.92(1) | 72 | 3.98 | 2.503 | | _ | _ | _ | | | | | |
| ŌĊ | 2.99(1) | 73 | 6.08 | 2.489 | 1.713(1) | 81 | 4.02 | 2.675 | | | | | |
| C_2H_2 | 5.39(2) | 74 | 9.24 | 2.440 | _ `` | _ | _ | _ | | | | | |
| C_2H_4 | 5.21(2) | 75 | 9.65 | 2.456 | — | | _ | — | | | | | |
| H_2S | 5.86(2) | 76 | 11.01 | 2.526 | 4.02(1) | 82 | 7.78 | 2.670 | | | | | |
| H ₃ P | 5.05(1) | 77 | 10.19 | 2.618 | 3.409(2) | 83 | 7.17 | 2.778 | | | | | |
| HCN | 7.64(2) | 78 | 14.92 | 2.161 | 4.44(1) | 84 | 10.42 | 2.319 | | | | | |
| H_2O | 10.06(15) | 79 | 17.53 | 1.969 | 6.64(1) | 85 | 12.06 | 2.117 | | | | | |
| H_3N | 13.4(3) | 80 | 28.60 | 1.800 | 7.18(6) | 86 | 19.82 | 1.926 | | | | | |

^{*a*} Values of k_{σ} are either taken directly from the reference having the number indicated in columns 3 or 7, as appropriate, or are recalculated from the centrifugal distortion constant D_{J} or Δ_{J} given therein by using eqn (1) or (2). The quoted error is that transmitted by the error in the distortion constant. D_{σ} and $r(Z \cdots H)$ are equilibrium values calculated *ab initio* at the CCSD(T)(F12*)/cc-pVDZ-F12 level of theory (see text). $r(Z \cdots H)$ is the distance from the acceptor atom/centre Z of the Lewis base B to the hydrogen atom H.

hydrogen-bonded complexes $B \cdots HX$.^{53–86} All complexes considered here, except for those involving H_2O and H_2S , are either linear molecules, symmetric-top molecules or have C_{2v} symmetry, so that eqn (1) and (2) are strictly applicable at equilibrium. All complexes of H_2O with either HX or XY are effectively planar, that is although the equilibrium geometry has a pyramidal conformation at O (C_s symmetry) there is rapid inversion in the zero-point state between the two equivalent conformers and the vibrational wavefunctions have C_{2v} symmetry. Eqn (2) is then a reasonable approximation. Complexes $H_2S \cdots HX/XY$, on the other hand, all have C_s symmetry and are non-inverting in the zero-point state. They have a rightangled geometry in which HX or XY lies along an axis that passes through the H_2S centre of mass and is very nearly perpendicular to the H_2S plane. Nevertheless, eqn (2) is probably an acceptable approximation for the H_2S complexes.

2.2 Calculation of D_{σ} values

Values of the energy change, D_{σ} , accompanying the dissociation B···XY = B + XY of each of the complexes B···XY into the components B and XY, all in their (hypothetical) equilibrium electronic ground states, were calculated at the explicitly correlated level⁸⁷ CCSD(T)(F12*)/cc-pVDZ-F12 by using the *ab initio* program MOLPRO.⁸⁸ This involved geometry optimisations of B···XY, B and XY. Corrections for basis set superposition error (BSSE) were applied using the Boys–Bernardi⁸⁹ method. An advantage of using basis functions of the type cc-pVDZ-F12, that is functions optimised for use at the explicitly correlated level of theory, is that the BSSE corrections are relatively small, typically a few percent of D_{σ} . The basis functions for Br and I were of the type cc-pVDZ-F12-PP, where

PP indicates that a pseudo-potential is used for core electrons, and were provided by J. G. Hill of the University of Sheffield prior to their public release.90 For some complexes B...XY it was possible to conduct calculations at the CCSD(T)(F12*)/ cc-pVTZ-F12 level. This increased D_{σ} by approximately 5% in each case. Unfortunately, for a few complexes in the series B...BrCl, $B \cdots Br_2$ and $B \cdots ICl$, use of the cc-pVTZ-F12 basis functions would have proved too demanding of computer time. Therefore, in view of the systematic nature of the present investigation, it was decided to employ the highest level of theory that could be applied uniformly to all complexes considered, namely the level CCSD(T)(F12*)/ cc-pVDZ-F12. Values of D_{σ} so calculated for halogen-bonded complexes are included in Tables 1 and 2 while those calculated by the same approach for the hydrogen-bonded analogues B···HX are in Table 3. The atoms/points Z···X−Y and Z···H−X are required by symmetry to be collinear for all B...XY and B···HX except those involving H_2O and H_2S . It has been shown, however, that the deviation from collinearity will be negligible for H₂O and H₂S complexes involving both types of non-covalent interaction^{31,70} and therefore collinearity was enforced for these complexes during the geometry optimisations.

