Introduction

The fabrication of functional nanomaterials by molecular self-assembly is attracting increasing attention.1,2 The use of short peptide derivatives as versatile building blocks offers a suitable platform due to the chemical versatility of amino acid building blocks combined with simplicity and chemical accessibility of peptides.3–8 There is an on-going effort to develop molecular design rules for the basic building blocks. Many new self-assembling structures are discovered by serendipity, rather than by rational design. One approach that may accelerate discovery is combining molecular self-assembly with dynamic combinatorial chemistry (DCC).9–12 This technique allows for the continuous interconversion of building blocks that leads to the formation of several dynamic combinatorial library (DCL) members under thermodynamic control through a reversible, yet covalent, chemical reaction. Recently, DCLs involving peptide derivatives have been developed by using disulphide,13–15 metal binding16,17 and amide18 exchange reactions. We have focused on a fully reversible enzymatic amide exchange approach driven by molecular self-assembly, demonstrating an enhanced self-selection and amplification of effective energy transfer nanostructures from complex mixtures of dipeptide derivatives. By taking advantage of an enzyme-catalysed fully reversible amide formation reaction, we show how gelation shifts the equilibrium in favor of the formation of short aromatic dipeptide derivatives in the DCL system, as confirmed by reversed-phase high pressure liquid chromatography (HPLC), fluorescence emission spectroscopy, atomic force microscopy (AFM), transmission force microscopy (TEM) and circular dichroism (CD) spectroscopy. This approach enabled us to identify a two-component donor–acceptor hydrogel, which forms within minutes and exhibits efficient energy transfer.

Peptide self-assembly provides a useful approach to control the organization of functional molecular components, as relevant to future opto-electronic or photonic nanostructures. In this article, we report on the discovery of efficient energy transfer nanostructures using a dynamic combinatorial library (DCL) approach driven by molecular self-assembly, demonstrating an enhanced self-selection and amplification of effective energy transfer nanostructures from complex mixtures of dipeptide derivatives. By taking advantage of an enzyme-catalysed fully reversible amide formation reaction, we show how gelation shifts the equilibrium in favor of the formation of short aromatic dipeptide derivatives in the DCL system, as confirmed by reversed-phase high pressure liquid chromatography (HPLC), fluorescence emission spectroscopy, atomic force microscopy (AFM), transmission force microscopy (TEM) and circular dichroism (CD) spectroscopy. This approach enabled us to identify a two-component donor–acceptor hydrogel, which forms within minutes and exhibits efficient energy transfer.

Discovery of energy transfer nanostructures using gelation-driven dynamic combinatorial libraries†

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Chemical Science

Results and discussion

The first objective was to develop a DCL based on the self-assembly of short aromatic dipeptide amphiphiles, to identify the most stable structures, in the absence of the acceptor. The self-assembly of naphthoxy-substituted dipeptides driven by an enzyme-catalysed condensation reaction was selectively chosen for this purpose. As shown in Fig. 1, the precursor solution consisted of naphthoxy-substituted tyrosine derivative (Nap-Y; 20 mM) with a 4-fold excess of an amino acid amide nucleophile (X-NH₂, where “X” denotes phenylalanine, F; leucine, L; valine, V; tyrosine, Y; alanine, A and/or glycine, G) in 100 mM phosphate buffer (pH 8). A non-specific endopeptidase, thermolysin from Bacillus thermoproteolyticus rokko, was used (1 mg ml⁻¹)⁴³,⁴⁴ to catalyse the condensation reaction of the amino acid amide derivatives. This ultimately converts the non-assembling precursors (Nap-Y and X-NH₂) into self-assembly building blocks (Nap-YX-NH₂), provided that the reaction is thermodynamically driven by the free energy contribution of the self-assembly process.

