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Co(II)/Co(I) reduction-induced axial histidine-flipping in myoglobin reconstituted with a cobalt tetrahydrocorrin as a methionine synthase model†

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A conjugate between apomyoglobin and cobalt tetrahydrocorrin was prepared to replicate the coordination behavior of cob(I)alamin in methionine synthase. X-ray crystallography reveals that the tetra-coordinated Co(I) species is formed through the cleavage of the axial Co–His93 ligation after the reduction of the penta-coordinated Co(II) cofactor in the heme pocket.

Methionine synthase, which is widely found in *Prokaryota* and *Eukaryota*, is responsible for the methylation of homocysteine to yield methionine. It is known that the enzyme requires *N*⁵-methyl-tetrahydrofolate as a methyl group donor and that the methyl group transfer to homocysteine is promoted by a cobalamin cofactor, a vitamin B₁₂-like cobalt corrinoid complex.^{1–4} The entire enzyme structure has a molecular weight of *ca.* 136 kDa and includes four domains. The complexity of this enzyme system poses significant challenges with respect to the characterization of the entire structure. The crystal structure of the cobalamin-binding domain of methionine synthase has been determined by Ludwig and co-workers, and it was found that the axial ligand of cobalamin is a protein-derived histidine residue and not the benzimidazole moiety of cobalamin. The base-off/His-on coordination state is shown in Fig. 1a.⁵ In the catalytic cycle, two key intermediates have been proposed to exist transiently; (i) a highly nucleophilic tetra-coordinated Co(I) species known as cob(I)alamin and (ii) a methylated Co(III) species known as methylcobalamin.^{1–3} However, the structure of cob(I)alamin has never been determined directly.^{5,6} Herein, we report a hybrid model consisting of a cobalamin model complex with a suitable protein matrix to

Fig. 1 Structure of holoproteins and cofactors. (a) The crystal structure of the cobalamin-binding domain (PDB: 1BMT) of methionine synthase and the molecular structure of cobalamin. (b) The crystal structure of horse heart myoglobin (PDB: 2V1K) and molecular structures of native heme and Co^{II}(TDHC). Native heme is replaceable with Co^{II}(TDHC) as a cobalamin model.

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improve our understanding of the formation of the intermediate with the axial ligand rearrangement in the protein.

The hybrid model for methionine synthase was constructed as follows (Fig. 1b): First, horse heart apomyoglobin was used as a substitute protein matrix for the cobalamin-binding domain of methionine synthase. It is a relatively small and well-known

protein which provides a histidine residue as an axial ligand to the heme cofactor in the heme-binding pocket as well as the cobalamin-binding domain. Second, a cobalt tetrahydrocorrin complex was designed as an artificial cofactor to mimic the cobalamin moiety because native cobalamin consists of a tetrapyrrole macrocyclic ligand known as a corrinoid. A corrinoid is structurally similar to a porphyrin, except that it lacks one of the four *meso* carbons which are present in the porphyrin framework. Dolphin and co-workers synthesized a cobalt complex of 1,19-dimethyltetrahydrocorrin and detected the stable Co(I) species by UV-vis spectroscopy.⁶ This inspired us to insert such a complex into the apomyoglobin heme pocket to obtain ligation by the His93 residue which coordinates to the native heme iron of myoglobin.

The cobalt corrinoid complex, Co(TDHC) (TDHC = 8,12-dicarboxyethyl-1,2,3,7,13,17,18,19-octamethyltetrahydrocorrin) (Fig. 1b), was prepared as a model of cobalamin by cyclization of the corresponding biladiene-*a,c* dihydrobromide derivative in the presence of Co(OAc)₂ (Scheme S1, ESI†).⁷ Two methyl substituents at the C1- and C19-positions are on opposite faces of the corrinoid ring and the compound exists as two enantiomers, (1*R*,19*R*)-TDHC and (1*S*,19*S*)-TDHC, in a 1 : 1 ratio. The two propionate side chains introduced at the C8- and C12-positions are analogous to those of native heme and are included for the purpose of fixing the cofactor in the proper position *via* a hydrogen bonding network with polar amino acid residues in the heme pocket. The Co^{II}(TDHC) complex was added to apomyoglobin obtained by removal of the heme.⁸ ESI-TOF MS data, the CD spectrum (Fig. S1 and S2, ESI†) and ICP analysis of the reconstituted protein, rMb(Co^{II}(TDHC)), were used to confirm that Co^{II}(TDHC) binds to the protein in a 1 : 1 ratio. The UV-vis spectrum of rMb(Co^{II}(TDHC)) has a sharp absorption band at 510 nm (Fig. 2a), which is similar to that observed for the Co^{II}(TDHC) dimethyl ester in pyridine.⁹ Furthermore, the EPR measurement of rMb(Co^{II}(TDHC)) indicates a typical spectrum (Fig. 2b): the interaction of the cobalt nucleus (*I* = 7/2) with the unpaired electron in a pentacoordinated low-spin *d*⁷ configuration of Co(II) results in hyperfine splitting of the signal into an octet