3. Results

3.1 Halogen-bonded complexes B···XY

Fig. 1 shows the calculated values of D_{σ} plotted as the ordinate against the experimental values of k_{σ} along the abscissa for the series of complexes $B \cdots Cl_2$ and $B \cdots Br_2$ in which B is CO, C_2H_2 , C_2H_4 , H_2S , HCN, H_2O , PH₃ and NH₃. Also included on the same graph are $B \cdots F_2$ for $B = H_2S$, H_2O , HCN and NH₃, which are the

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Fig. 1 A plot of D_{σ} versus k_{σ} for complexes of the type $B \cdots X_2$, where B is one of the series of Lewis bases CO, C_2H_2 , C_2H_4 , H_2S , H_2O , PH_3 or NH_3 and X_2 is one of the nonpolar dihalogen molecules F_2 , Cl_2 or Br_2 . The continuous line represents the straight line fitted to the points (including the origin) by linear regression and is given as eqn (3) in the text.

only difluorine complexes known in the gas phase for the B listed and for which experimental k_{σ} are available. The point (0, 0) has been included under the reasonable assumption that when there is no interaction between a pair of molecules forming a complex both measures of the binding strength become zero. The points in Fig. 1 fall on a straight line through the origin, indicating direct proportionality of D_{σ} and k_{σ} . The equation for the line obtained by means of linear regression is

$$D_{\sigma}/(\text{kJ mol}^{-1}) = 1.47(3)\{k_{\sigma}/(\text{N m}^{-1})\} - 0.21(21)$$
 (3)

The choice of $B \cdot \cdot \cdot X_2$ molecules for the initial demonstration of the direct proportionality of D_{α} and k_{α} was dictated by the fact that in general the nonpolar molecules X₂ form weaker complexes than do their polar counterparts ClF, BrCl and ICl. Given the limitations of the model (see Section 2.1) for the experimental determination of k_{σ} from the centrifugal distortion constants D_I and Δ_I , especially the assumption of monomer geometries unchanged on complex formation, it is likely that the model will better apply to the $B \cdots X_2$ than to those involving the polar dihalogens. Fig. 2 shows the points (k_{σ}, D_{σ}) for B···XY, when XY includes all dihalogen molecules, both polar and nonpolar. The continuous straight line in Fig. 2 corresponds to that defined in eqn (3), *i.e.* that fitted to the points for $B \cdots X_2$ only. Note that the scatter from the straight line of eqn (3) tends to increase as the binding strength increases. The points with largest deviation correspond to those for complexes of H₃N with each of BrCl and ICl. For H₃N···ClF there is experimental evidence from the nuclear quadrupole coupling constants for a significant charge redistribution (and probably geometrical rearrangement) on complex formation³⁴ and therefore the point (k_{σ}, D_{σ}) for this complex was excluded from the graph. When all the points shown in Fig. 2 are fitted by means of linear regression, the result is

$$D_{\sigma}/(\text{kJ mol}^{-1}) = 1.45(3)\{k_{\sigma}/(\text{N m}^{-1})\} - 0.06(35),$$
 (4)

that is, a straight line through the origin and of slope within experimental error of that obtained (eqn (3)) when only homonuclear dihalogen molecules act as the halogen-bond donor.



Fig. 2 A plot of D_{σ} against k_{σ} for complexes of the type $B \cdots X_2$ and $B \cdots XY$, where B is one of the Lewis bases N₂, CO, C₂H₂, C₂H₄, H₂S, H₂O, PH₃ or NH₃, X₂ is one of the nonpolar dihalogen molecules F₂, Cl₂ or Br₂, and XY is one of the polar dihalogens CIF, BrCl or ICl. The continuous line is that represented by eqn (3) and shown in Fig. 1, that is the straight line fitted to the points arising from $B \cdots X_2$ complexes only.