At first, the enzymatic formation of various Nap-YX-NH₂ derivatives (also simply represented as “XY”) from each combination of Nap-Y and X-NH₂ was investigated in isolation. The molecular self-assembly of the building blocks upon the addition of thermolysin was directly identified by the naked eye observation of a rapid transformation from a clear and transparent solution into a self-supporting hydrogel (as recognized by vial inversion). The percentage conversion was investigated by reversed-phase high-pressure liquid chromatography (HPLC). As shown in Fig. 2A, some of the dipeptide derivatives, YF (89%), YL (72%) and YY (81%) were formed in very good yields, while the other derivatives, YY (5%), YA (0.9%) and YG (0.7%) were formed in very low yields after 48 h. The high yields obtained for YF, YL and YY indicate that gelation-driven self-assembly of these dipeptide derivatives is thermodynamically favourable, as a result of the presence of strong π–π stacking between naphthalene chromophores and the hydrogen bonding between peptide motifs. Notably, only YF, YL and YY derivatives were self-assembled to form self-supporting hydrogels (Fig. S1 in ESI†).⁴¹ Interestingly, these hydrogels were formed rapidly (within minutes after the addition of thermolysin).⁴⁴ The percentage formation of these dipeptide derivatives was further monitored over time and in each case the maximum conversions were reached after 24 h (Fig. S2 in ESI†). The low yields obtained for the other dipeptide derivatives (YY, YA and YG) are likely due to their less favourable self-assembly; these systems lack the thermodynamic driving force to overcome the bias for amide hydrolysis in aqueous systems.⁴ It should be noted that thermolysin has a kinetic preference⁴⁵ for more hydrophobic amino acids (X = phenylalanine in the case of YX) on the amine side of the peptide bond, whereas it is
derivatives over time (solid traces, Fig. 2B). It was also found that the dipeptide, we also investigated the evolution of all dipeptide the enzyme has any kinetic preference for a particular dipeptide sequence in DCL. This hypothesis was reinforced by the fact that the percentage evolution of YF in DCL is strictly dependent on the total donor to acceptor ratio. The relative percentage evolutions of YF in DCL are 61% at 10 : 1, 65% at 7 : 1 and 70% at 2 : 1 donor-acceptor ratios (Fig. S3 in ESI†). This further confirms that the product distribution in the library is dependent on the type and the extent of the interactions present and is entirely thermodynamically driven by the more favourable self-assembly of mixed donors and acceptors in the library. The availability of both donors and acceptors in DCL makes this approach more versatile and greatly enhances the potential ability not just for self-selection but also for the amplification of most stable energy transfer nanostructures at the expense of less stable self-assembling nanostructures.

Next, we moved on to determine the efficiency of energy transfer within the newly discovered hydrogels. The supramolecular self-assembly of Nap-YX-NH$_2$ amphiphiles in the DCL system was characterized by fluorescence emission spectroscopy both in the absence and presence of DA. As shown in Fig. 3A and B, in the absence of DA, the fluorescence emission spectrum of the starting mixture (Nap-Y + X-NH$_2$) exhibited two distinct characteristic peaks. One is a sharp naphthalene
monomer emission peak at 350 nm and the other one is a broad shoulder peak at around 420 nm which corresponds to the excimer emission, emerging from the intermolecular π-π stacking between naphthalene chromophores (solid black trace).

The addition of thermolysin to the starting mixture induced a slight red-shift in the peak at 350 nm accompanied by a massive increase in the relative emission intensity (red trace, Fig. 3A). These observations indicate that excimer naphthalene chromophores were formed which eventually emit at higher wavelength when compared to monomer naphthalene chromophores. Even after 24 h, the peak at 350 nm was further red-shifted accompanied by a further increase in the relative emission intensity (green trace, Fig. 3A). Such an increase in the relative emission intensity over time indeed indicates that the self-assembly of Nap-YX-NH₂ amphiphiles in DCL system shows an aggregation-induced emission (AIE) behaviour.

The fluorescence emission of Nap-YX-NH₂ amphiphiles in the presence of acceptor DA was then studied. The fluorescence emission of DA molecules (in solution state, λ_em = 560 nm and λ_ex = 330 nm) is strongly dependent on the surrounding environment such as solvent, morphology, etc. and usually exhibits a blue-shift in the emission peak when migrates from aqueous to non-aqueous environment (Fig. S4 in ESI†). As shown in Fig. 3B, the addition of DA to the starting mixture (Nap-Y + X-NH₂) at variable concentrations induced substantial changes in the fluorescence emission spectra. For instance, when the concentration of DA was 2 mM (i.e., 10:1 donor–acceptor ratio), the monomer emission peak at 350 nm was quenched completely; while the excimer emission peak at 420 nm was enhanced exclusively and interestingly, a new broad emission peak was appeared at 535 nm (solid red trace). When the concentration of DA was increased to 10 mM (i.e., 2:1 donor–acceptor ratio), the emission peak at 420 nm was also quenched completely; while the broad peak at 535 nm was even broadened (solid blue trace). The new emission peak at 535 nm (rather than the expected peak at 560 nm for DA in solution) may suggest that the surrounding environment of DA molecules is changing from aqueous to non-aqueous environment (corresponding to blue-shifted emission). At this stage, the direct visualization of the morphology by atomic force microscopy (AFM) and transmission electron microscopy (TEM) confirmed the presence of spherical aggregates with an average diameter of 100–400 nm (Fig. 4A and C and S5 in ESI†). These results suggest that the DA molecules were incorporated in these spherical aggregates formed by Nap-Y and further indicate that there is a partial energy transfer between the naphthalene chromophores of Nap-Y and the dansyl chromophores of DA. In the solution state, efficient energy transfer occurs above 10 mM concentration of DA and precursor molecules.