centered at *g*_z = 1.99, and the presence of a ¹⁴N (*I* = 1) ligand attached to the axial position of the Co^{II}(TDHC) complex shows superhyperfine splitting of each component of the octet into a triplet.^{5,10} This provides support for our proposal that the cobalt atom coordinates to the Ne2-atom of the His93 imidazole ring as an axial ligand.¹¹

The crystal structure of rMb(Co^{II}(TDHC)) was obtained at a resolution of 1.30 Å and reveals that Co^{II}(TDHC) is bound in the heme pocket with Co(II)-His93 coordination characterized by a Co-Ne2 bond length of 2.18 Å (Fig. 3). In the crystals, the ratio of proteins, rMb(Co^{II}((1*R*,19*R*)-TDHC)) and rMb(Co^{II}((1*S*,19*S*)-TDHC)), was found to be 13 : 7 by the occupancy of the enantiomers (Fig. S3b, ESI†) because the TDHC ligand provides a mixture of two enantiomers due to introduction of two methyl substituents at the C1- and C19-positions (*vide supra*). Although the TDHC ligand, whose macrocycle is smaller than that of a porphyrin, has two methyl substituents oriented perpendicularly to the corrin plane, the position of the TDHC framework is remarkably similar to that of the porphyrin framework observed in native myoglobin. As a result, the two propionate side chains located at the C8- and C12-positions are found to participate in hydrogen bonding networks with Lys45 and Ser92/His97, respectively. These residues originally interact with the heme-propionate side chains in myoglobin. In addition, the polypeptide Cα atoms of the native myoglobin and rMb(Co^{II}(TDHC)) are superimposable with a root-mean-square deviation (RMS value) of 0.223 Å (Fig. S3, ESI†).¹² Although there are several examples of hemo-proteins reconstituted with artificial porphyrinoid cofactors,⁸ the present work is the first example of incorporation of a metalocorin derivative into a hemoprotein.

Reduction of rMb(Co^{II}(TDHC)) upon addition of dithionite under anaerobic conditions leads to UV-vis spectral changes that include clear isosbestic points over 2 s (Fig. S4, ESI†). The characteristic

Fig. 2 Characteristic spectra of reconstituted myoglobins: rMb(Co^{II}(TDHC)) (blue) and rMb(Co^I(TDHC)) (red). (a) UV-vis spectra of the reconstituted proteins in 0.1 M potassium phosphate buffer (pH 7.0) solution at 25 °C. (b) EPR spectra of the reconstituted proteins (0.5 mM) in 0.1 M potassium phosphate buffer (pH 7.0) solution at 10 K. The signal of *g*_z = *g*_y is centered at *g* = 2.23. The coupling constants of $|A_{\text{Coz}}|$ and $|A_{\text{Nz}}|$ for rMb(Co^{II}(TDHC)) are 99.7 and 20.8 G, respectively. rMb(Co^I(TDHC)) is EPR-silent.

Fig. 3 Crystal structure of rMb(Co^{II}(TDHC)). (a) The whole protein structure at a resolution of 1.30 Å. Co^{II}(TDHC) and proximal His93 are indicated by stick models. The structure of the TDHC framework possesses two enantiomers with a ratio of 13 : 7 due to the configuration of two methyl substituents at C1- and C19-positions (Fig. 1b). (b) A comparison of the positions of cofactors in the heme pocket of myoglobin. The superimposed structures (top and side views) are obtained from rMb(Co^{II}(TDHC)) (yellow stick) and native myoglobin (gray stick). The whole protein structures are shown in Fig. S3 (ESI†). The spheres in the centre of the macrocycles are the cobalt and iron atoms. The structure of native myoglobin is that of 2V1K of the Protein Databank (PDB data).

