Diselenolodiselenole: a selenium containing fused heterocycle for conjugated systems†‡

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The synthesis of new conjugated building blocks, diselenolodiselenole (C₄Se₄) derivatives, is described for the first time. The structural and optoelectronic properties of C₄Se₄-derivatives are tuned by varying end-capping aromatic substituents. In cyclic voltammetry, all C₄Se₄-derivatives show two reversible oxidation peaks. 

Organoselenium compounds have attracted considerable interest due to their wide range of applications in many fields. Chalcogenophene containing building blocks are being given great importance to synthesize new conjugated organic materials. Their applications in the state-of-the-art technologies including field-effect transistors, flexible light emitting diodes and organic photo voltaics are being extensively studied. Though the development in this field has grown rapidly in the last two decades, the conjugated electroactive building blocks with promising properties are limited. Conjugated chalcogenophene based materials with reversible redox activity have found applications in organic electronic devices.

Thow the applications of thiophene based electroactive small molecules or polymers in organic electronics have attracted considerable research attention, their selenium counterparts are sparsely reported. Partly, it can be ascribed to the difficulties in the synthesis of selenophene derivatives and their instabilities in the charged states. The advantages of replacement of sulfur by selenium in conjugated systems are manifold: (a) intermolecular Se···Se interactions lead to a wide bandwidth in organic conductors, which should facilitate intermolecular charge transfer, (b) selenium containing organoheteroles have lower oxidation and reduction potentials than that of sulfur containing heterocycles, (c) due to higher polarizability of Se than that of S, compounds with the selenium atom attached to the conjugated backbone possess more polarizability than their sulfur analogues, (d) selenium containing compounds should have a lower band gap than their sulfur counterparts and, consequently, their optoelectronic properties also differ.

Blum and co-workers reported the synthesis of the first example of dithiolodithiole (C₄S₄) derivatives in 26% yield from 1,4-diphenylbutadiyne and elemental sulfur at 150 °C for 52 h. The growing interest in the C₄S₄ systems led to the development of a more facile synthetic strategy and a systematic study reported by Swager et al focusing on their interesting structural and optoelectronic properties. According to Hückel’s rule, the C₄S₄ system is formally anti-aromatic in the ground state and non-aromatic in the excited state; thus these molecules can lead to interesting electronic properties. The structural and electronic properties of these systems can be further tuned by an atomistic approach by replacing S with Se, which may lead to a more interesting and important fused conjugated system (C₄Se₄). Here, we present for the first time the synthesis of diselenolodiselenole (C₄Se₄) derivatives, bicyclic heterocycles, as a new class of conjugated building blocks.

We have synthesized a series of compounds containing C₄Se₄ as the central conjugated system. The precursor diyne compounds 1a, 10, 1e 11 and 1e 12 were prepared by a previously reported procedure. Dinyes 1b and 1d were synthesized by new synthetic methods (see ESI†). The conversion of dinyes to C₄Se₄ derivatives could proceed through the radical mechanism similar to the formation of C₄S₄ derivatives. However, the reaction of 1a with Se powder in the presence of solvent (1,2-dichloroethane (DCE)-o-dichlorobenzene (o-DCB)) and the radical initiator (azobisisobutyronitrile (AIBN)-2,2,6,6-tetramethylpiperidinolox (TEMPO)) at 190 °C in a pressure vessel with and without microwaves resulted in very low yields (Scheme 1, conditions I and II).

Moreover, heating of elemental selenium with 1a nearly at the melting point of selenium without any solvents afforded the best yield of 23% for 2a (Scheme 1, condition III). Therefore, condition III was considered as a general procedure to prepare C₄Se₄ derivatives 2a–2e from diyne precursors (Scheme 2). The yields were obtained in the range of 10–30%. Though the strongly electron donating substituents on the phenyl ring were reasoned to cause a
complex mixture of the product and low yield in the case of C₄S₄ derivatives,⁹ the reaction of di(p-hexyloxyphenyl) diacetylene (1c) afforded the highest yield (30%) in this series.

Crystals of 2a–2c were obtained by the slow evaporation method from their solution in dichloromethane (DCM). In the crystal structure of 2a, capped phenyl rings are twisted from the C₄Se₄ core by a dihedral angle of ~53° (C₈–C₇–C₆–C₅ = 126.8(4)) (Fig. 1a and b), which is significantly higher (by ~28°) than that of sulfur analogue 3a.⁹ Molecules of 2a form end-to-end dimers via intermolecular π–π interactions (C₄–C₅ = 3.91 Å and C₄–C₄ = 3.37 Å). Selenium atoms of the Se–Se bond of 2a form Se/C₁/C₁/C₁Se interactions (Se₁/C₁/C₁/C₁Se₂ = 3.64 Å) with the neighboring two molecules. This leads to the formation of a virtual (Se–Se/C₁/C₁/C₁Se–Se)ₙ polymeric chain along the c-axis, from which the phenyl rings are hanged like pendants (Fig. 1c), whereas in the crystal packing of 3a, face-to-face dimer formation was observed via S/C₁/C₁/C₁Si interactions.⁹