The subsequent addition of thermolysin to the starting mixture solution of DA (2 mM and 10 mM) induced the formation of a self-supporting hydrogel. As shown in Fig. 3B, both peaks at 350 nm and 420 nm were quenched completely; while the broad emission peak at 535 nm was largely blue-shifted and appeared at 510 nm (dashed traces). Interestingly, energy transfer induced a dramatic enhancement in fluorescence emission (more than a 150-fold increase in the relative emission intensity above 500 nm) over time (Fig. S6 and S7 in ESI†). AFM and TEM images showed that the addition of thermolysin triggered the rapid morphological transition of spherical aggregates into entangled nanofibres of up to several micrometres in length (Fig. 4B, D and E). Furthermore, similar experiments were also repeated for the most stable YF system in isolation (since it was discovered as a major component from the direct competition in DCL).
Nap-YF-NH₂ nanostructures were characterized by monitoring the signals which appear due to the supramolecular chiral environment of naphthalene chromophores in the gel phase over time. As shown in Fig. 5, the dansyl chromophore of DA (2 mM) by itself is achiral and showed no CD signals (black trace). Similarly, no CD signals were observed even when DA was added to the starting mixture solution (Nap-Y + F-NH₂) as naphthalene chromophores are also achiral (red trace). However, upon the addition of thermolysin, the corresponding CD signals were instantaneously starting to appear within the measurement time (<2 min, blue trace) which indicates that the chromophores were self-assembling into a supramolecular chiral structure.

The strength and intensity of CD signals at 288 nm and 328 nm, originating due to the $\pi-\pi^*$ and $n-\pi^*$ transitions, respectively, of naphthalene chromophores were increased over time. In this case, the CD activity of dansyl chromophores of DA should be an induced CD (ICD) which would appear at 330 nm and it is more likely to interfere with the CD signals originating due to the $\pi-\pi^*$ transition peaks of the naphthalene chromophores. Therefore, we monitored the changes observed in the CD signals of the naphthalene chromophores both in the absence and presence of DA. Notably, the comparison of CD signals in the absence (orange trace) and presence (purple trace) of DA suggested the following changes in CD signals after 24 h: the intensity of the CD signals was increased by more than 2 times, the signal at 278 nm was red-shifted by about 10 nm and the ratio of both these $\pi-\pi^*$ and $n-\pi^*$ transitions peaks changed from 3 to 1.3 (purple and orange traces). These findings indicate that the key presence of DA molecules would certainly provide the crucial strong donor–acceptor interactions in addition to the fundamental non-covalent interactions that are already present in the entangled nanofibres and hence facilitates the formation of the most stable nanostructures with enhanced chirality. All these results consistently indicate that DA molecules have been successfully incorporated in the entangled fibre networks of Nap-YF-NH₂ which ultimately leads to the formation of two-component hydrogels.\(^{21}\)

Remarkably, it was found that the energy transfer in the gel state is efficient even at 0.6 mM concentration of DA (i.e., 33:1 donor–acceptor ratio) and also in this case, such an energy transfer induced an enhanced fluorescence emission over time (Fig. 3C and also Fig. S8–S10 in ESI†). TEM images also showed the presence of entangled nanofibres of up to several micrometers in length (Fig. 4F). The striking differences observed in the emission spectra before (solid traces) and after (dashed traces) the addition of thermolysin would indeed suggest that the energy transfer occurs within the gel preparation and measurement time (<2 min), as well as being efficient in the gel state (fibre networks) when compared to the solution state (spherical aggregates).

We further investigated the successful incorporation of DA molecules into the entangled fibre networks of naphthoxy-substituted dipeptide (for instance, YF) by circular dichroism (CD) spectroscopy. The interactions of DA molecules with...
Conclusions

In summary, we have demonstrated that a two-component hydrogel composed of \textit{Nap-YF-NH$_2$} and DA emerged as an effective energy transfer gel from a library of eight competing amino acid non-assembling precursor components (\textit{Nap-Y, F, L, V, Y, A, G-NH$_2$}, amino acid amide derivatives and DA molecules) through a fully reversible thermolysin-triggered condensation reaction. Furthermore, we also showed that the discovery of a functional (rather than just structural) system was achieved through a self-selection and amplification mechanism from component mixtures. This approach opens up the new possibility of discovering electronically conductive nanomaterials with enhanced charge transfer properties and fewer defects.

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Notes and references

40 L. Chen, S. Revel, K. Morris and D. J. Adams, \textit{Chem. Commun.}, 2010, 46, 4267–4269 (this paper describes the first reported energy transfer in a peptide-based hydrogel system that occurs only at relatively low pH values (pH < 4); organic solvents are used to dissolve propyl dansylamide acceptor used in this system, however, the energy transfer is apparently rather slow and low efficient).
43 However, the chemically synthesized (instead of enzymatic) YF and YL dipeptide derivatives formed amorphous precipitates upon cooling a heated solution in buffer which highlights the significant role of the enzyme needed for the formation of self-supporting hydrogels in this case. The enzymatic reaction starts from freely soluble building blocks rather than an insoluble precursor, which facilitates overcoming the free energy barrier to enable access to the self-assembled state.
44 The presence of free terminal amide groups in dipeptide amphiphiles may accelerate the self-assembly process that eventually triggers rapid hydrogelation.
51 We cannot be conclusive at this point to exclude the possibility of co-assembly with the other dipeptide derivatives (for instance, YF and YL). The investigation of this would require labeling experiments which are currently ongoing.