Fig. 4 Crystal structures of two oxidation states of the reconstituted protein. (a) Superimposition of the cofactors and the heme pocket residues of rMb(Co^I(TDHC)) (light blue) and rMb(Co^{II}(TDHC)) (yellow). The $2F_o - F_c$ electron density is shown as a blue grid (contoured at 1.0σ) around the His93 residue and the cofactor of rMb(Co^I(TDHC)). (b) and (c) Variations in crystallographic B -factors represented as width of the backbone trace in rMb(Co^{II}(TDHC)) (left) and rMb(Co^I(TDHC)) (right), respectively.

spectrum of the Co(I) species then appears with an absorption maximum at 530 nm (Fig. 2a). The disappearance of the characteristic Co(II) low spin EPR signals also indicates the formation of the diamagnetic d^8 Co(I) species, rMb(Co^I(TDHC)) (Fig. 2b). At pH 7.0, the CD spectrum of rMb(Co^I(TDHC)) displays two prominent minima at 209 and 222 nm, which are typical of α -helical structures, having the consistent molar ellipticity values with those determined for rMb(Co^{II}(TDHC)) (Fig. S2, ESI[†]). The redox potential of rMb(Co^{II}(TDHC))/rMb(Co^I(TDHC)) was determined by spectroelectrochemical measurements to be -0.13 V vs. SHE at pH 7.0 (Fig. S5, ESI[†]), while the value of cobalamin in methionine synthase is -0.5 V.¹³ This indicates that rMb(Co^I(TDHC)) is more thermodynamically stable than native cob(I)alamin in the protein.

To understand the coordination behavior of one of the intermediates, cob(I)alamin, during the catalytic reaction, the crystals of rMb(Co^{II}(TDHC)) were directly reduced by soaking in a solution of 0.5% (w/v) dithionite, and the resulting spectral changes suggest that the Co(I) species is formed in the crystal state. The crystal structure of rMb(Co^I(TDHC)) was successfully determined at 1.35 Å-resolution (Fig. 4a and Fig. S6 and S7, ESI[†]). Interestingly, the proximal His93 residue swings out upon cleavage of the Co–His coordination and gains flexibility. In fact, the decreased electron density of the His93 residue is observed in the $2F_o - F_c$ map ($\sigma = 1.0$). As a result, the distance between the Co(I) and N ϵ 2(His93) atoms is 3.44 Å in the crystal structure of rMb(Co^I(TDHC)). From the crystal structure it was found that Co^I(TDHC) is bound in the heme pocket as a tetra-coordinate complex without axial ligands. This action then induces a steric interaction with the 12-propionate side chain,

causing it to deviate from its normal position. Furthermore, in rMb(Co^I(TDHC)), the backbone B -factors of the F-helix,¹⁴ which includes the His93 residue, are found to be remarkably large compared to those of rMb(Co^{II}(TDHC)) (Fig. 4b and c and Fig. S8, ESI[†]) because of the lack of axial ligation. In contrast, the whole protein structure of rMb(Co^I(TDHC)) does not deviate significantly, as indicated by the RMS value of all the C α carbons (0.204 Å). Additionally, the protein matrix of myoglobin is found to be capable of holding the tetra-coordinate cofactor in the heme pocket *via* hydrogen bonding and hydrophobic interactions. The CD spectrum of rMb(Co^I(TDHC)) provides further evidence that the α -helicity of the protein is maintained in solution even after release of the axial ligand (*vide supra*). Taken together, these results indicate that the remarkable structural changes occur locally in the proximal site upon reduction and do not exert a significant influence on the whole structure of the protein. To the best of our knowledge, this is the first example of a tetra-coordinate structure of a Co(I) species in a corrinoid framework.¹⁵ It is of particular interest to monitor the unbound form of an axial ligand (His-off) in the myoglobin heme pocket in this unique situation. Our data support the speculation that the species of methionine synthase referred to as “base-off” cobalamin only requires proximal histidine flipping in the process of oxidation/reduction of cobalamin.

In this study, we have designed and prepared myoglobin reconstituted with the cobalt corrinoid complex, Co(TDHC) as a simple model for the active site of a complicated cobalamin-dependent methyltransferase. In the heme pocket of myoglobin, Co^{II}(TDHC) is tightly bound and provides a model of the base-off/His-on state which is seen in the cobalamin binding domain of methionine synthase, and then an important intermediate, the tetra-coordinated Co(I) species, is detectable in the protein matrix. The results described herein provide important insights into understanding how the on/off switching of the axial ligation takes place in the native protein.