Proportionality of k_{σ} and D_{σ} for several $H_3N \cdots XY$ complexes was also noted by Hill and Xu.⁹¹

3.2 Hydrogen-bonded complexes B···HX

It is of interest to apply the same approach to hydrogen-bonded complexes $B \cdots HX$, where B is again one of the same series of simple Lewis bases N₂, CO, C₂H₂, C₂H₄, H₂S, H₂O, PH₃ and NH₃ used in the discussion of the halogen-bonded complexes in Section 3.1 and X is F, Cl, Br or I. Experimental values of k_{σ} obtained as before from centrifugal distortion constants and D_{σ} values calculated at the CCSD(T)(F12*)/cc-pVDZ level of theory are set out in Table 3. Fig. 3, in which D_{σ} is plotted *versus* k_{σ} for all members of the hydrogen-bonded series $B \cdots HX$ except H₃N···HBr and H₃N···HI, again reveals a reasonable straight line, with the following equation fitted by linear regression:

$$D_{\sigma}/(\text{kJ mol}^{-1}) = 1.53(3)\{k_{\sigma}/(\text{N m}^{-1})\} - 1.8(3)$$
 (5)



Fig. 3 A plot of D_{σ} against k_{σ} for hydrogen-bonded complexes of the type B···HX, where B is one of the Lewis bases N₂, CO, C₂H₂, C₂H₄, H₂S, H₂O, PH₃ or NH₃ and X is F, Cl, Br or I. The continuous line is that represented by eqn (5), that is the straight line fitted to the points (including the origin) by linear regression.

4. Discussion

It has been shown that it is possible to express the intermolecular dissociation energy D_{σ} in terms of the intermolecular stretching force constant k_{α} for a wide range of simple bimolecular halogen- and hydrogen-bonded complexes B. · · XY and $B \cdots HX$, where XY is a homo- or hetero-dihalogen molecule and X is a halogen atom, by means of the expression

$$D_{\sigma} = C_{\sigma} k_{\sigma}, \tag{6}$$

where the constant C_{σ} = 1.50(3) × 10³ m² mol⁻¹ = 2.49(5) × 10^{-21} m² It is not obvious why C_{σ} should have the same value for the hydrogen- and halogen-bonded series; it could be a coincidence. The fact that these two series of halogenand hydrogen-bonded complexes obey eqn (6) does suggest, however, that empirical radial potential energy functions leading to a direct proportionality between k_{σ} and D_{σ} might be used to calculate the energy levels associated with the intermolecular stretching vibration in such molecules. In order to derive eqn (1) and (2) for use in the determination of k_{σ} , it was necessary to assume that the motion associated with k_{σ} involved only a change in the intermolecular distance r between the two rigid, unperturbed components while maintaining the angular geometry. Two examples of simple functions⁹² that imply a relation of the type in eqn (6) are the Morse function

$$V(r) = D_{\sigma} [1 - e^{-a(r-r_{\rm e})}]^2, \qquad (7)$$

and the Rydberg function

$$V(r) = -D_{\sigma}[1 + b(r - r_{\rm e})]e^{-b(r - r_{\rm e})}.$$
 (8)

The quadratic force constant k_{σ} is related to V(r) by $k_{\sigma} =$ $\partial^2 V$ $=V^{\prime\prime}(r_{\rm e})$ so that differentiation of eqn (7) and (8) leads ∂r^2 to the expressions $D_{\sigma} = k_{\sigma}/2a^2$ and $D_{\sigma} = k_{\sigma}/b^2$, respectively, and hence to the identifications $a = (2C_{\sigma})^{-\frac{1}{2}}$ and $b = C_{\sigma}^{-\frac{1}{2}}$ respectively. Given values of a or b, the term values G(v), as wavenumbers, for the intermolecular stretching vibration (as defined above) for any of the complexes B···XY or B···HX considered here can then be estimated by means of the usual expression

$$G(\boldsymbol{\nu}) = \omega_{\sigma}(\boldsymbol{\nu} + 1/2) - \omega_{\sigma} \boldsymbol{x}_{\sigma}(\boldsymbol{\nu} + 1/2)^{2}, \qquad (9)$$

in which⁹²

$$\omega_{\sigma} = (2\pi c)^{-1} [V''(r_{\rm e})/\mu]^{\frac{1}{2}} \text{ and}$$
$$\omega_{\sigma} x_{\sigma} = \left(\frac{\hbar}{\pi c \mu}\right) \left\{ \frac{5}{96} \left[\frac{V'''(r_{\rm e})}{V''(r_{\rm e})}\right]^{2} - \frac{1}{32} \left[\frac{V'''(r_{\rm e})}{V''(r_{\rm e})}\right] \right\}$$

For the Morse function, ω_{α} and $\omega_{\alpha} x_{\alpha}$ can then be related to the constant a in the exponential term (and thence to C_{σ}) by differentiation to give

$$\omega_{\sigma} = \frac{1}{2\pi c} \left(\frac{2a^2 D_{\sigma}}{\mu} \right)^{\frac{1}{2}} \text{ and } \omega_{\sigma} x_{\sigma} = \left(\frac{\hbar a^2}{4\pi c \mu} \right), \tag{10}$$