Interestingly, in the case of thiophene capped C₄Se₄ (2b) the torsional angle between the outer thiophene ring and the central C₄Se₄ unit is found to be only ~10° (C₁–C₂–C₃–C₄ = 169.3(7)) (Fig. 2a and b). The nearly planar conjugated backbone of 2b exhibited resolute intermolecular interactions through heteroatoms. In the crystal structure of 2b, a pair of Se atoms of each diselenole unit forms two perpendicular dimers by three different Se/C₁/C₁/C₁Se interactions (Se₁/C₁/C₁ = 3.648(1) and Se₂/C₁/C₁ = 3.525(1), Se₁/C₁/C₁ = 3.726(1)) with the diselenole unit of neighboring molecules (Fig. 2c). Thus four molecules connected by Se–Se and π–H–C interactions (C₆–H₆ = 2.81 Å) form a 2D brick-like structure in bulk. This 2D crystal packing with several nonbonding interactions could facilitate the intermolecular charge transport. In 2c the dihedral angle between the hexyloxy substituted phenyl ring and the central C₄Se₄ unit is ~58° (C₈–C₃–C₂–C₁ = −122.4(5)) (Fig. S1, ESIF).

DFT optimized structures (at B3LYP/6-31G(d)) of 2a and 2b showed dihedral angles of 47° and 0°, respectively, between the central C₄Se₄ unit and end-capping substituents. Corresponding values for C₄S₄ derivatives are 39° and 0°, respectively.

Compounds 2a–2e exhibited two sets of absorption peaks with λmax ranging from 236 to 277 nm and 427 to 484 nm, respectively (Fig. 3 and Table 1). Compound 2a showed λmax at 440 nm in solution, which was blue shifted compared to its C₄S₄ analogues. This may be ascribed to the large dihedral angle between the outer phenyl rings and the central C₄Se₄ unit that reduces the effective overlap between these two conjugated
Dihedral angles between C-Se interaction and Se-Se interaction, diselenolodiselenoles, was successfully synthesized simply by heating diaryl diselenides with elemental selenium. The structural and optoelectronic properties of diselenolodiselenene derivatives can be tuned by the judicious choice of the capped alkyl groups. The thiophene capped C₄Se₄ displayed a nearly planar structure with its absorption at the highest wavelength among the compounds in the present series. Therefore, it is a promising candidate to be exploited for application in organic electronics. Due to the presence of Se···Se interactions, diselenolodiselenene derivatives can arrange into interesting crystalline motifs. Thus, we have shown that the structural engineering and atomistic approach could be beneficial to synthesize meaningful building blocks for conjugated systems.

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Notes and references

![Fig. 3](image_url) UV-vis spectra of 2a–2e in DCM.

![Fig. 4](image_url) Electrochemical properties of compounds 2a–2e in 0.1 M TBAPF₆ in dry DCM as solvent using a Pt-disk working electrode, a Pt-wire counter electrode and an Ag/AgCl reference electrode.

**Table 1.** Yields, absorption and electrochemical properties of 2a–2e

<table>
<thead>
<tr>
<th>Compound</th>
<th>Yield (%)</th>
<th>λ_{max} (nm)</th>
<th>E_{1/2 \text{vs.} \text{Ag/Ag}⁺} (V)</th>
<th>E_{\text{opt}} (eV)</th>
<th>HOMO⁻</th>
<th>LUMO⁺</th>
</tr>
</thead>
<tbody>
<tr>
<td>2a</td>
<td>23</td>
<td>240, 440</td>
<td>0.35, 0.86</td>
<td>2.39</td>
<td>−4.79</td>
<td>−2.40</td>
</tr>
<tr>
<td>2b</td>
<td>12</td>
<td>275, 484</td>
<td>0.47, 0.98</td>
<td>2.21</td>
<td>−4.91</td>
<td>−2.70</td>
</tr>
<tr>
<td>2c</td>
<td>30</td>
<td>277, 428</td>
<td>0.25, 0.77</td>
<td>2.46</td>
<td>−4.69</td>
<td>−2.23</td>
</tr>
<tr>
<td>2d</td>
<td>15</td>
<td>253, 427</td>
<td>0.49, 1.03</td>
<td>2.43</td>
<td>−4.93</td>
<td>−2.50</td>
</tr>
<tr>
<td>2e</td>
<td>19</td>
<td>236, 475</td>
<td>0.50, 1.04</td>
<td>2.15</td>
<td>−4.94</td>
<td>−2.79</td>
</tr>
</tbody>
</table>

$E_{\text{opt}} = 1240/\lambda_{\text{max}}, \quad E_{\text{HOMO}} = E_{\text{LUMO}} = E_{\text{LUMO}} + E_{\text{opt}}.$