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Notes and references

- 1 B. Kräutler, in *Metal Ions in Life Science*, ed. A. Sigel, H. Sigel, R. K. O. Sigel, RSC, Cambridge, 2009, ch. 1, vol. 6, pp. 1–51.
- 2 R. G. Matthews, in *Metal Ions in Life Science*, ed. A. Sigel, H. Sigel, R. K. O. Sigel, RSC, Cambridge, 2009, ch. 2, vol. 6, pp. 53–113.
- 3 B. Kräutler, B. Puffer, in *Handbook of Porphyrin Science*, ed. K. M. Kadish, K. M. Smith, R. Guilard, World Scientific, Singapore, 2012, ch. 117, vol. 25, pp. 131–263.
- 4 (a) K. L. Brown, *Chem. Rev.*, 2005, **105**, 2075–2149; (b) R. Banerjee, C. Gherasim and D. Padovani, *Curr. Opin. Chem. Biol.*, 2009, **13**, 484–491; (c) K. Gruber, B. Puffer and B. Kräutler, *Chem. Soc. Rev.*, 2001, **40**, 4346–4363; (d) R. G. Matthews, *Acc. Chem. Res.*, 2001, **34**, 681–689; (e) Y. Kung, N. Ando, T. I. Doukov, L. C. Blasiak, G. Bender, J. Seravalli, S. W. Ragsdale and C. L. Drennan, *Nature*, 2012, **484**, 265–269; (f) C. H. Hagemeyer, M. Krüer, R. K. Thauer, E. Warkentin and U. Ermler, *Proc. Natl. Acad. Sci. U. S. A.*, 2006, **103**, 18917–18922.

- 5 C. L. Drennan, S. Huang, J. T. Drummond, R. G. Matthews and M. L. Ludwig, *Science*, 1994, **266**, 1669–1674.
- 6 C.-J. Liu, A. Thompson and D. Dolphin, *J. Inorg. Biochem.*, 2001, **83**, 133–138.
- 7 D. Dolphin, R. L. N. Harris, J. L. Huppertz, A. W. Johnson and I. T. Kay, *J. Chem. Soc. C*, 1966, 30–40.
- 8 T. Hayashi, in *Handbook of Porphyrin Science*, ed. K. M. Kadish, K. M. Smith, R. Guilard, World Scientific, Singapore, 2010, ch. 23, vol. 5, pp. 1–69.
- 9 Y. Murakami and Y. Aoyama, *Bull. Chem. Soc. Jpn.*, 1976, **43**, 683–688.
- 10 (a) S. A. Cockle, H. A. O. Hill, J. M. Pratt and R. J. P. Williams, *Biochim. Biophys. Acta*, 1969, **177**, 686–688; (b) Y. Murakami, Y. Hisaeda and A. Kajihara, *Bull. Chem. Soc. Jpn.*, 1983, **56**, 3642–3646.
- 11 N. S. Hush and I. S. Woolsey, *J. Chem. Soc., Dalton Trans.*, 1974, 24–34.
- 12 H.-P. Hersleth, T. Uchida, A. K. Röhr, T. Teschner, V. Schünemann, T. Kitagawa, A. X. Trautwein, C. H. Görbitz and K. K. Andersson, *J. Biol. Chem.*, 2007, **282**, 23372–23386.
- 13 R. V. Banerjee, S. R. Harder, S. W. Ragsdale and R. G. Matthews, *Biochemistry*, 1990, **29**, 1129–1135.
- 14 (a) P. Picotti, A. Marabotti, A. Negro, V. Musi, B. Spolaore, M. Zambonin and A. Fontana, *Protein Sci.*, 2004, **13**, 1572–1585; (b) C. Nishimura, H. J. Dyson and P. E. Wright, *J. Mol. Biol.*, 2011, **411**, 248–263.
- 15 (a) P. Engel, G. Rytz, L. Walder, U. Vögeli and R. Scheffold, in *Vitamin B₁₂*, ed. B. Zagalak and W. Friedrich, Walter de Gruyter & Co., Berlin, 1979, pp. 171–172; (b) P. Doppelt, J. Fischer and R. Weiss, *Inorg. Chem.*, 1984, **23**, 2958–2962.