For the Rydberg function the corresponding expressions in terms of its constant b are

$$\omega_{\sigma} = \frac{1}{2\pi c} \left(\frac{b^2 D_{\sigma}}{\mu} \right)^{\frac{1}{2}} \text{ and } \omega_{\sigma} x_{\sigma} = \frac{11}{24} \left(\frac{\hbar b^2}{4\pi c \mu} \right).$$
(11)

The quantities ω_{σ} and $\omega_{\sigma} x_{\sigma}$ have proved difficult to obtain experimentally, but Bevan, Lucchese and co-workers93-95 have determined accurate values of both for the complexes ${}^{16}\text{O}^{12}\text{C}\cdots\text{H}^{19}\text{F}$, ${}^{16}\text{O}^{12}\text{C}\cdots\text{H}^{35}\text{Cl}$, $\text{OC}\cdots{}^{35}\text{Cl}_2$ and ${}^{16}\text{O}^{12}\text{C}\cdots{}^{79}\text{Br}^{35}\text{Cl}$ with the aid of morphed potential energy functions for these molecules. Their values for the pairs of quantities $[\omega_{\sigma} \text{ and } \omega_{\sigma} x_{\sigma}]$ are [107.99(2) and 3.79 cm⁻¹], [62.88(3) and 1.61 cm⁻¹], [56.43(4) and 2.91 cm^{-1}] and [58(3) and 1.87 cm^{-1}], respectively. Those calculated from the Morse function by means of eqn (10) when $C_{\sigma} = 1.50(3) \times 10^3 \text{ m}^2 \text{ mol}^{-1} = 2.49(5) \times 10^{-21} \text{ m}^2$ is used are [117(2) and 2.9(1) cm⁻¹], [74.8(15) and 2.15(9) cm⁻¹], [54.2(10) and $1.69(7) \text{ cm}^{-1}$, and $[68.0(14) \text{ and } 1.51(6) \text{ cm}^{-1}]$, respectively. The agreement between the morphed values of Bevan, Lucchese et al. and those generated by eqn (10) is satisfactory, but the former are more accurate. Values of $\omega_{\sigma} x_{\sigma}$ implied by the Rydberg function are smaller by the factor of 0.917 than those predicted by the Morse function.

The morphed potential energy functions for OC···HF, OC···HCl, OC···Cl₂ and OC···BrCl reported in ref. 93–95 allow a severe test of the use of zero-point spectroscopic constants in eqn (1) and (2) in the absence of equilibrium values, as advertised in Section 2.1. The six-dimensional morphing described in ref. 93 leads to the prediction $\omega_{\sigma} = 107.99(2) \text{ cm}^{-1}$ for the (experimentally unknown) equilibrium wavenumber associated with the intermolecular stretching mode σ of the isotopologue ${}^{16}O^{12}C\cdots H^{19}F$, which implies the value $k_{\sigma} = 8.017(2)$ N m⁻¹ for the equilibrium quadratic force constant of that mode. By comparison, use of the centrifugal distortion constant D_I of ${}^{16}O^{12}C \cdots H^{19}F$ in eqn (1) in place of the unavailable equilibrium values gives $k_{\sigma} = 8.48(9)$ N m⁻¹ (see Table 3), thereby providing some confidence in the approximations alluded to. This confidence is reinforced by the similar quality of agreement found between the values k_{σ} = 3.668(3), 3.751(5) and 4.5(4) implied by the ω_{σ} from morphed potential energy functions^{94,95} of ¹⁶O¹²C···H³⁵Cl, ¹⁶O¹²C···³⁵Cl₂ and ¹⁶O¹²C···⁷⁹Br³⁵Cl, respectively, and those 3.88(1), 3.68(1) and 4.40(2) N m⁻¹ calculated from zero-point D_I values via eqn (1) (see Tables 1-3, respectively).

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Finally, it has been shown⁹⁶ that k_{σ} values for complexes B. $\cdot \cdot$ HX can be predicted with the aid of eqn (12)

$$k_{\sigma} = c' N_{\rm B} E_{\rm XY},\tag{12}$$

where c' is a constant, from numerical nucleophilicities $N_{\rm B}$ assigned to the Lewis bases B and numerical electrophilicities $E_{\rm HX}$ assigned to the acids HX. A similar expression⁹⁷ holds when the Lewis acids are dihalogen molecules. In view of the direct proportionality $D_{\sigma} = C_{\sigma}k_{\sigma}$ established here, it follows that it is also possible to use D_{σ} values in a similar manner to establish nucleophilicities for Lewis bases B and electrophilicities for Lewis acids HX or XY.

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